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Controlling the crystallinity and solubility of functional PCL with efficient post-polymerisation modification†

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Poly(ϵ -caprolactone) is a semi-crystalline biocompatible polymer with good mechanical properties. Its crystallinity also uniquely enables poly(ϵ -caprolactone) to be used in different applications, from the development of 3D scaffolds for tissue engineering to advanced nanoparticle design. However, the lack of side-chain functionality in the polymer backbone prevents further functionalisation, thereby limiting the potential to alter physico-chemical properties. Herein, we report the well-controlled ring-opening polymerisation (ROP) of ϵ -allyl caprolactone using $\text{Mg}(\text{BHT})_2(\text{THF})_2$ to furnish an allyl-functionalised PCL suitable for post-polymerisation modification. The isolated polymer could then be quantitatively post-functionalised via the efficient thiol–ene addition using a variety of commercially available alkyl thiols. The effect of alkyl chain length on bulk polymer properties was investigated, and demonstrates the potential to tune solubility and crystallinity of poly(ϵ -caprolactone) in a simple and efficient fashion. Lipophilic, functional polyesters that possess crystallinity are rare but this feature is crucial to exploit further applications of biocompatible polyesters such as for drug delivery or as a robust scaffold for tissue engineering.

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Introduction

The use of poly(ϵ -caprolactone) (PCL) for biomedical applications has attracted increasing interest, owing to its excellent biocompatibility, controlled biodegradation and robust mechanical properties.^{1–5} In particular, its semi-crystalline nature enables PCL to be used in various applications ranging from the development of 3D scaffolds for tissue engineering to the design of nanoparticles of controlled shape and precisely defined dimensions.^{5–8} The preparation of PCL *via* ring-opening polymerisation (ROP) has been widely investigated using various catalytic systems such as organic,^{9–12} metal-based,^{13–17} and enzymatic catalysts,^{18–20} achieving polymers with controlled molecular weights and narrow dispersities.²¹ However, in order for PCL to increase applicability in material and pharmaceutical sciences, it is essential to achieve precise control over its physical–chemical properties in an accessible manner. While copolymerisation with other lactones or physical blending with other polymers can be used to tailor pro-

perties of PCL formulations,^{22,23} the lack of functional pendant groups within the polymer backbone precludes further functionalisation, limiting the potential to engineer novel materials in a specific and precise fashion.

Post-polymerisation modifications represent an attractive strategy for the preparation of polymers bearing a diverse range of functionalities from a single monomer, enabling the introduction of functional moieties that are incompatible with ring-opening polymerisation conditions.^{24–26} In particular, the use of simple and robust thiol-based click chemistries has attracted considerable attention owing to their rapid reactivity, orthogonality and versatility of conjugation, allowing the incorporation of a wide range of substituents on the polymer backbone through reaction with terminal alkenes (thiol–ene)^{27–29} and terminal alkynes (thiol–yne).^{30,31} Moreover, the thiol–ene addition has been explored in materials chemistry for the preparation of elastomers,³² nanoparticles,³³ and tissue engineering scaffolds³⁴ exploiting the versatility of this chemistry to obtain different materials with tuneable properties. Notably, several studies have investigated the introduction of alkene functionalities on polyester backbones, however access to polyesters in which the semi-crystallinity of the native polyesters is retained is limited. For example, Hawker and co-workers reported the preparation of a series of alkene-functionalised polymers using thermal and photo-initiated thiol–ene chem-

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istry, comparing the efficiency and orthogonality of thiol-ene additions on PCL using a wide range of thiols, but the thermal properties of the resultant structures were not thoroughly investigated.³⁵ Heise and co-workers employed thiol-ene chemistry for the preparation of side-chain functional poly (macro)lactones, however functionalisation with primary alcohol and amide substituents disrupted the semi-crystalline structure of the native poly(macro)lactone, leading to amorphous materials.³⁶

Herein, we report a facile method for the preparation of post-functionalised PCL to provide lipophilic polyesters, which are semi-crystalline for some formulations containing sufficiently long alkyl side chains. The synthesis of the lipophilic polyesters was obtained *via* ring-opening polymerisation of an allyl-functional caprolactone monomer which was subsequently functionalised *via* thiol-ene addition using a range of commercially available alkyl thiols. Polymer solubility and thermal properties were evaluated and the effect of the alkyl chain length on these properties was investigated. This study presents a straightforward method to tune lipophilicity and crystallinity of PCL by post-polymerisation functionalisation in a simple and efficient manner.

Experimental section

Materials

All chemicals and solvents, unless otherwise stated, were purchased from Sigma-Aldrich, Fisher Scientific and Alfa Aesar and used without further purification. Silica gel (pore size 40 Å) was obtained from Fisher Scientific and used as received. Deuterated benzene was ordered from Apollo Scientific, degassed by repeated freeze-pump-thawing and then dried under reflux over sodium and benzophenone before storing under an inert atmosphere of N₂. Pentane was dried over CaH₂ for two days, distilled under N₂, degassed by freeze-pump-thawing (3 cycles) and stored in a N₂-filled glovebox. Dry toluene and dichloromethane were deoxygenated with N₂ and passed over an alumina column in a solvent tower delivery system. ϵ -Allyl- ϵ -caprolactone was synthesised as reported,³⁷ dried over CaH₂ for 24 h (and this was repeated 1× with fresh CaH₂), distilled under reduced pressure and stored under an inert atmosphere. 4-Methoxybenzyl alcohol was dissolved in dry dichloromethane and dried over CaH₂ for 24 h. Excess CaH₂ was removed by cannula filtration and the dichloromethane was distilled off under reduced pressure before storing the dry alcohol under an inert atmosphere of N₂. Diphenyl phosphate was dried under reduced pressure at 50 °C for 3 days. Irgacure 819 photoinitiator was obtained from BASF and stored in a light-free environment prior to use. Amberlyst A21 free base was washed repeatedly with methanol and dried under reduced pressure prior to use.

