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## Theoretical insights into the electroreduction of nitrate to ammonia on graphene-based single-atom catalysts†

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Electrocatalytic reduction of harmful nitrate ( $\text{NO}_3^-$ ) to valuable ammonia ( $\text{eNO}_3\text{RR}$ ) is critical and attractive for both environmental remediation and energy transformation. A single atom catalyst (SAC) based on graphene represents one of the most promising  $\text{eNO}_3\text{RR}$  catalysts. However, the underlying catalytic mechanism and the intrinsic factors dictating the catalytic activity trend remain unclear. Herein, using first-principles calculations,  $\text{eNO}_3\text{RR}$  on  $\text{TMN}_3$  and  $\text{TMN}_4$  (TM = Ti–Ni) doped graphene was thoroughly investigated. Our results reveal that  $\text{FeN}_4$  doped graphene exhibits excellent  $\text{eNO}_3\text{RR}$  performance with a low limiting potential of  $-0.38$  V, agreeing with the experimental finding, which can be ascribed to the effective adsorption and activation of  $\text{NO}_3^-$  via the charge “acceptance–donation” mechanism and its moderate binding due to the occupation of the d–p antibonding orbital. In particular, we found that  $\text{eNO}_3\text{RR}$  activities are well correlated with the intrinsic properties of TM centers and their local environments. With the established activity descriptor, several other graphene-based SACs were efficiently screened out with excellent  $\text{eNO}_3\text{RR}$  performance. Our studies could not only provide an atomic insight into the catalytic mechanism and activity origin of  $\text{eNO}_3\text{RR}$  on graphene-based SACs, but also open an avenue for the rational design of SACs for  $\text{eNO}_3\text{RR}$  towards ammonia by regulating the metal center and its local coordination environment.

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### 1. Introduction

Ammonia ( $\text{NH}_3$ ) is not only one of the most fundamental industrial chemicals but also acts as an emerging energy storage medium and carbon-free energy carrier.<sup>1,2</sup> For more than one century,  $\text{NH}_3$  synthesis in industry has mainly relied on the Haber–Bosch technology operating under harsh conditions. Considering that the Haber–Bosch process is both energy- and capital-intensive, in recent years the electrochemical synthesis of  $\text{NH}_3$  through the nitrogen reduction reaction (NRR), driven by renewable energy, has attracted great

research interest.<sup>3,4</sup> However, generally the NRR suffers from an unsatisfactory conversion rate and low faradaic efficiency, resulting from the extremely inert  $\text{N}_2$  molecule.<sup>5,6</sup> Given the much weaker  $\text{N}=\text{O}$  bond ( $204 \text{ kJ mol}^{-1}$ ) and  $\text{N}-\text{O}$  bond ( $176 \text{ kJ mol}^{-1}$ ) than the  $\text{N}\equiv\text{N}$  bond ( $941 \text{ kJ mol}^{-1}$ ), very recently, other nitrogen resources, including nitrate ( $\text{NO}_3^-$ ), nitrite ( $\text{NO}_2^-$ ), and nitric oxide (NO), have attracted more and more attention for  $\text{NH}_3$  synthesis.<sup>7,8</sup> As the third-largest accessible form of nitrogen<sup>9</sup> and most oxidized reactive nitrogen,<sup>10</sup>  $\text{NO}_3^-$  has been considered as one of the most widespread water pollutants in the world.<sup>11,12</sup> Under this circumstance, the electrocatalytic  $\text{NO}_3^-$  reduction reaction ( $\text{eNO}_3\text{RR}$ ) to produce  $\text{NH}_3$  is considered as a promising alternative to the NRR because it can simultaneously remediate environmental pollution and yield valuable  $\text{NH}_3$  potentially more efficiently than the NRR.<sup>13,14</sup>

Rational design of the electrocatalyst with excellent activity, high selectivity, and long durability is critical for the development of a practical  $\text{eNO}_3\text{RR}$  device. Encouragingly, various catalysts, such as molecular solids,<sup>13</sup> pure metals,<sup>7,15,16</sup> metallic alloys,<sup>17–19</sup> metal compounds,<sup>20–23</sup> and atomic catalysts,<sup>24–34</sup> have been demonstrated to be highly efficient for  $\text{eNO}_3\text{RR}$ . For example, it is shown that, at  $-0.4$  V versus the reversible hydro-

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gen electrode, an  $\text{NH}_3$  production rate of  $436 \pm 85 \mu\text{g h}^{-1} \text{cm}^{-2}$  and ultrahigh faradaic efficiency of 85.9% can be achieved on the copper–molecular solid catalyst.<sup>13</sup> Single atom catalysts (SACs), such as  $\text{TMN}_3$  and  $\text{TMN}_4$ -doped graphene (denoted as  $\text{TMN}_3@G$  and  $\text{TMN}_4@G$ , respectively), have attracted huge attention in electrocatalysis due to their unique electronic and geometric structures as well as high stability.<sup>35,36</sup> For  $\text{eNO}_3\text{RR}$ , several SACs have been experimentally reported to be excellent,<sup>24–28</sup> and theoretically predicted to be promising.<sup>29–33</sup> For example,  $\text{FeN}_4@G$  can exhibit an  $\text{NH}_3$  yield rate of up to  $\sim 20\,000 \mu\text{g h}^{-1} \text{mg}_{\text{cat}}^{-1}$  and a maximal faradaic efficiency of  $\sim 75\%$ .<sup>25</sup>

Besides the common merits of SACs, graphene-based SACs can be considered as an ideal platform for understanding the fundamental physical chemistry of a specific reaction from first-principles calculations, due to the distinct geometric and electronic characteristics of the graphene support.<sup>37</sup> Moreover, the coordination environment of the TM center embedded in graphene can be purposely tuned *via* direct and indirect coordination engineering, which produces huge numbers of SAC systems and makes graphene-based SACs find various potential applications, especially for electrocatalysis.<sup>38,39</sup> As discussed above, until now only  $\text{FeN}_4@G$  has been investigated for  $\text{eNO}_3\text{RR}$  towards  $\text{NH}_3$ .<sup>25</sup> Thus, the investigation of  $\text{eNO}_3\text{RR}$  catalyzed by graphene-based SACs is still in its infancy, and it is urgent to unravel the underlying catalytic mechanism and activity origin, and identify the intrinsic factors dictating their catalytic activity trend, guided by which high performance graphene-based SACs for  $\text{eNO}_3\text{RR}$  can be rationally designed effectively.

