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Achievement of visible-light-driven Z-scheme overall water splitting using barium-modified Ta₃N₅ as a H₂-evolving photocatalyst†

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Ta₃N₅ is one of the most promising photocatalyst candidates for solar water splitting, but it still remains challenging to achieve overall water splitting *via* Ta₃N₅-based photocatalysts regardless of whether it uses a one step or two step method. Here we will address the relatively poor photocatalytic proton reduction of Ta₃N₅ with an effort for the promotion of charge separation *via* barium modification. One-pot nitridation of barium nitrate-impregnated Ta₂O₅ precursor was adopted here for the synthesis of Ta₃N₅ accompanied with the creation of a Ta₃N₅/BaTaO₂N heterostructure and surface passivation. Due to the synergetic effect of the improved interfacial charge separation and the decreased defect density, the photocatalytic H₂ evolution rate of barium-modified Ta₃N₅ is effectively promoted. Encouraged by this, a visible-light-driven Z-scheme overall water splitting system was successfully constructed by using the barium-modified Ta₃N₅ as a H₂-evolving photocatalyst, together with a PtO_x/WO₃ and IO₃⁻/I⁻ pair as an O₂-evolving photocatalyst and a redox mediator, respectively.

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Introduction

Semiconductor-based photocatalytic overall water splitting for hydrogen production is an ideal way to convert solar energy to chemical energy and has inspired extensive interest in the past few decades.^{1–5} Towards this, hundreds of semiconductors have been reported for potential solar water splitting, but most of them are only active under UV light irradiation.^{6–10} To achieve highly efficient solar-to-chemical energy conversion, overall water splitting on photocatalysts harvesting visible light with longer wavelength is desirable. To date, however, the number of wide visible-light-driven overall water splitting systems, regardless of whether they use a one step or two step method, is limited.^{5,11–18}

Tantalum nitride (Ta₃N₅), with a theoretical solar-to-hydrogen conversion efficiency of 15.9%, is one of the most promising candidates for solar water splitting, considering its matched band edge positions (conduction band and valence band edges at *ca.* –0.4 V and +1.7 V *vs.* NHE, respectively, at pH = 0), wide visible light harvesting ability (up to 600 nm) and

good photo-stability.^{19–37} It was first synthesized in 1973,³⁸ but was not found to be active for the photocatalytic water splitting reaction until 2002.¹⁹ Afterwards, Ta₃N₅ has been widely investigated for water splitting in terms of particulate photocatalysts^{22–25} and photoanodes.^{26–31}

The increasing research interest and efforts have greatly promoted the water oxidation performance of Ta₃N₅ for both particulate photocatalyst and photoanode systems. For example, Li *et al.* fabricated a 1D Ta₃N₅ nanorod photoanode to achieve a STH of 1.5%.³⁰ Liu *et al.* achieved Ta₃N₅ photoanode stability for hours²⁷ and obtained nearly close to the theoretical photocurrent at a potential of 1.23 V *vs.* RHE under AM 1.5G simulated sunlight.³¹ Chen *et al.* reported that the apparent quantum efficiency of the photocatalytic water oxidation activity of the Ta₃N₅-based particulate photocatalyst can reach 11.3% at 500–600 nm *via* an interface engineering strategy.²⁴ Compared to the water oxidation, however, the activity of photocatalytic proton reduction from water is much lower or even undetectable in most cases, even though extensive investigations such as forming polymorphic macroporous Ta₃N₅, reducing the particle size through templates (*i.e.* SiO₂, C₃N₄) and surface modification have been made.^{32–37} As a result of the poor proton reduction ability, Z-scheme overall water splitting using particulate Ta₃N₅ as a H₂-evolving photocatalyst is still not reported.

Fabricating nanocomposites with another semiconductor to form heterostructures has been extensively adopted for the promotion of photocatalytic performances.^{16,39–42} A heterostructure can create external bias through interfacial junctions to spatially separate the photogenerated electrons and holes.

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However, it should be pointed out that most of the (oxy)nitride photocatalysts are thermally instable in air, so the fabrication of a heterostructure for (oxy)nitride commonly confronts technical challenges, rendering feasible examples very limited.¹⁶

In this work, a barium modification strategy is introduced to address the relatively poor photocatalytic proton reduction activity of Ta₃N₅ under visible light irradiation. A simple one-pot nitridation route was adopted for the synthesis of pristine Ta₃N₅ and barium-modified Ta₃N₅, in which a barium nitrate-impregnated Ta₂O₅ was used as a precursor. It is found that some Ba²⁺ ions could be doped into Ta₃N₅ to decrease its defect density. On the other hand, excessive Ba²⁺ ions will produce BaTaO₂N *in situ* on the surface of Ta₃N₅ to create a Ta₃N₅/BaTaO₂N heterostructure. As a result, the photogenerated carrier separation efficiency of Ta₃N₅ can be promoted after the barium modification, causing an effectively enhanced H₂ evolution rate in the presence of methanol. Finally, the first example of a visible-light-driven photocatalytic Z-scheme overall water splitting system using the modified Ta₃N₅ as a H₂-evolving photocatalyst was successfully constructed.

Experimental

Materials and reagents

For the preparation of Ba(*n*)-Ta₃N₅ samples, Ta₂O₅ (99.99%, Amresco Chemical), and Ba(NO₃)₂ (99.5%, Sinopharm Chemical) were used. WO₃ (99.99%, High Purity Chemical) was used as a water oxidation photocatalyst. H₂PtCl₆·6H₂O (99.5%, Sinopharm Chemical) was employed as the precursor for the reduction cocatalyst. CH₃OH (99.5%, Sinopharm Chemical) and NaI (99.5%, Guangfu Chemical) were used as sacrificial electron donors. La₂O₃ (99.95%, Sinopharm Chemical) was applied as a pH buffer agent. All chemicals were used as purchased without further purification.

