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# Fluorane sensitive supercapacitive microcrystalline MoO<sub>3</sub>: dual application in energy storage and HF detection†

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Exploring materials and device paradigms for multifunctional electrochemical applications such as supercapacitors and sensing makes materials more suitable for real-life applications. In this study, microcrystalline MoO<sub>3</sub> powder has been synthesized using a simple sol–gel method, and its suitability for energy storage devices and HF sensing performance has been studied. The MoO<sub>3</sub> microcrystallites, well-characterized using electron microscopy, X-ray diffraction, and Raman spectroscopy, have been tested for HF sensitivity on a glassy carbon electrode as well as on a carbon cloth electrode. Similarly, a solid-state prototype asymmetric supercapacitor has been demonstrated that displays its charge storage capabilities. The specific capacitance of MoO<sub>3</sub> increases linearly with the increase of HF concentration. Additionally, the sensing performance of MoO<sub>3</sub> can be seen by monitoring changes in current passing through the electrode in the presence of HF. High stability with good repeatability was displayed. *In situ* Raman spectroscopy, recorded during the charging and discharging process, has been used to understand the charge storage mechanism. A high sensitivity of 6656 mF mM<sup>-1</sup> g<sup>-1</sup> with a low limit of detection of 1.2 μM was observed, which makes this material suitable for sensing as well as charge storage.

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## Introduction

Chemical waste from an industry that is illegally disposed of or improperly treated has massively contaminated the environment around such industrial facilities leading to several

hazards and consequent health issues.<sup>1,2</sup> Therefore, the need for sensors to detect contamination, *e.g.*, in flowing waters, has risen drastically.<sup>3–5</sup> Hydrofluoric acid, or HF, is a commonly used material in various industries related to pharmaceuticals, herbicides, glass manufacturing, ceramics and metals (as surface treatment agents), photovoltaics, and the food industry due to its use as a reducing agent.<sup>6</sup> Exposure to HF even in very low concentrations can cause severe health issues such as skin damage, eye irritation, visual disturbances, *etc.*<sup>7</sup> Due to all the above factors, detection of HF is very crucial.<sup>8</sup> In recent years, many sensors have been fabricated based on calorimetry,<sup>9</sup> chemiluminescence,<sup>10</sup> fluorescence,<sup>11</sup> chromatography,<sup>12</sup> colorimetry sensors,<sup>13</sup> *etc.* Typically sensors have many disadvantages, such as a difficult preparation process, poor sensitivity, an unacceptable limit of detection, low reproducibility, limited storage time, *etc.*<sup>14</sup> To overcome the above disadvantages, electrochemical sensing methods could prove to be a better alternative due to their advantages such as high repeatability, high sensitivity, a low limit of detection, real time monitoring, easy fabrication, *etc.*<sup>15</sup> Moreover, electrochemical sensing based on supercapacitive behaviour is expected to have further high sensitivity and a low limit of detection as compared to electrochemical sensing using the *I–V* curve, which reads current at the voltage for various concentrations.

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Electrochemical sensing using specific capacitance reads the area under the CV curve, a combination of current values for a potential range for various concentrations. Due to this property as mentioned above the area under the CV curve is more sensitive to material concentration as compared to  $I$ - $V$  based electrochemical sensing. In other words, electrochemical response from appropriate materials in terms of area under CV while under an HF environment can be used as the experimental observable to detect HF in trace amounts as has been studied here.

The performance of an electrochemical sensor is governed mainly by the redox reactions occurring at the working electrode. The working electrode materials can be either anodic or cathodic depending upon the need for the type of sensing material, *e.g.*,  $H_2O_2$  is an oxidizing agent, so it requires an anodic type material to be detected, whereas HF is a reducing agent, so it requires a cathodic material to be detected. Transition metal oxides (TMOs) such as NiO,<sup>16</sup>  $Co_3O_4$ ,<sup>17,18</sup>  $WO_3$ ,<sup>19</sup> ZnO,<sup>20</sup>  $MoO_3$ ,<sup>21</sup>  $TiO_2$ ,<sup>18</sup> *etc.* have been explored for electrochemical sensing,<sup>22</sup> electrochromic devices,<sup>23,24</sup> supercapacitors<sup>25–27</sup> and battery performance.<sup>28</sup> The supercapacitor can be either an electric double-layer supercapacitor<sup>29</sup> (non-faradaic or EDLC) or a pseudocapacitor<sup>30,31</sup> (faradaic or redox active). Materials with pseudocapacitive properties can also be used as sensors, depending upon their redox activity. Another metal oxide,  $MoO_3$ , with an orthorhombic phase, has been studied over the last two decades for potential energy storage devices<sup>32,33</sup> (like lithium capacitors). However, the applications of  $MoO_3$  in a device form are not yet established due to low specific capacitance and poor charging/discharging rates due to kinetic issues linked with poor electrical conductivity, structural stability during the charge–discharge process, and limited scalability in its preparation.<sup>34,35</sup> Additionally,  $MoO_3$  is a cathodic pseudocapacitive material and hence can be used to detect reducing agents<sup>36–38</sup> like HF. Looking at this combination, being a charge storage material as well as a possible HF detector, it will be interesting to investigate the electrochemical properties of suitably designed (micro-)crystalline  $MoO_3$  for bifunctional applications.

