



# Targeted modifications in ionic liquids – from understanding to design†

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Ionic liquids are extremely versatile and continue to find new applications in academia as well as industry. This versatility is rooted in the manifold of possible ion types, ion combinations, and ion variations. However, to fully exploit this versatility, it is imperative to understand how the properties of ionic liquids arise from their constituents. In this work, we discuss targeted modifications as a powerful tool to provide understanding and to enable design. A 'targeted modification' is a deliberate change in the structure of an ionic liquid. This includes chemical changes in an experiment as well as changes to the parameterisation in a computer simulation. In any case, such a change must be purposeful to isolate what is of interest, studying, as far as is possible, only one concept at a time. The concepts can then be used as design elements. However, it is often found that several design elements interact with each other – sometimes synergistically, and other times antagonistically. Targeted modifications are a systematic way of navigating these overlaps. We hope this paper shows that understanding ionic liquids requires experimentalists and theoreticians to join forces and provides a tool to tackle the difficult transition from understanding to design.

## Introduction

Ionic liquids are liquid salts. The term usually implies a low melting point, often around or below room temperature. They are now well established as a separate class of solvents, distinctively different from conventional molecular solvents.<sup>1–5</sup> Ionic liquids have reached virtually every area of science and technology where solvents are required – from large scale industrial processes<sup>6–8</sup> and battery electrolytes<sup>9–13</sup> to pharmaceuticals,<sup>14–16</sup> from metal deposition<sup>17,18</sup> to protein crystallisation.<sup>19–21</sup>

One reason for this wide spread is the sheer number of possible ionic liquids.<sup>6,22</sup> Every ionic liquid has at least two constituents, cation and anion, which can be chosen independently. Different types of ions and variations thereof allow for near limitless fine tuning of properties. For example, the choice of a cation involves the choice of a type such as imidazolium, phosphonium or guanidinium. Each of these cation types itself can be modified by changing side chain lengths, substitution patterns, functional groups, *etc.* The properties of the resulting ionic liquids are diverse and can thus be tailored to a given application.

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† Electronic supplementary information (ESI) available: Ion abbreviations used in this work, details of potential energy, volume and charge arm calculations, literature data for methylated imidazolium [NTF<sub>2</sub>]<sup>−</sup> melting points, pressure and viscosity estimates for isodensity conditions of [C<sub>4</sub>C<sub>1</sub>im][NTF<sub>2</sub>], structures of cynoborates and z-matrices of imidazoliums used in the *ab initio* calculations. See DOI: 10.1039/d1cp00216c

Michael Freemantle recognised the potential of ionic liquids to serve as what he called 'designer solvents' in a visionary article in 1998.<sup>23</sup> The term 'designer solvents' captures the high expectations of the community that ionic liquids can be optimised and adapted to meet the practical requirements of many diverse applications.<sup>24–32</sup> In this context, ionic liquids are also referred to as 'neoteric solvents'. Neoteric solvents such as ionic liquids or supercritical CO<sub>2</sub> are those that emerged at the end of the 20th century, with the prospect to solve the challenges of the 21st.<sup>33–36</sup> In some cases, functional groups can be covalently attached to an ionic liquid. Such 'task specific ionic liquids' actively participate on a molecular level rather than just providing a supporting solvent, for example by coordinating a metal or by providing an active site for a reaction.<sup>37–42</sup>

To fulfil all these ambitious expectations, we require a thorough understanding of how the properties of an ionic liquid arise from its molecular structure. This is the basic idea behind the search for structure–property relationships: ionic liquids with systematic structural variations are characterised (experimentally and/or theoretically), and the observed changes are then rationalised using chemical concepts.<sup>43–49</sup> However, few properties are shared by most ionic liquids, such as a negligible vapour pressure or good thermal and chemical stability.<sup>24,50</sup> At the same time, some ionic liquids have undesirable properties, such as high viscosity or low biodegradability.<sup>22,51–54</sup> Practical implementations of ionic liquids must aim to exploit the advantages and overcome the issues.



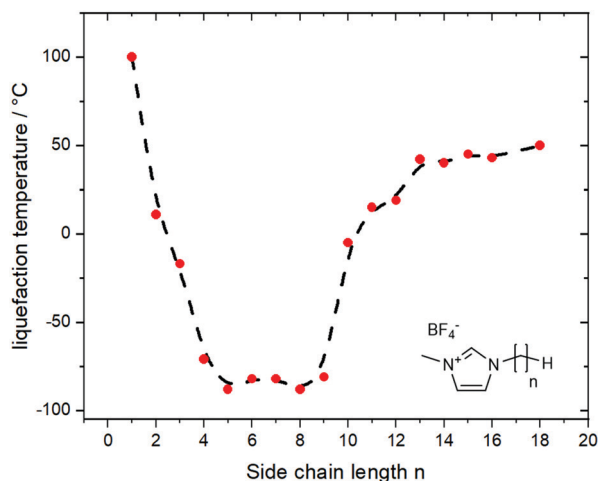


Fig. 1 The thermal transitions for imidazolium ionic liquids show an irregular behaviour with increasing length of the side chain. Data from ref. 56, the dashed line is a guide for the eye.

A thorough understanding of structure–property relationships paves the way for the design of new derivatives and ionic liquids without having to rely on trial and error. However, in the real world, structure–property relationships are rarely independent of each other. An example is the relationship between side chain length (structure) and melting points (property) in 1-alkyl-3-methylimidazolium ionic liquids, Fig. 1.<sup>55,56</sup> Increasing the length of the alkyl side chains reduces Coulombic interactions and introduces configurational entropy, thus decreasing the melting points.<sup>56</sup> At the same time, longer alkyl side chains lead to increased van der Waals interactions, eventually leading to an increase in melting point and the formation of a wax. Thus, ionic liquids are always a compromise, a sweet spot between high melting ‘conventional’ salt and highly viscous oil or wax.

It is important that experimental and theoretical studies consider the dependencies between underlying concepts if they are to be used as design elements. To this end, it is advantageous to choose the objects of study in a way that isolates one underlying concept. We will use the term ‘targeted modification’ for this approach. An example of a targeted modification would be the replacement of an element with a heavier isotope to isolate the influence of mass on molecular properties. Naturally, targeted modifications have some degree of ambiguity, since they depend on the researcher to choose an appropriate underlying concept. The length of a side chain has often been used as a key variable for ionic liquids and could be regarded as such a choice. However, volume, shape, mass and flexibility all depend on the length of the side chain. One needs to be aware of the multiple dependencies where it is not possible to eliminate these. Theoretical methods such as classical MD simulations are at an advantage, since they allow for targeted modifications beyond what is experimentally feasible, such as changing the mass of any atom or molecule without changing other interactions.

In this work, ‘targeted’ is not intended to suggest the pursuit of specific applications, since such an application driven

interpretation is already encompassed in the concept of task specific ionic liquids. Due to the diversity of ionic liquids and their potential applications, the fine tuning of properties for these must be done on a case-by-case basis. The reader interested in (industrial) applications is referred to the literature, including a recently published book.<sup>6,14,38,57–60</sup> Instead, we focus on a general methodology to enable purposeful design, using properties of general importance such as viscosity or thermal transitions as our examples.

