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Designable assembly of atomically precise Al_4O_4 cubane supported mesoporous heterometallic architectures†

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Heterometallic cluster-based framework materials are of interest in terms of both their porous structures and multi-metallic reactivity. However, such materials have not yet been extensively investigated because of difficulties in their synthesis and structural characterization. Herein, we reported the designable synthesis of atomically precise heterometallic cluster-based framework compounds and their application as catalysts in aldol reactions. By using the synergistic coordination protocol, we successfully isolated a broad range of compounds with the general formula, $[\text{Al}_4\text{M}_4\text{O}_4(\text{L})_{12}(\text{DABCO})_2]$ (L = carboxylates; DABCO = 1,4-diazabicyclo[2.2.2]-octane; M^{2+} = Co^{2+} , Mn^{2+} , Zn^{2+} , Fe^{2+} , Cd^{2+}). The basic heterometallic building blocks contain unprecedented main-group γ -alumina moieties and surrounding unsaturated transition metal centers. Interestingly, the porosity and interpenetration of these frameworks can be rationally regulated through the unprecedented strategy of increment of the metal radius in addition to general introduction of sterically bulky groups on the ligand. Furthermore, these porous materials are effective catalysts for aldol reactions. This work provides a catalytic molecular model platform with accurate molecular bonding between the supporters and catalytically active metal ions.

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Introduction

Aluminosilicate zeolites with 4-connected metal centers are a class of inorganic porous solids with a range of applications in catalysis, ion exchange and separation.^{1–6} Compared with these traditional 4-connected nodes,^{7–18} the use of polynuclear metal clusters as 4-connected nodes is highly desirable because of the availability of large pore sizes and the possibility of combining unique cluster-based properties.^{19–22} Such an idea has been realized by Feng *et al.*^{22,23} via the assembly of supertetrahedral metal–chalcogenide clusters. Most significantly, elements without tetrahedral geometry can be incorporated, thus providing access to multifunctional framework materials with zeolite type topology.^{24–26}

Heterometallic oxo clusters may exhibit unique properties resulting from the synergy effect of the two substantially different types of metal ions. Therefore, it is of great interest to

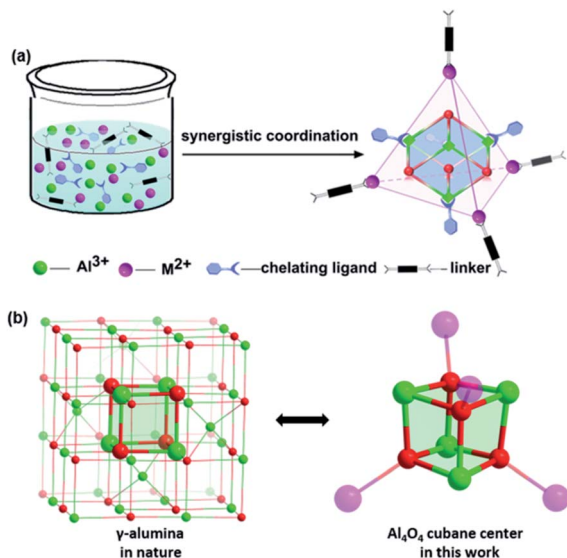
incorporate heterometallic clusters as 4-connected centers. However, extensive studies based on 4-connected heterometallic clusters remain scarce because of the lack of a strategy for efficient synthesis. There are two possible synthetic challenges: (1) competitive coordination and integration of two kinds of metal ions, and (2) metal clusters with high nuclearity and positive charge are prone to accepting more organic ligands, thus leading to highly-connected dense topologies. Once the building blocks are established, effective interpenetration regulation is another issue that should be considered.^{27,28} In 2019, Zheng *et al.*²⁹ summarized advances in the design and synthesis of cluster organic frameworks based on 4-connected metal–halide clusters,³⁰ metal–oxygen clusters,³¹ and metal–chalcogen clusters.³² In the domain of metal–oxygen clusters, only transition metal clusters (such as $\epsilon\text{-Zn}_4\text{PMo}_{12}$ ²⁶ or Cd_4Cu_6 ³³) have been reported as 4-connected nodes. However, main group element-based examples are challenging and still lacking.

