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## *iR* drop correction in electrocatalysis: everything one needs to know!

Sengeni Anantharaj \*<sup>ab</sup> and Suguru Noda <sup>ab</sup>

Direct current (DC) voltammetric techniques in which the potential is controlled while the response is measured as current are in use for decades now in the screening of electrocatalysts used in energy conversion devices, sensors, and electrolyzers despite being inferior in many ways to alternating current (AC) techniques and other advanced voltammetric techniques because of their simplicity to use and handle resultant data. One of the major issues with the controlled-potential DC techniques where the applied potential is controlled by means of a reference electrode is the potential drop at the interface. This drop in potential affects the obtained results to an inconceivably greater extent in some cases. Hence, this drop in potential is manually corrected by following Ohm's law. However, there exists a pile of miscalculated potential drop corrections in the literature because of the challenges in understanding them and the unawareness of the protocols that need to be followed. To help the beginners of this field, this article is dedicated to defining and elaborating a potential drop, the factors influencing it, commonly made mistakes, and best practices.

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### Introduction

The activity of every electrocatalyst used in energy conversion devices (batteries and fuel cells), electrolyzers (H<sub>2</sub>O, CO<sub>2</sub>, N<sub>2</sub>, etc.), and electrochemical sensors that involve electron transfer is determined using controlled-potential DC voltammetric techniques.<sup>1–4</sup> Among the variety of such techniques, linear sweep

voltammetry (LSV), cyclic voltammetry, and chronoamperometry (CA) are the most commonly used ones.<sup>5,6</sup> In sensors, pulse voltammetry and differential pulse voltammetry (DPV) are also used for analytes that have low detection limits and with inseparable (by LSV and CV) redox peaks.<sup>7</sup> Regardless of the technique used during the controlled-potential operation, there are numerous sources for potential drop at the WE-electrolyte interface.<sup>5,8</sup> The

<sup>a</sup>Department of Applied Chemistry, School of Advanced Science and Engineering, Waseda University, 3-4-1 Okubo, Shinjuku-ku, Tokyo 169-8555, Japan. E-mail: ananth@aoni.waseda.jp; anantharaj1402@gmail.com

<sup>b</sup>Waseda Research Institute for Science and Engineering, Waseda University, 3-4-1 Okubo, Shinjuku-ku, Tokyo 169-8555, Japan



Dr Sengeni Anantharaj obtained his PhD in 2018 from CSIR-CECRI, Karaikudi, Tamil Nadu, India, following which he worked with Prof. Noda as a JSPS postdoc fellow from January 2019 to March 2021. Currently, he is an Assistant Professor (Junior Researcher) at RISE, Waseda University since April 2021. He was the first 'ECS India Section S. K. Rangarajan Graduate Student Award'

winner of the year 2017 presented by The Electrochemical Society, USA. He places his interest on everything where there is a transfer of electrons resulting in the transformation of matter, especially at the electrode–electrolyte interface.



Prof. Suguru Noda received his PhD in 1999 from The University of Tokyo, Japan, and became an assistant professor and associate professor there. Then, joined Waseda University in 2012 as a full professor. He is a chemical engineer conducting research in the field of materials processes. He is recently focusing on the practical production of carbon and silicon nanomaterials, such as carbon

nanotubes and silicon films/nanoparticles, and applying these materials to energy and electronic devices including rechargeable batteries.



possible sources can be low electronic conductivity of the catalyst, the low electrolytic conductivity of the solution, the distance between the RE and WE, the resistance of the contacts, and the roughness of the electrode surface.<sup>7</sup> When the surface of the electrode is large and roughened in the DC mode, the passage of current from CE to WE initially leads the charging of the capacitor through the resistor in the circuit.<sup>9,10</sup> This charging takes up some time and results in a time lag before the interface can reach the desired potential. This time lag is called the RC time constant, which is mathematically defined as the time needed to charge the capacitor of the DC circuit to 63.2%.<sup>11</sup> Beyond this point, electron transfer pertaining to the catalysis of the reaction of interest occurs. Hence, an electrocatalyst is anticipated to have the lowest possible RC time constant as this is one of the main sources of potential drop.<sup>12</sup> Thus, the area of the double layer also contributes to the potential drop. In a three-electrode system, at the open circuit potential (OCP), the current that flows between RE and WE is usually in the range of nA or even lower. However, when the current passes at WE (*i.e.*, during operation), there is a significant current flow between WE and the tip of the RE. Hence, the electrolytic conductivity and the distance between RE and WE are the major contributors to the potential drop.<sup>7</sup> This infers that an inevitable drop in the potential is witnessed in all cases of DC operations in electrochemistry.

