

Cite this: *J. Mater. Chem. C*,
2024, 12, 14893Received 16th April 2024,
Accepted 8th August 2024

DOI: 10.1039/d4tc01561d

rsc.li/materials-c

3D-(p/p/n) NiO/NiWO₄/WO₃ heterostructures for the selective detection of ozone†

Navpreet Kaur * and Elisabetta Comini 

We propose 3D branched-like NiO/NiWO₄/WO₃ nano-heterostructures for the selective detection of ozone (O₃) at ppb levels, which is crucial for environmental and public health protection. These complex nano-heterostructures were fabricated using a combination of the vapor–liquid–solid and vapor–solid mechanisms, during which the reaction between NiO and WO₃ leads to the formation of an intermediate seed, *i.e.*, NiWO₄. By controlling the charge transport within the nano-heterostructure through modulation of the operating temperature, sensors demonstrated highly selective sensing performances toward O₃ compared to NiO and WO₃ nanowire sensors. At 300 °C, a response as high as 4709 ± 9 was observed for 300 ppb of O₃ gas. In fact, we were able to achieve high selectivity toward O₃ compared to other highly reactive oxidizing compounds such as NO₂. Due to their remarkable sensing performance, these heterostructures are leading candidates for the fabrication of future-generation miniaturized sensing devices for environmental and/or health monitoring.

Introduction

Improving the gas-sensing properties of metal oxides (MOXs) by fabricating heterostructures from two different materials is an innovative and efficient strategy.^{1–3} Heterostructures provide unique properties such as efficient charge carrier transport, large reaction surface, and superior gas sensing performance.^{4,5} However, constructing heterostructures using 1D-nanostructures such as nanowires (NWs) is a challenging task, but it opens new ways to improve the electronic properties of the materials at the nanoscale level. This novel idea brings two different 1D-nanostructured metal oxides into one single sensing platform through the formation of a heterojunction. Due to this junction formation, the Fermi levels of dissimilar metal oxides reach equilibrium, resulting in charge transfer and depletion region formation at the interface.^{6,7} In particular, 3D branched-like nano-heterostructures provide a surface with a substantial number of reactive sites and distinct charge carrier transport that enhance the sensitivity and selectivity and alter the operating temperature of the sensors compared to pristine material NWs.^{6,8}

In our previous report,⁸ we have seen that by controlling the charge transport in complex 3D p–p–n nano-heterostructures of NiO/NiWO₄/WO₃ with temperature, their sensing performance can be tuned toward oxidizing and reducing gases. In the

current work, we focus on the selective detection of ozone (O₃) using these nano-heterostructures. O₃ is a crucial atmospheric component that plays a dual role, serving as both a beneficial shield against harmful ultraviolet radiation in the stratosphere and a pollutant at ground level.^{9,10} Ozone, a highly reactive molecule composed of three oxygen atoms, plays a dual role in our atmosphere.^{11,12} The ozone layer in the stratosphere shields the Earth from harmful ultraviolet (UV) radiation,¹³ ground-level ozone, primarily formed through photochemical reactions involving nitrogen oxides (NO_x) and volatile organic compounds (VOCs), poses serious risks to human health and the environment.¹⁴ Exposure to elevated concentrations of ground-level ozone, even at parts per billion (ppb) levels, can result in adverse health effects, including respiratory problems, exacerbation of asthma, and cardiovascular issues.^{15–17} Therefore, the development of sensitive and selective ozone-sensing technologies is of paramount importance for environmental monitoring and public health protection.

In this work, we presented 3D-(p/p/n) nano-heterostructures for the detection of ozone at ppb levels. The core NiO nanowires were fabricated using the vapor-phase growth method *via* the vapor–liquid–solid (VLS) mechanism. WO₃ was used as the secondary MOX, grown in the form of nanowires using the same technique *via* the vapor–solid (VS) mechanism onto NiO NWs. Intermediate growth of the ternary material (nickel tungstate NiWO₄) was observed as seed formation on top of NiO NWs prior to the WO₃ NWs creating branch-like heterostructures of NiO/NiWO₄/WO₃ (NWO). Afterward, a set of conductometric sensing devices based on pristine NiO and heterostructure sensors were prepared and tested for different concentrations

SENSOR Laboratory, Department of Information Engineering (DII), University of Brescia, Via D. Valotti 9, Brescia 25133, Italy. E-mail: navpreet.kaur@unibs.it

† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d4tc01561d>



of oxidizing and reducing compounds mainly focusing on the selective detection of ozone. The fabricated NWO heterostructure sensors exhibit higher sensitivity and selective behavior towards ozone even at ppb levels.

Experimental

Growth and characterization of NWO 3D-heterostructures

Polycrystalline alumina substrates 2 mm × 2 mm (Kyocera, Japan, with 99% purity) were used to fabricate the nanostructures. Before deposition, all substrates underwent ultrasonic cleaning with acetone and were subsequently dried with synthetic air. An ultrathin layer of the Au catalyst was deposited onto the Al₂O₃ substrates using RF magnetron sputtering (Kenotec Sputtering system, Italy). The deposition of Au was carried out at a power of 70 W with an Ar flow rate of 7 sccm at a pressure of 5 × 10⁻³ mbar for 5 seconds.

