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# Synthesis, fluorescence property and cell imaging of a perylene diimide-based NIR fluorescent probe for hypochlorite with dual-emission fluorescence responses†

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In this study, for the first time, a novel near-infrared and ratiometric fluorescent probe was conveniently synthesized by reacting PDI-based salicylaldehyde with 2-(hydrazonomethyl)phenol. The probe was designed based on the intramolecular charge transfer (ICT) mechanism and the intramolecular electron transfer (IET) mechanism, which can be blocked by the hypochlorite and a larger conjugation is formed within the probe. Due to the turn-on fluorescence responses triggered by the PET and ICT processes after the ClO<sup>-</sup> addition, the probe produces simultaneous emission peaks at 600 nm and 820 nm. In addition, the probe shows a rapid fluorescence response towards the ClO<sup>-</sup> ions within 5 s, a low detection limit (0.8 × 10<sup>-7</sup> M), intense color changes (from purple to light green), excellent selectivity, and reversibility. Importantly, cell imaging experiments show that the probe can identify endogenous ClO<sup>-</sup> successfully.

## Introduction

Fluorescent chemosensors because of their fast detection speed, high sensitivity and simple operation have been widely used in biological tracing, ion detection, environmental protection and other fields. Today, fluorescent chemosensors used as detection tools for toxic ions or small molecules are getting increasing attention. In recent years, a large number of fluorescent sensors for the detection of ions or small molecules have been reported.<sup>1–5</sup> Among these ions or small molecules, hypochlorite (ClO<sup>-</sup>) as a reactive oxygen species plays an important role in our daily life. However, excessive hypochlorite easily causes a variety of diseases.<sup>6–8</sup> Therefore, it is necessary to develop fluorescent chemosensors for the real-time monitoring of hypochlorous acid/hypochlorite.

To date, numerous fluorescent probes for ClO<sup>-</sup> have been reported.<sup>9–25</sup> Most of these fluorescence probes for ClO<sup>-</sup> were a type of short wavelength response probes. Such probes have some prominent disadvantages, such as strong background effect, shallow imaging depth, and poor anti-infection ability. In order to overcome these defects, near-infrared ratiometric fluorescent probes were developed quickly.<sup>26–30</sup> Compared with

the common fluorescent probes, near-infrared ratiometric fluorescent probes are overwhelming because of their unique advantages: deeper imaging depth, stronger anti-interference energy, and less tissue damage.

It is known that perylene diimide (PDI) derivatives as excellent fluorophores have been widely studied, most of them are designed based on the PET mechanism, and almost none are designed based on both the PET and ICT mechanism.<sup>31–36</sup> In this study, a new near-infrared ratiometric fluorescence probe, namely **PDI-SY**, was successfully designed and easily obtained by a direct condensation reaction between the PDI-based salicylaldehyde and 2-(hydrazonomethyl)phenol (Fig. 1). **PDI-SY** produces simultaneous emission peaks at 600 nm and 820 nm, when ClO<sup>-</sup> is added. In addition, **PDI-SY** shows high sensitivity and selectivity in a fluorescence “turn-on” manner, the probe can efficiently monitor ClO<sup>-</sup> with rapid fluorescence responses within 5 s, with a low detection limit (0.8 × 10<sup>-7</sup> M) and intense color changes (from purple to light green). Further <sup>1</sup>H-NMR study also suggests that the PET and ICT mechanisms were involved after the addition of ClO<sup>-</sup> ions. Moreover, the cell imaging experiments show that the probe can identify endogenous ClO<sup>-</sup> successfully.