Preparation of Ba(*n*)-Ta₃N₅ samples

Typically, Ta₂O₅ powder was impregnated in the Ba(NO₃)₂ aqueous solution with a calculated molar ratio of Ba/Ta, and the dried mixture was then annealed in air at 1073 K for 2 h. The as-prepared powder was treated with “one-pot” nitridation under ammonia flow (250 mL min⁻¹) at 1223 K for 20 h. The as-obtained samples are correspondingly denoted as Ba(*n*)-Ta₃N₅, where “*n*” stands for the molar ratio of Ba/Ta and when *n* = 1 it stands for BaTaO₂N. As a comparison, the pure phase of Ta₃N₅ and BaTaO₂N powder was mechanically mixed at a Ba/Ta molar ratio of 0.3, which is denoted as Ta₃N₅/BaTaO₂N (0.3)-mix.

Deposition of cocatalysts

0.2 g of the as-obtained sample was dispersed in a calculated amount of H₂PtCl₆ aqueous solution, and sonicated for *ca.* 5 min. After the solution was completely evaporated in a water bath at 353 K, the resulting powder was collected and reduced at 473 K for 1 h under a flow of 5% H₂/Ar (200 mL min⁻¹). As for the deposition of PtO_x on the surface of WO₃ for water oxidation, typically, 0.3 g of WO₃ was annealed in the air at 773 K for

2 h, and then 0.2 g of the annealed sample was immersed in a calculated amount of H₂PtCl₆ aqueous solution with sonicating for *ca.* 5 min. After complete evaporation in a water bath at 353 K, the resulting powder was collected and annealed in air at 798 K for 0.5 h.

Electrochemical analysis

For the Mott–Schottky (M–S) measurement, Ta₃N₅ and BaTaO₂N powder were deposited on FTO conducting glass *via* electrochromic deposition (EPD). Typically, the powder samples (50 mg) and iodine (20 mg) were dispersed in acetone solution (50 mL), and continuously sonicated for 10 min. Afterwards, the FTO electrode was immersed, parallel to another FTO electrode, with a distance of about 1 cm. The duration time was 1 min with 20 V and 1 A applied using a potentiostat (ITECH IT6834), and then the prepared electrodes were calcined under an ammonia flow (250 mL min⁻¹) at 723 K for 0.5 h.

The M–S measurement was carried out using a Princeton Applied Research PARSTAT 2273, using 0.5 M Na₂SO₄ aqueous solution as electrolyte with a pH value of 8.5 adjusted using NaOH. The frequency was 1 kHz.

Characterizations of catalysts

XRD measurements were carried out using a Rigaku D/Max-2500/PC powder diffractometer (Cu K α radiation) with an operating voltage of 40 kV and an operating current of 200 mA. A scan rate of 5° min⁻¹ was applied in the range of 10–60°. UV-vis diffuse reflectance spectra (DRS) were recorded using a UV-vis spectrophotometer (JASCO V-550) equipped with an integrating sphere, and BaSO₄ powder was used as the reference for the baseline correction. The morphologies and particle sizes were examined using field emission scanning electron microscopy (FESEM; S-5500, Hitachi). High-resolution transmission electron microscopy (HRTEM) images were obtained using a Tecnai G2 F30 S-Twin (FEI Company) with an accelerating voltage of 300 kV. For the time-resolved IR spectroscopic study, the photocatalyst was fixed on a CaF₂ plate at a density of 2 mg cm⁻² and placed in a gas cell evacuated at 10⁻⁵ Torr. The Brunauer–Emmett–Teller (BET) surface area was measured at 77 K using a Micromeritics ASAP 2000 adsorption analyzer. Transient IR absorption signals were recorded on a Nicolet 870 FTIR spectrometer with a MCT detector. A pulse laser at 355 nm (1 Hz, 3 mJ per pulse) was used to excite the samples. The width of the laser pulse was 6–8 ns and no deconvolution on the data was carried out.

Photocatalytic reactions

Photocatalytic reactions were carried out in a Pyrex top-irradiation type reaction vessel connected to a closed gas circulation system. Before photoirradiation, the reaction system was evacuated to completely remove air, and then irradiated from the top side using a 300 W xenon lamp with a filtration mirror equipped with an optical filter (Hoya, L-42; $\lambda > 420$ nm) to cut off the ultraviolet light. A flow of cooling water was used to keep the reaction suspension at room temperature. Gas chromatography (Agilent; GC-7890A, MS-5A column, TCD, Ar carrier) was used to



analyze the evolved gases. The pH value before and after the photocatalytic overall water splitting reaction was similarly kept at *ca.* 6.

Measurement of AQE

The AQE measurement was carried out using a Pyrex top-irradiation-type reaction vessel and a 300 W xenon lamp fitted with a 420 nm band-pass filter. The number of photons reaching the reaction solution was measured using a calibrated Si photodiode (LS-100, EKO Instruments Co., LTD.), and the AQE (φ) was calculated according to the following equation:

$$\varphi(\%) = (AR/I) \times 100$$

where A represents a coefficient (4 for H₂ evolution; 8 for O₂ evolution), R represents the evolution rate of H₂ or O₂ in the initial one hour irradiation and I represents the absorption rate of incident photons. It was assumed that all incident photons were absorbed by the suspension. The total number of incident photons at a wavelength of 420 nm was measured to be 4.76×10^{20} photons per h.

