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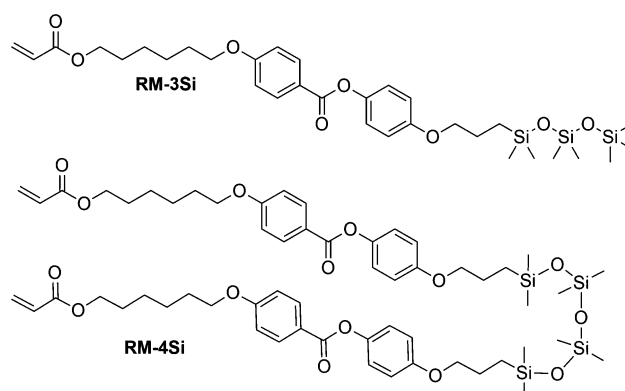
Reactive oligo(dimethylsiloxane) mesogens and their nanostructured thin films†

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Oligo(dimethylsiloxane)-based reactive mesogens were prepared and shown to form room-temperature smectic phases which were 'frozen-in' by photopolymerization. Homeotropically aligned, nanostructured thin films were obtained by spincoating, and micropatterning was demonstrated. These hybrid reactive mesogens are suitable for the preparation of aligned nanostructured polymer thin films with potential applications ranging from stimuli responsive coatings to nanoporous membranes.

While flat panel displays were the initial driving force behind the development of reactive mesogens (RMs), these materials are currently finding applications as functional nanostructured materials,^{1–3} for lab-on-a-chip systems,⁴ responsive coatings,⁵ actuators,⁶ membranes,^{7–9} and biosensors.¹⁰ In general, these applications rely on the self-assembly of reactive mesogenic precursors into an ordered LC state, their alignment into macroscopically oriented nanostructures, and the 'freezing-in' of the ordered state through polymerization to form temperature stable nanostructured polymers.^{11–14} In principle, the properties and functionality of the polymers can be tuned by the modification of the structures of the reactive mesogens.

Hybrid reactive mesogens containing organic and inorganic segments such as oligo(dimethylsiloxane) (ODMS) are of interest for various reasons.¹⁵ For example, due to the incompatibility and large surface tension mismatch of ODMS with most organic materials, nanophase segregation is promoted^{16–18} which leads to gel formation,¹⁹ and highly ordered nanostructured morphologies.^{20–23} Since ODMS can act as a precursor to inorganic silicon oxides through plasma oxidation, these materials could act as templates in porous membranes or nanolithography.^{20,24} Furthermore, the low glass transition temperature leads to self-healing properties,²⁵ and advantageous thermoplastic behavior for polymeric actuator materials requiring a large response.^{26–29} While these materials are typically prepared in cells using alignment layers,⁶ the alignment



Scheme 1 Molecular structures of the reactive ODMS mesogens **RM-3Si** and **RM-4Si**.

and polymerization in thin films is crucial for smart coating applications. ODMS has been incorporated into calamitic,^{30–33} discotic,^{34–36} and bent-core mesogens,^{21,37} but anisotropic thin film coatings consisting of ODMS RMs have not been reported.

In this work, we prepared a mono-reactive liquid crystal with an ODMS end group, and a di-reactive liquid crystal based on an ODMS bridge, **RM-3Si** and **RM-4Si**, respectively (Scheme 1). The synthesis and characterization of the novel liquid crystals is presented, as well as a characterization of the liquid crystalline behavior of the mesogens and their derived polymers. The reactive mesogens were aligned in thin films driven by the low surface energy of oligo(dimethylsiloxane), and photo-polymerized into nanostructured polymer films.

Results and discussion

Reactive mesogens **RM-3Si** and **RM-4Si** (Scheme 1) were synthesized according to Scheme 2. 4-(Allyloxy)phenyl acetate **1** was prepared by acetylation of 4-(allyloxy)phenol.³⁸ Protection of

