

RESEARCH ARTICLE

View Article Online

View Journal | View Issue

Cite this: *Org. Chem. Front.*, 2023, **10**, 1915Synthesis of α -amino acid derived (1,2,3-triazol-4-yl)-picolinamide (tzpa) ligands and their corresponding luminescent Tb(III) complexes†Isabel N. Hegarty,^a Chris S. Hawes^b and Thorfinnur Gunnlaugsson^b *Received 14th February 2023,
Accepted 8th March 2023

DOI: 10.1039/d3qo00232b

rsc.li/frontiers-organic

The formation of chiral ligands for use in metal directed synthesis of coordination complexes¹ and supramolecular self-assemblies^{2,3} has become a highly active area of research in recent times.⁴ This development has led to the synthesis of a wide range of organic ligands,^{5,6} which in conjunction with either d- or f-metal ions^{7–9} can result in the formation of both structurally versatile and beautiful structures,¹⁰ the synthesis of which cannot be achieved without the use of such a template synthesis.^{11,12} Furthermore, the physical properties of the ions themselves can be transferred to such structures, often resulting in the generation of systems possessing novel structural, physical and material properties.^{13,14} We have developed many examples of such chiral organic ligands and chiral metal templated self-assemblies in the past.^{15–18} In particular, we have probed their structural properties, as well as their physical organic and material properties in both solution and the solid state.¹⁹ We recently developed a new family of ligands based on the (1,2,3-triazol-4-yl)-picolinamide (**tzpa**) core, as depicted schematically in Fig. 1.^{20,21} The **tzpa** family are designed by combining two effective ligand motifs, namely 2,3-bis(1,2,3-triazol-4-yl)picolinamide (**btp**) and dipicolinic acid/amide (**dpa**), both of which we have synthesized and employed in our studies over the years.^{22,23} Here, we present the synthesis of two chiral **tzpa** ligands (shown as **4** and **5** in

Scheme 1) from L-amino acids (as protected esters), and investigate their interactions with the lanthanide ion Tb(III) by probing their ground and excited state properties in solution. The lanthanides have high coordination requirements which can be fulfilled with such tridentate ligands.²⁴ As depicted in Fig. 1, the formation of the 1 : 3 (M : L) stoichiometry would give rise to a fully saturated coordination environment, where each ligand coordinates with the Ln(III) ion *via* the pyridine nitrogen, as well as the carboxylic unit, and the triazole moiety. However, other stoichiometries can also exist which can be determined by carrying out spectroscopic titration studies and fitting the resulting data using non-linear regression analysis. Here, we give a full account of our investigation for ester ligands **4** and **5** as well as the carboxylic acid derivative **6**, formed from **4**. We also show that in the presence of the lanthanide ion Tb(III), the ligands form luminescent complexes both under thermodynamic and kinetic conditions.

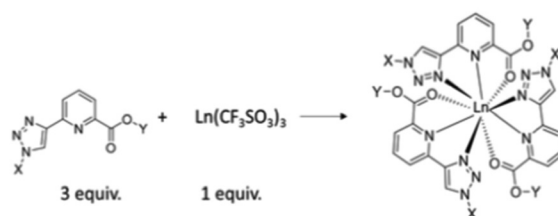


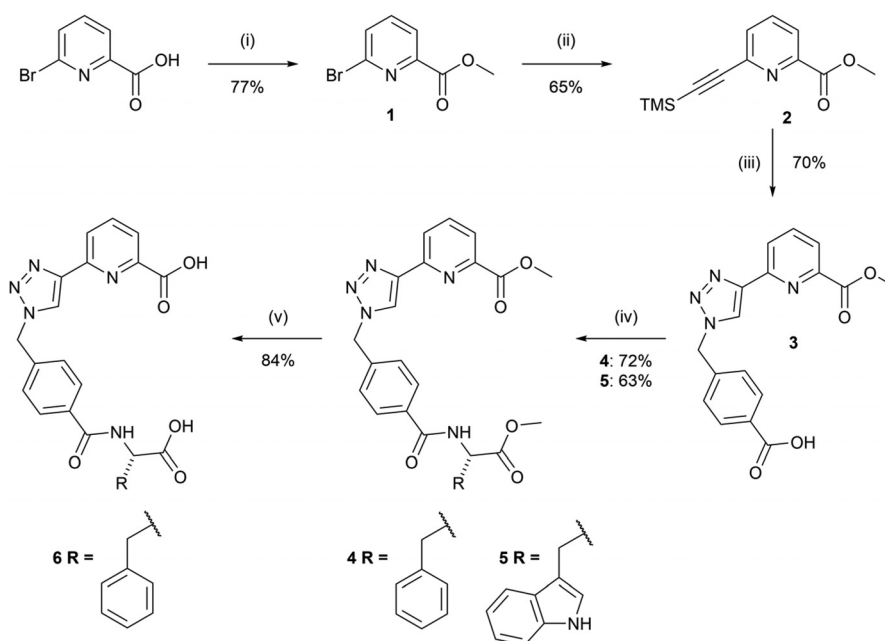
Fig. 1 Graphical representation of the (1,2,3-triazol-4-yl)-picolinamide (**tzpa**) core (X = phenyl, amino acid derived phenyl linker, etc.; Y = H, ester, amide, etc.), and the formation of the 1 : 3 (M : L) complex with lanthanide (Ln) salts. Such complexes (e.g. using ligands **4** and **5** developed herein) can be formed either under thermodynamic or kinetic control.

^aSchool of Chemistry and Trinity Biomedical Science Institute, University of Dublin, Trinity College Dublin, Dublin 2, Ireland. E-mail: gunnlaut@tcd.ie

^bSchool of Chemical and Physical Sciences, Keele University, Keele ST5 5BG, UK

†Electronic supplementary information (ESI) available: Additional spectroscopic data and crystallographic data table. CCDC 2239920. For ESI and crystallographic data in CIF or other electronic format see DOI: <https://doi.org/10.1039/d3qo00232b>





Scheme 1 The structure and synthesis of compounds 1–6. Reagents and conditions: (i) CH_3OH , H_2SO_4 , (ii) ethynyltrimethylsilane, $\text{THF}:\text{NEt}_3$, $\text{Pd}(\text{PPh}_3)_4$, CuI (iii) NaN_3 , $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$, 4-(bromomethyl)benzoic acid, sodium ascorbate, K_2CO_3 , $\text{DMF}/\text{H}_2\text{O}$ (iv) **4**: L-phenylalanine HCl **5**: L-tryptophan HCl, HOBt , $\text{DMF}/\text{CH}_2\text{Cl}_2$, NEt_3 , and $\text{EDC} \cdot \text{HCl}$. (vi) NaOH , 100°C .

Results and discussion

Synthesis and characterisation of ligands 4–6 and the $\text{Tb}(\text{III})$ complexes of 4 and 5

The syntheses of the chiral **tzpa**-derived ligands **4** and **5** from (S) phenylalanine or tryptophan methyl esters were undertaken (Scheme 1). These amino acids were chosen as they are known to be able to populate the $\text{Tb}(\text{III})$ excited state through sensitisation, or the so-called antenna effect.²⁵ Having developed **btp** based amino acid derived structures in the past with the view of making ‘peptide bundles’,²⁶ our initial objective was to synthesise the methyl ester-protected ligands, which could then be hydrolysed and further synthetically modified. In that work we also showed that the $\text{Tb}(\text{III})$ complexes are significantly more emissive than their corresponding $\text{Eu}(\text{III})$ complexes. Although we did study in part the interaction of **4** with $\text{Eu}(\text{III})$, the focus herein is on $\text{Tb}(\text{III})$ systems.

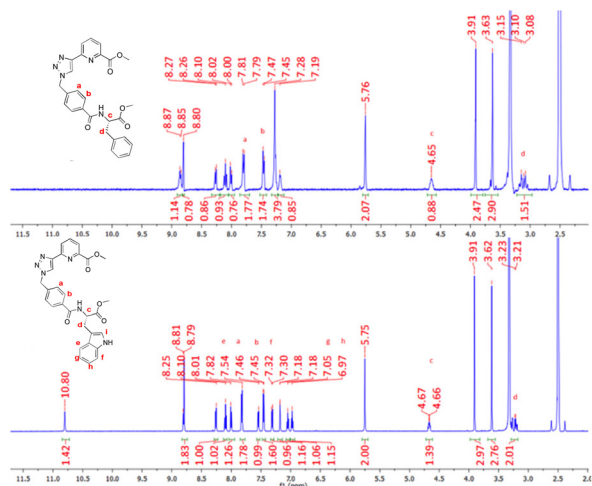
The route to **4** and **5** included four steps starting with the formation of methyl-6-bromopyridinate, **1**, from 6-bromopyridine-2-carboxylic acid in the presence of sulfuric acid and CH_3OH under reflux conditions. This compound was then subjected to Sonogashira coupling with ethynyltrimethylsilane in a THF/NEt_3 mixture using $\text{Pd}(\text{PPh}_3)_4$ as a catalyst. The crude product **2** was isolated as a brown oil, and purified using column chromatography to give **2** as a brown solid in 65% yield. 4-(Bromomethyl)benzoic acid was then converted to the relevant azide using sodium azide and employed in a CuAAC reaction *in situ* with compound **2** to generate compound **3**. **3** was isolated by initial washing with an EDTA solution in aqueous ammonia, followed by acidification using 1M HCl

which resulted in the precipitation of the desired product as a white solid in 70% yield. Compound **3** was then reacted with L-phenylalanine methyl ester hydrochloride or L-tryptophan methyl ester hydrochloride in a peptide coupling reaction to generate compounds **4** and **5** as white solids in yields of 72% and 63%, respectively.