For the core of the 3D-heterostructures, the NiO nanowires were fabricated using a tubular furnace by employing the vapor-liquid-solid (VLS) growth mechanism. The growth process was initiated at an evaporation temperature for the NiO powder (Sigma-Aldrich, 99% purity, CAS number: 1313-99-1) of 1400 °C, with an Ar flow rate of 100 sccm at a pressure of 1 mbar. The Au-catalyzed Al₂O₃ substrates were maintained at a temperature of 930 °C, and the deposition time was set for 15 minutes.¹⁸ After the fabrication of the core NiO NWs, the secondary material WO₃ was used to fabricate small nanowires directly onto the NiO nanowires *via* the vapor-solid (VS) growth mechanism. The source material, WO₃ powder (Sigma-Aldrich, 99% purity, CAS number: 1314-35-8), was heated up to 1100 °C, while the substrates with NiO nanowires were placed at a lower substrate temperature of 580 °C. During the deposition process, the argon flow was maintained at 100 sccm, and the pressure was held at 1 mbar, with a deposition time of 15 minutes.

The final 3D-heterostructures consisted of a long core of NiO NWs followed by the formation of the NiWO₄ seed on the top and afterward the formation of small WO₃ nanowires. During the process of WO₃ NW growth, as the vapors of WO₃ reach the surface of NiO NWs a reaction occurs, resulting in the formation of a ternary phase NiWO₄. The well-known phase transition diagram of N-W-O suggested that at elevated temperatures of around 680 °C (pressure = 1 atm), the W-O system consists of different phases like WO₂, WO_{2.9}, and WO₃.¹⁹ Moreover, in the presence of Ni and oxygen, these phases transit to the formation of the most stable phase of NiWO₄.¹⁹ Thus, under the given experimental conditions, the binary components of NiO and WO₃ first lead to the formation of a seed of NiWO₄ as long as the WO₃ vapors were able to reach the NiO surface. Afterward, further incoming WO₃ vapors are not able to reach the NiO surface and the seed acts as a nucleation site for the growth of WO₃ NWs resulting in the final 3D-heterostructures of NiO/NiWO₄/WO₃ (p/p/n).

The final heterostructure morphology was investigated using a field-emission scanning electron microscope (FE-SEM, model LEO 1525, ZEISS). The detailed investigation of pristine

NiO NWs and 3D-heterostructures using GI-XRD, HR-TEM, RAMAN, and XPS has been provided in our previous article.⁸

Fabrication of gas sensing devices and electrical measurements

The conductometric sensing devices were fabricated by DC magnetron sputtering using NWO 3D-heterostructures and NiO NWs as the active transduction material. First, on the top of nanostructures, Pt interdigitated electrodes (IDEs) were deposited, and a platinum heater on the backside of the Al₂O₃ substrate. The complete process consists of deposition of the TiW adhesion layer using DC magnetron sputtering (70 W, Ar plasma, 100 nm, 300 °C, ≈ 5.3 × 10⁻³ mbar) followed by Pt electrodes, using the same parameters used for the adhesion layer (thickness ≈ 1 μm). The final devices were mounted on the transistor outline (TO) packages using electro-soldered gold wires.

The conductometric response of the fabricated heterostructure based sensors was investigated using a flow-through technique.²⁰ A custom stainless-steel chamber (1 L volume) located inside a climatic chamber (Angelantoni, Italy, model MTC 120) set at 20 °C was used to investigate the sensor response toward different gases. To simulate real ambient conditions, all the measurements were performed at a relative humidity of 50% with a constant synthetic air flow of 200 sccm. The atmosphere composition was controlled using mass flow controllers (MKS, Germany), mixing flows coming from certified gas bottles (SOL, Italy) containing a precise concentration of target analytes diluted in synthetic air. A constant bias of 1 V was applied to the sensing material to measure the output signal, recorded using a picoammeter (Keithley, USA). All the sensors were thermally stabilized at a desired temperature for 6 hours before starting the sensing measurement. The response was calculated by the variation of the conductance using the following formulas, for a p-type metal oxide under exposure to an oxidizing gas,²¹

$$\text{Response} = \frac{G_{\text{gas}} - G_{\text{air}}}{G_{\text{air}}} \quad (1)$$

and for the reducing gas

$$\text{Response} = \frac{G_{\text{air}} - G_{\text{gas}}}{G_{\text{gas}}} \quad (2)$$

where G_{gas} and G_{air} are the sensor conductance in the presence of gas and in synthetic air, respectively. Different concentrations of ozone gas [50; 150; 300 ppb] were generated using an ozone generator. The sensors were tested towards different gas analysts at different temperatures [200, 300, 400, and 500 °C]. Other interfering reducing and oxidizing analytes such as ethanol, acetone, hydrogen, nitrogen dioxide, carbon monoxide, and toluene were tested during the sensing measurements. Furthermore, the experimental data from calibration curves were fitted using typical power trend relationships for metal oxide sensors.²²

$$\text{Response} = A(\text{gas concentration})^B \quad (3)$$





Fig. 1 (a) and (b) SEM images of NiO nanowires grown at an evaporation temperature of 1400 °C, a pressure of 1 mbar, and 100 sccm argon flow using alumina substrates with the Au catalyst substrate at a temperature of 930 °C. (c) and (d) SEM images of NWO heterostructures grown at an evaporation temperature 1100 °C, a pressure of 1 mbar, and 100 sccm argon flow using alumina substrates with the Au catalyst substrate at a temperature of 580 °C.

where, A and B are constants typical of the sensor material and the stoichiometry of the involved reaction. The detection limits

for ozone were calculated by considering a minimum response of 1 to have a detectable and reliable signal.