$\gamma\text{-Al}_2\text{O}_3$ is one of the most common kinds of alumina, which has a large surface area, and strong adsorption capacity and has been widely used in adsorbents, catalysts and catalyst supports.³⁴ With our recent research in Al(III) chemistry, we herein demonstrate a synergistic coordination strategy towards the assembly of 4-connected heterometallic cluster-based framework materials. The strategy is based on spontaneous selective coordination between two kinds of metal centers (Al centers and d-transition metals) and two types of ligands (chelating L with O-donors and bridging DABCO with N donors,

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† Electronic supplementary information (ESI) available: Experimental section, detailed structure charts, stability analysis, sorption isotherms, characterization, catalytic mechanism, and tables. The structures reported in this article have been deposited at the Cambridge Crystallographic Data Centre (CCDC). CCDC 2056099–2056118, 2056127, 2156657 (AIOC-99–AIOC-120). For ESI and crystallographic data in CIF or other electronic format see <https://doi.org/10.1039/d2sc00526c>



Scheme 1 (a) Scheme demonstrating the synergistic coordination synthetic strategy; (b) comparison between the structures of γ - Al_2O_3 and the Al_4O_4 cubane in a tetrahedral cluster node.

respectively) (Scheme 1a).³⁵ Such an approach is universal and we successfully isolate a broad range of compounds [$\text{Al}_4\text{O}_4\text{-M}_4(\text{L})_{12}(\text{DABCO})_2$] (denoted as AIOCs; $\text{M}^{2+} = \text{Mn}^{2+}, \text{Fe}^{2+}, \text{Co}^{2+}, \text{Zn}^{2+}, \text{Cd}^{2+}$, $\text{L} = \text{carboxylates}$; $\text{DABCO} = 1,4\text{-diazabicyclo}[2.2.2]\text{octane}$) (Table 1). We were surprised to find that the Al_4O_4 cubane (denoted as the γ -alumina moiety) is located in the center of the 4-connected nodes (Scheme 1b).^{36,37} In addition, both the inorganic and organic shells of the Al_4O_4 cubane can be modified to a large extent. Interestingly, controllable interpenetration is realized through the introduction of sterically bulky groups, increment of the ionic radius of transition metal ions, and variation of the $\text{M}^{2+}/\text{Al}^{3+}$ concentration. Furthermore, these porous materials are effective catalysts for aldol reactions.

Experimental

Materials and general methods

All the reagents and solvents employed were commercially available and used without further purification. $\text{Al}(\text{O}^i\text{Pr})_3$ was acquired from Aladdin Chemical Reagent Shanghai. $\text{Fe}(\text{Ac})_2$ was obtained from TCI. Acetonitrile (99%), benzoic acid, $\text{Co}(\text{Ac})_2 \cdot 4\text{H}_2\text{O}$, $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$, $\text{Zn}(\text{Ac})_2 \cdot 2\text{H}_2\text{O}$ and $\text{Cd}(\text{Ac})_2 \cdot 2\text{H}_2\text{O}$ were bought from Sinopharm Chemical Reagent Beijing. $\text{Ti}(\text{O}^i\text{Pr})_4$ and other carboxylic acids were purchased from Adamas-beta.

The energy dispersive spectroscopy (EDS) analyses of single crystals were performed on a JEOL JSM6700F field-emission scanning electron microscope equipped with an Oxford INCA system. IR spectra (KBr pellets) were recorded on an ABB Bomem MB102 spectrometer over a range of $400\text{--}3900\text{ cm}^{-1}$. Powder X-ray diffraction (PXRD) data were collected on a Rigaku Mini Flex II diffractometer using $\text{CuK}\alpha$ radiation ($\lambda = 1.54056\text{ \AA}$) under ambient conditions. The UV-vis diffuse reflection data were recorded at room temperature using a powder sample with BaSO_4 as a standard (100% reflectance) on a PerkinElmer

