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Au-decorated Sb₂Se₃ photocathodes for solar-driven CO₂ reduction†

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Photoelectrodes with FTO/Au/Sb₂Se₃/TiO₂/Au architecture were studied in photoelectrochemical CO₂ reduction reaction (PEC CO₂RR). The preparation is based on a simple spin coating technique, where nanorod-like structures were obtained for Sb₂Se₃, as confirmed by SEM images. A thin conformal layer of TiO₂ was coated on the Sb₂Se₃ nanorods *via* ALD, which acted as both an electron transfer layer and a protective coating. Au nanoparticles were deposited as co-catalysts *via* photo-assisted electrodeposition at different applied potentials to control their growth and morphology. The use of such architectures has not been explored in CO₂RR yet. The photoelectrochemical performance for CO₂RR was investigated with different Au catalyst loadings. A photocurrent density of $\sim 7.5 \text{ mA cm}^{-2}$ at -0.57 V vs. RHE for syngas generation was achieved, with an average Faradaic efficiency of $25 \pm 6\%$ for CO and $63 \pm 12\%$ for H₂. The presented results point toward the use of Sb₂Se₃-based photoelectrodes in solar CO₂ conversion applications.

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Broader context

Direct photoelectrochemical (PEC) CO₂ reduction offers an alternative pathway for producing chemicals by directly integrating the light absorption and CO₂ conversion in one device. Most devices use In-based and Ga-based semiconductors, which might be a potential hurdle to technological development because of the scarcity and expensiveness of these materials. In direct PEC water splitting, Sb₂Se₃ has gained attention due to its interesting properties and the simplicity of its fabrication *via* spin coating. We have shown that an FTO/Au/Sb₂Se₃/TiO₂ architecture can be utilized for direct photoelectrochemical CO₂ reduction using Au nanoparticles as cocatalyst. We found that the final architecture displayed a three times increase in current density and an earlier onset potential as compared to Au foil in dark CO₂ electroreduction. The product distribution showed a syngas mixture of CO and H₂. The results obtained suggest a favorable research direction towards exploring existing architectures for direct PEC CO₂ conversion studies.

Introduction

Photoelectrochemical (PEC) hydrogen evolution (HER) and CO₂ reduction (CO₂RR) are promising routes for the production of alternative fuels (*e.g.*, hydrogen, methanol, methane) to replace fossil fuels. Simultaneously, the utilization of solar fuels allows society to move a step closer to the decarbonization of various industrial processes and the energy sector. Direct PEC fuel generation, in theory, offers lower overall material cost compared to systems based on electrolyzers coupled to photovoltaic panels (PV-EC).¹ In the former case, the integration of the light

absorber and the particular redox reaction is achieved in a single system, making the use of wires and AC/DC converters unnecessary.¹ As there is direct contact between the light absorber and the electrolyte in PEC configuration, thermal effects can also play a role in determining the rate of electrochemical reactions.² The technological readiness level of direct PEC fuel generation is still low^{3,4} because: (i) the direct light absorber/electrolyte interface makes these systems prone to corrosion, and necessitates the use of protection layers, (ii) up to this point there is no “wonder material” that can integrate all required functionalities (*e.g.*, efficient light absorption, carrier transport, product generation) of the photoelectrode.¹

Sb₂Se₃ has gained significant attention in the PEC community as a photoabsorber especially in HER related applications because of its favorable band positions.^{5–9} It is a p-type semiconductor, which has a bandgap between 1.1 and 1.3 eV,¹⁰ similar to that of Si. It has only a single stable phase, which

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eliminates the possibility of secondary phase formation.^{11,12} Sb_2Se_3 can be synthesized through simple solution-based methods (e.g., spin coating).^{5,13} Aside from solution processing, other common synthesis methods of this material are vapor transport deposition,¹⁴ close-space sublimation,¹⁵ thermal evaporation,¹⁶ sputtering,¹⁷ electrodeposition,^{18,19} and selenization of Sb thin films.²⁰ Solution processing offers advantage in its simplicity and high throughput in preparing large area films. Nanostructured photoabsorbers provide better light absorption than planar films, because of lower reflection losses.²¹ A simple spin coating technique can be used to combine these two merits to produce high quality Sb_2Se_3 for PEC applications.

TiO_2 has been widely explored as a protective coating,^{20–23} because of its stability in a wide pH range (as supported by its Pourbaix diagram).²⁴ Additionally, it also functions as an electron transport layer that promotes the formation of a p–n junction²² when combined with a p-type semiconductor such as Sb_2Se_3 . This dual functionality of TiO_2 makes it an ideal material to pair with Sb_2Se_3 in fabricating a photocathode. However, making a conformal coating of TiO_2 especially at the nanoscale is challenging. Atomic layer deposition (ALD) can deliver this specific requirement, without compromising the nano-morphology of the Sb_2Se_3 . In spite of this conformality, the drawback of this specific type of TiO_2 is its lack of stability during long-term measurements.²⁵ Electron accumulation at the surface of the thin TiO_2 layer was shown to trigger the dissolution of the protective TiO_2 layer.⁹ Strategies to mitigate this issue were proposed such as adding an additional electron transport layer (e.g., fullerene⁹) or a hydrogel layer²⁶ that can kinetically enhance the electron transport to the catalyst.

Recent advances in PEC CO_2RR have shown that different products are possible depending on the cocatalyst used.^{27–29} The most commonly used cocatalysts are Au,^{30–36} Cu,^{37–39} Ag,^{39–41} and different molecular catalysts.^{42–45} Aside from choosing the appropriate cocatalyst, it was reported that tuning the interface exposed to the electrolyte can also have an impact by stabilizing the reaction intermediate (i.e., $^*\text{COOH}$ from CO_2), which can lead to lower overpotential and higher Faradaic efficiency.⁴⁶ This beneficial interface effect was shown experimentally using a $\text{n}^+\text{–p}$ Si/GaN substrate and was also supported by DFT calculations.^{31,47} Compared to only using Pt, a mixed Pt– TiO_2 interface showed an increase from 2% to 32% in the Faradaic efficiency for CO production (FE_{CO}) under 8 sun.⁴⁷ This interface engineering approach is further supported when Au– TiO_2 is used, where the FE_{CO} increased from $\sim 32\%$ (pure Au interface) to $\sim 55\%$ (mixed Au– TiO_2 interface) under 1 sun. A similar Au– TiO_2 interface was investigated but with InP nanopillar array as the substrate and the results showed a FE_{CO} of 84%.³⁰ In spite of this high selectivity towards CO, GaN-based and InP-based photoelectrodes are not ideal for commercial applications because of the high cost and scarcity of Ga and In, unlike the relative abundance and cost-effectiveness of Sb or Se.^{11,48} Nevertheless, these studies have demonstrated that the TiO_2 /metal interface is beneficial in PEC CO_2RR , particularly the TiO_2 /Au interface.

In this work, we prepared TiO_2 -protected Sb_2Se_3 nanorod photoelectrodes decorated with Au cocatalyst for direct PEC

CO_2RR . The preparation of the Sb_2Se_3 nanorod layers was performed by a simple spin coating technique followed by a conformal coating of TiO_2 via ALD. Subsequently, we carried out the deposition of Au cocatalyst on the electrode surfaces by PEC deposition. Finally, the PEC properties of the hybrid electrode assemblies were studied in CO_2RR .