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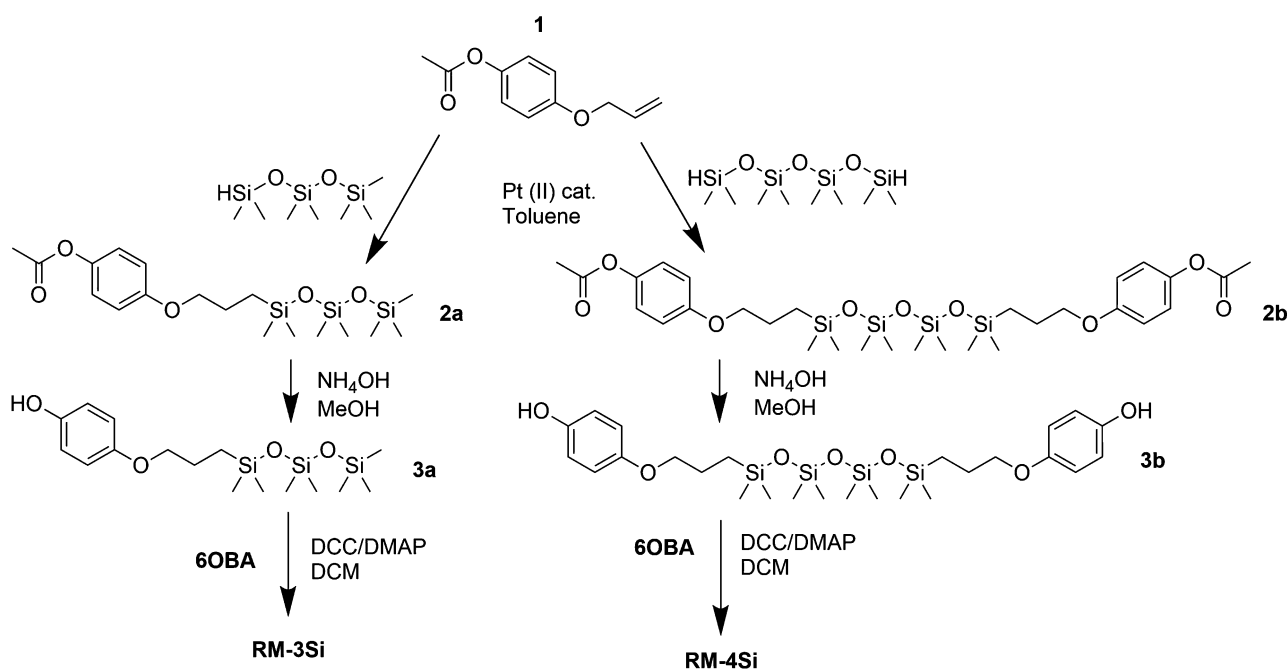


this phenol as an acetate was required for the subsequent catalytic hydrosilylation reaction with siloxane hydrides. To prepare **RM-3Si**, 1.1 eq. of **1** was reacted with 1.0 eq. of 1,1,1,3,3,5,5-heptamethyltrisiloxane in the presence of dichloro(1,5-cyclooctadiene)-platinum(II) in toluene, overnight at 60 °C in a Schlenk setup, to give **2a** (yield 50%). The acetate protecting group was removed with ammonium hydroxide in methanol to obtain **3a** as a clear viscous fluid (yield 71%). Procedures described in the literature were used to obtain 4-(6-acryloyloxyhexyloxy)benzoic acid¹² (6OBA), which was reacted with 1.0 eq. of **3a** in an esterification reaction using dicyclohexylcarbodiimide (DCC) and 4-dimethylaminopyridine (DMAP) to obtain **RM-3Si** (yield 28%). For preparation of the **RM-4Si**, 2.2 eq. **1** was reacted with 1.0 eq. of 1,1,3,3,5,5,7,7-octamethyltetrasiloxane under identical conditions as the preparation of **2a** to obtain **2b** (yield 67%). After deprotection **3b** was obtained (yield 87%) which was reacted with 2.0 eq. 6OBA to form the desired product **RM-4Si** (yield 89%). The ODMS RMs were purified by column chromatography and characterized by TLC, NMR, TGA, and MALDI-TOF MS (characterization data are gathered in the ESI†).

To investigate the LC phase behavior, the LCs were analyzed by differential scanning calorimetry (DSC), X-ray diffraction (XRD), and polarized optical microscopy (POM). Fig. 1a shows the DSC data obtained on second cooling of the two compounds. The phase transition temperatures of **RM-3Si** are $I \rightarrow 31.7\text{ °C}$ (7.93 J g^{-1}) $\rightarrow S_A \rightarrow 0.39\text{ °C}$ (36.79 J g^{-1}) $\rightarrow K$. Meanwhile, the phase transitions for **RM-4Si** are $I \rightarrow 58.4\text{ °C}$ (11.22 J g^{-1}) $\rightarrow S_A \rightarrow -4.5\text{ °C}$ (13.93 J g^{-1}) $\rightarrow K$. Here, K, S_A , and I denote the crystalline, smectic A, and isotropic phases, respectively, and the associated enthalpy changes are given in parentheses. In both cases, the compounds are liquid crystalline at room temperature. Compared to similar LCs without an ODMS segment, crystallization is