Compound **4** could also be hydrolysed with 1M NaOH under reflux conditions overnight. As $\text{Ln}(\text{III})$ coordination would be around the **tzpa** pocket and not involve the terminal amino acid moiety, this provides a handle for further modification, for instance in the generation of soft materials, where the acid unit is employed to extend the overall supramolecular assembly. This we have demonstrated for both **btp** and **dpa** ligands in the past, which in the presence of $\text{Ln}(\text{III})$ ions formed soft materials.²⁷ With this in mind, we hydrolysed the ester of **4** to give **6** in a 1M NaOH solution under reflux for 3 hours. Upon cooling and subsequent addition of 1M HCl, compound **6** precipitated out of solution as a pure white solid.

All compounds discussed above were characterised by NMR spectroscopy, HRMS and IR spectroscopy (see select examples in the ESI†). The ^1H NMR spectra of compounds **4** and **5** are shown in Fig. 2 (see the ESI† for their ^{13}C NMR spectra). The resonance corresponding to the pyridyl protons of both ligands reside at 8.26, 8.10 and 8.01 ppm while the triazole protons lie downfield at 8.81 ppm. The aromatic protons H_a and H_b resonate at 7.82 and 7.47 ppm as two doublets and the adjacent CH_2 protons are located at 5.75 ppm in both compounds. The multiplet upfield at 4.65 ppm corresponds to the proton H_c while the methyl protons associated with the pyridyl ring and the amino acid moiety lie at 3.91 and 3.62 ppm,





appeared at 1278 cm^{-1} in the case of $[\text{Tb}(\mathbf{4})_3]^{3+}$, further indicating the successful complexation of **4** with Tb(III) . Similar changes were observed in the IR spectra of compound **5** and $[\text{Tb}(\mathbf{5})_3]^{3+}$.

Photophysical characterisation of complexes $[\text{Tb}(\mathbf{4})_3]^{3+}$ and $[\text{Tb}(\mathbf{5})_3]^{3+}$

After obtaining the two Tb(III) complexes ($[\text{Tb}(\mathbf{4})_3]^{3+}$ and $[\text{Tb}(\mathbf{5})_3]^{3+}$) their photophysical properties were investigated in CH_3CN solution. The UV-visible absorption and the excitation spectra of the two complexes were dominated by two main broad absorption bands that were centred at $\lambda = 245\text{ nm}$ and $\lambda = 293\text{ nm}$ (see the ESI†). These bands can be assigned to $\pi-\pi^*$ transitions in the ligand structure. The fluorescence emission spectra were also recorded (see further discussion below), showing a ligand centred emission centred at around 350 nm . The excitation spectra (using Tb(III) λ_{em} at 545 nm) were structurally similar to the absorption spectra (see overlapped spectra in the ESI†), indicating that the excitation of the ligand gave rise to energy transfer to the Tb(III) excited state. Furthermore, this Tb(III) -centred emission was also clearly visible to the naked eye as seen in the inset in Fig. 4.

The sensitisation of the Tb(III) excited state was confirmed by observing the delayed emission at $\lambda = 255\text{ nm}$ which showed the presence of Tb(III) -centred luminescence in both complexes indicating effective population of the $^5\text{D}_4$ excited state of the Ln(III) ion and subsequent deactivation of the $^7\text{F}_J$ ($J = 6-2$) ground states with line-like emission bands appearing at 490 nm , 545 nm , 584 nm , 622 nm and 645 nm , the latter of which was very weak, as can be seen in Fig. 4 for $[\text{Tb}(\mathbf{4})_3]^{3+}$ (see the ESI† for $[\text{Tb}(\mathbf{5})_3]^{3+}$).

To understand further the coordination environment around the lanthanide, the Tb(III) centred luminescence lifetimes (τ_{obs}) were recorded for $[\text{Tb}(\mathbf{4})_3]^{3+}$ and $[\text{Tb}(\mathbf{5})_3]^{3+}$ in H_2O and D_2O in which the complexes were sparingly soluble. The emission decay (at $\lambda_{\text{em}} = 545\text{ nm}$) was measured at various gate

times upon excitation of the ligand, and the average radiative lifetimes of the delayed emission were calculated from the decay. The lifetimes of the $^5\text{D}_4$ excited state of Tb(III) were best fitted as a bi-exponential decay for $[\text{Tb}(\mathbf{4})_3]^{3+}$ and $[\text{Tb}(\mathbf{5})_3]^{3+}$ in H_2O and D_2O which indicates that there was more than one emissive species in solution. These results are summarised in Table 1.

Upon repetition of the experiment in CH_3OH and CD_3OD , in which the complexes were fully soluble, the lifetimes were the best fit to a mono-exponential decay for both complexes, indicating that there was a single emissive species in solution. This is likely due to the improved stability of $[\text{Tb}(\mathbf{4})_3]^{3+}$ and $[\text{Tb}(\mathbf{5})_3]^{3+}$ in methanolic solution which likely prevents any degree of dissociation of the tris complex, as occurring in aqueous solution.

From the fit decay curves, the excited state lifetimes in H_2O and D_2O as well as CH_3OH and CD_3OD were determined (Table S1†). From these data the hydration state (the number of metal ion bound water molecules), q , was also determined.^{28,29} From this analysis, the q value determined for $[\text{Tb}(\mathbf{4})_3]^{3+}$ in an aqueous suspension was ~ 3 . While this determination has large errors due to the poor solubility, the analysis would imply that three water molecules were bound to the Tb(III) -centre. This could also suggest the presence of the ML_2 species in the suspension as we had observed before in HRMS. Being highly competitive media, this could also indicate that the complex undergoes ligand dissociation in aqueous media. As mentioned, then the fitting of the excited state decay was best fit to bi-exponential decay, and from the fit, a second shorter lifetime was observed from which a q value of approximately 6 was determined, which could be speculated to be due to the presence of a ML species in solution. This would also contribute to the poor luminescence of this system in water where the $^5\text{D}_4$ excited state of Tb(III) is quenched by O–H oscillators.³⁰ However, in methanolic solution, in which the complex was stable, the decay was mono-exponential, and using Horrocks' equation, the hydration state was determined to be approximately zero. This suggests that in this less competitive solution, the full coordination requirement of the Tb(III) was fulfilled by the ligand and the ML_3 species was formed. The same trend and outcome were also observed upon determining the hydration states obtained for $[\text{Tb}(\mathbf{5})_3]^{3+}$ (see ESI Table S2†).

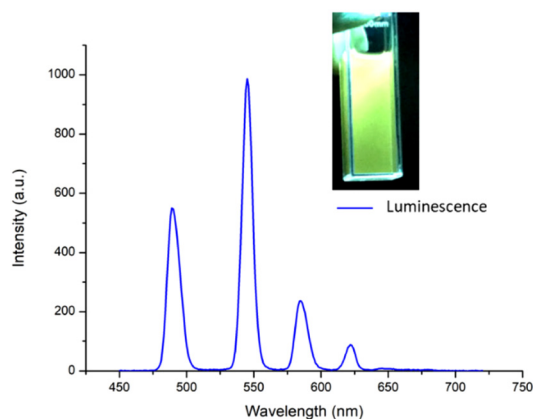


Fig. 4 The delayed luminescence spectra of $[\text{Tb}(\mathbf{4})_3]^{3+}$ ($2.7 \times 10^{-5}\text{ M}$) recorded in CH_3CN showing the characteristic Tb(III) transitions $^5\text{D}_4 \rightarrow ^7\text{F}_{6,5,4,3,2}$. Inset: green Tb(III) emission that was visible to the naked eye upon placing the sample under a UV-Vis lamp.