Experimental

Chemicals used

Fluorine-doped tin oxide-coated glass slides (FTO) (2.2 mm thickness, surface resistivity $\sim 7 \Omega \text{ sq}^{-1}$) were purchased from Sigma Aldrich. All aqueous solutions were prepared from ultrapure water (DI water written as H_2O ; Milli-Q, 18.2 M Ω cm). For the ink preparation, SbCl_3 (99.99%, Alfa Aesar), Se powder (99.5%, Sigma Aldrich), 2-methoxyethanol (anhydrous, 99.8%, Sigma Aldrich), thioglycolic acid (98%, Sigma Aldrich), and ethanolamine (99.5%, Sigma Aldrich) were used. For the cocatalyst deposition, chloroplatinic acid hexahydrate (ACS reagent, $\geq 37.50\%$ Pt basis, Sigma Aldrich) and potassium tetrachloroaurate(III) (99.995% trace metals basis, Sigma Aldrich) were used. Anhydrous methanol (max. 0.005% H_2O ; $\geq 99.8\%$, VWR), potassium bicarbonate (reagent grade, VWR), and sulfuric acid (ACS reagent, VWR) were used. Tetrakis-(dimethylamino)titanium(IV) (99% TDMAT (99.99%–Ti)) was used as the ALD precursor and purchased from Strem Chemicals, Inc.

Cleaning of the substrates

FTO slides were cut into 3 cm $\times$ 3 cm pieces and were cleaned sequentially in acetone and isopropanol for 10 minutes each in an ultrasonic bath. A 60 nm layer of gold was deposited via thermal evaporation. After which, the FTO/Au substrates were soaked in piranha solution (3:1 v/v H_2SO_4 : H_2O_2) to remove organic residues from the layer surfaces. Additional ultrasonic cleaning was performed in acetone and isopropanol for 10 minutes each. The substrates were subjected to plasma cleaning for 10 minutes immediately before use.

Preparation of spin coating solution

The layer preparation was performed in a N_2 filled glovebox (<1 ppm H_2O , 1–5 ppm O_2). 0.258 g of SbCl_3 was dissolved in 12 mL of 2-methoxyethanol. Then, 0.384 g of Se powder was dissolved in a solvent mixture of thioglycolic acid (TGA) and ethanolamine. The ratio of this solvent mixture dictates the morphology of the Sb_2Se_3 nanorods. A volume ratio of 0.45 mL of TGA and 7.45 mL of ethanolamine was used. Ethanolamine was first added to the Se powder, followed by the dropwise addition of TGA (exothermic reaction) under vigorous stirring. Then, the SbCl_3 precursor solution was carefully added dropwise to the Se precursor solution (similarly exothermic reaction). The mixture was stirred and heated at 80 $^\circ\text{C}$ for 6–10 hours prior to use.

Before spin coating, the SbCl_3 –Se solution was filtered using a syringe with a PTFE membrane filter (0.2 μm pore size).



The spin coating was performed with a rotation speed of 2500 rpm for 30 seconds using a static spin coating technique. After spin coating, the substrate was immediately placed on a hotplate preheated at 180 °C for 3 minutes, and was then treated at 300 °C for 3 minutes. To ensure complete coverage of the substrate surface, this spin coating and heating was performed for 10 cycles. Sufficient time was ensured to allow the substrates to cool down before the subsequent spin coating cycle. The first 2 layers were deposited using 120 µL of the precursor mixture, while for the remaining layers, only 100 µL of precursor was used. After 10 spin coating and heating cycles, the substrates were heated at 350 °C for 20 minutes inside the glovebox, followed by a final heat treatment at 200 °C for 30 minutes outside the glovebox in a muffle furnace to burn off the residual organics.⁹ The substrates were cut into 3 pieces each with a dimension of 3 cm × 1 cm. Immediately after, the substrates were loaded inside the ALD chamber for the TiO₂ deposition.

ALD TiO₂ deposition

The instrument used was from Beneq with a model number of TFS 200. TDMAT was used as a precursor in a hot source container and H₂O was used as the co-reactant. The TDMAT hot source container was heated to 40 °C to have enough vapor pressure, but also to avoid the thermal degradation of TDMAT. The ALD chamber itself was heated at 140 °C during the whole duration of the deposition. A 0.3 s pulse of the TDMAT was used followed by a 10 s N₂ purge and then a 0.2 s pulse of the H₂O also followed by a 10 s N₂ purge. A total of 1000 cycles were used, which corresponds to around ~60 nm of TiO₂ thickness.

Catalyst deposition

For the benchmark PEC HER, Pt nanoparticles were obtained *via* photodeposition. For the target CO₂ reduction reaction, Au nanoparticles were grafted on the surface *via* PEC deposition. The electrolyte for both the Pt and Au solutions was 0.25 mM H₂SO₄ in 5% v/v methanol-containing aqueous solution, adapting a procedure for photodeposition of metallic Pt on TiO₂.⁴⁹ For the Pt deposition, 5 mM H₂PtCl₆ was used, while for the Au deposition, a 0.5 mM KAuCl₄ was employed. The counter electrode used for Pt deposition was a Pt wire while for Au deposition, a Au wire was used. Pt was deposited under 1 sun illumination (AM 1.5G, 100 mW cm⁻²) with a UV filter for a total duration of 60 minutes under continuous stirring. For the PEC deposition of Au, three different applied potentials were studied in conjunction with the 1 sun illumination equipped with a UV filter. To control the amount of deposited Au particles, the amount of charge passed during deposition was monitored and controlled.

PEC CO₂RR measurements

A solar simulator was used from Sciencetech SL-50A-WS, (AM 1.5G, 100 mW cm⁻²). All electrochemical experiments were performed using a Metrohm Autolab PGSTAT302 type potentiostat/galvanostat in a standard three-electrode setup at room temperature (25 °C). All measurements were performed for at

least 3 times. A custom-made cell holder in an H-cell configuration with a quartz window connected to a gas chromatograph was employed for the PEC CO₂RR experiments (Fig. S1, ESI†). The cathode and anode compartments of the H-cell were separated by a Nafion membrane. The illuminated area of the photoelectrodes was 0.28 cm². Stirring was also performed to help the mass transport of the CO₂ gas. The counter electrode was a rectangular piece of glassy carbon, and the reference electrode was Ag/AgCl (3 M NaCl). CO₂-saturated solution of 0.5 M KHCO₃ with a measured pH of 7.44 was employed as the electrolyte. CO₂ was continuously bubbled throughout the experiments. Linear sweep voltammograms were recorded at a sweep rate of 5 mV s⁻¹ while the light was interrupted with 10 second intervals. The Ag/AgCl reference scale was converted to the RHE scale knowing the pH of the solution with this equation:

$$E_{\text{RHE}} = E_{\text{Ag/AgCl}} + 0.059\text{pH} + 0.1976 \text{ V} \quad (1)$$

Product detection

The gas products were detected using Shimadzu GC-2010 Plus gas chromatograph equipped with a Barrier Discharge Ionization Detector (BID). A Restek ShinCarbon ST column was employed for the separation with 6.0 grade helium carrier gas. An automatized 6-port valve was used to take samples in specified time. The flow rate used was 20 mL min⁻¹.

Materials characterization

Diffuse reflectance spectra (DRS) were recorded with a Shimadzu UV-3600 Plus spectrophotometer between 400 and 1400 nm. The optical bandgap value of Sb₂Se₃ was obtained from Tauc analysis of the DRS data. Raman measurements were carried out using a SENTERRA II Compact Raman microscope, at a 532 nm laser excitation wavelength with a laser power of 2.5 mW. Top-down scanning electron microscopic (SEM) images were recorded with a Thermo Fisher Scientific Apreo C instrument equipped with Everhart-Thornley detector (ETD). The microscope was operated at 25 pA current and 10 kV acceleration voltage. Cross-section SEM images were recorded with a Thermo Scientific Scios 2 SEM-FIB instrument, using a Ga ion beam. X-ray diffraction (XRD) for a sample deposited on glass substrate was performed with a Rigaku Mini Flex II desktop diffractometer equipped with a Cu Kα (λ = 1.5118 Å) X-ray source. The diffraction patterns were recorded in the 10–60° 2θ range, with 0.2° min⁻¹ scanning speed. Additional XRD measurements were performed with a Rigaku SmartLab diffractometer for the samples deposited on FTO/Au. Diffraction patterns were recorded using Cu Kα irradiation (λ = 1.54 Å). Scans were performed in theta–2θ scanning geometry and the X-ray tube was operated at 200 mA and 45 kV. Diffraction patterns were measured between 2θ = 10 and 60° degrees. To minimize reflections from the FTO substrate and the gold layer, the sample was measured with parallel beam (PB) in grazing-incidence configuration. We chose 0.4° as the angle of incidence (ω), which allows the profiling of the Sb₂Se₃ layer avoiding signal from the FTO glass substrate. X-ray photoelectron



spectroscopy (XPS) measurements were performed with a SPECS instrument equipped with a PHOIBOS 150 MCD 9 hemispherical analyzer. To acquire survey scans, 40 eV pass energy was employed, while for high resolution scans, 20 eV was used. For spectra acquisitions, Al K α radiation ($h\nu = 1486.6$ eV) was used as an excitation source and operated at 150 W power. Ultraviolet photoelectron spectroscopy (UPS) was performed with He(I) excitation (21.22 eV) with 10 V of external bias applied to the samples. The recorded UPS spectra were corrected for additional He(I) lines. For XPS and UPS spectrum evaluation, CasaXPS commercial software package was used.⁵⁰