suppressed,³⁹ an effect likely due to the high flexibility of the siloxane oligomers.^{33,40} For the same reason, the isotropization temperature is suppressed.¹⁶ Fig. 1b shows the room temperature X-ray diffraction (XRD) spectra measured in both wide-angle and medium-angle configurations on kapton tape (see methods). For **RM-3Si**, two sharp diffraction peaks are found at medium angles ($q = 1.58\text{ nm}^{-1}$ and $q = 3.15\text{ nm}^{-1}$), which corresponds to a periodic layered structure with an interlayer distance of $d = 4.0\text{ nm}$. Since this value is approximately equal to the molecular length (determined by MM2 molecular simulation), it follows that the LC phase of **RM-3Si** can be assigned as a smectic A phase. For **RM-4Si**, two sharp diffraction peaks are found at medium angles ($q = 1.71\text{ nm}^{-1}$ and $q = 3.42\text{ nm}^{-1}$), corresponding to a layered structure with an interlayer distance of $d = 3.6\text{ nm}$. Since the interlayer distance is approximately equal to half the length of the extended dimer, the LC phase of **RM-4Si** is assigned as a smectic A phase, in which the siloxane bridge and organic tail moieties are partially mixed.^{32,41,42} For both LCs, a broad peak is observed in the wide-angle region, signifying a combination of siloxane-siloxane interactions at 0.70 nm ,³⁷ organic-organic interactions at 0.45 nm , and intermediate organic-siloxane interactions at 0.55 nm . This indicates that there is partial nanophase segregation between siloxane and organic units, which is presumably the driving force behind the formation of the smectic layered morphologies,¹⁶ rather than the lower-order nematic ordering.³⁹ Schematic representations of the liquid crystal phases are included in Fig. 1c, in which the mesogenic core (red) form layers with the partially phase separated siloxane (blue) and reactive aliphatic tails (black).

To photopolymerize the reactive hybrid mesogenes, a cell ($\sim 7\text{ }\mu\text{m}$) was filled with the LCs containing a small amount of photoinitiator. When observed under crossed polarizers in a



Scheme 2 Synthetic routes for the preparation of reactive mesogens **RM-3Si** and **RM-4Si**.



