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Inexpensive polyphenylene network polymers with  
enhanced microporosity†

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Greatly enhanced microporosity is obtained for the amorphous porous polymers produced from the  $\text{AlCl}_3$ -mediated coupling of aromatic hydrocarbons by using dichloromethane as the reaction solvent. A polymer of average BET surface area =  $2435 \text{ m}^2 \text{ g}^{-1}$  was obtained reproducibly from 1,3,5-triphenylbenzene with the additional porosity being provided as ultramicroporosity as demonstrated by very high  $\text{CO}_2$  adsorption at 273 K/1 bar.

The potential for applications in gas storage, carbon capture, adsorption, heterogeneous catalysis and molecular separations has inspired intensive research activity focused on the synthesis of microporous materials using molecular building units.<sup>1</sup> From a structural perspective, there is an important distinction between crystalline porous materials derived from molecular precursors, such as the much-studied Metal Organic Frameworks (MOFs), and amorphous materials that possess a broader range of pore size. However, amorphous porous materials are attractive because they can be prepared using a greater diversity of chemistry and may offer enhanced stability as they do not require crystallisation *via* reversible bond formation.<sup>2</sup> With a few notable exceptions,<sup>3</sup> the generation of microporosity within amorphous polymers relies on the formation of a rigid and highly cross-linked network of covalent bonds. Among the rapidly growing number of reported porous network polymers, the Porous Aromatic Frameworks (*e.g.* PAF-1) and the structurally related Porous Polymer Networks (*e.g.* PPN-4) are particularly impressive with apparent Brunauer–Emmett–Teller (BET) surface areas rivalling those of the most porous MOFs ( $\text{SA}_{\text{BET}} > 5000 \text{ m}^2 \text{ g}^{-1}$ ).<sup>4</sup> PAFs and PPNs are prepared using the highly efficient Yamamoto aryl–aryl coupling reaction, which relies on the stoichiometric mediation of the expensive bis(1,5-cyclooctadiene)nickel(0) reagent. In addition, network polymer synthesis *via* Yamamoto coupling requires bromine-containing

monomers which are either expensive to purchase or are prepared using a multi-step synthesis, such as tetra-(4-bromophenyl) methane.<sup>5</sup> However, for many of the proposed applications of porous polymers, such as gas storage, water purification or carbon capture, low-cost and large-scale manufacture are necessary to provide competition with conventional inexpensive microporous materials such as activated carbons.<sup>1,6</sup> Therefore, network forming reactions that involve cheap monomers and reagents such as the Friedel–Craft reaction<sup>7</sup> or direct oxidative aryl–aryl coupling<sup>8</sup> (*i.e.* the Scholl reaction<sup>9</sup>) are of interest. Unfortunately, to date, these reactions have failed to provide highly porous network polymers. Here we report a simple modification to an aromatic coupling polymerisation that provides highly microporous polymers from readily available and inexpensive starting materials and reagent.

As part of our on-going research programme on triptycene-based porous polymers,<sup>10</sup> we performed an acylation of triptycene, following a literature procedure using acetyl chloride and  $\text{AlCl}_3$  in DCM.<sup>11</sup> An insoluble material was isolated from the reaction that proved to be a network polymer with a surprisingly high surface area ( $\text{SA}_{\text{BET}} > 1000 \text{ m}^2 \text{ g}^{-1}$ ). Optimisation studies, without acetyl chloride, gave polymer with  $\text{SA}_{\text{BET}} = 1750 \text{ m}^2 \text{ g}^{-1}$ , which is a significantly higher value than previously obtained from polymers derived from triptycene using the Friedel–Craft reaction ( $\text{SA}_{\text{BET}} = 1250\text{--}1430 \text{ m}^2 \text{ g}^{-1}$ )<sup>6c,7d</sup> and is comparable to results obtained using Yamamoto coupling of tribromo- or triiodo-triptycenes ( $\text{SA}_{\text{BET}} = 1300\text{--}1990 \text{ m}^2 \text{ g}^{-1}$ ).<sup>12</sup>

This result encouraged the polymerisation of a range of readily available monomers, from which porous polymers had previously been prepared, using these reaction conditions (*i.e.*  $\text{AlCl}_3$ , DCM at reflux). With the exception of tetraphenylmethane, highly porous polymers were obtained from each monomer (Table 1; Fig. 1). Indeed, for biphenyl, 1,3,5-triphenylbenzene (TPB), hexaphenylbenzene (HPB), spirobifluorene (SBF) and tetraphenylporphyrin (TPP) the  $\text{SA}_{\text{BET}}$  of the resulting polymer was greater than those reported for the polymers obtained previously from the same monomer using Friedel–Craft or Scholl reactions (ESI Table S1†). A particularly impressive result was obtained from the polymer derived from