Internal-photon-to-electron-conversion efficiency measurements

The PEC activity of a FTO/Au/Sb₂Se₃/TiO₂/Au sample was evaluated by internal-photon-to-electron-conversion efficiency (IPCE) measurements as well (0.5 M KHCO₃, continuous stirring, CO₂ saturation and bubbling). IPCE experiments were carried out with a Newport Quantum Efficiency Measurement System (QEPVSI-B) in the same H-type, three-electrode electrochemical cell. The wavelength range lied between 1100–400 nm ($\Delta\lambda = 50$ nm step size). The electrode was held at -1.0 V vs. Ag/AgCl constant potential during the measurements in parallel with periodically interrupted irradiation (30 s ON/OFF).

Kelvin probe measurements

The construction of FTO/Au/Sb₂Se₃ band diagram, determination of Fermi-level values, and surface photovoltage (SPV) measurements were carried out with a KP Technology APS04 Kelvin-probe instrument equipped with a 2 mm diameter gold alloy-coated tip. The Fermi level (E_F) of the tip was determined relative to a silver reference target ($E_F(\text{Au tip}) = -4.60$ eV). The KP system is coupled with a glove box (Mbraun, Labmaster Pro) in Ar atmosphere adjusting the O₂ and H₂O levels to <0.1 ppm.

Contact potential difference measurements (CPD), ambient-pressure photoemission spectroscopy (APS), and surface photovoltage spectroscopy (SPS) were applied to determine the Fermi-level, the valence band, and conduction band positions, respectively.

Results and discussions

Preparation of FTO/Au/Sb₂Se₃/TiO₂/Au photoelectrodes

The workflow of the preparation of the FTO/Au/Sb₂Se₃/TiO₂/Au photoelectrodes is summarized in Fig. 1. The Sb₂Se₃ nanorod films were prepared by a multi-step spin coating method, detailed in the Experimental section.⁵ To conformally coat the TiO₂ protection/electron extraction layer on top of the Sb₂Se₃ nanorod films (aspect ratio = 3.5), we employed ALD. The deposition was carried out for 1000 cycles, corresponding to a TiO₂ layer thickness of ~ 60 nm, which is considered ideal for PEC applications (offers complete blocking).⁵¹ A mild temperature procedure was chosen to ensure that no chemical (surface transformations) or physical (cracking of the films) alteration occurs. This also ensures that an amorphous TiO₂ is formed that can provide better passivation.⁵¹ The ALD deposition was followed by the photo-electrodeposition of the Au cocatalyst on the surface of these assemblies. With this technique, the growth of the catalyst particles is ensured on the most active sites of the photoelectrodes. Our main focus was to deposit Au cocatalyst and perform PEC CO₂RR. By regulating the parameters of this deposition procedure, control over the size and density of the deposited catalyst particles was achieved.

Characterization of FTO/Au/Sb₂Se₃/TiO₂ assemblies

Materials characterization was performed at the different stages of the preparation of the FTO/Au/Sb₂Se₃/TiO₂ assemblies. To confirm the formation of Sb₂Se₃ from the precursor

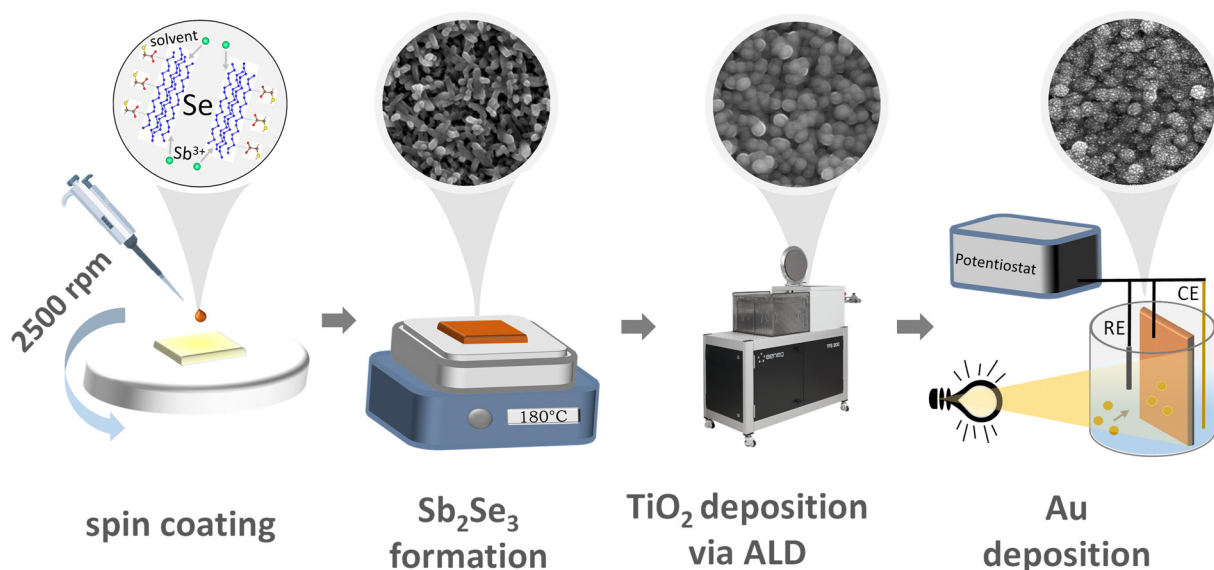


Fig. 1 Schematic representation of the photoelectrode preparation from the Sb₂Se₃ layer synthesis via spin coating, up to the Au catalyst grafting using photo-electrodeposition, with the corresponding SEM images of the thin films. The illustration for the ink was adopted from ref. 25.



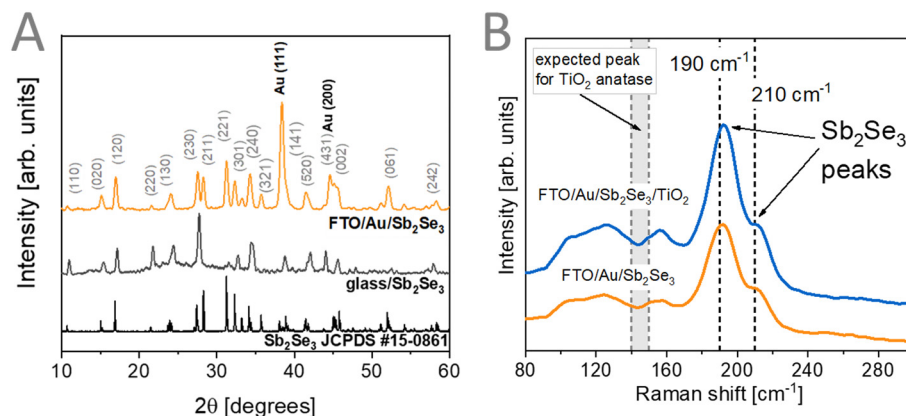


Fig. 2 (A) XRD patterns of Sb_2Se_3 thin films prepared on glass and FTO/Au. (B) Raman spectra of FTO/Au/ Sb_2Se_3 films before and after TiO_2 deposition.