Results and discussion

Fig. 1A shows XRD patterns of the Ba(n)-Ta₃N₅ samples ($n = 0-1$), in which all of them exhibit a well-crystallized feature. When the Ba/Ta molar ratio is below 0.03, only diffraction peaks assigned to a single phase of Ta₃N₅ are observed. With a further enhanced molar ratio of Ba/Ta, additional diffraction peaks attributed to BaTaO₂N appear, the intensities of which are continuously increased. Compared with the diffraction peaks of the unmodified Ta₃N₅ sample, a little shift in the diffraction peaks toward a lower angle is observed for the barium-modified Ta₃N₅ samples (Fig. S1†). This demonstrates that the six-coordinated Ba²⁺ may be partly doped into Ta₃N₅ to substitute the Ta⁵⁺ sites, similar to the previous report.³⁰

Fig. 1B shows the UV-vis spectra of the Ba(n)-Ta₃N₅ samples, in which all of the samples similarly exhibit a wide visible light absorption at around 600 nm. The absorption edge is continuously red-shifted with the increasing molar ratio of Ba/Ta, which should be the result of the formed BaTaO₂N species. Compared to the pristine Ta₃N₅ sample, the absorption background originating from the formation of reduced tantalum species (*e.g.*, Ta⁴⁺ and Ta³⁺)^{43,44} on the Ba(n)-Ta₃N₅ samples undergoes an initial decrease and a subsequent increase with the enhancing molar ratio of Ba/Ta. To understand the UV-Vis results, a single phase of BaTaO₂N was prepared *via* the same preparation procedure. As shown in Fig. 1B, the absorption edge of BaTaO₂N is at about 660 nm, and its absorption background is the highest among all of the Ta₃N₅-based samples. It is generally understood that the UV-vis absorption background of the mechanically mixed sample containing two phases should be located between that of the corresponding single phases. That is to say, if the Ba(n)-Ta₃N₅ samples are just a simple mixture of Ta₃N₅ and BaTaO₂N, the absorption background of the Ba(n)-Ta₃N₅ samples will lie between those of the Ta₃N₅ and BaTaO₂N. In this work, however, the samples with

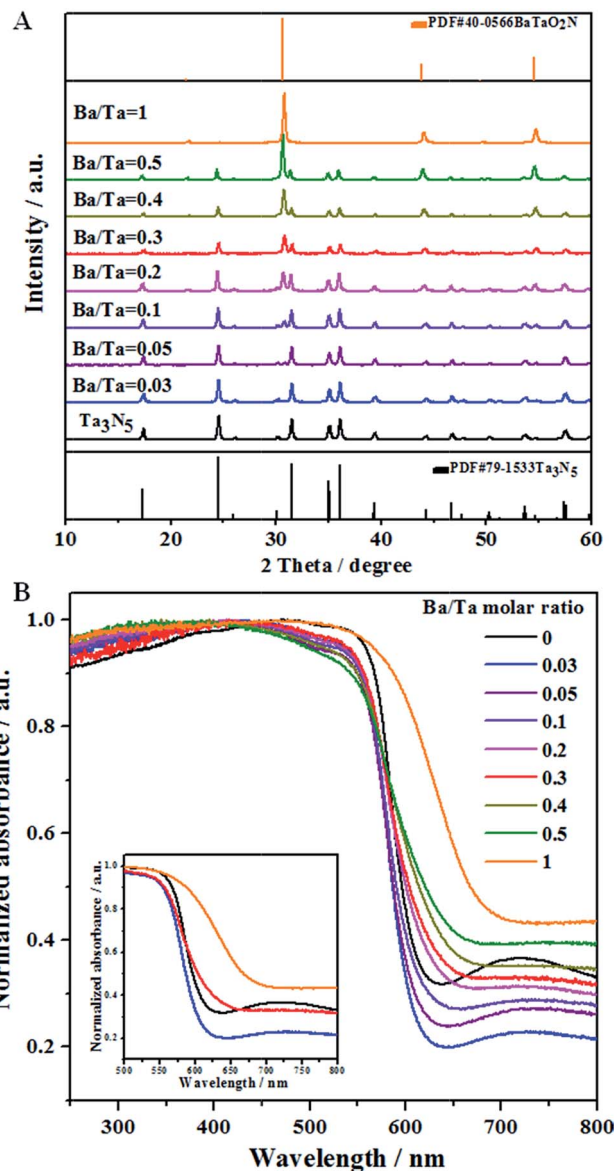


Fig. 1 Structural characterizations of typical Ba(n)-Ta₃N₅ samples ($n = 0-1$): (A) XRD patterns and (B) UV-vis spectra. " n " stands for the molar ratio of Ba/Ta. The inset figure is enlarged for the wavelength range of 500–800 nm.

a molar ratio of Ba/Ta below 0.3 exhibit much lower absorption backgrounds than those of both Ta₃N₅ and BaTaO₂N. This means that the as obtained Ba(n)-Ta₃N₅ samples do not exist as a simple mixture of Ta₃N₅ and BaTaO₂N, but exist as a nanocomposite. In consideration of that, here the Ta₃N₅ and BaTaO₂N phases are one-pot synthesized, thus the BaTaO₂N is expected to be formed *in situ* on the surface of Ta₃N₅ to partly eliminate the surface dangling bonds of Ta₃N₅. On the other hand, the partial barium ions are doped into Ta₃N₅ to inhibit the defect formation. Both of them cause the decrease of defect density. However, it needs to be pointed out that BaTaO₂N itself exhibits the highest defect density among all of the samples. Thus, when the molar ratio of Ba/Ta is excessively enhanced, the defect density of the Ba(n)-Ta₃N₅ sample will become higher



than that of Ta_3N_5 . All of these factors should be integrally responsible for the initial decrease and subsequent increase of the absorption background (*i.e.* defect density) with the increasing molar ratio of Ba/Ta in the UV-Vis results (Fig. 1B).