General considerations

All polymerisations were performed under an inert nitrogen atmosphere in a glovebox unless otherwise stated. All other

chemical manipulations were performed using standard Schlenk-line techniques. ¹H NMR and ¹³C NMR spectra were recorded on a Bruker DPX-300, DPX-400 or HD500 spectrometer. Chemical shifts are reported in ppm. Solvent residual signals were used as reference (CDCl₃, ¹H: δ = 7.26 ppm for CHCl₃, ¹³C: δ = 77.16 ppm for CDCl₃; C₆D₆, ¹H: δ = 7.16 ppm for C₆D₅H and ¹³C: δ = 128.06 ppm for C₆D₆). The signals are characterised as s = singlet, d = doublet, dd = doublet in doublet, ddd = doublet in doublet in doublet, t = triplet, q = quartet and m = multiplet. The coupling constant *J* is recorded in Hertz (Hz). Size exclusion chromatography (SEC) measurements were taken on an Agilent 1260 Infinity II Multi-Detector GPC/SEC System fitted with RI, ultraviolet (UV, λ = 309 nm), and viscometer detectors. The polymers were eluted through an Agilent guard column (PLGel 5 μ M, 50 × 7.5 mm) and two Agilent mixed-C columns (PLGel 5 μ M, 300 × 7.5 mm) using CHCl₃ (buffered with 0.5% v/v NEt₃) as the mobile phase (flow rate = 1 mL min⁻¹, 40 °C). Number average molecular weights (*M_n*), weight average molecular weights (*M_w*) and dispersities (*D_M* = *M_w*/*M_n*) were determined using Agilent GPC/SEC software (vA.02.01) against a 15-point calibration curve (*M_p* = 162–31 87 000 g mol⁻¹) based on poly(styrene) standards (Easivial PS-M/H, Agilent). Molecular weights (*M_n*), weight average molecular weights (*M_w*) are given in g mol⁻¹. Matrix-assisted laser desorption/ionisation time-of-flight mass spectrometry (MALDI-ToF MS) analysis was performed on a Bruker Ultraflex Extreme mass spectrometer using a nitrogen laser delivering 2 ns pulses at 337 nm with positive ion ToF detection performed using an accelerating voltage of 25 kV. *trans*-2-[3-(4-*tert*-Butylphenyl)-2-methyl-2-propenylidene] malononitrile (DCTB) was used as a matrix (40 g L⁻¹ solution in THF), with sodium iodide (NaI) used as a cationic agent (1 g L⁻¹ solution in THF). Analyte (1 g L⁻¹ solution in THF) was mixed with DCTB and NaI solutions (20 μ L of each) and 1 μ L was applied to form a thin matrix-analyte film. All samples were calibrated against a 3000 g mol⁻¹ poly(ethylene glycol) standard. The thermal characteristics of the polymers were determined using differential scanning calorimetry (DSC) on a STARE DSC3 system from Mettler Toledo and analysed in 40 μ L aluminium pans from -100–200 °C at a heating rate of 10 °C min⁻¹ for two heating/cooling cycles or 1 °C min⁻¹ for one heating/cooling cycle. UV-Vis spectroscopy data was collected on an Evolution 350 UV-Vis spectrophotometer equipped with a Xenon Flash Lamp light source and Dual Matched Silicon Photodiodes detector. Quartz cells (170–2000 nm) from Hellma with two polished sides were used for examining the absorption spectral data by using Thermo INSIGHT-2 v.10.0.30319.1 software. A thermostat and 8-cell Peltier system with precise temperature control between 0 °C and 90 °C was coupled with an Evolution 350 UV-Vis spectrophotometer to record the temperature-dependent transmittance spectra of the sample. For all measurements, *n*-dodecane was used as the solvent.

General procedure for synthesis of ϵ -allyl- ϵ -caprolactone (ϵ CL)

To a solution of 2-allylcyclohexan-1-one (5.00 g, 36.1 mmol) in dry CH₂Cl₂ (180 mL) was added NaHCO₃ (4.50 g, 54.3 mmol)



followed by *m*-CPBA (77%, 9.77 g, 43.6 mmol) at 0 °C. The resulting mixture was allowed to warm to room temperature and stirred for 48 h. The reaction was cooled to 0 °C, filtered over Celite and washed with 10% solution of sodium thiosulfate (3 × 100 mL), sodium hydrogen carbonate (3 × 100 mL), brine (3 × 100 mL) and dried over MgSO₄. The mixture was concentrated *in vacuo* and the crude product was purified by column chromatography using hexane/ethyl acetate (gradient elution, 4:1 to 1:1) to yield the desired product (3.88 g, 25.2 mmol, 70%) as a light-yellow oil. Analytical data is in agreement with literature:³⁷ ¹H NMR (400 MHz, CDCl₃, 299 K, ppm): δ = 5.82 (ddt, *J* = 17.2, 10.3, 7.1 Hz, 1H, CH=CH₂), 5.24–5.00 (m, 2H, CH=CH₂), 4.27 (dt, *J* = 8.6, 6.2 Hz, 1H, (CO)OCH), 2.75–2.51 (m, 2H, CH₂(CO)O), 2.46 (ddd, *J* = 14.1, 7.3, 5.8 Hz, 1H, CHCH₂CH=CH₂), 2.31 (dt, *J* = 14.0, 6.8 Hz, 1H, CHCH₂CH=CH₂), 2.01–1.80 (m, 3H, CH₂CH₂CH₂), 1.64–1.35 (m, 3H, CH₂CH₂CH₂); ¹³C NMR (101 MHz, CDCl₃, 299 K, ppm) δ = 175.9 ((CO)O), 133.4 (CH₂CH=CH₂), 118.2 (CH₂CH=CH₂), 79.9 ((CO)OCH), 40.6 (CH₂CH=CH₂), 34.9 (CH₂(CO)OCH), 33.8 ((CO)OCHCH₂), 28.3 (CHCH₂CH₂CH₂), 23.0 (CHCH₂CH₂CH₂).