ink, we performed XRD measurements (Fig. 2A). Reflections from the used FTO/Au substrate were visible together with reflections related to the thin Sb_2Se_3 (JCPDS #15-0861) films. To deconvolute the contribution of the used FTO/Au substrates, we prepared Sb_2Se_3 thin films on glass substrates (Fig. 2A). In this case, better-resolved reflections were observed confirming the formation of Sb_2Se_3 with this preparation method. To reveal the phase purity of the formed Sb_2Se_3 films, Raman-spectroscopic measurements were carried out (Fig. 2B). The comparison of the observed vibrations with literature values is shown in Table S1 (ESI†). This comparison reveals the presence of Sb_2Se_3 without the formation of impurity phases, such as Sb_2O_3 or different forms of Se which could be observed at 250 cm^{-1} if present.⁹ After the deposition of the ALD TiO_2 coating, no change in the Raman spectra of the samples was observed. This signals that no impurity phases were formed after the mild deposition. Furthermore, no vibrational modes related to TiO_2 were observed on the recorded spectra revealing that under these deposition conditions (140°C), an amorphous TiO_2 coating is formed.^{51–53} Amorphous TiO_2 layers formed under mild conditions are considered beneficial for PEC applications, as they offer proper electron conductivity to drive reduction reactions.⁵⁴

XPS measurements were carried out (Fig. 3) to assess the surface chemical composition of the prepared samples. The binding energies for the different chemical elements and their assignment is shown in the ESI† (Table S2). The survey spectra (Fig. 3A) together with the core-level spectra of Sb 3d (Fig. 3B) reveal the formation of Sb_2Se_3 (Sb $3d_{5/2}$ at 529.17 eV , 57 at%) together with the presence of a surface Sb_2O_3 layer (Sb $3d_{5/2}$ at 538.51 eV , 43 at%). As Raman spectroscopy probes the bulk of the sample, this surface oxide layer is confined to the upmost part of the Sb_2Se_3 films. After ALD deposition of TiO_2 , the Sb and Se related peaks disappear completely, indicating that the TiO_2 had conformal coverage over the Sb_2Se_3 film. The Ti 2p (Fig. 3C) and the O 1s (Fig. 3B) core-level spectra indicate the formation of TiO_2 without the presence of reduced forms of Ti.

Top-down SEM images were recorded (Fig. S2A and B, ESI†) at the different stages of the synthetic protocol to verify the conformality of the ALD TiO_2 coating. During the multi-step

spin coating procedure, Sb_2Se_3 nanorods were formed with an average diameter of $130 \pm 25\text{ nm}$. The thickness of the Sb_2Se_3 films was determined by cross-section SEM-FIB measurements (Fig. S2D–F, ESI†). An average thickness of $302 \pm 90\text{ nm}$ was obtained. After the deposition of the ALD TiO_2 layer, the nanorod structure was preserved, however with more rounded edges (Fig. S2B, ESI†). These samples showed a slightly larger thickness of $361 \pm 77\text{ nm}$ caused by the deposited TiO_2 layer. The light absorption properties of the Sb_2Se_3 films were characterized by UV-vis-NIR spectroscopy (Fig. S3A, ESI†). To extract the bandgap energy of the prepared Sb_2Se_3 films, Tauc analysis (Fig. S3B, ESI†), considering an indirect bandgap (supported by DFT calculations), was performed.⁵⁵ A bandgap energy of 1.22 eV was determined which is in good agreement with literature values.^{48,55}

To assess the PEC performance of the FTO/Au/ Sb_2Se_3 / TiO_2 assemblies, we have benchmarked our photoelectrodes in PEC hydrogen evolution reaction (HER). We photo-deposited Pt as a cocatalyst on top of the ALD TiO_2 layer (Fig. S4A, ESI†). The maximum photocurrent that our system could deliver was 19 mA cm^{-2} at -0.35 V vs. RHE (Fig. S4B, ESI†) which is in good agreement with the highest reported examples in the literature.^{5,25} This benchmarking was performed to remove the effect of variations from the FTO/Au/ Sb_2Se_3 / TiO_2 architecture, and therefore allow meaningful comparison of the CO_2RR data for the different photoelectrode assemblies.

Deposition of the Au catalyst

Fig. 4A illustrates the desired sample morphology we sought to obtain. The individual Sb_2Se_3 nanorods are conformally coated with the TiO_2 layers. The conformality of the TiO_2 layers is critically important to avoid the exposure of the bottom contact (*i.e.*, FTO/Au) to the electrolyte. PEC methods can be used as a next step to decorate this architecture with the Au cocatalyst.

Linear sweep voltammograms (LSVs) (Fig. 4C) were recorded under different conditions to optimize the deposition procedure of the Au cocatalyst on the Sb_2Se_3 / TiO_2 structure. As observed in the dashed gray LSV curve recorded in the absence of the Au salt, at an applied potential more negative than $-0.3\text{ V vs. Ag/AgCl}$, a reduction process occurs. When performing



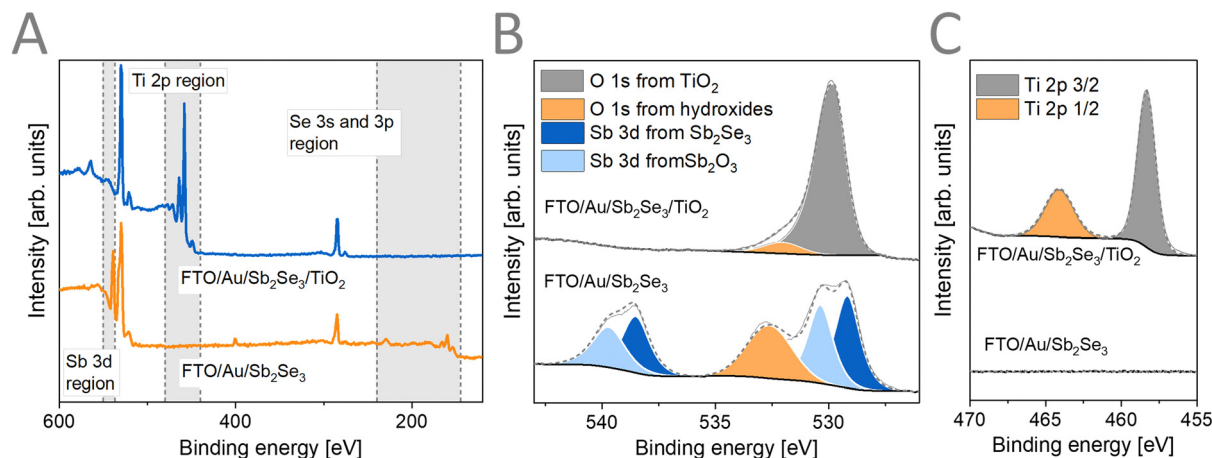


Fig. 3 (A) XPS survey spectra, (B) Sb 3d with O 1s, and (C) Ti 2p high resolution core-level XPS spectra of FTO/Au/Sb₂Se₃ films before and after TiO₂ deposition.