Fig. 2 shows FESEM images of typical samples. The Ta_3N_5 sample is porous (Fig. 2a), while the BaTaO_2N sample has a shortage of porosity (Fig. 2b). The difference in their morphology can be easily judged from their mixed sample (Fig. 2c). However, the morphology feature of the chosen $\text{Ba}(0.3)\text{-Ta}_3\text{N}_5$ sample (Fig. 2d) prepared in this work is quite different from those of the corresponding single phases (Fig. 2a and b) or their mixed sample (Fig. 2c). As can be seen in Fig. 2d, the $\text{Ba}(0.3)\text{-Ta}_3\text{N}_5$ sample exhibits a homogeneous morphology with the two phases difficult to distinguish, demonstrating their interaction with each other as a nanocomposite. The formation of the $\text{Ta}_3\text{N}_5/\text{BaTaO}_2\text{N}$ nanocomposite can be further supported by the elemental mapping results (Fig. 2e–h). In Fig. 2f, the Ta element originating from both Ta_3N_5 and BaTaO_2N is dispersed everywhere, while the Ba element that can only result from BaTaO_2N is only found in some specific places (Fig. 2g). This can be easily understood to show that the places with Ba element mapping mainly reveal the existence of BaTaO_2N , while the locations with Ta element mapping but a shortage of Ba element mapping stand for the Ta_3N_5 species. Based on the elemental mapping images, we can reasonably give a simulation of the composite state of Ta_3N_5 and BaTaO_2N for the $\text{Ba}(0.3)\text{-Ta}_3\text{N}_5$ sample (Fig. 2h). For comparison, the element mapping results of mechanically mixed Ta_3N_5 and BaTaO_2N ($\text{Ta}_3\text{N}_5/\text{BaTaO}_2\text{N}$ (0.3)-mix) are given in Fig. S2,† from which the Ta_3N_5 and BaTaO_2N phases are mainly separated, different from that of the $\text{Ba}(0.3)\text{-Ta}_3\text{N}_5$ sample. It needs to be pointed out that the composite of Ta_3N_5 and BaTaO_2N does not exist in a core-shell configuration. In addition, the surface areas of the $\text{Ba}(n)\text{-Ta}_3\text{N}_5$ samples are similar to that of BaTaO_2N ($7\text{ m}^2\text{ g}^{-1}$) but a little lower than that of Ta_3N_5 ($9\text{ m}^2\text{ g}^{-1}$), which should result from their shortage of porous structure (Table 1).

To further confirm the formation of the nanocomposite, we carried out a (HR)TEM characterization. Fig. 3 gives the

representative images of the $\text{Ba}(0.3)\text{-Ta}_3\text{N}_5$ sample, in which the interface of the nanocomposite can be clearly observed. As shown in Fig. 3, the obvious lattice fringes indicate that the sample synthesized in this work is well-crystallized, in accordance with the XRD patterns (Fig. 1A). Based on the measurement of lattice distance, we can easily judge the BaTaO_2N and Ta_3N_5 phases. Strikingly, the interfacial contact between BaTaO_2N and Ta_3N_5 is very intimate, revealing the formation of the nanocomposite. The formation of the intimate interface should originate from the one-pot high temperature route and their similar Ta-based octahedron units. In this case, BaTaO_2N is expected to be formed *in situ* on the surface of Ta_3N_5 during the one-pot nitridation process.

The relative band positions of Ta_3N_5 and BaTaO_2N were analysed by combining their Mott-Schottky (M-S) plots and UV-Vis results. In Fig. 4a, the flat band potentials of BaTaO_2N and Ta_3N_5 were evaluated according to M-S measurement results to be *ca.* -0.41 V and -0.32 V vs. NHE, respectively. In consideration of the fact that the bottom of the conduction band (CB) for one n-type semiconductor is normally more negative by *ca.* 0.2 V than the flat band potential,^{24,45,46} the CB positions of the n-type Ta_3N_5 and BaTaO_2N are estimated to be -0.52 eV and -0.61 eV , respectively. By combining their bandgaps achieved from the UV-Vis results (Fig. 1B), the relative band positions of BaTaO_2N and Ta_3N_5 are then deduced and given in Fig. 4b. Accordingly, the nanocomposite exists as a type II heterostructure, where the excited electrons are expected to transfer from the conduction band of BaTaO_2N to that of Ta_3N_5 , while the photogenerated holes will transfer in an opposite way, leading to the spatial charge separation.

The photocatalytic H_2 evolution rates on the pristine and modified Ta_3N_5 samples were examined using the deposited platinum nanoparticle as the reduction cocatalyst in the presence of CH_3OH under visible light irradiation ($\lambda > 420\text{ nm}$). No reaction takes place in the dark, and H_2 is evolved only under light irradiation. As given in the half reaction part of Table 1, the rate of H_2 evolution undergoes an initial increase and subsequent decrease with the increasing molar ratio of Ba/Ta, and the optimal value of the Ba/Ta molar ratio