In this paper, an economic sol–gel technique was used to synthesize microcrystalline  $MoO_3$  to be used as an electrode material for charge storage and HF sensing applications. Prior to this above-mentioned application, this material was characterized by scanning electron microscopy (SEM), elemental diffraction X-ray spectroscopy (EDX), transmission electron microscopy (TEM), X-ray diffraction (XRD), and Raman spectroscopy. The  $MoO_3$  powder was deposited on a glassy carbon electrode (GCE) and carbon cloth (CC) for charge storage supercapacitive and sensing performance. When measured in the three-electrode cell, the  $MoO_3$  deposited on the GCE shows a high sensitivity and low detection limit (LOD), and better charge storage capabilities. Electrochemical impedance spectroscopy (EIS) was employed to understand the charge storage mechanism. The chronoamperometric response of  $MoO_3$  on CC shows high selectivity, fast response time, and high current retention. Electrochemical measurements of

different batches of  $MoO_3$  samples show high reproducibility and repeatability. Furthermore, a prototype solid-state supercapacitor has been fabricated onto which *in situ* Raman spectroscopy has also been carried out to understand the charge storage mechanism by recording Raman spectra during the charging and discharging processes.

## Experimental details

### Sample preparation

Sodium molybdate dihydrate ( $Na_2MoO_4 \cdot 2H_2O$ ), citric acid ( $C_6H_8O_7$ ), lithium perchlorate ( $LiClO_4$ ), hydrofluoric acid (HF), propylene carbonate (PC), and nitric acid ( $HNO_3$ ) were procured commercially and used as received from Alfa Aesar™/Sigma Aldrich™. Sodium molybdate dihydrate, nitric acid, and citric acid were used to synthesize molybdenum trioxide ( $MoO_3$ ) powder using a sol–gel synthesis method. The precursor solution was prepared by dissolving 0.01 M (50 mg) sodium molybdate in 20 mL deionized water (DI). Furthermore, 0.016 M of  $HNO_3$  (20  $\mu$ L) was added dropwise in the precursor solution under constant stirring for over 30 min. 0.1 M of citric acid as a surfactant was added to the solution and stirred for 1 h. The resulting solution was kept in an oven at 120 °C for 6 h. The resultant precipitate from the heating process was annealed at 450 °C for 3 h under an inert atmosphere. The powder after annealing was collected for further characterization without any purification.

### Device fabrication

The GCE was chosen as the working electrode to check the electrochemical performance, and CC ( $2 \times 1 \text{ cm}^2$ ) was chosen as the working electrode for electrochemical sensing experiments and symmetric supercapacitive device performance measurements of  $MoO_3$ . A homogeneous solution consisting of active material ( $MoO_3$ ), activated charcoal, and Nafion in a mass ratio of 80 : 10 : 10 was prepared in DI. Afterward, 5  $\mu$ L and 50  $\mu$ L of the above homogeneous solution were drop cast on GCE ( $MoO_3$ /GCE) and CC ( $1 \times 1 \text{ cm}^2$ ) ( $MoO_3$ /CC), respectively and dried for 24 h in a desiccator. To prepare the liquid electrolyte less asymmetric supercapacitor device ( $MoO_3$ //AC),  $MoO_3$ /CC was chosen as the current collector, and Whitman filter paper soaked in 0.5 M  $LiClO_4$  + 1 mM HF in propylene carbonate (PC) electrolyte was used as a separator. The wet separator was sandwiched between  $MoO_3$ /CC and wrapped by two glass slides to give stability. The edges of the device were covered by Teflon tape to ensure no exchange of oxygen ions from the atmosphere to the device.