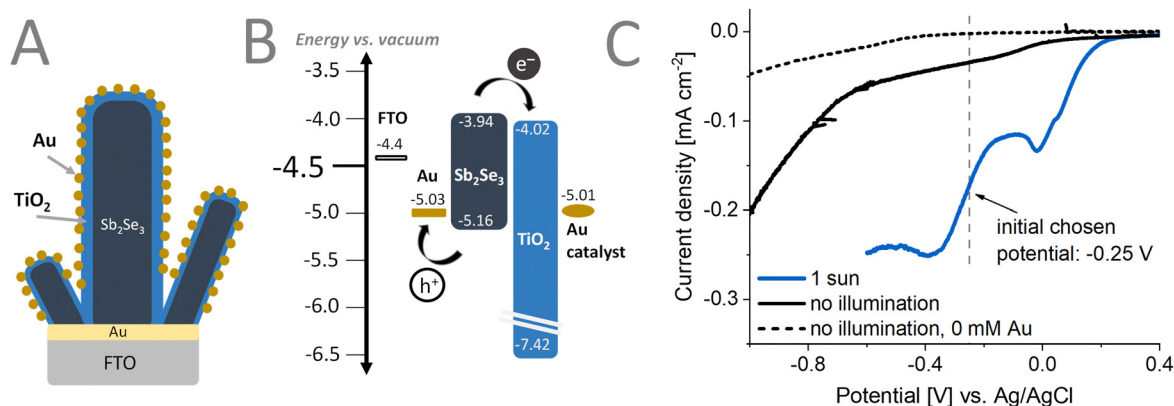


Fig. 4 (A) Schematic representation of the photoelectrode assembly with the corresponding (B) band diagram. (C) Linear sweep voltammogram of the photoelectrode with and without 1 sun illumination with UV filter (AM1.5G, 100 mW cm⁻², $\lambda > 400$ nm) in a solution of 0.5 mM KAuCl₄ and 0.25 mM H₂SO₄ with 5% v/v methanol.

the Au deposition (Fig. 4C) under 1 sun illumination (blue line), an earlier onset potential of the deposition of Au can be observed, compared to the electrodeposition without illumination (black line). This is particularly helpful in avoiding other redox phenomena such as corrosion that might happen at more negative applied potentials. Furthermore, the applied potential suppresses recombination by extracting one of the charge carriers (in our case, holes) generated by the illumination. The long-lived carriers can better reach the surface of Sb₂Se₃ and participate in the chemical reaction (*i.e.*, reduction of Au³⁺). By applying different deposition potentials under illumination, the properties (*e.g.*, size, morphology, distribution) of the Au catalyst particles can be fine-tuned.

To understand the electronic coupling between the materials in the photoelectrode assembly, we performed UPS and contact potential difference measurements (Fig. S5–S7 and Table S3, ESI[†]), determining their Fermi-level and the valence band position. Together with the determined bandgap energies, the band diagram can be constructed (Fig. 4B). There was good agreement in the determined Fermi-level, band positions of the constituents

except for the deposited Au catalyst. This can be ascribed to the inhomogeneity of the Au catalyst layer, that influences the work-function determination by UPS.⁵⁶ After light excitation, the generated electrons in the Sb₂Se₃ can be efficiently extracted by the ALD deposited TiO₂ layer. A conduction band offset of only 0.08 eV indicates an intimate contact (good interface quality) between these materials. In a similar manner, the used FTO/Au substrate, apart from orienting the growth of the Sb₂Se₃ nanorods, is capable of efficiently extracting the generated holes. The deep lying valence band of ALD TiO₂ can also act as a hole-blocking layer, further suppressing electron-hole recombination at the surface of the samples.

Top-down SEM images (Fig. 5) show that the applied potential and passed charge both have an impact on the morphology and density of the deposited Au catalyst. We performed these depositions at three different potentials (−0.4 V, −0.25 V, and 0.1 V vs. Ag/AgCl) for three different total charge densities (25, 50, and 75 mC cm⁻²). These values correspond to 17, 34, and 51 ng cm⁻² of Au loading respectively. When −0.4 V vs.



Fig. 5 SEM images of FTO/Au/Sb₂Se₃/TiO₂ with Au cocatalyst deposited at different biases -0.4 V (left: A and D), -0.25 V (middle: B and E), and 0.1 V (right: C and F) [referenced vs. Ag/AgCl] and at varying amounts of charge passed in mC cm^{-2} : 25 (top: A–C), 75 (bottom: D–F). The scale bar corresponds to $1 \mu\text{m}$.

Ag/AgCl (Fig. 5A and D) or -0.25 V vs. Ag/AgCl (Fig. 5B and E) was used with the same passed charge, small sized, finely dispersed Au catalysts were deposited on the sample surface. Notably, the average diameter of the Au nanoparticles at the same passed charge of 25 mC cm^{-2} were $26 \pm 5 \text{ nm}$ and $29 \pm 6 \text{ nm}$ for deposition potentials of -0.4 V vs. Ag/AgCl and -0.25 V vs. Ag/AgCl, respectively. For depositions performed at 0.1 V vs. Ag/AgCl, the Au catalysts are larger and more sparsely spaced (Fig. 5C and F) with an average diameter of $72 \pm 17 \text{ nm}$ at 25 mC cm^{-2} of passed charge (see size distribution in Fig. S8, ESI†). This signals that the nucleation of the Au catalyst is enhanced at deposition potentials that are more negative, leading to Au particles that are small and finely distributed throughout the Sb₂Se₃/TiO₂ surface. By increasing the passed charge during deposition, the Au particles increased in size, ultimately leading to aggregation as observed in the 75 mC cm^{-2} case (Fig. 5D–F). Interestingly, when compared to the benchmark photoelectrodes with photo-deposited Pt having an average particle diameter of $12 \pm 2 \text{ nm}$ (Fig. S4, ESI†), the photo-electrodeposited Au particles are larger.

Photoelectrochemical CO₂RR

To assess the PEC performance of the prepared Sb₂Se₃/TiO₂/Au assemblies in CO₂RR, we recorded linear sweep photovoltammograms under chopped illumination with 1 sun (Fig. 6A), in CO₂-saturated 0.5 M KHCO_3 electrolyte. Photoelectrodes where the Au deposition was carried out at -0.25 V vs. Ag/AgCl for a total charge passed of 75 mC cm^{-2} delivered a maximum photocurrent of $\sim 7.5 \text{ mA cm}^{-2}$ at -0.57 V vs. RHE. This performance successfully competes with other single junction photoelectrodes using Au as the catalyst (Table S4, ESI†). We performed IPCE measurements (Fig. S9, ESI†) and observed that the rise of the photocurrent corresponds to the optical bandgap of Sb₂Se₃ which shows that there are negligible

electronic losses in the assembly. In contrast, when performing dark CO₂RR on a Au foil under similar conditions, the magnitude of the current was only $\sim 2 \text{ mA cm}^{-2}$ at -0.57 V vs. RHE. Using a lower loading amount of 10 mC cm^{-2} leads to a decrease in the photocurrent (Fig. S10A and Fig. 6C, ESI†). Surprisingly, we found that after carrying out the deposition of Au catalyst above 25 mC cm^{-2} , no further increase in the measured photocurrent (Fig. S11, ESI†) could be observed. This points toward other limiting factors (*e.g.*, interface between the TiO₂/Au) in these assemblies. Fig. 6B illustrates that without having all components in the photoelectrode, no PEC activity in CO₂RR could be observed. Adding the Au catalyst on top of the bare Sb₂Se₃ without TiO₂ leads to a poor rectifying behavior, which primarily acts like the standard Au foil under dark (orange line Fig. 6B similar to the star line in Fig. 6A; long-term performance Fig. S12, ESI†). This is because the photogenerated carriers are not extracted efficiently without any charge-selective contact (hole-extraction or electron extraction layer) allowing charge recombination to prevail. On the other hand, with solely a TiO₂ protection layer, the rectifying behavior can be observed (blue line Fig. 6B). As no catalyst is deposited on the surface, and TiO₂ itself drives PEC CO₂RR poorly, no photocurrent could be observed. Achieving a satisfactory PEC performance can only be observed when the Au cocatalyst is deposited on the TiO₂ surface.