Many studies intuitively use (targeted) modifications to understand existing and design new ionic liquids. In this work, we present important concepts which guide the design of ionic liquids and demonstrate how targeted modifications have been used successfully to further our understanding of this design process. In addition, we identify gaps in the literature and discuss the limits of targeted modifications. Naturally, we cannot cover the whole body of literature on ionic liquids, thus our selection is representative, not exhaustive. Where bulk dynamics are relevant, we will use viscosity or diffusion coefficients vicariously for other transport properties, focusing on room temperature behaviour.

## Electrostatics and coulombic compaction

The common denominator shared by all ionic liquids is the inevitable dominance of Coulombic interactions.<sup>61</sup> The electrostatic forces obey Coulomb’s inverse-square law and are thus of considerably longer range than, *e.g.*, dispersion interactions, whose dependency is  $1/R^6$ .<sup>62</sup> Since the charge carriers in ionic liquids are mobile, charge neutrality is fulfilled even at a relatively short distances. The charge of an ion in the bulk ionic liquid is largely screened within a few nanometres,<sup>63–65</sup> but the exact nature and range of electrostatic screening in ionic liquids is still a subject of current investigation.<sup>66–73</sup>

Two major effects of the charged nature of ionic liquids can be identified. First, the bulk ionic liquid experiences a compaction as a direct consequence of the strong electrostatic forces.<sup>74,75</sup> This compaction is equivalent to an increase of pressure or a decrease in molar volume, and is thus expected to dampen the dynamics within the ionic liquid. It is indeed well known that viscosity increases with pressure for most liquids.<sup>76–78</sup> This problem is related to the broader concept of density scaling, that is, the dynamics of glass forming liquids scale not only with temperature  $T$ , but also with density  $\rho$ .<sup>79–82</sup> Thus, transport properties such as viscosity, electrical conductivity, or self-diffusion scale with  $\rho^2/T$ , an approach that has been found to work well for ionic liquids.<sup>83–86</sup> The parameter  $\gamma$  is specific for a given material; however different transport properties for the same material often reveal a common scaling parameter  $\gamma$ .<sup>84,85,87</sup>

Apart from the compaction, the second mechanism by which electrostatic forces influence the properties of ionic liquids is the requirement of charge neutrality. In terms of static properties, this leads to the formation of the shell structure found in ionic liquids, confining ions to an ion cage



formed of oppositely charged ions.<sup>63,88–90</sup> In terms of dynamic properties, the charge neutrality condition in combination with momentum conservation give rise to correlated dynamics.<sup>91</sup> In other words: the net movement of one ion in a given direction must eventually be compensated for by the net movement of a like-charge ion in the opposite direction or the net movement of an oppositely charged ion in the same direction, or a combination of both.

To explore the influence of electrostatics, ionic liquids can be compared to neutral analogues, for which we will use the term ‘molecular mimics’ coined by Niemann *et al.*<sup>92</sup> The concept is illustrated in Fig. 2 and 3. The ions in an ionic liquid are replaced with neutral, (valence) isoelectronic analogues. For example, the central nitrogen atom in the bis(trifluoromethanesulfonyl)imide anion shown in Fig. 3a could be replaced with oxygen, leading to triflic anhydride as a possible molecular mimic. Similarly, replacement of the central nitrogen atom in an ammonium ionic liquid with carbon or silicon leads to synthetically accessible molecular mimics (Fig. 3c).

However, the motifs required to stabilise the negative charge in an anion – such as fluorination and delocalisation – also render their molecular mimics highly reactive. Similar problems arise with some molecular mimics of cations, such as the examples shown for imidazolium in Fig. 3d. Preliminary experiments in our group showed that in many cases, the two components of the molecular mimic of an ionic liquid do not mix (such as triflic anhydride with butylbenzene or methyltripentyl silane) or react immediately (such as triflic anhydride with 2,5-dimethylfuran or *N*-butylpyrrole). It is thus not always feasible to use a binary molecular mimic for an ionic liquid, in which case it is still possible to resort to theoretical methods such as molecular dynamics simulations.<sup>93</sup>

In some cases, the addition of a molecular mimic for one of either the cation or anion will suffice to answer a specific scientific question. For example, Niemann *et al.* have used 2-phenylethanol as the molecular mimic of *N*-(2-hydroxyethyl)pyridinium to

investigate competition between hydrogen bonding and electrostatic interactions in an ionic liquid.<sup>92</sup> The anion in this work was tetrafluoroborate, whose obvious molecular mimics are gaseous at ambient conditions (tetrafluoromethane or silicon tetrafluoride). In a similar approach, Zaitsau *et al.* used amines as molecular mimics of ammonium ionic liquids to quantify the contributions of Coulomb interactions, hydrogen bonding, and dispersion interactions to the overall interaction energy in the bulk ionic liquid.<sup>94</sup>

Despite the difficulties, the concept of binary molecular mimics has been applied successfully by Shirota and Castner.<sup>75</sup> The ionic liquid and the corresponding binary molecular mimic used in their work are shown in Fig. 3e. By means of optical Kerr effect spectroscopy, Shirota and Castner obtained lifetimes for reorientational relaxation. They observed that the ratio of the longest lifetimes matched the predicted ratio from the Stokes–Einstein–Debye theory well, thus corroborating a hydrodynamic interpretation of reorientational relaxation in ionic liquids.<sup>75</sup>

The specific volume of the ionic liquid in Fig. 3e is reduced by approximately 16% compared to its binary molecular mimic, at the same time the viscosity is higher by a factor of 29 for the ionic liquid.<sup>75</sup> The question arises as to what degree compaction contributes to this increase in viscosity. This would require the measurement of transport properties in a binary molecular mimic at a pressure which leads to the same molar volume as the corresponding ionic liquid, Fig. 2. To the best of our knowledge, such a study has not yet been published. Using ethanol as an example liquid, reducing the specific volume by 16% requires pressures of approximately 4 kbar.<sup>95</sup> Viscosities can be measured at significantly higher pressures,<sup>76</sup> thus the proposed approach appears to be feasible and will be pursued in future work.

## Size and mass of ions

Not only the absolute values of size and mass have to be considered for ionic liquids, but also their relation. The concept



Fig. 2 The effect of Coulombic compactisation on the viscosity of ionic liquids can be investigated with molecular mimics, *i.e.* a targeted modification that replaces cations and anions with isostructural neutral molecules.



