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Semi-biological approaches to solar-to-chemical conversion

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This review presents a comprehensive summary of the recent development in semi-artificial photosynthesis, a biological-material hybrid approach to solar-to-chemical conversion that provides new concepts to shape a sustainable future fuelled by solar energy. We begin with a brief introduction to natural and artificial photosynthesis, followed by a discussion of the motivation and rationale behind semi-artificial photosynthesis. Then, we summarise how various enzymes can be combined with synthetic materials for light-driven water oxidation, H2 evolution, CO₂ reduction, and chemical synthesis more broadly. In the following section, we discuss the strategies to incorporate microorganisms in photocatalytic and (photo)electrochemical systems to produce fuels and chemicals with renewable sources. Finally, we outline emerging analytical techniques to study the bio-material hybrid systems and propose unexplored research opportunities in the field of semi-artificial photosynthesis.

1 Introduction

The consequences of anthropogenic carbon emissions call for innovative strategies to develop renewable energy technologies. Photovoltaics (PV) is the leading technology for solar energy conversion, 1,2 but it shows disadvantages in intermittency and long-distance electricity transmission. Artificial photosynthesis is a process that converts solar energy into fuels and it thereby circumvents these drawbacks by storing solar energy in chemical bonds using synthetic light absorbers and catalysts.3,4 While synthetic

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catalysts still encounter challenges in solar-to-chemical conversion, nature provides evolutionarily-optimised biocatalysts. This review summarises how this biological machinery can be leveraged to catalyse light-driven reactions by an emerging technology termed "semi-artificial photosynthesis", in which biocatalysts in the form of enzymes or microorganisms are incorporated with particles or structurally-crafted electrodes to produce fuels or chemicals.

1.1 Natural photosynthesis

Photosynthesis occurs in photoautotrophs such as cyanobacteria, algae and higher plants, which harvest solar energy to produce biomass and O₂ from CO₂ and H₂O. Photosynthesis is accomplished



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Xin Fang obtained his PhD degree from the University of Cambridge under the supervision of Prof. Erwin Reisner. His doctoral research focused on establishing efficient solar-to-chemical conversion avenues with biocatalysts in the form of enzymes and whole cells. His research broadly traverses the fields of biocatalysis, materials science, electrochemistry, and photocatalysis, with the aim to shift the paradigm of artificial photosynthesis by rationally integrating

biocatalysts with synthetic components. Before that, he studied at Fudan University and Sichuan University for a Master and Bachelor's degree in polymer science, respectively.



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by two phases of reactions: the light reaction uses solar energy to generate the energy carrier adenosine triphosphate (ATP) and the reduced nicotinamide adenine dinucleotide phosphate (NADPH) as the reducing agent; the dark reaction then uses the ATP and NADPH to reduce atmospheric CO2 into carbohydrates through the Calvin cycle.5

1.1.1 Light reactions. The light reaction is carried out in the thylakoid membrane by the photosynthetic apparatus, including photosystem II (PSII) and photosystem I (PSI), cytochrome $b_6 f(\text{cyt } b_6 f)$, the plastoquinone (PQ) pool, plastocyanin (PC), ferredoxin (Fd), ferredoxin-NADP⁺ reductase (FNR) and ATP synthase (ATPase) (Fig. 1a).6 The two light absorbing photosystems both contain antenna complexes and reaction centres. The antenna complex comprises hundreds of pigments (chlorophylls and carotenoids), which capture photons and funnel the light energy to the reaction centre at a high quantum efficiency of >95%. The reaction centre is a transmembrane protein-pigment complex, where electrons are excited and transferred to the downstream chemical synthesis pathway. The reaction centre chlorophylls in PSII and PSI are referred to as P680 and P700, respectively. The two photosystems function in concert to efficiently transfer electrons from H2O to NADP through a "Z-scheme" (Fig. 1b). The electron transfer in the thylakoid membrane is initiated by PSII, which absorbs photons and extracts electrons from H₂O to reduce its terminal electron acceptor, plastoquinone B (Q_B) . The fully reduced plastohydroquinone (Q_BH₂) dissociates from the reaction centre and transfers its electrons to cyt $b_6 f$, which forwards the electrons to PSI via a small copper-containing protein PC. At PSI, the arriving electrons are energised by a second excitation and finally delivered to FNR to reduce NADP+ to NADPH (Fig. 1a).⁵ Simultaneously, the water oxidation and the electron transport induce proton translocation from the stroma to the lumen, which generates a proton gradient and chemiosmosis driving the synthesis of ATP by ATPase.



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biology, synthetic chemistry, materials science and engineering relevant to the development of solar-driven processes for the sustainable synthesis of fuels and chemicals.

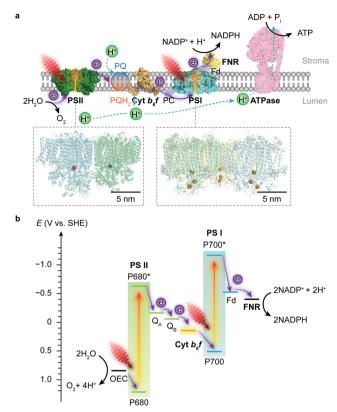


Fig. 1 Schematic illustration of the light reaction in natural photosynthesis. (a) Electron and proton transfer pathways in the thylakoid membrane. At the start of the photosynthetic chain is PSII that oxidises H_2O and releases O_2 and protons upon light absorption. The electrons are then delivered to PSI via a plastoquinone (PQ) pool, cyt b_6f and plastocyanin (PC). Electrons are photoexcited for a second time at PSI to reduce NADP+ to NADPH via a ferredoxin (Fd) and a ferredoxin-NADP+ reductase (FNR). The water oxidation and electron transport also induce proton translocation from the stroma to the lumen, which generates a proton gradient and chemiosmosis driving the synthesis of ATP by ATP synthase (ATPase). Protein data bank ID: PSII (4ub6), cyt b₆f (4h44), PC (1bxu), PSI (5oy0), Fd-FNR (2yvj), ATPase (6fkf). (b) Energy level diagram of the Z-scheme electron transfer in the thylakoid membrane. The redox carriers are placed at their midpoint redox potentials at pH 7.0

1.1.2 Dark reactions. The NADPH and ATP produced by the light reaction are then used to reduce CO2 to carbohydrates in the Calvin cycle. During the process, CO2 and H2O are combined with the carbon acceptor ribulose-1,5-bisphosphate to yield two molecules of 3-phosphoglycerate. This intermediate is then reduced to glyceraldehyde-3-phosphate with ATP and NADPH. The cycle closes with the regeneration of ribulose-1,5bisphosphate, the initial CO₂ acceptor. The Calvin cycle accounts for more than 90% of carbon assimilation on earth. Central to this reaction is ribulose-1,5-bisphosphate carboxylase/oxygenase (Rubisco) that covalently fixes CO₂ to a carbon skeleton.^{7,8} Rubisco is abundant in nature and operates using atmospheric CO2 concentrations (\sim 415 ppm). However, it also displays slow catalysis and low selectivity. The turnover frequency (TOF) of Rubisco is typically less than 10 s⁻¹, rendering the photosynthetic CO₂ fixation inefficient under optimal conditions.8 It also confuses the substrate (CO₂) and the product (O₂) of photosynthesis,

which offsets the CO₂ uptake and saddles oxygenic phototrophs with energy-dissipating photorespiration.⁹

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Photorespiration can consume 25% of carbon photosynthetically fixed by C_3 plants and this carbon loss is more severe at higher temperature. Some tropical plants, namely C_4 plants, overcome the disadvantageous photorespiration by deploying an additional C_4 pathway to increase the local concentration of CO_2 . In these plants, phosphoenolpyruvate is first carboxylated into a four-carbon compound oxaloacetate, which undergoes further transformations and decarboxylation to release CO_2 . The released CO_2 is then scavenged by Rubisco and drawn into the Calvin cycle with minimal photorespiration.

1.1.3 Photosynthetic efficiency. Photosynthetic organisms have evolved highly efficient light-harvesting systems with a quantum efficiency of more than 90%. Nevertheless, the overall photosynthetic efficiency is low: the highest efficiency measured under field conditions is 2.9% for C₃ and 4.2% for C₄ plants and typical efficiencies are 0.2-1% for crop plants. 10,11 Several reasons are responsible for the low photosynthetic efficiency: 10,12 (1) As the pigments in photosystems and antennae are primarily sensitive to visible light ($\lambda = 400-700$ nm), photosystems can only intercept 48.7% of the incident solar energy; (2) light reflection and transmission due to weak absorption of green light ($\lambda = 500-600$ nm) cause at least 4.9% of energy loss; 13 (3) 6.6% of energy will be dissipated as heat during the transfer in antenna complexes, before reaching the red lightabsorbing chlorophylls in the reaction centres; (4) energy expenditure associated with carbohydrate synthesis accounts for 24.6% and 28.7% for C₃ and C₄ plants, respectively; (5) oxygenation and ensuing photorespiration cost extra 6.1% of energy in C₃ plants; (6) respiratory metabolism for maintenance and growth consumes a minimum of 1.9% (C3 plants) and 2.5% (C4 plants) of the entire energy input from solar irradiation. Summing all of these energy losses, the maximal theoretical solar-to-biomass conversion efficiencies are 4.6% for C₃ plants and 6.0% for C₄ plants (30 °C). 10,12

1.1.4 Photobiological H_2 production and N_2 fixation. Cyanobacteria and microalgae contain thylakoid membranes to carry out photosynthetic reactions, 14 and they can produce H_2 in anaerobic conditions by partially diverting photo-generated electrons from PSI towards H_2 -producing hydrogenases (H_2 ases). $^{15-17}$ However, inhibition of H_2 ase by evolved O_2 renders this process transient and economically unfeasible. In addition, electron transport from PSI to H_2 ase requires diffusional electron carriers such as ferredoxins, which impose a kinetic limit on the overall process. 18

Some cyanobacteria can also fix atmospheric N_2 using solar energy when nitrogen-containing substrates (N_3 , N_3 , etc.) are not available. However, nitrogenase (N_2 ase), the enzyme responsible for N_2 fixation, is sensitive to oxygen. Therefore, cyanobacteria have developed strategies to separate the oxygenic photosynthesis and N_2 fixation spatially (in different cells) or temporally (during the night) to protect the N_2 ase from irreversible damage.

1.2 Artificial photosynthesis

Artificial photosynthesis does not reproduce the exact reactions occurring in photoautotrophs, but exploits light absorbers and catalysts to produce fuels and chemicals with earth-abundant feedstock chemicals such as H_2O and CO_2 . A prototype reaction is light-driven water splitting, which produces H_2 as a suitable energy carrier to overcome the intermittency of solar irradiation. Solar water splitting has been under intense investigation over several decades, but it still faces challenges arising from solar light harvesting, catalyst stability, interfacial charge transfer and the water oxidation reaction. He four-electron water oxidation to oxygen involves multiple bond rearrangements and concerted proton release, and thus causes both a thermodynamic barrier and kinetic sluggishness that confront the existing catalytic chemistry.

Thermodynamics determines the suitability of light absorbers from an energetic perspective (Fig. 2), as illustrated by the water splitting reaction, where the energy band positions should straddle the electrochemical potentials of the half reactions, *i.e.*, hydrogen evolution reaction (HER) and oxygen evolution reaction (OER) (Fig. 3a). The band gap of the light absorbers should not be too wide to enable the visible light (400–800 nm, band gap: 1.53–3.10 eV) utilisation.²⁶ The thermodynamic redox potentials of light absorbers should also be properly aligned with respect to their energy band positions to prevent themselves from photo-induced corrosion.²⁷

Several semiconductors whose electronic structures suffice in theory for overall water splitting with visible light have been under investigation, such as C₃N₄, CdS and Y₂Ti₂O₅S₂.²² An alternative strategy is to align the band position of two light absorbers and catalyse HER and OER with the aid of electron mediators (Fig. 3b), 22,28 which resembles the photosynthetic electron transfer ("Z-scheme") in the thylakoid membrane (Fig. 1b). This two-step excitation system can expand the light absorption by harvesting lower-energy photons, 22,29 and can bring together materials that can only drive either HER or OER alone.30-32 Such tandem systems enable larger driving forces for water splitting, and permit higher theoretical solar-tohydrogen conversion efficiencies than the single light-absorber systems.33 Yet two light absorbers also multiply chances of charge recombination and pose challenges to balance the electron transfer kinetics of the half reactions.3,22

The implementation of artificial photosynthesis is often envisioned in two forms: first, light absorbers are loaded with electrocatalysts and then suspended in a photoreactor (Fig. 3c).³⁴ Second, solar fuel synthesis is performed in a photoelectrochemical (PEC) cell, where catalysts are immobilised on photoelectrodes and photoinduced electrons flow across the external circuit (Fig. 3d).³⁵ From a techno-economical viewpoint, the photocatalytic system may be more advantageous, as PEC systems may be required to achieve up to ~25% solar-to-hydrogen (STH) conversion efficiency to rival with petrol in energy prices, primarily due to higher investment in installations, whereas a STH efficiency of 5-10% would possibly suffice to render photocatalytic reactors costcompetitive. 4,34 However, the highest STH conversion efficiency of particle-based systems is around 1%,28 whereas that of electrodebased systems can typically achieve more than 10%. 36,37 PEC systems also have several merits that are appealing for fundamental research and practical applications: (1) immobilisation of photocatalysts on electrodes enables in-depth studies of half-reactions

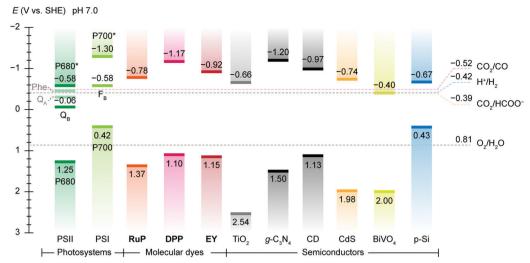


Fig. 2 Energy band edges of light absorbers used in semi-artificial photosynthesis. The band edge locations of photosystems are represented by redox potentials of their reaction centre chlorophylls, i.e. P680 (PSII) and P700 (PSI) and terminal electron acceptors, i.e., Q_B (PSII) and a [Fe-S] cluster F_B (PSI).¹¹⁹ For molecular dyes, the lowest unoccupied molecular orbital (LUMO) and the highest occupied molecular orbital (HOMO) are used as conceptual equivalents of conduction band edge and valence band edge in semiconductors. Data source: DPP, 147 RuP, 147 Eosin Y (EY), 292 TiO2, 293 g-C₃N₄, 294 carbon dot (CD), 283,295 CdS, 296 BiVO₄, 157 p-Si. 162 The redox potentials CO₂/CO (catalysed by CODH), 186 H⁺/H₂ (catalysed by H₂ase), 133 CO₂/ $HCOO^-$ (catalysed by FDH), ¹⁸³ and O_2/H_2O are displayed as references. The energy band edges and redox potentials are corrected with respect to SHE at pH 7.0 via the Nernstian relationship (25 °C).

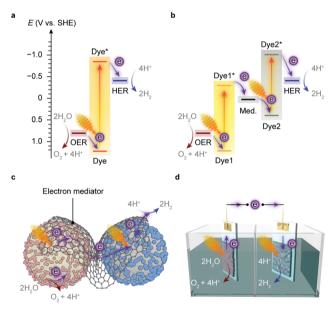


Fig. 3 Artificial approaches to solar-to-chemical conversion. (a and b) Energy level diagrams of light-driven water splitting at pH 7.0: single-step excitation (a) and two-step excitation (b). Dye and Dye*: the light absorber in the ground and excited state, respectively. Med.: redox mediator. (c) Representation of a two-step photocatalytic water-splitting system with semiconductor particles and HER and OER co-catalysts. The HER and OER photocatalysts are interfaced with a solid-state electron mediator. (d) Representation of a PEC tandem water-splitting system.

individually without sacrificial reagents and allows for in situ probing of photoredox chemistry through spectroscopic methodologies;³⁸⁻⁴⁰ (2) from an application's perspective, the electrode-based configuration enables separation of H2 and O2 in a two-compartment configuration, 41 and can be transformed into a continuous fuel-production system with a flowing electrolyte solution. 42 On the other side, drawbacks emerge: (1) the preparation of photoelectrodes requires conductivity and high film-forming capability of semiconductors; (2) the operation of a PEC cell can generate pH gradients and cause mass transfer limitations that account for substantial potential loss and subject electrodes to increasingly corrosive environments; 43,44 (3) the internal resistance often necessitates external bias voltages to drive the electron flow from a photoanode to a photocathode.

Sustainable transformation of CO2 to chemicals and fuels would provide a means to close the loop of the anthropogenic carbon cycle and offer a viable solution for carbon capture and utilisation.45 The linear CO2 molecule is chemically stable, and thus activating the C=O bond for the ensuing endothermic reactions may incur a substantial energy penalty, which is reflected in the negative potential (-1.9 V vs. SHE, pH 7.0) for the oneelectron reduction to CO2 •- 46-48 Although such a high energy barrier can be overcome by large electrochemical potentials or high temperature, it results in poor energetic efficiencies for subsequent fuel synthesis. The activation of CO2 is followed by a stepwise proton and/or electron transfer, which gives rise to furcate reaction pathways towards miscellaneous products. 45,49 In addition, due to the low solubility of CO₂ in water (0.033 mol L⁻¹ at 25 °C, 100 kPa),⁴⁶ the reduction of CO₂ usually competes with H₂ evolution from H₂O, which further reduces the selectivity and the conversion efficiency. 48 Current photocatalytic CO2 reduction systems can produce CO and formate, but still struggle to produce multicarbon chemicals such as ethylene and ethanol selectively, which challenges the economic exploitation of this approach.50

Artificial photosynthesis has recently been extended to drive reactions such as N2 reduction. Today, industrial N2 reduction Chem Soc Rev **Review Article**

relies on the energy-demanding Haber-Bosch process that requires H₂, high temperature (300-500 °C) and high pressure (150-200 atm). 51,52 The Haber-Bosch process also entails considerable CO₂ emission during the production of the inlet H₂ from natural gas (1.87 kg CO₂ per 1 kg NH₃).⁵¹ The intense energy consumption and environmental stress arising from the Haber-Bosch process have spurred explorations in sustainable routes for No-to-NH2 transformation. Photocatalytic N2 reduction uses solar light to replace fossil fuels as the energy source and water instead of natural gas as the hydrogen source and operates at room temperature and atmospheric pressure. 53-56 The prospect of this approach largely depends on the photocatalysts that generate photoelectrons to reduce the stable and inert N=N bond. However, the existing photocatalytic N2 reduction systems still display low solar-to-chemical conversion efficiencies (typically below 0.02%), 51,55,57 and current studies encounter technical difficulties in reliable NH3 quantification and elimination of probable experimental artefacts.⁵⁸

2 Semi-artificial photosynthesis

Artificial photosynthesis establishes solar-to-fuel pathways to address the global energy challenge, but the bottlenecks are key-step reactions during the fuel-forming process, such as water oxidation and selective CO₂ activation. Biology is highly capable of tackling these synthetic challenges through the naturally-refined biocatalytic machinery. Enzymes comprise only a handful of earth-abundant metal atoms as catalytic centres and cofactors, which are embedded in polypeptide chains (Fig. 4a).46 Enzymes perfectly orchestrate the electron and proton transfer, reactant delivery, bond rearrangement and product removal at the active site, thereby reducing the energy threshold and accelerating the reaction kinetics. Enzymes widely employ steric hindrance and electrostatic or hydrogen bond interactions to stabilise selected intermediates and transition states, and therefore provide an efficient reaction path towards a single product.⁵⁹ The catalytic prowess of enzymes has inspired synthetic endeavours to mimic their active sites and ligand environments, and to exploit the biological strategies to stabilise intermediates and control selectivity.60-62 Although synthetic mimics are making progress, they have not yet reproduced the performance of their natural models in terms of catalytic rate, selectivity and electrochemical overpotentials under benign aqueous conditions.