To evaluate the stability of the assemblies, photoelectrolysis was carried out for 30 minutes at an applied potential of -0.57 V vs. RHE with three different loadings of the Au catalyst (Fig. 6D). The lowest loading of 25 mC cm^{-2} showed relatively higher photocurrent at the first 1200 s and then reached the same performance as the other two loadings. The initial decay of the photocurrent is apparent, and this is attributed to the inherent behavior of Au under similar CO₂RR conditions (Fig. S13, ESI†). The used Au foil also displays a similar decay trend as the prepared photoelectrodes. This decay can be



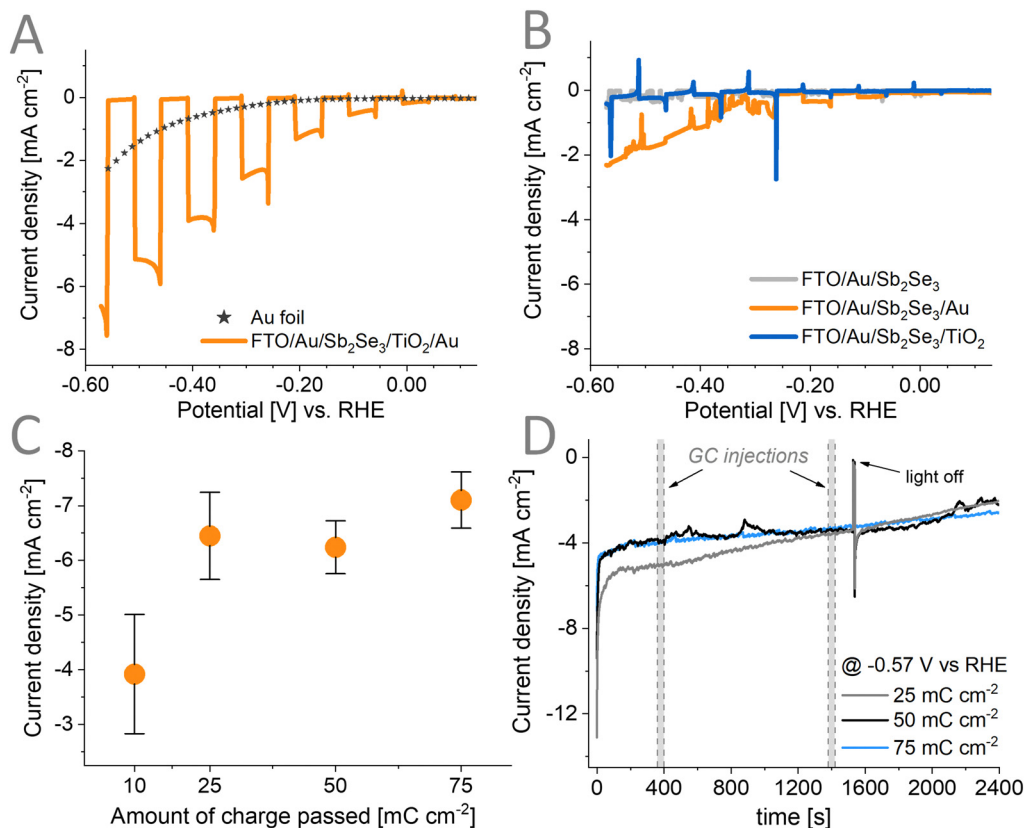


Fig. 6 Linear sweep photovoltammogram under 1 sun illumination (AM1.5G , 100 mW cm^{-2}) of (A) the complete architecture (Au catalyst loading of 75 mC cm^{-2} deposited at $-0.25 \text{ V vs. Ag/AgCl}$) and (B) the set of incomplete electrode architectures in a CO_2 -saturated solution of 0.5 M KHCO_3 . (C) Recorded current density at -0.57 V vs. RHE for different catalyst loadings deposited at $-0.25 \text{ V vs. Ag/AgCl}$. (D) Long-term chronoamperometry measurement of the architecture with different catalyst loadings deposited at $-0.25 \text{ V vs. Ag/AgCl}$.

divided into two regions: (1) a quick initial drop for the first 10 s and (2) a constant slow decay. We have performed XPS and Raman spectroscopy to characterize the chemical composition of the photoelectrodes in these two regions. Both techniques suggest that the chemical composition of the electrodes remain unchanged within the 1st region (Fig. S14A and S15A–C, ESI†).

Therefore, this quick drop at the start might be due to the double layer charging effect resulting from the sudden step from OCP to an applied $-1.2 \text{ V vs. Ag/AgCl}$ potential during these measurements. For the second region multiple effects can contribute to the observed behavior such as: (i) the change in alkalinity of the electrolyte as the $\text{CO}_2\text{RR/HER}$ proceeds,⁵⁷

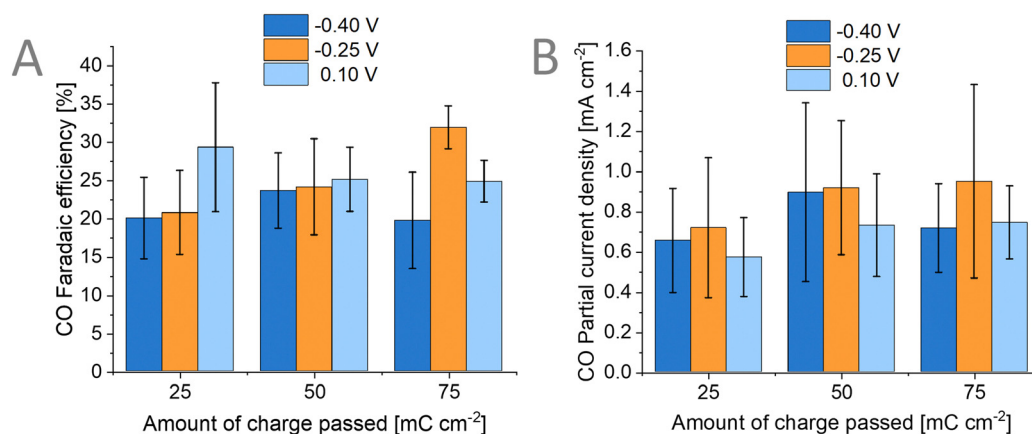


Fig. 7 (A) Faradaic efficiency for CO of the $\text{FTO/Au/Sb}_2\text{Se}_3/\text{TiO}_2/\text{Au}$ photoelectrode with corresponding (B) partial current density for CO at different deposition bias (-0.4 V , -0.25 V , $0.1 \text{ V vs. Ag/AgCl}$) and different amount of charge passed (25 , 50 , 75 mC cm^{-2}) for the Au catalyst deposition.



(ii) photocorrosion of the TiO_2 layer, (iii) the corrosion of the Sb_2Se_3 once TiO_2 fully delaminates, and (iv) partial removal of the Au catalysts. Fig. S10B (ESI†) also shows that redeposition of the Au catalyst does not retrieve the initial photocurrent. This points to the fact that part of the CO_2RR activity fading is permanent and not only caused by the catalyst removal, and the photocorrosion of the TiO_2 layer is also at play (shown by Raman and XPS results in Fig. S14 and S15, ESI†).

Product distribution during PEC CO_2RR

The main product when performing CO_2RR on Au surfaces is CO .⁵⁸ Aside from CO , hydrogen is also reported in many studies as a by-product. We performed an Ar-bubbled photoelectrolysis experiment at -0.57 V vs. RHE and we found that no CO product was formed but only hydrogen (Fig. S16, ESI†). To verify the source of the CO in our PEC experiments (where CO_2 was bubbled), we have performed CO_2RR with labelled $^{13}\text{CO}_2$ and $\text{KH}^{13}\text{CO}_3$ (Fig. S17, ESI†). These experiments proved that the observed CO is not from any additional source used in the photoelectrode or the setup itself. We also carried out liquid product detection experiments using ion chromatography and NMR spectroscopy which showed that no liquid products were formed during the photoelectrolysis. We explored the effects of varying both (1) the deposition potential of the Au catalyst and (2) the amount of Au catalyst on the performance of the photoelectrodes for CO_2RR . Although the particle size and morphology of the Au catalyst differ as previously seen in the SEM images (Fig. 5), we found that this has no correlation to the products detected and that there is no clear trend to be observed (Fig. 7). For the amount of 50 mC cm^{-2} of charge passed, the average FE_{CO} is around 24% independent of the applied potential during the deposition of Au. The size range of the Au nanoparticles ($26 \pm 5 \text{ nm}$, $29 \pm 6 \text{ nm}$, up to $72 \pm 17 \text{ nm}$) is too large to have a significant contribution to the performance as reported in the literature for similar Au nanoparticle-based photoelectrodes.^{31,59}