Based on the above considerations, herein we systematically studied the  $\text{eNO}_3\text{RR}$  on  $\text{TMN}_3@G$  and  $\text{TMN}_4@G$  (TM = Ti, V, Cr, Mn, Fe, Co, Ni, and Cu), through first-principles calculations. The detailed reaction pathway for the electroreduction of  $\text{NO}_3^-$  into  $\text{NH}_3$  was explored thoroughly, to illustrate the activity and selectivity. It is found that  $\text{FeN}_4@G$  stands at the top of the volcano among  $\text{TMN}_4@G$  with a limiting potential ( $U_L$ )  $-0.38$  V, and exhibits the highest  $\text{eNO}_3\text{RR}$  performance, in agreement with the experimental findings.<sup>25</sup> In addition, among  $\text{TMN}_3@G$ ,  $\text{CoN}_3@G$  exhibits the highest  $\text{eNO}_3\text{RR}$  performance with  $U_L$  of  $-0.65$  V. Then, the  $\text{eNO}_3\text{RR}$  activity origin was investigated through electronic structure analysis. Finally, an intrinsic descriptor to uncover the  $\text{eNO}_3\text{RR}$  activity trend for the graphene-based SAC was proposed, which correlates with the local environment of the TM atom, *i.e.*, its electronegativity and coordination number, and the electronegativity of its nearest neighbor atoms. The identified descriptor is distinct from the previously proposed one, that is, the binding free energy of  $\text{NO}_3^-$  ( $\Delta G(\text{NO}_3^*)$ ),<sup>29–33</sup> which is based on energies that are hard to measure and modulate in experiments. In addition, guided by the proposed descriptor, several other graphene-based SACs with high  $\text{eNO}_3\text{RR}$  performance were efficiently screened out, and especially the non-noble metal Ni SAC coordinated with one nitrogen and two carbon atoms exhibits a high intrinsic activity with  $U_L$  of  $-0.41$  V.

## 2. Computational details

In this work, spin-polarized first-principles calculations were carried out using Vienna *Ab initio* Simulation Package (VASP).<sup>40</sup> The projector augmented wave (PAW) method was employed to account for the electron–ion interactions.<sup>41</sup> The exchange–correlation effect was described by the Perdew–Burke–Ernzerhof (PBE) functional within the generalized gradient approximation (GGA).<sup>42</sup> The cutoff energy for the plane-wave basis set was taken as 450 eV. The Grimme’s DFT-D3 method was adopted to account for the van der Waals interaction.<sup>43</sup> The energy and force convergence criteria for the geometry optimizations were set as  $10^{-5}$  eV and  $0.02 \text{ eV \AA}^{-1}$ , respectively. The Brillouin zone of the  $7 \times 7 \times 1$  graphene supercell was sampled by  $3 \times 3 \times 1$  and  $6 \times 6 \times 1$  *k*-point meshes for structural optimization and calculation of densities of states (DOS), respectively. A vacuum thickness of  $\sim 20 \text{ \AA}$  was placed along the *z* direction to minimize the interaction between periodic units. The thermal stability was estimated through *ab initio* molecular dynamics (AIMD) simulations.

According to the computational hydrogen electrode (CHE) model,<sup>44</sup> the Gibbs free energy change ( $\Delta G$ ) for each elemental step was calculated by

$$\Delta G = \Delta E + \Delta E_{\text{ZPE}} - T\Delta S.$$

In the equation,  $\Delta E$  was directly obtained from DFT calculations.  $\Delta E_{\text{ZPE}}$  and  $T\Delta S$  ( $T = 298.15$  K) are the contributions of zero-point energy and entropy, respectively, which were acquired by computing the vibrational frequency of intermediates, as listed in Table S1,† while  $\Delta E_{\text{ZPE}}$  and  $T\Delta S$  of the free molecules were taken from the NIST database (Table S2†).<sup>45</sup> For the CHE model, for the reaction  $\text{A}^* + \text{H}^+ + \text{e}^- \rightarrow \text{AH}^*$ , the free energy change is calculated according to  $G(\text{AH}^*) - [\mu(\text{H}^+) + \mu(\text{e}^-)] - G(\text{A}^*)$ . The CHE model shows that the chemical potential of a proton–electron pair  $[\mu(\text{H}^+) + \mu(\text{e}^-)]$  at zero potential can be regarded as half the chemical potential of the free  $\text{H}_2$  molecule  $[1/2\mu(\text{H}_2)]$ , *i.e.*,  $\mu(\text{H}^+) + \mu(\text{e}^-) = 1/2\mu(\text{H}_2)$ . Based on this model,  $G(\text{AH}^*) - [\mu(\text{H}^+) + \mu(\text{e}^-)] - G(\text{A}^*)$  can be calculated as  $G(\text{AH}^*) - 1/2\mu(\text{H}_2) - G(\text{A}^*)$ .  $U_L$  is formalized with  $U_L = \Delta G_{\text{max}}/e$ , where  $\Delta G_{\text{max}}$  is the free energy change of the potential-determining step (PDS).<sup>46</sup> When computing the binding free energy of  $\text{NO}_3^-$  ( $\Delta G(\text{NO}_3^*)$ ), the gaseous  $\text{HNO}_3$  was adopted as a reference, and due to that directly calculating the energy of charged  $\text{NO}_3^-$  is difficult.<sup>47,48</sup> Then  $\Delta G(\text{NO}_3^*)$  was obtained by

$$\Delta G(\text{NO}_3^*) = G(\text{NO}_3^*) - G^* - [G(\text{HNO}_3) - 1/2G(\text{H}_2)] + \Delta G_{\text{correct}}$$

in which  $\Delta G(\text{NO}_3^*)$  and  $G^*$  are the Gibbs free energy of the  $\text{NO}_3^-$  adsorbed SACs and the pristine SACs, respectively.  $G(\text{HNO}_3)$  and  $G(\text{H}_2)$  denote the Gibbs free energy of gaseous  $\text{HNO}_3$  and  $\text{H}_2$  molecules, respectively. The free energy correction,  $\Delta G_{\text{correct}}$ , is set as 0.392 eV.<sup>47,48</sup> The solvation effect, realized by the implicit solvation model,<sup>49</sup> on the free energy during the reaction process has been tested for  $\text{FeN}_4@G$ . As presented below, the  $U_L$  and reaction pathway are insensitive

to solvation correction. Therefore, we did not consider the solvation effect for the other systems, as done in previous works.<sup>29,31</sup>

### 3. Results and discussion

#### 3.1. $\text{NO}_3^-$ adsorption

$\text{TMN}_3@\text{G}$  and  $\text{TMN}_4@\text{G}$  have been widely studied both experimentally and theoretically for various catalytic reactions, especially the electrocatalysis reaction because the graphene support has a large surface area, excellent conductivity, and high (electro)-chemical stability.  $\text{TMN}_3@\text{G}$  and  $\text{TMN}_4@\text{G}$  feature TM single atoms that are coordinated with three and four pyridine nitrogen atoms, respectively, in the graphene matrix, which could significantly influence their activity, selectivity, and stability, due to the strong electronic-metal support interaction.<sup>50,51</sup> Typical atomic configurations of  $\text{TMN}_3@\text{G}$  and  $\text{TMN}_4@\text{G}$  are presented in Fig. 1 and S1a, S1b,† and seven TMs (Ti, V, Cr, Mn, Fe, Co, and Ni) have been considered as the catalytic active center. The key parameters of these systems are compiled in Table S3,† such as TM–N bond lengths, charge transfers,<sup>52</sup> and spin magnetic moments, which agree well with previous studies.<sup>53</sup> The total DOS for all these SACs presented in Fig. S1c and S1d† suggest their metallic characteristics, beneficial to the charge transfer during  $\text{eNO}_3\text{RR}$ .<sup>30</sup>