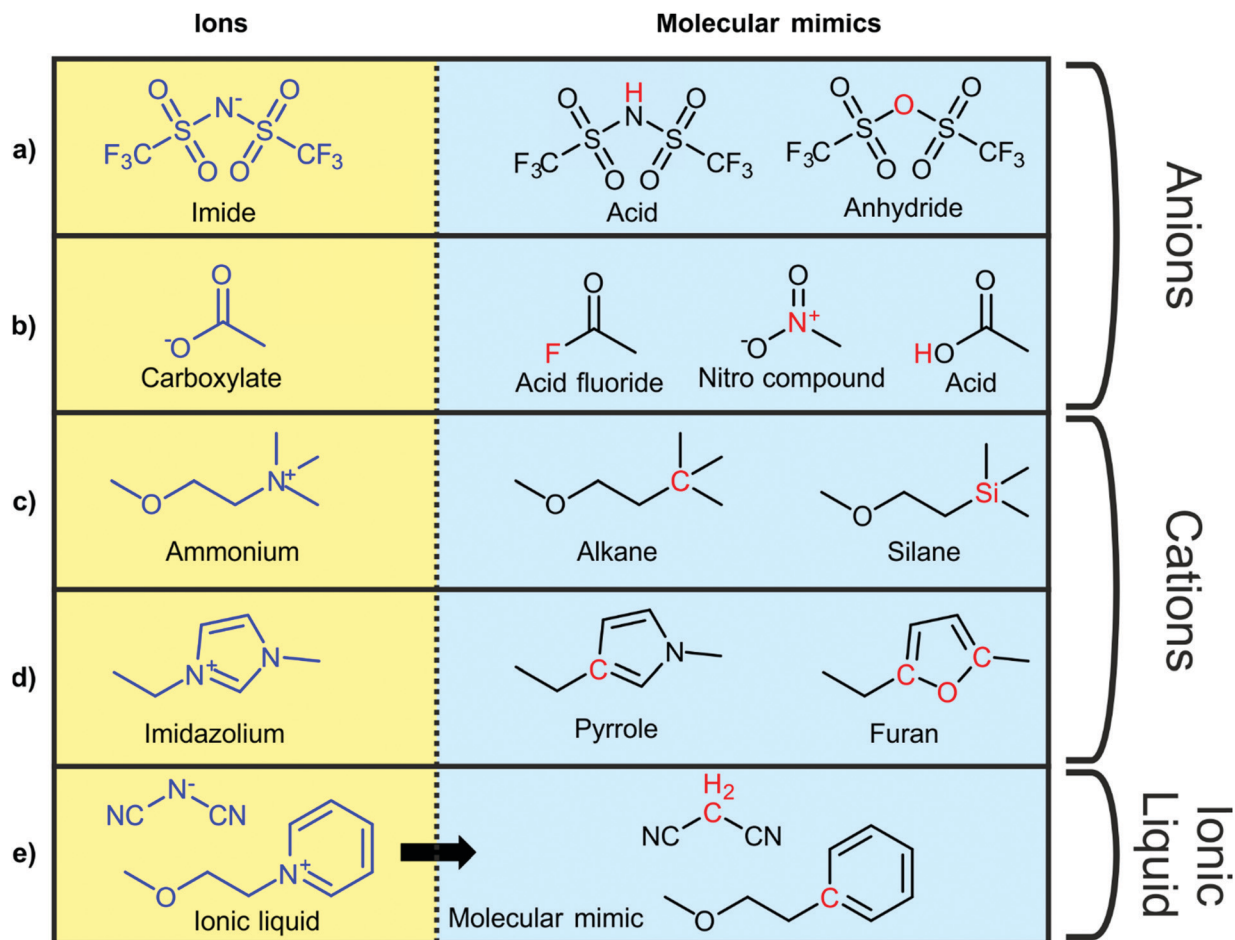


Fig. 3 Typical ionic liquid ions (left) and their targeted modifications to yield molecular mimics (right).

is best illustrated using the simple alkali halides as examples. Fig. 4 shows the solubility of different alkali halides in water in relation to the free enthalpy of hydration of the ions.<sup>96</sup> Salts composed of ions of similar size are poorly soluble and tend to associate in solution, whereas those of different size tend to

dissolve well.<sup>96,97</sup> The explanation of this general trend is not straightforward even for alkali metal halides and involves an interplay of electrostatics, entropic contributions and dispersion forces.<sup>96–98</sup>

It is challenging to separate the pure geometrical influence of size, since a change in size is coupled to changes in mass, dispersion interactions, functionalisation, *etc.* In general, larger ions reduce electrostatic interactions and introduce entropy into the melt, thus lowering the melting points of ionic liquids.<sup>99</sup> There are several approaches based on ion volumes that can be used to provide an educated guess as to whether a given ion combination will form an ionic liquid (at room temperature) or not.<sup>99,100</sup> Some cations, like phosphonium or ammonium, are benign in the regard that their size can be modified *via* the side chain length without changing the character of the ion significantly. However, especially for anions, small changes in the structure can lead to significant changes in specific interactions. For example, qualitative differences in hydrogen bonding and  $\pi$ - $\pi$  stacking (of the cation) were reported for the analogous series  $[\text{SCN}]^-$ ,  $[\text{N}(\text{CN})_2]^-$ , and  $[\text{B}(\text{CN})_4]^-$ , all paired with  $[\text{C}_2\text{C}_1\text{im}]^+$ .<sup>101</sup>

More so than size, the absolute mass can be modified experimentally by using isotopologues. Barthen *et al.* used this

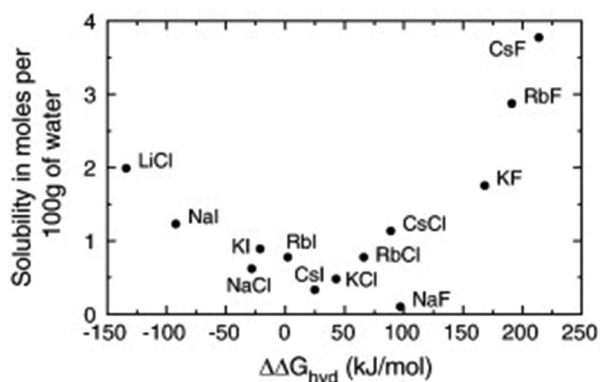


Fig. 4 The solubility of simple alkali halides in water can be rationalised with size differences. Reprinted from J. Gujt, M. Bešter-Rogač and B. Hribar-Lee, *J. Mol. Liq.*, **190**, 34–41, Copyright 2014, with permission from Elsevier.



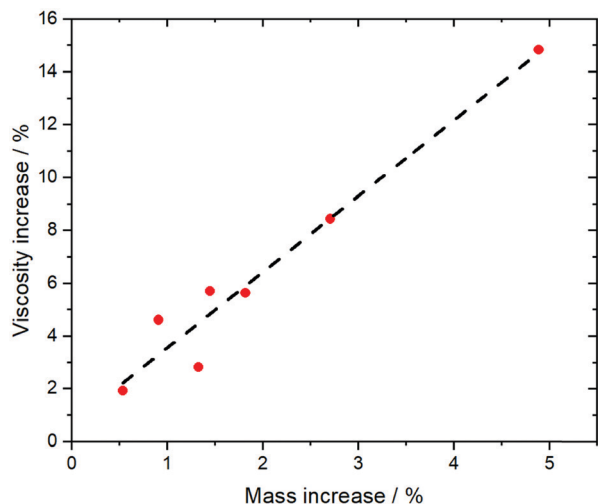


Fig. 5 The comparison of isotopologues with different mass reveals a correlation between viscosity and mass for a variety of different ionic liquids. Experimental data points from ref. 102, the dashed line is a linear fit.

concept to isolate the impact of ion mass on ionic liquids' viscosities.<sup>102</sup> To this end, they determined the change in viscosity for partial deuteration of 4 different ionic liquids, Fig. 5. A clear trend is visible despite the very small changes in mass. However, it must be kept in mind that the influence of mass on transport properties is often obscured by other effects. For example, the viscosity of  $[\text{C}_2\text{C}_1\text{im}][\text{NTf}_2]$  (1-ethyl-3-methylimidazolium bis(trifluoromethanesulfonyl)imide, 34 mPa s)<sup>103</sup> is considerably lower than that of  $[\text{C}_2\text{C}_1\text{im}][\text{NMs}_2]$  (1-ethyl-3-methylimidazolium bis(methanesulfonyl)imide, 787 mPa s),<sup>103</sup> despite the mass of the  $[\text{NTf}_2]^-$  ionic liquid being higher by 38%. The weakly coordinating properties of the fully fluorinated anion are in this case more important than size and mass.<sup>103,104</sup> Due to the greater basicity,  $[\text{NMs}_2]^-$  is a better hydrogen bond acceptor, thus forming stronger hydrogen bonds with the cation. In addition, the  $[\text{NTf}_2]^-$  anion has a much higher intrinsic flexibility than  $[\text{NMs}_2]^-$ ,<sup>105</sup> an effect that will be discussed in the next section.