Conclusions

In summary, we have studied Sb_2Se_3 -based photoelectrodes for CO_2 reduction by harnessing the photogenerated charge carriers of the semiconductor and the catalytic activity of photo-electrodeposited Au cocatalysts. Using a materials-focused approach of photocathode fabrication, we have developed (i) a nanostructured semiconductor that has good light absorption, (ii) a conformal amorphous TiO_2 coating that acted as both protection layer and electron extraction layer, and (iii) a highly dispersed Au nanoparticle cocatalyst. This assembly was explored in PEC CO_2RR , where the photoelectrode architecture delivered an average photocurrent density of $\sim 7.5 \text{ mA cm}^{-2}$ at -0.57 V vs. RHE, which is a threefold increase from the performance of a Au foil in the same CO_2RR conditions (2 mA cm^{-2}). This improved performance compared to a Au foil validates the research directions towards employing direct PEC CO_2RR . The FE_{CO} was not affected significantly by the deposition bias nor the amount of Au catalyst.

The average FE_{CO} is $25 \pm 6\%$ while FE_{H_2} is $63 \pm 12\%$, indicating roughly a $\sim 1 : 2.5 \text{ CO/H}_2$ ratio. Overall, we have shown that Sb_2Se_3 -based photoelectrode assemblies can also be used for PEC CO_2RR by employing Au cocatalyst. This study highlights the importance of utilizing already existing architectures and repurposing/re-engineering them for the specific applications intended.

Author contributions

John Mark Christian M. Dela Cruz: methodology, investigation, validation, data curation, writing – original draft, visualization, project administration. Ádám Balog: investigation, data curation. Péter S. Tóth: investigation, data curation. Gábor Bencsik: investigation. Gergely F. Samu: conceptualization, investigation, data curation, writing – review & editing, supervision. Csaba Janáky: conceptualization, resources, writing – review & editing, supervision, funding acquisition.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- 1 G. Segev, J. Kibsgaard, C. Hahn, Z. J. Xu, W.-H. (Sophia) Cheng, T. G. Deutsch, C. Xiang, J. Z. Zhang, L. Hammarström, D. G. Nocera, A. Z. Weber, P. Agbo, T. Hisatomi, F. E. Osterloh, K. Domen, F. F. Abdi, S. Haussener, D. J. Miller, S. Ardo, P. C. McIntyre, T. Hannappel, S. Hu, H. Atwater, J. M. Gregoire, M. Z. Ertem, I. D. Sharp, K.-S. Choi, J. S. Lee, O. Ishitani, J. W. Ager, R. R. Prabhakar, A. T. Bell, S. W. Boettcher, K. Vincent, K. Takanabe, V. Artero, R. Napier, B. R. Cuenya, M. T. M. Koper, R. V. D. Krol and F. Houle, *J. Phys. Appl. Phys.*, 2022, 55, 323003.
- 2 R. van de Krol and B. A. Parkinson, *MRS Energy Sustain.*, 2017, 4, 13.