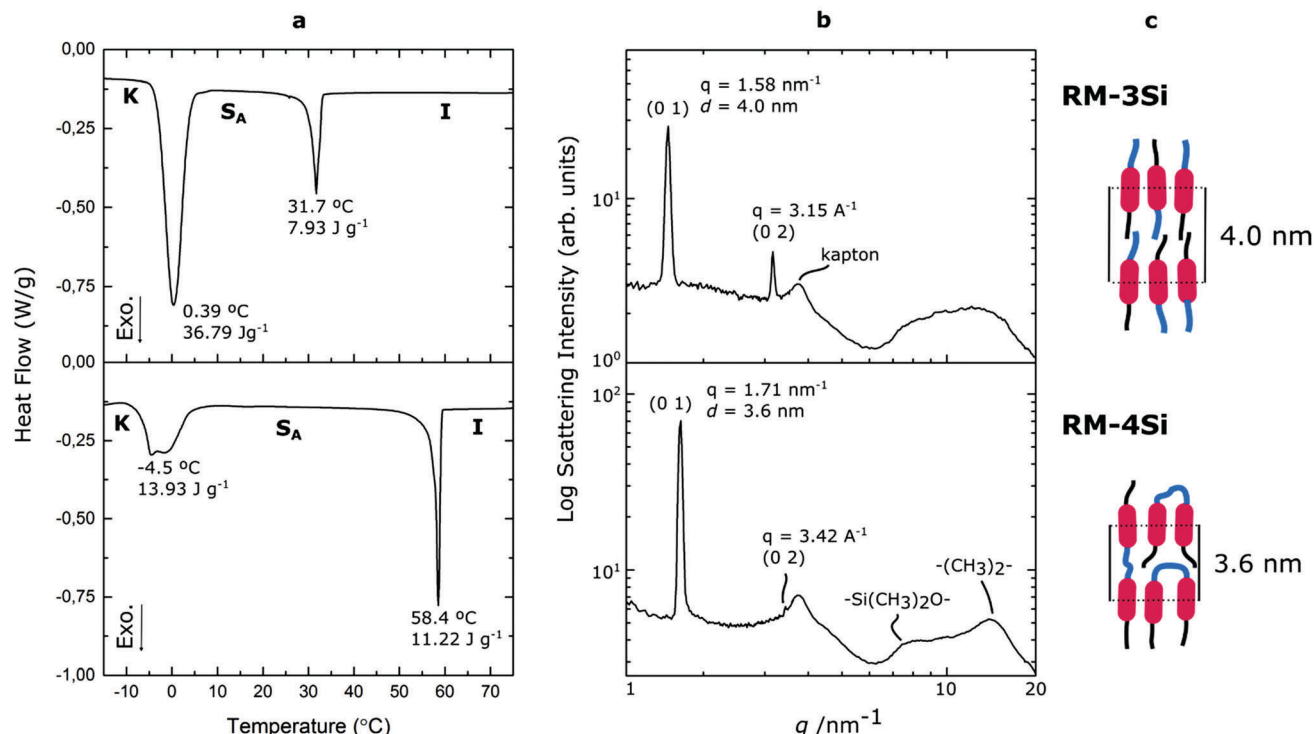


Fig. 1 (a) DSC thermographs and (b) room-temperature X-ray diffraction spectra of the reactive siloxane liquid crystals. The assigned Miller indices and the respective intermolecular scatterings are indicated, as well as the assigned lattices. (c) Schematic representation of the room-temperature smectic phases formed by the liquid crystals; mesogenic core (red), ODMS (blue) and reactive aliphatic tails (black).

POM experiment (Fig. 2a), the LCs exhibit a fan-shape texture characteristic of the smectic A phase. The cell was subsequently exposed to UV for 20 min at room temperature. Under these conditions, Fourier Transform Infrared Spectroscopy (FTIR) indicated the complete conversion of reactive acrylate groups, by the disappearance of the peak at 1635 cm^{-1} that corresponds to the C=C bond of the acrylates (Fig. S1, ESI†). The birefringence is largely maintained in the polymerized samples (Fig. 2b), indicating the 'freeze-in' of the SmA phase in the polymeric material. The polymerized samples were further heated under the POM to explore the stability of the LC phase (Fig. 2c). In this case, the linear **RM-3Si** polymer showed a reversible isotropic phase transition at 175 °C , which represents a significant increase compared to the corresponding monomer. The crosslinked liquid crystal network of **RM-4Si** did not become isotropic within the studied temperature range ($<250\text{ °C}$). However, it showed a reversible phase transition at 175 °C , accompanied by a lower birefringence. DSC performed on the polymeric material further confirm the stability of the polymerized films (Fig. S2, ESI†).

In order to prepare nanostructured thin films, dilute solutions of **RM-4Si** were spin coated on a Si wafer containing an Anti-Reflective Coating (ARC), and subsequently polymerized (see Experimental). The solution concentration was varied between 0.25 wt% and 1 wt%, to yield film thicknesses after polymerization between 11 nm (~ 3 smectic layers) and 60 nm (~ 17 smectic layers) (Fig. S3, ESI†). Across the thickness range, the static contact angle (H_2O) of the polymerized films was equal to 105° (Fig. S3, ESI†). Since this value is approximate to the contact

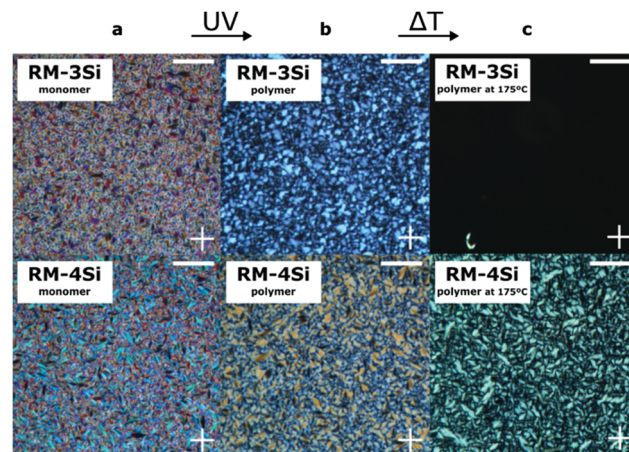


Fig. 2 POM images obtained between crossed polarizers of the LCs in a cell (a) before photopolymerization, (b) after photopolymerization, and (c) after heating the samples. Crossed polarizer axes are indicated with a cross. All scale bars: 100 μm.

angle for PDMS ($\sim 107^\circ$),⁴³ it suggests the predominant presence of siloxanes at the air interface.