A larger change in mass can be achieved with heavier analogues, *i.e.* by comparing valence isoelectronic cations or anions. However, a problem with this approach is that the change in mass is not isolated from other factors. A good example is the study of  $[\text{C}_4\text{C}_1\text{im}][\text{XF}_6]$  ionic liquids with  $\text{X} = \text{P}$ ,  $\text{As}$  or  $\text{Sb}$  by Shirota *et al.*<sup>106</sup> The authors observed that the viscosity decreased from 290 mPa s for  $[\text{C}_4\text{C}_1\text{im}][\text{PF}_6]$  to 134 mPa s for  $[\text{C}_4\text{C}_1\text{im}][\text{SbF}_6]$ . They attributed this decrease – despite the higher mass – to the larger volume of the anions. Another factor that might have contributed to the higher fluidity of the heavy analogues is the higher polarisability of the anions,<sup>107</sup> which we will discuss in the section ‘polarisability, ‘dipole moments’ and charge asymmetry’.

Notwithstanding the experimental difficulties, the influences of size and mass can be investigated in isolation with the help of simulations. To name an example related to the previous paragraph, the mass of the  $[\text{PF}_6]^-$  anion can be increased

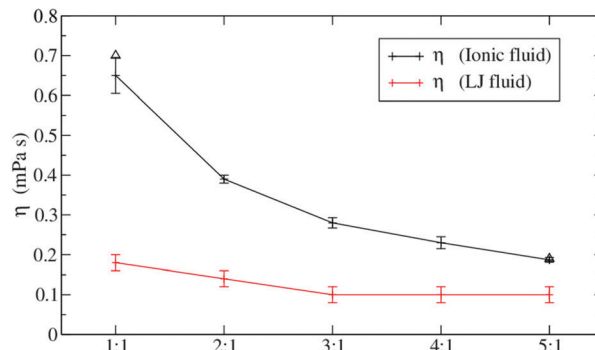


Fig. 6 Size differences can fluidize ionic liquids. The effect is less pronounced for the (uncharged) Lennard-Jones fluid. Reprinted from ref. 110, with the permission of AIP Publishing.

artificially in a simulation without changing any other parameter. In doing so, it was observed that the diffusion of  $[\text{PF}_6]^-$  in  $[\text{C}_2\text{C}_1\text{im}][\text{PF}_6]$  decreased as expected, apart from a small initial increase for a very small mass change.<sup>108</sup> The diffusion coefficient of the anion decreased by only 40% for a 20-fold increase in mass, which corresponds to an artificial molecular weight of the anion of 2900 Da.<sup>108</sup> It is immediately evident why such a small response to the mass of the anion is outweighed by other factors such as size, given that  $[\text{SbF}_6]^-$  is only 1.6 times heavier than  $[\text{PF}_6]^-$ . Similarly, Butler and Müller-Plathe used MD simulations of  $[\text{C}_4\text{C}_1\text{im}][\text{PF}_6]$  in which they artificially changed the ion volumes to demonstrate a clear relation between low viscosities and large volumes of both cation and anion.<sup>109</sup>

We will now focus on a selection of theoretical studies that use molecular dynamics simulations to investigate the influence of size (and mass) disparity. Spohr *et al.* compared simulations of charged and uncharged Lennard Jones fluids with two particle types of different size.<sup>110</sup> They observed that viscosity decreases (and diffusion increases) with greater difference in size, keeping other parameters constant, Fig. 6.<sup>110,111</sup> This effect was much more pronounced for the ionic liquid compared to the uncharged Lennard Jones fluid. The implications for ionic liquids are clear. Increased dynamics can be expected if ions of different size are used, but the effect will wear off eventually when another phenomenon such as dispersion interactions between long side chains begins to dominate the dynamics.

Structural effects are also evident in simulations of ionic liquids with significant size differences. A study of a binary Yukawa† ionic fluid with a size ratio of 5:1 of the two constituents showed that the larger particles crystallised upon cooling during the simulation.<sup>112</sup> Below this ‘freezing point’, the smaller particles remained mobile, showing regular diffusion at higher temperature and a hopping mechanism at lower temperature.<sup>112</sup> Such curious phase behaviours find application in solid state materials, such as ionic organic plastic crystals.<sup>113–115</sup> Although Spohr *et al.* did not encounter such a

† The Yukawa potential is a screened Coulomb potential, *i.e.* it includes an exponential decay on top of the  $r^{-1}$  dependence.



crystallisation of the larger particles during their simulation, they observed that the liquid structure is dominated by the larger particles.<sup>110</sup> The combined effects of size and charge displacements will be discussed in the section ‘polarisability, ‘Dipole moments’ and charge asymmetry’.

Ionic liquids are strongly correlated systems, and these (cross) correlations are further complicated by dependencies on size and mass. Non-negligible cross correlations are widely accepted as the reason for deviations from the Nernst–Einstein relation in ionic liquids, assuming integer charges on ions.<sup>91,116,117</sup> However, different – and not necessarily mutually exclusive – interpretations of these deviations exist, such as ion aggregation or the presence of dynamical heterogeneities.<sup>110,118–123</sup> In terms of the time evolution of the velocity correlation functions, both effects of relative mass and relative size have been found to be important, even for neutral binary liquids.<sup>124,125</sup> Lighter and smaller particles show pronounced backscattering, an effect which is also caused by the shell structure in ionic liquids, with ions rattling in a cage of counterions.<sup>88,126</sup>

With the design of ionic liquids in mind, it would be interesting to investigate the magnitude of cross contributions for systems similar to the ones of Butler and Müller-Plathe<sup>109</sup> as well as Spohr *et al.*<sup>110</sup> A central issue with the interpretation of distinct diffusion coefficients is their reference frame dependence, which can be avoided by using quantities such as the relative mean molecular velocity cross correlation coefficients that do not show a reference frame dependence.<sup>126–129</sup> Finally, technical details such as ion size dependency<sup>123</sup> or slow convergence of the cross contributions<sup>125</sup> need to be included in any investigation of these effects.

## Conformations and conformational flexibility

The vast majority of ionic liquids are composed of ions that can exist in different conformations. Flexible groups such as (alkyl) side chains in cations are ubiquitous design elements used to

lower melting points and accelerate the dynamics in ionic liquids.<sup>31,74,99,130–133</sup> However, side chains of different length are of limited use to probe the influence of conformational equilibria due to the competing effects of size and mass as well as dispersion interactions. We will thus focus on isostructural analogues which differ markedly in their conformational properties, but not in their size. We will first present examples involving cations (Fig. 7), followed by those involving anions (Fig. 10).