Table 1 Luminescence lifetimes (τ) recorded from aqueous suspensions and methanolic solutions of $[\text{Tb}(\mathbf{4})_3]^{3+}$ and $[\text{Tb}(\mathbf{5})_3]^{3+}$ with q -values (hydration states). The data for $[\text{Tb}(\mathbf{6})_3]^{3+}$ are also shown

Complex	τ (H_2O), ms	τ (D_2O), ms	τ (CH_3OH), ms	τ (CD_3OD), ms	q (± 0.5)
$[\text{Tb}(\mathbf{4})_3]^{3+}$			1.905	2.213	0.13
$[\text{Tb}(\mathbf{5})_3]^{3+}$			1.943	2.463	0.52
$[\text{Tb}(\mathbf{4})_3]^{3+} \tau_1$	0.401	0.751			5.53
$[\text{Tb}(\mathbf{4})_3]^{3+} \tau_2$	1.095	2.865			2.52
$[\text{Tb}(\mathbf{5})_3]^{3+} \tau_1$	0.376	0.749			6.32
$[\text{Tb}(\mathbf{5})_3]^{3+} \tau_2$	1.014	2.424			2.56
$[\text{Tb}(\mathbf{6})_3]^{3+}$	1.561	2.203			0.63



The quantum yield ($\Phi_{\text{tot}}^{\dagger}$) was also determined for both complexes (λ_{ex} at 255 nm), and found to be 84.0(7)% for $[\text{Tb}(\mathbf{4})_3]^{3+}$. This is similar to that observed for most $\text{Tb}(\text{III})$ based **btp** complexes that have high Φ_{tot} . However, in contrast, the $[\text{Tb}(\mathbf{5})_3]^{3+}$ complex was found to be significantly less emissive with $\Phi_{\text{tot}} = 15.3(7)\%$. This is, however, in line with our previous observation that a **btp** based-Trp ligand appears to be a poor sensitizer for $\text{Tb}(\text{III})$ when compared to its Phe counterpart.²⁶

The above results demonstrate that $\text{Tb}(\text{III})$ is in general a 'good lanthanide ion' for use in the formation of **tzpa** based complexes and that the photophysical properties of these systems mirror those generated by **btp** ligands. Unfortunately, the stability of these complexes in aqueous solution was not desirable, while they showed good stability in both CH_3CN and MeOH solutions where the coordination environment is fully saturated.

Observing the self-assembly formation between **4** and **5** and $\text{Tb}(\text{III})$ under kinetic control

Having probed the photophysical properties of the two complexes $[\text{Tb}(\mathbf{4})_3]^{3+}$ and $[\text{Tb}(\mathbf{5})_3]^{3+}$ that formed under microwave-aided synthesis, the $\text{Tb}(\text{III})$ directed self-assembly of **4** and **5** with $\text{Tb}(\text{CF}_3\text{SO}_3)_3$ was investigated *in situ* by carrying out a series of spectroscopic measurements (*i.e.* under kinetic control) under ambient conditions. Our previous work in this area has shown that these assemblies tend to be more dynamic and varied in stoichiometry. The main focus here will be on discussing the changes observed using **4**. (The changes seen for **5** are given in the ESI†) To probe these changes, specific aliquots of $\text{Tb}(\text{III})$ equivalents (in solution) were added to a solution of **4** (1×10^{-5} M) and the changes in the UV-visible absorption, fluorescence and delayed $\text{Tb}(\text{III})$ luminescence spectra were monitored in CH_3CN solution. All measurements were repeated three times to ensure full reproducibility.

Ground state and fluorescence emission changes. The UV-visible absorption of compound **4** showed two main bands in the absorbance spectrum. A high-energy band was centred at $\lambda = 245$ nm while a lower energy band was observed at $\lambda = 291$ nm ($\epsilon = 24\,100\text{ cm}^{-1}\text{ M}^{-1}$ at $\lambda_{\text{abs}} = 242$ nm). Upon addition of $\text{Tb}(\text{III})$, the band centred at $\lambda = 245$ nm experienced a hypochromic shift and the appearance of a new band at $\lambda = 225$ nm was also observed, as demonstrated in Fig. 5. These changes reached a maximum at approximately 0.5 equivalents of $\text{Tb}(\text{III})$, after which no further significant changes occurred. The low energy band at $\lambda = 291$ nm showed a hyperchromic shift and a concomitant redshift to $\lambda = 300$ nm upon the addition of $\text{Tb}(\text{III})$. Again, these changes reached a plateau at the addition of

0.5 equivalents of $\text{Tb}(\text{III})$. Two isosbestic points were observed at $\lambda = 236$ nm and $\lambda = 290$ nm, indicating the presence of multiple correlated species in solution. Similarly, for **5**, the main changes occurred up to the addition of 0.5 equivalents of $\text{Tb}(\text{III})$ after which they reached a plateau (see the ESI†).

Upon excitation of ligand **4** at $\lambda = 255$ nm, the ligand gave rise to broad fluorescence emission with $\lambda = 353$ nm, Fig. 5. Upon the addition of $\text{Tb}(\text{III})$, the ligand fluorescence was significantly affected by the complexation process, becoming gradually quenched, with *ca.* 70% reduction in the fluorescence emission intensity at the addition of 0.5 equivalents of $\text{Tb}(\text{III})$, Fig. 5. In fact, by the addition of 0.7 equivalents of $\text{Tb}(\text{III})$ the fluorescence was fully quenched. As before, this can be attributed to an energy transfer process from ligand **4** to the $\text{Tb}(\text{III})$ metal centre which gives rise to $\text{Tb}(\text{III})$ -centred emission. Indeed this we observed as discussed next. Again, similar results were observed in the fluorescence emission spectra of **5** upon titrating with $\text{Tb}(\text{III})$ (see the ESI†). However, the emission was fully quenched by the addition of 0.5 equivalents of $\text{Tb}(\text{III})$.

Lanthanide luminescence changes. The changes in the delayed $\text{Tb}(\text{III})$ -centred emission were recorded upon the excitation of **4** at $\lambda = 255$ nm. The spectra and binding isotherms can be seen in Fig. 6 demonstrating a gradual enhancement in the $\text{Tb}(\text{III})$ -centred transitions at $\lambda = 490, 545, 584, 622, 648, 667$ and 679 nm, upon deactivation of the $^5\text{D}_4$ state to the $^7\text{F}_J$ ($J = 6-0$) states, respectively. As shown in Fig. 6, emission gradually enhanced up to the addition of 0.5 equivalents of $\text{Tb}(\text{III})$. Subsequent additions resulted, however, in some degree of quenching of the $\text{Tb}(\text{III})$ emission. The changes in the $\text{Tb}(\text{III})$ emission upon titrating **5** were also monitored concomitantly (see the ESI†). However, unlike that seen for **4**, no significant enhancement was observed in the $\text{Tb}(\text{III})$ -centred emission before the addition of 0.5 equivalents of $\text{Tb}(\text{III})$, after which, the emission increased dramatically and continued to increase beyond 5 equivalents of $\text{Tb}(\text{III})$. This could be due to slow kinetics under ambient conditions.

Fitting the spectroscopic changes using non-linear regression analysis. With the view of shedding light on the changes above, and to explore which stoichiometries dominated in solution, the above data were analysed by using non-linear regression analysis, and the speciation in solution as well as the stability constants estimated from the global analysis (using ReactLab EQUILIBRIA software). Initially, we fitted the changes to the UV-visible absorption spectrum. Analysis of the UV-visible absorption titration data for ligand **4** with $\text{Tb}(\text{CF}_3\text{SO}_3)_3$ pointed to the presence of four main species (**4**, 1 : 1 metal : ligand species, 1 : 2 metal : ligand species and 1 : 3 M : L species) in solution as shown in Fig. 7. The distribution of these species was estimated in the analysis, which showed that from $0 \rightarrow 0.25$ equivalents of $\text{Tb}(\text{III})$, both the 1 : 2 metal : ligand and 1 : 3 metal : ligand species co-exist in solution, both reaching approximately 25% abundance. However, the subsequent additions of $\text{Tb}(\text{III})$ resulted in a decrease in the abundance of the ML_3 species while the ML_2 species continued to increase, reaching 92% abundance at 0.5 equivalents of $\text{Tb}(\text{III})$. From $0.5 \rightarrow 3$ equivalents, the ML_2 underwent gradual dis-

[†]The photoluminescence quantum yield (Φ_{tot} %) was determined in CH_3CN using a relative method against a standard developed by Bünzli and co-workers.³⁰ This relative method for quantum yield determination is described in the following equation:

$$\Phi_{\text{REL}}^{\text{Tb,L}} = \frac{\phi_{\text{S}}}{\phi_{\text{R}}} \times \frac{E_{\text{S}}}{E_{\text{R}}} \times \frac{A_{\text{R}}(\lambda_{\text{R}})}{A_{\text{S}}(\lambda_{\text{S}})} \times \frac{I_{\text{R}}(\lambda_{\text{R}})}{I_{\text{S}}(\lambda_{\text{S}})} \times \frac{n_{\text{S}}^2}{n_{\text{R}}^2}$$



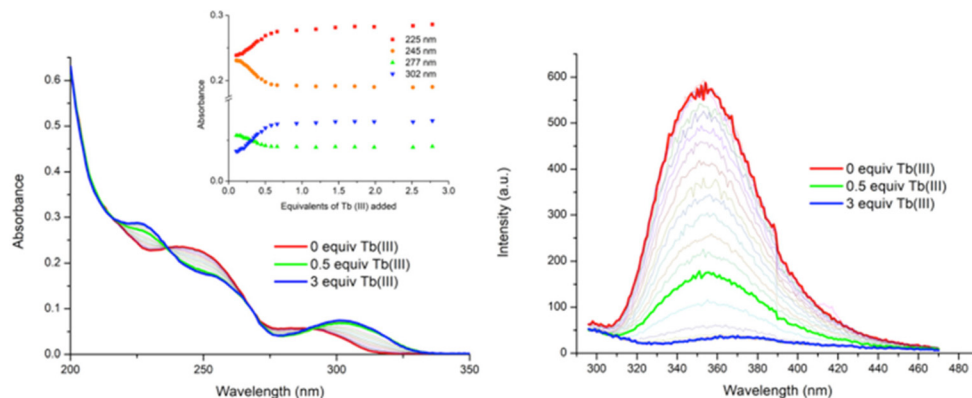


Fig. 5 The observed spectroscopic changes for **4** (1×10^{-5} M) against Tb(CF₃SO₃)₃ (0 → 3 equiv.) in CH₃CN at RT. (Left) the UV-visible absorption spectra. Inset: corresponding experimental binding isotherms with absorbance at $\lambda = 225, 245, 277$ and 300 nm. (Right) the corresponding changes in the fluorescence emission spectra, $\lambda_{\text{ex}} = 255$ nm.