Microorganisms can produce complex multicarbon compounds from simple feedstocks through their intracellular pathways (Fig. 4b). They can also maintain metabolic homeostasis at varying environmental and nutritional conditions. The enzymatic machinery embedded in its well-controlled and confined native environment benefits from the innate regulation mechanisms and can be repaired and regenerated when necessary. On the other side, microorganisms rank their physiological needs to survive ahead of synthetic efficiency to produce target chemicals. Carbon and electron fluxes in the metabolism are partially directed towards biomass synthesis for cell growth and maintenance, which reduces the

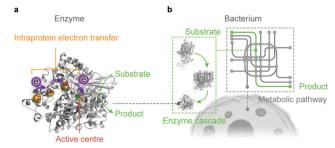


Fig. 4 Biocatalysts in the form of an enzyme (a) and a bacterium (b) used in semi-artificial photosynthesis.

pathway efficiency in chemical synthesis. Moreover, rerouting biochemical pathways towards desired products encounters resistance from the intracellular regulation, which renders such alteration problematic. 63,64

Artificial photosynthesis, however, permits more flexibility in system design and modification. Taking light harvesting as an example: most chlorophylls in nature have a minimum absorption to green light ($\lambda = 500-600 \text{ nm}$), ¹³ which partially accounts for the low photosynthetic efficiency, whereas in artificial photosynthesis, broadband absorption can be readily achieved with a variety of semiconductors and molecular dyes. 11 Furthermore, artificial systems allow for coupling fuelforming reduction reactions with useful oxidative chemical transformations. Several studies have already demonstrated that electrons could be extracted from biomass, organic compounds and even plastics to produce H2 with versatile light-absorbing systems. 65-69 Artificial photosynthesis systems reduce the dissipation of energy and electrons along the pathway, enabling high solar-to-fuel efficiencies routinely surpassing their natural counterparts. 70-73 Artificial photosynthesis is also empowered by an array of analytical techniques such as electrochemistry, spectroscopy and operando methodologies, 40,45,74,75 which are still less commonly employed for quantitative studies in biological research. The well-defined features of synthetic materials permit systematic investigations to understand reaction mechanisms and establish structure-function relationships for system design and optimisation.⁷⁶

Semi-artificial photosynthesis provides a hybrid approach to solar-to-chemical conversion, by integrating the biocatalytic machinery (enzymes and microbes) with synthetic materials (dyes, electrocatalysts, semiconductors, electrodes, mediators).⁷⁷ The photosynthetic biohybrid systems seek to outsource tasks to the components that can perform them best and thus combine strengths of both while bypassing limitations of each. In such hybrid systems, enzymes and microbes function often as catalysts to drive endergonic and complex chemical reactions, whereas synthetic materials are commonly scaffolds to immobilise biocatalysts and functional components that carry out light absorption, charge transfer, chemical transformations and product separation. Several synthetic materials such as polymers and metal organic frameworks can also provide additional protection to fragile enzymes for more sustained operation against environmental stressors. 78-80 According to the form of biocatalyst,

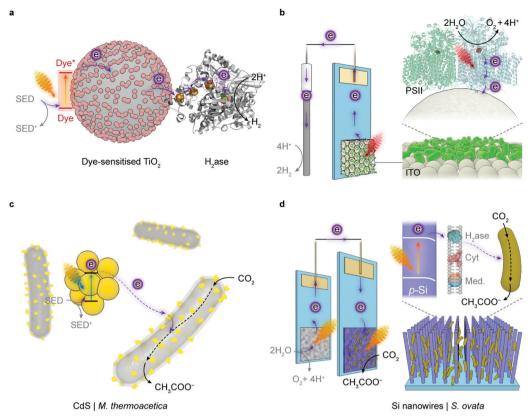


Fig. 5 Representative semi-artificial photosynthesis systems. (a) A colloidal system with dye-sensitised TiO₂ nanoparticles and H₂ase. ¹⁴⁵ Under irradiation, the dye is excited and donates electrons to the conduction band of TiO2. The photoelectrons are then delivered to the catalytic centre of H₂ase for proton reduction via intraprotein electron relays. The oxidised dye is regenerated by extracting electrons from a sacrificial electron donor (SED). (b) An electrode system with a three-dimensional indium tin oxide (ITO) electrode and PSII.^{102,120} PSII is wired by the porous ITO scaffold and photocatalyses water oxidation, aided by an applied potential. (c) A colloidal system with light absorbing CdS nanoparticles deposited on the CO₂-reducing bacterium *M. thermoacetica.*²²⁶ The excited CdS transfers photoelectrons to the bacterium, which generates reducing equivalents to produce acetate from CO2 via the Wood-Ljungdahl pathway. The holes in CdS are quenched by the SED. (d) A hybrid tandem system with TiO2 nanowires as the photoanode and p-Si nanowires integrated with acetogenic bacteria S. ovata as the photocathode.²⁴⁸ When irradiated, the TiO₂ electrode oxidises water while p-Si nanowires generate photoelectrons that are delivered to the interfacing S. ovata via extracellular electron transfer pathways. S. ovata uses these electrons to reduce CO2 into acetate through the Wood-Ljungdahl pathway.

semi-artificial photosynthesis falls into two categories: the enzymatic and the microbial system. Each of these systems can be further distinguished as either a colloidal suspension or a PEC cell (Fig. 5).

3 Enzymatic hybrid systems

An enzymatic hybrid system employs separated electroactive or photoactive enzymes as catalysts or light absorbers, to function in concert with synthetic components in driving reactions selectively at high rates and yields (Fig. 5a and b). Amongst a myriad of enzymes available in nature, only a handful of them are of interest in synthetic reactions starting from the simplest feedstocks (e.g., H₂O, CO₂ and N₂). Translation of their catalytic ability in vivo into advantages in vitro is confronted with several challenges. Whereas synthetic catalysts are either integral parts of electrodes or discrete entities (particles or molecules), these redox enzymes are proteins with molecular weights of 10⁴-10⁶ g mol⁻¹ and areal footprints of 50-400 nm². Their active sites

(catalytic centres) are embedded within the protein matrices, which prevent indiscriminate reactions with substrate analogues. These enzymes have evolved intraprotein electron relays such as haems, quinones, and [Fe-S] clusters that can carry electrons between active sites and redox partners.⁸¹ The electron relays define directional electron transfer pathways within the protein and allow electrochemistry to activate, monitor and modulate the enzymatic redox chemistry.82 The current knowledge of the structure, functionality and catalytic mechanism within the enzymes discussed in this section has already been reviewed previously and is therefore not discussed here. 59,83-86

To approach maximal catalytic rate, enzymes are expected to accelerate the chemical reactions towards diffusion controlled kinetics.87 As such, enzymes should engage with electrodes in an "electroactive" orientation to enable rapid electron transfer kinetics between the distal relay centre and the electrode surface. In this regard, the enzyme-electrode interface must be rationally engineered with respect to topology, porosity and surface chemistry (hydrophilicity, surface charge, functional moieties, etc.). From a kinetic perspective, redox enzymes

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reduce the activation energy and thus, minimise the electrochemical overpotential needed to drive a reaction. 59,88 Therefore, enzymes can often catalyse redox reactions very close to their thermodynamic potentials and reduce the energy penalty arising from large overpotentials. The low overpotentials can simplify system design by judiciously pairing with light absorbers and electron mediators according to their energy levels, and carry out reactions that were otherwise not possible with conventional electrocatalysts.

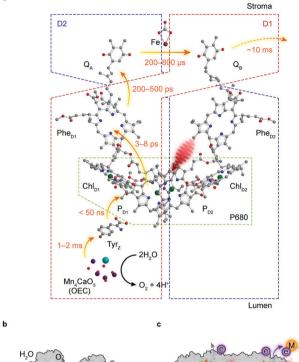
3.1 Water oxidation

Water is the "ideal" terminal electron source for the downstream fuel-forming reactions in artificial photosynthesis, but only few molecular catalysts with earth-abundant elements can handle this multielectron/proton process in an efficient fashion, which is further complicated by the combination with light absorbers. PSII, however, assigns the tasks of light harvesting and water oxidation to antenna proteins and the oxygen evolution centre (OEC), respectively, and orchestrates photochemistry in concert with catalysis in a stepwise manner via efficient intraprotein electron relays.⁸⁹ However, the intrinsic fragility of PSII under light irradiation limits the longevity of its catalytic activity ($\tau_{1/2}$ < 30 min) and thereby strangles the continuous downstream synthesis with isolated PSII.90

Antenna complexes in PSII harvest solar light and funnel the light energy to P680, where charge separation occurs within a few picoseconds.91 This is followed by an electron transfer towards a pheophytin molecule (PheD1) and later to plastoquinone A (Q_A) (Fig. 6a). The resulting highly oxidising P_{D1}^+ in P680 can extract an electron from a redox-active tyrosine Y_Z (Tyrz) and then initiate water oxidation at the Mn₄CaO₅ cluster. 91 On the stroma side, QA is tightly bound to the D2 protein and acts as a single-electron acceptor, whereas $Q_{\rm B}$ in the D1 protein can accept two electrons and be fully protonated. 92 The charge transfer from Q_A to Q_B is aided by a non-haem iron midway between them.93 The formed QBH2 departs from the reaction centre and is replenished by the plastoquinone pool in the thylakoid membrane.

PSII is a biological model OER photocatalyst that inspires synthetic endeavours to mimic its core components, 61,94-96 and stimulates mechanistic investigations to understand its functionality.97-99 Wiring PSII to an electrode allows the resulting biohybrid electrode to supply photoelectrons from water, and permits protein film-photoelectrochemistry (PF-PEC) to benchmark enzymatic activity in vitro, dissect photoinduced electron transfer pathways, and repurpose biogenic electrons to drive endergonic reactions. 90,100

Attempts to interface PSII with an electrode surface started several decades ago, 101 when isolated PSII was deposited on a Pt electrode and generated a photocurrent of a few microamperes. However, the minuscule amount of PSII loaded on the electrode renders in-depth studies and proof-of-concept demonstrations unfeasible. Underlying PSII-PEC is the electrode that provides a physical scaffold to immobilise enzymes and an artificial electron acceptor for the biogenic electrons. PSII's photochemistry and the ensuing current output can be greatly influenced by the material, morphology and physical property of



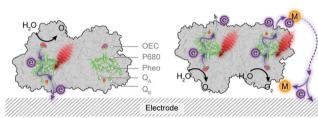


Fig. 6 Electron transfer pathways in PSII-photoelectrochemistry. (a) Schematic representation of the electron transfer pathways (indicated by the orange arrow) in the reaction centre complex of PSII. (b and c) Electron transfer at the PSII-electrode interface. Direct electron transfer can only take place provided that the stromal side of PSII intimately interfaces with the electrode (the electroactive orientation, b). PSII cannot donate electrons to the electrode in unsuitable orientations unless there are diffusional redox couples that can mediate electrons between Q_{B} and the electrode (c).

the electrode scaffold, which dictates the strength of protein binding, the capacity of protein loading, the accessibility to intraprotein electron relays, the depth of light penetration, and the transport of reactants and products therein. 102

The directionality of the intraprotein electron flow makes the interfacial electron transfer highly dependent on the protein orientation. $Q_{\rm B}$, the terminal electron acceptor in PSII, can only undertake outward electron transfer if the stromal side of PSII is in close proximity to an electrode surface (Fig. 6b). The electron transfer via intrinsic plastoquinones, namely, direct electron transfer (DET) can be registered as an anodic photocurrent by chronoamperometry under irradiation, and verified in control experiments by the removal of the Mn₄CaO₅ cluster or addition of 3-(3,4-dichlorophenyl)-1,1-dimethylurea (DCMU) that inhibits the Q_B site and interrupts the electron flow from Q_A. 103,104 DET often results in a relatively low photocurrent because PSII loading may be low and a significant portion of the enzymes may be in an unfavourable orientation or remote

from the electrode surface (Fig. 6c). This problem can be mitigated by using diffusional mediators such as 2,6-dichloro-1,4-benzoquinone (DCBQ, $E_{\rm m}$ = 0.32 V vs. SHE, pH 7.0) to shuttle electrons from the $Q_{\rm B}$ site to the electrode (Fig. 6c). ¹⁰³ This mediated electron transfer (MET), in biological and artificial systems, comes at an expense in both energetics and kinetics, as extra energy is needed to drive the redox turnover of electron mediators and the diffusion of mediators is likely to govern the overall rate of electron transfer.

From both fundamental and practical viewpoints, an electrode with a large number of proteins being wired in an electroactive orientation is desired to produce a high photocurrent that can afford reliable analysis and proof-of-concept demonstrations (Fig. 7a). The making of such electrodes enables new possibilities to create solarto-fuel pathways unattainable in biology. The chronology of the development of PSII electrode design will be outlined below. 100

In the nascent stage of PSII-PEC, PSII with a polyhistidine tag (His-tag) at its stromal side was immobilised on flat Au electrodes by a self-assembled monolayer (Fig. 7b). 105-108 The interaction between electrode-anchored nickel-nitrilotriacetic acid and His-tags permitted site-selective binding of PSII on Au electrodes. However, the flatness of the electrode surface limited the protein loading typically below 10 pmol cm⁻² and thus resulted in submicroampere photocurrents. 108 Electrodes using a redox polymer matrix as a non-diffusional mediator to electrically wire a high amount of proteins enable redox centres on the polymer backbone to access intraprotein electron relays, regardless of their orientations and distances (Fig. 7c). 109 A polymer with the Os3+/Os2+ redox couple $(E_{1/2} = 0.39 \text{ V } \text{vs. SHE})$ was employed as both an immobilisation matrix and an electron mediator to PSII, which gave rise to a photocurrent density up to 45 $\mu A~cm^{-2}$. ¹¹⁰

The second generation of electrodes are mesoporous films made with metal oxides such as Fe_2O_3 , TiO_2 and ITO. $^{104,111-115}$ A mesoporous ITO (meso-ITO) film with a thickness of 2-10 μm and a pore size up to 100 nm can be made with a high degree of tunability (Fig. 7d). 104,111,116,117 The meso-ITO afforded a PSII

loading of 19 pmol cm⁻², which resulted in a DET photocurrent of 1.6 μA cm⁻² and a mediated photocurrent of 22 μA cm⁻². ¹⁰⁴ The surface chemistry of ITO can be modified to covalently bind PSII in favour of an electroactive orientation. 111 The meso-ITO film can also be deposited onto the disk of a rotating ring disk electrode to detect the O₂ production on the ring and study the oxygenic photoreactivity in PSII. 118

The state-of-the-art third generation electrodes for PSII-PEC feature a hierarchical inverse opal structure with both macroporosity and mesoporosity (Fig. 7e). 102,119-127 This electrode architecture has interconnected macropores that provide a large surface area accessible by protein diffusion. The skeleton is made of a mesoporous structure consisting of (semi)conducting and hydrophilic nanoparticles (e.g., ITO or TiO2) where proteins are adsorbed. The inverse opal electrodes can be fabricated via a co-assembly method using polystyrene beads as the structural template and nanoparticles as the electrode material. Such method makes the electrode structure easily variable to fit biocatalysts with different dimensions. 128,129 The hierarchical structure also benefits both mass transport and light transmission, and permits high PSII loading in the range of 30–1000 pmol cm⁻² (depending on the film thickness).119-122

A 40 µm thick inverse opal-ITO (IO-ITO) electrode with PSII attained a high photocurrent of 17 μA cm⁻² for DET and 930 µA cm⁻² in the presence of a DCBQ mediator. 120 Os-based polymers can further be incorporated in the IO-ITO scaffold to improve the electrical wiring between PSII and the electrode to eliminate the need for diffusional additives. 121 Compared to flat electrodes, 110 the IO-ITO electrode with the polymer matrix achieved a higher photocurrent of 381 µA cm⁻² due to higher PSII (336 pmol cm⁻²) loading. The high photocurrent allowed for reliable O2 quantification to calculate the Faraday efficiency, which was not possible with flat electrodes.

PSII immobilisation on porous ITO electrodes also allowed PF-PEC to dissect unexpected electron transfer pathways at the enzyme–electrode interface. 100 DET from Q_A to the ITO electrode

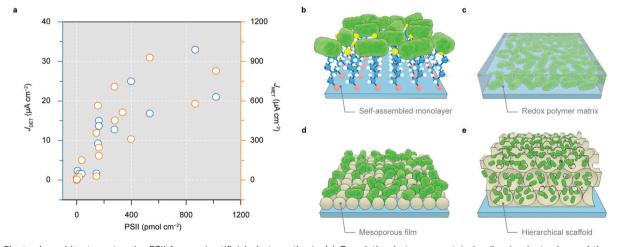


Fig. 7 Electrode architectures to wire PSII for semi-artificial photosynthesis. (a) Correlation between protein loading in electrodes and the resulting photocurrent by DET (blue dots, the left axis) and MET (orange dots, the right axis). (b-e) PSII wired to a flat electrode via a self-assembled monolayer (b), and a redox-polymer matrix (c); PSII immobilised in a mesoporous film with nanoparticles (d), and a three-dimensional hierarchical scaffold (e).

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was observed when inhibiting PSII with DCBO, showing the possibility to release electrons upstream of $Q_{\rm B}$. ^{104,130} Furthermore, analysing cathodic current contributions at low potentials allowed identifying a competing pathway stemming from photoinduced O2 reduction, presenting a short-circuit to the natural water oxidation process. 118,130

Although PSII is the only enzyme that operates in nature for water oxidation to O_2 , it is not the only enzyme that can perform this reaction in vitro. Laccase, a blue multicopper oxidase that couples the oxidation of organic substrates with the reduction of O2 to H2O, has been shown to catalyse the reverse water oxidation reaction when immobilised on an electrode (E > 1.2 V vs.SHE, pH 7.4). 131 In an attempt to use this enzyme for light-driven water oxidation, laccases were adsorbed on a semiconducting In₂S₃ electrode (band gap: 2.0 eV) as a visible-light absorber. ¹³²

3.2 Hydrogen production

H₂ases are metalloenzymes that can catalyse the interconversion between H₂ and H $^+$ /e $^-$ with a TOF of > 1000 s $^{-1}$. 84,133 The HER activity of H₂ases per active site is therefore comparable with that of the benchmark Pt catalyst. 134,135 There are three phylogenic classes of H2ases termed according to the metallic centres at their active sites: [NiFe]-H2ase, [FeFe]-H2ase and [Fe]-H2ase, but only [NiFe]- and [FeFe]-H2ases incorporate [Fe-S] clusters as electron relay centres and can catalyse the reversible proton reduction to H₂. These [Fe-S] clusters are spaced 10-14 Å apart, which permits sequential electron tunnelling towards the active sites at a rate (10⁷ s⁻¹) faster than catalysis. 137,138

These intraprotein electron relays can couple the catalytic turnover at the active site with the electron exchange through the electrode, and thus enable protein film electrochemistry (PFE) to probe the H₂-H⁺ interconversion by deciphering resulting voltammograms and chronoamperograms. 135,139 PFE reveals that [FeFe]-H2ases are active for both proton reduction and H₂ oxidation yet extremely sensitive to O₂, whereas the more O2-tolerant [NiFe]-H2ases usually show strong catalytic bias towards H2 oxidation and the reverse H+ reduction is often susceptible to H₂ inhibition. 136

[NiFeSe]-H2ases are a subclass of [NiFe]-H2ases with a selenocysteine residue coordinated to the nickel at the active site instead of a cysteine. 138 [NiFeSe]-H2ases are kinetically more biased than other [NiFe]-H2ases towards H2 evolution without substantial product inhibition (Fig. 8i and j). The presence of selenium provides additional protection to the nickel centre from oxidative attacks, and hence confers better oxygen tolerance. 138,140,141

[NiFeSe]-H₂ase is therefore a common model biocatalyst and has been rationally coupled with various light absorbers for photocatalytic HER (Fig. 8). Early attempts employed dyes to generate photoelectrons, which were carried to H2ases by diffusional electron mediators such as methyl viologen (MV), in the presence of a sacrificial electron donor (SED). 142,143 A more recent study shows that an organic dye Eosin Y could directly transfer photoelectrons to [NiFeSe]-H2ases without any electron mediators (Fig. 8a). 144 The homogeneous system produced 0.5 µmol h⁻¹ H₂ for up to 15 h, corresponding to a

 $TOF_{H,ase}$ of 13.9 s⁻¹. Notably, this **Eosin Y**-H₂ase system also demonstrated some tolerance towards oxygen: more than 80% of photoactivity was retained with 5% O2 and the system remained photoactive even at aerobic conditions (21% O2) for 3 h. The system performance was limited by the interfacial electron transfer between dye molecules and H2ases.