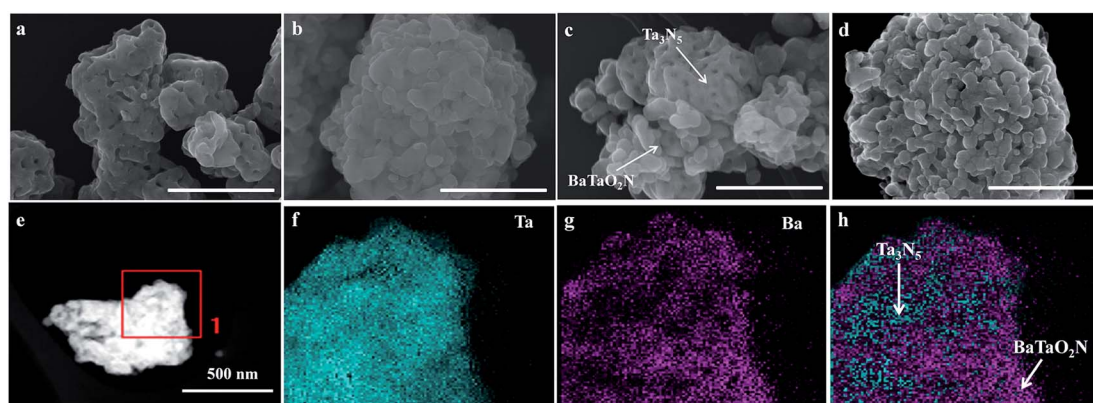


Fig. 2 FESEM images of typical samples: (a) Ta_3N_5 , (b) BaTaO_2N , (c) a mixture of Ta_3N_5 and BaTaO_2N , and (d) $\text{Ba}(0.3)\text{-Ta}_3\text{N}_5$ (the scale bar is 500 nm). Elemental mappings of $\text{Ba}(0.3)\text{-Ta}_3\text{N}_5$: (e) TEM image, (f) Ta element, (g) Ba element, and (h) simulated dispersion of Ta_3N_5 and BaTaO_2N .



Table 1 Photocatalytic performances of typical photocatalysts under visible light irradiation ($\lambda > 420$ nm)

Entry	H_2 -evolving photocatalysts	Surface area ($m^2 g^{-1}$)	Half reaction ^a	Overall water splitting ^b	
				Gas evolution rates ($\mu mol h^{-1}$)	
			H_2 evolution rate ($\mu mol h^{-1}$)	H_2	O_2
1	Ba(0)-Ta ₃ N ₅	9	0.05	0	0
2	Ba(0.03)-Ta ₃ N ₅	7	0.1	Trace	Trace
3	Ba(0.05)-Ta ₃ N ₅	7	4.2	0.8	0.4
4	Ba(0.1)-Ta ₃ N ₅	7	6.6	2.0	1.0
5	Ba(0.2)-Ta ₃ N ₅	7	19.3	2.5	1.3
6	Ba(0.3)-Ta ₃ N ₅	7	30.2	3.2	1.6
7	Ba(0.4)-Ta ₃ N ₅	7	28.2	3.0	1.5
8	Ba(0.5)-Ta ₃ N ₅	7	24.6	2.1	1.1
9	BaTaO ₂ N	7	9.5	0.3	0.15
10	Ta ₃ N ₅ /BaTaO ₂ N (0.3)-mix	8	16.5	0.6	0.3

^a Reaction conditions: 0.15 g of 0.5 wt% Pt/Ba(*n*)-Ta₃N₅ (*n* = 0–1) and 0.5 wt% Pt/Ta₃N₅/BaTaO₂N (0.3)-mix samples; 0.15 g of La₂O₃; aqueous methanol solution (150 mL, 20 vol%); 300 W xenon lamp ($\lambda > 420$ nm); 1 h irradiation. ^b Reaction conditions: 50 mg of 0.5 wt% Pt-modified H₂-evolving photocatalysts; 50 mg of 0.45 wt% PtO_x/WO₃ as O₂-evolving photocatalyst; 100 mL of aqueous NaI solution (1.0 mM); Pyrex top-irradiation type; 300 W xenon lamp ($\lambda > 420$ nm); 1 h irradiation.

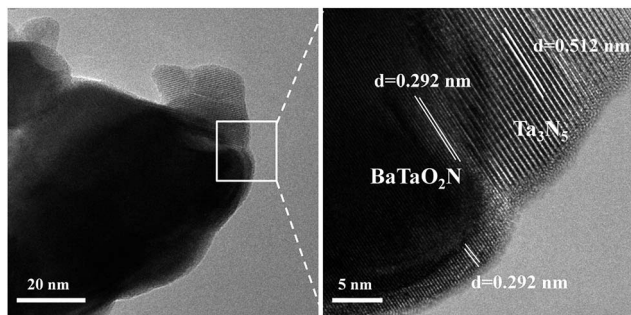


Fig. 3 Representative TEM (left) and locally enlarged HRTEM (right) images of the chosen Ba(0.3)-Ta₃N₅ sample.

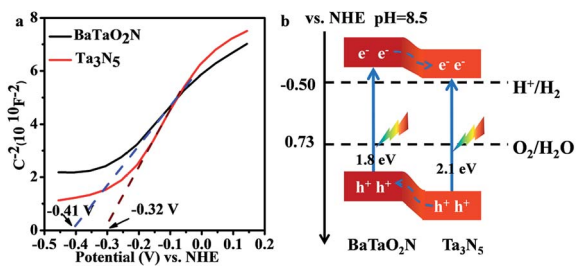


Fig. 4 Band structure characterizations of the Ta₃N₅ and BaTaO₂N samples. (a) Mott-Schottky plot for the Ta₃N₅ and BaTaO₂N electrodes. Electrolyte: 0.5 M Na₂SO₄ solution (pH = 8.5, adjusted using NaOH). Frequency: 1000 Hz. (b) The relative band positions of the Ta₃N₅/BaTaO₂N heterostructure.

is 0.3. Compared to the Ta₃N₅ (entry 1), BaTaO₂N (entry 9) or Ta₃N₅/BaTaO₂N (0.3)-mix (entry 10) sample, the H₂ evolution rate on the Pt/Ba(0.3)-Ta₃N₅ photocatalyst is remarkably promoted. The typical time curve of H₂ evolution on the Pt/

Ba(0.3)-Ta₃N₅ sample is given in Fig. S3,† in which it is almost linearly increased in the experimental region, demonstrating its good photochemical stability. In addition, only a small amount of N₂ (less than 1 μ mol) was detected in the initial stage of irradiation. The dependence of the H₂ evolution rate on the Pt/Ba(0.3)-Ta₃N₅ photocatalyst as a function of irradiation wavelength is well consistent with that of the UV-vis spectra (Fig. S4†), indicating that the H₂ evolution process is driven by the incident light.