General procedure for synthesis of poly(ε-allyl-ε-caprolactone) (AεCL)

Using standard glovebox techniques, a solution of Mg (BHT)₂(THF)₂ (293 mg, 0.48 mmol, 0.5 mol%) and 4-methoxy benzyl alcohol (26.0 mg, 0.19 mmol, 0.02 eq.) in dry toluene (1 mL) was added to a solution of ε-allyl-ε-caprolactone (1.50 g, 9.73 mmol) in dry toluene (8.70 mL). After stirring for 4 h at 50 °C, the ampoule was removed from the glove box and the reaction was quenched by the addition of two drops trifluoroacetic acid (TFA). The solution was precipitated into cold *n*-hexane (3×) and dried under vacuum to afford the polymer as a light-yellow oil. ¹H NMR and ¹³C NMR are in agreement with literature:³⁸ ¹H NMR (300 MHz, CDCl₃, 299 K, ppm): δ = 6.93–6.84 (m, 2H, CH_{aromatic}), 5.72 (ddt, *J* = 16.4, 10.5, 7.1 Hz, 49H, CH=CH₂), 5.18–5.01 (m, 109H, CH=CH₂), 4.91 (p, *J* = 6.2 Hz, 49H, (CO)OCH), 3.81 (s, 3H, OCH₃), 2.27 (m, 206H, CHCH₂CH=CH₂, (CO)OCH₂), 1.73–1.46 (m, 202H, CH₂CH₂CH₂), 1.44–1.14 (m, 98H, CH₂CH₂CH₂); ¹³C NMR (101 MHz, CDCl₃, 299 K, ppm): δ = 172.1 (O(CO)), 137.9 (CH=CH₂), 122.0 (CH=CH₂), 77.2 (OCHCH₂), 42.9 (CH₂CH=CH₂), 38.7 (O(CO)CH₂), 37.5 (OCHCH₂), 29.2 (O(CO)CH₂CH₂, O(CO)CH₂CH₂CH₂); SEC (CHCl₃): *M*_n = 12 800, *M*_w = 15 900, *D*_M = 1.24.

General procedure for radical thiol-ene click reaction

Poly(ε-allyl-ε-caprolactone) (500 mg, 59.0 μmol) was dissolved in chloroform (4 mL) and thiol (9.55 mmol, 3 equiv. per olefin) was added, followed by Irgacure 819 (134 mg, 0.32 mmol, 0.1 equiv. per olefin). The solution was irradiated with a UV lamp (315–400 nm) at room temperature for 2–50 min, precipitated into cold MeOH (3×) and dried under vacuum to obtain the thiol-functionalised poly(ε-allyl-ε-caprolactone):

Poly(ε-allyl-ε-caprolactone) functionalised with 2-ethyl hexanethiol: ¹H NMR (300 MHz, CDCl₃, 299 K, ppm): δ = 6.90–6.84

(m, 2H, CH_{aromatic}), 5.02 (s, 2H, CH_{aromatic}-CH₂), 4.85 (p, *J* = 6.0 Hz, 60H, (CO)OCH), 3.79 (s, 3H, OCH₃), 2.46 (dd, *J* = 7.7, 6.2 Hz, 263H, -CH₂SCH₂), 2.25 (t, *J* = 7.6 Hz, 133H, (CO)OCH₂), 1.88–1.15 (m, 1596H, CH₂), 0.87 (dt, *J* = 9.8, 6.7 Hz, 497H, CH₃); ¹³C NMR (101 MHz, CDCl₃, 299 K, ppm): δ = 173.0 (O(CO)), 73.1 (CHO), 39.4 (SCH₂CH), 36.5 (CH₂SCH₂CH), 34.2 ((CO)CH₂), 33.6 (CH₂SCH₂CH), 32.9 (CH₂), 32.4 (CH₂), 32.2 (CH₂), 28.6 (CH₂), 25.3 (CH₂), 25.2 (CH₂), 24.7 (CH₂), 22.8 (CH₂), 13.9 (CH₃), 10.6 (CH₃); SEC (CHCl₃): *M*_n = 16 400, *M*_w = 21 300, *D*_M = 1.30.

Poly(ε-allyl-ε-caprolactone) functionalised with 1-hexanethiol: ¹H NMR (300 MHz, CDCl₃, 299 K, ppm): δ = 6.82 (d, *J* = 8.7 Hz, 2H, CH_{aromatic}), 4.97 (s, 2H, CH_{aromatic}-CH₂), 4.87–4.70 (m, 55H, (CO)OCH), 3.74 (s, 3H, OCH₃), 2.42 (td, *J* = 7.3, 4.9 Hz, 242H, CH₂SCH₂), 2.21 (t, *J* = 7.6 Hz, 125H, COOCH₂), 1.67–1.37 (m, 550H, CH₂), 1.40–1.11 (m, 361H, CH₂), 0.88–0.72 (m, 144H, CH₃); ¹³C NMR (101 MHz, CDCl₃, 299 K, ppm): δ = 173.3 (O(CO)), 73.5 (CHO), 34.6 (CH₂SCH₂), 34.0 (CH₂SCH₂), 33.3 (CH₂), 32.3 (CH₂), 32.1 (CH₂), 31.6 (CH₂), 29.8 (CH₂), 28.8 (CH₂), 25.51 (CH₂), 25.1 (CH₂), 22.7 (CH₂), 14.2 (CH₃); SEC (CHCl₃): *M*_n = 14 700, *M*_w = 19 500, *D*_M = 1.33.