Table 1 A summary of the 4-connected AIOCs^a

Complex	Composition	Sp. Gr.	R_1 Value	Void ratio
Two-fold diamond (<i>dia</i>) interpenetration				
AIOC-99	$\text{Al}_4\text{Co}_4\text{O}_4(\text{BA})_{12}(\text{DABCO})_2$	$Fd\bar{3}c$	0.0801	1.2%
AIOC-100	$\text{Al}_4\text{Co}_4\text{O}_4(2\text{-FBA})_{12}(\text{DABCO})_2$	$Fd\bar{3}c$	0.0759	No
AIOC-101	$\text{Al}_4\text{Co}_4\text{O}_4(3\text{-FBA})_{12}(\text{DABCO})_2$	$Fd\bar{3}c$	0.1149	7.6%
AIOC-102	$\text{Al}_4\text{Co}_4\text{O}_4(4\text{-FBA})_{12}(\text{DABCO})_2$	$Fd\bar{3}c$	0.1474	2.5%
AIOC-103	$\text{Al}_4\text{Co}_4\text{O}_4(2\text{-FA})_{12}(\text{DABCO})_2$	$Fd\bar{3}c$	0.0998	2.8%
AIOC-104	$\text{Al}_4\text{Co}_4\text{O}_4(2\text{-TA})_{12}(\text{DABCO})_2$	$Fd\bar{3}c$	0.1447	1.6%
AIOC-105	$\text{Al}_4\text{Co}_4\text{O}_4(3\text{-TA})_{12}(\text{DABCO})_2$	$Fd\bar{3}c$	0.1274	2.4%
AIOC-106	$\text{Al}_4\text{Mn}_4\text{O}_4(\text{BA})_{12}(\text{DABCO})_2$	$Fd\bar{3}c$	0.0858	No
AIOC-107	$\text{Al}_4\text{Fe}_4\text{O}_4(\text{BA})_{12}(\text{DABCO})_2$	$Fd\bar{3}c$	0.0566	1.0%
AIOC-108	$\text{Al}_4\text{Zn}_4\text{O}_4(\text{BA})_{12}(\text{DABCO})_2$	$Fd\bar{3}c$	0.0548	1.0%

Non-interpenetrated diamond (*dia*) topology

AIOC-109	$\text{Al}_4\text{Co}_4\text{O}_4(3\text{-MBA})_{12}(\text{DABCO})_2$	$F4_132$	0.0736	42.5%
AIOC-110	$\text{Al}_4\text{Fe}_4\text{O}_4(\text{BA})_{12}(\text{DABCO})_2$	$F4_132$	0.0634	52.8%
AIOC-111	$\text{Al}_4\text{Cd}_4\text{O}_4(\text{BA})_{12}(\text{DABCO})_2$	$F4_132$	0.0751	55.6%

Non-interpenetrated lonsdaleite (*lon*) topology

AIOC-112	$\text{Al}_4\text{Co}_4\text{O}_4(\text{BA})_{12}(\text{DABCO})_2$	$P\bar{6}2c$	0.0629	53.9%
AIOC-113	$\text{Al}_4\text{Co}_4\text{O}_4(3\text{-FBA})_{12}(\text{DABCO})_2$	$P\bar{6}2c$	0.0809	47.0%
AIOC-114	$\text{Al}_4\text{Co}_4\text{O}_4(4\text{-CBA})_{12}(\text{DABCO})_2$	$P\bar{6}2c$	0.0757	44.9%
AIOC-115	$\text{Al}_4\text{Co}_4\text{O}_4(4\text{-MBA})_{12}(\text{DABCO})_2$	$P\bar{6}2c$	0.0736	44.3%
AIOC-116	$\text{Al}_4\text{Co}_4\text{O}_4(4\text{-ABA})_{12}(\text{DABCO})_2$	$P\bar{6}2c$	0.0972	48.4%
AIOC-117	$\text{Al}_4\text{Mn}_4\text{O}_4(\text{BA})_{12}(\text{DABCO})_2$	$P\bar{6}2c$	0.0576	54.3%
AIOC-118	$\text{Al}_4\text{Fe}_4\text{O}_4(\text{BA})_{12}(\text{DABCO})_2$	$P\bar{6}2c$	0.1081	53.3%
AIOC-119	$\text{Al}_4\text{Zn}_4\text{O}_4(\text{BA})_{12}(\text{DABCO})_2$	$P\bar{6}2c$	0.0542	53.5%
AIOC-120	$\text{Al}_4\text{Cd}_4\text{O}_4(\text{BA})_{12}(\text{DABCO})_2$	$P\bar{6}2c$	0.0784	56.4%