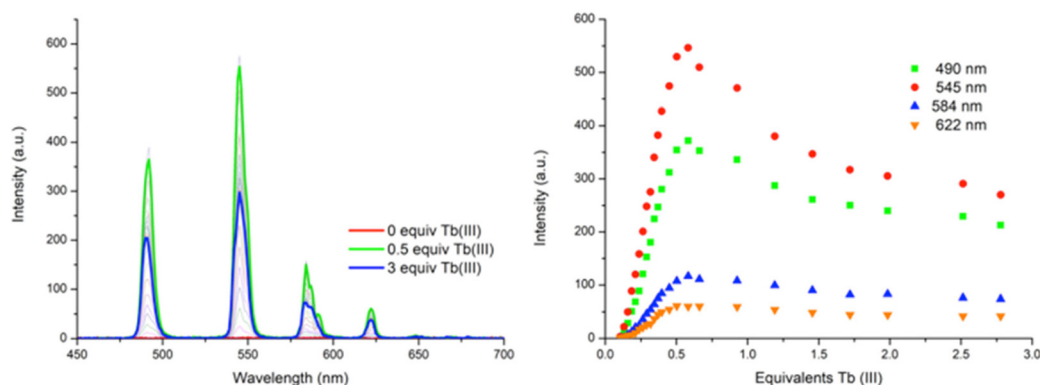


Fig. 6 (Left) The overall changes to the Tb(III)-centred emission spectra upon titrating **4** (1×10^{-5} M) against Tb(CF₃SO₃)₃ (0 → 3 equiv.) in CH₃CN at RT. (Right) The corresponding experimental binding isotherms for the Tb(III) emission at $\lambda = 492, 545, 585$ and 623 nm.

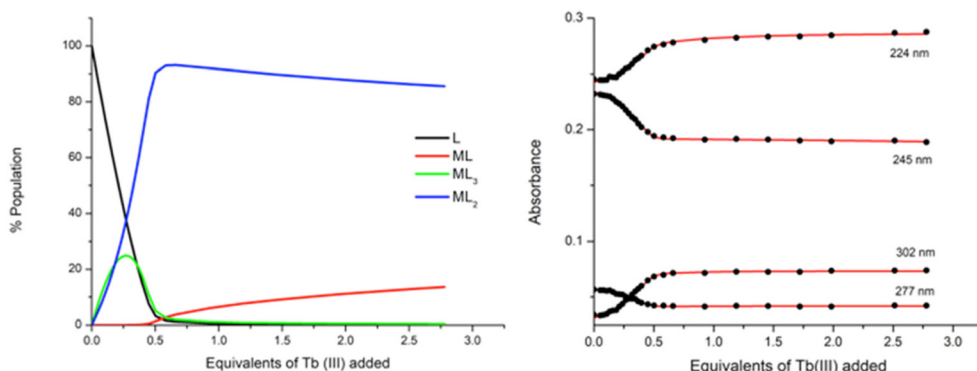


Fig. 7 (Left) the speciation distribution diagram obtained from the fit of the UV-visible absorption titration data of ligand **4** against Tb(CF₃SO₃)₃ in CH₃CN. (Right) the fit of the experimental binding isotherms using non-linear regression analysis software ReactLab.

sociation into the ML species due to the presence of excess metal ions in solution. And while the ML₃ species is likely the most emissive species in solution due to the exclusion of solvent molecules from the inner coordination sphere of the

Tb(III) ion, the ML₂ species is formed more favourably at this low concentration under kinetic control.

The stepwise binding constants were estimated for each species from the changes in the UV-visible absorption spectra



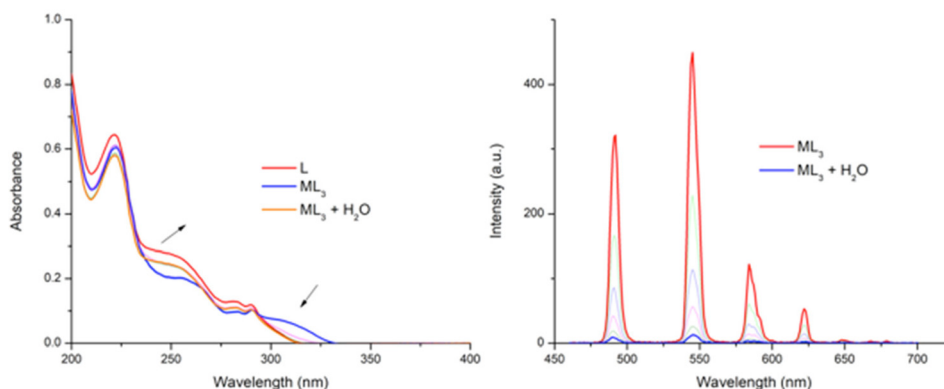


Fig. 8 (Left) Absorbance spectra of ligand **6** upon addition of Tb(III) at 0 and 3 equivalents of Tb(III), and upon addition of H₂O. (Right) the Tb(III) emission spectrum of the 1 : 3 complex showing decrease in luminescence intensity upon the addition of H₂O.

throughout the titration and are expressed here as $\log \beta_{ML}$. The 1 : 2 metal : ligand species formed with $\log \beta_{12} = 14.4$ and the 1 : 3 metal : ligand species had a calculated binding constant of $\log \beta_{13} = 19.5$ while the ML species was estimated as $\log \beta_{11} = 6.3 \pm 0.1$.

In a similar manner, the analysis of **5** with Tb(III) pointed to the presence of four main species in solution upon fitting the changes in the ground state (see the ESI†). The distribution of these species was also estimated from the analysis and indicates that from 0 $\rightarrow$ 0.5 equivalents of Tb(III), both the ML₂ and ML₃ species coexist in solution. However, the ML₃ species reaches a maximum abundance of 12% at 0.25 equivalents of Tb(III), while the ML₂ species is significantly more dominant at 94% abundance upon the addition of 0.5 equivalents of Tb(III). Subsequent additions of Tb(III) resulted in a gradual dissociation of the ML₂ species into the less emissive ML species. Binding constants were estimated from the analysis and the ML₃ species was found to form with $\log \beta_{13} = 19.2 \pm 0.2$ while the ML₂ species had a calculated binding constant of $\log \beta_{12} = 14.6 \pm 0.2$ and finally, the ML species was calculated as $\log \beta_{11} = 6.6 \pm 0.1$. These values are in good agreement with those seen for **4** above indicating that the amino acid moiety does not have a large influence on the self-assembly in solution.

Unfortunately, while attempts were made to obtain fits for either the fluorescence or the Tb(III) luminescence titration data obtained for **4** and **5**, they did not result in data convergence and reliable binding constants could not be obtained.

Probing the interaction of Tb(III) with ligand **6**

Having examined the amino ester derivatives **4** and **5**, we also investigated the ability of the hydrolysed ligand **6** to interact with Tb(III), and to see if the carboxylic acid had any effect on the self-assembly formation with Tb(III) under kinetic control, the measurements were carried out in CH₃CN. The changes in the absorption and the Tb(III) emission are shown in Fig. 8 upon the addition of 3 equivalents of lanthanide. Initially, the 1 : 3 complex was found to be stable in CH₃CN solution over 24 hours (by monitoring the UV-visible absorption spectra), with no changes being observed over the duration.

Furthermore, lifetime and q -value determination (Table 1) showed that the excited state decay was best fitted to a mono-exponential decay, and that the q value was ~ 0 . Given the improved solubility of the ligand **6** in competitive media we titrated water into this solution. However, this, as the changes in Fig. 8 demonstrate, resulted in the dissociation of the self-assembly, where the UV-visible absorption spectrum returns to the absorbance profile of ligand **6**. Similarly, the Tb(III) emission intensity was also gradually quenched upon the addition of H₂O.