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**Table 1** Properties of porous polymers prepared from given monomer using  $\text{AlCl}_3$  in refluxing DCM as solvent (alternatively  $\text{CHCl}_3$  or DCE, if stated)

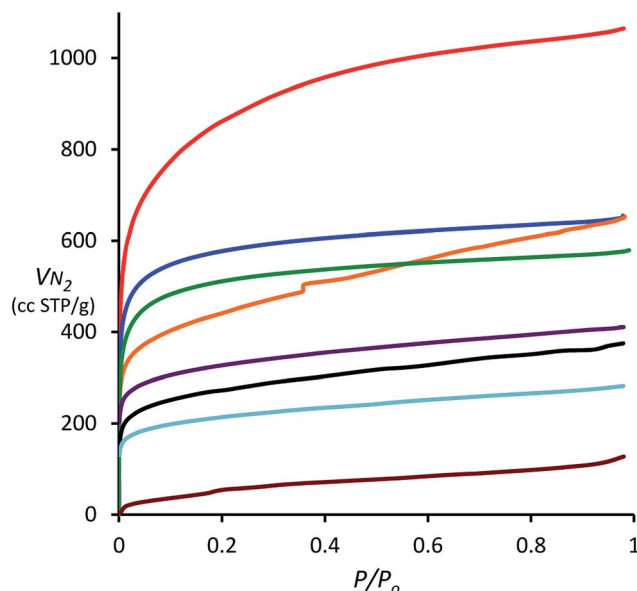
Monomer <sup>a</sup>	$\text{SA}_{\text{BET}}$ ( $\text{m}^2 \text{g}^{-1}$ )	$V_{\text{total}}^b$ ( $\text{cm}^3 \text{g}^{-1}$ )	$V_{\text{micro}}^c$ ( $\text{cm}^3 \text{g}^{-1}$ )	$\text{CO}_2$ uptake <sup>d</sup> ( $\text{mmol g}^{-1}$ )
Triptycene	1750	1.0	0.6	5.8 [3.1]
TPB	2435	1.6	1.0	5.9 [3.6]
TPB ( $\text{CHCl}_3$ )	1415	0.7	0.5	5.0 [3.2]
TPB (DCE)	725	0.7	0.2	1.7 [1.1]
Spirobifluorene	2035	1.0	0.7	5.8 [3.0]
HPB	1790	0.9	0.6	4.5 [2.7]
Biphenyl	1555	1.0	0.5	4.0 [2.7]
Biphenyl ( $\text{CHCl}_3$ )	800	0.5	0.2	3.3 [2.4]
Biphenyl (DCE)	453	0.3	0.1	1.7 [1.1]
Triphenylene	1180	0.7	0.4	4.0 [2.9]
DPA	1020	0.6	0.3	3.2 [2.0]
TPP	905	0.4	0.3	2.5 [2.0]
TPM	125	0.2	0.0	2.0 [1.3]

<sup>a</sup> TPB = 1,3,5-triphenylbenzene; HPB = hexaphenylbenzene; TPP = tetraphenyl porphyrin; DPA = 9,10-diphenylanthracene; TPM = tetraphenylmethane. <sup>b</sup> Pore volume estimated from  $\text{N}_2$  uptake at 77 K and  $P/P_0 = 0.98$ . <sup>c</sup> Micropore volume estimated from  $\text{N}_2$  uptake at 77 K and  $P/P_0 = 0.01$ . <sup>d</sup> At 1 bar/273 K [1 bar/298 K].

TPB that demonstrated reproducibly a  $\text{SA}_{\text{BET}}$  in the range 2328–2520  $\text{m}^2 \text{g}^{-1}$  greater than that of the equivalent porous polymer prepared from tri-brominated TPB using Yamamoto coupling ( $\text{SA}_{\text{BET}} = 1500\text{--}1870 \text{m}^2 \text{g}^{-1}$ ; ESI Table S1†).<sup>13</sup> This result allows access to an extremely porous polymer using cheap  $\text{AlCl}_3$  as reagent and from a monomer that is both inexpensive to buy and easily prepared on a large scale from acetophenone.<sup>14</sup> Initial scale-up attempts showed that a 10 g batch of the TPB polymer was readily obtained and demonstrated the same  $\text{SA}_{\text{BET}}$  as polymers from smaller (1 g) batches. The lack of success from tetraphenylmethane is likely due its insolubility in DCM combined with low reactivity due to relatively poorly activated phenyl groups.