## Results and discussion

### Synthesis of the probe PDI-SY

The preparation of **PDI-SY** is shown in Scheme 1, the key intermediate compound **PDI-Y** was prepared *via* a simple electrophilic

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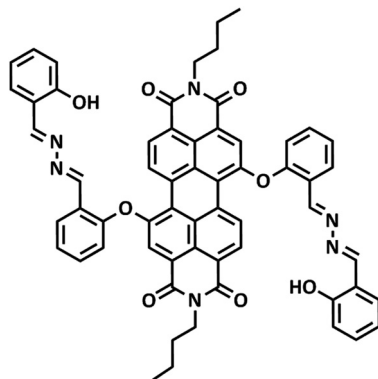


Fig. 1 Structure of the probe **PDI-SY**.

substitution reaction between **PDI-S** and salicylaldehyde, affording **PDI-Y** with 95% yield. All the intermediates and materials used in the synthesis process do not need to be further purified, the materials are cheap and readily available, which greatly reduce the synthesis cost of the probe **PDI-SY**. All the intermediates and the probe molecule **PDI-SY** were fully characterized *via*  $^1\text{H}$  NMR,  $^{13}\text{C}$  NMR, and high-resolution mass spectrometry (HRMS-MALDI-TOF); the detailed synthesis procedure of the **PDI-SY** has been discussed in the synthesis part.

#### Ratiometric absorption and dual emission of **PDI-SY** before and after reaction with $\text{ClO}^-$

The absorption and fluorescence emission properties of **PDI-SY** ( $2\ \mu\text{M}$ ) before and after the reaction with  $\text{ClO}^-$  were studied in an optimized mixture of the PBS buffer solution (pH 7.4). All the measurements were carried out in PBS buffer (pH 7.4). As shown in Fig. 2, **PDI-SY** exhibited maximum absorption peaks at around 354 nm and 556 nm without any other ions.

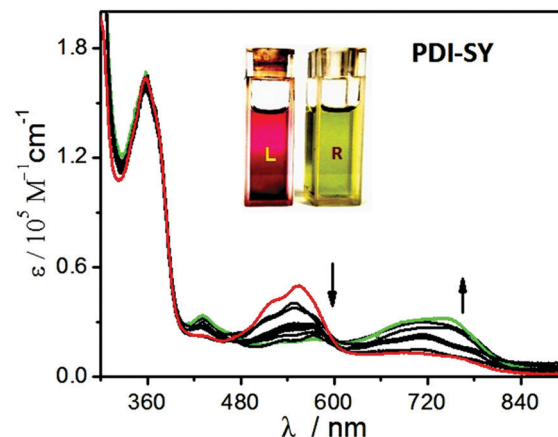
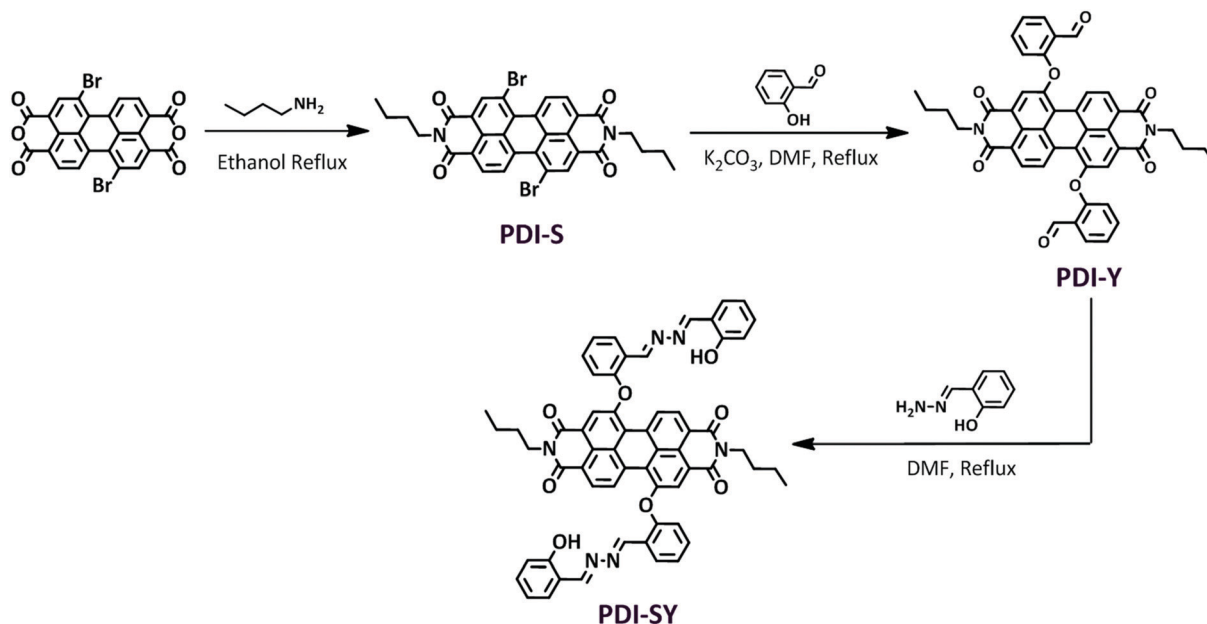