Conclusion

The synthesis of two new **tzpa** ligands **4** and **5** from the α -amino acid (S) L-phenylalanine or L-tryptophan (as their methyl esters) was achieved in a few step synthesis. The solid-state crystal structure of **4** was also obtained. The hydrolysis of the amino ester **4** also produced ligand **6**. Both ligands **4** and **5** were shown to form 1 : 3 complexes with Tb(III) under thermodynamic control, which were highly luminescent, and the green Tb(III) emission was clearly visible to the naked eye. In contrast to this, under kinetic control, the global analysis of the titration data obtained from the ground state changes with Tb(III) showed that for both **4** and **5**, the 1 : 2 stoichiometry dominated for both systems after the addition of 0.5 equivalents of the ion. This was also confirmed by investigating the excited state lifetimes of the Tb(III) centred emission. From these changes we were able to determine both the binding affinity and the stoichiometry in solution. This 'in solution' (titration) behaviour is different to that obtained normally for either **btp** or **dpa** ligands based on which the **tzpa** structure is modelled, where the 1 : 3 stoichiometry is dominant under both thermodynamic and kinetic conditions. Both ligands provide a platform to explore the chemistry of the **tzpa** structure further in supramolecular and coordination chemistries, an endeavour which we are currently engaged in. By structurally modulating both the carboxylic site as well as the triazole moiety (e.g. both X and Y sites, Fig. 1), the formation of



dimeric ligand structures that can be used in the formation of dimetallic helicates and other supramolecular structures is currently being undertaken in our laboratory.

Experimental

Methyl-6-bromopicolinate (1)

6-Bromopicolinic acid (3.0 g, 14.85 mmol) was suspended in MeOH (5 mL) and 4 drops of sulfuric acid were added. The mixture was heated to reflux for four hours and quenched with NaHCO₃ in water. The product was extracted into DCM, washed with sodium hydrogen carbonate solution and dried over magnesium sulphate to yield a white solid (2.095 g, 9.65 mmol, 77%). The product decomposed over 91 °C. HRMS (*m/z*) (ESI+): calculated for C₇H₇BrNO₂⁺ *m/z* = 215.9660 [M + H]⁺. Found *m/z* = 215.9654; ¹H NMR (400 MHz, CDCl₃-d₆): δ (ppm) = 8.10 (dd, *J* = 6.7, 1.8 Hz, 1H, 3-pyr H), 7.69 (t, *J* = 6.7, 2H, 4-pyr H), 7.70 (dd, *J* = 6.7, 1.8 Hz, 1H, 5-pyr H), 4.00 (s, 3H, OCH₃). ¹³C NMR (150 MHz, DMSO-*d*₆): δ (ppm) = 163.8, 148.1, 141.3, 140.9, 132.1, 124.5, 52.8; IR ν_{max} (cm⁻¹): 3106, 2968, 1718, 1556, 1437, 1422, 1310, 1246, 1132, 1113, 985, 955, 757, 732.

Methyl 6-((trimethylsilyl)ethynyl)picolinate (2)

To a solution of 1 (1.650 g, 7.60 mmol) in THF/NEt₃ 4:1, 40 (mL), in the presence of CuI (0.2 mmol) and Pd(PPh₃)₄ (0.2 mmol), ethynyltrimethylsilane (1.1 mL, 7.60 mmol) was added dropwise at 0 °C under an argon atmosphere. The resulting solution was left to stir at room temperature for 48 h, after which the solution was filtered through a plug of Celite and concentrated under reduced pressure to yield a brown solid which was purified by flash column chromatography (RediSep® 12 g, gradient elution 0 → 90% EtOAc in hexane). This gave 2 as a brown solid in 65% yield (1.158 g, 4.94 mmol). The product decomposes over 85 °C. HRMS (*m/z*) (ESI+): calculated for C₁₂H₁₅NO₂Si⁺ *m/z* = 234.0950 [M + H]⁺. Found *m/z* = 234.0963; ¹H NMR (400 MHz, CDCl₃-d₆): δ (ppm) = 8.06 (d, *J* = 7.6 Hz, 1H, 3-pyr H), 7.80 (t, *J* = 8.1, 7.6 Hz, 1H, 4-pyr H), 7.63 (d, *J* = 8.1 Hz, 1H, 5-pyr H), 4.00 (s, 3H, OCH₃), 0.26 (s, 9H, Si-(CH₃)₃). ¹³C NMR (150 MHz, DMSO-*d*₆): δ (ppm) = 164.6, 147.8, 141.8, 138.5, 130.5, 124.6, 103.3, 95.1, 52.6, 0.4; IR ν_{max} (cm⁻¹): 3106, 3082, 2092, 1864, 1736, 1577, 1437, 1251, 840, 764.

4-((4-(6-(Methoxycarbonyl)pyridin-2-yl)-1H-1,2,3-triazol-1-yl)methyl)benzoic acid (3)

To a solution of 4-bromomethyl benzoic acid (1.418 g, 6.60 mmol) in 10 mL 4:1 DMF/water was added sodium azide (0.428 g, 6.60 mmol) and the reaction mixture was stirred for one hour to yield the azide intermediate that was not isolated and therefore used without further purification. To this solution was added ligand 2 (1.539 g, 6.560 mmol), CuSO₄·5H₂O (0.329 g, 1.32 mmol), sodium ascorbate (0.522 g, 2.64 mmol) and anhydrous K₂CO₃ (0.912 g, 6.60 mmol) and stirred at room temperature under an argon atmosphere overnight. 1 M

EDTA/NH₄OH solution was added to the mixture and the product was precipitated using 1M HCl and filtered to yield an off-white solid. The product was used without further purification. (1.562 g, 4.62 mmol, 70%). HRMS (*m/z*) (ESI+): calculated for C₁₇H₁₄N₄O₄H⁺ *m/z* = 339.1093 [M + H]⁺. Found *m/z* = 339.1087; ¹H NMR (400 MHz, DMSO-*d*₆): δ (ppm) = 8.82 (s, 1H, triazole CH), 8.26 (dd, *J* = 7.8, 1.1 Hz, 1H, 5-pyr H), 8.10 (s, *J* = 7.8, 1H, 4-pyr H), 8.00 (dd, *J* = 7.8, 1.1 Hz, 1H, 3-pyr H), 7.95 (d, *J* = 8.4 Hz, 2H, ArH-COOH), 7.47 (d, *J* = 8.4 Hz, 2H, ArH-CH₂), 5.79 (s, 2H, CH₂-Ar), 3.90 (s, 3H, COO-CH₃). ¹³C NMR (150 MHz, DMSO-*d*₆): δ (ppm) = 170.1, 165.6, 150.2, 147.6, 146.8, 137.2, 138.7, 135.3, 129.2, 128.1, 124.1, 123.9, 123.1, 52.7; IR ν_{max} (cm⁻¹): 1725, 1612, 3215, 1282, 1126, 1011, 776, 707.

General experimental procedure for amino acid coupling to 3

To a solution of the relevant amine (1 equivalent) in a DMF:DCM (4:1) mixture under argon were added HOBt (1 equivalent), NEt₃ (1.2 equivalents) and the relevant acid (1.1 equivalents). The reaction mixture was stirred for 30 min and cooled to 0 °C. EDC-HCl (1.5 equivalents) was then added to the suspension and stirred at 0 °C for a further 30 min. The mixture was then allowed to warm to room temperature and stirred for a further 48 h.

Methyl(S)-6-(1-(4-((1-methoxy-1-oxo-3-phenylpropan-2-yl)carbamoyl)benzyl)-1H-1,2,3-triazol-4-yl)picolinate (4)

Synthesised according to the general procedure mentioned above from phenylalanine hydrochloride (0.99 g, 4.91 mmol) and compound 3 (1.512 g, 4.47 mmol) to yield a white solid (1.611 g, 3.26 mmol, 72%). The product decomposes over 207 °C. HRMS (*m/z*) (ESI+): calculated for C₂₇H₂₅N₅O₅Na⁺ *m/z* = 522.1753 [M + Na]⁺. Found *m/z* = 522.1767; ¹H NMR (600 MHz, DMSO-*d*₆): δ (ppm) = 8.86 (d, *J* = 8.1 Hz, 1H, NH), 8.80 (s, 1H, triazole H), 8.26 (d, *J* = 1.1, 7.7 Hz, 1H, 5-pyr H), 8.10 (m, *J* = 1.1 Hz, 1H, 4-pyr H), 8.01 (d, *J* = 1.1, 7.7 Hz, 1H, 3-pyr H), 7.80 (d, *J* = 8.3 Hz, 2H, Ar H-C=O), 7.46 (d, *J* = 8.3 Hz, 2H, Ar H-CH₂), 7.28 (m, 4H, phe H), 7.19 (s, 1H, phe H), 5.76 (s, 2H, CH₂), 4.65 (s, *J* = 8.1 Hz, 1H, CH-NH), 3.91 (s, 3H, pyr-COOCH₃), 3.63 (s, 3H, COOCH₃), 3.23–2.97 (m, 2H, CH₂-Phe). ¹³C NMR (150 MHz, DMSO-*d*₆): δ (ppm) = 172.1, 165.9, 165.0, 150.1, 147.6, 146.7, 139.2, 138.7, 137.6, 133.5, 129.0, 128.2, 127.9, 127.9, 126.5, 124.1, 123.8, 122.9, 54.2, 52.6, 52.5, 51.9, 36.2; IR ν_{max} (cm⁻¹): 3314, 1732, 1719, 1637, 1600, 1539, 1438, 1349, 1246, 1135, 1221, 1006, 825, 776, 752, 706, 644.

