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Revisiting poly(vinyl chloride) reactivity in the context of chemical recycling

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Poly(vinyl chloride) (PVC) is one of the highest production volume polymers due to its many applications, and it is one of the least recycled due to its chemical structure and frequent formulation with additives. Developing efficient PVC recycling techniques would enable PVC waste to be reused or repurposed in other processes. Within this context, the literature on PVC modification offers considerable insight into versatile reaction pathways, potentially inspiring new approaches for repurposing PVC waste into value-added products. This perspective provides an overview of PVC functionalization through a lens of chemical recycling, discussing various PVC reactivity trends and their applications with a critical assessment and future outlook of their recycling implications.

Introduction

Plastics are integrated into all facets of modern society, and their ubiquity has led to waste accumulation that threatens both environmental and ecological systems.^{1,2} The United States (U.S.) generates 44 million tons of plastic waste each year—the majority of which is comprised of polyethylene (PE, 52%), polypropylene (PP, 19%), polyethylene terephthalate (PET, 14%), polystyrene (PS, 7%), and poly(vinyl chloride) (PVC, 2%).³

Each year, 86% (38 million tons) of this plastic waste is disposed in landfills, prompting efforts to expand plastic reuse and recycling through technology, policy, and humanitarian efforts.

For some plastics (e.g., PET), closed-loop recycling can be achieved by chemically deconstructing the polymer into its constituent monomers and repolymerizing them to generate pristine material.⁴ For others, such as those with all-carbon backbones (e.g., PE, PP, PS, PVC), the polymers are more thermodynamically and kinetically stable, making depolymerization more difficult.⁵ In most cases, mechanical recycling—where plastics are separated by their chemical identity and melt reprocessed—is used.⁶ However, the efficacy is limited and the resultant recycled plastics typically exhibit poorer material properties. An alternative is open-loop chemical recycling—

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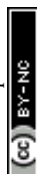
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where polymers are converted into alternative products through chemical reactions.⁷ While this approach is feasible for most polyolefins, PVC presents additional challenges to existing recycling methods due to facile cleavage of the C–Cl bond and extensive use of additives.^{8,9} Consequently, only *ca.* 18 000 tons (3% of total municipal PVC waste) is recycled in the U.S. per year, presumably *via* mechanical recycling methods.³

PVC is a uniquely versatile polymer, owing to its variability in material strength, durability, chemical stability, flexibility, and processability when compatibilized with various plasticizers, stabilizers, and lubricants. These qualities enable its use in wide-ranging applications, including construction materials, fluid piping, cable insulation, floor and wall coverings, and healthcare devices, some of which have product lifetimes of 30–50 years.^{8,10} However, the diverse applications and long lifetimes make it difficult to recycle, as the additives frustrate a one-size-fits-all approach while corrosive chlorine byproducts present hazardous conditions for thermochemical recycling processes.¹¹

Despite the challenges associated with PVC recycling, several efforts by industrial and governmental groups have

demonstrated promising results.¹² For instance, the European PVC industry has worked for over 20 years to minimize the environmental impact of PVC production and manufacturing through voluntary commitments (Vinyl 2010, VinylPlus, and VinylPlus 2023).¹³ Using primarily mechanical recycling methods, they have recycled 8.9 million tons of PVC since 2000, including 896 000 tons in 2022 (*ca.* 27% of total PVC waste generated in Europe).¹⁴ Meanwhile, VinylLoop—a recycling project developed by Solvay that used solvent-based separation—ceased operation at its Ferrara plant in 2018 following European Union regulations on phthalates, which stymied market demand for their recycled PVC containing phthalate plasticizers.¹⁵ In the U.S., the Vinyl Institute recently launched the Vinyl Sustainability Council, an industry-wide initiative that aims to drive sustainable economic development in the PVC industry, citing 550 000 tons of PVC recycled in 2020, 13% of which is post-consumer materials.¹⁶ Indeed, there are certainly promising directions within post-industrial PVC recycling, yet the challenges of post-consumer PVC waste remain. Mechanical methods are particularly effective and economical for recycling within production facilities and clean waste streams. However, the diversity of post-consumer PVC waste—featuring varying amounts and classes of additives—requires chemical processing which can tolerate or separate these complex mixtures. To this end, significant opportunities remain to develop sustainable and economical chemical recycling routes for PVC.