Results and discussion

The morphology of the nanostructures was examined using FE-SEM. The bare NiO samples (Fig. 1(a) and (b)) on alumina substrates show a dispersed and uniform nanowire structure with diameters ranging from 10 to 70 nm and lengths along with a few large structures. In Fig. 1(c) and (d), the morphology of the heterostructure illustrates the small WO_3 nanowires growing on top of each core NiO nanowire, resulting in the formation of 3D-branch-like nanostructures comprised of three different materials NiO (core), NiWO_4 (seed) and WO_3 (secondary NWs).

Gas sensing performance

The sensing performance of the fabricated NWO heterostructures and pristine NiO NWs was tested for different concentrations of ozone (50, 50, 150 and 300 ppb) at different temperatures (200, 300, 400, and 500 °C). Fig. 2(a)–(d) show the isothermal dynamic-transient response of NWO heterostructures compared to that of the NiO NWs at temperatures



Fig. 2 Dynamic response curves of NiO and NWO sensing devices toward ozone gas (50, 50, 150, and 300 ppb) at temperatures of (a) 500 °C, (b) 400 °C, (c) 300 °C, and (d) 200 °C. All the measurements were performed at a relative humidity of 50% at 20 °C.



of 500, 400, 300 and 200 °C, respectively, and at a relative humidity level of 50%. Evidently, the baseline resistance of the NWO heterostructures is higher than that of the NiO NW sensors at all the operating temperatures. Specifically, the drop in the conductance baseline is near to one order of magnitude at a lower temperature (200 °C). The increased electrical resistance observed in NWO heterostructures is attributed to the built-in potential at the two interfaces (depletion region): one at the junction of NiO/NiWO₄ (p/p) and the other at the junction of NiWO₄/WO₃ (p/n). The charge transport mechanism in NWO heterostructures will be discussed in detail later.

Furthermore, the dynamic response of both NiO and NWO sensors shows an increase in electrical conductance under exposure to an oxidizing compound like ozone. This is the standard behavior of a p-type metal oxide sensor when exposed to an oxidizing analyte. This confirms that the hole is acting as the majority charge carrier in the heterostructure, which is due to the presence of two p-type metal oxide semiconductors

(NiO and NiWO₄) in the heterostructure system. The pristine NiO sensor shows low to no response when operated at a temperature of 500 °C compared to the NWO heterostructures. However, the dynamic response starts to increase at lower temperatures, with the highest dynamic response at 300 °C. Hence, 300 °C is the optimal working temperature for both the sensors. On the other hand, at lower temperatures such as 200 °C, both NiO and NWO sensors showed a significant response towards the O₃; however, the heterostructure sensor dynamic kinetics are much faster. The NiO sensors show a lower recovery rate compared to the NWO sensors. Furthermore, the response values were calculated from eqn (1) and (2). Fig. 3(a) reports the calculated response values of NWO and NiO sensors at different operating temperatures towards 300 ppb of O₃. The response of NWO sensors is comparably higher than that of NiO at temperatures of 200, 300, and 400 °C, while exhibiting comparable values at 500 °C. Evidently, at the optimal working temperature (300 °C), the response of NWO