The greatly enhanced porosity of the TPB and biphenyl network polymers – twice the  $\text{SA}_{\text{BET}}$  of those reported in previous studies of  $\text{AlCl}_3$ -mediated polymerisation – is surprising because the only difference is the use of DCM as solvent instead of  $\text{CHCl}_3$  (ESI Table S1†).<sup>8a,8c</sup> This prompted us to perform a direct comparison of the effect of DCM with commonly used  $\text{CHCl}_3$  or dichloroethane (DCE) solvent on the porosity of the resulting polymers from these two monomers (Table 1). It was found that the  $\text{SA}_{\text{BET}}$  values obtained from the TPB and biphenyl polymers prepared using  $\text{CHCl}_3$  as solvent were similar to those from previous studies.<sup>8a,8c</sup> Reactions using DCE as solvent produced polymer with only modest porosity. Therefore, it appears that the choice of DCM as solvent provides the dramatic enhancement of  $\text{SA}_{\text{BET}}$ . Consequently, it was important to determine whether DCM acts only as a solvent or also provides methylene cross-links *via* a Friedel–Craft reaction. The  $^{13}\text{C}$  solid-state NMR spectrum of all the polymers prepared using DCM shows a small peak between 30 and 40 ppm that may indicate the presence of methylene cross-links, originating from the DCM (ESI†). However, this peak is much smaller than that found previously from a network polymer prepared from a formal Friedel–Craft reaction between TPB and formyl

dimethyl acetal as the methylene source.<sup>7a</sup> In addition, the non-aromatic peaks observed for the TPB polymer are of a similar size, relative to that of the much larger peak of the aromatic carbons, to those previously observed from a TPB network obtained using  $\text{CHCl}_3$ .<sup>8a,8c</sup> Therefore, it appears that network formation is predominately based on direct aryl–aryl coupling *via* the Scholl reaction, however, it seems likely that the DCM may contribute some cross-links to the network but



**Fig. 1**  $\text{N}_2$  adsorption isotherms obtained at 77 K for porous polymers derived from 1,3,5-triphenylbenzene (TPB, red), spirobifluorene (blue), triptycene (green), biphenyl (orange), triphenylene (purple), 9,10-diphenylanthracene (DPA, black), tetraphenylporphyrin (TPP, light blue) and tetraphenylmethane (TPM, brown) using DCM and  $\text{AlCl}_3$ . Note that the isotherm of hexaphenylbenzene (HPB) is superimposable on that of triptycene and has been omitted for clarity (see ESI†).



only to a similar extent as  $\text{CHCl}_3$  or DCE. Therefore, the effect of DCM in enhancing  $\text{SA}_{\text{BET}}$  is predominately related to its role as a solvent ensuring a more efficient network formation and/or facilitating the generation of a more microporous structure by behaving as a better porogen.<sup>15</sup> Similarly large solvent effects on  $\text{SA}_{\text{BET}}$  have been reported for other types of reactions to make porous network polymers.<sup>7e,16</sup>

The  $\text{N}_2$  adsorption isotherms also provide information about the pore size distribution within the porous polymers. For the polymers derived from TPB, spirobifluorene and triptycene, the relatively large uptake of  $\text{N}_2$  at 77 K at low relative pressure ( $P/P_0 < 0.01$ ), demonstrates that micropores (<2 nm) contribute a large portion (>60%) of the total pore volume (Table 1). The absence of significant hysteresis between the adsorption and desorption isotherms (ESI†) for all polymers, with the exception of that from biphenyl, indicates both low degree of mesoporosity and that the pore structure appears fixed on adsorption of  $\text{N}_2$  with no significant swelling.

$\text{CO}_2$  adsorption isotherms obtained at 273 K (ESI and Table S1†) also provide information on pore size distribution. Such data are also useful as an indicator of the suitability of a polymer for use in carbon capture using pressure-swing adsorption.<sup>6a,b</sup>  $\text{CO}_2$  uptake for the triptycene, SBF and TPB network polymers are all approximately  $5.8 \text{ mmol g}^{-1}$  at 1 bar/273 K ( $3.0\text{--}3.6 \text{ mmol g}^{-1}$  at 1 bar/295 K), which is impressive in comparison with other polymers derived from these building units (ESI Table S1†) and are similar to the best performing network polymers ( $\sim 6.0 \text{ mmol g}^{-1}$  at 1 bar/273 K and  $3.6 \text{ mmol g}^{-1}$  at 1 bar/295 K).<sup>17</sup> The high uptake of  $\text{CO}_2$  suggests that the use of DCM provides additional porosity in the form of ultramicropores (*i.e.* pores of diameter less than 0.7 nm), which are the physisorption sites for  $\text{CO}_2$  at 273 K/1 bar. The porosity and gas adsorption of the polymer derived from TPB is strikingly similar to that of one of the best performing activated carbons for  $\text{CO}_2$  capture.<sup>18</sup>