Poly(ε-allyl-ε-caprolactone) functionalised with 1-octanethiol: ¹H NMR (300 MHz, CDCl₃, 299 K, ppm): δ = 6.88 (d, *J* = 8.7 Hz, 2H, CH_{aromatic}), 5.03 (s, 2H, CH_{aromatic}-CH₂), 4.87 (p, *J* = 6.7 Hz, 57H, (CO)OCH), 3.80 (s, 3H, OCH₃), 2.57–2.42 (m, 244H, CH₂SCH₂), 2.27 (t, *J* = 7.6 Hz, 132H, (CO)OCH₂), 1.77–1.45 (m, 638H, CH₂), 1.44–1.16 (m, 795H, CH₂), 0.94–0.80 (m, 200H, CH₃); ¹³C NMR (101 MHz, CDCl₃, 299 K, ppm): δ = 173.6 (O(CO)), 73.8 (CHO), 34.8 (CH₂SCH₂), 34.3 (CH₂SCH₂), 33.6 (CH₂), 32.6 (CH₂), 32.3 (CH₂), 32.2 (CH₂), 30.1 (CH₂), 29.6 (CH₂), 29.6 (CH₂), 29.4 (CH₂), 25.8 (CH₂), 25.3 (CH₂), 23.1 (CH₂), 14.7 (CH₃); SEC (CHCl₃): *M*_n = 13 800, *M*_w = 18 500, *D*_M = 1.34.

Poly(ε-allyl-ε-caprolactone) functionalised with 1-decanethiol: ¹H NMR (300 MHz, CDCl₃, 299 K, ppm): δ = 4.87 (p, *J* = 6.1 Hz, 59H, (CO)OCH), 3.81 (s, 3H, OCH₃), 2.49 (td, *J* = 7.2, 5.2 Hz, 253H, CH₂SCH₂), 2.27 (t, *J* = 7.6 Hz, 131H, COOCH₂), 1.74–1.45 (m, 718H, CH₂), 1.27 (d, *J* = 4.9 Hz, 1207H, CH₂), 0.96–0.78 (m, 210H, CH₃); ¹³C NMR (101 MHz, CDCl₃, 299 K, ppm): δ = 173.6 (O(CO)), 73.9 (CHO), 34.5 (CH₂SCH₂), 34.0 (CH₂SCH₂), 32.3 (CH₂), 32.1 (CH₂), 29.9 (CH₂), 29.7 (CH₂), 29.5 (CH₂), 29.1 (CH₂), 25.5 (CH₂), 25.1 (CH₂), 22.8 (CH₃); SEC (CHCl₃): *M*_n = 17 500, *M*_w = 123 500, *D*_M = 1.34.

Poly(ε-allyl-ε-caprolactone) functionalised with 1-dodecanethiol: ¹H NMR (300 MHz, CDCl₃, 299 K, ppm): δ = 5.03 (s, 2H, CH_{aromatic}-CH₂), 4.86 (p, *J* = 6.1 Hz, 53H, (CO)OCH), 3.80 (s, 3H, OCH₃), 2.59–2.38 (m, 234H, CH₂SCH₂), 2.27 (t, *J* = 7.6 Hz, 125H, (CO)OCH₂), 1.84–1.11 (m, 1970H, CH₂), 1.00–0.76 (m, 266H, CH₃); ¹³C NMR (101 MHz, CDCl₃, 299 K, ppm): δ = 173.6 (O(CO)), 73.71 (CHO), 34.8 (CH₂SCH₂), 34.3 (CH₂SCH₂), 33.6 (CH₂), 32.6 (CH₂), 32.3 (CH₂), 30.1 (CH₂), 30.0 (CH₂), 29.8 (CH₂), 29.7 (CH₂), 29.6 (CH₂), 29.4 (CH₂), 28.9 (CH₂), 25.7 (CH₂), 25.3 (CH₂), 25.2 (CH₂), 23.1 (CH₂), 14.3 (CH₃); SEC (CHCl₃): *M*_n = 19 300, *M*_w = 24 800, *D*_M = 1.29.

Poly(ε-allyl-ε-caprolactone) functionalised with 1-hexadecanethiol: ¹H NMR (300 MHz, CDCl₃, 299 K, ppm): δ = 6.88 (d,



$J = 8.6$ Hz, 2H, $\text{CH}_{\text{aromatic}}$), 5.04 (s, 2H, $\text{CH}_{\text{aromatic}}\text{-CH}_2$), 4.88 (m, 56H, (CO)OCH), 3.81 (s, 3H, OCH_3), 2.58–2.41 (m, 229H, CH_2SCH_2), 2.27 (t, $J = 7.6$ Hz, 143H, (CO)OCH₂), 1.25 (s, 2900H, CH_2), 0.98–0.75 (m, 205H, CH_3); ^{13}C NMR (101 MHz, CDCl_3 , 299 K, ppm): $\delta = 173.2$ (O(CO)), 73.4 (CHO), 34.5 (CH_2SCH_2), 34.0 (CH_2SCH_2), 33.3 (CH_2), 32.3 (CH_2), 32.0 (CH_2), 29.8 (CH_2), 29.7 (CH_2), 29.5 (CH_2), 29.4 (CH_2), 29.3 (CH_2), 29.1 (CH_2), 28.6 (CH_2), 25.5 (CH_2), 25.1 (CH_2), 24.9 (CH_2), 22.8 (CH_2), 14.2 (CH_3); SEC (CHCl_3): $M_n = 22\,500$, $M_w = 30\,150$, $D_M = 1.34$.

Results and discussion

Monomer synthesis and polymerisation

The allyl-functionalised lactone, ϵ -allyl- ϵ -caprolactone (A ϵ PCL), was synthesised from the corresponding 2-allylcyclohexanone *via* Bayer-Villiger oxidation using *meta*-chloroperoxybenzoic acid (*m*-CPBA) as an oxidant (ESI, Fig. S1[†]). We initially screened the ROP of A ϵ PCL using 4-methoxybenzyl alcohol as initiator ($[M_0]/[I_0] = 50$) and DPP as a catalyst (10 mol%), at 25 °C. DPP was chosen as a catalyst since it was also found to efficiently catalyse the ROP of CL and afford polymers with high molecular weight and narrow dispersities.¹² The polymerisation was monitored by ^1H NMR spectroscopy using the disappearance of the diagnostic $\text{O}=\text{COCH}$ monomer resonance ($\delta = 4.2$ ppm) and the appearance of $\text{O}=\text{COCH}$ polymer resonance ($\delta = 4.9$ ppm). Under these reaction conditions, the polymerisation was sluggish, reaching a monomer conversion of 70% after 330 h (Table 1, entry 1 & ESI, Fig. S2[†]). The kinetics of the polymerisation could be slightly improved by