Encouraged by the significantly enhanced H₂ evolution rate, we tried to use the pristine or barium-modified Ta₃N₅ samples as H₂-evolving photocatalysts to construct a Z-scheme overall water splitting system together with a PtO_x/WO₃ and IO₃⁻/I⁻ pair as an O₂-evolving photocatalyst and redox mediator, respectively. As shown in the overall water splitting part of Table 1, when using pristine Ta₃N₅ as the H₂-evolving photocatalyst (entry 1), no obvious H₂ evolution is detected, demonstrating the infeasibility of Ta₃N₅ itself to drive the Z-scheme overall water splitting process. However, using the barium-modified Ta₃N₅ samples as H₂-evolving photocatalysts (entries 2–8), overall water splitting with H₂/O₂ molar ratios of close to 2 : 1 is achieved, and the photocatalytic activity is dependent on the Ba/Ta molar ratio with an optimal value of *ca.* 0.3. The Z-scheme activities using the barium-modified samples as H₂-evolving photocatalysts (entries 2–8) are all higher than those using BaTaO₂N (entry 9) or the mixed sample (entry 10). The AQE was measured using the Ba(0.3)-Ta₃N₅ sample as the H₂-evolving photocatalyst to be 0.1% at 420 nm. The activity trend is similar to the result of the photocatalytic proton reduction reaction, indicating that the overall water splitting performance is rate-determined by the H₂-evolving side. In addition, the multiple cycles of time course curves further demonstrate its photochemical stability in the experimental region (Fig. 5). No obvious Ba²⁺ ion residue is observed in the centrifuged solution after reaction.



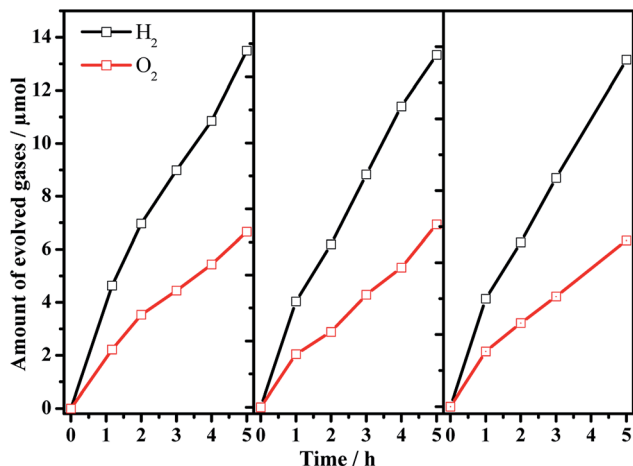


Fig. 5 Multiple cycles of Z-scheme overall water splitting with 0.5 wt% Pt/Ba(0.3)-Ta₃N₅ and 0.45 wt% PtO_x/WO₃ as H₂-evolving and O₂-evolving photocatalysts, respectively. Reaction conditions: 50 mg of Pt/Ba(0.3)-Ta₃N₅ and 50 mg of PtO_x/WO₃; 100 mL of NaI aqueous solution (1.0 mM); 300 W xenon lamp ($\lambda > 420$ nm), top-irradiation.

Photocatalytic overall water splitting commonly confronts huge challenges from both thermodynamic and kinetic aspects.^{2,47} In the past few decades, many candidate materials have satisfied the thermodynamics requirement, but overall water splitting on them is unfeasible because of the constraint of insufficient reaction kinetics in the H₂ and/or O₂-evolving side. Accordingly, developing strategies to address the water splitting reaction kinetics, which is greatly affected by the charge separation and surface catalytic process, is highly valuable. In this work, we adopt a simple one-pot nitridation approach with an ammonia flow (250 mL min⁻¹) at high temperature (1223 K) to address the key issue of the charge separation *via* barium modification of Ta₃N₅. Based on our modification, not only is the defect density of Ta₃N₅ decreased, but also a Ta₃N₅/BaTaO₂N heterostructure with intimate interfacial contact is formed for the promotion of spatial charge separation. Both of these are reasonably responsible for promoting photogenerated charge separation, contributing to the enhanced proton reduction performance as well as the feasible overall water splitting process. The promotion of charge separation is confirmed by comparing the time-resolved infrared spectra (TRIR) of the typical Ta₃N₅, BaTaO₂N and Ba(0.3)-Ta₃N₅ samples (Fig. 6). The effective formation of the Ta₃N₅/BaTaO₂N heterostructure probably originates from their similar structure units containing Ta-based octahedra. The decreased defect density of Ta₃N₅ originates from the part doping of Ba ions and the formation of BaTaO₂N on the surface of Ta₃N₅ leading to surface passivation. It needs to be pointed out that with the increasing Ba/Ta molar ratio, the content of BaTaO₂N with the highest defect density (see UV-Vis results in Fig. 1B) is enhanced, resulting in the increase of recombination centres, which is unfavourable for the photocatalytic H₂ evolution process. As an integral factor of the heterostructure and the defect centres, the photocatalytic activity exhibits an initial increase and a subsequent decrease with the increasing molar ratio of Ba/Ta.