Phosphonium and ammonium ionic liquids offer this similarity, with a wide variety of ionic liquid precursors and even ionic liquids readily available to the synthetic chemist. Phosphonium ionic liquids are in general less viscous than their ammonium counterparts, which is one of the many advantages that have drawn attention to the class of phosphonium ionic liquids.<sup>134–137</sup> Fig. 7a shows typical examples of ionic liquids based on triethylphosphine and triethylamine.<sup>138</sup> For the five pairs of phosphonium/ammonium functionalised ionic liquids, the ammonium ionic liquids are more viscous by a factor of 1.7 to 2.0.<sup>138</sup> The higher conformational flexibility of phosphonium ionic liquids has been identified as the root cause for this difference between phosphonium and ammonium ionic liquids.<sup>139–141</sup> Of central importance appears to be the C–P–C–C (C–N–C–C) dihedral angle corresponding to rotation around the P–C (N–C) bonds. The *ab initio* energy barrier for this rotation is found to be approximately 10 kJ mol<sup>−1</sup> higher for ammonium ions compared to phosphonium ions.<sup>139,140</sup>

In addition to the higher conformational flexibility of the phosphonium cations, it has also been reported that the positive charge is more localised on the phosphonium centre.<sup>139,142</sup> Scarbath-Evers *et al.* used molecular dynamics simulations to differentiate between these effects and observed that the localised charge in phosphonium ionic liquids acts in the opposite direction to the higher conformational flexibility, decreasing the dynamics.<sup>139</sup> Thus, given the current state of knowledge, conformational flexibility is the dominant cause for the improved dynamics of phosphonium based ionic liquids compared with ammonium based ionic liquids.

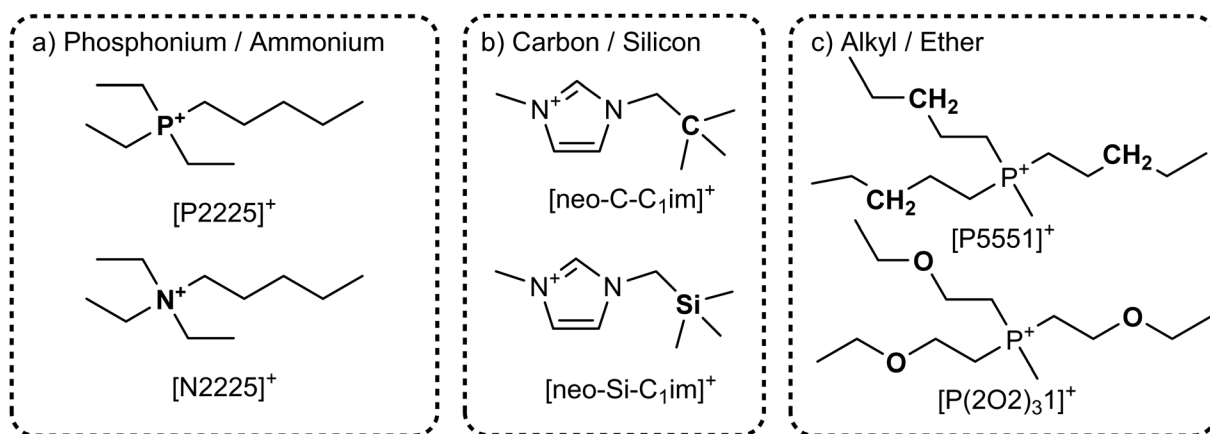


Fig. 7 Isostructural cation analogues serve as a targeted modification to probe the effects of conformation and conformational flexibility. All ionic liquids have [NTf<sub>2</sub>]<sup>−</sup> as anion.<sup>138,140,143</sup>



Shirota and Castner applied the concept of isostructural analogues to the side chain of imidazolium ionic liquids, Fig. 7b.<sup>143</sup> This approach is very similar to the comparison of phosphonium and ammonium ionic liquids; the quaternary carbon is replaced by the heavier silicon. The resemblance to the case of phosphonium *vs.* ammonium is striking: the viscosities were 98 mPa s for [neo-Si-C<sub>1</sub>im][NTf<sub>2</sub>]<sup>-</sup> and 161 mPa s for [neo-C-C<sub>1</sub>im][NTf<sub>2</sub>]<sup>-</sup>, *i.e.* the viscosity of the lighter analogue is higher by a factor of 1.6.<sup>143</sup> Again, this effect is attributed to the higher conformational flexibility of silicon compounds.<sup>74,144</sup>

The higher flexibility of the silicon analogues can be shown with a straightforward calculation of the potential energy as a function of the dihedral angles defining the ion geometry, see Fig. 8 and ESI.† The barriers for rotation around the N–C and C–E bonds are higher for the lighter element. The energy difference between barrier heights is approximately 10 kJ mol<sup>-1</sup>, in line with the observations for phosphonium *vs.* ammonium. Furthermore, the barriers for the two analogues are well separated by the 15 kJ mol<sup>-1</sup> line, which we have introduced to identify rigid (barrier above the threshold) and flexible (barrier below the threshold) ions.<sup>105,140</sup>

Ether groups are well-known as a design element to improve the properties of ionic liquids, in particular to achieve lower viscosities.<sup>145–147</sup> Taking the example in Fig. 7c, the viscosity of the purely alkyl functionalised ionic liquid [P5551][NTf<sub>2</sub>]<sup>-</sup> is a factor of 6 higher than the viscosity of the ether functionalised ionic liquid [P(2O2)<sub>3</sub>1][NTf<sub>2</sub>]<sup>-</sup>.<sup>140</sup> A number of thorough studies can be found that report similar effects for a variety of other cation structures such as ammonium (including cyclic), pyridinium, and imidazolium cations.<sup>147–150</sup>

Similar to phosphonium or silicon functionalisation, ether groups are another structural motif whose beneficial properties

are often attributed to their high conformational flexibility.<sup>151–154</sup> However, for ether groups, this is only one side of the coin. Equally important to the conformational flexibility is the preferred conformer geometry. Several groups have reported evidence that ether side chains tend to coil up towards the positively charged part of a cation, Fig. 9.<sup>140,155–158</sup> Similarly, for imidazolium ionic liquids, ether groups in a (sufficiently long) side chain tend to coordinate to hydrogen atoms of the imidazolium ring, with both intra- and intermolecular interactions observed.<sup>159,160</sup> The presence of these two overlapping phenomena, *i.e.* changes in both flexibility and geometry, lead to the rich diversity of ether functionalised ionic liquids, but also to a somewhat irregular behaviour.

The concept of conformational flexibility is equally important for anions, the best example being [NTf<sub>2</sub>]<sup>-</sup>, introduced as an anion for ionic liquids by Pierre Bonhôte in 1996.<sup>161</sup> The [NTf<sub>2</sub>]<sup>-</sup> anion can exist in two different conformations, *cis* and *trans*.<sup>162</sup> These two conformers are close in energy (difference  $\approx kT$  at 25 °C), and separated by a small barrier ( $\approx 2kT$  at 25 °C).<sup>162</sup> The conformational equilibrium in [NTf<sub>2</sub>]<sup>-</sup> and its relation to macroscopic properties is well documented by both theoretical<sup>105,163,164</sup> and experimental<sup>165–172</sup> approaches.