Fig. 2 Ratiometric absorption spectra of **PDI-SY** as a function of the  $\text{ClO}^-$  concentration. ( $[\text{PDI-SY}] = 2\ \mu\text{M}$ ,  $V_{\text{THF}/\text{H}_2\text{O}} = 1 : 3$ , pH = 7.4, PBS buffer). The inset picture shows the color change of **PDI-SY** upon the addition of  $\text{ClO}^-$ , L: **PDI-SY**, R: **PDI-SY** +  $\text{ClO}^-$ .

Moreover, **PDI-SY** showed a weak emission at 520 nm, as shown in Fig. 3. However, upon the addition of  $\text{ClO}^-$  ( $0\text{--}20\ \mu\text{M}$ ), a new absorption peak at around 720 nm increased dramatically, while the absorption signals at 556 nm decreased gradually. When  $\text{ClO}^-$  was added to 2 equivalents, the reaction basically completed, which was concomitant with a remarkable color change from purple to light green, as shown in Fig. 2 inset. Moreover, upon the addition of  $\text{ClO}^-$ , the emission peak at 520 nm decreased gradually, while a new sharp enhancement of the fluorescence emission at 600 nm ( $\Phi_{1\text{Free}} < 0.1$ ,  $\Phi_{1\text{Hypochlorite}} = 0.58$ ) and 820 nm ( $\Phi_{2\text{Free}} < 0.1$ ,  $\Phi_{2\text{Hypochlorite}} = 0.18$ ) was observed, as shown in Fig. 3. When  $\text{ClO}^-$  was added to 2 equivalents, the reaction was basically completed; the new emission peaks at 600 nm and 820 nm reached a maximum value. From the spectral property of **PDI-SY**, it could be



Scheme 1 Synthesis route of the probe **PDI-SY**.



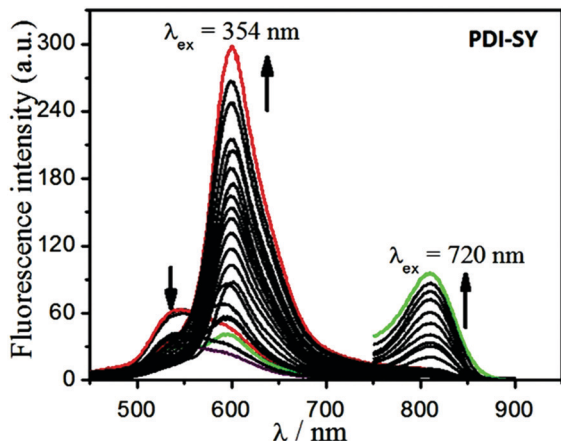


Fig. 3 Dual emission fluorescence spectra of **PDI-SY** as a function of the  $\text{ClO}^-$  concentration, ( $[\text{PDI-SY}] = 2 \mu\text{M}$ ,  $\lambda_{\text{ex}} = 354 \text{ nm}$  and  $720 \text{ nm}$ ,  $V_{\text{THF}/\text{H}_2\text{O}} = 1:3$ ,  $\text{pH} = 7.4$ ,  $\text{PBS}$  buffer).

confirmed that **PDI-SY** can be used as a ratiometric fluorescent probe for sensing  $\text{ClO}^-$ .