The polymerized film of 60 nm thickness was studied with wide-angle grazing-incidence X-ray diffraction (GIWAXS) to determine the through-film morphology. In the small-angle region, scattering peaks are observed at $q = 1.84\text{ nm}^{-1}$ and $q = 3.66\text{ nm}^{-1}$, perpendicular to the substrate, corresponding to a layered structure with a periodicity of 3.4 nm. Hence, the smectic LC phase is maintained in the thin polymerized film,

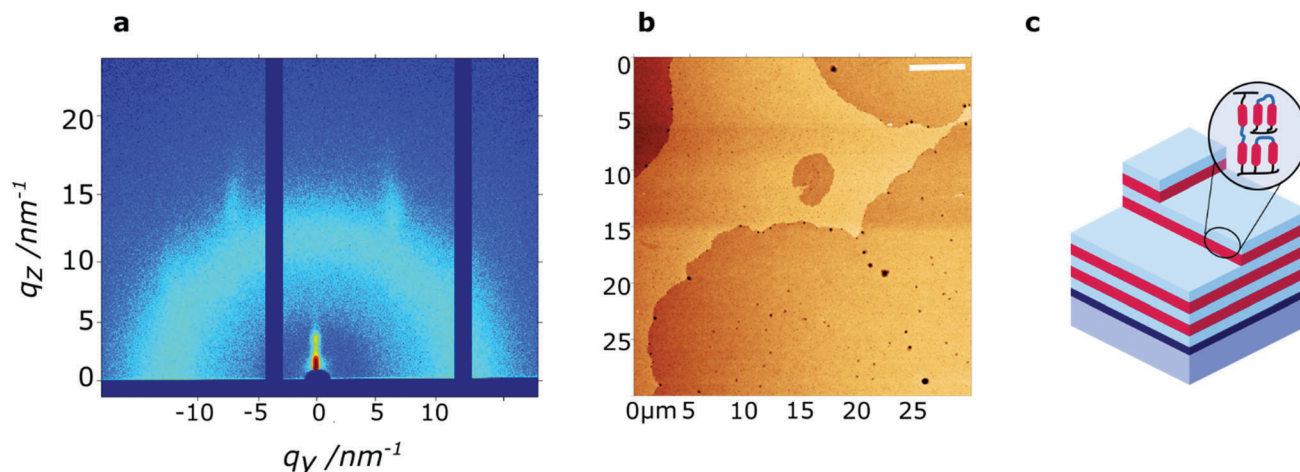


Fig. 3 Characterization of a polymerized **RM-4Si** thin film of 60 nm thickness. (a) grazing-incidence X-ray diffraction data, measured around the critical angle, obtained at room temperature. (b) Low-resolution AFM height image, obtained at room temperature. Scale bar: 5 μm . (c) Schematic representation of the nanostructured thin film.

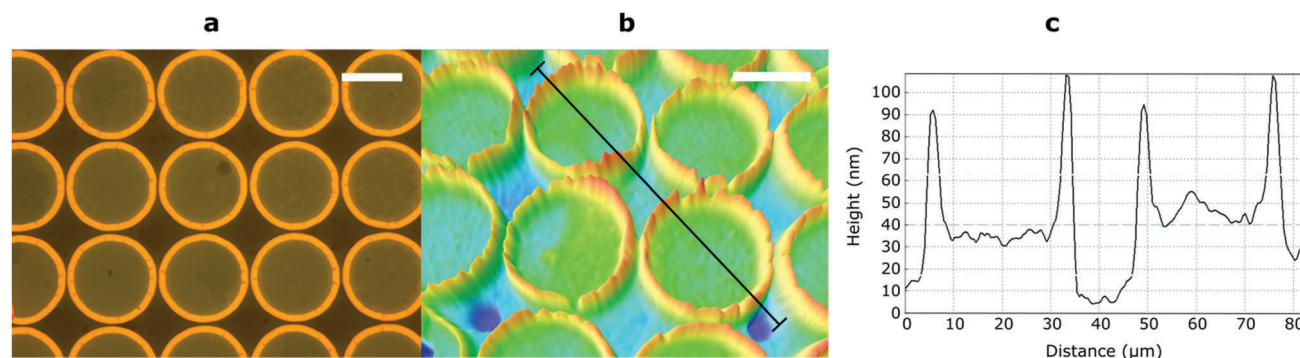


Fig. 4 Characterization of a micropatterned **RM-4Si** thin film of 60 nm thickness. (a) Optical microscopy image. (b) Optical profilometric image with cross-section (black line) (c) cross-section profile corresponding to (b). Scale bars: 20 μm .

and organized preferentially in a homeotropic orientation. Likely, the homeotropic orientation is driven by the low surface energy of PDMS ($\gamma_{\text{PDMS}} = 20.5 \text{ mN m}^{-1}$),⁴³ as previously reported for siloxane-based LC thin films.²⁰ The broad reflections at $7 \text{ nm}^{-1} < q < 14 \text{ nm}^{-1}$ signify that the isotropic nature of the respective intermolecular interactions (siloxane-siloxane and organic-organic) is maintained upon photopolymerization. Comparison of the samples before and after photopolymerization (Fig. S4, ESI[†]), reveals that the smectic LC order is maintained upon photopolymerization.