An ideal chemical recycling process would enable complete utilization of the carbon, chlorine, and hydrogen content of PVC without negative environmental impacts (*e.g.*, emissions, process waste) and prohibitive costs. While this ideal has yet to be fully achieved, numerous research groups have proposed potential routes for recycling waste PVC. Conventional approaches apply thermochemical methods (*e.g.*, combustion, pyrolysis, gasification) to deconstruct PVC, and these developments have been previously reviewed.^{8,10,17} Broadly, these processes dechlorinate PVC and subsequently deconstruct the resulting hydrocarbon to generate fuels and small molecules or



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recover energy. However, the evolution of gaseous hydrogen chloride—among other reactive species—requires specialized process equipment (*e.g.*, scrubbers) to capture or neutralize these products. Synthetic methods offer a compelling alternative, wherein PVC may be dechlorinated or converted directly to alternative products within less harsh reaction environments. As such, these targeted transformations may require fewer unit operations and cheaper reactor equipment than existing thermochemical routes and may provide facile pathways for chlorine utilization rather than disposal.

Here we aim to focus on more direct chemical transformations that circumvent deconstruction to monomeric and oligomeric constituents. There is an enormous body of existing research on PVC reactivity, which has largely been conducted outside the scope of recycling. We believe that many of these studies could be leveraged to develop new chemical recycling methods for PVC waste. This perspective highlights select papers from an extensive literature search of over 3000 PVC papers published since 2010.[†] We then consider possible extensions of this prior research in the context of recycling and emphasize studies which specifically target post-consumer PVC materials. Ultimately, we aim to invigorate the chemical science community to identify and evaluate potentially sustainable recycling routes which can facilitate complete PVC utilization.

Structure and reactivity

The repeat unit structure informs how PVC plastics can be recycled or repurposed. Three reaction types are most often utilized based on the C–Cl bond reactivity (Scheme 1): (1) heterolytic dechlorination (*i.e.*, nucleophilic substitution), (2) homolytic dechlorination, and (3) dehydrochlorination (*i.e.*, elimination). The resulting products can be used directly or leveraged in further transformations as discussed throughout this perspective. Before delving into the specific reactions, we start by providing some general structure and reactivity principles gleaned from the literature.

Chlorine is a modest leaving group and can take part in substitution and elimination reactions under many conditions.

[†] Papers indexed in Clarivate's Web of Science were searched using the following criteria: using the "all fields" search with multiple rows, the keywords "PVC", "polyvinyl chloride", "poly(vinyl chloride)", and "polyvinylchloride" were entered with the "or" option, which returned 47 845 documents. This list was narrowed down by selecting specific "citation topics", including "polymer sciences", "polymers and macromolecules", "electrochemistry", "sustainability science", and "manufacturing", yielding 11 350 papers. Next, "years published" was narrowed to only include papers published in 2010–2023, which gave 4224 papers. These papers were further filtered by document type, including "articles", "reviews", "book chapters", "early access", "letters", and "books", to give 3866 papers. Next, papers on the following topics were removed to generate 2600 papers: "lithium-sulfur batteries", "flame retardancy", "protein adsorption", "lithium-ion batteries", "climbing robot", "proton conductivity", "thermal comfort", and "gene delivery". Next, the title, abstract, and conclusion of these papers were read and non-relevant ones (*e.g.*, construction applications) were removed from the list, bringing the total to 214 papers. Focusing further on the chemical modification of PVC, and omitting the papers related to pyrolysis, gasification, hydrothermal, plasticizer stability, *etc.*, provided 112 papers (85 articles and 27 review papers), which are cited herein. Additional papers were discovered and added during the writing of this work.