The interconversion between the *cis* and *trans* conformers can be prevented experimentally by formally connecting the trifluoromethyl end groups to form a cyclic anion such as [NCyF]<sup>-</sup>, Fig. 10a. The [NCyF]<sup>-</sup> anion is an analogue of [NTf<sub>2</sub>]<sup>-</sup> which is effectively locked into a *cis* conformation, thus lacking the flexibility of [NTf<sub>2</sub>]<sup>-</sup>. Some internal structural freedom – comparable to cyclohexane – can be expected, however reorganisation of one end group relative to the other is naturally not possible. The effect of this increased rigidity is evident in the physical properties of ionic liquids containing this anion.



Fig. 8 Energy as a function of the dihedral angles corresponding to rotation around the indicated bonds. Both the N–C and C–E bond rotations benefit from substitution of carbon with silicon; thus reducing the energy barriers by approximately 10 kJ mol<sup>-1</sup>.





Fig. 9 Cations with alkyl side chains (a) prefer a linear conformation, whereas those with ether side chains (b) coil.<sup>140</sup> (c) For both ammonium and phosphonium ionic liquids, the distance between terminal carbon and quaternary centre is reduced considerably when ether groups are introduced in the side chain.<sup>156</sup> (a) and (b) adapted from ref. 140 – Published by the PCCP Owner Societies. (c) Reprinted with permission from ref. 156. Copyright 2013 American Chemical Society.

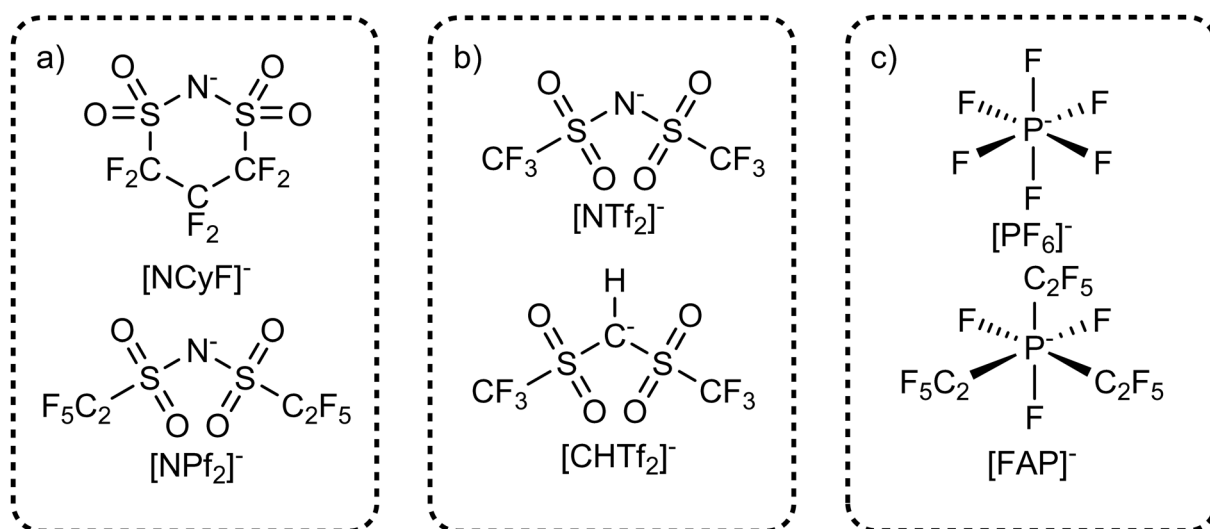


Fig. 10 Targeted modifications which influence the conformational flexibility via (a) ring closure, (b) degrees of freedom for hybridization, (c) introducing flexible groups.

A viscosity of 174 mPa s has been reported for  $[C_4C_1im][NCyF]$ , whereas the viscosity of  $[C_4C_1im][NTf_2]$  is only 48 mPa s.<sup>173</sup> The higher viscosity of  $[C_4C_1im][NCyF]$  is not caused by its higher mass compared to  $[C_4C_1im][NTf_2]$ , which becomes clear when comparing  $[C_4C_1im][NCyF]$  with the heavier  $[C_4C_1im][NPF_2]$  ( $NPF_2 = N(SO_2C_2F_5)_2$ , also known as BETI). The viscosity of  $[C_4C_1im][NPF_2]$  is 104 mPa s, significantly lower than that of  $[C_4C_1im][NCyF]$ . The impact of the rigidity of the cyclic anion on macroscopic transport has also been shown for  $[C_2C_1pyrr][NCyF]$ , specifically for the ionic conductivity.<sup>174</sup> Compared to  $[C_2C_1pyrr][NPF_2]$ ,  $[C_2C_1pyrr][NCyF]$  is more conductive by a factor of 83.<sup>174</sup>

A disadvantage of using cyclic analogues is that only the 'cis-like' conformer is present; thus the observed effects might

also be a characteristic of said conformer. For example, it is well known that the 'dipole moment' of *cis*- $[NTf_2]^-$  is higher than that of *trans*- $[NTf_2]^-$ .<sup>175,176</sup> Studies using cyclic analogues are a valuable and needed addition to those with the common  $[NTf_2]^-$  anion and bear the potential to reveal both the influence of conformational flexibility and the characteristic effects of a particular conformer – provided that a separation of those contributions can be made. Future studies might also include the five- and even the four-membered ring, both of which are stable anions that are yet to be used in ionic liquids.<sup>177,178</sup>

§ The dipole moment of charged species is reference frame dependent, see the section 'Polarisability, 'dipole moments' and charge asymmetry'.





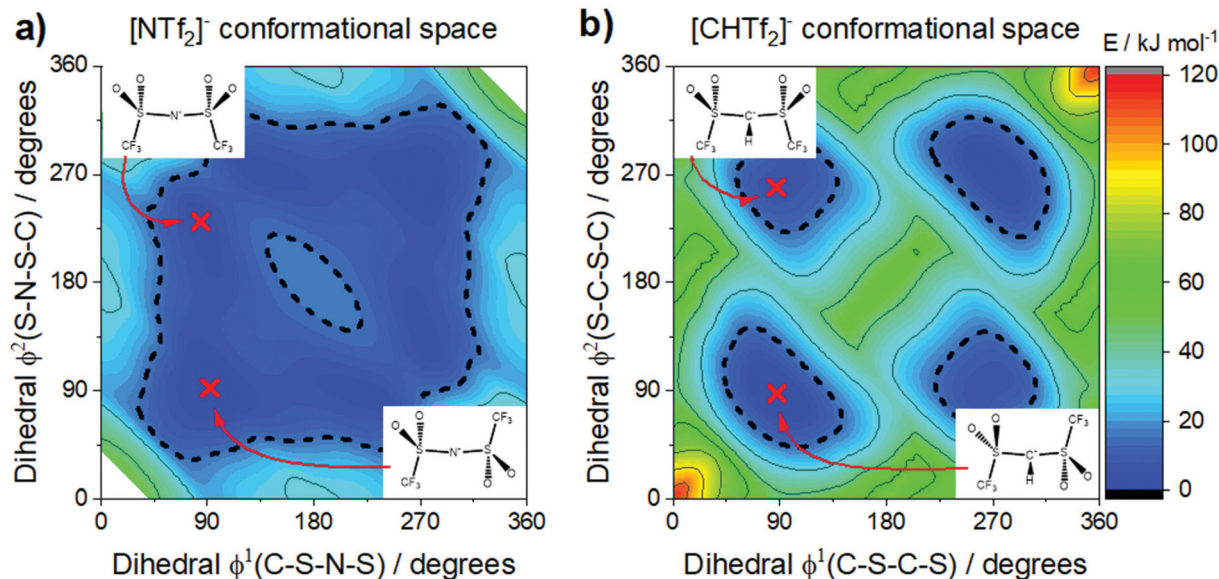


Fig. 11 The  $[\text{NTf}_2]^-$  anion (a) is extremely flexible, easily changing between *cis* and *trans* conformers. In contrast, replacing the imide nitrogen with a methanide group leads to the rigid  $[\text{CHTf}_2]^-$  anion (b), which experiences a high barrier for *cis*–*trans* interconversion. (a) and (b) adapted from ref. 140 – Published by the PCCP Owner Societies.