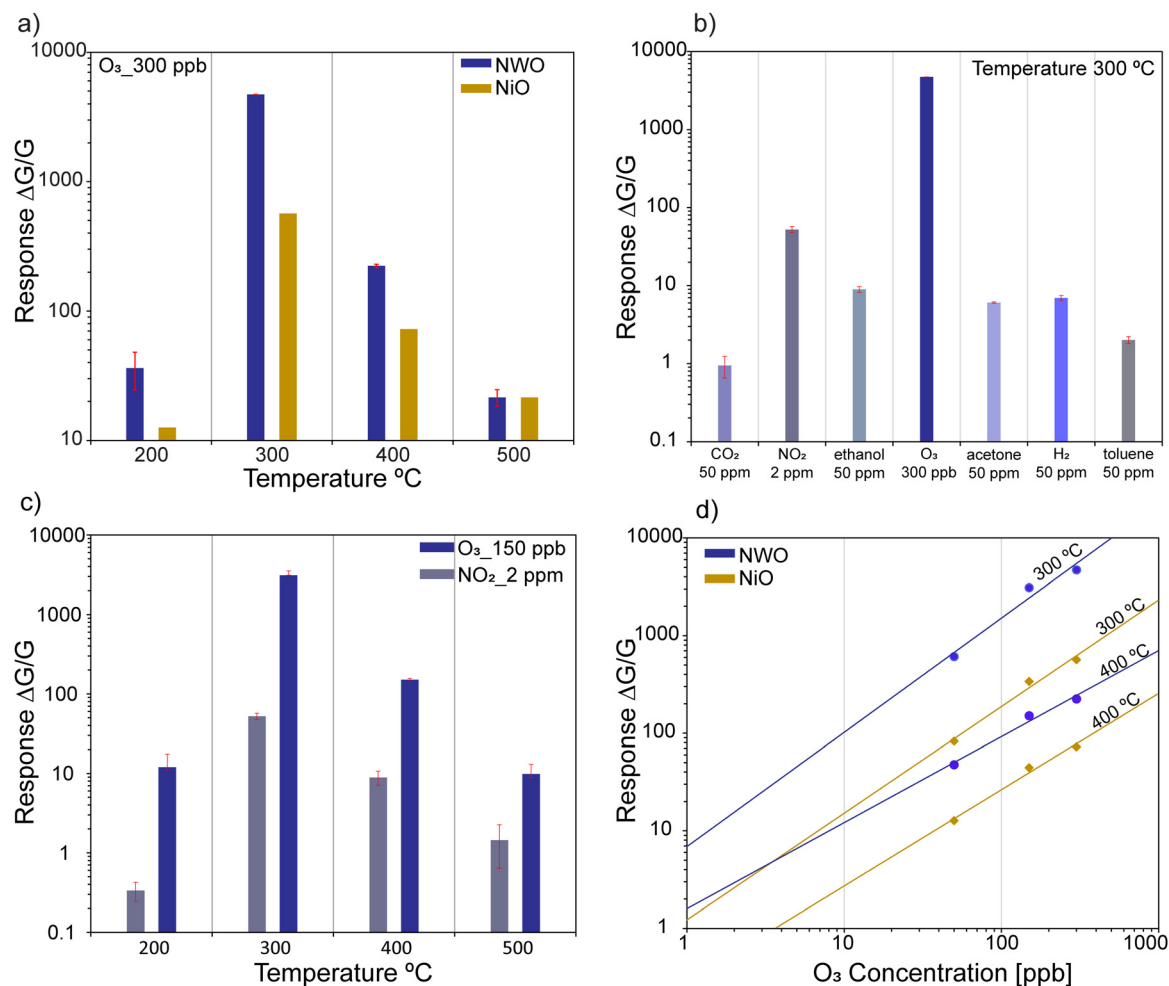


Fig. 3 (a) Temperature dependence responses of the NWO heterostructure and NiO sensors towards 300 ppb of O₃. (b) Response of the NWO heterostructure towards different interfering gases, ethanol (50 ppm), CO (50 ppm), NO₂ (2 ppm), H₂ (50 ppm), acetone (50 ppm), and toluene (50 ppm) with O₃ (300 ppb) at an optimal working temperature of 300 °C. (c) Temperature dependence comparison of NWO heterostructure sensors towards O₃ (150 ppb) and NO₂ (2 ppm) in the temperature range from 200 to 500 °C. (d) Calibration plot for the NWO and NiO sensors toward different concentrations of O₃ (50, 150, and 300 ppb) at temperatures of 300 °C and 400 °C.



sensors is 10 times higher compared to that of the pristine NiO NW sensors.

Furthermore, the sensing performance of NWO sensors was tested towards different reducing and oxidizing gas analytes at their optimal working temperatures. Fig. 3(b) reports the response of NWO sensors towards various gases such as hydrogen (50 ppm), CO (50 ppm), NO₂ (2 ppm), ethanol (50 ppm), acetone (50 ppm), and toluene (50 ppm) at 300 °C. Different concentrations of the gases were chosen according to their exposure limits according to the European Union (EU) Air Quality Standards.²³ For instance, testing at very high concentrations, such as 50 ppm of NO₂ and O₃, would not be significant for real-world conditions. According to the European Union (EU) Air Quality Standards, the permissible exposure limit for ozone is 0.1 ppm, with guidelines from organizations like the World Health Organization (WHO). Testing ozone at 50 ppm would not reflect practical exposure scenarios and thus would not be meaningful. Clearly, as shown in Fig. 3(b), the sensors showed the highest response towards ozone compared to the other analytes, proving their selective behavior towards ozone.

Investigating the selectivity of the sensors toward two major oxidizing gases like NO₂ and O₃ is very important. In Fig. 3(c), the response of the NWO sensors was compared at a range of temperatures from 200 to 500 °C towards 2 ppm of NO₂ and 150 ppb of O₃. Fig. 3(c) clearly shows that the NWO heterostructure sensors are more sensitive and selective towards O₃ at all the operating temperatures. In particular, the response of NWO sensors is almost 59 times higher towards O₃ compared to that of NO₂ sensors at 300 °C. Fig. 3(d) reports the calibration curves of NWO and NiO sensors, *i.e.*, response *versus* concentration, for O₃ (50, 150 and 300 ppb) at 300 °C and 400 °C. The calculated coefficients with power fitting (eqn (3)) of the curve and detection limits for both sensors are reported in Table S1 in the ESI.† Considering the minimum response of 1, the detection limits of NWO sensors were found to be 0.19 and 0.59 ppb at 300 and 400 °C, respectively, indicating that the NWO heterostructure sensors are highly sensitive under the given operating conditions.

Additionally, Fig. 4(a)–(f) compare the response dynamics and values under optimal working conditions towards O₃ for NWO heterostructure sensors with the host NiO NWs and to our previously reported WO₃ nanowires¹⁷ fabricated by using the VLS mechanism. The response dynamics of WO₃ follow an opposite trend compared to the NWO and NiO sensors due to its n-type semiconducting behavior.

Due to the electron-capturing nature of ozone, when it interacts with the surface of WO₃ NWs, the concentration of majority charge carriers (electrons) decreases and in turn decreases the conductance of the WO₃ sensor (Fig. 4(c)). This is a typical behavior of an n-type metal oxide semiconductor upon exposure to an oxidizing compound. Evidently, WO₃ sensors exhibit the lowest sensing performance compared to the pristine NiO and NWO sensors (Fig. 4(d)–(f)). The highest response was found in the order of 10, 100 and 1000 for WO₃, NiO and NWO 3D-heterostructure sensors, respectively.