As a result, any activity evidenced as a function of the applied potential often has a slightly lower potential at the interface than the applied potential. Hence, to present the actual activity of an electrocatalyst<sup>13</sup> and to use the polarization curves for Tafel analysis,<sup>14</sup> the complete compensation for potential drops is done manually.<sup>5</sup> This drop in potential is also known as ohmic drop or

$iR_u$  drop. As the main sources of this drop are the low electrolytic conductivity and the distance between RE and WE, there have been several successful attempts to minimize the contribution of resistance from the solution ( $R_s$ ) to the overall uncompensated resistance ( $R_u$ ). Two of the most significant strategies ever used are the use of redox inactive supporting electrolytes to improve the electrolytic conductivity and the deployment of Luggin capillary.<sup>7</sup> The latter enables us to place the reference electrode (RE) at the closest possible distance from the working electrode (WE). Since there are also other aforementioned sources of potential drop, avoiding the  $iR_u$  drop completely is impossible and it must be corrected to present the intrinsic activity and to use the polarization curves for Tafel analyses. However, there are several challenges in correcting  $iR_u$  drop (Fig. 1) and there also exists a pile of incorrectly  $iR_u$  drop compensated polarization curves in the literature because of the unawareness of the protocols and the best practices to be followed. This article details what is  $iR_u$  drop? why should it be corrected? how should it be corrected? the challenges, commonly made mistakes, and the best practices.

## Factors influencing $iR_u$ drop

There are several things that can influence  $iR_u$  drop as mentioned briefly above. Here, we explain a significant few of them in detail.<sup>7,15</sup>

### Electrolytic conductivity

A lower conductivity of the electrolyte is apparently a prime contributor to the overall  $R_u$ . Improving electrolytic conductivity

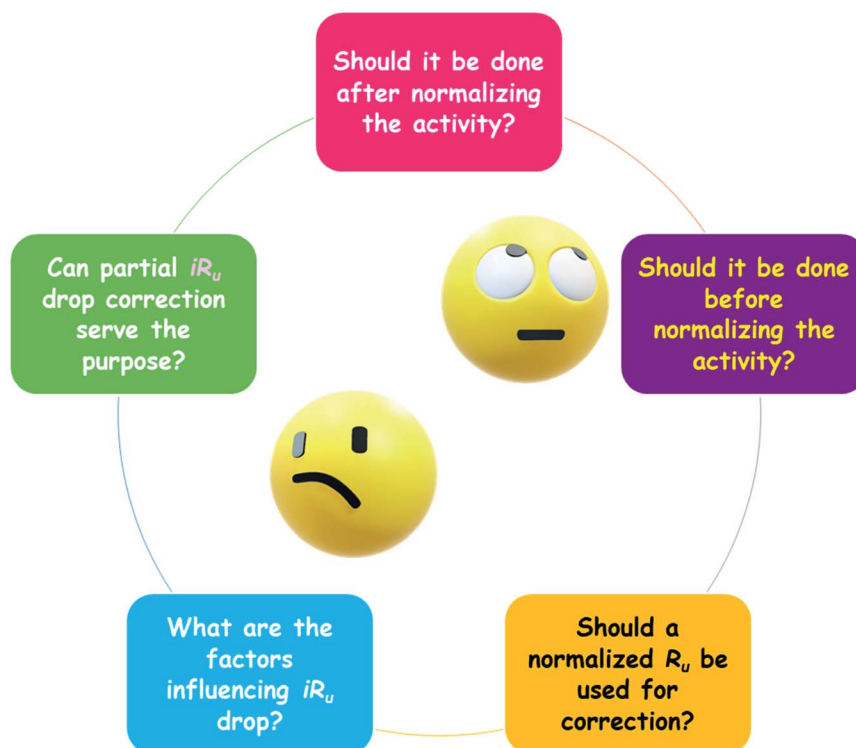


Fig. 1 Commonly encountered challenges in understanding and correcting  $iR_u$  drop in controlled-potential electrolysis.