Scheme 1 Common reaction modalities for PVC and relative reactivities of different C–Cl bonds within PVC.

Competition between these processes can diminish product selectivity, presenting a challenge for chemical recycling strategies, which require high purity and/or control for practical applications. Various factors influence the propensity for substitution or elimination, including polymer structure and reaction environment. For instance, isotactic sequences preferentially undergo substitution over elimination as compared to atactic and syndiotactic polymers.^{18,19} Within the context of reaction conditions, aprotic polar solvents such as dimethylformamide favor substitution over elimination as compared to protic polar solvent like ethylene glycol;^{20,21} higher temperatures favor elimination over substitution;^{22,23} higher nucleophile concentrations improve dechlorination efficiency with increased substitution and elimination;²⁴ and less basic nucleophiles favor substitution over elimination as compared to more basic ones.²⁵

In addition, structural defects, which stem from uncontrolled free radical polymerization, can impact reactivity.²⁶ Specifically, tertiary and allylic chlorine-substituted carbons that result from backbiting and subsequent rearrangements, hydride shifts, and β -scission reactions are often the most reactive sites in PVC. The concentration of defects varies with polymerization conditions,²⁷ but on average, free radical polymerization results in 3–5 tertiary defects per 1000 repeat units²⁸ and 1 allylic defect (either internal or terminal) per polymer chain.²⁹ Although sparse, these defects are responsible for initiating autocatalytic thermal dehydrochlorination because of their weaker C–Cl bond strengths.³⁰ Though efforts to reduce the defect sites^{31,32} and autocatalysis³³ are underway, some researchers have taken advantage of these more labile chlorine atoms in reactions (*vide infra*).

Heterolytic dechlorination

Chlorine substitution by nucleophiles can convert PVC into alternative polymeric materials that may be useful in specialty



applications. Nevertheless, these studies rarely leverage waste PVC as their starting materials. We discuss herein potential extensions of these reaction pathways, highlighting particularly interesting or promising research across synthetic routes and end-use applications. Throughout, we generally organize the discussion by reaction pathway.

As noted above, PVC readily undergoes substitution reactions with nucleophiles. Previous work with PVC has primarily utilized S-, N- and O-based nucleophiles to generate materials for targeted applications such as catalysis,^{34–41} separations,^{42–47} metal-complexation,^{48–55} biodegradable polymers,^{56,57} thermally stable and photostable polymers,^{58,59} and antimicrobial polymers.^{59,60} In their 2010 and 2015 reviews, Moulay outlines many transformations that occur *via* a heterolytic, S_N2 mechanism.^{61,62}

As an example, Reinecke and coworkers developed a method to attach phthalate- and isophthalate-based plasticizers to PVC *via* thiol substitution, which would reduce plasticizer contamination in other waste streams during recycling (Scheme 2A).⁶³ When compared to commercial plasticized PVC, the covalently linked plasticizer completely reduced plasticizer migration, albeit at the expense of flexibility. Further optimizing these strategies by evaluating the reaction efficiency with other molar masses could lead to competitive, flexible PVC products (*e.g.*, packaging) without plasticizer migration. In a different report, Yagci and coworkers utilized elemental sulfur and sodium sulfide nonahydrate to form polysulfide, which was used to vulcanize commercial PVC at room temperature.⁶⁴ The resulting PVC-polysulfide network exhibited self-healing behavior and could be reprocessed after use due to the dynamic S–S linkages.