### Sample characterization

A JEOL IT 500 field emission scanning electron microscope (SEM) was used for surface morphology study. The transmission electron microscope (TEM) images were taken with a FEI Tecnai G2 F20 TMP operated using an emission field gun at 200 kV to investigate the particle morphology. The dried  $MoO_3$  sample was collected on the surface of carbon covered copper



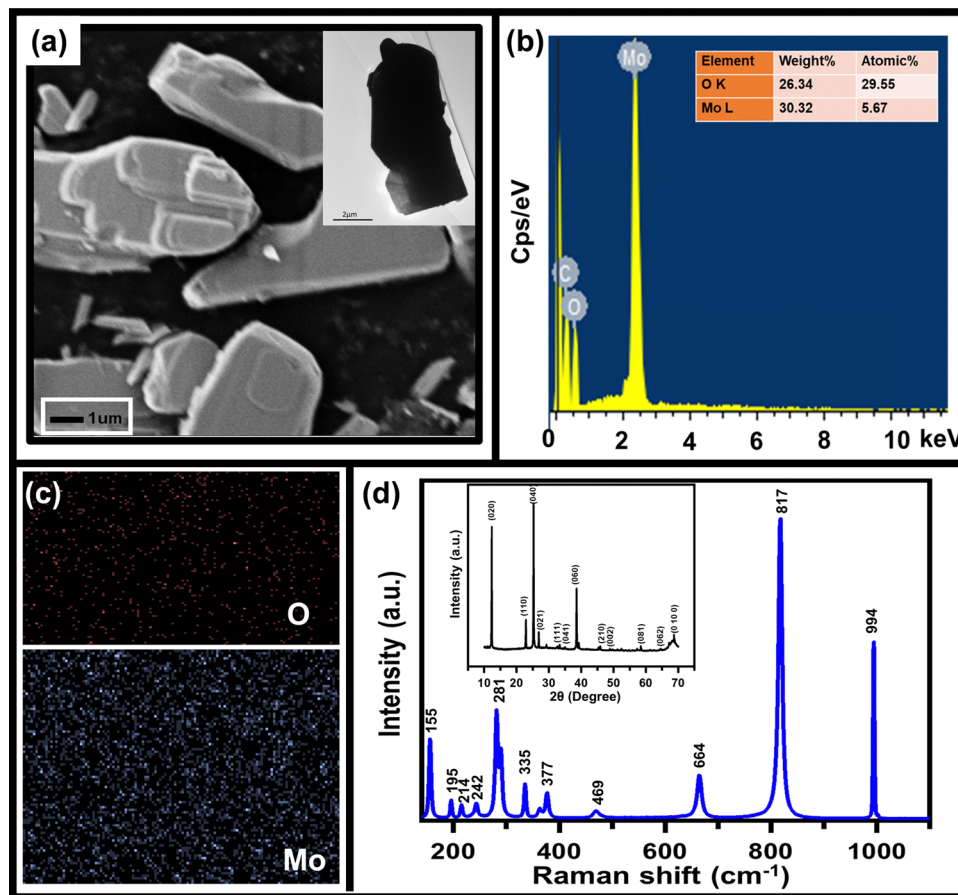


Fig. 1 Characterization of as-synthesized  $\text{MoO}_3$  powder as seen using (a) SEM micrograph along with TEM image (inset), (b) EDX spectrum for elemental confirmation (inset: elemental composition table), (c) elemental composition, and (d) Raman spectrum and XRD pattern (inset).

grids (300 mesh) from Quantifoil by wiping the grid over the powdered sample. Energy dispersive X-ray (EDX) was applied for elemental compositional analysis. A Horiba Jobin-Yvon Raman spectrometer with a 633 nm excitation laser was used to record the Raman spectrum. A Bruker D2-Phaser diffractometer was used for X-ray diffraction (XRD) using a  $\text{Cu K}\alpha$  wavelength of 1.54 Å. A Metrohm-Multi Autolab M204 potentiostat was used for cyclic voltammetry, electrochemical impedance spectroscopy, galvanostatic charge discharge, and chronoamperometric response measurements using the parameters as discussed later on.