Moreover, a linear relationship between the concentrations of  $\text{ClO}^-$  and the fluorescence intensity values of **PDI-SY** at  $600 \text{ nm}$  was observed (as shown in Fig. 3), the linear fitting curve of **PDI-SY** at  $600 \text{ nm}$  is shown in Fig. 4. According to this fitted equation, the detection limit of **PDI-SY** was measured to be  $0.8 \times 10^{-7} \text{ M}$  ( $3\sigma/\text{slope}$ ,  $\sigma$  is the standard deviation of the blank measurement). The above studies showed that **PDI-SY** can accurately and sensitively detect  $\text{ClO}^-$  (Table 1).

#### Rapid time responses of **PDI-SY** for $\text{ClO}^-$ ions

The reaction rate of **PDI-SY** with the  $\text{ClO}^-$  ions is an important parameter to value its real-time tracking capability. As shown in Fig. 5, the reaction of **PDI-SY** with  $\text{ClO}^-$  can be triggered within  $5 \text{ s}$  without stirring, and completed in  $60 \text{ s}$  with a maximum emission intensity, indicating that **PDI-SY** can be employed as a rapid analytical chemosensor for the  $\text{ClO}^-$  detection. Simultaneously, the color of the solution containing **PDI-SY** changed from purple to light green gradually, as shown in Fig. 5 inset.



Fig. 4 Linear fitting curve of the probe **PDI-SY** with  $\text{ClO}^-$ , ( $[\text{PDI-SY}] = 2 \mu\text{M}$ ,  $[\text{ClO}^-] = 2$  equivalents,  $\lambda_{\text{em}} = 600 \text{ nm}$ ,  $V_{\text{THF}/\text{H}_2\text{O}} = 1:3$ ).

Table 1 Fluorescence properties of **PDI-SY** in  $V_{\text{THF}/\text{H}_2\text{O}} = 1/3$  at  $298 \text{ K}$

Compounds	$\lambda_{\text{max}}^a$ (nm)	$\lambda_{\text{em}}^b$ (nm)	$\Delta\nu_1/\Delta\nu_2^c$ ( $\text{cm}^{-1}$ )	$\Phi_1^d$ at $600 \text{ nm}$	$\Phi_2^d$ at $820 \text{ nm}$
<b>PDI-SY</b>	354/556	600	3947	<0.1	<0.1
<b>PDI-SY-ClO<sup>-</sup></b>	556/720	600/820	1692	0.58	0.18

<sup>a</sup> Absorption maximum. <sup>b</sup> Emission maximum. <sup>c</sup> Stoke's shift =  $(1/\lambda_{\text{abs}} - 1/\lambda_{\text{em}}) \times 10^7$ . <sup>d</sup> Fluorescence quantum yield, the fluorescence quantum yields were determined using rhodamine 6G and PDI ( $\Phi = 0.95$ ) as a standard, the max absorbance was used as the excited wavelength.<sup>37</sup>