(S)-6-(1-(4-((1-Carboxy-2-phenylethyl)carbamoyl)benzyl)-1H-1,2,3-triazol-4-yl)picolinic acid (6)

Compound 4 (1.611 g, 3.23 mmol) was suspended in 1M NaOH and refluxed for 3 hours. The product was isolated by neutralization of the solution with acetic acid and filtration to yield a white solid (1.278 g, 2.72 mmol, 84%). HRMS (*m/z*) (ESI+): calculated for C₂₅H₂₁N₅O₅Na⁺ *m/z* = 494.1440 [M + Na]⁺. Found *m/z* = 494.1459; ¹H NMR (400 MHz, DMSO-*d*₆): δ (ppm) = 8.77 (s, 1H, triazole H), 8.25 (s, 1H, NH), 8.00 (d, *J* = 8.1 Hz, 1H, 5-pyr H), 7.91 (m, *J* = 7.6, 8.1 Hz, 1H, 4-pyr H), 7.85 (d, *J* =



7.6 Hz, 1H, 3-pyr H), 7.73 (d, $J = 8.2$ Hz, 2H, Ar H-C=O), 7.42 (d, $J = 8.2$ Hz, 2H, Ar H-CH₂), 7.23–7.07 (m, 5H, phe H), 5.74 (s, 2H, CH₂), 4.33 (s, 1H, CH-NH), 3.04 (dd, 2H, CH₂-Ph); ¹³C NMR (150 MHz, DMSO-*d*₆): δ (ppm) = 173.10, 167.22, 165.79, 154.12, 149.30, 149.27, 147.75, 139.22, 139.13, 138.37, 134.84, 129.69, 128.40, 128.08, 126.48, 124.37, 123.41, 121.20, 55.29, 53.16, 37.24; IR $\nu_{\max}$ (cm⁻¹): 3355, 3059, 2436, 1960, 1736, 1634, 1535, 1503, 1381, 1180, 1052, 851, 812, 737.

Methyl (S)-6-(1-(4-((3-(1H-indol-3-yl)-1-methoxy-1-oxopropan-2-yl)carbamoyl)benzyl)-1H-1,2,3-triazol-4-yl)picolinate (5)

Synthesised according to the general procedure mentioned above using tryptophan hydrochloride (0.15 g, 0.59 mmol) and compound 3 (0.20 g, 0.59 mmol) to yield compound 5 as a white solid (0.20 g, 63%); HRMS (m/z) (ESI⁺): C₂₉H₂₆N₆O₅⁺ m/z = 538.1965 [M + H]⁺. Found m/z = 539.2037; ¹H NMR (600 MHz, DMSO-*d*₆): δ (ppm) = 10.80 (s, 1H, NH), 8.81 (s, 1H, NH), 8.79 (s, 1H, triazole H), 8.26 (d, $J = 7.9$ Hz, 1H, Ha), 8.10 (t, $J = 7.8$ Hz, 1H, Hb), 8.00 (d, $J = 7.7$ Hz, 1H, Hc), 7.82 (d, $J = 8.2$ Hz, 2H, Phe-CH₂), 7.54 (d, $J = 7.9$ Hz, 1H, Hd), 7.46 (d, $J = 8.2$ Hz, 2H, Phe-C=O), 7.31 (d, $J = 8.1$ Hz, 1H, Hg), 7.18 (s, 1H, Hh), 7.05 (t, $J = 7.5$ Hz, 1H, Hg), 6.97 (t, $J = 7.4$ Hz, 1H, He), 5.75 (s, 2H, CH₂-triazole), 4.67 (d, $J = 5.4$ Hz, 1H, CH), 3.91 (s, 3H, CH₃-pyr), 3.62 (s, 3H, CH₃-Tryp), 3.24 (ddd, $J = 23.9, 14.6, 7.3$ Hz, 2H, CH₂-Tryp); ¹³C NMR (150 MHz, DMSO-*d*₆): δ (ppm) = 172.9, 166.5, 165.5, 150.6, 148.1, 147.1, 139.6, 139.2, 136.6, 134.1, 128.4, 124.6, 124.4, 124.1, 123.4, 121.4, 118.9, 118.5, 111.9, 110.4, 54.3, 53.1, 52.9, 52.4, 27.1; IR $\nu_{\max}$ (cm⁻¹): 3403, 3297, 3150, 3049, 2948, 2923, 2841, 1739, 1627, 1602, 1531, 1450, 1369, 1253, 1217, 1147, 1055, 1010, 974, 772, 737, 696, 600.

Synthesis of Tb-4₃

Ligand 4 (10 mg, 0.02 mmol) was added to Tb(CF₃SO₃)₃ (5 mg, 0.007 mmol) in CH₃OH (5 mL) and heated at 70 °C under microwave irradiation for 15 minutes. The resulting solution was dried under vacuum to yield a colourless oil which was taken up in CH₃OH and dropped slowly into a large excess of swirling Et₂O to afford Tb-4₃ as a white solid in 56% yield. The ¹H NMR spectra were broad and are shown in the ESI. HRMS (m/z) (ESI⁺): C₅₆H₅₀N₁₀O₁₆F₆S₂Tb m/z = 1455.2005 [Tb(4)₂(CF₃SO₃)₂]⁺. Found m/z = 1455.2075; elemental analysis for C₈₄H₇₅N₁₅O₂₄F₉S₃Tb·4H₂O·CH₃OH calculated: C 46.22 H 3.97 N 9.51; found: C 45.70 H 3.45 N 9.55.

Synthesis of Tb-5₃

Ligand 5 (10 mg, 0.02 mmol) was added to Tb(CF₃SO₃)₃ (5 mg, 0.007 mmol) in CH₃OH (5 mL) and heated at 70 °C under microwave irradiation for 15 minutes. The resultant solution was dried under vacuum to yield colorless oil which was taken up in CH₃OH and dropped slowly into a large excess of swirling Et₂O to afford Tb-5₃ as a white solid in 43% yield. The ¹H NMR spectrum was broad and is shown in the ESI.† HRMS (m/z) (ESI⁺): C₆₀H₅₂N₁₂O₁₆F₆S₂Tb m/z = 1533.2223 [Tb(5)₂(CF₃SO₃)₂]⁺ found m/z = 1533.2262.

Conflicts of interest

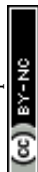
There are no conflicts to declare.

Acknowledgements

The authors gratefully acknowledge the Science Foundation Ireland (SFI PI Award 13/1A/1865 to TG) and the School of Chemistry, Trinity College Dublin for financial support. We thank Dr Gary Hessman for MS characterisation, Dr John O'Brien for NMR spectroscopy and Dr Manuel Rueher for instrumental support.

References

- (a) J. Dong, Y. Liu and Y. Cui, Supramolecular chirality in metal-organic complexes, *Acc. Chem. Res.*, 2021, **54**, 194; (b) C.-T. Yeung, K.-H. Yim, H.-Y. Wong, R. Pal, W.-S. Lo, S.-C. Yan, M. Yee-Man Wong, D. Yufit, D. E. Smiles, L. J. McCormick, S. J. Teat, D. K. Shuh, W.-T. Wong and G.-L. Law, Chiral transcription in self-assembled tetrahedral Eu4L6 chiral cages displaying sizable circularly polarized luminescence, *Nat. Commun.*, 2017, **8**, 1128; (c) D. A. Roberts, B. S. Pilgrim and J. R. Nitschke, Covalent post-assembly modification in metallosupramolecular chemistry, *Chem. Soc. Rev.*, 2018, **47**, 626; (d) A. J. Savyasachi, O. Kotova, S. Shanmugaraju, S. J. Bradberry, G. M. O'Máille and T. Gunnlaugsson, Supramolecular Chemistry: A Toolkit for Soft Functional Materials and Organic Particles, *Chem*, 2017, **3**, 764.
- (a) X.-Y. Luo and M. Pan, Metal-organic materials with circularly polarized luminescence, *Coord. Chem. Rev.*, 2022, **468**, 214640; (b) J. Gong and X. Zhang, Coordination-based circularly polarized luminescence emitters: Design strategy and application in sensing, *Coord. Chem. Rev.*, 2022, **453**, 214329; (c) D. E. Barry, D. F. Caffrey and T. Gunnlaugsson, Lanthanide-directed synthesis of luminescent self-assembled supramolecular structures and mechanically bonded systems from acyclic coordinating organic ligands, *Chem. Soc. Rev.*, 2016, **45**, 3244.
- (a) L.-J. Chen, H.-B. Yang and M. Shionoya, Chiral metallo-supramolecular architectures, *Chem. Soc. Rev.*, 2017, **46**, 2555; (b) A. M. Castilla, W. J. Ramsay and J. R. Nitschke, Stereochemistry in subcomponent self-assembly, *Acc. Chem. Res.*, 2014, **47**, 2063; (c) T. Sawada, A. Matsumoto and M. Fujita, Coordination-driven folding and assembly of a short peptide into a protein-like two-nanometer-sized channel, *Angew. Chem., Int. Ed.*, 2014, **53**, 7228; (d) C. M. G. dos Santos, A. J. Harte, S. J. Quinn and T. Gunnlaugsson, Recent developments in the field of supramolecular lanthanide luminescent sensors and self-assemblies, *Coord. Chem. Rev.*, 2008, **252**, 2512.
- (a) Y. B. Tan, M. Yamada, S. Katao, Y. Nishikawa, F. Asanoma, J. Yuasa and T. Kawai, Self-assembled tetra-