^a Abbreviations: BA = benzoate; 2-FBA = 2-fluorobenzoate; 3-FBA = 3-fluorobenzoate; 4-FBA = 4-fluorobenzoate; 2-FA = 2-furan formate; 2-TA = 2-thiophene formate; 3-TA = 3-thiophene formate; 3-MBA = 3-methylbenzoate; 4-MBA = 4-methylbenzoate; 4-CBA = 4-chlorobenzoate; 4-ABA = 4-aminobenzoate.

Lambda-950 UV spectrophotometer and scanned at $200\text{--}1200\text{ nm}$. The absorption data were calculated from the Kubelka-Munk function, $(F(R) = (1 - R)^2/2R)$, where R represents the reflectance.³⁸ Surface chemical analyses were performed by X-ray photoelectron spectroscopy (XPS, Thermo Fisher, ESCALAB 250Xi). Metal proportional analyses were performed on an Ultima-2 inductively coupled plasma (ICP) spectrometer. The TGA curves were recorded in the region of $30\text{--}800\text{ }^\circ\text{C}$ at a heating rate of $10\text{ }^\circ\text{C min}^{-1}$ in a flowing N_2 atmosphere on a Mettler Toledo TGA/SDTA 851^e analyzer.

Synthesis of two-fold interpenetrated diamond (*dia*) frameworks AIOC-99–AIOC-108

$\text{Al}(\text{O}^i\text{Pr})_3$ (0.918 g, 4.5 mmol), $\text{Co}(\text{Ac})_2 \cdot 4\text{H}_2\text{O}$ (0.75–2.5 g, 3–10 mmol), DABCO (0.224 g, 1 mmol) and benzoic acid (0.978 g, 8 mmol) were dissolved in acetonitrile (10 mL), and then $\text{Ti}(\text{O}^i\text{Pr})_4$ (0.5 mL, 1.5 mmol) was added to the bottle and mixed at room temperature. The resultant solution was sealed in a Teflon-lined autoclave (23 mL) and heated at $160\text{ }^\circ\text{C}$ for three days. After being cooled to room temperature, red octahedral crystals of AIOC-99 were obtained (maximum yield: $\sim 64\%$ (1.5 g) for AIOC-99 based on $\text{Al}(\text{O}^i\text{Pr})_3$). Synthesis of AIOC-100–AIOC-105 is similar to that of AIOC-99 except that benzoic acid was replaced by 2-FBA, 3-FBA, 4-FBA, 2-FA, 2-TA, and 3-TA. Red octahedral



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successfully isolated (Table 1, **AIOC-99–AIOC-120**, Fig. S1–S22†). Each Al atom in all compounds adopts a six-coordination geometry, bonding to three μ_4 -O atoms and three carboxylate oxygen atoms (O_{COO}) (Fig. S23†). Every external M ion resided in a trigonal bipyramid environment with three O_{COO} atoms, one μ_4 -O bridge and one N atom from DABCO. Each triangular face of the supertetrahedral core is consolidated by three bridging bidentate chelating ligands (Fig. S24 and S25†). Compared with the charged tetrahedral cages,^{47–49} the supertetrahedral cluster $[\text{Al}_4\text{Co}_4\text{O}_4(\text{L})_{12}]$ herein is neutral. Four DABCO linkers on the cluster node allow further connections, which is similar to the 4-connected TO_4 (T = Si, Al, P, etc.) node in inorganic zeolite materials.

Interpenetration regulation

As shown in Fig. 2, the 4-connected frameworks are categorized into three types. $[\text{Al}_4\text{M}_4\text{O}_4(\text{L})_{12}(\text{DABCO})_2]$ undergoes a change in the degree of interpenetration from a dense doubly interpenetrated framework (**AIOC-99–AIOC-108**, Fig. S1–S10†) to two kinds of highly porous non-interpenetrated framework with diamond (*dia*) topology: **AIOC-109–AIOC-111**, Fig. S11–S13;† and lonsdaleite (*lon*) topology (**AIOC-112–AIOC-120**, Fig. S14–S22†). It is reported that factors that affect interpenetration include pressure, temperature, solvent, ligand design, etc.^{27,50,51} Herein, precise suppression of interpenetration by an increment of the metal radius is found in addition to the ligand design and concentration ratio of $\text{M}^{2+}/\text{Al}^{3+}$.