Effective adsorption of  $\text{NO}_3^-$  is a prerequisite to trigger the  $\text{eNO}_3\text{RR}$  process. Hence, we first investigated the adsorption behavior of  $\text{NO}_3^-$  on these SACs. Various adsorption configurations have been considered with one or two O atoms bonded with TM centers, and the most stable ones are shown in Fig. 2a, for which the key structural parameters are displayed in Table S4.† We can see that on  $\text{TMN}_3@\text{G}$ ,  $\text{NO}_3^-$  tends to be adsorbed *via* two O atoms, while on  $\text{TMN}_4@\text{G}$ ,  $\text{NO}_3^-$  prefers to

be adsorbed *via* one O atom, except  $\text{TiN}_4$  and  $\text{VN}_4@\text{G}$ . Accordingly, the values of  $\Delta G(\text{NO}_3^*)$  (Table S4†) indicate that for the same TM,  $\text{TMN}_3@\text{G}$  possesses higher  $\text{NO}_3^-$  binding strength than  $\text{TMN}_4@\text{G}$ , and for the same coordination the adsorption strength gradually decreases with the active center changing from Ti to Ni. Then the charge transfer from the supports to  $\text{NO}_3^-$  was calculated (Table S4†). From Fig. 2b, we can see that there is a good linear correlation between the charge gained by the adsorbates and their binding strengths to  $\text{TMN}_3@\text{G}$ , *i.e.*, the more strongly  $\text{NO}_3^-$  binds, the more negatively charged they are.<sup>29,30</sup> In contrast, the correlation between the charge of the adsorbates and the binding strengths for  $\text{TMN}_4@\text{G}$  is rather poor, which may result from the distinct adsorption configurations of  $\text{NO}_3^-$  on different  $\text{TMN}_4@\text{G}$ , as shown in Fig. 2a. The different correlations between  $\text{NO}_3^-$  binding strength on  $\text{TMN}_3@\text{G}$  and  $\text{TMN}_4@\text{G}$  reflect the coordination environment effects.

Moreover, in solution, the hydrogen evolution reaction (HER) is the main competing reaction against  $\text{eNO}_3\text{RR}$ . To effectively suppress the HER,  $\text{NO}_3^-$  should be bonded more strongly than H.<sup>29,30</sup> We, therefore, studied the adsorption of H atoms on all the investigated systems, for which the optimized structures are presented in Fig. S2.† As shown in Fig. 2c, for all the cases the binding free energy of  $\text{H}^*$  ( $\Delta G(\text{H}^*)$ ) is much more positive than  $\Delta G(\text{NO}_3^*)$ , indicating that the active site will be preferentially occupied by  $\text{NO}_3^-$  rather than H, and thus the HER can be effectively suppressed.

#### 3.2. Reaction mechanism of $\text{eNO}_3\text{RR}$

Although  $\text{eNO}_3\text{RR}$  is complicated and potentially has a variety of products (such as  $\text{NO}_2$ ,  $\text{NO}$ ,  $\text{N}_2\text{O}$ , and  $\text{N}_2$ ), the Pourbaix diagram of nitrogen species indicates that  $\text{NH}_3/\text{NH}_4^+$  is the most stable product in thermodynamics under the negative electrode potential.<sup>29,54</sup> Thus, it is quite necessary to fully explore the detailed reaction process of  $\text{NO}_3^-$  electroreduction

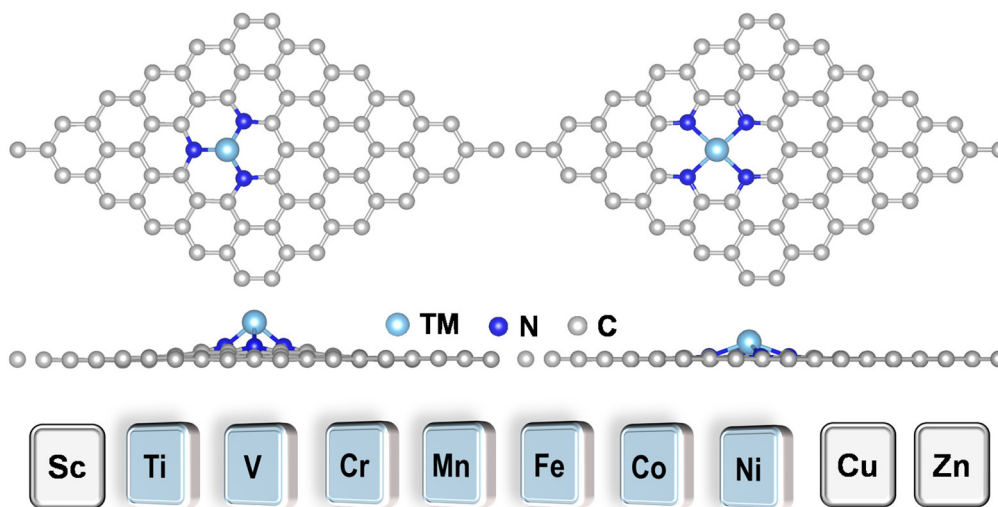


Fig. 1 Typical atomic configurations of the  $\text{TMN}_3@\text{G}$  (left) and  $\text{TMN}_4@\text{G}$  (right) systems. The considered TM atoms include Ti, V, Cr, Mn, Fe, Co, and Ni.

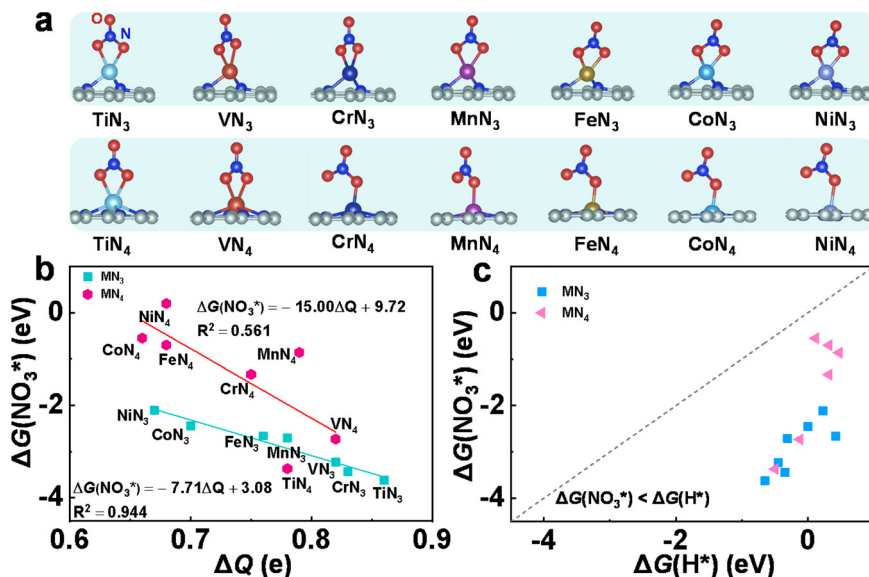


Fig. 2 (a) Atomic configurations of the TMN<sub>3</sub>@G (upper panel) and TMN<sub>4</sub>@G (lower panel) with the adsorbed NO<sub>3</sub><sup>-</sup>. (b) The relationship between charge transfer (ΔQ) of the NO<sub>3</sub><sup>-</sup> and its binding free energy (ΔG(NO<sub>3</sub><sup>\*</sup>)). (c) Comparison of the binding free energies of NO<sub>3</sub><sup>-</sup> (ΔG(NO<sub>3</sub><sup>\*</sup>)) and H atom (ΔG(H<sup>\*</sup>)).

into NH<sub>3</sub>. Given that there are multiple adsorption structures and hydrogenation positions of intermediates, during the eNO<sub>3</sub>RR process the most stable intermediates in each elementary step have been selected for further reaction to efficiently screen out the optimal eNO<sub>3</sub>RR pathway on TMN<sub>3</sub>@G and TMN<sub>4</sub>@G. The adsorption configurations of the key intermediate, NO<sup>\*</sup>, are displayed in Fig. S3.† In most cases, NO adsorption adopts the N-end pattern with the N atom bonded with the TM atom, suggesting that the N-end pathway proposed previously should be feasible.<sup>29,30</sup> The calculated U<sub>L</sub> values and the corresponding PDS are summarized in Fig. 3. Interestingly, for most systems, OH<sup>\*</sup> + H<sup>+</sup> + e<sup>-</sup> → \* + H<sub>2</sub>O is the PDS, and the strongly bonded OH<sup>\*</sup> group could lead to catalyst poisoning.