To this end, semiconductors were used as light absorbers with immobilised Hases. Photogenerated electrons can thereby stream directly from the conduction band towards the enzyme's distal relay centre (Fig. 8b). Co-adsorbing [NiFeSe]-H2ases on **RuP**-TiO₂ particles through the putative interaction between the TiO₂ surface and side-chain carboxylates near the distal [Fe-S] cluster yielded 3.56 µmol H₂ in the first hour of irradiation and a benchmark TOF_{H,ase} of 50 s⁻¹. 145,146 Replacing RuP with a metal-free diketopyrrolopyrrole (DPP) dye to sensitise TiO2 also attained similar HER activity (Fig. 8b). 147 Likewise, [FeFe]-H₂ases were immobilised on CdTe nanoparticles or CdS nanorods that had their surfaces modified with 3-mercaptopropionic acid. 148,149 The negatively-charged surface induced [FeFe]-H₂ases to interface with the semiconductors at their positive patch close to the distal [Fe-S] cluster. Transient absorption spectroscopy elucidated that the electron transfer occurred between the conduction band of CdS and the distal [Fe-S] cluster in [FeFe]-H₂ase, regardless of the enzyme activity, at a rate (\sim 100 ns) comparable with that of the charge recombination in CdS. 150 The rate of electron transfer is several orders of magnitude slower than that in synthetic systems (10⁻³-1 ns), ¹⁵¹ probably due to the longer tunnelling distance (>10 Å) between the semiconductor surface and the distal [Fe-S] cluster. Nevertheless, the interfacial electron transfer is fast enough to outpace the incident photon flux, rendering the light absorption by CdS performancelimiting.150

The H₂ evolution rate of the above systems declined over a few hours, presumably due to dye degradation or an unstable material-enzyme interface. Thus, robust carbon-based light absorbers such as carbon nitride (CNx) and carbon dots were utilised to improve the long-term stability (Fig. 8c). 152-154 The CN_x-H₂ase hybrids remained photoactive for 48 h, totalling 2.5 µmol of H₂ production and a turnover number (TON) of > 50 000. 152 But the weak interaction between CN_x and H₂ases rendered the TOF_{H, ase} much lower (1.5 s⁻¹, in the first 4 h) than that of RuP-TiO2-H2ase hybrids. This problem can be partially resolved by blending CN_r with TiO₂ that has a high affinity for H₂ases.¹⁵³ The ternary hybrids exhibited better photocatalytic activity with a higher TOF_{H, ase} of 7.8 s⁻¹ within the first 4 h of irradiation and a longer H₂ evolution span of 72 h. The photostability of H2ases was even better than their synthetic analogue, a molecular DuBois-type Ni bis(diphosphine) catalyst, that decomposed within 3 h of irradiation. 152 [NiFeSe]-H2ases adsorbed on amine-capped carbon dots attained similar activity: a TOFH, ase of 1.1 s⁻¹, H_2 production of 2.6 µmol and a total TON of 52 000 in 48 h. 154 A comparison with carboxylic acid-capped carbon dots echoes with the previous finding that [NiFeSe]-H2ases prefer to electrostatically interact with positive surfaces, and underlines the importance of semiconductor surface chemistry that controls the electronic communication with enzymes. 148,149,155

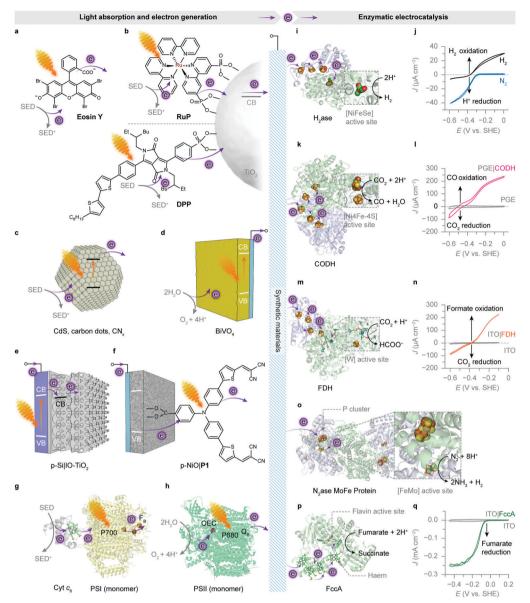


Fig. 8 Enzymatic hybrid systems for semi-artificial photosynthesis. (a-h) Light harvesting and low-potential electron generation by abiotic and biotic light absorbers. (i–q) Enzymatic electrocatalysis. (a) An excited xanthene dye **Eosin Y** ($\lambda_{max} = 539 \text{ nm}^{297}$). (b) Photoexcited electrons from the **RuP** or diketopyrrolopyrrole (DPP) dye are injected into the conduction band (CB) of TiO₂ and further directed to biocatalysts. (c) Irradiation of semiconducting particles, such as polymeric carbon nitride (CN_x), carbon dots and CdS, results in photoinduced electron transfer to biocatalysts. (d) A monoclinic BiVO₄ electrode absorbs light and generates photoelectrons while using water as the electron donor. (e) An IO-TiO2 scaffold on a p-Si electrode as the photocathode for biocatalysts. Under irradiation, the excited p-Si electrode injects electrons into the CB of IO-TiO2, which are further delivered to the interfacing biocatalysts. (f) A P1 dye-sensitised p-type NiO electrode as the photocathode for biocatalysts. Light absorption by the P1 dye provokes electron transfer from the valence band (VB) of NiO, followed by an electron injection into the biocatalysts adsorbed on the NiO scaffold. (g) The reaction centre chlorophylls (P700, λ_{max} = 700 nm) in PSI are excited by visible light irradiation. The photogenerated electrons are transferred to the terminal [Fe-S] cluster F_B via intraprotein electron relays. The luminal side of PSI is tethered with a cyt c_6 . Shown here is the core complex of a PSI monomer (protein data bank ID: 50y0). (h) An excited PSII monomer (protein data bank ID: 4ub6) transfers photoinduced electrons to the terminal electron acceptor (Q_B) at the stromal side. The generated holes oxidise water to O_2 at the OEC. (i) The protein structure of a [NiFeSe]- H_2 ase (protein data bank ID: 1cc1). Exogenous electrons are delivered via [Fe-S] clusters to the [NiFeSe] active site for proton reduction. (j) Protein film voltammogram of H₂ase adsorbed on a TiO₂ electrode in N₂ or H₂, at pH 6.0. Reproduced from ref. 162 with permission from Wiley-VCH, copyright 2016. 162 (k) The protein structure of a CODH (protein data bank ID: 1jqk). Electrons are transferred via [Fe-S] clusters to the [NiFeS] active sites to drive the CO2 reduction into CO. (l) Protein film voltammetry scans of CODH adsorbed on a pyrolytic graphite edge (PGE) electrode (control: a bare PGE electrode, grey trace) in CO2/CO (50%/50%) at pH 6.0. Reproduced from ref. 159 with permission from the American Chemical Society, copyright 2013.¹⁵⁹ (m) The protein structure of a FDH (protein data bank ID: 1h0h). Electrons are transferred through the intraprotein [Fe-S] relays to the [W] active site for the reduction of CO2 to formate. (n) Protein film voltammogram of reversible reduction of CO2 to formate by an FDH-loaded mesoporous ITO electrode (control: a bare ITO electrode, grey trace) in CO2/NaHCO3 containing formate, at pH 6.5. Reproduced from ref. 193 with permission from Wiley-VCH, copyright 2019. (a) The protein structure of a [MoFe] protein from N₂ase (protein data bank ID: 4wes). Exogenous electrons access the [MoFe] active site via a P cluster ([8Fe7S]) to transform N₂ into NH₃. (p) The protein structure of fumarate reductase (FccA, protein data bank ID: 1d4c). Haem cofactors function as intraprotein electron relays to direct electrons towards the flavin active site for fumarate reduction to succinate. (g) Protein film voltammogram of FccA adsorbed on an IO-ITO electrode (control: a bare IO-ITO electrode, grey trace) in 1 mM fumarate at pH 7.0. Reproduced from ref. 117 with permission from the Royal Society of Chemistry, copyright 2016.¹¹⁷

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The aforementioned systems employ SEDs such as triethanolamine (TEOA) or ethylenediaminetetraacetic acid (EDTA) to expediently bypass the kinetic difficulty of water oxidation. To enable water to supply electrons for reductive reactions, synthetic chemistry and biology utilise different approaches. Artificial photosynthesis can carry out water oxidation separately by a photoanode in a PEC system. BiVO4 has a band structure well-suited for solar water oxidation, but it is insufficient to reduce protons (Fig. 2 and 8d). 156,157 A photocathode hosting HER catalysts is thus required to further energise the electrons that are withdrawn from water by the BiVO₄ photoanode. The electrochemical window of ITO (-0.6 to +2 V vs. SHE, pH 7.0) limits its application at negative potentials needed to drive reducing reactions. 158 In view of this, ITO can be substituted with semiconductive TiO2 to fabricate porous electrodes that are applicable under cathodic conditions. 159-161

To this end, [NiFeSe]-H₂ases were interfaced with p-type silicon that was stabilised by a mesoporous TiO2 layer. 162,163 The TiO₂ layer is necessary to protect the underlying p-Si from unwanted surface passivation, 163,164 but also to assist in binding [NiFeSe]-H2ases. An ensuing problem is that enzymes having a large footprint show a low loading capacity on the mesoporous TiO₂ layer. 163,165 As such, a hierarchical IO-TiO₂ structure was built atop the p-Si to increase the H2ase loading (up to 120 pmol cm⁻²) (Fig. 8e). 166 A 4 nm TiO₂ interlayer was deposited on the surface of p-Si by atomic layer deposition to protect the electrode from the formation of an insulating SiO₂ layer, followed by the construction of the IO-TiO2 scaffold. The p-Si|IO-TiO₂|H₂ase electrode gave a photocurrent onset potential of 0.0 V vs. SHE, pH 6.0, which is more positive than that of a water oxidising $BiVO_4$ -TiCo photoanode (-0.1 V vs. SHE, pH 6.0). Thus the BiVO₄-TiCo||p-Si|IO-TiO₂|H₂ase tandem PEC system afforded overall solar water splitting without an external voltage, resulting in a stoichiometric production of H₂ (0.47 μ mol, $\eta_{\rm F}$ = 98%) and O₂ (0.20 μ mol, $\eta_{\rm F}$ = 84%) after 5 h of irradiation. 166 [FeFe]-H2ases have also been interfaced with black Si electrodes with submicron porosity, but they are prone to desorb from the electrode surface due to weak interfacial interactions. 167

Instead of using a p-type semiconductor, the H2ase-loaded IO-TiO₂ scaffold was also interfaced with an encapsulated lead halide perovskite solar cell that generated a photovoltage of 0.9 V. In a tandem configuration with a BiVO₄-TiCo photoanode, the resulting system produced H2 coupled to O2 evolution with a STH efficiency of 1.1%. 168

To streamline the electron transfer in biological H₂ production, photosystems and H2ases could be directly wired via molecular wires, redox polymers, protein subunits or external circuits to eliminate the diffusion-governing steps. 18,120,166,169-171 For example, PSI has been connected with [FeFe]-H2ases by 1,6-hexanedithiol to allow photogenerated electrons to tunnel between the two proteins (Fig. 8g). 18,169,172

Alternatively, PSII and H2ases were immobilised on IO-ITO electrodes and compartmentalised in a PEC cell (Fig. 8h). 120 The onset potential of an IO-ITO $|H_2|$ as cathode (-0.4 V vs. SHE, pH 6.5) remained close to the thermodynamic value of proton reduction, whereas electrons from P680 were attenuated during the intraprotein electron transfer in PSII (Fig. 2).

The energy is further offset by the use of electron mediators such as DCBQ that aid in the electron transfer from PSII to the electrodes. As such, a minimum bias voltage of 0.6 V was required to drive the photoelectrons from PSII towards H₂ase. At an applied voltage of 0.9 V, this biohybrid PEC cell (IO-ITO|PSII||IO-ITO|H₂ase) generated 0.52 μ mol O₂ (η _E = 104%) and 0.96 μ mol H₂ (η_F = 98%). Note that the system performance with respect to photocurrent, onset potential and longevity was limited by the IO-ITO PSII photoanode. To reduce the bias voltage in the future, one could possibly extract electrons from the upstream electron relays such as Q_A or pheophytin (Fig. 2). 104,130

An easier way is to substitute for an electron mediator with a more negative redox potential than DCBO whilst also introducing a second light absorber to reduce the bias voltage. This has been achieved by the use of 3,5-di-tert-butyl-1,2-benzoquinone (DTBoQ, $E_{\rm m}$ = 0.29 V vs. SHE) as the electron mediator, which shifted the photocurrent onset by 0.1 V earlier than DCBQ. When paired with a p-Si|IO-TiO₂|H₂ase photocathode, the required voltage was reduced to 0.24 V. At 0.4 V, the tandem PEC cell produced 0.125 μ mol H₂ (η _F = 91%) within 3 h. ¹⁶⁶

At the anodic side, a second light absorber can be judiciously selected to imitate the function of PSI or to complement the light absorption of PSII. A working example is DPP that has a negative excited-state redox potential (-1.15 V vs. SHE) and strong light absorption between 400 and 550 nm (λ_{max} = 496 nm). 147 DPP and PSII were loaded in an IO-TiO2 scaffold as the light absorbers and electrically connected by an Os-based redox polymer (POs) hydrogel. Thus, a vectorial electron transfer from water to the TiO2 electrode was established to prime electrons with sufficient reducing power for HER. Moreover, DPP conferred supplementary green-light absorption on this PSII-containing photoanode, thereby enabling a panchromatic sensitivity to visible light. By directly wiring the IO-TiO₂-**DPP**|P_{Os}-PSII photoanode with an IO-ITO|H₂ase cathode, autonomous solar water splitting can be carried out using wired enzymes without a bias voltage. 119 After 1 h of irradiation at zero bias, 0.015 µmol H2 was detected with a Faraday efficiency of 76%. H2ases at the cathode were not driven to their full capacity in catalysing proton reduction, due to the photodegradation of PSII in vitro. An initial applied bias STH conversion efficiency of 0.14% was obtained at a voltage of 0.3 V during the first hour. 119 PSI, the native second light absorber in the Z-scheme, has also been rewired with PSII and H2ase by redox polymers in an attempt to reconstruct a biological H2 production pathway in vitro. 173

3.3 Carbon dioxide reduction

From a viewpoint of carbon products, the biomass produced by the Calvin cycle is not ideal to substitute for petrochemicals as feedstocks for the chemical industry. 174 However, biology has evolved more than one pathway for CO₂ fixation, ¹⁷⁵ and the reductive acetyl-coenzyme A (acetyl-coA) pathway (Wood-Ljungdahl pathway) is an energy-conserving CO2-fixing pathway, through which inorganic carbon is assimilated into cellular metabolism in the form of formate and CO by carbon monoxide dehydrogenase (CODH) and formate dehydrogenase (FDH). 176-178 Both [NiFe]-CODH and metal-dependent FDH can catalyse the interconversion between CO_2 and their two-electron reduced carbon products, CO and formate, respectively. They control the binding of CO_2 and the transfer of electrons and protons to partition intermediates on the reduction pathway towards CO or its hydrated form, formate. ⁴⁶

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[NiFe]-CODH is an O2-sensitive metalloenzyme with a [NiFe]active site from anaerobes such as Moorella thermoacetica and Carboxydothermus hydrogenoformans (Fig. 8k). It efficiently catalyses CO oxidation with a TOF up to 40 000 s⁻¹ and CO₂ reduction with a TOF up to 45 s⁻¹, at a nearly thermodynamic potential (-0.5 V vs.SHE, pH 6.7) (Fig. 8l). 46,179,180 CO₂ is bound to the [NiFe]-cluster and undergoes a bifunctional attack by the electrophilic iron centre and the nucleophilic nickel centre, which enables a twoelectron pathway for CO₂ activation. Such two-electron reduction is also applied by [Mo]- or [W]-FDH, a metalloenzyme reversibly interconverting CO2 and formate with minimal overpotentials (-0.39 V vs. SHE, pH 7.0) (Fig. 8m and n). 183,184 It has been determined that [W]-FDH isolated from Syntrophobacter fumaroxidans exhibited a TOF up to 3400 s⁻¹ for formate oxidation and 280 s⁻¹ for CO₂ reduction (with MV as the redox partner). 183 Both [W]- and [Mo]-FDH can achieve ~100% Faraday efficiency for electrocatalytic CO2 reduction with formate as the sole product at modest overpotentials. 183,184 These enzymes provide an efficient two-electron CO2 reduction pathway that bypasses the thermodynamically-uphill formation of CO₂•-.185

In view of the close thermodynamic redox potentials for CO_2/CO (-0.52 V vs. SHE, pH 7.0), proton reduction (-0.42 V vs. SHE, pH 7.0) and $CO_2/HCOO^-$ (-0.39 V vs. SHE, pH 7.0) (Fig. 2), similar strategies can be employed to drive CO₂ reduction with solar energy. [NiFe]-CODH was adsorbed on RuP-TiO2 nanoparticles for light-driven CO₂ reduction to CO in the presence of an SED (Fig. 8b). During 4 h of irradiation, the RuP-TiO₂-CODH hybrids produced \sim 5 μ mol CO, corresponding to an average TOF_{CODH} of 0.15 s⁻¹. ¹⁸⁶ The relatively low turnover rate compared with H2ase hybrids was partly due to a smaller driving force for CO2 reduction than proton reduction, as the conduction band edge of TiO₂ (-0.52 V vs. SHE, pH 6.0) nears the thermodynamic potential for CO₂ reduction to CO. Factors such as interaction with TiO₂ and enzyme orientation also affect performance. The enzyme activity began to decrease after 3 h, and the low activity might be partly due to inefficient dye regeneration. 186,187 CODH has also been integrated in a PEC system comprising a P1 dye-sensitised p-type NiO cathode (Fig. 8f). 188 The excited P1 dye receives an electron from the valence band of the NiO electrode, followed by electron injection into an adjacent CODH, via intraprotein electron relay centres, down to its [NiFe]-active site, where CO2 was reduced to CO (Fig. 8f).

Light-driven CO_2 -to-formate conversion has been previously realised by FDH-based photoredox systems, where electrons from SEDs are energised by photosensitisers and delivered to FDH via electron mediators, $^{78,189-191}$ or by an FDH-based PEC system, where photoelectrons from the semiconductor photocathode p-InP were mediated to FDH by the MV/MV²⁺ redox couple. 192 The diffusional mediator can be eliminated by immobilising FDH on **RuP**-TiO₂ nanoparticles (Fig. 8b). 193 FDH showed a TOF of 11 s⁻¹ in the first 6 h of photocatalysis

and yielded 2.6 µmol formate. 193 Quartz crystal microbalance and infrared spectroscopy revealed that TiO2 binds FDH strongly through both electrostatic interaction and chemisorption. FDH can also be wired with an electrode to work in tandem with a photoanode. 122,194,195 A recent study coupled an FDH-functionalised IO-TiO2 cathode with an IO-TiO2-DPP|POs-PSII tandem photoanode to drive CO2-to-formate reaction with the aid of a small bias voltage. 122 After 1 h of irradiation with a bias of 0.3 V, 0.046 μ mol of formate ($\eta_F = 70\%$) was detected as the only product in the cathodic chamber of the PEC cell. The photocurrent decayed by half after 8 min, arising from the photodamage of PSII in the anode. Self-driven CO2 reduction was recently achieved by a PV-PEC tandem system with a perovskite solar cell, a BiVO₄-FeOOH photoanode and a TiN nanoshell|FDH cathode. 194 The resulting PV-PEC tandem cell yielded formate at a rate of 1.06 μmol h⁻¹ with a Faraday efficiency of 83% in 8 h, corresponding to a solar-to-fuel efficiency of 0.08%. The formate production could last for three days, and the decrease in activity was ascribed to the degradation of the perovskite solar cell outside the PEC cell. FDH can also be loaded on a TiO2-deposited CuFeO2-CuO photocathode to couple with a BiVO₄-FeOOH photoanode for autonomous light-driven ${\rm CO_2}$ reduction. This bias-free PEC cell allowed CO₂-to-formate conversion at a rate of 0.098 µmol h⁻¹ with a Faraday efficiency of 33.5% in 8 h.