Fig. 4 Dynamic response curves of NiO (a), NWO (b), and WO₃ (c) sensing devices toward ozone gas (50, 50, 150, and 300 ppb) at 300 °C. Response curves of NiO (d), NWO (e), and WO₃ (f) with respect to the concentration of O₃ at 300 °C.

These results prove that integrating a sensing platform consisting of different materials in the form of nanostructures enhances the sensing performance. Finally, the calculated response/recovery time of the NWO heterostructures is presented in Table S1 (ESI†). Moreover, all fabricated sensors demonstrated good repeatability of the sensor signal, as presented in the statistical data. Additionally, Fig. S2 (ESI†) shows that the NWO heterostructure sensing device maintained a stable response even after three months from the initial measurements at 300 °C towards 150 ppb. The variation in the response of the sensors to ozone was less than 20%.

Furthermore, a comparison of the sensing performance of the NWO nano-heterostructure sensors with the literature is presented in Table S2 (ESI†). Drawing direct comparisons between our fabricated sensors and existing data poses challenges, as no similar type of heterostructure has been previously tested for ozone detection. However, compared to sensors based on bare and modified WO₃^{24,25} and NiO²⁶ nanostructures, the NWO sensors exhibited superior performance. For instance, compared to other reported heterostructures such as In₂O₃/ZnO,²⁷ although their operating temperature is lower, their response is around 6.7, which is significantly lower. Additionally, for heterostructures based on V₂O₅ NWs and TiO₂,²⁸ despite having the same operating temperature as reported for NWO heterostructures, the response is much lower (1.4), even at a higher concentration of O₃ (1 ppm). Furthermore, L. Shao *et al.*²⁹ reported an ozone sensor based on Zn:MoS₂ nanosheets, which offers the advantage of room temperature operation. However, testing such a high concentration of O₃ (5000 ppb) with a response lower than 1 has very little significance for real-time sensing. All the aforementioned comparisons, along with additional reports from the literature, are presented in Table S2 (ESI†).



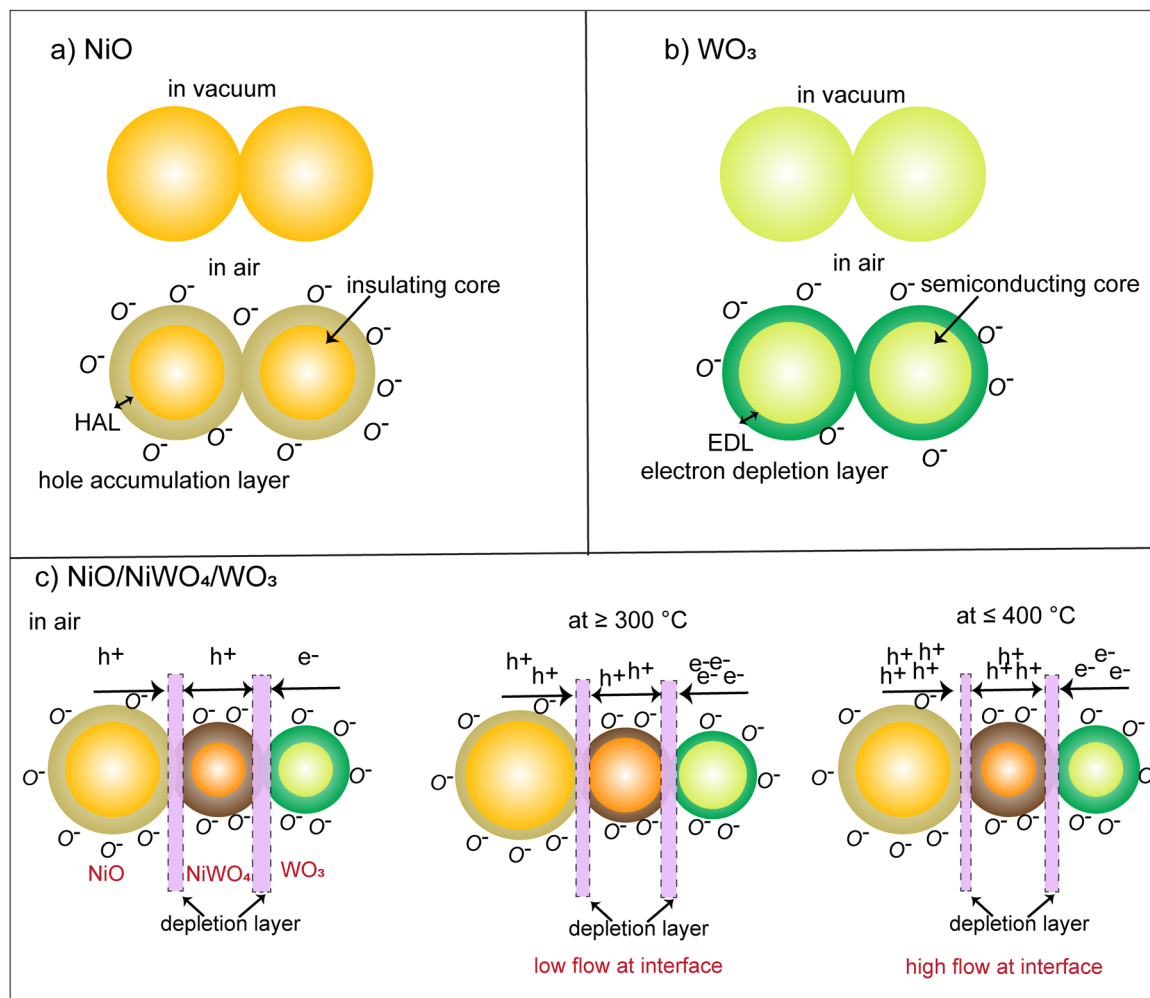


Fig. 5 (a) The effect of chemisorption on the conductance of NiO, (b) WO_3 and (c) $\text{NiO}/\text{NiWO}_4/\text{WO}_3$ heterostructures.