Low-resolution AFM height images of the 60 nm thin film show a layered hole-island morphology (Fig. 3b) with a step-height of approximately 3.5 nm, confirming the presence of smectic layers at the film surface. While the films are generally smooth (RMS roughness $\sim 0.55 \text{ nm}$), holes of approximately 10–100 nm diameter are present on the film surface. High-resolution AFM images, obtained in peakForce tapping mode, further indicate the presence of small defects (5–20 nm diameter) on the film surface (Fig. S5, ESI[†]).

Finally, micropatterning was achieved by exposing a 60 nm film of **RM-4Si** through a mask (25 μm diameter circular holes). After the mask exposure, the nonpolymerized material was

removed by developing the film in the casting solvent. Optical microscopy revealed the presence of circular features corresponding to the dimensions of the mask (Fig. 4a). Interestingly, the features are surrounded by an additional structure at the edges, observable as a yellow ring. Since the optical contrast arises due to height differences in the film, the film was further characterized by optical profilometry (Fig. 4b). Here, the rings appear as protrusions which rise up to 50 nm above the circular features, which is remarkable considering the initial height of the film (60 nm). This phenomenon is likely due to the diffusion of reactive mesogens to areas of low concentration during the polymerization process.⁴⁴ A cross-section across the circular features (black line) (Fig. 4c) indicates that the height of the circular features is approximately 30–40 nm, which is lower than the initial height of the film (60 nm). This may be due to (limited) polymerization in the unexposed areas due to diffuse light.

Conclusions

In this study, an oligo(dimethylsiloxane)-based reactive monoacrylate and a similar diacrylate were prepared and fully characterized.



The room-temperature smectic phases formed by these compounds were shown to be 'frozen-in' through photopolymerization, resulting in the preparation of stable nanostructured polymers. Aligned thin films were fabricated by spin-coating and subsequent photopolymerization. These films exhibit smectic layers aligned parallel throughout the film (GIWAXS) and at the surface (AFM), likely driven by the low surface energy of the ODMS. Finally, the *in situ* polymerization approach allows for micropatterning. Hence, this fabrication method is highly suitable for the preparation of aligned nanostructured polymers with applications as thin-film actuators or well-defined membrane materials.

Experimental

The reactive LC mixtures used for the experiments contain the reactive mesogens as well as 2 wt% of photo initiator 1-hydroxycyclohexyl-phenyl ketone (Irgacure 184, Ciba Specialty Chemicals). To make homogenous blends, the components were dissolved in chloroform and the solvent was evaporated. For thin film preparation, dilute solutions were spin coated at 3000 rpm for 45 seconds. For photopolymerization, the LCs were irradiated while under nitrogen atmosphere for 20 minutes, at room temperature with 365 nm light from a Mercury lamp ($\sim 5 \text{ mW cm}^{-2}$). Film thickness was obtained using a Bruker DektakXT stylus profiler. Film thickness was obtained from polymerized films only due to the very soft nature of the as-deposited films. Polarized optical microscopy (POM) was performed with crossed polarizers using a Leica DM6000M equipped with a DFC420C camera and a Linkam THMS600 hot-stage for temperature control. For characterization of the LC textures, a small amount of LC material was placed between microscope slides, heated to the isotropic phase, and cooled ($5 \text{ }^{\circ}\text{C min}^{-1}$) to room temperature. Phase transition temperatures of the reactive mesogens were determined using a TA Instrument Q1000 differential scanning calorimeter (DSC). 3–4 mg of material was hermetically sealed in aluminum pans. The heating and cooling rate was $5 \text{ }^{\circ}\text{C min}^{-1}$ and the second cooling curve was used to determine the transition temperatures. X-ray scattering measurements were performed on a Ganesha lab instrument equipped with a GeniX-Cu ultralow divergence source producing X-ray photons with a wavelength of $1.54 \text{ \AA}$ and a flux of $1 \times 10^8 \text{ ph s}^{-1}$. Scattering patterns were collected using a Pilatus 300 K silicon pixel detector. The beam center and the q range were calibrated using the diffraction peaks of silver behenate. Due to the sticky nature of the LCs, XRD samples were prepared by placing a small amount of LC between kapton tape, leading to a background signal at $q = 0.39$. The sample to detector distance was 91 mm for wide-angle measurements and 441 mm for medium-angle measurements, and the spectra were concatenated at $q = 0.36$. Grazing incidence XRD was performed at the critical angle ($\sim 0.18^{\circ}$). The AFM data were recorded in ambient conditions using a Dimension ICON atomic force microscope (Bruker Nano Inc., Santa Barbara, CA) fitted with an NCHV silicon probe (Bruker, spring constant 42 N m^{-1} and a resonance frequency 320 kHz). Fourier transform infrared spectroscopy (FTIR) spectra were obtained using a FTS 6000 spectrometer