The product of CO₂ reduction can be extended from formate to methanol via an enzyme cascade. 197-199 A prototypical PEC system employed CoP_i-modified α-Fe₂O₃ and BiFeO₃ as the photoanode and photocathode, respectively, and contained a mixture of enzymes, i.e., FDH, formaldehyde dehydrogenase and alcohol dehydrogenase, in the cathodic chamber. 198 Electrons extracted from water are energised by photoelectrodes and vectored through enzyme cascades to reduce CO2 to methanol via NADH and diffusional electron mediators. Instead of being directly wired on the photocathode, enzymes in this system were dispersed in the catholyte, and depended on NADH, the reducing agent, to carry out redox reactions. As such, despite the well-aligned band positions, an external voltage (0.8 V) was needed to efficiently regenerate NADH and produce methanol (1.31 mM in 6 h). 198 An enzyme cascade was also immobilised in a silica matrix to work with an "artificial thylakoid" in a colloidal system. 199 The "artificial thylakoid" constituted a microporous protamine-TiO2 hollow sphere (microcapsule) with CdS nanoparticles deposited on its luminal surface. The TiO2 microcapsule received electrons from the excited CdS and exported electrons to regenerate NADH that furnished the enzyme cascade with reducing equivalents. In this way, photocatalytic oxidation and enzymatic CO2 reduction were decoupled so that the enzymes were protected from photoinduced reactive oxygen species (ROS). This biohybrid system has demonstrated a high rate and yield in renewing NADH and thus generated 85 μM methanol from CO₂ in 2 h. ¹⁹⁹

3.4 Chemical synthesis

With enzymes as biocatalysts, photogenerated electrons from light absorbers can realise a broad spectrum of reactions of Chem Soc Rev **Review Article**

synthetic interest, such as N2 reduction, reduction of C=C bonds, formation of C-C bonds and hydroxylation of C-H bonds etc. Several recent studies are briefly highlighted here.

Biology converts N2 into metabolically tractable NH3 using N₂ases under physiological conditions. N₂ase is a two-component system comprising a [MoFe] protein and an electron-transfer [Fe] protein. The two proteins associate and dissociate during catalysis, sequentially delivering electrons and energy to the [MoFe] active site, where two NH3 and one H2 are generated from one N2 and eight protons and electrons (Fig. 8o).86 To leverage N2ases for light-driven N₂ fixation, the isolated [MoFe] proteins were adsorbed on a CdS nanorod to form a biohybrid complex, where light energy replaces ATP hydrolysis to drive the enzymatic turnover in N2.200 The photoexcited CdS nanorods generated electrons with low reducing potential (-0.74 V vs. SHE, pH 7.0), which were then forwarded to the interfacing [MoFe] proteins for N₂ reduction. The CdS-[FeMo] protein hybrid attained an NH₃ production rate of 315 nmol $\text{mg}_{N_2 \text{ase}}^{-1} \ \text{min}^{-1}$ and an average TOF_{N_2ase} of 1.25 s⁻¹. The NH₃ production lasted up to 5 h with a TON of 1.1 \times 10⁴ mol_{NH₂} mol_{N₂ase}⁻¹. ²⁰⁰

The reduction of C=C bonds is exemplified by the conversion of fumarate to succinate, which can be catalysed by the flavoenzyme fumarate reductase (FccA) (Fig. 8p and q).201 The light-driven fumarate reduction can be either performed in a colloidal system or within a PEC cell, using RuP-sensitised TiO2 nanoparticles, and carbon dots or a W-BiVO₄-CoP_i electrode, respectively (Fig. 8b-d). 154,202 The FccA adsorbed on RuP-TiO₂ exhibited a TON of 5800 over 4 h of irradiation, corresponding to a TOF_{FccA} of 0.4 s^{-1} , 202 while that on amine capped carbon dots achieved a similar TOF_{FccA} of 0.47 s⁻¹ (averaged in the first 2 h of irradiation) with improved photostability up to 24 h. 154 Photoinduced enzyme degradation was likely to account for the slowdown of productivity. In a PEC system with a W-BiVO₄-CoP_i photoanode, the turnover rate of the immobilised FccA dropped down to $\sim 0.01 \text{ s}^{-1.202}$

Another example is the use of old yellow enzymes (NADPH dehydrogenases) for stereoselective C=C reduction. Light absorbers including Au-TiO2 or N-doped carbon dots were employed to produce photoelectrons that drive the regeneration of reducing equivalents such as NAD(P)H and its synthetic substitute, for trans-hydrogenation of conjugated C=C bonds. 203,204 With molecular xanthene dyes, photoelectrons could be directly transferred to the prosthetic flavin moiety in the old yellow enzyme, eliminating the need of using costly nicotinamide cofactors.²⁰⁵ Such light-driven enzymatic catalysis yielded enantioselective products (enantiomeric excess > 90%) with conversion yields up to 80-90%. 204,205

Photochemical C-C bond formation has recently been realised in a CdS-2-oxoglutarate:ferredoxin oxidoreductase (OGOR) hybrid system.²⁰⁶ The OGOR catalyses the amalgamation of CO₂ and succinate into 2-oxoglutarate and is part of the reductive tricarboxylic acid cycle responsible for CO2 fixation in many autotrophic microorganisms. The irradiated CdS generates reducing equivalents to drive the catalytic turnover at OGOR which involves large substrates, significant conformational changes during catalysis, and eventual formation of C-C bonds.206

Functionalisation of C-H bonds provides a straightforward and atom-economical access towards a plethora of organic products,207 where enzymes such as cyt P450 show their synthetic advantages. The hallmark reaction of cyt P450 is the hydroxylation of C-H bonds, during which the delivery of electrons from the NADH reductase is synchronised with the activation of oxygen at the haem centre. The disadvantage of activity dependence on NADH cofactors and redox partners can be overcome by a semi-biological solution. A Ru(II)-diimine dye was covalently attached to a mutant haem domain of a cyt P450, which enabled excited dyes to inject electrons into the haem centre to catalyse the hydroxylation of lauric acid at an initial rate of 2.1 s^{-1} with a TON of more than 900 in 2 h, without the aid of NADH and reductase.208

4 Microbial hybrid systems

Microbial cells include an entire system of biochemical pathways and can produce metabolites of synthetic interest with high specificity at physiological conditions. Microorganisms utilise numerous enzyme cascades to maintain the intracellular metabolism. Different metabolic pathways are spatially organised to divert metabolites or enzymes that can react promiscuously, to maintain selectivity, concentrate reactants to drive unfavourable reactions, and protect enzymes or unstable intermediates from harmful cytoplasmic contents.²⁰⁹ Moreover, these biosynthetic pathways are under dynamic regulation to keep cellular functionality in tune with physiological needs at different conditions. These features allow microbes to synthesise complex products from the simplest and stable feedstocks (e.g., H₂O, CO₂, N₂, etc.) and render microbial catalysis resilient to environmental variations. Appealing to synthetic chemistry is their ability to produce multicarbon products from CO₂ via inherent carbon assimilation pathways such as the Wood-Ljungdahl pathway and the Calvin cycle. 175,178

Such synthetic complexity is challenging to achieve with individual enzymes and a fixed reaction stoichiometry. Microbial catalysis can be further empowered by synthetic biology that employs genetic tools to access and engineer the intracellular metabolism, and create new pathways beyond the native metabolic pattern. 210 Compared with enzymatic counterparts, the microbial hybrid systems are less explored, but have provoked increasing interest in solar-to-chemical conversion, due to the possibility to synthesise complex chemicals, achieve better stability and prospects for scalability.211,212 In the following discussions, we will summarise the recent achievements that take advantage of the microbial metabolism for chemical synthesis in solar energy conversion with photocatalysis and (photo)electrochemistry (Table 1).

Whereas enzymes catalyse biochemical reactions with specific substrates, products and stoichiometry, microorganisms are more versatile due to a set of diverse metabolic pathways involving a multitude of enzymes. Moreover, the microbial hybrid systems are being developed to establish an efficient conversion strategy of solar energy to desired multicarbon chemicals, a catalytic prospect particularly suitable for microorganisms.²¹³

Table 1 Summary of selected microbial hybrid systems for solar-tochemical conversion

Microbe	Light absorber	Microbial reaction	Ref.
E. coli	TiO ₂	$H^+ \rightarrow H_2$	221
E. coli	CdS	$H^+ \rightarrow H_2^-$	222
R. capsulata	Bi_2O_3	$H^+ \rightarrow H_2$	218
R. capsulata	dye-TiO ₂	$H^+ \rightarrow H_2$	219
R. eutropha	PV	$CO_2 \rightarrow biomass$, isopropanol	240
M. barkeri	CdS	$CO_2 \rightarrow methane$	225
M. barkeri	TiO ₂ /InP	$CO_2 \rightarrow methane$	239
M. thermoacetica	CdS	$CO_2 \rightarrow acetate$	226
M. thermoacetica	Au	$CO_2 \rightarrow acetate$	231
M. thermoacetica	PDI/PFP	$CO_2 \rightarrow acetate$	234
S. ovata	Si nanowires	$CO_2 \rightarrow acetate$	248
S. ovata	Si nanowires	$CO_2 \rightarrow acetate$	250
Clostridium sp.	PV	$CO_2 \rightarrow butanol, hexanol$	212
R. palustris	CdS	$CO_2 \rightarrow PHB$, carotenoids	232
S. cerevisiae	InP	Glucose → shikimic acid	238
G. sulfurreducens	$BiVO_4$	Fumarate → succinate	262
X. autotrophicus	PV	$N_2 \rightarrow NH_3$	243

Thus, we summarise progress for microbial hybrid systems from a different perspective: instead of organising this section by individual reactions enabled by such systems, we put the emphasis on key questions imperative for research in this field: how microorganism and synthetic materials are integrated into hybrid structures, how non-photosynthetic microbes are empowered with light-absorbing capabilities, and how fluxes of reducing equivalents are guided into intracellular metabolic pathways.

Microbial photocatalysis in colloidal systems

The discussion begins with the microbial hybrids operating in colloidal systems (Fig. 5c). We consider the HER as a model reaction in the context of solar-to-fuel conversion. Photosynthetic H₂ production can transiently take place in nature by microalgae and cyanobacteria under anaerobiosis, during which electron flux on the photosynthetic chain is shifted away from its normal path towards H₂ase. ^{17,214} However, such process is O₂-sensitive and competes with NADPH-dependent biological processes. Although the O2-sensitivity can be partially alleviated by protective encapsulation, 215,216 the rate and yield of H2 evolution are intrinsically limited by the light reaction, which saturates at fairly low light intensities (10-20% of solar light).²¹⁴ Whereas these drawbacks can be partially overcome by streamlining the solar-to-H₂ pathway with isolated H₂ases and light absorbers, laborious purification and fragility of the enzymes diminish the practicality of this approach.

An alternative are microbial hybrids where non-photosynthetic live cells are employed as biocatalysts, and the task of light harvesting is outsourced to synthetic light absorbers. Several early studies illustrate this strategy: H2ase- (and N2ase-) containing bacteria such as Clostridium butyricum, Rhodopseudomonas capsulata, Rhodospirillum rubrum were coupled with semiconductor nanoparticles, specifically, TiO2 and Bi2O3 via MV as the electron mediator (Fig. 9a). In the presence of SEDs, photogenerated electrons were delivered into cells where H₂ evolved.^{217–219} The low growth rate of these bacteria limits the availability of biocatalysts. As such, Escherichia coli (E. coli), a well-established bacterial chassis for synthetic biology, 220 was engineered to heterologously express genes encoding [FeFe]-H₂ases from Clostridium acetobutylicum.²²¹

The recombinant E. coli was mixed with TiO2 nanoparticles and MV, and such hybrid system produced 216 µmol H2 during 15 h of irradiation.

To enable visible light absorption, CdS was employed, and instead of inducing heterologous expression of [FeFe]-H2ases, E. coli can produce H₂ in anaerobic conditions with endogenously synthesised [NiFe]-H2ases. Moreover, CdS nanoparticles were precipitated on the bacterial surface to directly supply photoelectrons without diffusional mediators (Fig. 9b).²²² It has been found that photogenerated electrons interacted with the intracellular metabolism and guided more reducing equivalents along the pathway for H₂ production. To extend its aerobic uses, the CdS-E. coli hybrid was further encapsulated with biomimetic silica to protect the bacterium from oxygen.216 Under aerobic conditions, the hybrid system continuously produced H₂ within 96 h of irradiation.

The catalytic ability of microorganisms is better exemplified in CO2 reduction, where they can produce complex organic products not readily accessible through synthetic methods.²¹¹ One example is CO2 reduction to CH4, which involves eight electron transfers that can facilely diverge to form a wide range of products. 223 In the realm of biology, this reaction is handled by methanogenic archaea as part of their energy-conserving metabolism.²²⁴ In light of this, Methanosarcina barkeri (M. barkeri), an anaerobic methanogen, was employed as the catalytic machinery for CO₂-to-CH₄ conversion.²²⁵ Like the aforementioned CdS-E. coli hybrids, ^{216,222} CdS nanoparticles were deposited on the surface of the M. barkeri as the light absorber. Photoelectrons produced by the excited CdS were delivered to the intracellular pathways via membrane-bound cytochromes or H2ases and were used by the bacteria for methanogenesis (Fig. 9b). The CdS-M. barkeri hybrid produced 13.7 μmol CH₄ after three-day irradiation with a quantum efficiency of 0.34%. 225 Isotopic labelling experiments confirmed that the CH₄ was derived from CO₂ reduction. The CH₄ evolution halted after three days, likely due to the depletion of the SED cysteine.

CdS nanoparticles can also be interfaced with Moorella thermoacetica (M. thermoacetica), an acetogenic and electrotrophic bacterium. M. thermoacetica could directly use photogenerated electrons from the excited CdS to reduce CO2 into acetate via the intracellular Wood-Ljungdahl pathway, while holes produced by CdS were quenched by cysteine (Fig. 9b). 226 A peak quantum yield of 2.44% was recorded at low-intensity simulated sunlight, which is comparable to the year-long averages (\sim 0.2-1.6%) for plants and algae.²²⁷

The electron transfer pathways between CdS and M. thermoacetica was further elucidated by transient absorption spectroscopy: photoelectrons were transported via membrane-bound electron acceptors such as cytochromes, ferredoxin, flavoproteins and menaquinones at the initial stage (3 h) of photosynthesis; at longer duration (24 h), a membrane-bound H2ase-mediated pathway dominated, which correlated with increased rates of acetate generation and thus a higher quantum efficiency.²²⁸ A latest proteomic study revealed that the photoexcited CdS induced changes to the metabolic pattern of M. thermoacetica: a host of enzymes associated with the tricarboxylic acid cycle and glycolysis were up-regulated. Both pathways together with the

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a

TiO₂

B

NADP+

NADPH | H-I

SED+

NADPH | H-I

SED+

C

C

d

C

CO₂

SED+

CCO₂

SED+

CCO₂

SED+

CCO₂

SED+

CCO₂

SED+

CCO₂

CCO₃

CCO₄

CCO₂

CCO₂

CCO₃

CCO₄

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CCO₄

CCO₅

CCO₆

CCO₇

CCO₈

CCO₈

CCO₉

CC

Fig. 9 Photosynthetic microbial hybrid systems. (a) Irradiated semiconducting nanoparticles (e.g., TiO_2) transfer photoexcited electrons to microbes to produce H_2 or other products via the MV^2+MV^+ or $NADP^+/NADPH$ redox couples in the presence of a SED. (b) CdS nanoparticles are deposited on the surface of microbes and deliver photoelectrons directly to the intracellular pathways for solar-to-chemical conversion and oxidise the SED under illumination. (c) Au nanoclusters are absorbed by microbes as intracellular light absorbers. Photoinduced electrons are delivered by cytoplasmic redox mediators to the Wood–Ljungdahl pathway that reduces CO_2 into acetate. (d) Microbes are introduced in the cathodic chamber of an electrolyser and leverage the produced H_2 as the reducing equivalents to transform CO_2 into CO_2 into CO_2 into alcohols.

CH,COO

Wood–Ljungdahl pathway constitute the energy-conservation scheme of the CdS–*M. thermoacetica* hybrid system. ²²⁹

To make the hybrid system more sustainable, the added cysteine, the SED to CdS, could be regenerated photocatalytically by TiO₂ nanoparticles loaded with Mn(II) phthalocyanine co-catalysts, which ultimately extracted electrons from water and delivered them to the CdS-M. thermoacetica hybrids via the cysteine/cystine redox couple.²³⁰ The light absorbers can be replaced with those that are more biocompatible and accessible, for example, Au nanoclusters. 231 Au nanoclusters could be ingested by M. thermoacetica and directly deliver photoexcited electrons with longer lifetime to the Wood-Ljungdahl pathway via cytoplasmic redox mediators (Fig. 9c). The intracellular interplay between photochemistry and metabolism bypassed kinetic and energetic drawbacks stemming from transmembrane diffusion. The Au nanoclusters also could reduce ROS formed during irradiation, thus improving the cellular viability in this hybrid system and increasing acetate production with a higher quantum yield of $\sim 2.9\%$.²³¹

A similar protocol is applicable to various microorganisms with different physiological functionality. For example, CdS nanoparticles were precipitated on the photosynthetic bacterium *Rhodopseudomonas palustris* for CO_2 fixation. Under visible light, the excited CdS generated additional reducing equivalents for the innate Calvin cycle, which promoted the production of biomass and valuable multicarbon compounds such as carotenoids and poly- β -hydroxybutyrate (PHB), as well as the photosynthetic efficiency. When CdS was hybridised with *Thiobacillus denitrificans* (*T. denitrificans*), an autotrophic denitrifier, the resulting CdS–*T. denitrificans* hybrid could use the photoelectrons for NO_3 reduction to N_2O_* .

In light of its cytotoxicity, CdS has been substituted with organic semiconductors, *i.e.*, perylene diimide derivative (PDI) and poly(fluorene-*co*-phenylene) (PFP) as the photosensitiser.²³⁴ PDI and PFP could be immobilised on the surface of *M. thermoacetica* through electrostatic and hydrophobic interactions, which enabled efficient electron transport across the cellular membrane. Such microbial hybrid system registered a quantum efficiency of 1.6%.