Charge transport properties of NWO heterostructures

Before we go into the details of the sensing mechanism in NWO heterostructures, it is important to discuss their charge transport properties. It is a well-known fact that the chemisorption of oxygen ions on the surface of MOXs determines their charge transport properties when operated in air. Specifically, the chemisorption of oxygen ions on the surface of p- and n-type MOXs extracts electrons and forms a core-shell structure.³⁰ In particular, a hole accumulation layer (HAL) around the resistive core and an electron depletion layer (EDL) around the conducting core are formed on the p- and n-type MOXs, respectively, as shown in Fig. 5(a) and (b). The formation of the HAL in p-type (NiO) MOXs increases their conductivity, while the conductivity of n-type MOXs (WO_3) falls due to the formation of the EDL.³¹ On the other hand, NWO is a double junction device with the n/p/p configuration in which the conduction is controlled by the holes due to its dominating p-character.⁸ Specifically, two interfaces or junctions *i.e.* one between $\text{WO}_3/\text{NiWO}_4$ (n-p) and the second at NiWO_4/NiO (p-p) are formed (Fig. 5(c)). In a vacuum, an interfacial potential barrier is formed at both interfaces due to the diffusion of

charge carriers. The existence of these interface barriers in the heterostructure is the main reason behind their lower conductance from the NiO NWs. Furthermore, when NWO is operated in air, the chemisorption of oxygen ions leads to the formation of an EDL on the WO_3 surface, while a HAL will be formed on NiO and NiWO_4 surfaces. With the increase in the operating temperature, the charge carriers (electrons and holes) gain enough energy and start crossing both potential barriers, leading to the phenomenon of electron-hole recombination especially at the $\text{WO}_3/\text{NiWO}_4$ interface.⁸ The rate of recombination is increased with an increase in temperature in which electrons from WO_3 are consumed significantly. Meanwhile, the density of holes is only slightly affected due to the dominating p-character of NWO. In fact, as the temperature increases, more and more holes are injected from NiO and NiWO_4 and hence start dominating the overall p-type character of NWO. This also implies that, at $300\text{ }^\circ\text{C}$, a high density of electrons will be present in the NWO compared to at $400\text{ }^\circ\text{C}$ and the density progressively decreases due to the electron-hole recombination as the temperature increases up to $500\text{ }^\circ\text{C}$.⁸ The whole phenomenon is pictorially depicted in Fig. 5(c). On the other hand, at $200\text{ }^\circ\text{C}$, the NWO heterostructure showed highly





Fig. 6 (a) Ozone sensing mechanism of pure NiO nanowires. (b) Ozone sensing mechanism of the NWO heterostructure sensor system. (c) Variation of the response and baseline conductance of the NWO heterostructure with respect to the applied temperature.

resistive behavior (see the dynamic response in Fig. 2(d)) and the charge transport is mainly controlled by the built-in potential at the two interfaces. It should be noted that the overall characteristics of the NWO heterostructure is always p-type due to the presence of two p-type materials, *i.e.* NiO and NiWO₄. This has already been indicated by the sensing data shown in Fig. 2, in which the conductivity of NWO heterostructures increases upon interacting with oxidizing compounds like O₃, which is a typical p-type MOX semiconductor behavior.

Sensing mechanism

As mentioned earlier, the gas sensing properties of MOXs are determined by the chemisorbed oxygen ions whose density and types vary with temperature.³¹ Specifically, in the temperature range from 100 to 300 °C, the chemisorbed oxygen ions are in the form of O⁻, which further changes to O²⁻ at temperatures > 300 °C *via* capturing the electrons from the surface of MOXs. To understand the ozone interaction mechanism with the p-type MOX like NiO, let us first discuss the properties of O₃. Ozone is a powerful oxidizing agent and a highly unstable gas molecule. Indeed, O₃ possesses a higher electron affinity $E_A = 2.14 \pm 0.15$ eV (ref. 32) compared to chemisorbed oxygen ions ($E_A = 0.43$ eV),³³ thus exhibiting a higher tendency to extract electrons deep from the surface of NiO NWs. Moreover, O₃ has a distinct dipole moment due to its molecular structure. In its configuration, the central atom carries a positive charge, while the surrounding atoms hold a negative charge.³⁴ When NiO NWs are exposed to O₃, the central positive atom extracts electrons from NiO, leading to the generation of free holes (wider than the thickness of the HAL), and increasing their conductance (see the dynamics in Fig. 2). The overall reaction of O₃ with NiO NWs is as follows:



The above relationship shows that the capturing of electrons by surface-adsorbed O₃ leads to the formation of O⁻ ions and the release of oxygen gas. Fig. 6(a) depicts the different stages in which NiO NWs are in air and exposed to ozone.