from Bio-Rad equipped with Specac Golden gate diamond ATR and were signal-averaged over 50 scans at a resolution of 1 cm^{-1} . Optical profilometric measurements were performed on a Sensofar Plu 2300 microscope equipped with a lens with $150\times$ zoom using blue light.

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References

- 1 D. J. Broer, C. M. W. Bastiaansen, M. G. Debije and A. P. H. J. Schenning, *Angew. Chem. – Int. Ed.*, 2012, **51**, 7102–7109.
- 2 J. P. F. Lagerwall and G. Scalia, *Curr. Appl. Phys.*, 2012, **12**, 1387.
- 3 T. Kato, N. Mizoshita and K. Kishimoto, *Angew. Chem., Int. Ed.*, 2006, **45**, 38–68.
- 4 C. L. van Oosten, C. W. M. Bastiaansen and D. J. Broer, *Nat. Mater.*, 2009, **8**, 677–682.
- 5 D. Liu and D. J. Broer, *Nat. Commun.*, 2015, **6**, 8334.
- 6 H. Shahsavan, S. M. Salili, A. Jákli and B. Zhao, *Adv. Mater.*, 2015, **27**, 6828–6833.
- 7 X. D. Feng, M. E. Tousley, M. G. Cowan, B. R. Wiesenauer, S. Nejati, Y. Choo, R. D. Noble, M. Elimelech, D. L. Gin and C. O. Osuji, *ACS Nano*, 2014, **8**, 11977–11986.
- 8 C. L. Gonzalez, C. W. M. Bastiaansen, J. Lub, J. Loos, K. Lu, H. J. Wondergem and D. J. Broer, *Adv. Mater.*, 2008, **20**, 1246–1252.
- 9 H. P. C. van Kuringen, G. M. Eikelboom, I. K. Shishmanova, D. J. Broer and A. P. H. J. Schenning, *Adv. Funct. Mater.*, 2014, **24**, 5045–5051.
- 10 S. J. Woltman, G. D. Jay and G. P. Crawford, *Nat. Mater.*, 2007, **6**, 929–938.
- 11 J. Lub, V. Recaj, L. Puig, P. Forcén and C. Luengo, *Liq. Cryst.*, 2004, **31**, 1627–1637.
- 12 D. J. Broer, J. Boven, G. N. Mol and G. Challa, *Makromol. Chem.*, 1989, **190**, 2255–2268.
- 13 J. Lub and D. Broer, *Cross-Linked Liquid Crystalline Systems*, CRC Press, 2011, pp. 3–47.
- 14 I. Dierking, *Adv. Mater.*, 2000, **12**, 167–181.
- 15 For silsesquioxane based hybrid liquid crystals and functional materials, see: G. H. Mehl and I. M. Saez, *Appl. Organomet. Chem.*, 1999, **13**, 261–272; Q. Ye, H. Zhou and J. Xu, *Chem. – Asian J.*, 2016, **11**, 1322–1337; X. Wang, C. M. Cho, W. Y. Say, A. Y. X. Tan, C. He, H. S. O. Chan and J. Xu, *J. Mater. Chem.*, 2011, **21**, 5248–5257.
- 16 C. Pugh, J. Y. Bae, J. Dharia, J. J. Ge and S. Z. D. Cheng, *Macromolecules*, 1998, **31**, 5188–5200.