Despite the large number of reported porous polymers, prepared using a wide variety of monomers and network-forming reactions,<sup>2</sup> only a few demonstrate  $\text{SA}_{\text{BET}}$  in excess of  $2000 \text{ m}^2 \text{ g}^{-1}$  and all of these require preparation from monomers that need a multistep synthesis or expensive precursors.<sup>4a,b,5,13b,17a,19</sup> This work now allows access to highly microporous network polymers, of surface areas in excess of those of commercial activated carbons, from readily obtained starting materials.

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## Notes and references

- 1 A. G. Slater and A. I. Cooper, *Science*, 2015, **348**, 988.
- 2 (a) N. B. McKeown and P. M. Budd, *Macromolecules*, 2010, **43**, 5163–5176; (b) R. Dawson, A. I. Cooper and D. J. Adams, *Prog. Polym. Sci.*, 2012, **37**, 530–563; (c) D. Wu, F. Xu, B. Sun, R. Fu,

- H. He and K. Matyjaszewski, *Chem. Rev.*, 2012, **112**, 3959–4015; (d) Y. Zhang and S. N. Riduan, *Chem. Soc. Rev.*, 2012, **41**, 2083–2094; (e) Z. Chang, D.-S. Zhang, Q. Chen and X.-H. Bu, *Phys. Chem. Chem. Phys.*, 2013, **15**, 5430–5442; (f) Z. Xiang and D. Cao, *J. Mater. Chem. A*, 2013, **1**, 2691–2718; (g) M. Rose, *ChemCatChem*, 2014, **6**, 1166–1182.
- 3 Polymers of intrinsic microporosity and some other high free volume polymers exhibit microporosity without a network structure – see N. B. McKeown and P. M. Budd, *Chem. Soc. Rev.*, 2006, **35**, 675–683.
- 4 (a) T. Ben, H. Ren, S. Q. Ma, D. P. Cao, J. H. Lan, X. F. Jing, W. C. Wang, J. Xu, F. Deng, J. M. Simmons, S. L. Qiu and G. S. Zhu, *Angew. Chem., Int. Ed.*, 2009, **48**, 9457–9460; (b) D. Yuan, W. Lu, D. Zhao and H.-C. Zhou, *Adv. Mater.*, 2011, **23**, 3723–3725; (c) C. Pei, T. Ben and S. Qiu, *Mater. Horiz.*, 2015, **2**, 11–21.
- 5 W. Lu, D. Yuan, D. Zhao, C. I. Schilling, O. Plietzsch, T. Muller, S. Braese, J. Guenther, J. Bluemel, R. Krishna, Z. Li and H.-C. Zhou, *Chem. Mater.*, 2010, **22**, 5964–5972.
- 6 (a) R. Dawson, E. Stoeckel, J. R. Holst, D. J. Adams and A. I. Cooper, *Energy Environ. Sci.*, 2011, **4**, 4239–4245; (b) C. F. Martin, E. Stockel, R. Clowes, D. J. Adams, A. I. Cooper, J. J. Pis, F. Rubiera and C. Pevida, *J. Mater. Chem.*, 2011, **21**, 5475–5483; (c) R. T. Woodward, L. A. Stevens, R. Dawson, M. Vijayaraghavan, T. Hasell, I. P. Silverwood, A. V. Ewing, T. Ratvijitvech, J. D. Exley, S. Y. Chong, F. Blanc, D. J. Adams, S. G. Kazarian, C. E. Snape, T. C. Drage and A. I. Cooper, *J. Am. Chem. Soc.*, 2014, **136**, 9028–9035; (d) H. A. Patel, F. Karadas, A. Canlier, J. Park, E. Deniz, Y. Jung, M. Atilhan and C. T. Yavuz, *J. Mater. Chem.*, 2012, **22**, 8431–8437.
- 7 (a) B. Li, R. Gong, W. Wang, X. Huang, W. Zhang, H. Li, C. Hu and B. Tan, *Macromolecules*, 2011, **44**, 2410–2414; (b) G. Liu, Y. Wang, C. Shen, Z. Ju and D. Yuan, *J. Mater. Chem. A*, 2015, **3**, 3051–3058; (c) K. Schute and M. Rose, *ChemSusChem*, 2015, **8**, 3419–3423; (d) C. Zhang, P.-C. Zhu, L. Tan, J.-M. Liu, B. Tan, X.-L. Yang and H.-B. Xu, *Macromolecules*, 2015, **48**, 8509–8514; (e) N. Fontanals, R. M. Marce, F. Borrull and P. A. G. Cormack, *Polym. Chem.*, 2015, **6**, 7231–7244.
- 8 (a) B. Li, Z. Guan, X. Yang, W. D. Wang, W. Wang, I. Hussain, K. Song, B. Tan and T. Li, *J. Mater. Chem. A*, 2014, **2**, 11930–11939; (b) L. Li, H. Ren, Y. Yuan, G. Yu and G. Zhu, *J. Mater. Chem. A*, 2014, **2**, 11091–11098; (c) L. Li, K. Cai, P. Wang, H. Ren and G. Zhu, *ACS Appl. Mater. Interfaces*, 2015, **7**, 201–208.
- 9 M. Grzybowski, K. Skonieczny, H. Butenschoen and D. T. Gryko, *Angew. Chem., Int. Ed.*, 2013, **52**, 9900–9930.
- 10 (a) B. S. Ghanem, M. Hashem, K. D. M. Harris, K. J. Msayib, M. Xu, P. M. Budd, N. Chaukura, D. Book, S. Tedds, A. Walton and N. B. McKeown, *Macromolecules*, 2010, **43**, 5287–5294; (b) M. Carta, M. Croad, R. Malpass-Evans, J. C. Jansen, P. Bernardo, G. Clarizia, K. Friess, M. Lanc and N. B. McKeown, *Adv. Mater.*, 2014, **26**, 3526–3531; (c) R. G. D. Taylor, C. G. Bezzu, M. Carta, K. J. Msayib, J. Walker, R. Short, B. M. Kariuki and N. B. McKeown, *Chem.–Eur. J.*, 2016, **22**, 2466–2472.