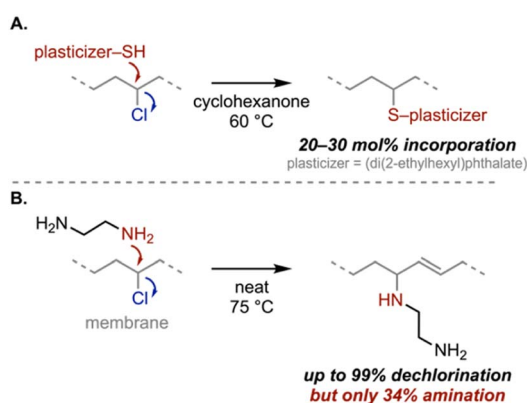
While these are promising examples, nucleophilic substitution pathways are limited in practice due to incomplete dechlorination and competition with elimination, which lead to residual chlorine atoms and alkenes, respectively. These structural “impurities” will likely alter the polymer properties and potentially limit their utility. Nevertheless, these synthetic challenges represent opportunities for innovation. For instance, Hwang and coworkers achieved up to 99% dechlorination of

PVC membranes using ethylenediamine; however, only 34% amination (substitution) was observed (Scheme 2B).⁶⁵ Another example, introduced by Fieser and coworkers, uses (xantphos) Rh catalysts and hydrogen donors (*i.e.*, sodium hydride, sodium formate, triethylsilane) to enable selective hydrodechlorination to polyethylene-like materials with 63% substitution efficiency.⁶⁶ Optimizing dechlorination efficiency and selectivity simultaneously, while challenging, could be instrumental in yielding versatile recycling approaches. Previous studies have shown that choosing appropriate reaction conditions, including nucleophile identity, temperature, concentration, and solvent polarity (summarized in the Structure and reactivity section), could lead to reactions with higher substitution efficiency.

Other, less common routes to substitute PVC include Friedel–Crafts alkylation,^{67,68} cross-coupling,⁶⁹ and the Barbier reaction.⁷⁰ For example, Arslan and coworkers utilized magnesium-mediated activation of the C–Cl bonds to generate a nucleophile that subsequently reacted with electrophilic aldehydes and epoxide derivatives.⁷⁰ An impressively high functionalization efficiency (up to 90%) was achieved by optimizing the ratio of magnesium to chlorine. Such methods could enable polymers with higher degrees of functionalization; however, there are few examples, likely due to the low reactivity of alkyl chlorides in organometallic reactions. One exciting approach is to directly convert PVC into an existing high-value polymer by substituting chlorine for other functionalities (*e.g.*, a phenyl moiety to form PS). For example, Fieser and coworkers recently reported a method for converting PVC to poly(ethylene-co-styrene) *via* tandem hydrodechlorination/Friedel–Crafts alkylation.⁷¹

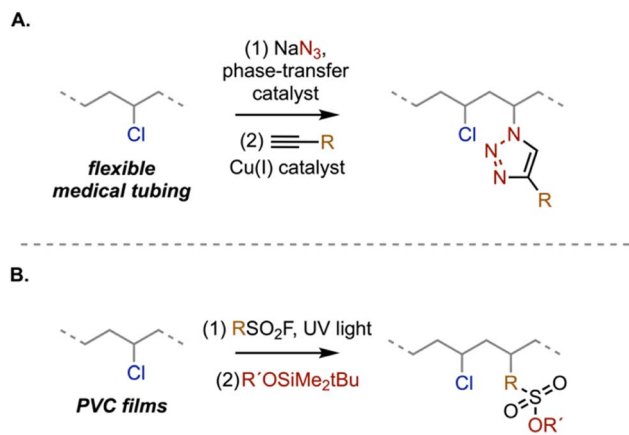
An alternative approach to direct chlorine substitution is to install a more versatile reactive moiety that enables subsequent functionalization. While the need for and cost associated with additional processing steps could stifle the practical efficacy of this approach, these secondary reactions provide access to a wider array of functionality. As an example, an azide can be installed onto the polymer backbone using conventional S_N2 chemistry and then subsequently reacted with an alkyne to yield a triazole linkage in a “click” reaction. This strategy has been used to affix different moieties onto PVC for wide-ranging applications, including covalently linked plasticizers,^{72–77} electrochemical systems,^{78–83} metal detection,⁸⁴ energy storage,⁸⁵ photochemistry,^{86,87} and anti-bactericides.^{88,89} One potential limitation is the lower degree of azide incorporation (<15%) due to both incomplete substitution and competing elimination pathways.