Interpenetration regulation through ligand design. The regulation of the topology is greatly related to the types of ligands. For benzoate or fluorobenzoate, both interpenetrated and non-interpenetrated frameworks can be obtained (interpenetrated diamond (*dia*) topology: **AIOC-99–AIOC-102**; non-interpenetrated lonsdaleite (*lon*) topology: **AIOC-112** and **AIOC-**

113). However, if large steric groups ($-\text{Cl}$, $-\text{CH}_3$ and $-\text{NH}_2$) are introduced, only non-interpenetrated frameworks can be generated (**AIOC-109** and **AIOC-114–AIOC-116**) (Fig. 2c, d and S11, S16–S18†). Intriguingly, the presence of the substituents can also enhance the intermolecular and intramolecular hydrogen bond interactions (Fig. S26†). For smaller pentacyclic carboxylates, only doubly interpenetrated frameworks (**AIOC-103–AIOC-105**) were synthesized (Fig. S5–S7†).

Interpenetration regulation through increasing the ionic radius of TM ions. Cd-containing frameworks (**AIOC-111** and **AIOC-120**) present non-interpenetrated topologies (Fig. 2e, f and S13, S22†), which are different from those of Mn, Fe, Co, and Zn involving compounds. A possible reason is attributed to its larger radius, and the size of $\text{Cd}_4\text{Al}_4\text{O}_4$ tetrahedral clusters is obviously larger than that of other ones (Fig. S27†). Although a large number of factors influencing interpenetration have been reported, the effect of the ionic radius was discovered for the first time.

Interpenetration regulation through altering the ratio of $\text{M}^{2+}/\text{Al}^{3+}$. The variation of the mole ratio of raw materials has a great impact on the structural types.^{8,50} Herein, the $\text{M}^{2+}/\text{Al}^{3+}$ concentration plays an important role in interpenetration regulation. For example, when the ratio of $\text{Co}^{2+}/\text{Al}^{3+}$ is greater than 1.11, the pure octahedral crystals with a doubly interpenetrated diamond (*dia*) framework can be obtained (Fig. 2a and b, **AIOC-99**, *Fd3c*; Fig. 3a, left), in which the 6^6 diamond (*dia*) cage is exclusively made up of six chair-like six-membered rings (Fig. 3b, left). When the ratio of $\text{Co}^{2+}/\text{Al}^{3+}$ is lower than 0.44, only pure hexagonal flaky crystals with non-interpenetrated lonsdaleite (*lon*) topologies can be generated (Fig. 2g, h and S14,† **AIOC-112**, *P62c*; Fig. 3a, right). Different from the 6^6 diamond (*dia*) cage, the 6^5 lonsdaleite (*lon*) cage in **AIOC-112** is composed of chair-like and boat-like six-membered rings