In the case of NWO heterostructure sensors, different factors are involved behind their superior performance. The very first reason is that the 3D nano-heterostructure provides a significantly higher reactive surface area for O₃ in comparison to 1D-NiO NWs. The symmetry of these heterostructures is also highly important to consider when discussing their performance. Using WO₃ as a secondary material on the core of NiO provides a highly electron-rich material on the outer surface of the whole heterostructure system. Furthermore, the temperature-dependent response of NWO heterostructures can be explained based on the charge transport mechanism described earlier. At 300 °C, the density of electrons is higher compared to higher temperatures due to the electron-hole recombination mechanism, as shown in Fig. 6(b). This can also be seen from the baseline conductance of NWO sensors in air (Fig. 6(c)) which is three times higher at 400 °C than at 300 °C. Due to the thermal excitation at higher temperatures, the electron-hole recombination through the different interfaces increases, leading to a higher conductance resulting in the reduction of the density of free electrons. Thus, when NWO heterostructures are exposed to O₃ at 300 °C, it is absorbed on the surface. These adsorbed O₃ then dissociate and capture bound electrons from the surface of the heterostructure that results in the formation of O⁻ species and the release of O₂ gas. On the other hand, at 200 °C, the NWO heterostructure is apparently highly resistive (Fig. 6(c)), thus exhibiting poor sensing performance, but retains a higher response value in comparison to NiO NWs due to the presence of n-type WO₃ and enhanced reactive sites in heterostructures.



Furthermore, even though both O₃ and NO₂ are oxidizing and highly reactive agents, the NWO heterostructure exhibited better performance toward O₃. The answer to this question lies in their molecular structure that highly affects the reaction mechanism. Ozone is a triatomic molecule with three oxygen atoms arranged in a O–O–O manner creating an overall bent molecular geometry, with an angle of approx. 116° between the oxygen atoms.^{34,35} On the other hand, in the case of NO₂, the diatomic molecule consists of one nitrogen atom and two oxygen atoms, resulting in a linear molecular geometry.^{36,37} Comparing both, the O₃ is much more reactive due to the presence of three oxygen atoms that provide a higher degree of molecular strain leading to its high reactivity and is prone to undergo chemical reactions in the presence of other substances and with the metal oxide surface. Thus, the oxidizing potential of O₃ is much higher than that of NO₂, as it can readily react with the MOX surface and lead to the formation of adsorbed oxygen species. It is important to note that these two are gaseous pollutants with similar chemical properties, making it difficult to differentiate between them using traditional sensing MOX sensors. However, using advanced tailored heterostructures offers a promising solution, enabling the development of highly sensitive and selective sensors capable of distinguishing between different analytes with higher precision.

Conclusions

A novel highly selective ozone sensor was fabricated based on complex NiO/NiWO₄/WO₃ heterostructures. The NWO heterostructures were successfully synthesized using the single vapor-phase growth method employing a two-step growth technique. In the first step, NiO NWs were grown directly on the Au-catalysed alumina substrate using the VLS mechanism. Secondly, the small WO₃ NWs were grown directly on the NiO core *via* the VS mechanism from a seed of the ternary material NiWO₄, creating 3D branch-like nano-heterostructures. Furthermore, the fabricated NWO sensors exhibited superior performance in comparison to pristine NiO and WO₃ NWs towards the detection of ozone at the ppb level. In particular, the heterostructure sensors demonstrated a higher and more selective response at 300 °C towards ozone compared to other oxidizing and reducing gas analytes. The main reason behind the superior performance is attributed to the formation of different interfaces between the different MOXs (p/p of NiO/NiWO₄ and p/n of NiWO₄/WO₃). By controlling the charge carrier flow and density through these interfaces with applied temperature, the sensor exhibited a selective response towards O₃ at 300 °C. In particular, the presence of WO₃ provides free electrons on the surface of heterostructures, which makes them more reactive and selective towards an oxidizing gas like O₃. Additionally, fabricated sensors based on 3D nano-heterostructures provide larger reactive surface areas and more reactive sites compared to the pristine NW sensors, leading to their enhanced and superior sensing performance. Finally, the NWO heterostructure sensor exhibited a response value of 4705 ± 9 towards 300 ppb of ozone, with a

detection limit of 0.19 ppb at 300 °C. This makes them promising candidates for advancing cutting-edge environmental/health monitoring technologies.

Author contributions

Navpreet Kaur: conceived the idea, planned the work, did all the experiments and data analysis, and wrote the manuscript. Elisabetta Comini: as the Head of the sensor lab supervised, provided funding for the research work, and edited the manuscript.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This study was carried out within the MOST – Sustainable Mobility National Research Center and received funding from the European Union Next-Generation EU (PIANO NAZIONALE DI RIPRESA E RESILIENZA (PNRR) – MISSIONE 4 COMPONENTE 2, INVESTIMENTO 1.4 – D.D. 1033 17/06/2022, CN00000023), Spoke 5 “Light Vehicle and Active Mobility”. This manuscript reflects only the authors’ views and opinions, neither the European Union nor the European Commission can be considered responsible for them.