is an utmost goal in all solution-based electrocatalysis. The most common way to improve the electrolytic conductivity is to add redox inactive salts. These salts are ideally expected to unalter the pH, and hence, should be formed by the neutralization of a strong acid and a strong base (e.g.  $\text{KNO}_3$ ). In some cases, buffers are used to tackle the issue of changing pH upon the addition of supporting electrolytes. If adding a supporting electrolyte is not an option, the alternative options to improve the electrolytic conductivity to a notable extent are stirring the solution, rotating the electrode, heating the solution, and applying ultrasonic waves. However, as water electrolysis and the electrolysis of  $\text{CO}_2$ ,  $\text{N}_2$ ,  $\text{O}_2$ , etc., are carried out in extreme pH conditions, the addition of supporting electrolytes or the other mentioned means are mostly not necessary and may only be needed for improving the selectivity and yield in  $\text{N}_2$  and  $\text{CO}_2$  electrolysis.<sup>16–18</sup> However, as the current of electrolysis is usually several orders of magnitude higher than what flows between the WE and the tip of the RE during operation, high electrolytic conductivity achieved at extreme pH conditions is also insufficient to completely avoid  $R_s$ , which is a major contributor to the overall  $R_u$ .

### The distance between RE and WE

Since the highest electrolytic conductivity that one can realize even at extreme pH conditions is not enough to alleviate  $R_s$  completely, the distance between RE and WE could have further influence in increasing the  $R_s$ . Hence, it is essential to place RE and WE as close as possible.<sup>19,20</sup> The Luggin capillary is a setup that allows one to bring the tip of the RE to a closer proximity to the surface of the WE. However, as a general rule of thumb, it is said that a RE of diameter  $d$  should be placed at a distance of  $2d$  from the surface of the WE. This may be suitable for highly sensitive sensor studies but not for electrolyzers that afford gaseous products and have a thicker double layer (in the cases of a large WE and WE with a high specific surface area (SSA)). Hence, the optimal distance between RE and WE should be determined by one's own discretion depending on the reaction, cell design, and the nature of the catalyst. At all costs, one should avoid touching the WE with the tip of the RE as it would cause the potential difference to cease to 0.0 V and mislead measurements. However, by compensating the  $iR_u$  drop by 100%, one can easily eliminate the effect of WE and RE distances on  $iR_u$  drop since  $R_s$  and the intrinsic resistance of the electrode (electrocatalyst-substrate electrode) contribute more than 90% to the overall  $R_u$ .

### Electronic conductivity of catalyst/electrode

A poor electronic conductivity can add up its own resistance to the overall  $R_u$ . Hence, it is important to have high electronic conductivity with WE. For catalytic electrodes with self-supported catalysts grown directly on the substrate electrodes, this issue is minimized to a greater extent. However, modified electrodes, such as catalyst ink-coated substrate electrodes, suffer from high intrinsic resistance. In such cases, ensuring the substrate to be a better current collector and the catalyst ink to be a better electronic conductor is necessary. In instances where an electrocatalyst suffers from poor electronic

conductivity, carbonization is the most commonly used strategy to improve its electronic conductivity. This works well with batteries and supercapacitors and even in some sensors too. However, carbon is not stable enough at high anodic potentials with which water electrolyzers and electrolyzers of  $\text{CO}_2$  and  $\text{N}_2$  work. In those cases, it is essential to modify the catalysts to have excellent electronic conductivity right from the synthesis/fabrication step.

### The roughness of WE and double-layer thickness

As stated earlier, the charging capacitor takes up some time in all DC circuits and the time required to charge the capacitor up to 63.2% of its total capacity *via* the resistor is known as the RC time constant.<sup>21</sup> The lesser the RC time constant, the faster the interface can reach the desired applied potential and can begin the charge transfer associated with the reaction of interest. As high surface roughness and large area can result in a thicker double-layer, the RC time constant can also be higher, thus leading to more  $iR_u$  drop. Hence, upon fabricating or synthesizing high surface area catalysts, one should be mindful of this factor as well. The factors governing the thickness of the double-layer are not only the surface roughness and active area but also the cell dimensions, the quantity of electrolyte, and  $A/V$  (area to volume ratio) values. The thickness of the double-layer can vary from a few nanometers to a few millimeters depending on the cell dimensions, the quantity of electrolyte, and  $A/V$  values. In general, a large area WE and a low volume of electrolyte always result in the formation of a thicker double-layer. These conditions are met with most thin-film cells. As for the electrolyzers that evolve gaseous products, such a high surface area could increase the resistance indirectly as a result of activity sites blocking by the gas bubbles formed at the surface. This issue is predominant with catalytic materials of high porosity such as 3D foam and felt-type catalytic electrodes.