- 17 I. Dierking, L. Komitov, S. T. Lagerwall, T. Wittig and R. Zentel, *Liq. Cryst.*, 1999, **26**, 1511–1519.
- 18 S. Westphal, S. Diele, A. Mädicke, F. Kuschel, U. Scheim, K. Rühlmann, B. Hisgen and H. Ringsdorf, *Makromol. Chem., Rapid Commun.*, 1988, **9**, 489–493.
- 19 A. Noro, M. Hayashi, A. Ohshika and Y. Matsushita, *Soft Matter*, 2011, **7**, 1667–1670.
- 20 K. Nickmans, J. N. Murphy, B. de Waal, P. Leclère, J. Doise, R. Gronheid, D. J. Broer and A. P. H. J. Schenning, *Adv. Mater.*, 2016, **28**, 10068–10072.
- 21 R. Achten, A. Koudijs, M. Giesbers, A. Marcelis and E. Sudholter, *Mol. Cryst. Liq. Cryst.*, 2007, **477**, 169–176.
- 22 M. Funahashi, *Mater. Chem. Front.*, 2017, **1**, DOI: 10.1039/C6QM00263C.
- 23 M. Funahashi and A. Sonoda, *J. Mater. Chem.*, 2012, **22**, 25190–25197.
- 24 C. A. Ross, Y. S. Jung, V. P. Chuang, F. Ilievski, J. K. W. Yang, I. Bitá, E. L. Thomas, H. I. Smith, K. K. Berggren, G. J. Vancso and J. Y. Cheng, *J. Vac. Sci. Technol., B: Microelectron. Nanometer Struct. – Process., Meas., Phenom.*, 2008, **26**, 2489–2494.
- 25 S. J. Clarson and J. E. Mark, in *Polymer science and technology*, ed. S. J. Clarson and J. A. Semlyen, Prentice Hall PTR, Englewood Cliffs, NJ, 1993, pp. 616–648.
- 26 M. Wang, L.-X. Guo, B.-P. Lin, X.-Q. Zhang, Y. Sun and H. Yang, *Liq. Cryst.*, 2016, 1–10.
- 27 S. Guo, K. Matsukawa, T. Miyata, T. Okubo, K. Kuroda and A. Shimojima, *J. Am. Chem. Soc.*, 2015, **137**, 15434–15440.
- 28 A. Sanchez-Ferrer, A. Merekalov and H. Finkelmann, *Macromol. Rapid Commun.*, 2011, **32**, 671–678.
- 29 H. Finkelmann, H.-J. Kock and G. Rehage, *Makromol. Chem., Rapid Commun.*, 1981, **2**, 317–322.
- 30 D. J. Gardiner and H. J. Coles, *J. Appl. Phys.*, 2006, **100**, 124903.
- 31 L. Li, C. D. Jones, J. Magolan and R. P. Lemieux, *J. Mater. Chem.*, 2007, **17**, 2313–2318.
- 32 W. K. Robinson, P. S. Kloess, C. Carboni and H. J. Coles, *Liq. Cryst.*, 1997, **23**, 309–312.
- 33 E. Corsellis, D. Guillon, P. Kloess and H. Coles, *Liq. Cryst.*, 1997, **23**, 235–239.
- 34 T. J. Zhang, D. M. Sun, X. K. Ren, L. L. Liu, G. Y. Wen, Z. J. Ren, H. H. Li and S. K. Yan, *Soft Matter*, 2013, **9**, 10739–10745.
- 35 M. García-Iglesias, B. F. M. de Waal, I. de Feijter, A. R. A. Palmans and E. W. Meijer, *Chem. – Eur. J.*, 2015, **21**, 377–385.
- 36 A. Zelcer, B. Donnio, C. Bourgogne, F. D. Cukiernik and D. Guillon, *Chem. Mater.*, 2007, **19**, 1992–2006.
- 37 C. Keith, R. A. Reddy, U. Baumeister and C. Tschierske, *J. Am. Chem. Soc.*, 2004, **126**, 14312–14313.
- 38 R. Murashige, Y. Hayashi, S. Ohmori, A. Torii, Y. Aizu, Y. Muto, Y. Murai, Y. Oda and M. Hashimoto, *Tetrahedron*, 2011, **67**, 641–649.
- 39 M. Portugall, H. Ringsdorf and R. Zentel, *Makromol. Chem.*, 1982, **183**, 2311–2321.
- 40 M. Ibn-Elhaj, A. Skoulios, D. Guillon, J. Newton, P. Hodge and H. Coles, *J. Phys. II*, 1996, **6**, 271–279.
- 41 N. Olsson, B. Helgee, G. Andersson and L. Komitov, *Liq. Cryst.*, 2005, **32**, 1139–1150.
- 42 N. Olsson, M. Schröder, S. Diele, G. Andersson, I. Dahl, B. Helgee and L. Komitov, *J. Mater. Chem.*, 2007, **17**, 2517–2525.
- 43 S. Wu, *Polymer Handbook*, Wiley-Interscience, New York, 3rd edn, 1989, vol. VI.
- 44 B. M. I. van der Zande, J. Steenbakkers and J. Lub, *J. Appl. Phys.*, 2005, **97**, 123519.