Light-driven fumarate reduction to succinate has been achieved by an enzymatic hybrid system with flavoenzymes (FccA), 154,202 but such reaction can also be accomplished by a

dye-sensitised whole cell containing flavocytochromes.²³⁵ This hybrid system comprised of photosensitisers (Eosin Y or proflavine), Shewanella oneidensis MR-1 (S. oneidensis), an electron mediator (MV) and a SED. Another model reaction is C-H bond functionalisation. Cyt P450 was employed to catalyse the hydroxylation of lauric acid with a covalently attached Ru(II) light absorber that was meant to deliver photoelectrons to the haem domain without the need of redox partners. 208 The same strategy also applies in a microbial hybrid system: the cyt P450 was expressed in an engineered E. coli bacterial chassis and Eosin Y was absorbed into the cytoplasm and directly bound to the haem domain of the enzyme.²³⁶ Under visible light irradiation and with an SED, photoexcited Eosin Y can continuously supplement electrons to sustain the catalytic turnover of the cyt P450 more than 18 h (average TOF_{P450}: $2.5 \times 10^{-4} \text{ s}^{-1}$). 236

A comparison with the previous P450-Ru(II) dye hybrids illustrates the disparity that epitomises general differences between enzymatic and microbial systems. First, the P450-Ru(II) dye hybrids need post-translational modification to link the Ru(II) dye molecule to the specific cysteine residue, and such modification is specific to a particular type of cyt P450 and may affect the activity when applied to different cyt P450 variants.208 In contrast, the whole cell system has demonstrated success with both bacterial and human cyt P450 in producing high value chemicals such as drugs and steroids.²³⁶ Lastly, the enzyme hybrids showed an initial TOF_{P450} (2.1 s⁻¹) several orders of magnitude higher than the whole cell system, but the activity of the former began to languish within 2 h (average TOF_{P450} : 1.25 \times 10⁻³ s⁻¹), while productivity of the latter can be compensated by the prolonged longevity (>18 h) and the increased quantity of biocatalysts. 208,236

Besides bacteria, eukaryotic cells such as yeasts are also favoured as a platform cell factory for biosynthesis of fuels and chemicals, as their well-studied biology primes genetic and analytical tools to unravel the mechanisms governing electron transport and metabolic flux in biohybrid systems. ⁶⁴ The intracellular redox reactions are dependent on cofactors, namely NADH or NADPH, as hydride sources and their generation is strongly intertwined with the central carbon metabolism. ²³⁷

The expense of NAD(P)H precludes a stoichiometric supply to sustain the biosynthesis, and an effective cofactor regeneration was achieved photocatalytically by hybridising the yeast strain Saccharomyces cerevisiae (S. cerevisiae) with semiconducting InP nanoparticles (Fig. 9a). 238 S. cerevisiae was genetically engineered to suppress the cytosolic synthesis of NADPH and to enhance the carbon flux towards shikimic acid production. The NADPH was regenerated by the inward electron transport by irradiated InP nanoparticles bound on the cell surface. A maximum light-to-chemical conversion efficiency of 1.58% was attained within 12 h of irradiation. 238

4.2 Electrolysis coupled to homogeneous microbial synthesis

Another biohybrid system integrated planktonic M. barkeri with a PEC cell system consisting of a water oxidising TiO2 photoanode and a p-InP-Pt photocathode, to couple artificial H2 evolution with biological methanogenesis. The unassisted water splitting under irradiation generated H₂ in the cathodic chamber, which was consumed by the methanogens therein as reducing equivalents to reduce CO₂ to CH₄ (Fig. 9d). This hybrid system produced 68.8 nmol CH₄ over three days, corresponding to a Faraday efficiency of 74%. 239

Similar strategies can be extended to other strains of bacteria to obtain CO2-reducing products that are more feasible for storage and transport. 240,241 For example, in a water electrolyser inoculated with genetically-engineered Ralstonia eutropha, a water oxidising CoPi anode and a Co-P alloy cathode worked together at an applied voltage of 2.0 V to furnish the bacteria with H₂.²⁴¹ The bacteria transformed CO2 and H2 following a desired route into an array of fuels and chemical products such as poly(3hydroxybutyrate), isopropanol, isobutanol and 3-methyl-1-butanol (Fig. 9d). The energy input can be readily outsourced to a photovoltaic device with a routinely attained efficiency of 18%, and the resulting system could anticipate a solar-to-chemical conversion efficiency of > 7%. ^{241,242} Such a modular platform has furthermore been leveraged beyond CO2 reduction, towards more complex ambient NH3 synthesis from N2 and H2O with the H2-uptake diazotrophic microorganism Xanthobacter autotrophicus.²⁴³ In the above hybrid systems, H2 acted as the vector of reducing power to connect microbial metabolism with artificial electrolysis. However, the low solubility of H₂ (0.79 mM, 25 °C) and inefficient H₂ delivery from the electrode limit the rate of throughput. A recent study shows that this bottleneck can be alleviated by introducing perfluorocarbon nanoemulsions as biocompatible H2 carriers to increase the H2 availability around cells, accelerate the H2 transfer and thereby improve the productivity of CO₂ reduction.²⁴⁴

4.3 Photoelectrosynthesis with immobilised microorganisms

Instead of suspending microbes in solution, they can also be immobilised on electrodes to catalyse redox reactions driven by light or electricity (Fig. 10a). Such systems are referred to as microbial (photo)electrosynthesis. Underlying this process is the electron exchange ability of microorganisms with the interfacing electrode (Fig. 10b-d).245 This can either occur directly between the electrode and microorganism or be mediated by H2 and diffusional redox couples. The direct electron transfer for microbial electrosynthesis has the advantage of bypassing the handicap of low solubility of H2 and the diffusional limitations of electron shuttles. Besides, it avoids the potential loss from intermittent H2 generation and the potential toxicity of some mediators (e.g., MV).246

From a practical perspective, the advantages over a planktonic system largely resemble the benefits of heterogeneous catalysis over homogeneous systems such as easy catalyst separation and recycling. One can readily and flexibly apply the microbial electrode for photosynthesis, either by pairing with a photoelectrode in a PEC cell or connecting with a solar cell for PV-driven electrolysis. Of more interest to fundamental studies is that the hybrid electrode system permits electrochemical toolsets to access the intracellular redox chemistry via the current exchange at the biointerface.

Oxygenic photosynthetic microorganisms such as cyanobacteria have been immobilised on IO-ITO electrodes with tailored porosity to perform semi-artificial photosynthesis. 128,129 The cyanobacterium encloses the entire photosynthetic apparatus within the cell, rendering themselves more robust than isolated PSII to carry out light-driven water oxidation. The cyanobacterial biofilm produced much lower photocurrent densities than those of PSII, but with increased longevity (>5 days). 128 Compared with the IO-ITO|PSII electrodes, the IO-ITO|biofilm systems exhibited different photocurrent profiles, which is derived from different electron transfer mechanisms at the interface. The cyanobacterium likely relies on intracellular redox mediators that are released under irradiation to carry electrons towards the electrode surfaces.247

There are only few reports to directly interface nonphotosynthetic microorganisms with photoelectrodes. 248,249 The difficulty is in part due to the large footprint of cells, which requires appropriate electrode geometry to achieve a high loading density while maintaining light penetration. A notable example is an array of Si nanowires that served both as a lightabsorbing semiconductor and a physical scaffold for the acetogenic bacterium, Sporomusa ovata (S. ovata) (Fig. 5d).²⁴⁸ The bacteria were integrated with the Si nanowires by taking photoexcited electrons from the conduction band to sustain their metabolism. The resulting biohybrid electrode was assembled with a photoanode made of water oxidising TiO2 nanowires for autonomous photoreduction of CO2 to acetate. The PEC cell produced a stable photocurrent of 0.3 mA cm⁻² for more than 120 h, yielding an acetate titre of 20 mM with a Faraday efficiency of 86%. The overall solar-to-chemical conversion efficiency registered 0.38% during 200 h operation. Acetate could further be activated into acetyl-coA, a common biochemical intermediate, to access a variety of biosynthetic fine chemicals. The downstream synthesis could be performed by genetically-engineered E. coli that transformed acetate into n-butanol, polyhydroxybutyrate polymer, and isoprenoid compounds.²⁴⁸ Recently, the acetate production was accelerated as a result of a larger bacterial population and a better bacterium-nanowire interface, which was realised by improving the local pH environment inside the electrode.250 As such, the solar-to-chemical conversion efficiency was increased to 3.6% over seven days when coupling to a photovoltaic device.

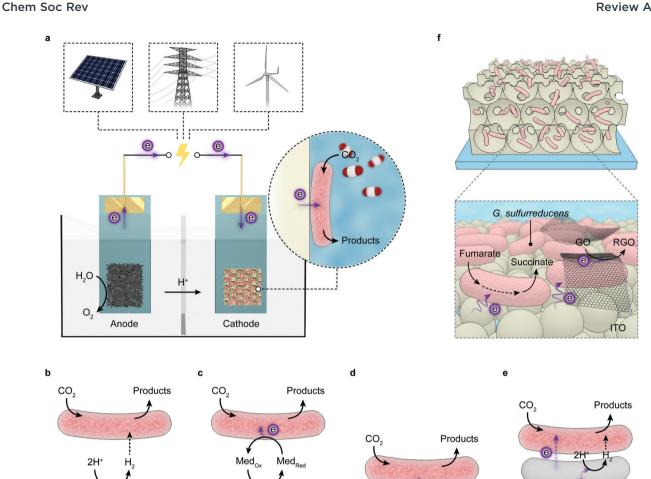


Fig. 10 Microbial electrosynthesis. (a) Schematic representation of a microbial electrosynthetic cell powered with electricity that can be derived from renewable sources (solar light, wind, etc.). At the anode chamber, an electrode donor (e.q., H₂O) is oxidised, providing electrons. The cathode chamber contains a microbial hybrid electrode where microorganisms receive electrons from the interfacing electrode scaffold and drive the reduction reaction (e.g. CO₂ reduction). (b-e) Mechanisms for electron transfer from electrodes to microorganisms. Electrons from the electrode can be delivered to the microorganisms by H2 (b), redox mediators (Med), (c), direct electron transfer (d) or via interspecies electron/H2 transfer (e). (f) An IO-ITO|G. sulfurreducens hybrid electrode, in which G. sulfurreducens can directly receive electrons from the cathode and reduce fumarate or graphene oxide (GO), Reproduced from ref. 262 with permission from the National Academy of Sciences of the United States of America (NAS), copyright 2020.

Electrode

Few light-absorbing semiconductors have been optimised to fit the dimension of microorganisms. Integration of microbes into the electrode scaffold also poses challenges on the chemical stability, biocompatibility, conductivity and surface chemistry of the semiconducting materials, not least the impact on light transmission. A simpler strategy is to immobilise microorganisms on electrodes with tailored morphologies while outsourcing light harvesting to photoanodes or external PV devices. This approach decouples light absorption from microbial metabolism, and thereby suppresses unwanted parasitic reactions and eliminates possible oxidative stresses stemming from photoinduced ROS. The simplified design also grants more flexibility in the making of electrode architectures to enhance the loading capacity and engineer the electrode surface to improve the interaction. Therefore, more efforts converge on interfacing microbial cells with tailored electrodes and improving the interfacial electron transfer and mass transport.

Such exploitation can capitalise on decades of research in microbial fuel cells, 251 and be guided by recent understandings of the interaction between synthetic materials and microorganisms. 252,253 As these hybrid electrodes can readily operate with light absorbing electrodes or devices, to broaden the scope of discussion, this section will continue with a survey of recent progress in microbial electrosynthesis, where sessile bacteria exchange electrons with electrodes and carry out synthetic reactions.

An early study of microbial electrosynthesis employed graphite electrodes to supply Geobacter sulfurreducens (G. sulfurreducens) with electrons for fumarate reduction.²⁵⁴ This is reminiscent of a previously described electrode-bacterium hybrid system, where H₂ acted as the electron vector (Fig. 10b). Yet this scenario has been ruled out because the electrode barely produced H₂ at the applied potential (-0.3 V vs. SHE, pH 6.8). Instead, G. sulfurreducens was able to directly receive electrons from the electrode (Fig. 10d).²⁵⁵ The same reaction could also be accomplished by S. oneidensis

cultured on a graphite electrode or in a three-dimensional reduced graphene oxide (RGO) scaffold. 256,257

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An envisioned application for microbial electrosynthesis is bioremediation, namely, decontaminating the wastewater possibly powered by solar energy. 258 For example, under applied negative potentials, G. sulfurreducens could precipitate soluble U(v1) ions on electrodes, which can be easily collected. 259 Bacteria such as Geobacter lovlevi and Anaeromyxobacter dehalogenans with the ability to respire with chlorinated compounds can also be exploited to dechlorinate contaminants (e.g., tetrachloroethene, 2-chlorophenol) in this manner. 260,261

More recently, an IO-ITO electrode scaffold was employed to culture a large population of G. sulfurreducens (Fig. 10f). 262 The resulting biohybrid electrode catalyses the reduction of soluble fumarate and heterogenous graphene oxide (GO), with electrons from an external power source or an irradiated dye-sensitised (**RuP**-TiO₂) or water oxidising BiVO₄ photoanode. The microbe-modified IO-ITO electrode enabled high current densities and also served as a platform underpinning a variety of analytical methods to decipher the electrical interplay at the biointerface.²⁶³

However, Geobacter and Shewanella strains lack CO₂-fixation pathways. Autotrophic microorganisms that can source energy from inorganic chemical reactions (chemolithotrophs) or light (photoautotrophs) for carbon assimilation are more promising in this regard, as they can adapt to use an electrode as the electron donor to generate reducing power.²⁶⁴ Acetogens such as S. ovata, Clostridium ljungdahlii, and M. thermoacetica have been shown able to directly receive electrons from electrodes and reduce CO_2 into acetate with Faraday efficiencies of > 80%. Moreover, the electron uptake and CO2 fixation can also be decoupled using two bacteria strains (Acetobacterium woodii and Desulfobacterium corrodens), which can separately receive electrons from the electrode and reduce CO2, respectively, and communicate by means of interspecies H₂ transfer (Fig. 10e).²⁶⁵

As microbial electrosynthesis systems are anticipated to operate together with PV devices, the ultimate solar-to-chemical efficiency will to a large extent depend on the biochemical pathway that dictates the energetics and stoichiometry of the substrateproduct conversion. Whereas a sulphur-oxidising phototroph Prosthecochloris aestaurii can sustain its photosynthesis with electrons supplied from a cathode, 266 its application for photoelectrosynthesis is largely unfeasible because the Calvin cycle is rather inefficient for solar fuel production and genetic tools to expand their synthetic diversity are not yet developed for some less studied microorganisms. Compared with methanogens of which CH₄ is the only attainable product, acetogens enable a wide repertoire of chemicals such as ethanol, isopropanol, n-butanol and 2,3-butanediol, which can be further extended by metabolic engineering. 267,268

The alliance of microbial metabolism with photocatalysis and (photo)electrochemistry is formed by virtue of the translocation of reducing power from abiotic materials to biotic entities. This can be achieved by interfacing microorganisms with light absorbers/electrodes, where electrons are either directly transferred to microorganisms or shuttled inward via in situ generated electron carriers (e.g. H2). Alternatively,

electrons can be delivered to planktonic cells by diffusive redox mediators. Sessile microorganisms on electrodes can manage a high Faraday efficiency up to 90%. Yet the slow electron uptake (typically $< 1 \text{ mA cm}^{-2}$), stemming from the kinetic disparity between the electrochemical and biochemical processes, which is less amenable to improvement by interface engineering, diminishes the economic feasibility of this approach. 269 Thus a technically favoured strategy may be to decouple the extracellular electrochemistry and intracellular biochemistry in a modular fashion, as exemplified by a recent microbial electrosynthesis setup that combined PV, electrolysis and fermentation modules.212 CO2 and H2O was electrochemically reduced into H₂ and CO in a PV-powered electrolyser at a current density of 50 mA cm⁻². The abiotically generated syngas was then fed into a microbial fermenter where a mixed culture of Clostridium autoethanogenum and Clostridium kluyveri worked together to produce butanol and hexanol at an energetic efficiency of 78% and an overall Faraday efficiency of $\sim 100\%$. ²¹² As such, the highest attainable solar-to-chemical efficiency of the integrated system could be projected at 15%, assuming that the existing PV modules deliver a routine energy conversion efficiency of 20%. 212

5 Future perspectives

Semi-artificial photosynthesis opens a wide spectrum for new solar-to-chemical conversion designs by rationally integrating biocatalysts in the form of proteins and whole cells with synthetic components, which traverses the fields of enzymology, microbiology, materials science, electrochemistry, and photocatalysis. The ultimate goal is to catalyse the transformation of H₂O, CO₂, and N₂ into fuels, chemicals and fertilisers via a sustainable path and synthesise value-added fine chemicals with high yield and selectivity with renewable energy sources. Here we briefly propose several promising opportunities for the future development.

At the biotic side, more enzymes or bacteria relevant to fuel/ chemical production can be added into the inventory. The methodologies established with model biocatalysts on bespoke electrodes and semiconductor particles can be extrapolated to new biocatalysts to create new reaction pathways. For example, RuP-TiO2 particles were used in early work to adsorb H2ases for photocatalytic HER. 145 The use of the same light absorber was later extended to CODH, FDH and FccA. 186,193,202 CdS nanorods developed for CdS-H2ase complexes have been employed to drive N2-to-NH3 conversion with N2ases. 149,200 Likewise, CdS nanoparticles have been in situ deposited on the surface of E. coli and M. thermoacetica. 222,226 It is reasonable to envision that a similar strategy can be applied to other acetogens such as S. ovata and methanogens like M. barkeri.

Biocatalysts are widely-appreciated for their high product selectivity, yet such catalytic prowess is confined to naturallyoccurring reactions. One way to expand the reaction scope is orchestrating various catalysts, including biotic and synthetic, to drive cascade reactions towards more complex and valuable products.²⁷⁰ For example, an enzyme cascade has been introduced

into a tandem PEC cell to reduce CO_2 into methanol. ¹⁹⁸ More opportunities are emerging in this respect to judiciously construct reaction pathways and sensibly direct electron and carbon flux towards desired products.

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The second approach arises when advancing genetic techniques meet with enzymology and microbiology. With the derived methodologies, enzymes can be engineered with improved promiscuity and stability. PSII has been introduced with a polyhistidine tag near the stromal side in order to enhance the linkage with Ni-terminated moieties on an electrode surface. 108 Such modifications and single site mutations more broadly are available for other enzymes to improve the interaction with light-absorbing materials. Pathways in microorganisms can be altered to maximise the chemical production or heterologously introduced to unlock unnatural reactions. Metabolic engineering can allow heterotrophic E. coli to synthesise sugar from CO₂, ²⁷¹ and cyanobacteria to drive the reduction of C=C bonds, ²⁷² or produce valuable aromatic compounds directly from CO₂.²⁷³ Engineered microbes with additional pathways can work with light absorbers to yield synthetically useful products beyond their natural metabolites.²³⁸

Moreover, the genetic toolkit can alter the existing or enable new physiological functionalities in model microorganisms. For example, current production of electrogenic microorganisms can be enhanced by increasing metabolic fluxes, promoting electron shuttle secretion, encouraging biofilm formation, *etc.*²⁷⁴ Heterologous electron conduits have been transplanted into *E. coli* to confer extracellular electron transfer capability.^{274,275} *S. oneidensis* has been engineered to metabolise glucose as carbon and energy source.²⁷⁶ Beyond this, these techniques also provide powerful means to generate insights into the poorly understood questions with regard to microbial metabolism strategies and electron uptake pathways, which, in turn, will facilitate the rational design of next-generation biohybrid systems.

In parallel, progress at the artificial side can be translated into strengths to stimulate the advancement of semi-artificial photosynthesis systems. Developments in materials science and synthetic chemistry will provide more opportunities in tuning a material's composition, morphology and property at various levels to improve the interaction with biocatalysts or adapt to different reaction conditions. A recent report demonstrated the application of a metal-organic framework to confine FDH within its precisely defined volume and porosity to stabilise enzymatic activity at non-physiological conditions. 277 A metal-organic framework was also employed to encapsulate anaerobic bacteria (M. thermoacetica) to protect them against oxidative stress arising from irradiation.215 Such a strategy is feasibly applicable to other enzymes and bacteria. Another emerging possibility is to foster interaction with biocatalysts at a sub-entity level, instead of forming heterogeneous biointerfaces. This has been realised by fusing dye molecules with cyt P450 to replace native redox partners, 208 or anchoring Pt nanoparticles with PSI for light-driven H_2 evolution. 278 At a subcellular scale, Au nanoclusters or molecular dyes have been exploited to intervene pathways within cells with light irradiation as bioorthogonal opportunities for intracellular modulation. 231,236,279

Furthermore, the intracellular metabolism can be coupled with extracellular redox transformations to manufacture functional materials such as polymers and inorganic particles.^{280–282} These materials may provide additional benefits to promote extracellular electron transfer or cytoprotection.²⁸¹

Advanced photochemistry provides another thrust in future development. Many photosynthetic biohybrid systems are limited by light absorbers that display non-ideal absorption and driving force to perform reactions, or suffer from degradation faster than the inactivation of biocatalysts. Potential solutions may arise from the iterative cycle of design, synthesis and characterisation in the field of artificial photosynthesis, which yields a variety of photocatalysts with fine-tuned properties and morphologies. Progress in this field can instantly contribute to the construction and optimisation of biohybrid systems. For example, carbon nitride and carbon dots emerged as an attractive visible-light absorber due to their well-fit band structure, low cost, and robustness. 66,283 Carbon nitride-H2ase and carbon dot-H2ase complexes have been shown effective in photocatalytic H₂ production. ^{152–154} As such, more applications can be anticipated by replacing H2ase with other enzymes, or substituting the toxic CdS to interface with acetogenic bacteria. Studies into these newly-derived systems will gain more knowledge of the biotic-abiotic interplay and generate guidelines for the assembly of better systems.