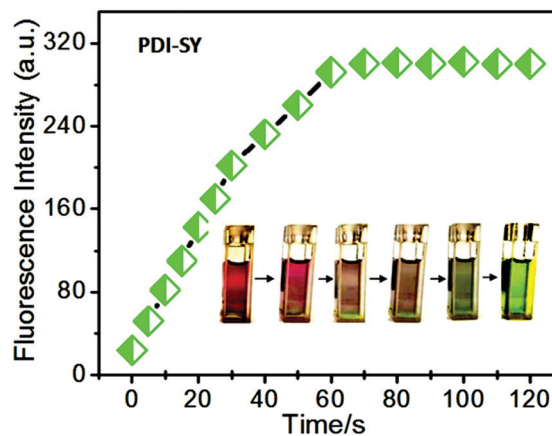


Fig. 5 Time-dependent fluorescence change of **PDI-SY** ( $2 \mu\text{M}$ ) treated with  $\text{ClO}^-$ .  $\lambda_{\text{ex}} = 354 \text{ nm}$ ,  $\lambda_{\text{em}} = 600 \text{ nm}$ ,  $\text{pH} = 7.4$ ,  $\text{PBS}$  buffer. (inset: photos of **PDI-SY** reacted with  $\text{ClO}^-$  ions in waste water from  $5$  seconds to  $120$  seconds).

The results demonstrate that **PDI-SY** has the ability for real-time monitoring of hypochlorite.

#### Selectivity and reversibility of **PDI-SY** to various ions

The interference experiments of the probe **PDI-SY** with other ions such as  $\text{NO}_2^-$ ,  $m\text{-Cl-ph-COOH}$ , glutathione, cysteine,  $\text{NO}_3^-$ ,  $\text{PO}_4^{3-}$ ,  $\text{H}_2\text{O}_2$ ,  $\text{Fe}^{3+}$ ,  $\text{Ag}^+$ ,  $\text{Ba}^{2+}$ ,  $\text{Co}^{2+}$ ,  $\text{Ni}^{2+}$ ,  $\text{Cu}^{2+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Cd}^{3+}$ ,  $\text{Pb}^{2+}$ ,  $\text{Mn}^{2+}$  and  $\text{Fe}^{2+}$  were taken into account. First, eighteen parts of the **PDI-SY** solution with the same concentration were prepared. Then, the other ions were added to the solution, respectively. The fluorescence spectra of **PDI-SY** at  $600 \text{ nm}$  were collected. At last,  $\text{ClO}^-$  ions were added to the mixed solutions of **PDI-SY** with the other ions. All the fluorescence spectra excited at  $354 \text{ nm}$  were collected. As shown in Fig. 6, none of the other ions caused any fluorescence emission, while an obvious fluorescence enhancement occurred only after the addition of  $\text{ClO}^-$ , indicating that **PDI-SY** can detect  $\text{ClO}^-$  specifically.

**PDI-SY** shows good reversibility to  $\text{ClO}^-$  (as oxidant) and  $\text{S}^{2-}$  (as a reducing agent); the reversibility experiment was carried out by recording the emission intensity of **PDI-SY** at  $600 \text{ nm}$  with respect to the change of the ions  $\text{ClO}^-/\text{S}^{2-}$  up to 5 cycles. Fig. 7 shows the fluorescence intensity change of **PDI-SY** for 5 cycles. The relative standard deviations from five replicates were estimated to be  $<2\%$ .