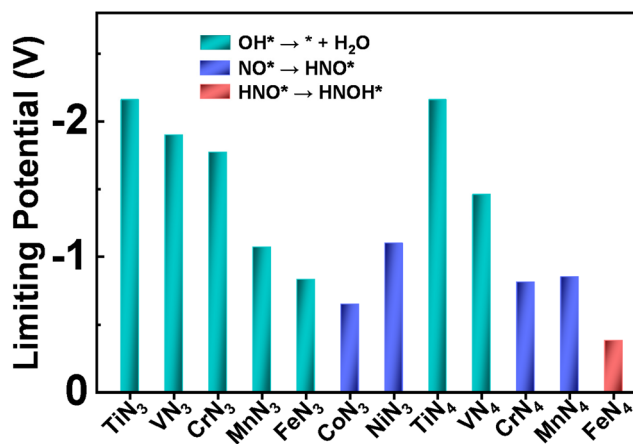


Fig. 3 Summary of the limiting potential and potential-determining step for eNO<sub>3</sub>RR on various SACs along the most favorable pathway.

Moreover, generally U<sub>L</sub> exhibits a general decreasing trend changed from early TM to late TM for both TMN<sub>3</sub> and TMN<sub>4</sub>@G. Importantly, CoN<sub>3</sub> and FeN<sub>4</sub>@G exhibit the highest eNO<sub>3</sub>RR activity among the considered TMN<sub>3</sub>@G and TMN<sub>4</sub>@G systems, respectively. The PDS of CoN<sub>3</sub> and FeN<sub>4</sub>@G are NO<sup>\*</sup> + H<sup>+</sup> + e<sup>-</sup> → HNO<sup>\*</sup> and HNO<sup>\*</sup> + H<sup>+</sup> + e<sup>-</sup> → HNOH<sup>\*</sup> with U<sub>L</sub> values of -0.65 and -0.38 V, respectively. Note that the full reaction pathways of CoN<sub>4</sub> and NiN<sub>4</sub>@G systems have not been studied, due to that they are unable to effectively capture the key reaction intermediate (NO<sub>3</sub>H<sup>\*</sup>), as shown in Fig. S4.†

To get insight into the reaction mechanism of eNO<sub>3</sub>RR, the reaction pathways on CoN<sub>3</sub> and FeN<sub>4</sub>@G are presented in Fig. 4, and those for the other systems in Fig. S5–S14.† Firstly, NO<sub>3</sub><sup>-</sup> is adsorbed on CoN<sub>3</sub>@G (Fig. 4a) and FeN<sub>4</sub>@G (Fig. 4b) with two O–Co bonds and one O–Fe bond, respectively, with ΔG(NO<sub>3</sub><sup>\*</sup>) values of -2.45 and -0.70 eV. Distinct binding strengths of NO<sub>3</sub><sup>-</sup> will lead to different activation and different free energy changes in the first hydrogenation step. Interestingly, the first proton–electron pair prefers to attack the O atom of NO<sub>3</sub><sup>\*</sup> bonded with Co (Fe) rather than the terminal O atom. Consequently, one N–O bond is cleaved to form NO<sub>2</sub><sup>\*</sup> + OH<sup>\*</sup> on CoN<sub>3</sub>@G with an energy release of -0.72 eV, while on FeN<sub>4</sub>@G, NO<sub>3</sub>H<sup>\*</sup> is formed with an energy input of 0.23 eV. The second hydrogenation step leads to the formation of a H<sub>2</sub>O molecule. The intermediate, NO<sub>2</sub><sup>\*</sup>, binds with the Co atom *via* two Co–O bonds, while it binds with the Fe atom with the Fe–N bond, and the processes are exothermic by 0.82 and 2.28 eV, respectively, on CoN<sub>3</sub> and FeN<sub>4</sub>@G. The subsequent proton–electron pair transfer leads to the hydrogenation of NO<sub>2</sub><sup>\*</sup>. On CoN<sub>3</sub>@G, NO<sup>\*</sup> + OH<sup>\*</sup> is formed with a free energy uphill of 0.44 eV, while, on FeN<sub>4</sub>@G, NO<sub>2</sub>H<sup>\*</sup> is formed with a free energy downhill of 0.08 eV. Following this, they are