By optimizing the electrolytic conductivity, the electronic conductivity, the distance between WE and RE, and the thickness of the double-layer, one can greatly minimize the  $iR_u$  drop. However, it is essentially impossible to completely avoid the  $iR_u$  drop with controlled-potential operations. Apart from these factors, corroded clips and other heterojunctions with their own interfacial potentials along the circuit also add significant resistance to the charge transfer. This makes the complete elimination of the  $iR_u$  drop a literally impossible task. Hence, one should learn to correct  $iR_u$  drop appropriately. As highlighted in Fig. 1, there are ambiguities among researchers in both understanding  $iR_u$  drop and correcting it. The following sections deal with these problems coherently.

## $iR_u$ drop correction

### Before or after normalizing the activity? or with normalized or as acquired $R_u$ ?

Errors in  $iR_u$  drop correction are usually made when one does not know when it is done and with which it needs to be done. Let's look the first case: mostly, in all energy conversion electrocatalytic reactions, the area of WE is usually less than  $1 \text{ cm}^2$





Fig. 2 Nyquist plot of the SS electrode ( $0.3 \text{ cm}^2$ ) acquired at  $1.674 \text{ V}$  vs. RHE from  $100 \text{ kHz}$  to  $0.1 \text{ Hz}$  with (red) and without (black) geometrical area normalization.

and hence, to get the apparent activity, the current response obtained is normalized by the geometrical surface area of the WE.<sup>22</sup> Similarly, the values of  $R_u$  obtained from electrochemical measurements are usually presented without normalizing by the geometrical area. In such instances, the as-acquired  $R_u$  value

without normalization is mostly used for  $iR_u$  drop correction, which is incorrect if the activity is already normalized. This often would result in abnormally corrected  $iR_u$  drop polarization curves. To show this, a stainless steel OER electrode of area  $0.3 \text{ cm}^2$  with stabilized activity (through extended potential cycling) is chosen as an example. The Nyquist plot of this electrode with and without geometrical area normalized impedance data points are shown in Fig. 2. It can be seen that after normalization, the  $R_u$  value decreased significantly.

Similarly, an LSV for examining the OER activity was also acquired. In Fig. 3a, as acquired activity is presented without normalizing using the geometrical area of the electrode in which the red datapoints are completely (*i.e.*, for 100%)  $iR_u$  drop corrected and the black ones are not. After the  $iR_u$  drop correction, the polarization curve is still valid and acceptable because both activity (*i.e.*, current) and  $R_u$  were not normalized. In cases where one wishes to use the as-acquired  $R_u$  for  $iR_u$  drop correction, the current should also not be normalized. After making  $iR_u$  drop correction using as acquired activity and as-acquired  $R_u$ , the activity can then be normalized using the geometrical area, as shown in Fig. 3c.

Fig. 3b shows another instance of error in  $iR_u$  drop correction, wherein one tries to correct the activity normalized polarization curve with the as-acquired  $R_u$  value. Since this type of error results in unacceptably  $iR_u$  drop corrected LSVs, as shown in Fig. 3b, people often perform partial  $iR_u$  drop