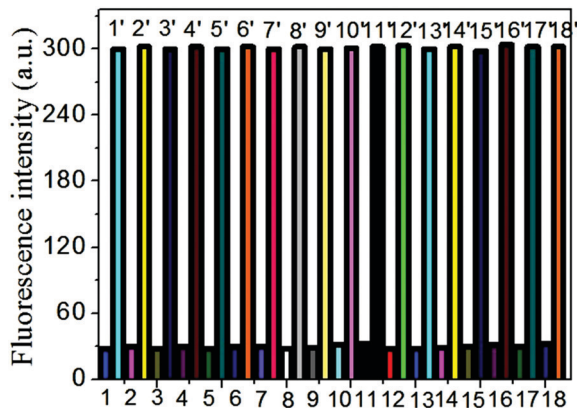


Fig. 6 Fluorescence spectra of the probe **PDI-SY** to numerous metal cations such as: (1–18)  $\text{NO}_2^-$ , *m*-Cl-ph-COOH, glutathione, cysteine,  $\text{NO}_3^-$ ,  $\text{PO}_4^{3-}$ ,  $\text{H}_2\text{O}_2$ ,  $\text{Fe}^{3+}$ ,  $\text{Ag}^+$ ,  $\text{Ba}^{2+}$ ,  $\text{Co}^{2+}$ ,  $\text{Ni}^{2+}$ ,  $\text{Cu}^{2+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Cd}^{3+}$ ,  $\text{Pb}^{2+}$ ,  $\text{Mn}^{2+}$  and  $\text{Fe}^{2+}$ , [**PDI-SY**] = 2  $\mu\text{M}$ ,  $\lambda_{\text{ex}}$  = 354 nm,  $\lambda_{\text{em}}$  = 600 nm,  $V_{\text{THF}/\text{H}_2\text{O}}$  = 1/3, pH = 7.4, PBS buffer.



Fig. 7 Fluorescence intensity of **PDI-SY** at 600 nm ([**PDI-SY**] = 2  $\mu\text{M}$ ,  $\lambda_{\text{ex}}$  = 354 nm, in  $V_{\text{THF}/\text{H}_2\text{O}}$  = 1:3, pH = 7.4, PBS buffer) upon addition of  $\text{ClO}^-$  (2 equiv.) and  $\text{S}^{2-}$  solution up to 5 cycles.

### PET and ICT-based mechanism discussion of **PDI-SY** with $\text{ClO}^-$ ions

The possible mechanism of the **PDI-SY** reaction with the  $\text{ClO}^-$  ions is shown in Fig. 8, a PET process from the dibenzylidene-hydrazine to the **PDI** core in the excited **PDI-SY** blocks the fluorescence emission of the **PDI** core. After the addition of the  $\text{ClO}^-$  ions, an oxidation reaction between **PDI-SY** and  $\text{ClO}^-$  ions occurred. The



Fig. 8 The possible mechanism based on the PET and ICT of **PDI-SY** for detecting  $\text{ClO}^-$  ions.

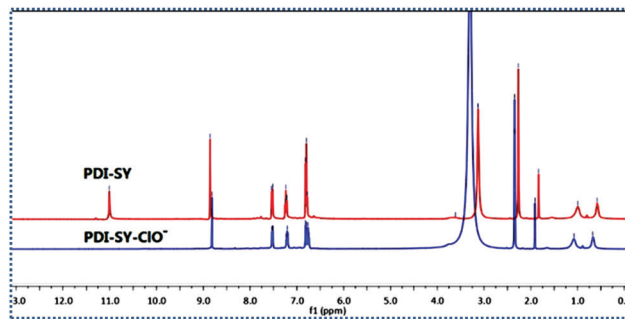


Fig. 9  $^1\text{H}$  NMR spectra of **PDI-SY** in the absence (top) and presence (bottom) of two equiv. of  $\text{ClO}^-$  in deuterated  $\text{DMSO-d}_6$ .

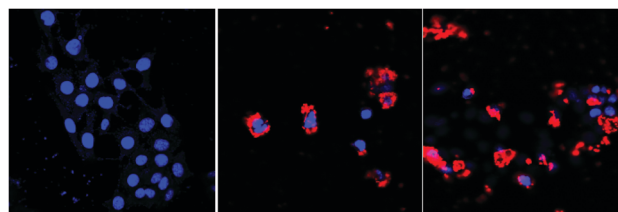


Fig. 10 Confocal laser scanning microscopy images of 4T1 cells incubated with **PDI-SY** (5  $\mu\text{M}$ ) for 30 min (left) and then further incubated with  $\text{NaClO}$  (10  $\mu\text{M}$ ) for 30 min (middle and right). Blue channel:  $\lambda_{\text{ex}}$  = 358 nm; red channel:  $\lambda_{\text{ex}}$  = 456 nm.