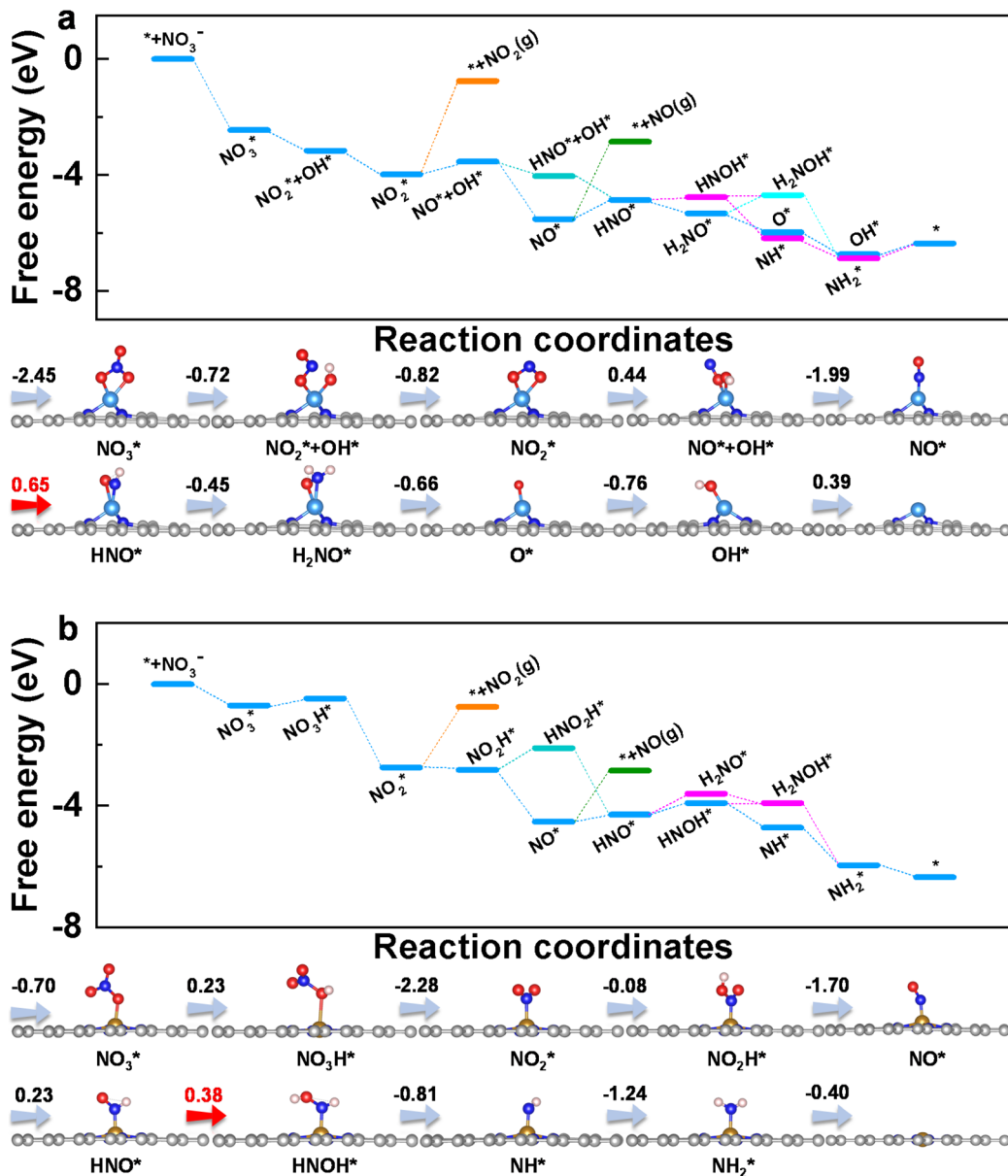


Fig. 4 Free energy diagrams and corresponding intermediates of eNO<sub>3</sub>RR on CoN<sub>3</sub>@G (a) and FeN<sub>4</sub>@G (b). The pathways to release of NO<sub>2</sub> and NO are also plotted for comparison.

attacked by the proton–electron pair to release the H<sub>2</sub>O molecule and leave NO\* on both CoN<sub>3</sub> and FeN<sub>4</sub>@G, with an energy change of -1.99 and -1.70 eV, respectively. On both systems, NO\* prefers the end-on adsorption pattern, and is hydrogenated to form HNO\* species. For CoN<sub>3</sub>@G, the hydrogenation of NO\* needs an energy input of 0.65 eV, and the following hydrogenation of HNO\* prefers to undergo the consecutive pathway to release the NH<sub>3</sub> molecule through HNO\* → H<sub>2</sub>NO\* → O\* + NH<sub>3</sub> with the energy release of 1.11 eV. In addition, after the formation of HNO\*, all the electrochemical steps are endothermic along the preferred pathways. Therefore, NO\* → HNO\* is the PDS and eNO<sub>3</sub>RR on CoN<sub>3</sub>@G possesses a limiting potential of -0.65 V. In contrast, as shown in Fig. 4b, NO\*

is hydrogenated to HNO\* with a free energy change of 0.23 eV, and then HNO\* is preferably hydrogenated to HNOH\* with a free energy uphill of 0.38 eV. Due to this all the remaining elemental reaction steps are exothermic along the preferred pathways, HNO\* → HNOH\* is the PDS with a low limiting potential of -0.38 V, which agrees well with the experimental findings.<sup>25</sup>

Above, we have demonstrated that the HER can be well suppressed on CoN<sub>3</sub> and FeN<sub>4</sub>@G. In addition to the HER, we further examined the possible formation of other byproducts along the eNO<sub>3</sub>RR reaction pathways. As shown in Fig. 4a, both NO and NO<sub>2</sub> are tightly bonded on CoN<sub>3</sub>@G with the desorption energy barriers of 2.68 and 3.22 eV, respectively, indi-

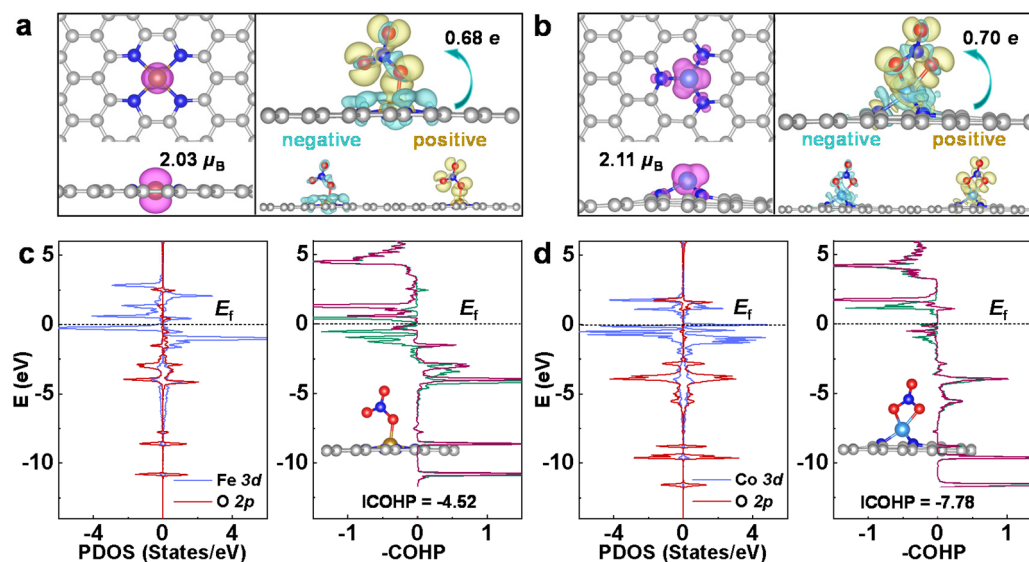
cating the significantly difficult formation of these byproducts. In contrast, the hydrogenation of  $\text{NO}_2^*$  and  $\text{NO}^*$  requires much less energy, further confirming that both species will be further hydrogenated rather than desorbed. Similarly, for  $\text{FeN}_4@G$  (Fig. 4b), the desorption of  $\text{NO}$  and  $\text{NO}_2$  is also very difficult, as indicated by the large desorption energy barriers of 1.69 and 1.99 eV, respectively, while their hydrogenation is even exothermic ( $-0.08$  eV for  $\text{NO}_2^* \rightarrow \text{NO}_2\text{H}^*$  and  $0.23$  eV for  $\text{NO}^* \rightarrow \text{HNO}^*$ ). Considering that  $\text{NO}^*$  is bonded with the metal center *via* the N atom only, the  $\text{NO}$  dimer may be formed through a couple of N–Co (Fe) bonds, which could lead to the production of  $\text{N}_2$ .<sup>29–33</sup> However, due to that  $\text{NO}$  is strongly adsorbed on  $\text{CoN}_3$  and  $\text{FeN}_4@G$ , the formation of the  $\text{NO}$  dimer will need to overcome a large energy barrier. In fact, in ref. 25 the Faraday efficiencies of  $\text{N}_2$  and  $\text{H}_2$  are less than 1%, and thus our theoretical studies on  $\text{eNO}_3\text{RR}$  selectivity provide a good explanation of the experimental findings. In addition, the desorption of  $\text{NH}_3$  is exothermic ( $-0.4$  eV) on  $\text{FeN}_4@G$  and slightly endothermic by  $0.4$  eV on  $\text{CoN}_3@G$ , indicating that both systems can be easily recovered for the  $\text{NO}_3^-$  to  $\text{NH}_3$  electroreduction cycle. Finally, the solvation effect has been investigated for  $\text{eNO}_3\text{RR}$  on  $\text{FeN}_4@G$ , simulated by the implicit solvation model implemented in VASPsol.<sup>49</sup> The free energy diagram is presented in Fig. S15.† It can be seen that the reaction pathway and PDS are the same as in the case without solvation correction, and  $U_L$  is only lowered by  $0.03$  V.