Finally, semi-artificial photosynthesis, in a broader context, also enables a suite of powerful analytical techniques to delve into fundamental questions relevant to physiological functionality, biocatalytic redox chemistry and biotic–abiotic interaction (Fig. 11).²⁸⁴

Wiring enzymes or cells to an electrode allows electrochemical methods to probe their redox chemistry and physiological functionality under turnover and non-turnover conditions, through efficient interfacial electron transfer (Fig. 11a). ^{59,82,285,286} The simplicity of the electrochemical apparatus (typically in a three-electrode electrochemical cell) renders it readily compatible with an extra light source to investigate the light-driven redox reactions of photoactive enzymes (Fig. 11b). ⁹⁰ Besides stationary electrodes, rotating electrodes, namely rotating disk electrodes or rotating ring disk electrodes, provide an effective way to eliminate the limitations in mass transport that often obscures catalytic features. ²⁸⁷ The rotating ring disk electrode, with an additional ring working electrode surrounding the central disk, enables quantification of reaction products formed at the disk electrode (Fig. 11c). ¹¹⁸

Complementary to electrochemical methods are spectroscopic techniques, which allow for *in situ* detection of electronic and vibrational changes with relevance to enzymatic or cellular functionalities under electrochemical control and provide structural or mechanistic insights that were otherwise intractable with electrochemistry alone. For instance, ATR-IR was previously deployed to track molecular catalysts and reaction intermediates during catalytic cycles, ²⁸⁸ while it also assisted to probe the enzyme–electrode interaction in H₂ase–TiO₂, FDH–TiO₂ and PSII–ITO systems (Fig. 11d). ^{102,166,193} Their versatility renders these tools readily transferable to biohybrid assemblies, and offers a high degree of flexibility to work together with (photo)-electrochemistry in *operando* studies (Fig. 11e).

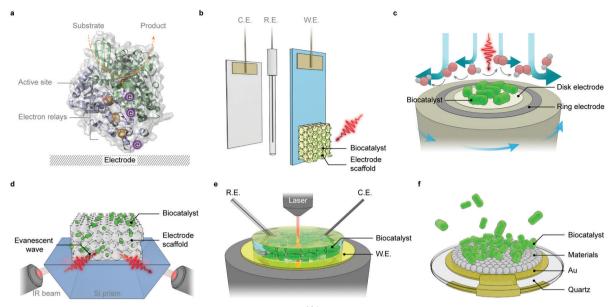


Fig. 11 Analytical techniques employed in semi-artificial photosynthesis.²⁸⁴ (a) An enzyme-electrode interface. Electrons from the electrode are delivered to the active site via intraprotein electron relays. (b) A three-electrode setup comprising a working electrode (W.E.), counter electrode (C.E.) and reference electrode (R.E.) for protein film (photo)electrochemistry and microbial biofilm (photo)electrochemistry. (c) A rotating ring disk electrode with a biocatalyst-adsorbed disk electrode and a Pt ring electrode. The rotation of the apparatus generates a convective flow that hydrodynamically carries reaction products from the disk to the ring electrode for product analysis. (d) An ATR-IR setup with an electrode scaffold deposited on an optical waveguide crystal (e.g., a Si prism). The internal reflection generates an evanescent wave that penetrates into the electrode scaffold by typically ~500 nm. (e) In situ spectroelectrochemistry that combines vibrational/electronic spectroscopy with biofilm electrochemistry. Spectroscopic signatures with mechanistic significance can be captured during the course of redox variation with accurate electrochemical control. (f) A typical piezoelectric chip (electrode materials deposited on a Au substrate) for quartz crystal microbalance analysis. Through the mass-frequency relationship, the mass variation stemming from the biocatalyst binding or desorption can be monitored, enabling quantitative insights into bio-electrode interaction and biocatalyst loading capacity.

Time-resolved pump-probe spectroscopy, namely transient absorption spectroscopy, is universally applied in artificial photosynthesis systems to look into kinetic aspects of charge separation and transfer, which take place on a timescale $(10^{-12}-10^{-3} \text{ s})$ intractable by other methods.²⁸⁹ With such temporal resolution, electron transfer at a semiconductorenzyme interface and the ensuing catalytic turnover can be deconvoluted, enabling unique strength in unravelling kinetic intricacies in photosynthetic biohybrid systems. This technique has been exploited to track the charge transfer in CdS nanorod-H₂ase hybrids and revealed that photogenerated electrons transferred from CdS to the distal relay centre in H2ase at a rate comparable to that of their relaxation process. 150 In a more complex microbial hybrid system (CdS-sensitised M. thermoacetica), transient absorption spectroscopic analysis provided kinetic insights into the electrical interplay between photoexcited semiconductors and acetogenic bacteria, and contributed to the elucidation of pathways connecting extracellular photochemistry and intracellular metabolism.²²⁸

Beyond electrochemistry and spectroscopy, quartz crystal microbalance (with dissipation) enables a sensitive method to measure the mass change on a piezoelectric quartz chip (Fig. 11f). 76 This allows quantification of enzyme loadings in an electrode scaffold, which is not easily accessible by electrochemical or spectroscopic means and can be further extrapolated to evaluate the nature of enzyme interaction with electrode materials. 166,193 More tools are continuously being added into the toolbox. Recently, electron paramagnetic resonance spectroscopy has been combined with protein film electrochemistry to study metalloenzymes (Cu-Zn superoxide dismutase) adsorbed on a mesoporous ITO electrode under precise potential control.²⁹⁰ Nanoscale secondary ion mass spectrometry was applied to study the metabolic activity of G. sulfurreducens biofilms.²⁹¹ These emerging analytic techniques can provide fresh insights from different perspectives to produce a more comprehensive picture of biohybrid systems.

Semi-artificial photosynthesis synergises functional components from different disciplines. Advances in each field will yield new opportunities to further the development and implementation of biohybrid systems. Conclusions drawn from these studies will deepen the current understanding of the biology-materials interplay and form concrete steps to expedite the evolution of semi-artificial photosynthesis towards maturity.

Conflicts of interest

There are no conflicts to declare.

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References

Chem Soc Rev

- 1 P. K. Nayak, S. Mahesh, H. J. Snaith and D. Cahen, *Nat. Rev. Mater.*, 2019, 4, 269–285.
- 2 A. Polman, M. Knight, E. C. Garnett, B. Ehrler and W. C. Sinke, *Science*, 2016, 352, aad4424.
- 3 Y. Tachibana, L. Vayssieres and J. R. Durrant, *Nat. Photonics*, 2012, 6, 511–518.
- 4 J. H. Montoya, L. C. Seitz, P. Chakthranont, A. Vojvodic, T. F. Jaramillo and J. K. Nørskov, *Nat. Mater.*, 2017, 16, 70–81.
- 5 N. Nelson and A. Ben-Shem, *Nat. Rev. Mol. Cell Biol.*, 2004, 5, 971–982.
- 6 R. E. Blankenship, *Molecular Mechanisms of Photosynthesis*, Wiley-Blackwell, 2nd edn, 2014.
- 7 G. G. B. Tcherkez, G. D. Farquhar and T. J. Andrews, *Proc. Natl. Acad. Sci. U. S. A.*, 2006, **103**, 7246–7251.
- 8 T. J. Erb and J. Zarzycki, Curr. Opin. Biotechnol., 2018, 49, 100–107.
- M. Hagemann and H. Bauwe, Curr. Opin. Chem. Biol., 2016, 35, 109–116.
- 10 X.-G. Zhu, S. P. Long and D. R. Ort, Annu. Rev. Plant Biol., 2010, 61, 235–261.
- 11 R. E. Blankenship, D. M. Tiede, J. Barber, G. W. Brudvig, G. Fleming, M. Ghirardi, M. R. Gunner, W. Junge, D. M. Kramer, A. Melis, T. A. Moore, C. C. Moser, D. G. Nocera, A. J. Nozik, D. R. Ort, W. W. Parson, R. C. Prince and R. T. Sayre, *Science*, 2011, 332, 805–809.
- 12 X.-G. Zhu, S. P. Long and D. R. Ort, *Curr. Opin. Biotechnol.*, 2008, **19**, 153–159.
- 13 R. Croce and H. van Amerongen, *Nat. Chem. Biol.*, 2014, 10, 492–501.
- 14 D. J. Lea-Smith, P. Bombelli, R. Vasudevan and C. J. Howe, *Biochim. Biophys. Acta, Bioenerg.*, 2016, **1857**, 247–255.
- 15 S. Rumpel, J. F. Siebel, C. Farès, J. Duan, E. Reijerse, T. Happe, W. Lubitz and M. Winkler, *Energy Environ. Sci.*, 2014, 7, 3296–3301.
- 16 S. I. Allakhverdiev, V. Thavasi, V. D. Kreslavski, S. K. Zharmukhamedov, V. V. Klimov, S. Ramakrishna, D. A. Los, M. Mimuro, H. Nishihara and R. Carpentier, *J. Photochem. Photobiol.*, C, 2010, 11, 101–113.
- 17 J. Benemann, Nat. Biotechnol., 1996, 14, 1101-1103.
- 18 C. E. Lubner, A. M. Applegate, P. Knörzer, A. Ganago, D. A. Bryant, T. Happe and J. H. Golbeck, *Proc. Natl. Acad. Sci. U. S. A.*, 2011, 108, 20988–20991.
- 19 H. Bothe, O. Schmitz, M. G. Yates and W. E. Newton, *Microbiol. Mol. Biol. Rev.*, 2010, 74, 529–551.
- 20 A. M. Muro-Pastor and W. R. Hess, *Trends Microbiol.*, 2012, 20, 548–557.
- 21 P. Fay, Microbiol. Rev., 1992, 56, 340-373.

- 22 S. Chen, T. Takata and K. Domen, *Nat. Rev. Mater.*, 2017, 2, 17050.
- 23 Q. Wang and K. Domen, Chem. Rev., 2020, 120, 919-985.
- 24 D. Kim, K. K. Sakimoto, D. Hong and P. Yang, Angew. Chem., Int. Ed., 2015, 54, 3259–3266.
- 25 J. K. Hurst, Science, 2010, 328, 315.
- 26 M. G. Walter, E. L. Warren, J. R. McKone, S. W. Boettcher, Q. Mi, E. A. Santori and N. S. Lewis, *Chem. Rev.*, 2010, 110, 6446–6473.
- 27 S. Chen and L.-W. Wang, *Chem. Mater.*, 2012, 24, 3659–3666.
- 28 Q. Wang, T. Hisatomi, Q. Jia, H. Tokudome, M. Zhong, C. Wang, Z. Pan, T. Takata, M. Nakabayashi, N. Shibata, Y. Li, I. D. Sharp, A. Kudo, T. Yamada and K. Domen, *Nat. Mater.*, 2016, 15, 611–615.
- 29 R. Kobayashi, T. Takashima, S. Tanigawa, S. Takeuchi, B. Ohtani and H. Irie, *Phys. Chem. Chem. Phys.*, 2016, 18, 27754–27760.
- 30 K. Maeda, M. Higashi, D. Lu, R. Abe and K. Domen, *J. Am. Chem. Soc.*, 2010, **132**, 5858–5868.
- 31 A. Iwase, Y. H. Ng, Y. Ishiguro, A. Kudo and R. Amal, *J. Am. Chem. Soc.*, 2011, **133**, 11054–11057.
- 32 Y. Sasaki, H. Kato and A. Kudo, *J. Am. Chem. Soc.*, 2013, 135, 5441–5449.
- 33 S. Keene, R. Bala Chandran and S. Ardo, *Energy Environ. Sci.*, 2019, 12, 261–272.
- 34 D. M. Fabian, S. Hu, N. Singh, F. A. Houle, T. Hisatomi, K. Domen, F. E. Osterloh and S. Ardo, *Energy Environ. Sci.*, 2015, **8**, 2825–2850.
- 35 K. Sivula and R. van de Krol, *Nat. Rev. Mater.*, 2016, 1, 15010.
- 36 J. Luo, J.-H. Im, M. T. Mayer, M. Schreier, M. K. Nazeeruddin, N.-G. Park, S. D. Tilley, H. J. Fan and M. Grätzel, *Science*, 2014, 345, 1593–1596.
- 37 C. R. Cox, J. Z. Lee, D. G. Nocera and T. Buonassisi, *Proc. Natl. Acad. Sci. U. S. A.*, 2014, 111, 14057–14061.
- 38 E. Pastor, F. Le Formal, M. T. Mayer, S. D. Tilley, L. Francàs, C. A. Mesa, M. Grätzel and J. R. Durrant, *Nat. Commun.*, 2017, 8, 14280.
- 39 O. Zandi and T. W. Hamann, Nat. Chem., 2016, 8, 778-783.
- 40 K. J. Lee, N. Elgrishi, B. Kandemir and J. L. Dempsey, *Nat. Rev. Chem.*, 2017, **1**, 0039.
- 41 A. Landman, H. Dotan, G. E. Shter, M. Wullenkord, A. Houaijia, A. Maljusch, G. S. Grader and A. Rothschild, *Nat. Mater.*, 2017, **16**, 646–651.
- 42 A. Vilanova, T. Lopes, C. Spenke, M. Wullenkord and A. Mendes, *Energy Storage Mater.*, 2018, **13**, 175–188.
- 43 M. R. Singh, K. Papadantonakis, C. Xiang and N. S. Lewis, *Energy Environ. Sci.*, 2015, **8**, 2760–2767.
- 44 J. Jin, K. Walczak, M. R. Singh, C. Karp, N. S. Lewis and C. Xiang, *Energy Environ. Sci.*, 2014, 7, 3371–3380.
- 45 A. D. Handoko, F. Wei, Jenndy, B. S. Yeo and Z. W. Seh, *Nat. Catal.*, 2018, **1**, 922–934.
- 46 A. M. Appel, J. E. Bercaw, A. B. Bocarsly, H. Dobbek, D. L. DuBois, M. Dupuis, J. G. Ferry, E. Fujita, R. Hille, P. J. A. Kenis, C. A. Kerfeld, R. H. Morris, C. H. F. Peden,

A. R. Portis, S. W. Ragsdale, T. B. Rauchfuss, J. N. H. Reek, L. C. Seefeldt, R. K. Thauer and G. L. Waldrop, Chem. Rev., 2013, 113, 6621-6658.

Review Article

- 47 X. Li, J. Wen, J. Low, Y. Fang and J. Yu, Sci. China Mater., 2014, 57, 70-100.
- 48 X. Chang, T. Wang and J. Gong, Energy Environ. Sci., 2016, 9, 2177-2196.
- 49 R. Kortlever, J. Shen, K. J. P. Schouten, F. Calle-Vallejo and M. T. M. Koper, J. Phys. Chem. Lett., 2015, 6, 4073-4082.
- 50 X. Li, J. Yu, M. Jaroniec and X. Chen, Chem. Rev., 2019, 119, 3962-4179.
- 51 S. L. Foster, S. I. P. Bakovic, R. D. Duda, S. Maheshwari, R. D. Milton, S. D. Minteer, M. J. Janik, J. N. Renner and L. F. Greenlee, Nat. Catal., 2018, 1, 490-500.
- 52 J. G. Chen, R. M. Crooks, L. C. Seefeldt, K. L. Bren, R. M. Bullock, M. Y. Darensbourg, P. L. Holland, B. Hoffman, M. J. Janik, A. K. Jones, M. G. Kanatzidis, P. King, K. M. Lancaster, S. V. Lymar, P. Pfromm, W. F. Schneider and R. R. Schrock, Science, 2018, 360, eaar6611.
- 53 A. Banerjee, B. D. Yuhas, E. A. Margulies, Y. Zhang, Y. Shim, M. R. Wasielewski and M. G. Kanatzidis, J. Am. Chem. Soc., 2015, 137, 2030-2034.
- 54 J. Liu, M. S. Kelley, W. Wu, A. Banerjee, A. P. Douvalis, J. Wu, Y. Zhang, G. C. Schatz and M. G. Kanatzidis, Proc. Natl. Acad. Sci. U. S. A., 2016, 113, 5530-5535.
- 55 H. Hirakawa, M. Hashimoto, Y. Shiraishi and T. Hirai, J. Am. Chem. Soc., 2017, 139, 10929-10936.
- 56 A. J. Medford and M. C. Hatzell, ACS Catal., 2017, 7, 2624-2643.
- 57 C. J. M. van der Ham, M. T. M. Koper and D. G. H. Hetterscheid, Chem. Soc. Rev., 2014, 43, 5183-5191.
- 58 S. Z. Andersen, V. Čolić, S. Yang, J. A. Schwalbe, A. C. Nielander, J. M. McEnaney, K. Enemark-Rasmussen, J. G. Baker, A. R. Singh, B. A. Rohr, M. J. Statt, S. J. Blair, S. Mezzavilla, J. Kibsgaard, P. C. K. Vesborg, M. Cargnello, S. F. Bent, T. F. Jaramillo, I. E. L. Stephens, J. K. Nørskov and I. Chorkendorff, Nature, 2019, 570, 504-508.
- 59 F. A. Armstrong and J. Hirst, Proc. Natl. Acad. Sci. U. S. A., 2011, 108, 14049-14054.
- 60 C. Wombwell and E. Reisner, Chem. Eur. J., 2015, 21, 8096-8104.
- 61 C. Zhang, C. Chen, H. Dong, J.-R. Shen, H. Dau and J. Zhao, Science, 2015, 348, 690-693.
- 62 D. Brazzolotto, M. Gennari, N. Queyriaux, T. R. Simmons, J. Pécaut, S. Demeshko, F. Meyer, M. Orio, V. Artero and C. Duboc, Nat. Chem., 2016, 8, 1054–1060.
- 63 J. W. Lee, D. Na, J. M. Park, J. Lee, S. Choi and S. Y. Lee, Nat. Chem. Biol., 2012, 8, 536-546.
- 64 J. Nielsen and Jay D. Keasling, Cell, 2016, 164, 1185-1197.
- 65 H. Kasap, D. S. Achilleos, A. Huang and E. Reisner, J. Am. Chem. Soc., 2018, 140, 11604-11607.
- 66 H. Kasap, C. A. Caputo, B. C. M. Martindale, R. Godin, V. W.-h. Lau, B. V. Lotsch, J. R. Durrant and E. Reisner, J. Am. Chem. Soc., 2016, 138, 9183-9192.
- 67 D. W. Wakerley, M. F. Kuehnel, K. L. Orchard, K. H. Ly, T. E. Rosser and E. Reisner, Nat. Energy, 2017, 2, 17021.