Click reactions have also been employed to modify PVC surfaces. For example, Finn and coworkers utilized the Cu-catalyzed azide–alkyne reaction to add ammonium, pyridyl, ether, and fluorinated amides to the surface of commercial, plasticized PVC tubing (Scheme 3A).⁹⁰ In complementary work, Chen and coworkers reported a one-pot photoreaction and sulfur-fluoride exchange (SuFEx) reaction to functionalize PVC films (Scheme 3B).⁹¹ SuFEx reactions are appealing due to their mild conditions and lack of metal catalyst. It would be interesting to extend this chemistry to conditions that favor PVC

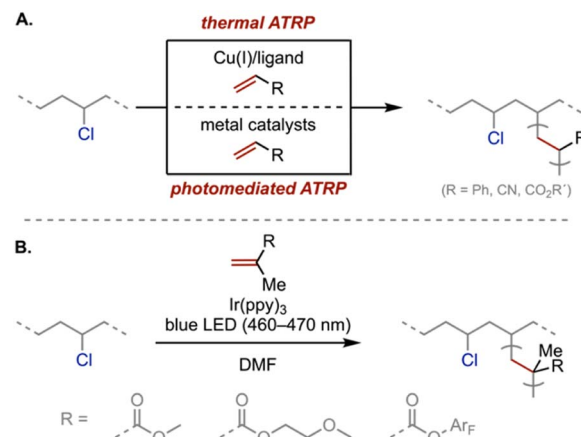


Scheme 2 Nucleophilic substitution reactions on PVC. (A) Covalent incorporation of plasticizer *via* substitution.⁶³ (B) Competing substitution and elimination observed when PVC is reacted with ethylenediamine.⁶⁵





Scheme 3 Click reaction pathways for a two-step PVC functionalization, including an (A) azide-alkyne cycloaddition⁹⁰ and (B) SuFEx reactions.⁹¹



Scheme 4 (A) Thermal and photomediated ATRP for PVC functionalization. (B) An Ir-catalyzed example for generating graft copolymers.¹⁰⁶

dissolution and evaluate the functionalization efficiencies under homogeneous reaction conditions.

Beyond click chemistry, other researchers have performed multi-step reactions on PVC *via* secondary steps such as acylation⁵⁷ or epoxidation.⁹² For example, Jaymand and coworkers used allylic alcohol-based substitution followed by epoxidation to generate epoxy-resin nanocomposites consisting of epoxidized novolac (a phenol-formaldehyde resin), epoxidized poly(vinyl chloride), and modified silica nanoparticles.⁹² The resin nanocomposite showed improved thermal stability compared to the neat epoxidized novolac; however, it was not compared to commercially available epoxy resins. Similar to other studies, a competing elimination reaction generates alkenes that can also undergo epoxidation, complicating structural control.

Homolytic dechlorination

Homolytic C–Cl bond cleavage generates reactive carbon-centered radicals, which have been used to generate graft copolymers from PVC. For example, atom transfer radical polymerization (ATRP) is the most common approach, leveraging tertiary alkyl and allylic chlorides in PVC as macro-initiator sites (Scheme 4A). Previous work has utilized numerous monomers—methyl methacrylate, styrene, acrylonitrile, glycidyl methacrylate—for various applications such as hydrogen-bonded macromolecules,^{93,94} dye-sensitized solar cells,^{95–97} membranes,^{98–102} covalent plasticizers,¹⁰³ and polymeric sorbents.¹⁰⁴ In particular, photoinitiated ATRP has gained attention due to its spatial and temporal control. For example, Yagci and coworkers reported a visible-light-mediated ATRP using dimanganese decacarbonyl as a light-absorbing component and parts per million levels of Cu-catalyst to synthesize graft poly(methyl methacrylate) copolymers from PVC.¹⁰⁵ In a separate report, Huang and coworkers developed an Ir-catalyzed visible-light-mediated ATRP to graft poly(methyl methacrylate), poly(pentafluorophenyl methacrylate), and poly(oligo(ethylene glycol) methyl ether methacrylate) onto PVC