result of the oxidation reaction produces a new 1,3,4-oxadiazole ring that increases the conjugate plane of the new molecule, blocking the PET process in the probe, and producing the ICT process. Thus, the fluorescence of the probe recovered, and a new absorption peak at 720 nm appeared. Moreover, the reaction mechanism of **PDI-SY** with the  $\text{ClO}^-$  ions has been proved by  $^1\text{H}$  NMR. As shown in Fig. 9, the  $^1\text{H}$  NMR shifts of  $\text{C}=\text{N}-\text{H}$  in the probe is 11.0 ppm. However, after two equiv. of  $\text{ClO}^-$  was added, the peak at 11 ppm disappeared completely; simultaneously, no other new peak appeared. This change of  $^1\text{H}$  NMR combined with the spectral property of **PDI-SY** strongly proves the possible mechanism of **PDI-SY** for the detection of  $\text{ClO}^-$  ions, as shown in Fig. 8. This discussion above is helpful for us to develop some new probes in the follow-up research work.

### Confocal laser scanning microscopy images of **PDI-SY** with 4T1 cells

4T1 cells ( $\approx 10\,000$  cells per well) plated on a 96-well plate were cultured in fresh media that contained 10% fetal calf serum. **PDI-SY** at various concentrations was added and cultivated at 37  $^\circ\text{C}$  for 0.5 h. Then, 20  $\mu\text{l}$  of  $\text{NaClO}$  (10  $\mu\text{M}$ ) was added to the well for 30 min. The CLSM imaging was performed on an Olympus FLUOVIEW FV1000 confocal laser scanning microscope. The cell nuclei were stained with DAPI (blue) and the cytoplasm was stained with  $\text{NaClO}$  modified **PDI-SY** (red), as shown in Fig. 10.

## Conclusions

In summary, a near-infrared and ratiometric fluorescent probe **PDI-SY** designed on the PET and ICT mechanism, which



features a robust fluorescence “light-up” responses to  $\text{ClO}^-$  was successfully synthesized and **PDI-SY** was fully characterized by  $^1\text{H}$  NMR,  $^{13}\text{C}$  NMR, and high-resolution mass spectrometry. The probe can efficiently monitor  $\text{ClO}^-$  with rapid spectral responses within 5 s, low detection limit ( $0.8 \times 10^{-7}$  M) and intense color changes (from purple to light green). In addition, **PDI-SY** can be used to detect the hypochlorite for 5 cycles, the new detection mechanism of **PDI-SY** for  $\text{ClO}^-$  ions was proved by  $^1\text{H}$  NMR for the first time. More importantly, the application of  $\text{ClO}^-$  detection in cell imaging has also been achieved. Therefore, we expect that **PDI-SY** may provide a new strategy for the detection of  $\text{ClO}^-$  in biological application.

## Experimental

### Materials and methods

$^1\text{H}$  NMR and  $^{13}\text{C}$  NMR spectra were recorded on a Bruker ADVANCE 300 NMR spectrometer in  $\text{CDCl}_3$  or DMSO. HRMS were recorded on an Ultraflex II MALDI-TOF mass spectrometer. UV-visible absorption spectra were recorded on a Shimadzu UV-3600 spectrophotometer. Fluorescence spectra were recorded on a HORIBA FL-4 Max spectrometer. FT-IR spectra were recorded on a Nicolet 750 series in the region of  $4000\text{--}400\text{ cm}^{-1}$ . The CLSM imaging was performed on Olympus FLUOVIEW FV1000 confocal laser scanning microscope. All the reagents used were purchased from Aladdin. The solvent used in the test was PBS buffer.

### Synthesis

Compound 2-(hydrazonomethyl)phenol was prepared our laboratory according to the modified published methods.<sup>38</sup> The structure of the compounds **PDI-S**, **PDI-Y** and the compound **PDI-SY** was characterized by  $^1\text{H}$  NMR and their spectra have been displayed in the ESI.†