- 68 T. Uekert, M. F. Kuehnel, D. W. Wakerley and E. Reisner, Energy Environ. Sci., 2018, 11, 2853-2857.
- 69 T. Uekert, H. Kasap and E. Reisner, J. Am. Chem. Soc., 2019, **141**, 15201-15210.
- 70 K. E. Dalle, J. Warnan, J. J. Leung, B. Reuillard, I. S. Karmel and E. Reisner, Chem. Rev., 2019, 119, 2752-2875.
- 71 N. Cox, M. Retegan, F. Neese, D. A. Pantazis, A. Boussac and W. Lubitz, Science, 2014, 345, 804-808.
- 72 J. Jia, L. C. Seitz, J. D. Benck, Y. Huo, Y. Chen, J. W. D. Ng, T. Bilir, J. S. Harris and T. F. Jaramillo, Nat. Commun., 2016, 7, 13237.
- 73 X. Zhou, R. Liu, K. Sun, Y. Chen, E. Verlage, S. A. Francis, N. S. Lewis and C. Xiang, ACS Energy Lett., 2016, 1, 764–770.
- 74 L. Buzzetti, G. E. M. Crisenza and P. Melchiorre, Angew. Chem., Int. Ed., 2019, 58, 3730-3747.
- 75 N. Heidary, K. H. Ly and N. Kornienko, Nano Lett., 2019, 19, 4817-4826.
- 76 N. Kornienko, N. Heidary, G. Cibin and E. Reisner, Chem. Sci., 2018, 9, 5322-5333.
- 77 N. Kornienko, J. Z. Zhang, K. K. Sakimoto, P. Yang and E. Reisner, Nat. Nanotechnol., 2018, 13, 890-899.
- 78 Y. Chen, P. Li, J. Zhou, C. T. Buru, L. Đorđević, P. Li, X. Zhang, M. M. Cetin, J. F. Stoddart, S. I. Stupp, M. R. Wasielewski and O. K. Farha, J. Am. Chem. Soc., 2020, 142, 1768-1773.
- 79 N. Plumeré, O. Rüdiger, A. A. Oughli, R. Williams, J. Vivekananthan, S. Pöller, W. Schuhmann and W. Lubitz, Nat. Chem., 2014, 6, 822-827.
- 80 H. Li, D. Buesen, S. Dementin, C. Léger, V. Fourmond and N. Plumeré, J. Am. Chem. Soc., 2019, 141, 16734-16742.
- 81 A. Heller, Acc. Chem. Res., 1990, 23, 128-134.
- 82 C. Léger, S. J. Elliott, K. R. Hoke, L. J. C. Jeuken, A. K. Jones and F. A. Armstrong, Biochemistry, 2003, 42, 8653-8662.
- 83 M. Can, F. A. Armstrong and S. W. Ragsdale, Chem. Rev., 2014, 114, 4149-4174.
- 84 W. Lubitz, H. Ogata, O. Rüdiger and E. Reijerse, Chem. Rev., 2014, 114, 4081-4148.
- 85 B. Mondal, J. Song, F. Neese and S. Ye, Curr. Opin. Chem. Biol., 2015, 25, 103-109.
- 86 B. M. Hoffman, D. Lukoyanov, Z.-Y. Yang, D. R. Dean and L. C. Seefeldt, Chem. Rev., 2014, 114, 4041-4062.
- 87 W. J. Albery and J. R. Knowles, Biochemistry, 1976, 15, 5631-5640.
- 88 A. Warshel, Proc. Natl. Acad. Sci. U. S. A., 1978, 75, 5250-5254.
- 89 W. Lubitz, M. Chrysina and N. Cox, Photosynth. Res., 2019, 142, 105-125.
- 90 M. Kato, J. Z. Zhang, N. Paul and E. Reisner, Chem. Soc. Rev., 2014, 43, 6485-6497.
- 91 T. Cardona, A. Sedoud, N. Cox and A. W. Rutherford, Biochim. Biophys. Acta, Bioenerg., 2012, 1817, 26-43.
- 92 S. De Causmaecker, J. S. Douglass, A. Fantuzzi, W. Nitschke and A. W. Rutherford, Proc. Natl. Acad. Sci. U. S. A., 2019, 116, 19458-19463.
- 93 J. Barber, Chem. Soc. Rev., 2009, 38, 185-196.
- 94 M. M. Najafpour, T. Ehrenberg, M. Wiechen and P. Kurz, Angew. Chem., Int. Ed., 2010, 49, 2233-2237.

S. Ye, C. Ding, R. Chen, F. Fan, P. Fu, H. Yin, X. Wang, Z. Wang,
 P. Du and C. Li, J. Am. Chem. Soc., 2018, 140, 3250–3256.

Chem Soc Rev

- 96 D. Wang, R. N. Sampaio, L. Troian-Gautier, S. L. Marquard, B. H. Farnum, B. D. Sherman, M. V. Sheridan, C. J. Dares, G. J. Meyer and T. J. Meyer, *J. Am. Chem. Soc.*, 2019, 141, 7926–7933.
- 97 M. Suga, F. Akita, K. Hirata, G. Ueno, H. Murakami, Y. Nakajima, T. Shimizu, K. Yamashita, M. Yamamoto, H. Ago and J.-R. Shen, *Nature*, 2015, 517, 99–103.
- 98 X. Wei, X. Su, P. Cao, X. Liu, W. Chang, M. Li, X. Zhang and Z. Liu, *Nature*, 2016, 534, 69–74.
- 99 J. Kern, R. Chatterjee, I. D. Young, F. D. Fuller, L. Lassalle, M. Ibrahim, S. Gul, T. Fransson, A. S. Brewster, R. Alonso-Mori, R. Hussein, M. Zhang, L. Douthit, C. de Lichtenberg, M. H. Cheah, D. Shevela, J. Wersig, I. Seuffert, D. Sokaras, E. Pastor, C. Weninger, T. Kroll, R. G. Sierra, P. Aller, A. Butryn, A. M. Orville, M. Liang, A. Batyuk, J. E. Koglin, S. Carbajo, S. Boutet, N. W. Moriarty, J. M. Holton, H. Dobbek, P. D. Adams, U. Bergmann, N. K. Sauter, A. Zouni, J. Messinger, J. Yano and V. K. Yachandra, Nature, 2018, 563, 421–425.
- 100 J. Z. Zhang and E. Reisner, Nat. Rev. Chem., 2020, 4, 6-21.
- 101 A. Agostiano, A. Ceglie and M. D. Monica, *Bioelectrochem. Bioenerg.*, 1984, **12**, 499–507.
- 102 X. Fang, K. P. Sokol, N. Heidary, T. A. Kandiel, J. Z. Zhang and E. Reisner, *Nano Lett.*, 2019, 19, 1844–1850.
- 103 K. Brinkert, F. Le Formal, X. Li, J. Durrant, A. W. Rutherford and A. Fantuzzi, *Biochim. Biophys. Acta, Bioenerg.*, 2016, **1857**, 1497–1505.
- 104 M. Kato, T. Cardona, A. W. Rutherford and E. Reisner, J. Am. Chem. Soc., 2012, 134, 8332–8335.
- 105 J. Maly, C. Di Meo, M. De Francesco, A. Masci, J. Masojidek, M. Sugiura, A. Volpe and R. Pilloton, *Bioelectrochemistry*, 2004, **63**, 271–275.
- 106 J. Maly, J. Krejci, M. Ilie, L. Jakubka, J. Masojídek, R. Pilloton, K. Sameh, P. Steffan, Z. Stryhal and M. Sugiura, Anal. Bioanal. Chem., 2005, 381, 1558–1567.
- 107 A. Badura, B. Esper, K. Ataka, C. Grunwald, C. Wöll, J. Kuhlmann, J. Heberle and M. Rögner, *Photochem. Photo-biol.*, 2006, 82, 1385–1390.
- 108 N. Terasaki, M. Iwai, N. Yamamoto, T. Hiraga, S. Yamada and Y. Inoue, *Thin Solid Films*, 2008, **516**, 2553–2557.
- 109 A. Ruff, Curr. Opin. Electrochem., 2017, 5, 66-73.
- 110 A. Badura, D. Guschin, B. Esper, T. Kothe, S. Neugebauer, W. Schuhmann and M. Rögner, *Electroanalysis*, 2008, 20, 1043–1047.
- 111 M. Kato, T. Cardona, A. W. Rutherford and E. Reisner, *J. Am. Chem. Soc.*, 2013, **135**, 10610–10613.
- 112 K. K. Rao, D. O. Hall, N. Vlachopoulos, M. Grätzel, M. C. W. Evans and M. Seibert, *J. Photochem. Photobiol.*, *B*, 1990, **5**, 379–389.
- 113 W. Wang, Z. Wang, Q. Zhu, G. Han, C. Ding, J. Chen, J.-R. Shen and C. Li, *Chem. Commun.*, 2015, **51**, 16952–16955.
- 114 J. Li, X. Feng, J. Fei, P. Cai, J. Huang and J. Li, *J. Mater. Chem. A*, 2016, 4, 12197–12204.

- 115 J. Li, X. Feng, Y. Jia, Y. Yang, P. Cai, J. Huang and J. Li, *J. Mater. Chem. A*, 2017, 5, 19826–19835.
- 116 B. Reuillard, K. H. Ly, P. Hildebrandt, L. J. C. Jeuken, J. N. Butt and E. Reisner, *J. Am. Chem. Soc.*, 2017, **139**, 3324–3327.
- 117 C.-Y. Lee, B. Reuillard, K. P. Sokol, T. Laftsoglou, C. W. J. Lockwood, S. F. Rowe, E. T. Hwang, J. C. Fontecilla-Camps, L. J. C. Jeuken, J. N. Butt and E. Reisner, *Chem. Commun.*, 2016, 52, 7390–7393.
- 118 N. Kornienko, J. Z. Zhang, K. P. Sokol, S. Lamaison, A. Fantuzzi, R. van Grondelle, A. W. Rutherford and E. Reisner, *J. Am. Chem. Soc.*, 2018, **140**, 17923–17931.
- 119 K. P. Sokol, W. E. Robinson, J. Warnan, N. Kornienko, M. M. Nowaczyk, A. Ruff, J. Z. Zhang and E. Reisner, *Nat. Energy*, 2018, 3, 944–951.
- 120 D. Mersch, C.-Y. Lee, J. Z. Zhang, K. Brinkert, J. C. Fontecilla-Camps, A. W. Rutherford and E. Reisner, *J. Am. Chem. Soc.*, 2015, **137**, 8541–8549.
- 121 K. P. Sokol, D. Mersch, V. Hartmann, J. Z. Zhang, M. M. Nowaczyk, M. Rögner, A. Ruff, W. Schuhmann, N. Plumeré and E. Reisner, *Energy Environ. Sci.*, 2016, 9, 3698–3709.
- 122 K. P. Sokol, W. E. Robinson, A. R. Oliveira, J. Warnan, M. M. Nowaczyk, A. Ruff, I. A. C. Pereira and E. Reisner, J. Am. Chem. Soc., 2018, 140, 16418–16422.
- 123 K. R. Stieger, S. C. Feifel, H. Lokstein, M. Hejazi, A. Zouni and F. Lisdat, *J. Mater. Chem. A*, 2016, 4, 17009–17017.
- 124 D. Ciornii, M. Riedel, K. R. Stieger, S. C. Feifel, M. Hejazi, H. Lokstein, A. Zouni and F. Lisdat, J. Am. Chem. Soc., 2017, 139, 16478-16481.
- 125 M. Riedel, W. J. Parak, A. Ruff, W. Schuhmann and F. Lisdat, *ACS Catal.*, 2018, **8**, 5212–5220.
- 126 M. Riedel, J. Wersig, A. Ruff, W. Schuhmann, A. Zouni and F. Lisdat, *Angew. Chem., Int. Ed.*, 2019, **58**, 801–805.
- 127 D. Ciornii, A. Kölsch, A. Zouni and F. Lisdat, *Nanoscale*, 2019, **11**, 15862–15870.
- 128 J. Z. Zhang, P. Bombelli, K. P. Sokol, A. Fantuzzi, A. W. Rutherford, C. J. Howe and E. Reisner, *J. Am. Chem. Soc.*, 2018, **140**, 6–9.
- 129 T. Wenzel, D. Härtter, P. Bombelli, C. J. Howe and U. Steiner, *Nat. Commun.*, 2018, **9**, 1299.
- 130 J. Z. Zhang, K. P. Sokol, N. Paul, E. Romero, R. van Grondelle and E. Reisner, *Nat. Chem. Biol.*, 2016, 12, 1046–1052.
- 131 M. Pita, D. M. Mate, D. Gonzalez-Perez, S. Shleev, V. M. Fernandez, M. Alcalde and A. L. De Lacey, *J. Am. Chem. Soc.*, 2014, 136, 5892–5895.
- 132 C. Tapia, S. Shleev, J. Conesa, A. L. De Lacey and M. Pita, *ACS Catal.*, 2017, 7, 4881–4889.
- 133 S. V. Hexter, F. Grey, T. Happe, V. Climent and F. A. Armstrong, *Proc. Natl. Acad. Sci. U. S. A.*, 2012, **109**, 11516–11521.
- 134 A. K. Jones, E. Sillery, S. P. J. Albracht and F. A. Armstrong, *Chem. Commun.*, 2002, 866–867.
- 135 K. A. Vincent, A. Parkin and F. A. Armstrong, *Chem. Rev.*, 2007, **107**, 4366-4413.

136 E. Reisner, Eur. J. Inorg. Chem., 2011, 1005-1016.

Review Article

- 137 C. C. Page, C. C. Moser, X. Chen and P. L. Dutton, Nature, 1999, **402**, 47-52.
- 138 C. Wombwell, C. A. Caputo and E. Reisner, Acc. Chem. Res., 2015, 48, 2858-2865.
- 139 F. A. Armstrong, N. A. Belsey, J. A. Cracknell, G. Goldet, A. Parkin, E. Reisner, K. A. Vincent and A. F. Wait, Chem. Soc. Rev., 2009, 38, 36-51.
- 140 A. Parkin, G. Goldet, C. Cavazza, J. C. Fontecilla-Camps and F. A. Armstrong, J. Am. Chem. Soc., 2008, 130, 13410-13416.
- 141 M. C. Marques, C. Tapia, O. Gutiérrez-Sanz, A. R. Ramos, K. L. Keller, J. D. Wall, A. L. De Lacey, P. M. Matias and I. A. C. Pereira, Nat. Chem. Biol., 2017, 13, 544-550.
- 142 I. Okura, Coord. Chem. Rev., 1985, 68, 53-99.
- 143 I. Okura, N. Kaji, S. Aono, T. Kita and A. Yamada, Inorg. Chem., 1985, 24, 451-453.
- 144 T. Sakai, D. Mersch and E. Reisner, Angew. Chem., Int. Ed., 2013, 52, 12313-12316.
- 145 E. Reisner, D. J. Powell, C. Cavazza, J. C. Fontecilla-Camps and F. A. Armstrong, J. Am. Chem. Soc., 2009, 131, 18457-18466.
- 146 E. Reisner, J. C. Fontecilla-Camps and F. A. Armstrong, Chem. Commun., 2009, 550-552.
- 147 J. Warnan, J. Willkomm, J. N. Ng, R. Godin, S. Prantl, J. R. Durrant and E. Reisner, Chem. Sci., 2017, 8, 3070-3079.
- 148 K. A. Brown, S. Dayal, X. Ai, G. Rumbles and P. W. King, J. Am. Chem. Soc., 2010, 132, 9672-9680.
- 149 K. A. Brown, M. B. Wilker, M. Boehm, G. Dukovic and P. W. King, J. Am. Chem. Soc., 2012, 134, 5627-5636.
- 150 M. B. Wilker, K. E. Shinopoulos, K. A. Brown, D. W. Mulder, P. W. King and G. Dukovic, J. Am. Chem. Soc., 2014, 136, 4316-4324.
- 151 K. Wu, H. Zhu, Z. Liu, W. Rodríguez-Córdoba and T. Lian, J. Am. Chem. Soc., 2012, 134, 10337-10340.
- 152 C. A. Caputo, M. A. Gross, V. W. Lau, C. Cavazza, B. V. Lotsch and E. Reisner, Angew. Chem., Int. Ed., 2014, 53, 11538-11542.
- 153 C. A. Caputo, L. Wang, R. Beranek and E. Reisner, Chem. Sci., 2015, 6, 5690-5694.
- 154 G. A. M. Hutton, B. Reuillard, B. C. M. Martindale, C. A. Caputo, C. W. J. Lockwood, J. N. Butt and E. Reisner, J. Am. Chem. Soc., 2016, 138, 16722-16730.
- 155 O. Rüdiger, J. M. Abad, E. C. Hatchikian, V. M. Fernandez and A. L. De Lacey, J. Am. Chem. Soc., 2005, 127, 16008-16009.
- 156 Y. Park, K. J. McDonald and K.-S. Choi, Chem. Soc. Rev., 2013, 42, 2321-2337.
- 157 T. W. Kim and K.-S. Choi, Science, 2014, 343, 990-994.
- 158 J. D. Benck, B. A. Pinaud, Y. Gorlin and T. F. Jaramillo, PLoS One, 2014, 9, e107942.
- 159 A. Bachmeier, V. C. C. Wang, T. W. Woolerton, S. Bell, J. C. Fontecilla-Camps, M. Can, S. W. Ragsdale, Y. S. Chaudhary and F. A. Armstrong, J. Am. Chem. Soc., 2013, 135, 15026-15032.

- 160 T. E. Rosser and E. Reisner, ACS Catal., 2017, 7, 3131-3141.
- 161 N. M. Muresan, J. Willkomm, D. Mersch, Y. Vaynzof and E. Reisner, Angew. Chem., Int. Ed., 2012, 51, 12749-12753.
- 162 C.-Y. Lee, H. S. Park, J. C. Fontecilla-Camps and E. Reisner, Angew. Chem., Int. Ed., 2016, 55, 5971-5974.
- 163 J. J. Leung, J. Warnan, D. H. Nam, J. Z. Zhang, J. Willkomm and E. Reisner, Chem. Sci., 2017, 8, 5172-5180.
- 164 B. Seger, T. Pedersen, A. B. Laursen, P. C. K. Vesborg, O. Hansen and I. Chorkendorff, J. Am. Chem. Soc., 2013, 135, 1057-1064.
- 165 F. Lakadamyali and E. Reisner, Chem. Commun., 2011, 47, 1695-1697.
- 166 D. H. Nam, J. Z. Zhang, V. Andrei, N. Kornienko, N. Heidary, A. Wagner, K. Nakanishi, K. P. Sokol, B. Slater, I. Zebger, S. Hofmann, J. C. Fontecilla-Camps, C. B. Park and E. Reisner, Angew. Chem., Int. Ed., 2018, 57, 10595-10599.
- 167 Y. Zhao, N. C. Anderson, M. W. Ratzloff, D. W. Mulder, K. Zhu, J. A. Turner, N. R. Neale, P. W. King and H. M. Branz, ACS Appl. Mater. Interfaces, 2016, 8, 14481-14487.
- 168 E. Edwardes Moore, V. Andrei, S. Zacarias, I. A. C. Pereira and E. Reisner, ACS Energy Lett., 2020, 5, 232-237.
- 169 C. E. Lubner, R. Grimme, D. A. Bryant and J. H. Golbeck, Biochemistry, 2010, 49, 404-414.
- 170 H. Krassen, A. Schwarze, B. Friedrich, K. Ataka, O. Lenz and J. Heberle, ACS Nano, 2009, 3, 4055-4061.
- 171 C. Tapia, R. D. Milton, G. Pankratova, S. D. Minteer, H.-E. Åkerlund, D. Leech, A. L. De Lacey, M. Pita and L. Gorton, ChemElectroChem, 2017, 4, 90-95.
- 172 C. E. Lubner, P. Knörzer, P. J. N. Silva, K. A. Vincent, T. Happe, D. A. Bryant and J. H. Golbeck, *Biochemistry*, 2010, 49, 10264-10266.
- 173 F. Zhao, P. Wang, A. Ruff, V. Hartmann, S. Zacarias, I. A. C. Pereira, M. M. Nowaczyk, M. Rögner, F. Conzuelo and W. Schuhmann, Energy Environ. Sci., 2019, 12, 3133-3143.
- 174 P. De Luna, C. Hahn, D. Higgins, S. A. Jaffer, T. F. Jaramillo and E. H. Sargent, Science, 2019, 364, eaav3506.
- 175 G. Fuchs, Annu. Rev. Microbiol., 2011, 65, 631-658.
- 176 S. W. Ragsdale and E. Pierce, Biochim. Biophys. Acta, Proteins Proteomics, 2008, 1784, 1873-1898.
- 177 K. Schuchmann and V. Müller, Nat. Rev. Microbiol., 2014, 12, 809-821.
- 178 C. A. R. Cotton, C. Edlich-Muth and A. Bar-Even, Curr. Opin. Biotechnol., 2018, 49, 49-56.
- 179 S. W. Ragsdale, Crit. Rev. Biochem. Mol. Biol., 2004, 39, 165-195.
- 180 A. Parkin, J. Seravalli, K. A. Vincent, S. W. Ragsdale and F. A. Armstrong, J. Am. Chem. Soc., 2007, 129, 10328-10329.
- 181 J. Fesseler, J.-H. Jeoung and H. Dobbek, Angew. Chem., Int. Ed., 2015, 54, 8560-8564.
- 182 M. W. Ribbe, Angew. Chem., Int. Ed., 2015, 54, 8337-8339.
- 183 T. Reda, C. M. Plugge, N. J. Abram and J. Hirst, Proc. Natl. Acad. Sci. U. S. A., 2008, 105, 10654-10658.
- 184 A. Bassegoda, C. Madden, D. W. Wakerley, E. Reisner and J. Hirst, J. Am. Chem. Soc., 2014, 136, 15473-15476.