- nuclear Eu(III) complexes with D₂- and C_{2h}-symmetrical square scaffold, *Inorg. Chem.*, 2020, **59**, 12867; (b) J. Anhäuser, R. Puttreddy, Y. Lorenz, A. Schneider, M. Engeser, K. Rissanen and A. Lützen, Chiral self-sorting behaviour of [2.2]paracyclophane-based bis(pyridine) ligands, *Org. Chem. Front.*, 2019, **6**, 1226; (c) T. Gorai, W. Schmitt and T. Gunnlaugsson, Highlights of the development and application of luminescent lanthanide based coordination polymers, MOFs and functional nanomaterials, *Dalton Trans.*, 2021, **50**, 770; (d) J. A. Kitchen, Lanthanide-based self-assemblies of 2,6-pyridyldicarboxamide ligands: Recent advances and applications as next-generation luminescent and magnetic materials, *Coord. Chem. Rev.*, 2017, **340**, 232.
- 5 (a) Y. Wang, Y. Zhou, Z. Yao, W. Huang, T. Gao, P. Yan and H. Li, Asymmetric induction in quadruple-stranded europium(III) helicates and circularly polarized luminescence, *Dalton Trans.*, 2022, **51**, 10973; (b) S. D. Bonsall, M. Houcheime, D. A. Straus and G. Muller, Optical isomers of N, N'-bis(1-phenylethyl)-2,6-pyridinedicarboxamide coordinated to europium(III) ions as reliable circularly polarized luminescence calibration standards, *Chem. Commun.*, 2007, **35**, 3676.
 - 6 (a) J. Jiao, J. Dong, Y. Li and Y. Cui, Fine-tuning of chiral microenvironments within triple-stranded helicates for enhanced enantioselectivity, *Angew. Chem., Int. Ed.*, 2021, **60**, 16568; (b) J.-K. Ou-Yang, Y.-Y. Zhang, M.-L. He, J.-T. Li, X. Li, X.-L. Zhao, C.-H. Wang, Y. Yu, D.-X. Wang, L. Xu and H.-B. Yang, Unexpected self-assembly of chiral triangles from 90° chiral di-Pt(II) acceptors, *Org. Lett.*, 2014, **16**, 664.
 - 7 (a) R. Chen, Q.-Q. Yan, S.-J. Hu, X.-Q. Guo, L.-P. Zhou and Q.-F. Sun, Dinuclear mono-bridged or polymeric lanthanide complexes from one ligand: structural transformation and chiral induction, *Dalton Trans.*, 2023, **52**, 37; (b) L.-L. Yan, C.-H. Tan, G.-L. Zhang, L.-P. Zhou, J.-C. Bünzli and Q.-F. Sun, Stereocontrolled Self-Assembly and Self-Sorting of Luminescent Europium Tetrahedral Cages, *J. Am. Chem. Soc.*, 2015, **137**, 8550.
 - 8 (a) D. Luo, X.-Z. Wang, C. Yang, X.-P. Zhou and D. Li, Self-assembly of chiral metal-organic tetartoid, *J. Am. Chem. Soc.*, 2018, **140**, 118; (b) E. Badetti, K. Wurst, G. Licini and C. Zonta, Multimetallic architectures from the self-assembly of amino acids and tris(2-pyridylmethyl)amine Zinc(II) Complexes: circular dichroism enhancement by Chromophores organization, *Chem. – Eur. J.*, 2016, **22**, 6515; (c) Y. Ye, T. R. Cook, S.-P. Wang, J. Wu, S. Li and P. J. Stang, Self-Assembly of chiral metallacycles and metal-lacages from a directionally adaptable BINOL-derived donor, *J. Am. Chem. Soc.*, 2015, **137**, 11896.
 - 9 (a) T. Granch, C. Tourbillon, J. Ferrando-Soria, M. Julve, F. Lloret, J. Pasán, C. Ruiz-Pérez, O. Fabelo and E. Pardo, Self-assembly of a chiral three-dimensional manganese (II)-copper(II) coordination polymer with a double helical architecture, *CrystEngComm*, 2013, **15**, 9312; (b) A. D. Lynes, J. I. Lovitt, C. Rotella, J. J. Boland, T. Gunnlaugsson and C. S. Hawes, Crystal engineering studies of a series of pyridine-3,5-dicarboxamide ligands possessing alkyl ester arms, and their coordination chemistry, *Results Chem.*, 2022, **4**, 100679; (c) S. E. Bodman and S. J. Butler, Advances in anion binding and sensing using luminescent lanthanide complexes, *Chem. Sci.*, 2021, **12**, 2716.
 - 10 (a) J. P. Leonard, P. Jensen, T. McCabe, J. E. O'Brien, R. D. Peacock, P. E. Kruger and T. Gunnlaugsson, Self-assembly of chiral luminescent lanthanide coordination bundles, *J. Am. Chem. Soc.*, 2007, **129**, 10986; (b) F. Stomeo, C. Lincheneau, J. P. Leonard, J. E. O'Brien, R. D. Peacock, C. P. McCoy and T. Gunnlaugsson, Metal directed synthesis of enantiomerically pure dimetallic lanthanide luminescent triple-stranded helicates, *J. Am. Chem. Soc.*, 2009, **131**, 9636.
 - 11 (a) O. Kotova, C. O'Reilly, S. T. Barwich, L. E. Mackenzie, A. D. Lynes, A. J. Savyasachi, M. Ruether, R. Pal, M. E. Möbius and T. Gunnlaugsson, Lanthanide luminescence from supramolecular hydrogels consisting of bio-conjugated picolinic-acid-based guanosine quadruplexes, *Chem*, 2022, **8**, 1395; (b) T. Gorai, J. I. Lovitt, D. Umadevi, G. McManus and T. Gunnlaugsson, Hierarchical supramolecular co-assembly formation employing multi-component light-harvesting charge transfer interactions giving rise to long-wavelength emitting luminescent microspheres, *Chem. Sci.*, 2022, **13**, 7805.
 - 12 (a) T. Tateishi, T. Kojima and S. Hiraoka, Chiral self-sorting process in the self-assembly of homochiral coordination cages from axially chiral ligands, *Commun. Chem.*, 2018, **1**, 18; (b) C. Gütz, R. Hovorka, G. Schnakenburg and A. Lützen, Homochiral supramolecular M₂L₄ cages by high-fidelity self-sorting of chiral ligands, *Chem. – Eur. J.*, 2013, **19**, 10890.
 - 13 (a) P. Stachelek, L. MacKenzie, D. Parker and R. Pal, Circularly polarised luminescence laser scanning confocal microscopy to study live cell chiral molecular interactions, *Nat. Commun.*, 2022, **13**, 553; (b) A. B. Aletti, S. Blasco, S. J. Aramballi, P. E. Kruger and T. Gunnlaugsson, Sulfate-Templated 2D Anion-Layered Supramolecular Self-Assemblies, *Chem*, 2019, **5**, 2617.
 - 14 (a) P. W. Zabierowski, O. Jeannin, T. Fix, J.-F. Guillemales, L. J. Charbonnière and A. M. Nonat, From mono- to polynuclear coordination complexes with a 2,2'-bipyrimidine-4,4'-dicarboxylate ligand, *Inorg. Chem.*, 2021, **60**, 8304; (b) S. Naseri, M. Mirzakhani, C. Besnard, L. Guénée, L. Briant, H. Nozary and C. Piguet, Preorganized polyaromatic soft terdentate hosts for the capture of [Ln(β-diketonate)₃] guests in solution, *Chem. – Eur. J.*, 2023, **29**, e202202727; (c) D. F. Caffrey, A. J. Savyasachi, K. Byrne, G. Tobin, B. D'agostino, W. Schmitt and T. Gunnlaugsson, Self-assembled bright luminescent hierarchical materials from a tripodal benzoate antenna and heptadentate Eu(III) and Tb(III) cyclen complexes, *Front. Chem. Sci. Eng.*, 2019, **13**, 171.
 - 15 O. Kotova, S. Comby, K. Pandurangan, F. Stomeo, J. E. O'Brien, M. Feeney, R. D. Peacock, C. P. McCoy and