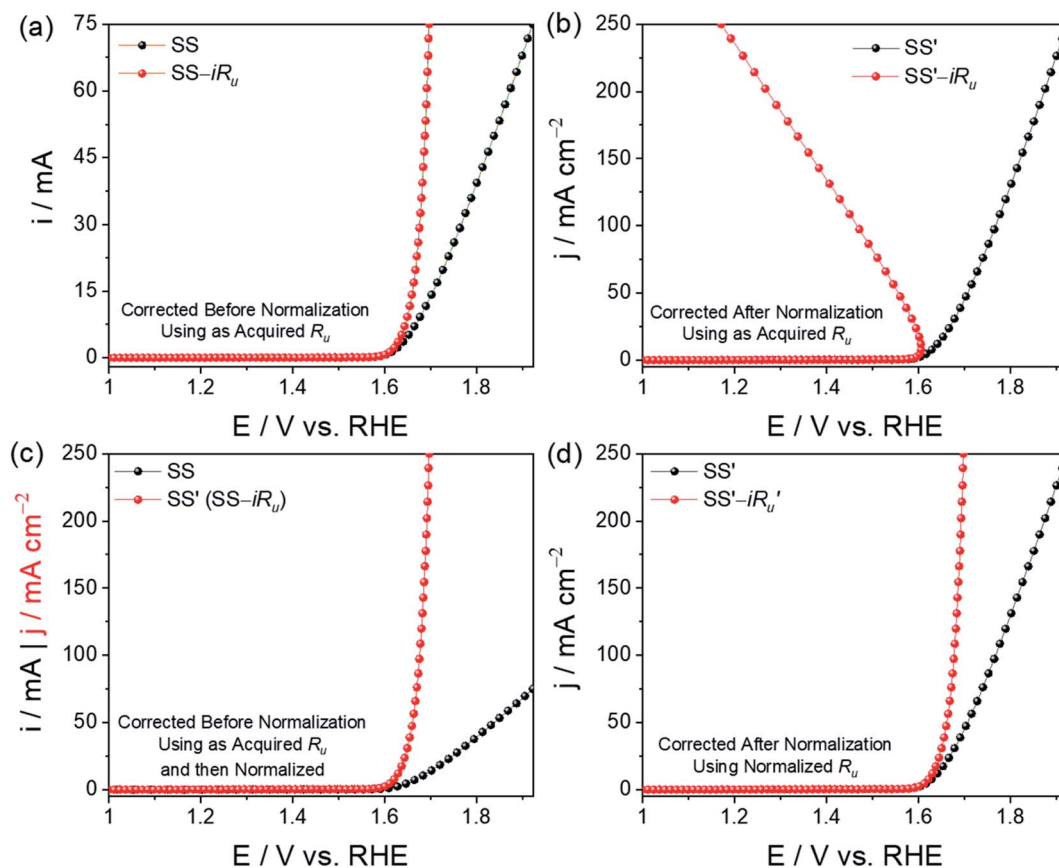


Fig. 3 (a–d) LSVs of the SS electrode with (red) and without (black)  $iR_u$  drop correction following different practices of  $iR_u$  drop correction.





correction in order to avoid this issue. However, this can simply be avoided if one does not use the as-acquired  $R_u$  for the  $iR_u$  drop correction of a normalized polarization curve. As seen from Fig. 3b, such practice would always result in unacceptable polarization curves with huge errors. If the electrode area is  $1 \text{ cm}^2$  or the activity obtained is already normalized by the system coupled with the workstation, it is essential to normalize the  $R_u$  too before it is used for  $iR_u$  drop correction. If not, the error witnessed in Fig. 3b will be encountered. Fig. 3d shows an acceptably  $iR_u$  drop-corrected polarization curve when both activity and  $R_u$  were normalized with the geometrical area. The correct way of correcting the  $iR_u$  drop is to either have both activity and  $R_u$  be normalized or not. If  $iR_u$  drop correction is done using the as-acquired activity and as-acquired  $R_u$  value, normalization can be done after  $iR_u$  drop correction without any error.

### Performing partial $iR_u$ drop correction even when 100% correction is possible!

Another error that is very commonly seen in the literature is the partial  $iR_u$  drop correction even when 100%  $iR_u$  drop correction is possible with the activity measured and  $R_u$  obtained. This has been followed by many without knowing why. In our recent works, we have stressed why partial  $iR_u$  drop correction is incorrect and up to what extent it could affect the results of Tafel analysis and overpotential determination.<sup>14</sup> The idea of  $iR_u$  drop correction is to report the activity free from the ohmic drop and to ensure that the Tafel analysis is done using completely  $iR_u$  drop-free potential and current responses. Performing partial  $iR_u$  drop correction is basically undermining the idea of Tafel analysis. Here, we show how partial  $iR_u$  drop correction would affect the polarization curve by taking the activity stabilized Ni foam in which both activity and  $R_u$  were normalized with their geometrical area (Fig. 4a and b). The inset of Fig. 4a is the normalized Nyquist plot of the Ni foam electrode from which the normalized  $R_u$  was taken for  $iR_u$  drop correction. When the normalized activity is corrected for  $iR_u$  drop for different percentages of normalized  $R_u$ , a change in overpotential can be seen at all current densities. Such changes will largely affect the obtained exchange current density values while using these partially  $iR_u$  drop-corrected LSVs. Hence, it is always essential to perform 100%  $iR_u$  drop correction when it is possible.

However, one can notice from Fig. 4b that if overpotential determination is what these LSVs are to be used for, an  $iR$  drop compensation ranging from 90% (sometimes even 85%) to 100% would be serving the purpose with no issues because the difference in overpotentials among the LSVs corrected for  $iR$  drop by 90–100% will be highly negligible.