## Results and discussion

As mentioned above, to achieve high electrochemical sensitivity, a porous structure is preferred.<sup>39</sup> Prior to being used as an electrochemical sensor and supercapacitor, the as-synthesized  $\text{MoO}_3$  electrode was characterized using different techniques (Fig. 1) to check its surface morphology, elemental composition, phase, and purity. The SEM image (Fig. 1a) indicates high surface porosity with irregularly shaped microcrystals of  $\text{MoO}_3$ , implying more diffusion of electrolyte ions for achieving high

electrochemical performance. Uniform deposition of the cuboidal structure of  $\text{MoO}_3$  powder of a few micrometres in size uniformly spread over the surface can be seen in the low magnification SEM image (Fig. S1, ESI<sup>†</sup>). The cuboidal geometry of  $\text{MoO}_3$  can also be seen in the TEM image (inset, Fig. 1a). To further confirm the elemental composition of  $\text{MoO}_3$  powder, EDX spectra were recorded (Fig. 1b). The EDX spectra confirm the presence of Mo and O atoms within the powder without any noticeable impurities. The elemental mapping of  $\text{MoO}_3$  (Fig. 1c) indicates the uniform atomic distribution of Mo and O atoms all over the surface. To further confirm the phase of  $\text{MoO}_3$ , the Raman spectrum was recorded (Fig. 1d), which shows peaks at 155, 195, 214, 242, 281, 335, 377, 469, 664, 817, and 994  $\text{cm}^{-1}$  indicating the orthorhombic phase of  $\text{MoO}_3$ .<sup>40,41</sup> The absence of other peaks in the Raman spectrum further indicates the high purity of the synthesized  $\text{MoO}_3$  powder. To further validate the crystal phase of the sample, XRD was carried out (inset, Fig. 1d), which shows diffraction peaks at  $2\theta = 12, 23, 25, 27, 29, 33, 35, 38, 46, 49, 58, 65,$  and  $69^\circ$  confirming the orthorhombic phase of  $\text{MoO}_3$  powder (JCPDS no 050508),<sup>42,43</sup> making it consistent with the Raman results. Overall, the morphological, spectroscopic, and diffraction techniques confirm the synthesis of the orthorhombic phase of



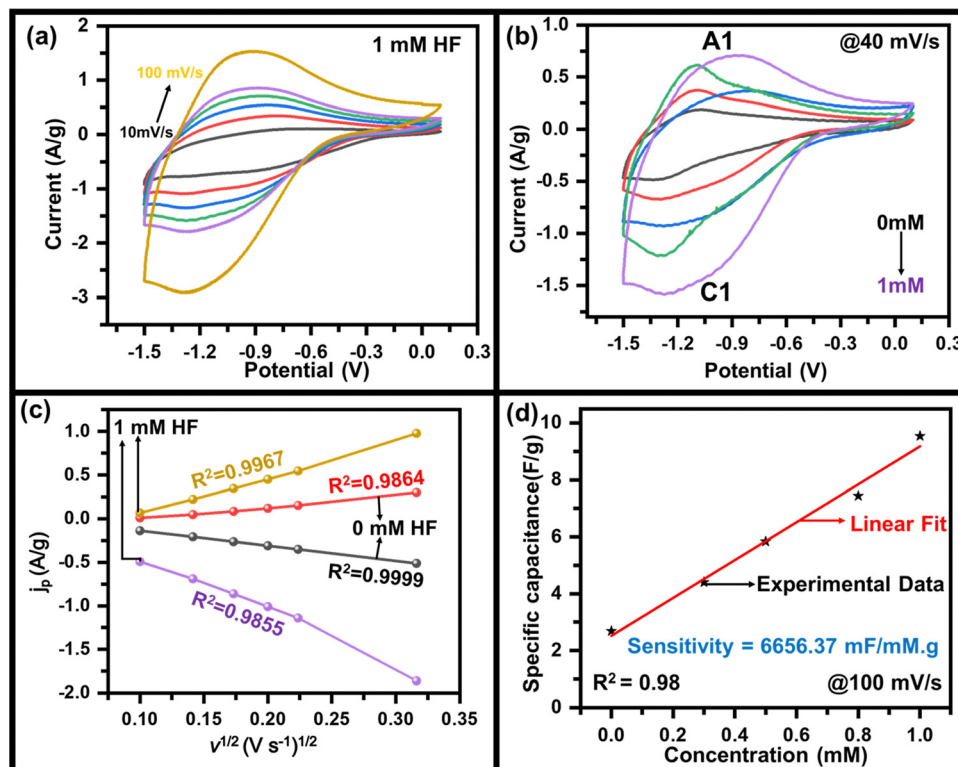
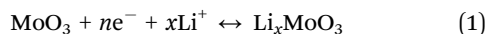


Fig. 2 CV graph of  $\text{MoO}_3$  at GCE with (a) 1 mM HF, (b) 0 mM, 0.2 mM, 0.5 mM, 0.8 mM, and 1.0 mM HF in 0.5 M  $\text{LiClO}_4$  in PC, (c) variation of current density ( $i_p$ ) vs.  $v^{1/2}$  graph with and without HF, and (d) variation of specific capacitance with concentration of HF including raw data and linear fit.