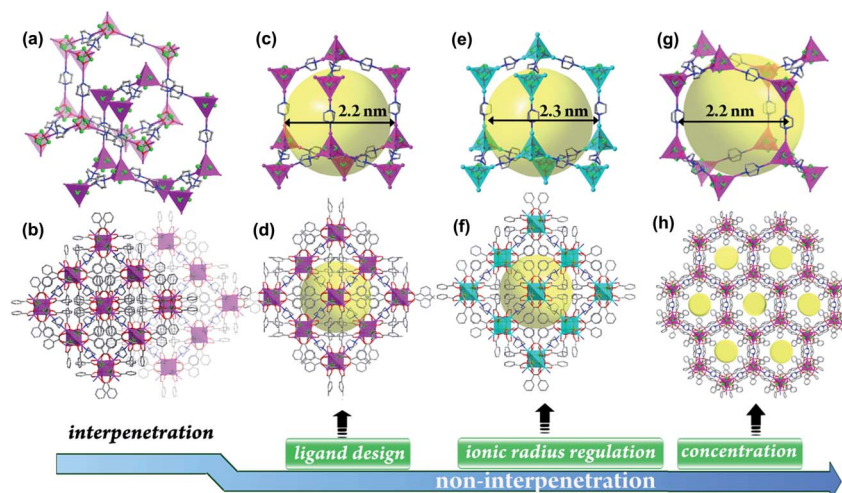


Fig. 2 The interpenetration regulation strategies: introduction of sterically bulky groups, increment of the ionic radius and decrease of the $\text{M}^{2+}/\text{Al}^{3+}$ concentration. The mesoporous cavities (a, c, e, and g) and overall frameworks (b, d, f, and h) in two-fold interpenetrated **AIOC-99**, and non-interpenetrated **AIOC-109**, **AIOC-111** and **AIOC-112**, respectively. The atom color code: O, red; C, dark grey; N, blue; Co, pink; Cd, turquoise; Al, green. Hydrogen atoms are omitted for clarity. The diamond (*dia*) frameworks present rhombic windows, while the lonsdaleite (*lon*) frameworks possess hexagonal channels (diameter: 3.761 nm; interior wall diameter: 0.909 nm) aligned parallel to the *c*-axis.



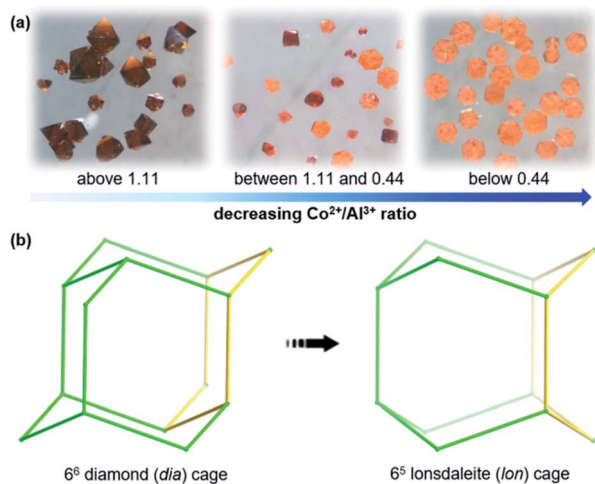


Fig. 3 (a) Photos of reaction products obtained from controlling the ratio of $\text{Co}^{2+}/\text{Al}^{3+}$. (b) The simplified diamond (*dia*) and lonsdaleite (*lon*) cages. Chair-like (left) and boat-like (right) six-membered rings have been highlighted with yellow bonds.

(Fig. 3b, right). However, when the ratio of $\text{Co}^{2+}/\text{Al}^{3+}$ is controlled between 0.44 and 1.11, the product was a mixture phase containing both octahedral crystals and hexagonal flaky crystals (Fig. 3a, middle).

Therefore, it is reasonable to infer that the low $\text{M}^{2+}/\text{Al}^{3+}$ ratio favored the non-interpenetrated lonsdaleite (*lon*) form,⁵² which is supported by the experimental results that come from the isolation of a series of isostructural Mn, Fe, Zn and Cd constructed AIOCs (AIOC-117–AIOC-120; Table 1, Fig. S19–S22†). For Fe constructed frameworks, the control of the $\text{Fe}^{2+}/\text{Al}^{3+}$ ratio is not easy due to the rapid oxidation of Fe^{2+} . Thus, the synthesis of the doubly interpenetrated diamond (*dia*) framework of AIOC-107 with a high $\text{Fe}^{2+}/\text{Al}^{3+}$ ratio should be performed in a glovebox under pumped nitrogen. Instead, non-interpenetrated *dia* framework AIOC-110 will be generated when synthesized under the same conditions without nitrogen.