### Challenges in $iR_u$ drop correction

Even with the protocols explained above, it may sometimes become impossible to make 100%  $iR_u$  drop corrections. In those cases, there is no other way than performing partial  $iR_u$  drop correction or not performing any at all. When either 100%  $iR_u$  drop correction is not possible or the  $iR_u$  drop

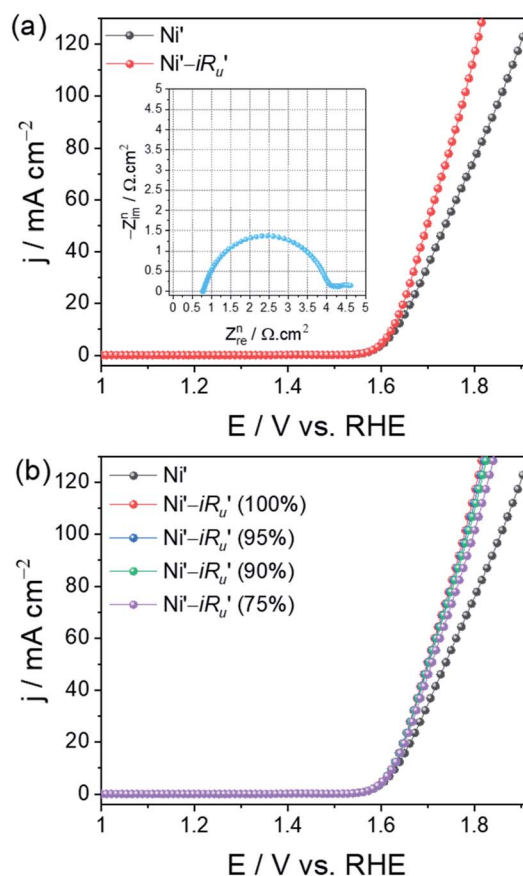


Fig. 4 (a) OER LSVs of the Ni foam electrode with and without  $iR_u$  drop correction. Inset of (a) is the normalized Nyquist plot of the same. (b) The same LSV with different percentages of  $iR_u$  drop correction.

correction is not done at all, one should never use such polarization curves for Tafel analysis.<sup>5</sup> For cases like these, we recommend the use of steady-state responses for Tafel analysis. Challenges in 100%  $iR_u$  drop correction occur when the activity is in the range of hundreds of milliamperes. In such cases, when the normalized  $R_u$  is as small as  $1 \text{ ohm cm}^2$ , large errors could occur. For example, consider a catalyst with an onset potential of  $1.5 \text{ V}$  delivering an activity of  $500 \text{ mA cm}^{-2}$  at  $1.8 \text{ V}$  and possesses a  $R_u$  of  $1 \text{ ohm cm}^2$ . The potential at  $500 \text{ mA cm}^{-2}$  after 100%  $iR_u$  drop correction would be  $1.3 \text{ V}$ , which is  $0.2 \text{ V}$  lesser than that of the onset potential. This is an unacceptable polarization curve. Similarly, when the value of  $R_u$  is higher, even for an electrocatalyst delivering just a few tens of milliamperes of activity, 100%  $iR_u$  drop correction will be impossible. For example, consider a catalyst with a normalized  $R_u$  of  $5 \text{ ohm cm}^2$  and an onset of  $1.5 \text{ V}$ , and an activity of  $80 \text{ mA cm}^{-2}$  at  $1.8 \text{ V}$ . The potential at  $80 \text{ mA cm}^{-2}$  would be  $1.4 \text{ V}$ , which is  $0.1 \text{ V}$  lesser than the onset potential, and hence, is not an acceptable polarization curve. In such cases, it is better to use steady-state responses obtained in a selective potential range for Tafel analysis and use galvanostatic measurements to find  $iR$  drop-free activity in which the potential is monitored as a function of applied current instead of not controlled.