The conformational flexibility of  $[\text{NTf}_2]^-$  is sensitive to seemingly small changes in its structure. The electronic structure of  $[\text{NTf}_2]^-$  allows for rehybridisation involving the central nitrogen atom (imide group), which leads to the high geometrical flexibility of this ion.<sup>105</sup> Thus, the flexibility is lost when the central nitrogen atom (imide) is replaced with a methanide group, Fig. 10b. The effect is clearly seen in the potential energy surface, Fig. 11, and can be used to experimentally demonstrate the difference between rigid and flexible anions.<sup>105</sup> The  $[\text{NTf}_2]^-$  anion, Fig. 11a, has a flat potential energy surface where *cis* and *trans* can easily interconvert to each other. The dashed line drawn at  $15 \text{ kJ mol}^{-1}$  encompasses minima and paths connecting these minima in a continuous basin. In contrast, the potential energy surface of the  $[\text{CHTf}_2]^-$  anion in Fig. 11b shows minima for *cis* and *trans* which are separated by large barriers (compared to the thermal energy). However, both conformers – *cis* and *trans* – are still present, which is a clear difference over the use of cyclic analogues.

The structural variation between  $[\text{NTf}_2]^-$  and  $[\text{CHTf}_2]^-$  is minimal, the degree of fluorination is the same, and the change in mass and size is negligible. Thus, changes in the macroscopic properties can be rationalised with the higher conformational flexibility of  $[\text{NTf}_2]^-$  compared to the rigid analogue,  $[\text{CHTf}_2]^-$ . The impact on the viscosities of the respective ionic liquids with the  $[\text{P5551}]^+$  cation is shown in Fig. 12a.<sup>140</sup> At room temperature, the viscosity of the ionic liquid with the rigid  $[\text{CHTf}_2]^-$  anion is almost twice as high as the viscosity of  $[\text{P5551}][\text{NTf}_2]$ , similar to the relative viscosities of  $[\text{C}_4\text{C}_1\text{im}][\text{NCyF}]$  and  $[\text{C}_4\text{C}_1\text{im}][\text{NTf}_2]$ . The effect is also observed for other anions of similar structure, for example comparing the ionic liquid  $[\text{C}_4\text{C}_1\text{im}][\text{NMsTFA}]$  (rigid anion, viscosity  $292 \text{ mPa s}$ ) to  $[\text{C}_4\text{C}_1\text{im}][\text{NTfAc}]$  (flexible anion, viscosity  $149 \text{ mPa s}$ ),<sup>105</sup> it is thus not limited to the methanide motif.

The influence of conformational flexibility decreases with increasing temperature, which is an important observation.

The factor by which  $[\text{P5551}][\text{CHTf}_2]$  is more viscous than  $[\text{P5551}][\text{NTf}_2]$  decreases from 1.9 (at  $25^\circ\text{C}$ ) to 1.3 (at  $100^\circ\text{C}$ ).<sup>140</sup> The higher the temperature, the more thermal energy is available to enable more and more mechanisms of structural relaxation.<sup>179,180</sup> Basin theory provides the framework for this rationalisation of viscous flow using the intuitive concept of the potential energy landscape.<sup>179–182</sup> The potential energy surface is the potential energy as a function of the degrees of freedom of the system. It is thus a rugged high-dimensional surface. Viscous flow, *i.e.* structural relaxation, requires the bulk system to pass from one preferred structure to another. The possible preferred structures are connected by different paths involving barriers of different height, see the ‘thermally activated’ process in Fig. 12b. Higher temperatures allow the system to explore a greater proportion (of the high energy region) of the potential energy surface. At lower temperatures, other pathways might still be feasible, see the ‘entropic’ process in Fig. 12b. Thus, conformational flexibility can be seen as a ‘last resort’ to enable viscous flow of a glass-forming ionic liquid at low temperatures before the systems becomes non-ergodic, and is an important consideration for applications around room temperature.<sup>171,183</sup>

Effects of size, shape, and flexibility are necessarily connected. A large, rigid ion needs a sufficiently large gap to pass through when moving through the cage of counterions in the bulk. Depending on the shape of the ion, a relatively small gap can be passed through if the orientation is right. Thus, it is easier for a prolate (rugby ball-shaped) ion to move along its axis, Fig. 13a, whereas movement perpendicular to this axis exposes a larger surface area and thus provides more resistance. Hole theory considers the static size of ions in relation to the probability to find a sufficiently large hole, thus offering quantitative prediction of dynamic properties.<sup>184–186</sup> On the other hand, a flexible ion





Fig. 12 (a) The ionic liquid with the more rigid anion, [P5551][CHTf<sub>2</sub>], has an almost twofold viscosity at room temperature compared to the ionic liquid with flexible anion, [P5551][NTf<sub>2</sub>].<sup>140</sup> (b) The viscosity difference decreases with increasing temperature, which can be explained with basin theory, as more and more pathways become thermally activated.<sup>180,378</sup>

with the same size and shape can sample its surroundings and dynamically adapt to available gaps, Fig. 13b. This molecular level reasoning is the basis for the 'slithering' mechanism of diffusion of [NTf<sub>2</sub>]<sup>-</sup> like anions.<sup>163,165,166</sup> The entropic process in Fig. 12b thus corresponds to the flexible ion in Fig. 13. When [CHTf<sub>2</sub>]<sup>-</sup> is used instead of [NTf<sub>2</sub>]<sup>-</sup>, the potential energy surface will be qualitatively similar, but the barriers are much higher – thus the previously entropic processes involving conformational changes become thermally activated.

Naturally, experimental approaches are limited by synthetic accessibility, while theoretical methods can and should be used to fill the gaps. An elegant and straightforward way to investigate the influence of conformational flexibility in isolation is to change the barrier height for interconversion between conformers artificially in an MD simulation. Increasing the *cis-trans* barrier in [C<sub>2</sub>C<sub>1</sub>im][NTf<sub>2</sub>], above 15 kJ mol<sup>-1</sup>, led to a decrease in diffusion by 40–50%.<sup>163</sup> This change in mobility agrees well with the experimental findings presented above.