Physical performance

The non-interpenetrated framework density (for example, AIOC-112 (0.823 g cm⁻³)) is approximately half that of doubly interpenetrated AIOC-99 (1.651 g cm⁻³). The formation of a non-interpenetrated framework is noteworthy given the mesoporous cavity channels (*ca.* 2.2–2.3 nm) (Fig. 2c, e, g and S30†). Correspondingly, it also exhibits obvious higher solvent-accessible volumes (*e.g.*, 53.9% for AIOC-112 calculated using PLATON).⁵³ Considering their potential application in gas adsorption and separation, N₂ sorption tests were conducted to evaluate porosity. Based on N₂ sorption isotherms, the BET surface areas were calculated to be 11.48, 779.6 and 797.2 m² g⁻¹ for AIOC-99, AIOC-109 and AIOC-112 (Fig. 4a and S31†). The pronounced H4 hysteresis loops of AIOC-109 and AIOC-112 are located at high P/P_0 ($P/P_0 = 0.4\text{--}1.0$), further confirming the presence of the mesopores in the materials.^{54,55} The experimental CO₂ uptake of AIOC-99, AIOC-109 and AIOC-112 was measured to be 10.2, 13.9 and 21.7 cm³ g⁻¹ at 273 K.

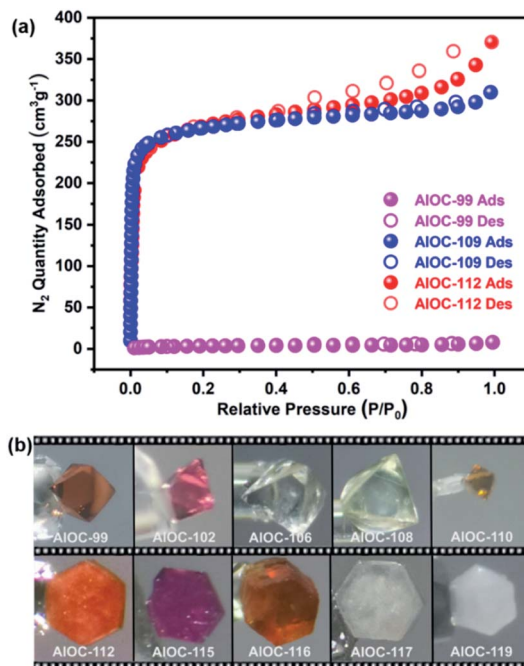


Fig. 4 (a) The N₂ sorption isotherms of AIOC-99, AIOC-109 and AIOC-112. (b) Crystal photos of representative samples.

The stability of frameworks plays an important role in research and industrial production, because it provides an applicable scope of the material. Their stability is related to interpenetration and topology. For example, the doubly interpenetrated framework is thermodynamically stable when heated to 300 °C, while the non-interpenetrated ones can only tolerate 250 °C (Fig. S32 and S33†), indicating that interpenetration in AIOC-99 obviously enhances the stability of the framework. In addition, AIOC-99 and AIOC-109 were stable after being immersed in an aqueous solution with different pH values (Fig. S34 and S35†). The lonsdaleite (*lon*) framework of AIOC-112 will collapse gradually in water, indicating that the diamond (*dia*) type framework is more stable than that of lonsdaleite (*lon*). Nevertheless, both of them were stable in common organic solvents, including MeOH, EtOH, MeCN, DMF and DMSO (Fig. S36†). Their solution stabilities are related to hydrophobic wettability. And the measured contact angles for AIOC-99, AIOC-109 and AIOC-112 are 160.58°, 140.13° and 153.23°, respectively (Fig. S37 and S38†).⁵⁶ Furthermore, all three materials can remain in the air for more than one year (Fig. S39†), which is related to the presence of the Al₄O₄ core because compounds like Co₂Bdc₂Dabco without aluminium (Fig. S40†) cannot be exposed to air for more than two weeks.^{57,58}

These compounds were all characterized by powder X-ray diffraction (PXRD), IR spectroscopy, and energy dispersive spectroscopy (EDS) analyses (Fig. S41–S68†). The atomic ratio of M to Al was determined to be approximately 1 : 1 by inductively coupled plasma (ICP) analysis, which is in agreement with the crystallographic results. The valence states of the metal ions were confirmed by bond-valence-sum (BVS) calculations and X-ray photoelectron spectroscopy (XPS) (Table S5, Fig. S69–S90†).