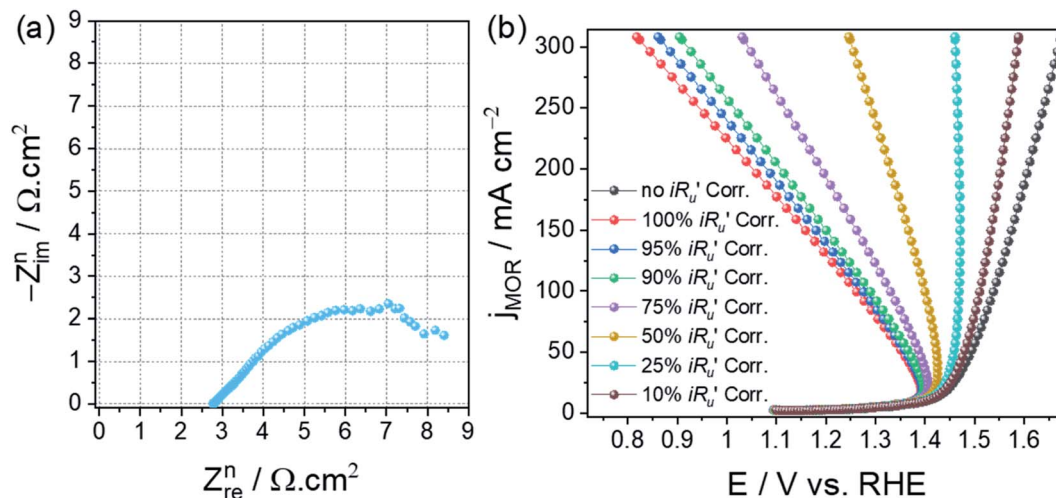
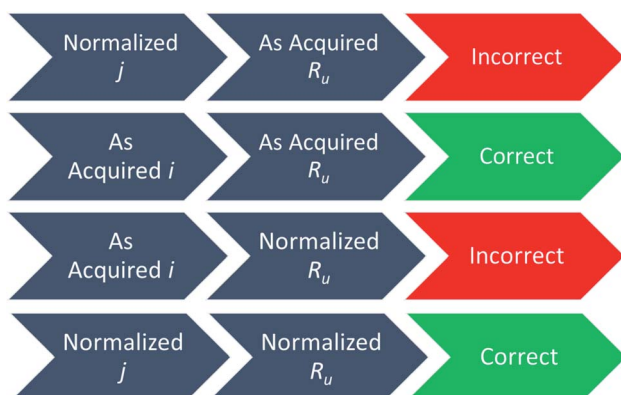


Fig. 5 (a) Normalized Nyquist plot of anodized Cu foam containing Ni impurities. (b) MOR LSVs of the same with different percentages of  $iR_u$  drop correction. Reproduced and modified from our earlier work (ref. 23) (Copyright 2021, American Chemical Society).

Fig. 5a shows the normalized Nyquist plot of anodized Cu foam with Ni impurity and Fig. 5b shows the methanol oxidation reaction (MOR) LSVs with different percentages of  $iR_u$  drop correction.<sup>23</sup> This is an example of the first kind mentioned above. This catalyst delivers high activity but also has a relatively higher  $R_u$  value. As a result, even a 50%  $iR_u$  drop compensation results in an unacceptable polarization curve. Hence, one should decide on his/her own discretion whether to correct their polarization curve for  $iR_u$  drop or not. If not they must opt for other electroanalytical tools (such as steady-state chronoamperometry and chronopotentiometry) that can avoid this issue. If one deems that 100%  $iR_u$  drop correction is possible for their catalyst with no issues, one should always go for 100%  $iR_u$  drop correction especially when the same polarization curve is to be used for Tafel analysis.

### Best practices in $iR_u$ drop correction

Since  $iR_u$  drop correction is an essential operation in all controlled-potential electrocatalytic conversions, it is important to ensure that it is done appropriately to avoid the



Scheme 1 Correct and incorrect ways of handling normalized and as-acquired activity and  $R_u$  data in  $iR_u$  drop correction.

miscalculated data in the literature. There are two instances where mistakes in  $iR_u$  drop correction can occur: (1) when confused with which (normalized or as acquired data)  $iR_u$  drop correction is to be done. And (2) performing partial  $iR_u$  drop correction even when 100%  $iR_u$  drop correction is possible. The second can be avoided by always remembering to perform 100%  $iR_u$  drop correction. The first issue can be avoided by using both activity and  $R_u$  before or after normalization. When the  $iR_u$  drop correction is done using the as acquired data, the normalization of activity is to be done afterwards to determine overpotential at benchmarking normalized activity. Scheme 1 shows the correct and incorrect ways of handling normalized and as-acquired activity and  $R_u$  data to appropriately perform  $iR_u$  drop correction.

### Summary

Wherever there is a control of applied potential in an electrolysis process and the response of the interface is monitored in terms of current passing through it, there is always a drop in potential at the interface known familiarly as the  $iR_u$  drop because of the uncompensated resistance  $R_u$  that exists in all DC circuits. Since most of the polarization curves obtained using voltammetric techniques suffer from the  $iR_u$  drop, it is essential to correct them before they can be inferred and used for further analysis (such as in Tafel analysis). However, there exist several ambiguities in understanding the  $iR_u$  drop and the practices of correcting it appropriately. This perspective defines what  $iR_u$  drop is, elaborates the factors influencing  $iR_u$  drop, highlights common errors made in  $iR_u$  drop correction, explains existing challenges in  $iR_u$  drop correction, and also suggests the best practices for performing appropriate  $iR_u$  drop correction. With the insights provided in this perspective, it is expected that the clouding miscalculated (for  $iR_u$  drop correction) polarization curves in the literature will be greatly minimized and the accuracy of data in the field of energy conversion electrocatalysis will be improved.