increasing the reaction temperature to 50 °C, affording a monomer conversion of 88% after 140 h (Table 2, entry 2). To further decrease the polymerisation time, we employed $\text{Mg}(\text{BHT})_2(\text{THF})_2$ as the catalyst, which has been previously reported to be highly active for the ROP of ϵ -caprolactone and other lactones.^{16,39,40} Similarly, $\text{Mg}(\text{BHT})_2(\text{THF})_2$ was found to be comparatively more active for the ROP of A ϵ PCL, leading to 95% monomer conversion in just 8 h, however the dispersity of the final polymer was slightly higher than the DPP-catalysed reactions (Table 1, entry 3). Nevertheless, the polymerisation kinetics displayed a linear relationship between M_n and conversion and low dispersity, indicative of a well-controlled polymerisation (Fig. 1).

Analysis of the polymer using ^1H NMR spectroscopy confirmed the PCL backbone with pendant allyl moieties ($\delta = 5.72$ and 5.18 ppm) and well-defined end-group signals ($\delta = 7.31$, 6.93 and 3.81 ppm) attributed to the 4-methoxybenzyl alcohol initiator. These were used to calculate molecular weight, $M_{n, \text{NMR}} = 8.1 \text{ kg mol}^{-1}$, which is close to the theoretical molecular weight, $M_{n, \text{theor}} = 7.9 \text{ kg mol}^{-1}$ (ESI Fig. S3[†]). Size exclusion chromatography (SEC) analysis of the polymer showed a higher dispersity ($D_M = 1.24$), compared to the DPP-catalysed reaction, and a larger than predicted molecular weight ($M_{n, \text{SEC}} = 12.8 \text{ kg mol}^{-1}$), which can be explained by poor compositional correlation between A ϵ PCL and the poly(styrene) standards (Fig. 2a).

To corroborate the end-group fidelity that was observed in ^1H NMR spectroscopic analysis, matrix-assisted laser ionisation time-of-flight mass spectrometry (MALDI-ToF MS) was performed on a short-chain A ϵ PCL sample (Degree of

Table 1 Ring-opening polymerisation of ϵ -allyl- ϵ -caprolactone catalysed by DPP and $\text{Mg}(\text{BHT})_2(\text{THF})_2$, respectively^a

Entry	Catalyst (mol%)	Reaction temperature (°C)	Reaction time (h)	Monomer conversion ^b (%)	M_n^c (kg mol ⁻¹)	D_M^c
1	DPP (10)	25	330	70	11.9	1.07
2	DPP (10)	50	141	88	11.8	1.10
3	$\text{Mg}(\text{BHT})_2(\text{THF})_2$ (5)	50	8	90	12.8	1.24

^a Polymerisations were performed in benzene- d_6 with $[M_0]/[I_0] = 50$ and $[M]_0 = 1 \text{ M}$, using 4-methoxybenzyl alcohol as initiator. ^b Determined by ^1H NMR spectroscopy. ^c Determined by SEC analysis in CHCl_3 (0.5% NEt_3), calibrated against polystyrene standards.

Table 2 Synthesis of thiol-functionalised poly(ϵ -allyl- ϵ -caprolactone)^a

Entry	Alkyl thiol	Irradiation time (min)	$M_{n, \text{theoretical}}$ (kg mol ⁻¹)	$M_{n, \text{NMR}}^b$ (kg mol ⁻¹)	$M_{n, \text{SEC}}^c$ (kg mol ⁻¹)	D_M^c	T_m^d (°C)	T_c^d (°C)	ΔH_m^e (J g ⁻¹)
1	1-Hexanethiol	50	15.1	15.1	14.7	1.33	—	—	—
2	1-Octanethiol	15	16.6	17.2	13.8	1.34	—	—	—
3	2-Ethyl hexanethiol	30	16.6	18.1	16.4	1.30	—	—	—
4	1-Decanethiol	15	18.1	19.4	17.5	1.34	−28	−43	−13.43
5	1-Dodecanethiol	5	19.3	19.0	19.3	1.29	4	−8	−31.76
6	1-Hexadecanethiol	2	22.8	23.2	22.5	1.34	35	29	−47.11

^a Thiol-ene addition of poly(ϵ -allyl- ϵ -caprolactone). (i) Irgacure 819 (0.1 eq. per olefin), alkyl-thiol (3 equiv. per olefin), chloroform, UV light ($\lambda = 315\text{--}400 \text{ nm}$), 22 °C. ^b Determined by integrating end-groups using ^1H NMR spectroscopy (CDCl_3 , 298 K). ^c Determined by SEC analysis in CHCl_3 (0.5% NEt_3) as eluent and calibrated against polystyrene standards. ^d T_m and T_c were determined by differential scanning calorimetry and was obtained from the 2nd heating cycle and cooling cycle, respectively. ^e Total enthalpy of melting (ΔH_m) was calculated by integration of the first-order endothermic transition.





Fig. 1 Number-average molecular weight (M_n) and dispersity ($D_M = M_w/M_n$) plotted against monomer conversion for the ROP of ϵ -allyl- ϵ -caprolactone using $Mg(BHT)_2(THF)_2$ and 4-methoxybenzyl alcohol as initiator ($[M_0]/[I_0] = 50$). Dispersity determined by SEC analysis in $CHCl_3$ (0.5% NEt_3), calibrated against polystyrene standards.

Polymerisation (DP) = 14). The distribution plot shows a primary distribution with a spacing of m/z 154.16 that is in excellent agreement with the repeating monomer unit. The signal at m/z 2318 represents a DP 14 polymer chain carrying a sodium ion and the 4-methoxybenzyl end-group (Fig. 2b).