- 11 P.-F. Li and C.-F. Chen, *J. Org. Chem.*, 2012, **77**, 9250–9259.
- 12 C. Zhang, Y. Liu, B. Li, B. Tan, C.-F. Chen, H.-B. Xu and X.-L. Yang, *ACS Macro Lett.*, 2012, **1**, 190–193.
- 13 (a) H. Ren, T. Ben, F. Sun, M. Guo, X. Jing, H. Ma, K. Cai, S. Qiu and G. Zhu, *J. Mater. Chem.*, 2011, **21**, 10348–10353; (b) Z. Xiang, X. Zhou, C. Zhou, S. Zhong, X. He, C. Qin and D. Cao, *J. Mater. Chem.*, 2012, **22**, 22663–22669.
- 14 (a) Y. Zhao, J. Li, C. Li, K. Yin, D. Ye and X. Jia, *Green Chem.*, 2010, **12**, 1370–1372; (b) S. Zhao, L. Kang, H. Ge, F. Yang, C. Wang, C. Li, Q. Wang and M. Zhao, *Synth. Commun.*, 2012, **42**, 3569–3578; (c) D. Prasad, A. Preetam and M. Nath, *C. R. Chim.*, 2013, **16**, 252–256; (d) H. R. Safaei, M. Davoodi and M. Shekouhy, *Synth. Commun.*, 2013, **43**, 2178–2190.
- 15 D. C. Sherrington, *Chem. Commun.*, 1998, 2275–2286.
- 16 R. Dawson, A. Laybourn, Y. Z. Khimyak, D. J. Adams and A. I. Cooper, *Macromolecules*, 2010, **43**, 8524–8530.
- 17 (a) S. Hug, L. Stegbauer, H. Oh, M. Hirscher and B. V. Lotsch, *Chem. Mater.*, 2015, **27**, 8001–8010; (b) Y. Zhu, H. Long and W. Zhang, *Chem. Mater.*, 2013, **25**, 1630–1635.
- 18 M. Sevilla and A. B. Fuertes, *Energy Environ. Sci.*, 2011, **4**, 1765–1771.
- 19 (a) J. H. Ahn, J. E. Jang, C. G. Oh, S. K. Ihm, J. Cortez and D. C. Sherrington, *Macromolecules*, 2006, **39**, 627–632; (b) P. Kuhn, M. Antonietti and A. Thomas, *Angew. Chem., Int. Ed.*, 2008, **47**, 3450–3453; (c) E. Preis, C. Widling, G. Bruncklaus, J. Schmidt, A. Thomas and U. Scherf, *ACS Macro Lett.*, 2013, **2**, 380–383; (d) Z. Xiang, R. Mercado, J. M. Huck, H. Wang, Z. Guo, W. Wang, D. Cao, M. Haranczyk and B. Smit, *J. Am. Chem. Soc.*, 2015, **137**, 13301–13307.