Intriguingly, these compounds present a variety of colors (Fig. 4b), that are affected by several factors. For example, Fe-containing compounds are easily oxidized, and their colors can be changed from yellow to red when exposed to air (Fig. S91†). The colors of Co constructed frameworks are related to the framework density and chelated ligands. Doubly interpenetrated **AIOC-99** presents darker color when compared with non-interpenetrated **AIOC-112**. Furthermore, the colors of 4-FBA coordinated frameworks (*e.g.*, **AIOC-102**) are also different from that of BA modified structures due to the electron-withdrawing effects of F centers (*e.g.*, **AIOC-99**). Diffuse reflectance spectroscopy measurements were performed on all complexes to better understand their optical absorption properties (Fig. S92–S95†).

Catalytic performance

The aldol reaction is a basic transformation in creating new carbon–carbon bonds and β -hydroxyl carbonyl compounds.⁵⁹ Considering the higher stability and catalytically active sites (unsaturated transition metals) on the Al_2O_3 support, three representative AIOCs (**AIOC-99**, **AIOC-109** and **AIOC-112**) were used as catalysts for the direct aldol reaction of acetone and nitro-substituted aromatic aldehyde in DMSO. The catalytic results are summarized in Table S6.†

In the presence of 10 mmol% catalysts, all three AIOCs show high catalytic activities giving desired aldol products in 79–88% yield. Upon completion of the reaction, all of them could be readily recovered by centrifugation and reused for the next cycle without significant loss of activity. The yield after three consecutive runs of **AIOC-109** can still remain 88% for isolated products. PXRD exhibited that the recycled sample retained high crystallinity after catalysis (Fig. S96†). In order to explore the impact of penetration on performance, the amount of catalyst is reduced to 5 mmol%, and mesoporous frameworks afford higher yield (67% for **AIOC-109**, and 68% for **AIOC-112**) compared with doubly interpenetrated **AIOC-99** (48%). This suggests that the catalytic reactions with a low content catalyst are related to the specific surface area. Besides, the control experiment without catalyst loading was also performed, and no condensation product can be observed, indicating that these crystals possess catalytic capability.

The probable catalytic mechanism is shown in Fig. 5.⁶⁰ First, the Co centers with unsaturated coordination sites polarize the carbonyl group of acetone, leading to the dissociation of the α -proton and generating a metal enolate. The oxygen atoms on BAs bond to the α -proton and leave the Co center transitorily. Then, 4-nitrobenzaldehyde combined with the metal enolate through C–C coupling, and condensation products can be obtained.

Conclusions

This is not only the first time that the γ -alumina moiety is found in crystalline porous compounds but it also serves as a support for loading various metal ions. The incorporation of hetero-metallic ions broadens the scope and potential applications of

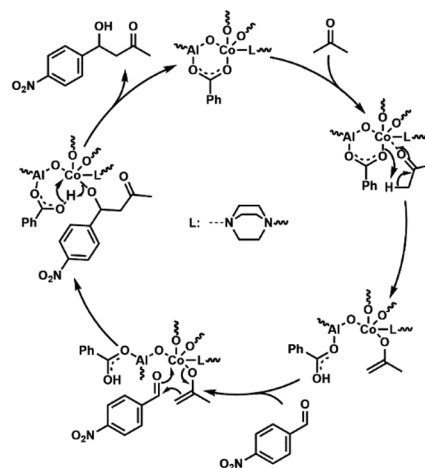


Fig. 5 Probable catalytic mechanism for the aldol condensation reaction by using **AIOC-99**, **AIOC-109** and **AIOC-112** as catalysts.

framework materials. This work provides a typical catalytic molecular model and demonstrates the unambiguous connections between the support (γ -alumina moiety) and catalytically active metal ions at the atomic level. Such γ -alumina moiety-supported materials can provide a reference for the catalytic system to achieve the hetero-disperse requirement meanwhile exposing as many catalytic sites as possible.

Data availability

All data supporting this study are available from article and ESI.† The raw data files are available from author upon reasonable request.

Author contributions

All authors contributed extensively to the work presented in this paper. J. Zhang and W.-H. Fang conceived the research project. Y.-J. Liu performed the synthesis and characterization experiments. Y. Yu performed catalytic experiments. Y.-F. Sun assisted in the data collection. W.-H. Fang and Y.-J. Liu wrote the manuscript and the ESI with input from the other authors.

Conflicts of interest

There are no conflicts to declare.

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