(Scheme 4B).¹⁰⁶ Although ATRP is quite versatile in functionalizing PVC, the reaction proceeds predominantly at the sparse structural defect sites (*i.e.*, allylic and tertiary chlorides), limiting grafting density and dechlorination efficiency.

Other graft copolymerization techniques have also been applied for PVC modification; however, their use is limited to only a few studies (*e.g.*, reversible addition fragmentation chain transfer,¹⁰⁷ emulsion polymerization,¹⁰⁸ as well as other radical-generating strategies including plasma,^{109–111} gamma rays,¹¹² and UV light¹¹³). Although these graft copolymers, along with those synthesized by ATRP, could potentially make cheaper end products by leveraging waste PVC feedstocks, the present market demand for such specialty copolymers is limited.

Beyond polymer grafting, others have used radical chemistry to install maleic anhydride substituents using benzoyl peroxide^{114,115} and fullerene substituents using azobisisobutyronitrile as initiators.¹¹⁶ Additionally, sodium trisulfide has been applied to crosslink PVC *via* monosulfide bridges.¹¹⁷ Our group recently introduced an electrochemical approach that combines homolytic C–Cl reduction with oxidative arene electrochlorination, allowing waste PVC to serve as a chlorine feedstock for organic syntheses at room temperature.¹¹⁸

Nevertheless, radical pathways are likely more limited on what functional groups can be installed when compared with nucleophilic substitution. While the generation of carbon-centered radicals presents unique opportunities to install new functionalities, their potent reactivity can result in undesired coupling reactions (*e.g.*, crosslinking). In controlled radical polymerizations like ATRP, crosslinking may not be prominent because of the low concentration of defect sites. However, generating higher radical concentrations could facilitate reactions with solvent and adventitious contaminants, including O₂ and H₂O.¹¹⁹ Additionally, the PVC macroradical may undergo chain-scission to form low molar-mass fragments that can limit applications.¹²⁰ Consideration and control of these multifaceted reaction pathways will be essential to enable such methods for broader applications in chemical recycling.



Dehydrochlorination

Dehydrochlorination involves the base-mediated elimination of hydrogen and chloride to generate an alkene. For example, aqueous alkaline solutions (e.g., NaOH, K₂CO₃, NH₃) can be used to dechlorinate PVC (Scheme 5A), forming polyolefins and inorganic chlorides (e.g., NaCl, KCl, NH₄Cl). To this end, Osada and Yana performed deplasticization and dechlorination in a single pot by varying the temperature—dioctyl phthalate hydrolysis at 150 °C and up to 100% dechlorination at 235 °C—*via* microwave heating in an 8 M NaOH solution.¹²¹ Using a more conventional heating source required higher temperatures and base concentrations. One advantage of this approach is that chlorine-containing products (e.g., NaCl) have existing markets. However, the resulting polyene has limited applications, owing to insolubility. In addition, further reaction optimization is required to reduce NaOH concentrations for chemical recycling given its high energy requirements for production.¹²² In other reports, NaOH was combined with diols (ethylene glycol, diethylene glycol, tetraethylene glycol)¹²³ and then ball milled^{124,125} to achieve up to 99% dechlorination for waste PVC at lower alkaline concentrations. However, competing substitution reactions led to mixed products. In related work, Pang and coworkers evaluated KOH-assisted dehalogenation of waste PVC *via* ballmilling to generate KCl and chlorine-free carbonaceous materials, which were thermally annealed to form porous carbon electrodes for supercapacitor applications.¹²⁶