$\text{MoO}_3$  with porous nature as required for better electrochemical performances.

The sensing of hydrofluoric acid was carried out by three-electrode electrochemical measurements by adding a definite amount of HF to the electrolyte. The CV curves of  $\text{MoO}_3/\text{GCE}$  (Fig. 2a) measured in 0.5 M  $\text{LiClO}_4$  with 1 mM HF in PC show a typical CV variation with varying scan rates ( $10 \text{ mV s}^{-1}$  to  $100 \text{ mV s}^{-1}$ ). The increase in the current density with increasing scan rate indicates the faster diffusion of electrolyte ions towards the surface at high scan rates.

The reaction governing the  $\text{Li}^+$  insertion due to charging and  $\text{Li}^+$  extraction during the discharge process is given by following the redox reaction eqn (1):



During the cathodic scan,  $\text{MoO}_3$  is reduced to  $\text{Li}_x\text{MoO}_3$ , whereas during the anodic scan, the  $\text{Li}_x\text{MoO}_3$  is converted back to  $\text{MoO}_3$ . It is crucial here to mention that the CV curve shape depends on the concentration of HF in the electrolyte, as can be seen in Fig. S2 (ESI<sup>†</sup>), which shows CV curves of  $\text{MoO}_3/\text{GCE}$  at various scan rates with different concentrations of HF (0 mM to 0.8 mM). The effect of HF concentration can be seen by comparing the CV graphs of  $\text{MoO}_3/\text{GCE}$  at various concentrations of HF (0 mM to 1 mM HF in the electrolyte) recorded at a given scan rate of  $40 \text{ mV s}^{-1}$  (Fig. 2b). The anodic and cathodic peak currents increase proportionally to the HF concentration. The cathodic peak (marked as C1) current increases with concentration and shows a linear change with HF

concentration, likely because HF is a reducing agent and can accept electrons from  $\text{MoO}_3$ . The CV curve also reveals that the oxidation of HF is taking place in a quasi-reversible regime for the anodic peak (marked A1) current, whereas for the cathodic peak (marked C1) current it is reversible. To further quantify the reduction of HF, peak current density ( $i_p$ ) as a function of scan rate ( $v^{1/2}$ ) was plotted as extracted from the CV curve for 0 mM and 1 mM HF (Fig. 2c) using the Dunn method<sup>44</sup> using eqn (2) given below.

$$\frac{i_p}{v^{1/2}} = k_1 v^{1/2} + k_2, \quad (2)$$

where  $k_2$  and  $k_1 v^{1/2}$  represent the current response for diffusion-controlled contributions, and surface capacitance, respectively. The linear dependence (Fig. 2c) of peak current density with  $v^{1/2}$  confirms the diffusion-controlled process. The  $k_2$  values are obtained from the intercept of the ( $i_p/v^{1/2}$ ) vs. ( $v^{1/2}$ ) graph plotted with and without HF and are found to be 0.038 and  $0.21 \text{ A g}^{-1} \text{ V s}^{-1/2}$ , respectively. In addition to HF, the increase in  $k_2$  value indicates the dominance of diffusion-controlled contribution in the presence of HF, as indicated by the CV curve of  $\text{MoO}_3$  at different concentrations (Fig. 2b).

The effect of HF on the specific capacitance was also studied to investigate it deeper, where the specific capacitance was estimated using the method mentioned in the ESI<sup>†</sup>. The variation of the specific capacitance as a function of scan rate at various concentrations of HF of  $\text{MoO}_3/\text{GCE}$  calculated from the CV curve was observed (Fig. S3, ESI<sup>†</sup>). A maximum specific