### 3.3. Activity origin of $\text{eNO}_3\text{RR}$

In the following, the  $\text{eNO}_3\text{RR}$  activity origin of  $\text{CoN}_3$  and  $\text{FeN}_4@G$  is discussed based on the electronic structure analysis, and then the key factors affecting the  $\text{eNO}_3\text{RR}$  activity trend of graphene-based SACs was revealed. The adsorption

and activation of  $\text{NO}_3^-$  is the prerequisite for its further hydrogenation, and the moderate binding strength is crucial for achieving high  $\text{eNO}_3\text{RR}$  activity, according to the Sabatier principle. As shown in Fig. 5a and b,  $\text{FeN}_4$  and  $\text{CoN}_3@G$  possess considerable spin magnetic moment localized on the embedded Fe ( $2.03\mu_B$ ) and Co ( $2.11\mu_B$ ) atoms, respectively. This localized spin magnetic moment of SACs and its associated spin-unpaired electrons could play an important role for the adsorption and activation of small molecules, as suggested by previous studies.<sup>55,56</sup> Accordingly, as shown in Table S4,†  $\text{NiN}_4@G$  that possesses a nonmagnetic ground state is unable to effectively bind  $\text{NO}_3^-$  with a binding free energy of  $0.20$  eV. The charge redistributions owing to  $\text{NO}_3^-$  adsorption for  $\text{FeN}_4$  and  $\text{CoN}_3@G$  are also displayed in Fig. 5a and b, respectively, where the yellow and cyan regions denote electron accumulation and depletion, respectively. On the whole, electrons are transferred from the catalysts to  $\text{NO}_3^*$ , in line with the Bader analysis,<sup>52</sup> which shows that  $\text{NO}_3^*$  obtains  $\sim 0.7e$  electrons for both systems. In detail, the electrons (yellow region) obtained by  $\text{NO}_3^*$  mainly occupy its antibonding orbitals, while besides the supports,  $\text{NO}_3^*$  also loses some charges (cyan region) from its bonding orbitals. Therefore,  $\text{NO}_3^*$  is effectively activated for the further hydrogenation reaction *via* the charge “acceptance–donation” mechanism, similar to the case of  $\text{N}_2$  activation in the NRR.<sup>57</sup>

Furthermore, the orbital interaction between  $\text{NO}_3^*$  and its bonded Fe or Co was investigated. The projected DOS for the  $\text{FeN}_4$  and  $\text{CoN}_3@G$  are presented in Fig. 5c and d, respectively, and the others in Fig. S16.† We can see that, in a large energy range, 3d states of the Fe or Co effectively interact with 2p states of O of  $\text{NO}_3^*$ . Furthermore, interestingly, crystal orbital Hamilton populations (COHP)<sup>58</sup> calculations show that for



**Fig. 5** (a) Spin density distribution of the pristine  $\text{FeN}_4@G$  and charge density difference of the  $\text{FeN}_4@G$  with the adsorbed  $\text{NO}_3^-$ . (b) is similar to (a), except that it is for  $\text{CoN}_3@G$ . (c) Projected densities of states (PDOS) of the Fe 3d states and O 2p states, which is bonded with Fe, and the COHP for the interaction between  $\text{NO}_3^-$  and Fe. (d) is similar to (c), except that it is for the  $\text{NO}_3^-$  adsorbed  $\text{CoN}_3@G$ . In (a) and (b), the yellow and cyan regions denote electron accumulation and depletion, respectively, and the isosurface value is set to  $0.002$  e per bohr<sup>3</sup>. In (c) and (d), Fermi level ( $E_f$ ) is set to  $0$  eV.

FeN<sub>4</sub>@G there are significant d-p antibonding states occupied below the Fermi level, while for CoN<sub>3</sub>@G the occupation of the d-p antibonding states is much less, which gives integrated COHP (ICOHP) of  $-4.52$  and  $-7.78$ , respectively, for FeN<sub>4</sub> and CoN<sub>3</sub>@G. The less negative ICOHP means weaker chemical bonding interaction, and thus the COHP analysis verifies the calculated binding strength trend of NO<sub>3</sub><sup>-</sup> on FeN<sub>4</sub> and CoN<sub>3</sub>@G. As discussed in previous works<sup>29–33</sup> and in the following herein, the moderate binding of NO<sub>3</sub><sup>-</sup> is crucial for the high eNO<sub>3</sub>RR activity. Thus, the occupied antibonding states near the Fermi level could play an important role in mediating the binding strength of NO<sub>3</sub><sup>\*</sup> on FeN<sub>4</sub>@G appropriately to achieve high eNO<sub>3</sub>RR activity.

According to the Sabatier principle, either too strong or too weak adsorption of the reaction intermediates goes against the whole catalytic process. This empirical principle has been turned into a quantitative predictive tool with the improvement of the accuracy of DFT calculations.<sup>59</sup> Currently, the binding free energy of the key intermediate has been often taken as the descriptor for various reactions to predict the activity trend and guide the catalyst design.<sup>60–62</sup> As for eNO<sub>3</sub>RR, previous studies indicate that  $\Delta G(\text{NO}_3^*)$  can be adopted as an effective descriptor.<sup>29–33</sup> Herein, as shown in Fig. 6a, although exhibiting different activity trend, the  $U_L$  values of both TMN<sub>3</sub>@G and TMN<sub>4</sub>@G are well correlated with  $\Delta G(\text{NO}_3^*)$ . For TMN<sub>3</sub>@G, CoN<sub>3</sub>@G locates at the peak of the activity volcano, while for TMN<sub>4</sub>@G, FeN<sub>4</sub>@G at the peak of the

activity volcano, considering that Co<sub>4</sub> and Ni<sub>4</sub>@G should have low catalytic activity due to the considerably low chemical activity to capture the key intermediate as discussed above. In addition, from Fig. 3, we note that too strong adsorption of NO<sub>3</sub><sup>-</sup> means that the active sites possess too high chemical activity, which tend to overbind OH\*, leading to OH\* → \* + H<sub>2</sub>O as the PDS and high  $U_L$  (including TiN<sub>3</sub>, VN<sub>3</sub>, CrN<sub>3</sub>, MnN<sub>3</sub>, FeN<sub>3</sub>, TiN<sub>4</sub>, and VN<sub>4</sub>@G). On the other hand, relatively weak adsorption of NO<sub>3</sub><sup>-</sup> on the catalysts leads to NO\* → HNO\* or HNO\* → HNOH\* as the PDS with lower  $U_L$  (such as CoN<sub>3</sub>, CrN<sub>4</sub>, MnN<sub>4</sub>, and FeN<sub>4</sub>@G).