Post-polymerisation functionalisation of poly(ϵ -allyl- ϵ -caprolactone)

Post-polymerisation functionalisation of A ϵ PCL was investigated *via* photoinitiated thiol–ene addition on the pendant allyl-functionality using a selection of commercially available alkyl thiols (Scheme 1, Table 2). To study the overall lipophilicity of the functionalised polyesters, thiols with a linear alkyl chain (6–16 carbon length) were primarily assessed. However, a branched thiol was also reacted with the polymer to compare the solubility and crystallinity to its linear analogue. After irradiation with UV light ($\lambda = 315$ –400 nm), initially for 15 min, the polymer was precipitated into methanol to afford a functionalised PCL (Scheme 1, Table 2).



Scheme 1 Synthesis of thiol-functionalised poly(ϵ -allyl- ϵ -caprolactone). Conditions: (i) IRGACURE 819, alkyl-thiol, chloroform, UV light (315–400 nm), RT.

Analysis of the thiol–ene reactions by 1H NMR spectroscopy indicated if a quantitative reaction had occurred by monitoring the disappearance of the allyl proton signals (ESI Fig. S4–S9†). While the branched 2-ethyl hexanethiol required an additional 15 min irradiation (30 min in total), the hexanethiol required a total of 50 min irradiation for complete conversion. Interestingly, analysis of the A ϵ PCL samples functionalised with 1-dodecane- and 1-hexadecanethiol observed an impurity around $\delta = 2.6$ ppm, which was likely formed as a consequence of reactions between the allyl groups on the polymer backbone. We therefore sought to optimise reaction conditions for the thiol–ene addition of 1-hexadecanethiol. Changing the solvent from $CHCl_3$ to THF or varying the thiol to alkene group ratio had little effect on the formation of reaction side products (ESI, Fig. S6–S9†). However, reducing the irradiation time ($t = 2$ min) significantly improved the purity of the functionalised polymer and still afforded complete conversion of A ϵ PCL (Table 2, entry 6). The functionalisation was also confirmed by SEC analysis where we observed a noticeable shift to higher molecular weight for the post-functionalised A ϵ PCLs compared to native polymer and without significant change in dispersity (Fig. 3a). Not surprisingly, the molecular weight of the functional polymers was positively correlated to the alkyl chain length of the thiol (for example, $M_{n, SEC} = 22.5$ kg mol^{-1} for C_{16} and $M_{n, SEC} = 14.7$ kg mol^{-1} for C_6) (Table 2). 1-Dodecane thiol-functionalisation was also complete at a lower irradiation time of only 5 min. These activities are in line with the more electron-rich thiols being more active for radical addition.⁴¹



Fig. 2 (a) Size exclusion chromatogram of the molecular weight distribution of A ϵ PCL using 4-methoxybenzyl alcohol as initiator ($[M_0]/[I_0] = 50$). Molecular weight was determined against poly(styrene) standards using $CHCl_3$ (0.5% NEt_3) as eluent. (b) MALDI ToF-MS of A ϵ PCL initiated by 4-methoxybenzyl alcohol.



Since the pendant-functionalised polyesters possess a crystallisable alkyl side chain, we investigated the thermal properties of the functional polyesters by differential scanning calorimetry (DSC) (Fig. 3b, ESI Fig. S10†). Even though the crystallinity is likely associated to the alkyl side chain, some of the bulk polymers were nonetheless semi-crystalline which is an atypical behaviour for any backbone functionalised PCL. While C₆-, C₈- and ethylhexanethiol functionalised PCL are completely amorphous, C₁₀-, C₁₂- and C₁₆-PCL exhibited a sharp first-order melting/crystallisation transition. The melting temperature increased with alkyl chain length, from *T_m* of −28 °C to 35 °C for C₁₀PCL and C₁₆PCL, respectively (Table 2, entries 4–6 & ESI Fig. S10†). The same trend was observed for the crystallisation temperatures (*T_c*), increasing from −43 °C to 29 °C for C₁₀PCL and C₁₆PCL, respectively (Fig. 3b). Unsurprisingly, the crystallinity increased with alkyl chain length, albeit a minimum length of C₁₀ carbon atoms was required to induce crystallinity. These findings are in agreement with previous reports, where cooperative organisation of alkyl pendant chains increase effective chain packing and overall order in the polymers to induce crystallinity.^{42,43} Being able to effectively tune polymer crystallisation by only varying the alkyl chain functionality demonstrates the possibility of customising the physical and mechanical properties of PCL, which is highly sought-after for a wide range of industrial applications.⁴⁴

Polymer solubility is an important aspect in allowing control over self-assembly behaviour, polymer microstructure and material performance.⁴⁵ While enhancing water solubility has been widely studied, increasing solubility in hydrophobic solvents is less well studied. Hence, we expected that the ability to quantitatively graft various alkyl thiols onto the PCL backbone would enable increased solubility in non-polar solvents compared to PCL. In order to study the effect of the alkyl chain

length on the solubility behaviour, polymer solutions prepared in *n*-dodecane (1–122 mg mL^{−1}) were analysed by UV-Vis spectroscopy using variable temperature sweeps (20–95 °C). A detection wavelength of 500 nm was chosen since the dissolved polymer does not show any absorption at this wavelength and hence a low transmittance would only indicate poor solubility (Fig. 4). The polyester functionalised with the shortest alkyl chain (C₆) showed poor solubility under 50 °C at all concentrations (Fig. 4a). Increasing the alkyl chain length by two carbon atoms to C₈ improved the solubility at low concentration, but solubility was still poor for 40 mg mL^{−1} and 100 mg mL^{−1}, only becoming soluble near 50 °C (Fig. 4b). The branched version (ethylhexanethiol) showed similar behaviour to the linear C₈ analogue with only a slightly improvement at 100 mg mL^{−1} which suggests that branching does not have a significant effect on solubility, which is surprising considering the greater disorder and poorer chain packing behaviour that would be expected as a consequence of the branching (Fig. 4c). The polyester functionalised with 1-decanethiol displayed full solubility at 1 mg mL^{−1} concentration, however it was only sparingly insoluble over the entire temperature range at higher concentrations (Fig. 4d). Such behaviour was unexpected, especially considering the good solubility of the C₆ and C₈ analogues at higher temperatures and as such we sought to probe this result further using dynamic light scattering (DLS). Initially, study of the 2-ethylhexane, C₈ and C₁₀ functionalised polymers in *n*-dodecane (at 40 mg mL^{−1}) at both 20 and 90 °C revealed that while 2-ethylhexane- and C₈-functionalised PCL solutions contain aggregates at 20 °C, these are greatly diminished at 90 °C (ESI, Fig. S11 and S12†). In contrast, the C₁₀-functionalised PCL solution contained aggregates at both 20 and 90 °C (ESI, Fig. S13†) and furthermore displayed a strong Tyndall effect (ESI, Fig. S14†) – observations that were repeated on *n*-octane and *n*-decane (ESI, Fig. S15 & 16†). These observations