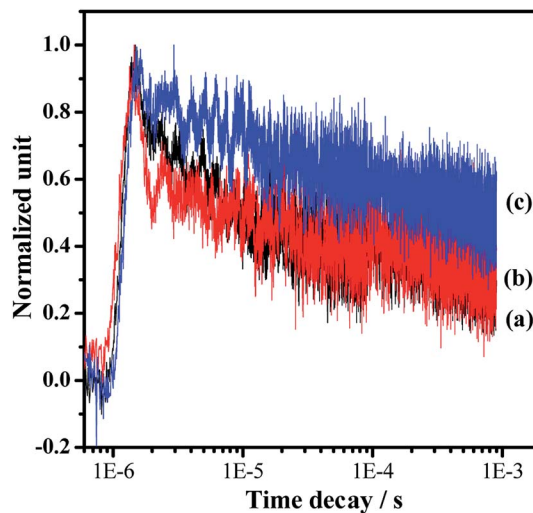


Fig. 6 Normalized transient absorption profiles of the representative samples in a vacuum: (a) Pt/Ta₃N₅, (b) Pt/BaTaO₂N and (c) Pt/Ba(0.3)-Ta₃N₅. The pulse laser at 355 nm was used to excite the samples for the IR tests. The cocatalyst of Pt with a loading amount of 0.5 wt% was deposited by impregnation and a subsequent H₂ reduction method.

Conclusions

In summary, a simple one-pot nitridation strategy is adopted for the barium modification of Ta₃N₅ photocatalyst to address its poor photogenerated carrier separation ability as well as H₂-evolving activity. The one-pot nitridation route overcomes well the challenge of low thermal stability in air for (oxy)nitride-related photocatalysts. Based on this, barium ions are partially doped into Ta₃N₅ to inhibit the formation of defects, and the residue amount of barium ions will cause the *in situ* formation of BaTaO₂N on the surface of Ta₃N₅ to create an intimate interface for the Ta₃N₅/BaTaO₂N heterostructure. Both of the structures favour the enhancement of charge separation efficiency as well as the promotion of the H₂-evolving rate. Finally, we successfully achieve a Z-scheme overall water splitting process under visible light irradiation using the Ba-modified Ta₃N₅ as a H₂-evolving photocatalyst. The fabrication of the heterostructure *via* a one-pot route is expected to be extended into more (oxy)nitride systems for promoted solar energy conversion.

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Notes and references

- 1 K. Maeda and K. Domen, *J. Phys. Chem. C*, 2007, **111**, 7851–7861.
- 2 A. Kudo and Y. Miseki, *Chem. Soc. Rev.*, 2009, **38**, 253–278.