**Compound PDI-S.** Compound 1,7-bis(4-*tert*-butylphenoxy)perylene-3,4:9,10-tetracarboxylic acid bisanhydride (4.0 g, 7.3 mmol), and *n*-butylamine (0.53 g, 1.5 mmol) was dissolved in ethanol (80 ml), the mixture was heated under reflux for 12 h. After cooled to room temperature, the solvent was removed under reduced pressure, a red-black solid was observed, which was washed with water (100 ml) and dried in  $60\text{ }^\circ\text{C}$ . It was used for the next reaction without further purification.  $^1\text{H}$  NMR (DMSO, ppm):  $\delta$  8.67–8.65 (d, 3H,  $J = 6$  Hz), 8.46–8.44 (d, 3H,  $J = 6$  Hz), 4.15–4.10 (m, 4H), 1.96–1.95 (m, 2H), 1.41–1.23 (m, 6H), 0.97–0.90 (m, 6H).

**Compound PDI-Y.** In  $\text{N}_2$  atmosphere, compound **PDI-S** (3.0 g, 4.5 mmol), *o*-hydroxybenzaldehyde (1.1 g, 9 mmol),  $\text{K}_2\text{CO}_3$  (1.5 g, 15 mmol), and DMF (40 ml) were mixed and the mixture was heated at  $100\text{ }^\circ\text{C}$  for 8 h. After cooled to the room temperature, water (40 ml) was added and the pH value was adjusted to weak acidic using 1N hydrochloric acid resulting in the precipitation of black solid. Subsequently it was dried in a vacuum drying oven at  $50\text{ }^\circ\text{C}$  for 24 hours, to give a yellow solid 3.2 g, yield: 95%. It was used for the next reaction without further purification.  $^1\text{H}$  NMR (DMSO, ppm):  $\delta$  10.25 (s, 2H), 8.99 (s, 4H),

7.99–7.94 (m, 2H), 7.64–7.52 (m, 3H), 7.49–7.25 (m, 3H), 6.97–6.94 (m, 6H), 4.15–4.10 (m, 4H), 1.98–1.96 (m, 2H), 1.41–1.23 (m, 6H), 0.97–0.90 (m, 6H). TOF-MS-ES:  $m/z$ . Calculated:  $[\text{M} + \text{H}]^+ = 743.238$ , found:  $[\text{M} + \text{H}]^+ = 743.236$ .

**Compound PDI-SY.** Compound **PDI-Y** (2.0 g, 2.7 mmol), compound salicylhydrazine (5.9 mmol excess), and DMF (60 ml) was mixed, the mixture was heated at  $80\text{ }^\circ\text{C}$  for 12 h. After cooled to room temperature, water (60 ml) was added resulting in the precipitation of a black solid which was filtrated and dried in a blast drying oven at  $60\text{ }^\circ\text{C}$ . Then, it was purified *via* recrystallization by ethanol to give a black solid 2.9 g, yield: 92%. FT-IR (KBr)  $\text{cm}^{-1}$ : 3452 ( $\nu_{\text{C}=\text{H}}$ ); 2932–2980 ( $\nu_{\text{C}-\text{H}}$ ); 1602 ( $\nu_{\text{as}}\text{H}-\text{C}=\text{C}$ ) 1498 ( $\nu_{\text{s}}\text{C}=\text{C}$ ) 1276 ( $\nu_{\text{C}-\text{O}}$ ).  $^1\text{H}$  NMR (DMSO, ppm):  $\delta$  11.12 (s, 4H), 8.99 (s, 4H), 7.68 (m, 5H), 7.42–7.38 (t, 5H,  $J = 6$  Hz), 6.99–6.95 (m, 10H), 3.86–3.80 (m, 4H), 2.07 (m, 2H), 1.25 (m, 6H), 0.88 (m, 6H).  $^{13}\text{C}$  NMR (DMSO, ppm):  $\delta$  192.17, 162.78, 161.20, 136.89, 129.65, 122.76, 117.69, 36.25, 31.24, 22.97, 14.39. TOF-MS-ES:  $m/z$ . Calculated:  $[\text{M} + \text{H}]^+ = 979.3450$ , found:  $[\text{M} + \text{H}]^+ = 979.3450$ .

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

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