Table 1 Comparison of the HF electrochemical sensor with other reported sensors

S. no	Material	Sensing material	Sensitivity	LOD ( $\mu\text{M}$ )	Ref.
1	$\text{Bi}_2\text{Te}_3$	$\text{H}_2\text{O}_2$	$4903 \mu\text{A mM}^{-1} \text{cm}^{-2}$	0.016	15
2	$\text{Co}_3\text{O}_4$ $\text{NiCo}_2\text{O}_4$ NSs	$\text{H}_2\text{O}_2$	$303.42 \mu\text{A mM}^{-1} \text{cm}^{-2}$	0.596	17
3	$\text{Co}_3\text{O}_4$ $\text{NiCo}_2\text{O}_4$ NSs	Glucose	$1463.13 \mu\text{A mM}^{-1} \text{cm}^{-2}$	0.596	17
4	$\text{CoO}_x$ NPs/cholesterol oxidase	$\text{H}_2\text{O}_2$	$43.5 \mu\text{A mM}^{-1} \text{cm}^{-2}$	4.2	45
5	$\text{Cu}_2\text{O}$ on ALOOH/reduced graphene oxide	Glucose	$155.1 \mu\text{A mM}^{-1} \text{cm}^{-2}$	2.6	46
6	$\text{Fe}_3\text{O}_4$ @Au/ $\text{MnO}_2$	Glucose	$2520 \mu\text{A mM}^{-1} \text{cm}^{-2}$	13.2	47
7	Nanoneedle- $\text{NiCo}_2\text{O}_4$	Glucose	$3000 \mu\text{A mM}^{-1} \text{cm}^{-2}$	2	48
8	$\text{MoS}_2/\text{S-Ti}_3\text{C}_2/\text{LGE}$	Aristolochic acid	$69.955 \mu\text{A mM}^{-1} \text{cm}^{-2}$	—	49
9	$\text{MoS}_2/\text{S-Ti}_3\text{C}_2/\text{LGE}$	Roxarsone	$57 \mu\text{A mM}^{-1} \text{cm}^{-2}$	—	49
10	$\text{MoO}_3$	Fluorane	$6656 \text{mF mM}^{-1} \text{g}^{-1}$	1.5	This work

capacitance of  $210 \text{ F g}^{-1}$ ,  $54 \text{ F g}^{-1}$  at  $10 \text{ mV s}^{-1}$  was calculated which decreased gradually to  $95 \text{ F g}^{-1}$ ,  $26 \text{ F g}^{-1}$  at  $100 \text{ mV s}^{-1}$  for 1 mM, and 0 mM HF concentration respectively (Fig. S3, ESI<sup>†</sup>). This also indicates that the capacitance values are sensitive to the presence of HF and are affected by its concentration. To further evaluate the charging–discharging properties of  $\text{MoO}_3$ , GCD tests have been performed on  $\text{MoO}_3/\text{GCE}$  in 0.5 M  $\text{LiClO}_4$  with 1 mM HF at various current densities (Fig. S4, ESI<sup>†</sup>), which also shows the variation of specific capacitance with current density (inset, Fig. S4, ESI<sup>†</sup>). A maximum specific capacitance of  $\sim 110 \text{ F g}^{-1}$  at  $2 \text{ A g}^{-1}$  was observed.

The electrochemical measurements show a variation in the specific capacitance as a function of HF concentration. A linear variation (Fig. 2d) is seen throughout the measured range and was used to calculate an HF sensitivity of  $\sim 6656 \text{ mF mM}^{-1} \text{g}^{-1}$ ,

which is very encouraging as compared to the values reported in the range of thousands.<sup>45–47</sup> In order to cross-check, the sensitivity has also been estimated by using the redox peak current as the observable quantity. A variation in peak current density with HF concentration (Fig. S5, ESI<sup>†</sup>) was used for this purpose, yielding a sensitivity of  $\sim 689 \mu\text{A mM}^{-1} \text{cm}^{-2}$ . In addition to the sensitivity, the limit of detection (LOD) is also estimated from the variation of specific capacitance (or current density) with HF concentration taking the signal to noise ratio of  $S/N = 3$ . The value of  $1.5 \mu\text{M}$  (or  $1.2 \mu\text{M}$ ) was the LOD calculated from variation of specific capacitance (or current density) with HF concentration implying a very low LOD. The almost same LOD from experiments validates the above capacitance-based experiments. The LOD and sensitivity of the fabricated HF sensor have been compared with other