Fig. 3 (a) Poly(ϵ -allyl- ϵ -caprolactone) functionalised with various thiols: 1-hexanethiol, 1-octanethiol, 2-ethyl hexanethiol, 1-decanethiol, 1-dodecanethiol and 1-hexadecanethiol, compared to poly(ϵ -allyl- ϵ -caprolactone). Size exclusion chromatogram with molecular weights determined against poly(styrene) standards using CHCl₃ (0.5% NEt₃) as eluent. (b) Differential scanning calorimetry (DSC) for 2nd cooling cycle (200 to −100 °C, 10 °C min^{−1}) showing crystallisation of the polymer.





Fig. 4 UV-Vis transmittance spectra as function of temperature and concentration in *n*-dodecane for poly(ϵ -allyl- ϵ -caprolactone) functionalised with (a) 1-hexanethiol, (b) 1-octanethiol, (c) 2-ethyl hexanethiol, (d) 1-decanethiol, (e) 1-dodecanethiol and (f) 1-hexadecanethiol (heating rate = 1 °C min⁻¹ from 20–95 °C).

are consistent with the UV-Vis solubility analyses and suggest that there is a delicate balance between solubility and aggregation behaviour (possibly due to favourable side chain/solvent interactions) that contribute to scattering and diminished transmittance.

For polymers with longer alkyl pendant chains (C₁₂ and C₁₆) good solubility in *n*-dodecane at all concentrations and temperatures was observed, as indicated by the high transmittance (Fig. 4e and f), most likely due to the enhanced solubility from increased alkyl chain length. This shows that functionalisation with alkyl substituents of different chain lengths can effectively tune the lipophilicity of PCL, with solubility in *n*-dodecane improving with alkyl chain length.

Conclusions

The ring-opening polymerisation of an allyl-functionalised ϵ -caprolactone to afford PCL with a pendant allyl group was explored using DPP and Mg(BHT)₂(THF)₂ as catalysts. The latter catalyst was more active and provided good control over molecular weight as indicated by the linear growth of the polymer chains and low dispersity. Subsequent post-polymerisation functionalisation of the pendant allyl motif by photo-initiated thiol-ene addition showed this to be an effective and facile method for the preparation of a range of alkyl functionalised polyesters with divergent physicochemical properties. As the alkyl chain length increased on the PCL backbone, we observed increasing solubility in *n*-dodecane and the polymers

became semi-crystalline when the alkyl length reached 10 carbon atoms in length which is likely a consequence of effective chain interactions among the alkyl pendant groups. These findings demonstrate the potential to influence and tune the physical and thermal properties of PCL by simply varying the alkyl functionality, opening a new route for the development of novel materials in a simple and precise fashion. In particular, this affords a method to tailor polymer microstructure and solubility⁴³ which should be important for exploring PCL in new nanostructures and (bio)materials applications.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- 1 R. A. Gross and B. Kalra, *Science*, 2002, **297**, 803–807.
- 2 V. R. Sinha, K. Bansal, R. Kaushik, R. Kumria and A. Trehan, *Int. J. Pharmacol.*, 2004, **278**, 1–23.
- 3 D. R. Chen, J. Z. Bei and S. G. Wang, *Polym. Degrad. Stab.*, 2000, **67**, 455–459.