References

- 1 D. R. Miller, S. A. Akbar and P. A. Morris, *Sens. Actuators, B*, 2014, **204**, 250–272.
- 2 F. J. Meng, R. F. Xin and S. X. Li, *Materials*, 2023, **16**, 263.
- 3 M. W. Alam, P. Pooja, M. Aamir, B. Souayeh, S. Mushtaq, M. S. Khan, M. N. Amin, K. Khan and S. Shajahan, *Membranes*, 2022, **12**, 555.
- 4 D. Zappa, V. Galstyan, N. Kaur, H. M. M. Munasinghe Arachchige, O. Sisman and E. Comini, *Anal. Chim. Acta*, 2018, **1039**, 1–23.
- 5 A. Singh, S. Sikarwar, A. Verma and B. Chandra Yadav, *Sens. Actuators, A*, 2021, **332**, 113127.
- 6 N. Kaur, D. Zappa, M. Ferroni, N. Poli, M. Campanini, R. Negrea and E. Comini, *Sens. Actuators, B*, 2018, **262**, 477–485.
- 7 N. Kaur, E. Comini, N. Poli, D. Zappa and G. Sberveglieri, *Procedia Eng.*, 2016, **168**, 1140–1143.
- 8 N. Kaur, D. Zappa, V.-A. Maraloiu and E. Comini, *Adv. Funct. Mater.*, 2021, 2104416.
- 9 Z. Zhu, J. L. Chang, C. H. Wu, T. L. Chou and R. J. Wu, *Sens. Actuators, B*, 2016, **232**, 442–447.
- 10 S. Vallejos, V. Khatko, J. Calderer, I. Gracia, C. Cané, E. Llobet and X. Correig, *Sens. Actuators, B*, 2008, **132**, 209–215.



- 11 T. Becker, L. Tomasi, C. Bosch-V.braunmühl, G. Müller, G. Sberveglieri, G. Fagli and E. Comini, *Sens. Actuators, A*, 1999, **74**, 229–232.
- 12 S. C. Mills, B. Lee and V. Misra, *ECS Meet. Abstr.*, 2015, **MA2015-01**, 2144.
- 13 J. A. Logan, *J. Geophys. Res.*, 1985, **90**, 10463–10482.
- 14 P. J. D. Peterson, A. Aujla, K. H. Grant, A. G. Brundle, M. R. Thompson, J. Vande Hey and R. J. Leigh, *Sensors*, 2017, **17**, 1653.
- 15 Z. Zhu, J. L. Chang and R. J. Wu, *Sens. Actuators, B*, 2015, **214**, 56–62.
- 16 R. Souissi, N. Bouguila, M. Bendahan, K. Aguir, T. Fiorido, M. Abderrabba, I. Halidou and A. Labidi, *J. Alloys Compd.*, 2022, **900**, 163513.
- 17 N. Kaur, D. Zappa, N. Poli and E. Comini, *ACS Omega*, 2019, **4**, 16336–16343.
- 18 N. Kaur, E. Comini, D. Zappa, N. Poli and G. Sberveglieri, *Nanotechnology*, 2016, **27**, 205701.
- 19 K. T. Jacob, *J. Mater. Sci.*, 1977, **12**, 1647–1652.
- 20 G. Sberveglieri, G. Faglia, C. Perego, P. Nelli, R. N. Marks, T. Virgili, C. Taliani and R. Zamboni, *Synth. Met.*, 1996, **77**, 273–275.
- 21 N. Kaur, M. Singh, A. Casotto, L. Sangaletti and E. Comini, *Chem. Commun.*, 2023, **59**, 1329–1332.
- 22 D. Zappa, A. Bertuna, E. Comini, N. Kaur, N. Poli, V. Sberveglieri and G. Sberveglieri, *Beilstein J. Nanotechnol.*, 2017, **8**, 1205–1217.
- 23 Standards - Air Quality - Environment - European Commission, <https://ec.europa.eu/environment/air/quality/standards.htm>, (accessed 28 January 2020).
- 24 M. Bendahan, R. Boulmani, J. L. Seguin and K. Aguir, *Sens. Actuators, B*, 2004, **100**, 320–324.
- 25 A. Labidi, E. Gillet, R. Delamare, M. Maaref and K. Aguir, *Sens. Actuators, B*, 2006, **120**, 338–345.
- 26 A. Paralikis, E. Gagaoudakis, V. Kampitakis, E. Aperathitis, G. Kiriakidis and V. Binas, *Appl. Sci.*, 2021, **11**, 3104.
- 27 N. Sui, Y. Xu, P. Zhang, S. Cao, T. Zhou and T. Zhang, *Sens. Actuators, B*, 2023, **380**, 133312.
- 28 W. Avansi, A. C. Catto, L. F. Da Silva, T. Fiorido, S. Bernardini, V. R. Mastelaro, K. Aguir and R. Arenal, *ACS Appl. Nano Mater.*, 2019, **2**, 4756–4764.
- 29 L. Shao, Z. Wu, H. Duan and T. Shaymurat, *Sens. Actuators, B*, 2018, **258**, 937–946.
- 30 N. Kaur, M. Singh and E. Comini, *Langmuir*, 2020, **36**, 6326–6344.
- 31 N. Kaur, M. Singh and E. Comini, *Langmuir*, 2020, **36**, 6326–6344.
- 32 E. W. Rothe, S. Y. Tang and G. P. Reck, *J. Chem. Phys.*, 1975, **62**, 3829–3831.
- 33 N. Kaur, D. Zappa and E. Comini, *Electron. Mater. Lett.*, 2019, **15**, 743–749.
- 34 E. Miliordos and S. S. Xantheas, *J. Am. Chem. Soc.*, 2014, **136**, 2808–2817.
- 35 D. Bermúdez-Aguirre and G. V. Barbosa-Cánovas, *Food Control*, 2013, **29**, 82–90.
- 36 P. Gray and A. D. Yoffe, *Chem. Rev.*, 1955, **55**, 1069–1154.
- 37 L. F. Gamon and U. Wille, *Acc. Chem. Res.*, 2016, **49**, 2136–2145.