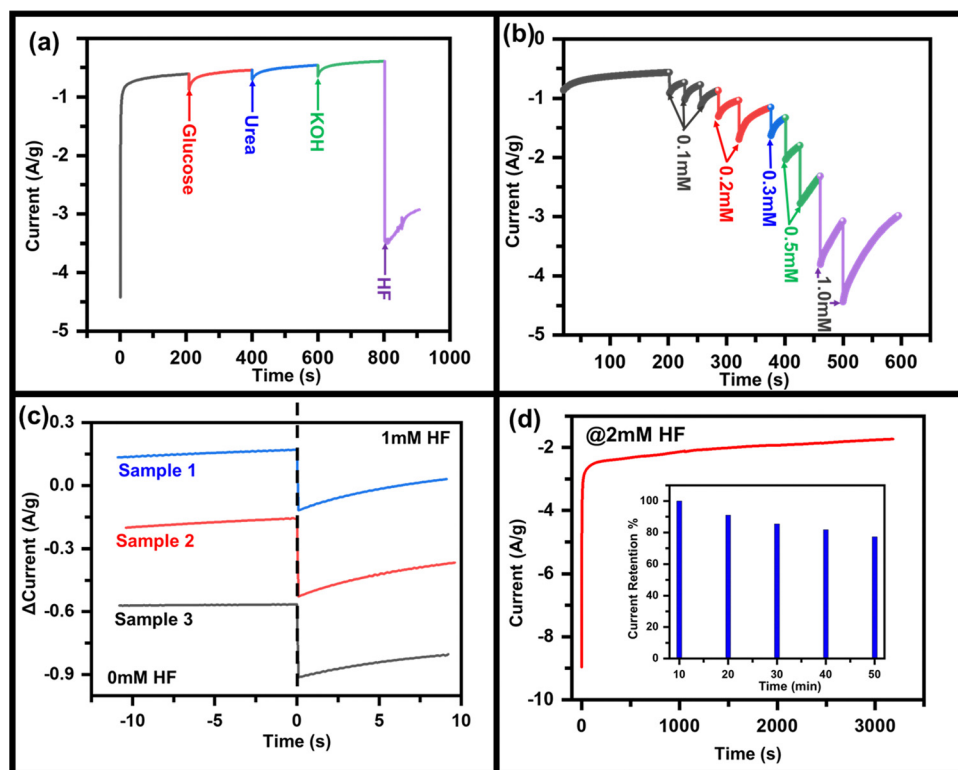


Fig. 3 Chronoamperometric response of  $\text{MoO}_3/\text{CC}$  (a) upon the addition of 1 mM HF, glucose, urea, and KOH, (b) upon the addition of different concentrations of HF (0.1, 0.2, 0.3, 0.5, and 1.0 mM). (c) Response time of the sensor for three different batches of the sample under 1 mM HF, and (d) long-term stability (inset: current retention versus time graph) test of the  $\text{MoO}_3/\text{CC}$  electrode.