As  $\Delta G(\text{NO}_3^*)$  can serve as a good activity descriptor, it is necessary to identify the properties of the catalysts that determine it. It has been well-established in surface science and heterogeneous catalysis that the binding strength of an adsorbate on a metal surface correlates with the d-band center of the atoms to which the adsorbate binds.<sup>63,64</sup> This conclusion also holds for the present case. As shown in Fig. 6b, there are good linear relationships between the d-band center of the embedded TM (Fig. S17†) and the binding free energies of NO<sub>3</sub><sup>-</sup> with different trends for TMN<sub>3</sub>@G and TMN<sub>4</sub>@G. It is noted that the binding energy of the adsorbate is hard to measure and control, while the d-band center can be accessible experimentally by using synchrotron-based high-resolution photoemission spectroscopy and tuned by various means.<sup>65,66</sup> However, the descriptor that only includes the intrinsic physical and chemical properties without resorting to theoretical calcu-

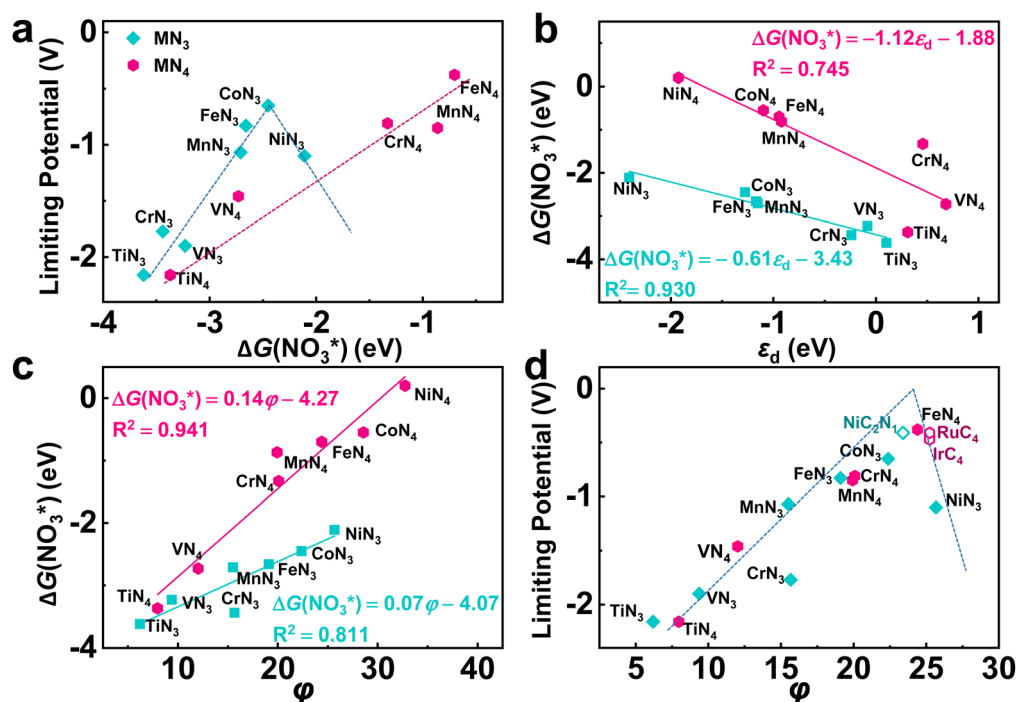


Fig. 6 (a) Relationship between the binding free energy of NO<sub>3</sub><sup>-</sup> ( $\Delta G(\text{NO}_3^*)$ ) and the limiting potential for the considered TMN<sub>3</sub>@G and TMN<sub>4</sub>@G systems. (b) Relationship between the d-band center ( $\epsilon_d$ ) of the embedded TM atom and  $\Delta G(\text{NO}_3^*)$ . (c) Relationship between the  $\phi$  values of the catalysts and  $\Delta G(\text{NO}_3^*)$ . (d) Activity volcano plot for the eNO<sub>3</sub>RR on the considered graphene-based SACs with  $\phi$  as the activity descriptor. The dashed lines in (a) and (d) is used to guide the eyes.

lations or experimental measurements is more preferable. So, we consider the descriptor,  $\varphi$ , proposed by Xu *et al.*,<sup>67</sup> which can be verified by the d-band center model and defined as

$$\varphi = \theta_d \times \frac{E_M + \alpha \times (n_N \times E_N + n_C \times E_C)}{E_O}$$

where  $E_M$ ,  $E_N$ , and  $E_C$  denote the electronegativity of metal, nitrogen, and carbon elements, respectively;  $n_C$  and  $n_N$  represent the number of C and N atoms coordinated with the TM atoms;  $\theta_d$  is the valence electrons in the occupied d orbital of the TM element derived from the periodic table of elements;  $E_O$  is the electronegativity of oxygen element; and  $\alpha$  is taken as 1. Interestingly, as shown in Fig. 6c,  $\Delta G(\text{NO}_3^*)$  for  $\text{TMN}_3@G$  and  $\text{TMN}_4@G$  are correlated linearly with  $\varphi$  having  $R^2$  values of 0.811 and 0.941, respectively. The strong linear correlation suggests that  $\varphi$  can function as an efficient descriptor to predict the eNO<sub>3</sub>RR catalytic activity. As expected, as shown in Fig. 6d, the  $U_L$  of eNO<sub>3</sub>RR indeed exhibits a volcano relationship with  $\varphi$ , and importantly  $\varphi$  is a unified descriptor for both  $\text{TMN}_3@G$  and  $\text{TMN}_4@G$  systems. Moreover, compared with the commonly adopted energy- and orbital-based descriptor,  $\varphi$  relates to the basic properties of catalysts, including the coordination number, and the electronegativity of the TM active center and its coordinated atoms. Thus, it is convenient to use  $\varphi$  to predict the eNO<sub>3</sub>RR activity of the graphene-based SACs, compared with  $\Delta G(\text{NO}_3^*)$  or the d-band center.

Taking  $\varphi$  as a guide, we further studied the eNO<sub>3</sub>RR activity of graphene-based SACs in other coordination environments, including  $\text{TMC}_{3-x}\text{N}_x@G$  ( $x = 0-2$ ) and  $\text{TMC}_{4-x}\text{N}_x$  ( $x = 0-3$ ). In particular,  $\text{NiC}_2\text{N}_1$ ,  $\text{RuC}_4$ , and  $\text{IrC}_4@G$ , possessing  $\varphi$  values of 23.4, 25.23, and 25.23 respectively, locate near the high-activity region. Therefore, the eNO<sub>3</sub>RR reaction pathways for these three systems were further explored in detail. Interestingly, our

calculated results show that  $\text{NiC}_2\text{N}_1$ ,  $\text{RuC}_4$ , and  $\text{IrC}_4@G$  indeed follow the volcano relationship presented in Fig. 6d, with  $U_L$  values of  $-0.41$ ,  $-0.41$ , and  $-0.47$  V respectively. The free energy diagram and atomic configuration of intermediates are presented in Fig. 7 for  $\text{NiC}_2\text{N}_1@G$ , and Fig. S18 and S19† for  $\text{RuC}_4$  and  $\text{IrC}_4@G$ , respectively. From Fig. 7,  $\text{NiC}_2\text{N}_1@G$  not only exhibits high intrinsic eNO<sub>3</sub>RR activity but also excellent selectivity against other products, and thus is a promising candidate with low cost for efficient electro conversion of  $\text{NO}_3^-$  to  $\text{NH}_3$ . Considering that  $\text{NiC}_2\text{N}_1@G$  has been synthesized,<sup>68</sup> this descriptor guided eNO<sub>3</sub>RR catalyst design strategy could be experimentally verified in the future.