Ionic liquids have been explored as both a solvent and base for dehydrochlorination reactions, owing to their high basicity, thermal stability, and low vapor pressure.¹²⁷ For example, De Vos and coworkers reported 98% dehydrochlorination of PVC in a PBu₄Cl ionic liquid after heating at 180 °C for 1 h.¹²⁸ Building on this work, Oster and coworkers evaluated similar phosphonium-based ionic liquids to separate the components of a multilayer blisterpack (containing PVC, aluminium, and oriented polyamide), with up to 99% dehydrochlorination.¹²⁹

Though often touted as “green” solvents, ionic liquids are made through energy- and material-intensive processes.^{130,131} As

such, developing approaches to reuse ionic liquids after the reaction could facilitate its large-scale usability. However, this approach has yet to be explored in PVC chemistry beyond a few reuses. In addition, the costs of recycling ionic liquids must be considered. A separate concern relates to how to dispose of ionic liquids after use, as they can be classified as persistent pollutants due to their chemical and thermal stability. Concerns about their ecological toxicity have already been well documented.¹³¹

One remaining challenge is what to do with the polyene products formed after dehydrochlorination. De Vos and coworkers developed an elegant tandem dehydrochlorination–hydrogenation process to convert PVC into saturated but nonlinear, soluble polymers with approximately 80% dechlorination (Scheme 5B).¹³² While the use of rhodium catalysts could prohibit large-scale demonstrations, cheaper hydrogenation catalysts could be explored.¹³³ For example, Goulas and coworkers reported a tandem dehydrochlorination–hydrogenation–hydrogenolysis to generate polyethylene waxes suitable for hot melt adhesive applications.¹³⁴ Using an organic amine base, a Pt/C heterogeneous catalyst, high pressures of H₂, and high temperatures, they obtained >70% yields of wax products with high conversions for both Cl elimination and alkene hydrogenation. To further improve this approach, lower catalyst loadings or recyclable catalysts will be needed, as well as milder reaction conditions.

The future of chemical PVC recycling

While there is a wealth of literature aimed at chemically modifying PVC, few of these studies utilize waste plastic as a feedstock. As such, further research will be required to understand how these methods would work with commercial PVC products, which often contain plasticizers, stabilizers, lubricants, and other plastic and non-plastic components. However, there are several emergent approaches demonstrating promising chemical recycling methods applied directly to post-consumer PVC materials.

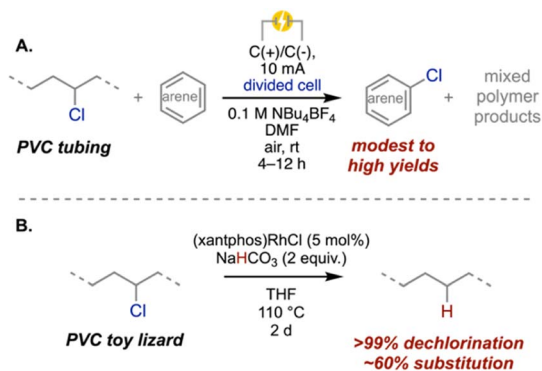
One approach is to leverage waste PVC as a chlorine feedstock by using dechlorination to extract chlorine on-demand for direct use as a synthetic reagent. Work by our group used plasticized PVC tubing as a chlorine source for an oxidative arene electrochlorination (Scheme 6A).¹¹⁸ A preliminary life cycle assessment showed a 68% reduction in greenhouse gas emissions when using waste PVC as the chlorine source compared to hydrochloric acid. More work is needed, however, to optimize the electrodechlorination because the mixed, partially dechlorinated polymer has low market value.