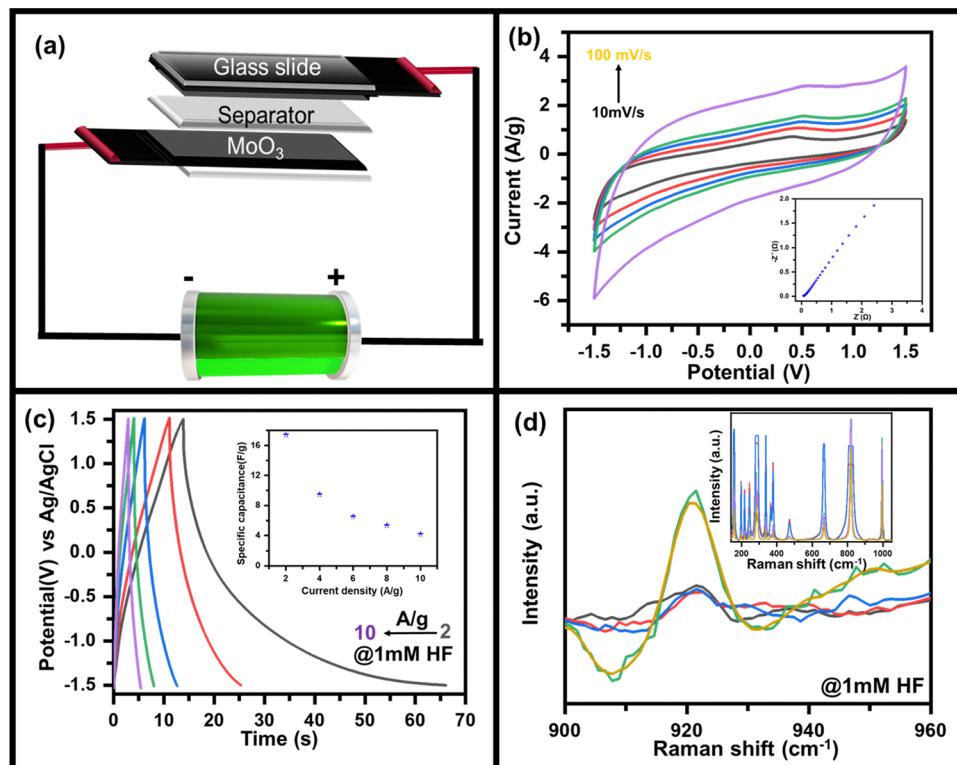


Fig. 4 (a) Schematic diagram of as-fabricated  $\text{MoO}_3//\text{AC}$  supercapacitor device, (b) CV curves measured at scan rates of 10 to  $100 \text{ mV s}^{-1}$  along with the corresponding EIS plot (ESI<sup>†</sup>), (c) GCD plots at  $2 \text{ A g}^{-1}$  to  $10 \text{ A g}^{-1}$  along with variation of specific capacitance as a function of current density (inset) and (d) *in situ* Raman spectra of asymmetric supercapacitor device during the charging and discharging process.

sensors in Table 1, indicating the superior performance of  $\text{MoO}_3/\text{GCE}$  as the HF sensor. For practical applications of a sensor, the material used in sensing must not degrade during the course of sensing so that the electrode can be used multiple times. To ensure the quality of pre- and post-sensing samples, Raman scattering was used. Fig. S6 (ESI<sup>†</sup>) indicates no change in the spectra after electrochemical measurements with (1.0 mM and 0.5 mM) and without HF in the electrolyte, confirming that the material is stable, and returns to its initial state after sensing is done.

From the above discussions, it is clear that the materials have characteristics for HF detection, but for practical applications, similar properties must be replicated on substrates that can be handled easily. To investigate the HF sensing property,  $\text{MoO}_3$  was drop casted on the CC electrode (say  $\text{MoO}_3/\text{CC}$ ) and further measurements were done. Fig. S7 (ESI<sup>†</sup>) represents the CV curve of  $\text{MoO}_3/\text{CC}$  at a  $50 \text{ mV s}^{-1}$  scan rate without HF. The CV test at  $50 \text{ mV s}^{-1}$  was performed on three identical batches of  $\text{MoO}_3/\text{CC}$  samples (Fig. S8, ESI<sup>†</sup>). No significant change in the anodic/cathodic peak current was observed, implying the  $\text{MoO}_3/\text{CC}$  has good repeatability. For an excellent electrochemical sensor, its performance should not interfere with other chemical compounds such as glucose, urea, and KOH. To show the anti-interference property of this material to the sensor, in a continuously stirred electrolytic solution, 1 mM of glucose, urea, KOH, and HF were added one after another (Fig. 3a), and the corresponding change in current was observed. With

the addition of materials other than HF, no significant change in the cathodic peak was observed. On the other hand, in addition to HF, cathodic peak current increases significantly, which also depends on the concentration of HF. Fig. 3b shows the chronoamperometry graph of  $\text{MoO}_3/\text{CC}$  at  $-1.3 \text{ V}$  with various HF concentrations. A significant change in the cathodic current was observed with the continuous addition of different HF concentrations. Fig. 3c indicates a fast response time (for three different samples to check reproducibility) of the sensor,  $\sim 2.5 \text{ s}$ , which is generally due to mixing HF droplets in the existing electrolyte solution. Fig. 3d indicates the long-term sensing stability of the  $\text{MoO}_3/\text{CC}$  electrode under 2 mM HF at room temperature, implying that the catalyst is stable even after a continuous performance of 50 min. The sensor retains  $\sim 80\%$  of its maximum current after continuous working of 50 min, even at a high concentration of HF (2 mM), indicating high current stability.

As discussed above, the charge storage capability of  $\text{MoO}_3$  has increased after HF addition in the electrolyte. Hence the suitability as the better charge storage device is tested by making an asymmetric supercapacitor device  $\text{MoO}_3//\text{AC}$  fabricated using the recipe discussed in the Experimental section as shown in the schematic (Fig. 4a). Since uniform deposition of  $\text{MoO}_3$  on CC is essential for better electrochemical performance, it was ensured prior to using it in the device, as seen in the SEM image (Fig. S9, ESI<sup>†</sup>). In addition, enough pores could be seen in  $\text{MoO}_3/\text{CC}$  to diffuse electrolyte ions into the substrate material interface. The CV curve of the device in 0.5 M  $\text{LiClO}_4$  with 1 mM HF shows a quasi-rectangular shape





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