- 185 W. E. Robinson, A. Bassegoda, E. Reisner and J. Hirst, *J. Am. Chem. Soc.*, 2017, **139**, 9927–9936.
- 186 T. W. Woolerton, S. Sheard, E. Reisner, E. Pierce, S. W. Ragsdale and F. A. Armstrong, *J. Am. Chem. Soc.*, 2010, **132**, 2132–2133.
- 187 T. W. Woolerton, S. Sheard, E. Pierce, S. W. Ragsdale and F. A. Armstrong, *Energy Environ. Sci.*, 2011, 4, 2393–2399.
- 188 A. Bachmeier, S. Hall, S. W. Ragsdale and F. A. Armstrong, J. Am. Chem. Soc., 2014, 136, 13518–13521.
- 189 S. Ikeyama and Y. Amao, Sustainable Energy Fuels, 2017, 1, 1730–1733.
- 190 Y. Amao, Sustainable Energy Fuels, 2018, 2, 1928-1950.
- 191 Y. Amao, Chem. Lett., 2017, 46, 780-788.

Chem Soc Rev

- 192 B. A. Parkinson and P. F. Weaver, *Nature*, 1984, **309**, 148–149.
- 193 M. Miller, W. E. Robinson, A. R. Oliveira, N. Heidary, N. Kornienko, J. Warnan, I. A. C. Pereira and E. Reisner, Angew. Chem., Int. Ed., 2019, 58, 4601–4605.
- 194 S. K. Kuk, Y. Ham, K. Gopinath, P. Boonmongkolras, Y. Lee, Y. W. Lee, S. Kondaveeti, C. Ahn, B. Shin, J.-K. Lee, S. Jeon and C. B. Park, *Adv. Energy Mater.*, 2019, **9**, 1900029.
- 195 S. Y. Lee, S. Y. Lim, D. Seo, J.-Y. Lee and T. D. Chung, *Adv. Energy Mater.*, 2016, **6**, 1502207.
- 196 S. K. Kuk, J. Jang, J. Kim, Y. Lee, Y. S. Kim, B. Koo, Y. W. Lee, J. W. Ko, B. Shin, J.-K. Lee and C. B. Park, *ChemSusChem*, 2020, DOI: 10.1002/cssc.202000459.
- 197 X. Ji, Z. Su, P. Wang, G. Ma and S. Zhang, *Small*, 2016, **12**, 4753–4762.
- 198 S. K. Kuk, R. K. Singh, D. H. Nam, R. Singh, J.-K. Lee and C. B. Park, *Angew. Chem., Int. Ed.*, 2017, **56**, 3827–3832.
- 199 S. Zhang, J. Shi, Y. Sun, Y. Wu, Y. Zhang, Z. Cai, Y. Chen, C. You, P. Han and Z. Jiang, ACS Catal., 2019, 9, 3913–3925.
- 200 K. A. Brown, D. F. Harris, M. B. Wilker, A. Rasmussen, N. Khadka, H. Hamby, S. Keable, G. Dukovic, J. W. Peters, L. C. Seefeldt and P. W. King, *Science*, 2016, 352, 448–450.
- 201 P. Taylor, S. L. Pealing, G. A. Reid, S. K. Chapman and M. D. Walkinshaw, *Nat. Struct. Biol.*, 1999, 6, 1108–1112.
- 202 A. Bachmeier, B. J. Murphy and F. A. Armstrong, *J. Am. Chem. Soc.*, 2014, **136**, 12876–12879.
- 203 M. Mifsud, S. Gargiulo, S. Iborra, I. W. C. E. Arends, F. Hollmann and A. Corma, *Nat. Commun.*, 2014, 5, 3145.
- 204 J. Kim, S. H. Lee, F. Tieves, D. S. Choi, F. Hollmann, C. E. Paul and C. B. Park, *Angew. Chem., Int. Ed.*, 2018, 57, 13825–13828.
- 205 S. H. Lee, D. S. Choi, M. Pesic, Y. W. Lee, C. E. Paul, F. Hollmann and C. B. Park, *Angew. Chem., Int. Ed.*, 2017, 56, 8681–8685.
- 206 H. Hamby, B. Li, K. E. Shinopoulos, H. R. Keller, S. J. Elliott and G. Dukovic, *Proc. Natl. Acad. Sci. U. S. A.*, 2020, 117, 135–140.
- 207 J. Wencel-Delord and F. Glorius, *Nat. Chem.*, 2013, 5, 369–375.
- 208 N.-H. Tran, D. Nguyen, S. Dwaraknath, S. Mahadevan, G. Chavez, A. Nguyen, T. Dao, S. Mullen, T.-A. Nguyen and L. E. Cheruzel, *J. Am. Chem. Soc.*, 2013, 135, 14484–14487.

- 209 C. M. Agapakis, P. M. Boyle and P. A. Silver, *Nat. Chem. Biol.*, 2012, 8, 527–535.
- 210 M. J. Smanski, H. Zhou, J. Claesen, B. Shen, M. A. Fischbach and C. A. Voigt, *Nat. Rev. Microbiol.*, 2016, 14, 135–149.
- 211 K. K. Sakimoto, N. Kornienko and P. Yang, Acc. Chem. Res., 2017, 50, 476–481.
- 212 T. Haas, R. Krause, R. Weber, M. Demler and G. Schmid, *Nat. Catal.*, 2018, **1**, 32–39.
- 213 S. Cestellos-Blanco, H. Zhang, J. M. Kim, Y.-X. Shen and P. Yang, *Nat. Catal.*, 2020, 3, 245–255.
- 214 M. L. Ghirardi, A. Dubini, J. Yu and P.-C. Maness, *Chem. Soc. Rev.*, 2009, **38**, 52–61.
- 215 Z. Ji, H. Zhang, H. Liu, O. M. Yaghi and P. Yang, *Proc. Natl. Acad. Sci. U. S. A.*, 2018, 115, 10582–10587.
- 216 W. Wei, P. Sun, Z. Li, K. Song, W. Su, B. Wang, Y. Liu and J. Zhao, *Sci. Adv.*, 2018, **4**, eaap9253.
- 217 A. A. Krasnovsky and V. V. Nikandrov, *FEBS Lett.*, 1987, **219**, 93–96.
- 218 P. Maruthamuthu, S. Muthu, K. Gurunathan, M. Ashokkumar and M. V. C. Sastri, *Int. J. Hydrogen Energy*, 1992, 17, 863–866.
- 219 K. Gurunathan, J. Mol. Catal. A: Chem., 2000, 156, 59-67.
- 220 S. Pontrelli, T.-Y. Chiu, E. I. Lan, F. Y. H. Chen, P. Chang and J. C. Liao, *Metab. Eng.*, 2018, 50, 16–46.
- 221 Y. Honda, H. Hagiwara, S. Ida and T. Ishihara, *Angew. Chem., Int. Ed.*, 2016, 55, 8045–8048.
- 222 B. Wang, C. Zeng, K. H. Chu, D. Wu, H. Y. Yip, L. Ye and P. K. Wong, Adv. Energy Mater., 2017, 7, 1700611.
- 223 K. P. Kuhl, E. R. Cave, D. N. Abram and T. F. Jaramillo, *Energy Environ. Sci.*, 2012, 5, 7050–7059.
- 224 R. K. Thauer, A.-K. Kaster, H. Seedorf, W. Buckel and R. Hedderich, *Nat. Rev. Microbiol.*, 2008, **6**, 579–591.
- 225 J. Ye, J. Yu, Y. Zhang, M. Chen, X. Liu, S. Zhou and Z. He, *Appl. Catal.*, *B*, 2019, 257, 117916.
- 226 K. K. Sakimoto, A. B. Wong and P. Yang, *Science*, 2016, **351**, 74–77.
- 227 A. W. D. Larkum, Curr. Opin. Biotechnol., 2010, 21, 271-276.
- 228 N. Kornienko, K. K. Sakimoto, D. M. Herlihy, S. C. Nguyen, A. P. Alivisatos, C. B. Harris, A. Schwartzberg and P. Yang, *Proc. Natl. Acad. Sci. U. S. A.*, 2016, 113, 11750–11755.
- R. Zhang, Y. He, J. Yi, L. Zhang, C. Shen, S. Liu, L. Liu,B. Liu and L. Qiao, *Chem*, 2020, 6, 234–249.
- 230 K. K. Sakimoto, S. J. Zhang and P. Yang, *Nano Lett.*, 2016, 16, 5883–5887.
- 231 H. Zhang, H. Liu, Z. Tian, D. Lu, Y. Yu, S. Cestellos-Blanco, K. K. Sakimoto and P. Yang, *Nat. Nanotechnol.*, 2018, 13, 900–905.
- 232 B. Wang, Z. Jiang, J. C. Yu, J. Wang and P. K. Wong, Nanoscale, 2019, 11, 9296–9301.
- 233 M. Chen, X.-F. Zhou, Y.-Q. Yu, X. Liu, R. J.-X. Zeng, S.-G. Zhou and Z. He, *Environ. Int.*, 2019, 127, 353–360.
- 234 P. Gai, W. Yu, H. Zhao, R. Qi, F. Li, L. Liu, F. Lv and S. Wang, *Angew. Chem., Int. Ed.*, 2020, **59**, 7224–7229.
- 235 S. F. Rowe, G. Le Gall, E. V. Ainsworth, J. A. Davies, C. W. J. Lockwood, L. Shi, A. Elliston, I. N. Roberts, K. W. Waldron, D. J. Richardson, T. A. Clarke, L. J. C. Jeuken, E. Reisner and J. N. Butt, ACS Catal., 2017, 7, 7558–7566.

236 J. H. Park, S. H. Lee, G. S. Cha, D. S. Choi, D. H. Nam,

Review Article

J. H. Lee, J.-K. Lee, C.-H. Yun, K. J. Jeong and C. B. Park, Angew. Chem., Int. Ed., 2015, 54, 969-973.

- 237 X. Wang, T. Saba, H. H. P. Yiu, R. F. Howe, J. A. Anderson and J. Shi, Chem, 2017, 2, 621-654.
- 238 J. Guo, M. Suástegui, K. K. Sakimoto, V. M. Moody, G. Xiao, D. G. Nocera and N. S. Joshi, Science, 2018, 362, 813-816.
- 239 E. M. Nichols, J. J. Gallagher, C. Liu, Y. Su, J. Resasco, Y. Yu, Y. Sun, P. Yang, M. C. Y. Chang and C. J. Chang, Proc. Natl. Acad. Sci. U. S. A., 2015, 112, 11461-11466.
- 240 J. P. Torella, C. J. Gagliardi, J. S. Chen, D. K. Bediako, B. Colón, J. C. Way, P. A. Silver and D. G. Nocera, Proc. Natl. Acad. Sci. U. S. A., 2015, 112, 2337-2342.
- 241 C. Liu, B. C. Colón, M. Ziesack, P. A. Silver and D. G. Nocera, Science, 2016, 352, 1210-1213.
- 242 D. K. Dogutan and D. G. Nocera, Acc. Chem. Res., 2019, 52, 3143-3148.
- 243 C. Liu, K. K. Sakimoto, B. C. Colón, P. A. Silver and D. G. Nocera, Proc. Natl. Acad. Sci. U. S. A., 2017, 114, 6450-6455.
- 244 R. M. Rodrigues, X. Guan, J. A. Iñiguez, D. A. Estabrook, J. O. Chapman, S. Huang, E. M. Sletten and C. Liu, Nat. Catal., 2019, 2, 407-414.
- 245 K. Rabaey and R. Rozendal, Nat. Rev. Microbiol., 2010, 8, 706-716.
- 246 D. R. Lovley, Environ. Microbiol. Rep., 2011, 3, 27-35.
- 247 A. J. McCormick, P. Bombelli, R. W. Bradley, R. Thorne, T. Wenzel and C. J. Howe, Energy Environ. Sci., 2015, 8, 1092-1109.
- 248 C. Liu, J. J. Gallagher, K. K. Sakimoto, E. M. Nichols, C. J. Chang, M. C. Y. Chang and P. Yang, Nano Lett., 2015, 15, 3634-3639.
- 249 F. Qian, H. Wang, Y. Ling, G. Wang, M. P. Thelen and Y. Li, Nano Lett., 2014, 14, 3688-3693.
- 250 Y. Su, S. Cestellos-Blanco, J. M. Kim, Y.-x. Shen, Q. Kong, D. Lu, C. Liu, H. Zhang, Y. Cao and P. Yang, Joule, 2020, 4, 800-811.
- 251 S. Li, C. Cheng and A. Thomas, Adv. Mater., 2017, 29, 1602547.
- 252 K. K. Sakimoto, N. Kornienko, S. Cestellos-Blanco, J. Lim, C. Liu and P. Yang, J. Am. Chem. Soc., 2018, 140, 1978–1985.
- 253 V. Flexer and L. Jourdin, Acc. Chem. Res., 2020, 53, 311–321.
- 254 K. B. Gregory, D. R. Bond and D. R. Lovley, Environ. Microbiol., 2004, 6, 596-604.
- 255 S. M. Strycharz, R. H. Glaven, M. V. Coppi, S. M. Gannon, L. A. Perpetua, A. Liu, K. P. Nevin and D. R. Lovley, Bioelectrochemistry, 2011, 80, 142-150.
- 256 D. E. Ross, J. M. Flynn, D. B. Baron, J. A. Gralnick and D. R. Bond, PLoS One, 2011, 6, e16649.
- 257 Y.-C. Yong, Y.-Y. Yu, X. Zhang and H. Song, Angew. Chem., Int. Ed., 2014, 53, 4480-4483.
- 258 B. E. Logan and K. Rabaey, Science, 2012, 337, 686-690.
- 259 K. B. Gregory and D. R. Lovley, Environ. Sci. Technol., 2005, 39, 8943-8947.
- 260 S. M. Strycharz, T. L. Woodard, J. P. Johnson, K. P. Nevin, R. A. Sanford, F. E. Löffler and D. R. Lovley, Appl. Environ. Microbiol., 2008, 74, 5943-5947.

- 261 S. M. Strycharz, S. M. Gannon, A. R. Boles, A. E. Franks, K. P. Nevin and D. R. Lovley, Environ. Microbiol. Rep., 2010, 2, 289-294.
- 262 X. Fang, S. Kalathil, G. Divitini, Q. Wang and E. Reisner, Proc. Natl. Acad. Sci. U. S. A., 2020, 117, 5074-5080.
- 263 N. Heidary, N. Kornienko, S. Kalathil, X. Fang, K. H. Ly, H. F. Greer and E. Reisner, J. Am. Chem. Soc., 2020, 142, 5194-5203.
- 264 N. J. Claassens, D. Z. Sousa, V. A. P. M. dos Santos, W. M. de Vos and J. van der Oost, Nat. Rev. Microbiol., 2016, 14, 692-706.
- 265 J. S. Deutzmann and A. M. Spormann, ISME J., 2016, 11, 704-714.
- 266 P. T. Ha, S. R. Lindemann, L. Shi, A. C. Dohnalkova, J. K. Fredrickson, M. T. Madigan and H. Beyenal, Nat. Commun., 2017, 8, 13924.
- 267 B. Schiel-Bengelsdorf and P. Dürre, FEBS Lett., 2012, 586, 2191-2198.
- 268 C. M. Humphreys and N. P. Minton, Curr. Opin. Biotechnol., 2018, 50, 174-181.
- 269 N. J. Claassens, C. A. R. Cotton, D. Kopljar and A. Bar-Even, Nat. Catal., 2019, 2, 437-447.
- 270 F. Rudroff, M. D. Mihovilovic, H. Gröger, R. Snajdrova, H. Iding and U. T. Bornscheuer, Nat. Catal., 2018, 1, 12-22.
- 271 N. Antonovsky, S. Gleizer, E. Noor, Y. Zohar, E. Herz, U. Barenholz, L. Zelcbuch, S. Amram, A. Wides, N. Tepper, D. Davidi, Y. Bar-On, T. Bareia, D. G. Wernick, I. Shani, S. Malitsky, G. Jona, A. Bar-Even and R. Milo, Cell, 2016, 166, 115-125.
- 272 K. Köninger, A. Gómez Baraibar, C. Mügge, C. E. Paul, F. Hollmann, M. M. Nowaczyk and R. Kourist, Angew. Chem., Int. Ed., 2016, 55, 5582-5585.
- 273 J. Ni, H.-Y. Liu, F. Tao, Y.-T. Wu and P. Xu, Angew. Chem., Int. Ed., 2018, 57, 15990-15994.
- 274 M. A. TerAvest and C. M. Ajo-Franklin, Biotechnol. Bioeng., 2016, 113, 687-697.
- 275 M. A. TerAvest, T. J. Zajdel and C. M. Ajo-Franklin, Chem-ElectroChem, 2014, 1, 1874–1879.
- 276 D. Choi, S. B. Lee, S. Kim, B. Min, I.-G. Choi and I. S. Chang, Bioresour. Technol., 2014, 154, 59-66.
- 277 Y. Chen, P. Li, H. Noh, C.-W. Kung, C. T. Buru, X. Wang, X. Zhang and O. K. Farha, Angew. Chem., Int. Ed., 2019, 58, 7682-7686.
- 278 I. J. Iwuchukwu, M. Vaughn, N. Myers, H. O'Neill, P. Frymier and B. D. Bruce, Nat. Nanotechnol., 2009, 5, 73-79.
- 279 Y. V. Lee and B. Tian, Nano Lett., 2019, 19, 2189-2197.
- 280 G. Fan, C. M. Dundas, A. J. Graham, N. A. Lynd and B. K. Keitz, Proc. Natl. Acad. Sci. U. S. A., 2018, 115, 4559-4564.
- 281 T. Song, H. Zhang, H. Liu, D. Zhang, H. Wang, Y. Yang, H. Yuan and J. Xie, Bioresour. Technol., 2017, 243, 573-582.
- 282 C. M. Dundas, A. J. Graham, D. K. Romanovicz and B. K. Keitz, ACS Synth. Biol., 2018, 7, 2726-2736.
- 283 B. C. M. Martindale, G. A. M. Hutton, C. A. Caputo and E. Reisner, J. Am. Chem. Soc., 2015, 137, 6018-6025.

284 N. Kornienko, K. H. Ly, W. E. Robinson, N. Heidary, J. Z. Zhang and E. Reisner, *Acc. Chem. Res.*, 2019, 52, 1439–1448.

Chem Soc Rev

- 285 M. del Barrio, M. Sensi, C. Orain, C. Baffert, S. Dementin, V. Fourmond and C. Léger, *Acc. Chem. Res.*, 2018, **51**, 769–777.
- 286 E. Marsili, J. B. Rollefson, D. B. Baron, R. M. Hozalski and D. R. Bond, *Appl. Environ. Microbiol.*, 2008, 74, 7329–7337.
- 287 T. Reda and J. Hirst, J. Phys. Chem. B, 2006, 110, 1394-1404.
- 288 B. Reuillard, K. H. Ly, T. E. Rosser, M. F. Kuehnel, I. Zebger and E. Reisner, *J. Am. Chem. Soc.*, 2017, **139**, 14425–14435.
- 289 J. Willkomm, K. L. Orchard, A. Reynal, E. Pastor, J. R. Durrant and E. Reisner, *Chem. Soc. Rev.*, 2016, 45, 9–23.
- 290 K. Abdiaziz, E. Salvadori, K. P. Sokol, E. Reisner and M. M. Roessler, *Chem. Commun.*, 2019, 55, 8840–8843.

- 291 G. L. Chadwick, F. Jiménez Otero, J. A. Gralnick, D. R. Bond and V. J. Orphan, *Proc. Natl. Acad. Sci. U. S. A.*, 2019, **116**, 20716–20724.
- 292 Z. S. Wang, K. Sayama and H. Sugihara, J. Phys. Chem. B, 2005, 109, 22449–22455.
- 293 R. Beranek, Adv. Phys. Chem., 2011, 786759.
- 294 J. Zhang, X. Chen, K. Takanabe, K. Maeda, K. Domen, J. D. Epping, X. Fu, M. Antonietti and X. Wang, *Angew. Chem., Int. Ed.*, 2010, **49**, 441–444.
- 295 B. C. M. Martindale, E. Joliat, C. Bachmann, R. Alberto and E. Reisner, *Angew. Chem., Int. Ed.*, 2016, 55, 9402–9406.
- 296 M. F. Kuehnel, K. L. Orchard, K. E. Dalle and E. Reisner, *J. Am. Chem. Soc.*, 2017, **139**, 7217–7223.
- 297 D. P. Hari and B. König, *Chem. Commun.*, 2014, **50**, 6688–6699.