One route aimed at recycling the carbon content focuses on dechlorinating PVC to form polyethylene-like materials that could be used directly or further recycled alongside waste PE using conventional thermochemical recycling methods. The previously described study by Fieser and coworkers (Scheme 6B), which leverages hydride donors, offers a pathway for simultaneously recycling the carbon content of the polymer and generating potentially valuable Cl-containing byproducts (e.g., chlorotriethylsilane when Et₃SiH is used).⁶⁶ However, the



Scheme 5 (A) Dehydrochlorination of PVC *via* elimination leading to insoluble char. (B) Tandem ionic liquid-mediated PVC dehydrochlorination with Rh-catalyzed hydrogenation to form highly saturated but non-linear soluble polymers.¹³²





Scheme 6 (A) PVC as a chlorine source for organic synthesis using electrochemistry.¹¹⁸ (B) Reductive hydrodechlorination using a metal catalyst and a hydride source.⁶⁶

required catalysts and hydrogen sources are expensive, potentially challenging economic viability. Additionally, a more selective substitution would help limit dehydrochlorination side reactions, adding value to the resulting polyethylene-like materials.

While numerous potential approaches have achieved partial utilization of carbon, chlorine, and hydrogen content, we know of few such non-thermochemical processes capable of fully dechlorinating PVC and reconstituting its substituents into new, useful products. Nevertheless, some of these chemical recycling pathways discussed above and throughout this perspective appear promising, and we posit that these methodologies could yield commercially viable PVC recycling technologies in the future with further optimization.

Conclusions

By evaluating recent literature on chemical modification to PVC, we have outlined promising future routes for waste recycling and highlighted potential challenges and limitations for such processes. These trends suggest that a given class of reaction may be more desirable depending on the end-goal of a particular recycling operation. For instance, nucleophilic substitution offers direct access to diverse polymer products, but these reactions display relatively low substitution efficiency and high competition with elimination. Future chemical modification schemes should look to improve reaction efficacy and product purity. Alternatively, reactions to form polyenes could be a more facile approach to achieve high dechlorination efficiencies, as elimination reactions can be conducted with higher selectivity and conversion. However, subsequent processing is typically needed to generate useful polymeric or chemical products from the dechlorinated PVC. New methods in this area should look to conduct reactions under more mild conditions (*e.g.*, lower NaOH concentrations) while considering the practical scalability of such pathways (*e.g.*, ionic liquid cost). In addition, finding applications for the char produced or identifying more mild hydrogenation conditions to instead generate polyethylene are needed. Finally, homolytic pathways forming graft copolymers may also feature a diverse range of unique product

structures, but the limited large-scale demand for such products currently frustrates the viability of these pathways for chemical recycling. Other homolytic pathways, including electrodechlorination, are promising but need more optimization to increase dechlorination efficiency and product selectivity.

Beyond the chemical processes themselves, the field would benefit greatly from complementary analyses of waste PVC products and their broader impacts. For instance, the chemical composition of PVC waste streams is not thoroughly characterized, making investigations in representative conditions more challenging at the lab-scale and preventing more detailed understanding of reactions that utilize post-consumer products. Life-cycle assessments, especially done in tandem with chemical recycling studies, could provide necessary context for the overall sustainability of PVC recycling methods. Those approaches that address challenges in plastic waste yet generate other environmental hazards (*e.g.*, CO₂ emissions) limit the overall benefits of such processes.

Taken together, there are numerous promising approaches for chemically recycling PVC waste beyond traditional thermochemical processes. Considering the breadth of opportunities within chemical discovery, reaction engineering, and auxiliary analyses, there exists considerable space for new entrants to impact the chemical recycling field.

Author contributions

The manuscript was written through contributions of all authors. RKJ was involved with conceptualization, data curation, formal analysis, methodology, validation, visualization, and writing – original draft, reviewing, and editing. BJN was involved with formal analysis, methodology, validation, visualization, and writing – original draft, reviewing, and editing. MSY and DEF were involved with early conceptualization and writing – reviewing and editing. AJM was involved with conceptualization, funding acquisition, project administration, resources, supervision, visualization, and writing – original draft, review, and editing.

Conflicts of interest

There are no conflicts to declare.

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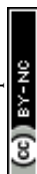
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