- T. Gunnlaugsson, The effect of the linker size in C2-symmetrical chiral ligands on the self-assembly formation of luminescent triple-stranded di-metallic Eu(III) helicates in solution, *Dalton Trans.*, 2018, **47**, 12308.
- 16 (a) D. E. Barry, J. A. Kitchen, L. Mercs, R. D. Peacock, M. Albrecht and T. Gunnlaugsson, Chiral luminescent lanthanide complexes possessing strong (SmIII) circularly polarised luminescence (CPL), and their self-assembly into Langmuir-Blodgett films, *Dalton Trans.*, 2019, **48**, 11317; (b) O. Kotova, B. Twamley, J. O'Brien, R. D. Peacock, S. Blasco, J. A. Kitchen, M. Martinez-Calvo and T. Gunnlaugsson, The application of chiroptical spectroscopy (circular dichroism) in quantifying binding events in lanthanide directed synthesis of chiral luminescent self-assembly structures, *Chem. Sci.*, 2015, **6**, 457.
 - 17 (a) A. Galanti, O. Kotova, S. Blasco, C. J. Johnson, R. D. Peacock, S. Mills, J. J. Boland, M. Albrecht and T. Gunnlaugsson, Exploring the Effect of Structural Isomerism in Langmuir-Blodgett Films of Chiral Luminescent Eu(III) Self-Assemblies, *Chem. – Eur. J.*, 2016, **22**, 9709; (b) J. A. Kitchen, D. E. Barry, L. Mercs, M. Albrecht, R. D. Peacock and T. Gunnlaugsson, Circular polarized lanthanide luminescence from Langmuir-Blodgett films formed from optically active and amphiphilic Eu(III) based self-assembly complexes, *Angew. Chem., Int. Ed.*, 2012, **51**, 704; (c) C. Lincheneau, R. D. Peacock and T. Gunnlaugsson, Europium Directed Synthesis of Enantiomerically Pure Dimetallic Luminescent “Squeezed” Triple-Stranded Helicates; Solution Studies, *Chem. – Asian J.*, 2010, **5**, 500.
 - 18 (a) D. E. Barry, J. A. Kitchen, K. Pandurangan, A. J. Savyasachi, R. D. Peacock and T. Gunnlaugsson, Formation of enantiomerically pure luminescent triple-stranded dimetallic europium helicates and their corresponding hierarchical self-assembly formation in protic polar solutions, *Inorg. Chem.*, 2020, **59**, 2646; (b) J. A. Kitchen, D. E. Barry, L. Mercs, M. Albrecht, R. D. Peacock and T. Gunnlaugsson, Circularly polarized lanthanide luminescence from Langmuir-Blodgett films formed from optically active and amphiphilic Eu(III)-based self-assembly complexes, *Angew. Chem., Int. Ed.*, 2012, **51**, 704.
 - 19 O. Kotova, J. A. Kitchen, C. Lincheneau, R. D. Peacock and T. Gunnlaugsson, Probing the effects of ligand isomerism in chiral luminescent lanthanide supramolecular self-assemblies; an europium “Trinity Sliotar” study, *Chem. – Eur. J.*, 2013, **19**, 16181.
 - 20 (a) I. N. Hegarty, H. L. Dalton, A. F. Henwood, C. S. Hawes and T. Gunnlaugsson, Unexpected linkage isomerism in chiral tetranuclear bis-tridentate (1,2,3-triazol-4-yl)-picolinamide (tzpa) grids, *Chem. Commun.*, 2019, **55**, 9523; (b) D. E. Barry, C. S. Hawes, J. P. Byrne, B. la Cour Poulsen, M. Ruether, J. E. O'Brien and T. Gunnlaugsson, A folded [2×2] metallo-supramolecular grid from a bis-tridentate (1,2,3-triazol-4-yl)-picolinamide (tzpa) ligand, *Dalton Trans.*, 2017, **46**, 6464.
 - 21 I. N. Hegarty, H. L. Dalton, A. D. Lynes, B. Haffner, M. E. Möbius, C. S. Hawes and T. Gunnlaugsson, Balancing connectivity with function in silver(I) networks of pyridyl-triazole (tzpa) ligands results in the formation of a metallogel, *Dalton Trans.*, 2020, **49**, 7364.
 - 22 (a) A. F. Henwood, I. N. Hegarty, E. P. McCarney, J. I. Lovitt, S. Donohoe and T. Gunnlaugsson, Recent advances in the development of the btp motif: A versatile terdentate coordination ligand for applications in supramolecular self-assembly, cation and anion recognition chemistries, *Coord. Chem. Rev.*, 2021, **449**, 214206; (b) J. P. Byrne, J. A. Kitchen and T. Gunnlaugsson, The btp [2,6-bis(1,2,3-triazol-4-yl)pyridine] binding motif: a new versatile terdentate ligand for supramolecular and coordination chemistry, *Chem. Soc. Rev.*, 2014, **43**, 5302.
 - 23 (a) E. P. McCarney, J. I. Lovitt and T. Gunnlaugsson, Mechanically interlocked chiral self-templated [2]catenanes from 2,6-bis(1,2,3-triazol-4-yl)pyridine (btp) ligands, *Chem. – Eur. J.*, 2021, **27**, 12052; (b) C. O'Reilly, S. Blasco, B. Parekh, H. Collins, G. Cooke, T. Gunnlaugsson and J. P. Byrne, Ruthenium-centred btp glycoclusters as inhibitors for *Pseudomonas aeruginosa* biofilm formation, *RSC Adv.*, 2021, **11**, 16318; (c) J. P. Byrne, S. Blasco, A. B. Aletti, G. Hessman and T. Gunnlaugsson, Formation of self-templated 2,6-bis(1,2,3-triazol-4-yl)pyridine [2]catenanes by triazolyl hydrogen bonding: selective anion hosts for phosphate, *Angew. Chem., Int. Ed.*, 2016, **55**, 8938; (d) J. P. Byrne, M. Martínez-Calvo, R. D. Peacock and T. Gunnlaugsson, Chiroptical Probing of Lanthanide-Directed Self-Assembly Formation Using btp Ligands Formed in One-Pot Diazo-Transfer/Deprotection Click Reaction from Chiral Amines, *Chem. – Eur. J.*, 2016, **22**, 486.
 - 24 (a) W.-L. Chan, C. Xie, W.-S. Lo, J.-C. G. Bünzli, W.-K. Wong and K.-L. Wong, Lanthanide-tetrapyrrole complexes: synthesis, redox chemistry, photophysical properties, and photonic applications, *Chem. Soc. Rev.*, 2021, **50**, 12189; (b) S. V. Eliseeva and J.-C. G. Bünzli, Lanthanide luminescence for functional materials and bio-sciences, *Chem. Soc. Rev.*, 2010, **39**, 189.
 - 25 (a) D. Parker, J. D. Fradgley and K.-L. Wong, The design of responsive luminescent lanthanide probes and sensors, *Chem. Soc. Rev.*, 2021, **50**, 8193; (b) E. M. Surender, S. Comby, B. Cavanagh, O. Brennan, T. C. Lee and T. Gunnlaugsson, Two-photon luminescent bone Imaging using europium nanoagents, *Chem.*, 2016, **1**, 438.
 - 26 (a) J. P. Byrne, J. A. Kitchen, J. E. O'Brien, R. D. Peacock and T. Gunnlaugsson, Lanthanide Directed Self-Assembly of Highly Luminescent Supramolecular “Peptide” Bundles from α -Amino Acid Functionalized 2,6-Bis(1,2,3-triazol-4-yl)pyridine (btp) Ligands, *Inorg. Chem.*, 2015, **54**, 1426; (b) C. S. Bonnet, M. Devocelle and T. Gunnlaugsson, Structural studies in aqueous solution of new binuclear lanthanide luminescent peptide conjugates, *Chem. Commun.*, 2008, 4552.
 - 27 (a) S. J. Bradberry, G. Dee, O. Kotova, C. P. McCoy and T. Gunnlaugsson, Luminescent lanthanide (Eu(III)) cross-



- linked supramolecular metallo co-polymeric hydrogels: The effect of ligand symmetry, *Chem. Commun.*, 2019, **55**, 1754;
- (b) S. J. Bradberry, J. P. Byrne, C. P. McCoy and T. Gunnlaugsson, Lanthanide luminescent logic gate mimics in soft matter: [H⁺] and [F⁻] dual-input device in a polymer gel with potential for selective component release, *Chem. Commun.*, 2015, **51**, 16565.
- 28 (a) W. De, W. Horrocks and D. R. Sudnick, Lanthanide ion probes of structure in biology. Laser-Induced luminescence decay constants provide a direct measure of the number of metal-coordinated water molecules, *J. Am. Chem. Soc.*, 1979, **101**, 334; (b) T. J. Sørensen and S. Faulkner, Multimetallic lanthanide complexes: Using kinetic control to define complex multimetallic arrays, *Acc. Chem. Res.*, 2018, **51**, 2493.
- 29 A. Beeby, I. Clarkson, R. Dickins, S. Faulkner, D. Parker, L. Royle, A. S. de Sousa, J. A. G. Williams and M. Woods, Non-radiative deactivation of the excited states of europium, terbium and ytterbium complexes by proximate energy-matched OH, NH and CH oscillators: an improved luminescence method for establishing solution hydration states, *J. Chem. Soc., Perkin Trans. 2*, 1999, 493.
- 30 A.-S. Chauvin, F. Gumy, D. Imbert and J.-C. G. Bünzli, Europium and terbium tris(dipicolinates) as secondary standards for quantum yield determination, *Spectrosc. Lett.*, 2004, **37**, 517.