## Conflicts of interest

There are no conflicts to declare.

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## References

- H. Yang, X. Wang, Q. Hu, X. Chai, X. Ren, Q. Zhang, J. Liu and C. He, *Small Methods*, 2020, **4**, 1900826.
- S. Anantharaj, S. Noda, V. R. R. Jothi, S. C. Yi, M. Driess and P. W. Menezes, *Angew. Chem., Int. Ed.*, 2021, **60**, 18981–19006.
- J. Wang, S. Chen, Z. Li, G. Li and X. Liu, *ChemElectroChem*, 2020, **7**, 1067–1079.
- S. C. Perry, D. Pangotra, L. Vieira, L. I. Csepei, V. Sieber, L. Wang, C. Ponce de León and F. C. Walsh, *Nat. Rev. Chem.*, 2019, **3**, 442–458.
- S. Anantharaj, S. R. Ede, K. Karthick, S. Sam Sankar, K. Sangeetha, P. E. Karthik, S. Kundu, E. K. Pitchiah and S. Kundu, *Energy Environ. Sci.*, 2018, **11**, 744–771.
- S. Anantharaj, S. Kundu and S. Noda, *J. Electrochem. Soc.*, 2022, **169**, 014508.
- A. J. Bard and L. R. Faulkner, *Electrochemical Methods : Fundamentals and Applications*, Wiley, 2001.
- L. Yu and Z. Ren, *Mater. Today Phys.*, 2020, **14**, 100253.
- X. Meng, C. Yu, X. Song, J. Iocozzia, J. Hong, M. Rager, H. Jin, S. Wang, L. Huang, J. Qiu and Z. Lin, *Angew. Chem., Int. Ed.*, 2018, **57**, 4682–4686.
- Y. Du, X. Meng, Z. Wang, X. Zhao and J. Qiu, *Acta Phys.-Chim. Sin.*, 2021, **38**, 2101009.
- S. Anantharaj and S. Noda, *ChemElectroChem*, 2020, **7**, 2297–2308.
- J. O. Bockris, A. K. N. Reddy and M. E. Gamboa-Aldeco, *Modern Electrochemistry*, Plenum Press, 1998.
- S. Anantharaj, S. R. Ede, K. Sakthikumar, K. Karthick, S. Mishra and S. Kundu, *ACS Catal.*, 2016, **6**, 8069–8097.
- S. Anantharaj, S. Noda, M. Driess and P. W. Menezes, *ACS Energy Lett.*, 2021, **6**, 1607–1611.
- A. M. Bond, R. G. Compton, D. A. Fiedler, G. Inzelt, H. Kahlert, H. Lohse, M. Lovric, S. K. Lovric, F. Marken, A. Neudeck, U. Retter, F. Scholz and Z. Stojek, *Electroanalytical Methods Guide to Experiments and Applications 2nd*, revised and extended edition, 2009.
- M. König, J. Vaes, E. Klemm and D. Pant, *iScience*, 2019, **19**, 135–160.
- G. Chen, S. Ren, L. Zhang, H. Cheng, Y. Luo, K. Zhu, L. Ding and H. Wang, *Small Methods*, 2019, **3**, 1800337.
- R. Chen, *ChemElectroChem*, 2019, **6**, 603–612.
- F. Zhang, J. Liu, I. Ivanov, M. C. Hatzell, W. Yang, Y. Ahn and B. E. Logan, *Biotechnol. Bioeng.*, 2014, **111**, 1931–1939.
- A. Battistel, M. Fan, J. Stojadinović and F. La Mantia, *Electrochim. Acta*, 2014, **135**, 133–138.
- K. B. Oldham, *Electrochem. Commun.*, 2004, **6**, 210–214.
- S. Anantharaj and S. Kundu, *ACS Energy Lett.*, 2019, **4**, 1260–1264.
- S. Anantharaj, T. Nagamatsu, S. Yamaoka, M. Li and S. Noda, *ACS Appl. Energy Mater.*, 2021, **5**, 419–429.