- 3 Advanced Alternative Fuels – Technology Development Report 2020.
- 4 S. Ardo, D. F. Rivas, M. A. Modestino, V. S. Greiving, F. F. Abdi, E. A. Llado, V. Artero, K. Ayers, C. Battaglia, J.-P. Becker, D. Bederak, A. Berger, F. Buda, E. Chinello, B. Dam, V. D. Palma, T. Edvinsson, K. Fujii, H. Gardeniers, H. Geerlings, S. M. H. Hashemi, S. Haussener, F. Houle, J. Huskens, B. D. James, K. Konrad, A. Kudo, P. P. Kunturu, D. Lohse, B. Mei, E. L. Miller, G. F. Moore, J. Muller, K. L. Orchard, T. E. Rosser, F. H. Saadi, J.-W. Schüttauf, B. Seger, S. W. Sheehan, W. A. Smith, J. Spurgeon, M. H. Tang, R. van de Krol, P. C. K. Vesborg and P. Westerik, *Energy Environ. Sci.*, 2018, **11**, 2768–2783.
- 5 J. Kim, W. Yang, Y. Oh, H. Lee, S. Lee, H. Shin, J. Kim and J. Moon, *J. Mater. Chem. A*, 2017, **5**, 2180–2187.
- 6 S. Chen, T. Liu, Z. Zheng, M. Ishaq, G. Liang, P. Fan, T. Chen and J. Tang, *J. Energy Chem.*, 2022, **67**, 508–523.
- 7 H. Lee, W. Yang, J. Tan, Y. Oh, J. Park and J. Moon, *ACS Energy Lett.*, 2019, **4**, 995–1003.
- 8 J. Park, W. Yang, J. Tan, H. Lee, J. W. Yun, S. G. Shim, Y. S. Park and J. Moon, *ACS Energy Lett.*, 2020, **5**, 136–145.
- 9 J. Tan, W. Yang, Y. Oh, H. Lee, J. Park, R. Boppella, J. Kim and J. Moon, *Adv. Energy Mater.*, 2019, **9**, 1900179.
- 10 X. Liu, J. Chen, M. Luo, M. Leng, Z. Xia, Y. Zhou, S. Qin, D.-J. Xue, L. Lv, H. Huang, D. Niu and J. Tang, *ACS Appl. Mater. Interfaces*, 2014, **6**, 10687–10695.
- 11 W. Yang, X. Zhang and S. D. Tilley, *Chem. Mater.*, 2021, **33**, 3467–3489.
- 12 Y. Zhou, L. Wang, S. Chen, S. Qin, X. Liu, J. Chen, D.-J. Xue, M. Luo, Y. Cao, Y. Cheng, E. H. Sargent and J. Tang, *Nat. Photonics*, 2015, **9**, 409–415.
- 13 J. Park, W. Yang, Y. Oh, J. Tan, H. Lee, R. Boppella and J. Moon, *ACS Energy Lett.*, 2019, **4**, 517–526.
- 14 X. Wen, C. Chen, S. Lu, K. Li, R. Kondrotas, Y. Zhao, W. Chen, L. Gao, C. Wang, J. Zhang, G. Niu and J. Tang, *Nat. Commun.*, 2018, **9**, 2179.
- 15 O. S. Hutter, L. J. Phillips, K. Durose and J. D. Major, *Sol. Energy Mater. Sol. Cells*, 2018, **188**, 177–181.
- 16 C. Chen, K. Li, S. Chen, L. Wang, S. Lu, Y. Liu, D. Li, H. Song and J. Tang, *ACS Energy Lett.*, 2018, **3**, 2335–2341.
- 17 G.-X. Liang, Z.-H. Zheng, P. Fan, J.-T. Luo, J.-G. Hu, X.-H. Zhang, H.-L. Ma, B. Fan, Z.-K. Luo and D.-P. Zhang, *Sol. Energy Mater. Sol. Cells*, 2018, **174**, 263–270.
- 18 T. T. Ngo, S. Chavhan, I. Kosta, O. Miguel, H.-J. Grande and R. Tena-Zaera, *ACS Appl. Mater. Interfaces*, 2014, **6**, 2836–2841.
- 19 M. B. Costa, F. W. S. Lucas and L. H. Mascaro, *J. Mater. Chem. A*, 2020, **8**, 13742–13753.
- 20 R. R. Prabhakar, W. Septina, S. Siol, T. Moehl and S. D. Tilley, *J. Mater. Chem. A*, 2017, **5**, 23139–23145.
- 21 L. Gao, Y. Cui, R. H. J. Vervuurt, D. van Dam, R. P. J. van Veldhoven, J. P. Hofmann, A. A. Bol, J. E. M. Haverkort, P. H. L. Notten, E. P. A. M. Bakkers and E. J. M. Hensen, *Adv. Funct. Mater.*, 2016, **26**, 679–686.
- 22 N. Guijarro, M. S. Prévot and K. Sivula, *Phys. Chem. Chem. Phys.*, 2015, **17**, 15655–15674.
- 23 M. Crespo-Quesada and E. Reisner, *Energy Environ. Sci.*, 2017, **10**, 1116–1127.
- 24 P. Acevedo-Peña, J. Vazquez-Arenas, R. Cabrera-Sierra, L. Lartundo-Rojas and I. González, *J. Electrochem. Soc.*, 2013, **160**, C277.
- 25 W. Yang, J. Ahn, Y. Oh, J. Tan, H. Lee, J. Park, H.-C. Kwon, J. Kim, W. Jo, J. Kim and J. Moon, *Adv. Energy Mater.*, 2018, **8**, 1702888.
- 26 J. Tan, B. Kang, K. Kim, D. Kang, H. Lee, S. Ma, G. Jang, H. Lee and J. Moon, *Nat. Energy*, 2022, 1–11.
- 27 X. Li, J. Yu, M. Jaroniec and X. Chen, *Chem. Rev.*, 2019, **119**, 3962–4179.
- 28 V. Kumaravel, J. Bartlett and S. C. Pillai, *ACS Energy Lett.*, 2020, **5**, 486–519.
- 29 J. L. White, M. F. Baruch, J. E. Pander, Y. Hu, I. C. Fortmeyer, J. E. Park, T. Zhang, K. Liao, J. Gu, Y. Yan, T. W. Shaw, E. Abelev and A. B. Bocarsly, *Chem. Rev.*, 2015, **115**, 12888–12935.
- 30 G. Liu, P. R. Narangari, Q. T. Trinh, W. Tu, M. Kraft, H. H. Tan, C. Jagadish, T. S. Choksi, J. W. Ager, S. Karuturi and R. Xu, *ACS Catal.*, 2021, **11**, 11416–11428.
- 31 S. Chu, P. Ou, R. T. Rashid, Y. Pan, D. Liang, H. Zhang and J. Song, *Green Energy Environ.*, 2022, **7**, 545–553.
- 32 Y. J. Jang, J.-W. Jang, J. Lee, J. H. Kim, H. Kumagai, J. Lee, T. Minegishi, J. Kubota, K. Domen and J. S. Lee, *Energy Environ. Sci.*, 2015, **8**, 3597–3604.
- 33 K. Wang, N. Fan, B. Xu, Z. Wei, C. Chen, H. Xie, W. Ye, Y. Peng, M. Shen and R. Fan, *Small*, 2022, **18**, 2201882.
- 34 J. T. Song, H. Ryoo, M. Cho, J. Kim, J.-G. Kim, S.-Y. Chung and J. Oh, *Adv. Energy Mater.*, 2017, **7**, 1601103.
- 35 Y. J. Jang, I. Jeong, J. Lee, J. Lee, M. J. Ko and J. S. Lee, *ACS Nano*, 2016, **10**, 6980–6987.
- 36 J. S. DuChene, G. Tagliabue, A. J. Welch, W.-H. Cheng and H. A. Atwater, *Nano Lett.*, 2018, **18**, 2545–2550.
- 37 I. Roh, S. Yu, C.-K. Lin, S. Louisia, S. Cestellos-Blanco and P. Yang, *J. Am. Chem. Soc.*, 2022, **144**, 8002–8006.
- 38 B. Zhou, P. Ou, N. Pant, S. Cheng, S. Vanka, S. Chu, R. T. Rashid, G. Botton, J. Song and Z. Mi, *Proc. Natl. Acad. Sci. U. S. A.*, 2020, **117**, 1330–1338.
- 39 H. A. Chaliyawala, S. Bastide, D. Muller-Bouvet, C. Pichon, K. Bah, A. Djoumoi, F. Marty, T. Bourouina and E. Torralba, *ACS Appl. Energy Mater.*, 2023, **6**, 8397–8409.
- 40 Y. Hu, F. Chen, P. Ding, H. Yang, J. Chen, C. Zha and Y. Li, *J. Mater. Chem. A*, 2018, **6**, 21906–21912.
- 41 C. Kim, S. Choi, M.-J. Choi, S. A. Lee, S. H. Ahn, S. Y. Kim and H. W. Jang, *Appl. Sci.*, 2020, **10**, 3487.
- 42 P. B. Pati, R. Wang, E. Boutin, S. Diring, S. Jobic, N. Barreau, F. Odobel and M. Robert, *Nat. Commun.*, 2020, **11**, 3499.
- 43 A. Call, M. Cibian, K. Yamamoto, T. Nakazono, K. Yamauchi and K. Sakai, *ACS Catal.*, 2019, **9**, 4867–4874.
- 44 V. Andrei, B. Reuillard and E. Reisner, *Nat. Mater.*, 2020, **19**, 189–194.
- 45 Q. Wang, J. Warnan, S. Rodríguez-Jiménez, J. J. Leung, S. Kalathil, V. Andrei, K. Domen and E. Reisner, *Nat. Energy*, 2020, **5**, 703–710.



- 46 H. A. Hansen, J. B. Varley, A. A. Peterson and J. K. Nørskov, *J. Phys. Chem. Lett.*, 2013, **4**, 388–392.
- 47 S. Chu, P. Ou, P. Ghamari, S. Vanka, B. Zhou, I. Shih, J. Song and Z. Mi, *J. Am. Chem. Soc.*, 2018, **140**, 7869–7877.
- 48 W. Yang and J. Moon, *J. Mater. Chem. A*, 2019, **7**, 20467–20477.
- 49 C. Tossi, L. Hällström, J. Selin, M. Vaelma, E. See, J. Lahtinen and I. Tittonen, *J. Mater. Chem. A*, 2019, **7**, 14519–14525.
- 50 N. Fairley, V. Fernandez, M. Richard-Plouet, C. Guillot-Deudon, J. Walton, E. Smith, D. Flahaut, M. Greiner, M. Biesinger, S. Tougaard, D. Morgan and J. Baltrusaitis, *Appl. Surf. Sci. Adv.*, 2021, **5**, 100112.
- 51 T. Moehl, J. Suh, L. Sévery, R. Wick-Joliat and S. D. Tilley, *ACS Appl. Mater. Interfaces*, 2017, **9**, 43614–43622.
- 52 A. Jolivet, C. Labbé, C. Frilay, O. Debieu, P. Marie, B. Horcholle, F. Lemarié, X. Portier, C. Grygiel, S. Duprey, W. Jadwisieniczak, D. Ingram, M. Upadhyay, A. David, A. Fouchet, U. Lüders and J. Cardin, *Appl. Surf. Sci.*, 2023, **608**, 155214.
- 53 A. Niilisk, M. Moppel, M. Pärs, I. Sildos, T. Jantson, T. Avarmaa, R. Jaaniso and J. Aarik, *Open Phys.*, 2006, **4**, 105–116.
- 54 S. Hu, M. R. Shaner, J. A. Beardslee, M. Lichterman, B. S. Brunschwig and N. S. Lewis, *Science*, 2014, **344**, 1005–1009.
- 55 R. Vadapoo, S. Krishnan, H. Yilmaz and C. Marin, *Phys. Status Solidi B*, 2011, **248**, 700–705.
- 56 J. W. Kim and A. Kim, *Curr. Appl. Phys.*, 2021, **31**, 52–59.
- 57 A. Goyal, G. Marcandalli, V. A. Mints and M. T. M. Koper, *J. Am. Chem. Soc.*, 2020, **142**, 4154–4161.
- 58 Y. Hori, in *Modern Aspects of Electrochemistry*, ed. C. G. Vayenas, R. E. White and M. E. Gamboa-Aldeco, Springer New York, New York, NY, 2008, vol. 42, pp. 89–189.
- 59 C. Li, T. Wang, B. Liu, M. Chen, A. Li, G. Zhang, M. Du, H. Wang, S. F. Liu and J. Gong, *Energy Environ. Sci.*, 2019, **12**, 923–928.