- 4 P. Joshi and G. Madras, *Polym. Degrad. Stab.*, 2008, **93**, 1901–1908.
- 5 D. W. Hutmacher, T. Schantz, I. Zein, K. W. Ng, S. H. Teoh and K. C. Tan, *J. Biomed. Mater. Res.*, 2001, **55**, 203–216.
- 6 M. C. Arno, M. Inam, Z. Coe, G. Cambridge, L. J. Macdougall, R. Keogh, A. P. Dove and R. K. O'Reilly, *J. Am. Chem. Soc.*, 2017, **139**, 16980–16985.
- 7 M. J. Jenkins, K. L. Harrison, M. M. C. G. Silva, M. J. Whitaker, K. M. Shakesheff and S. M. Howdle, *Eur. Polym. J.*, 2006, **42**, 3145–3151.
- 8 Y. Qian, Z. Yao, X. Wang, Y. Cheng, Z. Fang, W.-E. Yuan, C. Fan and Y. Ouyang, *Cell Proliferation*, 2020, **53**, e12730.
- 9 B. G. G. Lohmeijer, R. C. Pratt, F. Leibfarth, J. W. Logan, D. A. Long, A. P. Dove, F. Nederberg, J. Choi, C. Wade, R. M. Waymouth and J. L. Hedrick, *Macromolecules*, 2006, **39**, 8574–8583.
- 10 R. C. Pratt, B. G. G. Lohmeijer, D. A. Long, R. M. Waymouth and J. L. Hedrick, *J. Am. Chem. Soc.*, 2006, **128**, 4556–4557.
- 11 P. V. Persson, J. Schröder, K. Wickholm, E. Hedenström and T. Iversen, *Macromolecules*, 2004, **37**, 5889–5893.
- 12 K. Makiguchi, T. Satoh and T. Kakuchi, *Macromolecules*, 2011, **44**, 1999–2005.
- 13 L. F. Sánchez-Barba, A. Garcés, M. Fajardo, C. Alonso-Moreno, J. Fernández-Baeza, A. Otero, A. Antiñolo, J. Tejada, A. Lara-Sánchez and M. I. López-Solera, *Organometallics*, 2007, **26**, 6403–6411.
- 14 E. Martin, P. Dubois and R. Jérôme, *Macromolecules*, 2003, **36**, 7094–7099.
- 15 A. Kowalski, J. Libiszowski, T. Biela, M. Cypryk, A. Duda and S. Penczek, *Macromolecules*, 2005, **38**, 8170–8176.
- 16 H.-J. Fang, P.-S. Lai, J.-Y. Chen, S. C. N. Hsu, W.-D. Peng, S.-W. Ou, Y.-C. Lai, Y.-J. Chen, H. Chung, Y. Chen, T.-C. Huang, B.-S. Wu and H.-Y. Chen, *J. Polym. Sci., Part A: Polym. Chem.*, 2012, **50**, 2697–2704.
- 17 X. Wang, A. Thevenon, J. L. Brosmer, I. Yu, S. I. Khan, P. Mehrkhodavandi and P. L. Diaconescu, *J. Am. Chem. Soc.*, 2014, **136**, 11264–11267.
- 18 C. Bao, X. Xu, J. Chen and Q. Zhang, *Polym. Chem.*, 2020, **11**, 682–686.
- 19 S. Kobayashi, H. Uyama and M. Ohmae, *Bull. Chem. Soc. Jpn.*, 2001, **74**, 613–635.
- 20 J. Engel, A. Cordellier, L. Huang and S. Kara, *ChemCatChem*, 2019, **11**, 4983–4997.
- 21 O. Nuyken and D. S. Pask, *Polymers*, 2013, **5**, 361–403.
- 22 Q. Liu, S. Yuan, Y. Guo, A. Narayanan, C. Peng, S. Wang, T. Miyoshi and A. Joy, *Polym. Chem.*, 2019, **10**, 2579–2588.
- 23 A. Sarasam and S. V. Madihally, *Biomaterials*, 2005, **26**, 5500–5508.
- 24 R. J. Williams, I. A. Barker, R. K. O'Reilly and A. P. Dove, *ACS Macro Lett.*, 2012, **1**, 1285–1290.
- 25 A. W. Thomas and A. P. Dove, *Macromol. Biosci.*, 2016, **16**, 1762–1775.
- 26 S. Tempelaar, L. Mespouille, O. Coulembier, P. Dubois and A. P. Dove, *Chem. Soc. Rev.*, 2013, **42**, 1312–1336.
- 27 A. W. Thomas, P. K. Kuroishi, M. M. Pérez-Madrigal, A. K. Whittaker and A. P. Dove, *Polym. Chem.*, 2017, **8**, 5082–5090.
- 28 J. W. Chan, C. E. Hoyle, A. B. Lowe and M. Bowman, *Macromolecules*, 2010, **43**, 6381–6388.
- 29 B. D. Fairbanks, D. M. Love and C. N. Bowman, *Macromol. Chem. Phys.*, 2017, **218**, 1700073.
- 30 J. W. Chan, J. Shin, C. E. Hoyle, C. N. Bowman and A. B. Lowe, *Macromolecules*, 2010, **43**, 4937–4942.
- 31 D. Konkolewicz, A. Gray-Weale and S. Perrier, *J. Am. Chem. Soc.*, 2009, **131**, 18075–18077.
- 32 B. Nottet, G. Tambutet, Y. Bakkour and J. Coudane, *Polym. Chem.*, 2012, **3**, 2956–2963.
- 33 J. Zou, C. C. Hew, E. Themistou, Y. Li, C.-K. Chen, P. Alexandridis and C. Cheng, *Adv. Mater.*, 2011, **23**, 4274–4277.
- 34 Y. Zou, L. Zhang, L. Yang, F. Zhu, M. Ding, F. Lin, Z. Wang and Y. Li, *J. Controlled Release*, 2018, **273**, 160–179.
- 35 L. M. Campos, K. L. Killops, R. Sakai, J. M. J. Paulusse, D. Damiron, E. Drockenmuller, B. W. Messmore and C. J. Hawker, *Macromolecules*, 2008, **41**, 7063–7070.
- 36 Z. Ates, P. D. Thornton and A. Heise, *Polym. Chem.*, 2011, **2**, 309–312.
- 37 A. Chen, J. Xu, W. Chiang and C. L. L. Chai, *Tetrahedron*, 2010, **66**, 1489–1495.
- 38 D. Mecerreyes, R. D. Miller, J. L. Hedrick, C. Detrembleur and R. Jérôme, *J. Polym. Sci., Part A: Polym. Chem.*, 2000, **38**, 8700–8875.
- 39 J. A. Wilson, S. A. Hopkins, P. M. Wright and A. P. Dove, *Polym. Chem.*, 2014, **5**, 2691–2694.
- 40 J. A. Wilson, S. A. Hopkins, P. M. Wright and A. P. Dove, *Macromolecules*, 2015, **48**, 950–958.
- 41 C. E. Hoyle, A. B. Lowe and C. N. Bowman, *Chem. Soc. Rev.*, 2010, **39**, 1355–1387.
- 42 S. Zhou, Y. Zhao, Y. Cai, Y. Zhou, D. Wang, C. C. Han and D. Xu, *Polymer*, 2004, **45**, 6261–6268.
- 43 S. Mete, K. G. Goswami, E. Ksendzov, S. V. Kostjuk and P. De, *Polym. Chem.*, 2019, **10**, 6588–6599.
- 44 U. Kumar, T. Kato and J. M. J. Frechet, *J. Am. Chem. Soc.*, 1992, **114**, 6630–6639.
- 45 D. S. Boucher, *J. Appl. Polym. Sci.*, 2020, **137**, 48908.