- 3 R. Abe, *J. Photochem. Photobiol., C*, 2010, **11**, 179–209.
- 4 X. Zong and L. Z. Wang, *J. Photochem. Photobiol., C*, 2014, **18**, 32–49.
- 5 D. M. Fabian, S. Hu, N. Singh, F. A. Houle, T. Hisatomi, K. Domen, F. E. Osterloh and S. Ardo, *Energy Environ. Sci.*, 2015, **8**, 2825–2850.
- 6 F. E. Osterloh, *Chem. Mater.*, 2008, **20**, 35–54.
- 7 H. Tong, S. X. Ouyang, Y. P. Bi, N. Umezawa, M. Oshikiri and J. H. Ye, *Adv. Mater.*, 2012, **24**, 229–251.
- 8 X. B. Chen, S. H. Shen, L. J. Guo and S. S. Mao, *Chem. Rev.*, 2010, **110**, 6503–6570.
- 9 Y. Wang, X. C. Wang and M. Antonietti, *Angew. Chem., Int. Ed.*, 2012, **51**, 68–89.
- 10 Y. Ma, X. L. Wang, Y. S. Jia, X. B. Chen, H. X. Han and C. Li, *Chem. Rev.*, 2014, **114**, 9987–10043.
- 11 K. Maeda, K. Teramura, D. L. Lu, T. Takata, N. Saito, Y. Inoue and K. Domen, *Nature*, 2006, **440**, 295.
- 12 Y. Sasaki, H. Kato and A. Kudo, *J. Am. Chem. Soc.*, 2013, **135**, 5441–5449.
- 13 W. Y. Wang, J. Chen, C. Li and W. M. Tian, *Nat. Commun.*, 2014, **5**, 4647.
- 14 C. S. Pan, T. Takata, M. Nakabayashi, T. Matsumoto, N. Shibata, Y. Ikumura and K. Domen, *Angew. Chem., Int. Ed.*, 2015, **54**, 2955–2959.
- 15 Q. X. Jia, A. Iwase and A. Kudo, *Chem. Sci.*, 2014, **5**, 1513–1519.
- 16 S. S. Chen, Y. Qi, T. Hisatomi, Q. Ding, T. Asai, Z. Li, S. S. K. Ma, F. X. Zhang, K. Domen and C. Li, *Angew. Chem., Int. Ed.*, 2015, **54**, 8498–8501.
- 17 D. J. Martin, P. J. T. Reardon, S. J. A. Moniz and J. W. Tang, *J. Am. Chem. Soc.*, 2014, **136**, 12568–12571.
- 18 G. G. Zhang, Z. A. Lan, L. H. Lin, S. Lin and X. C. Wang, *Chem. Sci.*, 2016, **7**, 3062–3066.
- 19 G. Hitoki, A. Ishikawa, T. Takata, J. N. Kondo, M. Hara and K. Domen, *Chem. Lett.*, 2002, 736–737.
- 20 A. Ishikawa, T. Takata, J. N. Kondo, M. Hara and K. Domen, *J. Phys. Chem. B*, 2004, **108**, 11049–11053.
- 21 A. B. Murphy, P. R. F. Barnes, L. K. Randeniya, I. C. Plumb, I. E. Grey, M. D. Horne and J. A. Glasscock, *Int. J. Hydrogen Energy*, 2006, **31**, 1999–2017.
- 22 S. S. K. Ma, T. Hisatomi, K. Maeda, Y. Moriya and K. Domen, *J. Am. Chem. Soc.*, 2012, **134**, 19993–19996.
- 23 Z. Wang, J. G. Hou, S. Q. Jiao, K. Huang and H. Zhu, *J. Mater. Chem.*, 2012, **22**, 21972–21978.
- 24 S. S. Chen, G. J. Liu, Y. Qi, F. X. Zhang and C. Li, *Angew. Chem., Int. Ed.*, 2015, **54**, 3047–3051.
- 25 M. Tabata, K. Maeda, M. Higashi, D. L. Lu, T. Takata, R. Abe and K. Domen, *Langmuir*, 2010, **26**, 9161–9165.
- 26 M. J. Liao, J. Y. Feng, W. J. Luo, Z. Q. Wang, J. Y. Zhang, Z. S. Li, T. Yu and Z. Z. Zou, *Adv. Funct. Mater.*, 2012, **22**, 3066–3074.
- 27 G. J. Liu, J. Y. Shi, F. X. Zhang, Z. Chen, J. F. Han, C. M. Ding, S. S. Chen, Z. L. Wang, H. X. Han and C. Li, *Angew. Chem., Int. Ed.*, 2014, **53**, 7295–7299.
- 28 L. Wang, N. T. Nguyen, X. M. Zhou, I. Hwang, M. S. Killian and P. Schmuki, *ChemSusChem*, 2015, **8**, 2615–2620.
- 29 J. Seo, T. Takata, M. Nakabayashi, T. Hisatomi, N. Shibata, T. Minegishi and K. Domen, *J. Am. Chem. Soc.*, 2015, **137**, 12780–12783.
- 30 Y. B. Li, L. Zhang, A. Torres-Pardo, J. M. Gonzalez-Calbet, Y. H. Ma, P. Oleynikov, O. Terasaki, S. Asahina, M. Shima, D. Cha, L. Zhao, K. Takanabe, J. Kubota and K. Domen, *Nat. Commun.*, 2013, **4**, 2566.
- 31 G. J. Liu, S. Ye, P. L. Yan, F. Q. Xiong, P. Fu, Z. L. Wang, Z. Chen, J. Y. Shi and C. Li, *Energy Environ. Sci.*, 2016, **9**, 1327–1334.
- 32 M. Y. Tsang, N. E. Pridmore, L. J. Gillie, Y. H. Chou, R. Brydson and R. E. Douthwaite, *Adv. Mater.*, 2012, **24**, 3406–3409.
- 33 L. Yuliati, J. H. Yang, X. C. Wang, K. Maeda, T. Takata, M. Antonietti and K. Domen, *J. Mater. Chem.*, 2010, **20**, 4295–4299.
- 34 K. Maeda, N. Nishimura and K. Domen, *Appl. Catal., A*, 2009, **370**, 88–92.
- 35 D. A. Wang, T. Hisatomi, T. Takata, C. S. Pan, M. Katayama, J. Kubota and K. Domen, *Angew. Chem., Int. Ed.*, 2013, **52**, 11252–11256.
- 36 X. M. Liu, L. Zhao, K. Domen and K. Takanabe, *Mater. Res. Bull.*, 2014, **49**, 58–65.
- 37 S. S. Chen, Y. Qi, Q. Ding, Z. Li, J. Y. Cui, F. X. Zhang and C. Li, *J. Catal.*, 2016, **339**, 77–83.
- 38 V. Strahle, *Z. Anorg. Allg. Chem.*, 1973, **402**, 47–57.
- 39 J. Zhang, Q. Xu, Z. C. Feng, M. J. Li and C. Li, *Angew. Chem., Int. Ed.*, 2008, **47**, 1766–1769.
- 40 X. Wang, Q. Xu, M. R. Li, S. Shen, X. L. Wang, Y. C. Wang, Z. C. Feng, Y. S. Shi, H. X. Han and C. Li, *Angew. Chem., Int. Ed.*, 2012, **51**, 13089–13092.
- 41 J. Z. Su, L. J. Cuo, N. Z. Bao and C. A. Grimes, *Nano Lett.*, 2011, **11**, 1928–1933.
- 42 E. S. Kim, N. Nishimura, G. Magesh, J. Y. Kim, J. W. Jang, H. Jun, J. Kubota, K. Domen and J. S. Lee, *J. Am. Chem. Soc.*, 2013, **135**, 5375–5383.
- 43 J. Q. Wang, S. Y. Su, B. Liu, M. H. Cao and C. W. Hu, *Chem. Commun.*, 2013, **49**, 7830–7832.
- 44 K. Maeda, M. Higashi, D. L. Lu, R. Abe and K. Domen, *J. Am. Chem. Soc.*, 2010, **132**, 5858–5868.
- 45 Y. Matsumoto, *J. Solid State Chem.*, 1996, **126**, 227–234.
- 46 S. S. Chen, Q. Yi, G. J. Liu, J. X. Yang, F. X. Zhang and C. Li, *Chem. Commun.*, 2014, **50**, 14415–14417.
- 47 T. Hisatomi, J. Kubota and K. Domen, *Chem. Soc. Rev.*, 2014, **43**, 7520–7535.