In addition, small amounts of  $\text{NO}_2^-$  may coexist with  $\text{NO}_3^-$  in the solution.<sup>69</sup> Therefore, we studied the binding strengths of  $\text{NO}_2^-$  ( $\Delta G(\text{NO}_2^*)$ ) for the systems of  $\text{CoN}_3$ ,  $\text{FeN}_4$ ,  $\text{NiC}_2\text{N}_1$ ,  $\text{RuC}_4$ , and  $\text{IrC}_4@G$ , which exhibit relatively high eNO<sub>3</sub>RR activity. The values of  $\Delta G(\text{NO}_3^*)$  for  $\text{CoN}_3$ ,  $\text{FeN}_4$ ,  $\text{NiC}_2\text{N}_1$ ,  $\text{RuC}_4$ , and  $\text{IrC}_4@G$  are  $-2.45$ ,  $-0.70$ ,  $-1.86$ ,  $-1.85$ , and  $-1.73$  eV, respectively. The values of  $\Delta G(\text{NO}_2^*)$  for  $\text{CoN}_3$ ,  $\text{FeN}_4$ ,  $\text{NiC}_2\text{N}_1$ ,  $\text{RuC}_4$ , and  $\text{IrC}_4@G$  are  $-2.38$ ,  $-1.14$ ,  $-1.60$ ,  $-1.64$ , and  $-1.64$  eV, respectively. We can see for  $\text{FeN}_4@G$  that  $\text{NO}_2^-$  binds more strongly than  $\text{NO}_3^-$ . Therefore, there may be competition between the electrocatalytic reduction of  $\text{NO}_2^-$  and  $\text{NO}_3^-$ . However, as shown in Fig. 4b,  $\text{NO}_2^*$  is the intermediate of eNO<sub>3</sub>RR. In particular, the PDS for eNO<sub>3</sub>RR on  $\text{FeN}_4@G$  is the hydrogenation of  $\text{HNO}^*$ , which means that the reduction of  $\text{NO}_2^-$  follows the same paths as eNO<sub>3</sub>RR. Therefore, if there is  $\text{NO}_2^-$  in the solution,  $\text{NO}_2^-$  will be reduced first to produce  $\text{NH}_3$ , and then  $\text{NO}_3^-$  will be reduced to produce  $\text{NH}_3$ . In other words,  $\text{FeN}_4@G$  is a bifunctional electrocatalyst for reducing both  $\text{NO}_2^-$  and  $\text{NO}_3^-$  to yield  $\text{NH}_3$ .

Finally, it is noted that Ling *et al.* have performed a systematic theoretical study on the NRR on the graphene-based

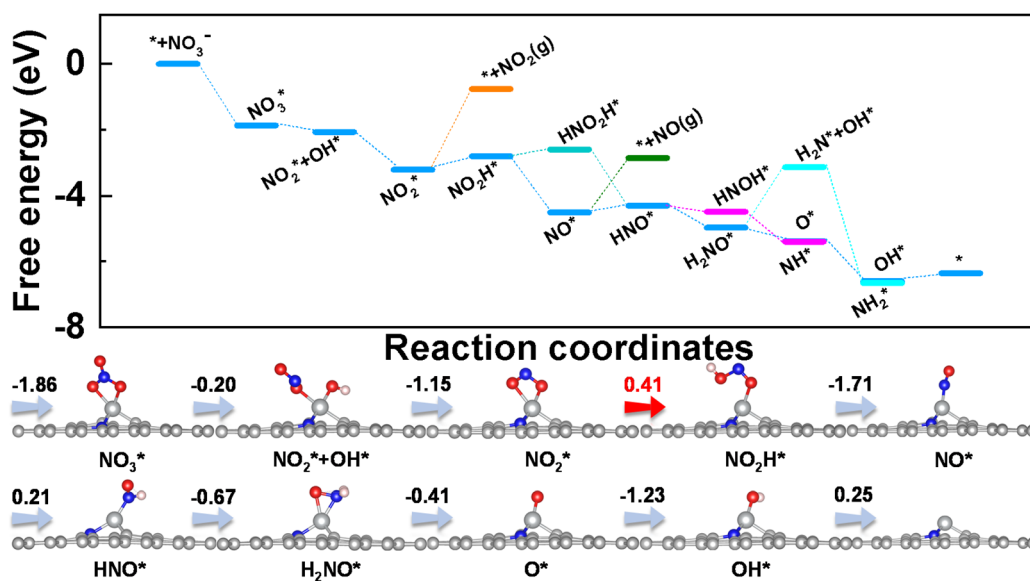


Fig. 7 Free energy diagrams and corresponding intermediates of eNO<sub>3</sub>RR on  $\text{NiC}_2\text{N}_1@G$ . The pathways to release of  $\text{NO}_2$  and  $\text{NO}$  are also plotted for comparison.



SACs with the computational settings consistent with ours.<sup>53</sup> Comparing this work with ours, the main differences between the NRR and eNO<sub>3</sub>RR on the graphene-based SACs can be summarized as follow. From ref. 53, we can see that for the graphene-based SACs the active sites for the NRR are early transition metals with less d electrons and high activity, including Nb, W, Re, V, Zr, and Mo, except that there is one system with Ru as the active site. This is in sharp contrast with the situation of eNO<sub>3</sub>RR. From our theoretical results, the active centers of the SACs with high NO<sub>3</sub>RR activity are late transition metals with more d electrons and low chemical activity, such as Fe, Ni, Ru, and Ir. The reason for this difference is that the N<sub>2</sub> molecule is extremely inert and hard to be activated, while NO<sub>3</sub><sup>-</sup> can be easily activated, due to the much larger bond energy of N<sub>2</sub> compared with NO<sub>3</sub><sup>-</sup>.

## 4. Conclusion

In conclusion, eNO<sub>3</sub>RR on TMN<sub>3</sub>@G and TMN<sub>4</sub>@G SACs has been systematically studied by means of first-principles calculations. The calculated results show that FeN<sub>4</sub>@G exhibits high eNO<sub>3</sub>RR activity ( $U_L = -0.38$  V) and excellent NH<sub>3</sub> selectivity to suppress the formation of byproducts, including other nitrogen-containing molecules and H<sub>2</sub>. The activity origin can be traced back to the effective adsorption and activation of NO<sub>3</sub><sup>-</sup> via the charge “acceptance–donation” mechanism and its moderate binding due to the occupation of the d–p antibonding orbital. These theoretical results provide a deep understanding of the excellent eNO<sub>3</sub>RR performance of FeN<sub>4</sub>@G demonstrated in the experiments.<sup>25</sup> More importantly, it is found that the eNO<sub>3</sub>RR activities of TMN<sub>3</sub>@G and TMN<sub>4</sub>@G SACs are well correlated with their intrinsic properties. The established activity descriptor,  $\varphi$ , including the coordination number, and the electronegativity of the TM center and its coordinated atoms, can be conveniently used to guide the design of graphene-based SACs for efficient eNO<sub>3</sub>RR. Finally, taking  $\varphi$  as a guide, several other graphene-based SACs were efficiently screened out with excellent eNO<sub>3</sub>RR performance. Our studies provide an atomic understanding of the catalytic mechanism and activity origin of the eNO<sub>3</sub>RR on graphene-based SACs, and further highlight the importance of the metal center and its local coordination environment for tuning the eNO<sub>3</sub>RR performance of SACs.

## Conflicts of interest

The authors declare no conflict of interest.

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