Another important example of the use of conformational flexibility is the [FAP]<sup>-</sup> anion, a hydrolytically and electrochemically more stable analogue of the simpler [PF<sub>6</sub>]<sup>-</sup>, Fig. 10c.<sup>187</sup> The use

of [FAP]<sup>-</sup> instead of [PF<sub>6</sub>]<sup>-</sup> considerably lowers the viscosity of ionic liquids. For example, the viscosity of [C<sub>6</sub>C<sub>1</sub>im][PF<sub>6</sub>] is 548 mPa s, whereas that of [C<sub>6</sub>C<sub>1</sub>im][FAP] is only 74 mPa s – despite the much higher volume of the latter.<sup>188</sup> Formally, this change translates to an increase in fluidity by a factor of approximately two per pentafluoroethyl side chain, again in line with results from other systems. A crucial difference between these two anions is that the [PF<sub>6</sub>]<sup>-</sup> anion is relatively rigid, whereas the pentafluoroethyl side chains in the [FAP]<sup>-</sup> anion can rotate around the P–C bonds, leading to increased conformational flexibility. However, other parameters such as size and mass, fluorination, coordination tendency, and symmetry change as well going from [PF<sub>6</sub>]<sup>-</sup> to [FAP]<sup>-</sup>.<sup>189</sup> Thus, this comparison alone would not be a convincing argument for conformational flexibility. Experimentally, it would be beneficial to have at hand the [P(CF<sub>3</sub>)<sub>6</sub>]<sup>-</sup> anion, which will have a similar mass and coordination tendency compared to [FAP]<sup>-</sup>. However, the syntheses of such compounds are involved, and the [P(CF<sub>3</sub>)<sub>6</sub>]<sup>-</sup> anion might not even be sufficiently stable.<sup>187,190</sup>

Notwithstanding experimental obstacles, the contribution of conformational flexibility within the [FAP]<sup>-</sup> anion can be

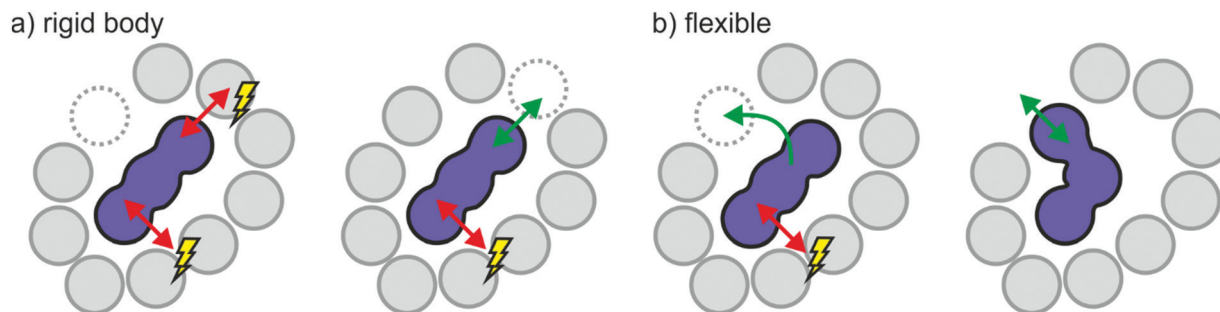


Fig. 13 Schematic representation of a prolate ion (blue) surrounded by a cage of counterions (grey). The dotted area represents an available gap in the cage. (a) The rigid ion can only move out of the cage when oriented correctly, whereas (b) the flexible ion can react dynamically to a change in the surroundings.





Fig. 14 Diffusion coefficients of  $[C_4C_1im][PF_6]$  and  $[C_4C_1im][FAP]$  obtained from molecular dynamics simulations.<sup>191</sup> The simulations were set up with scaled vs. unit charges and flexible vs. rigid anions. Arrows indicate the increase in diffusion when the FAP anion is allowed to be flexible, regardless if charges are scaled or not.

isolated by means of MD simulations. Kowsari and Ebrahimi performed such simulations, comparing  $[C_4C_1im][FAP]$  with  $[C_4C_1im][PF_6]$ . The authors determined a variety of dynamic properties from classical molecular dynamics simulations, including a simulation of  $[C_4C_1im][FAP]$  where the anion was simulated as a rigid body using QSHAKE constraints.<sup>191</sup> The diffusion coefficients from the mean squared displacement, Fig. 14, clearly show the acceleration of translational dynamics of both cation and anion when the  $[FAP]^-$  anion is allowed to be flexible. The diffusion of ions in  $[C_4C_1im][FAP]$  with a rigid anion is comparable to or even slower than in  $[C_4C_1im][PF_6]$ .

The experimental increases in fluidity at room temperature discussed in this section are summarised in Table 1. The importance of conformations and conformational flexibility is clearly visible, even though some systems such as  $[C_4C_1im][NCyF]/[C_4C_1im][NTf_2]$  and  $[C_6C_1im][PF_6]/[C_6C_1im][FAP]$  might also have other contributions. More work needs to be done to explore the conformational space of  $[FAP]^-$  and similar anions both experimentally and computationally. Furthermore, the concepts can be easily extended to mixed fluoroborates, which

Table 1 Summary of targeted modifications that aim to isolate the influence of conformations and conformational flexibility. For cases where more than one group is changed, brackets give the increase per motif

Rigid	Flexible	Fluidity increase $\eta_{rigid}/\eta_{flexible}$	Ref.
$[N2225][NTf_2]$	$[P2225][NTf_2]$	1.95	138
$[neo-C-C_1im][NTf_2]$	$[neo-Si-C_1im][NTf_2]$	1.64	143
$[P5551][NTf_2]$	$[P(2O)_3]_1[NTf_2]$	6.28 (2.09)	140
$[C_4C_1im][NCyF]$	$[C_4C_1im][NTf_2]$	3.63	173
$[C_4C_1im][NCyF]$	$[C_4C_1im][NPF_2]$	1.67	173
$[P5551][CHTf_2]$	$[P5551][NTf_2]$	1.89	140
$[C_4C_1im][NMsTFA]$	$[C_4C_1im][NTfAc]$	1.96	105
$[C_6C_1im][PF_6]$	$[C_6C_1im][FAP]$	7.41 (2.47)	188

are an exciting new platform for low-viscosity ionic liquids.<sup>192,193</sup> Finally, since the effects are most visible at low temperatures (*i.e.* room temperature) for ionic liquids with relatively high absolute viscosity values, design concepts such as conformational flexibility could enable fluorine-free ionic liquids of sufficient fluidity.

## Polarisability, 'dipole moments' and charge asymmetry

We have discussed earlier the effects of Coulombic compaction and how they might be addressed in a future study using targeted modifications. On top of this overall compaction, the charged nature of ionic liquids also produces strong local electric fields, which affect the ions in two ways. First, the electron cloud of a molecule (including ions and atoms) will be distorted more or less by the presence of an electric field. This inherent property of matter is termed polarisability, and in molecules the electric field leads to an induced dipole moment. Second, the charge distribution in a molecule can be asymmetric even without the presence of an electric field. For neutral molecules, this leads to a permanent dipole moment; for ions, which are formally monopoles, concepts such as the charge arm or charge lever moment are preferred.<sup>49,194,195</sup>

Polarisability in ionic liquids has been a key area of research in the molecular dynamics community over recent years.<sup>63,196–204</sup> The comparison of polarisable with non-polarisable simulations is an excellent example of how targeted modifications are used to understand which components define the properties of ionic liquids. However, care must be taken even for simulations to not (unknowingly) change more than one parameter at once. For example, simply adding polarisability to an existing non-polarisable force field will lead to biased results, as inductive effects already included in the original force field are then treated twice.<sup>205</sup> Polarisability in molecular dynamics simulations of ionic liquids is a very active area of research, and a thorough review has been published recently.<sup>202</sup> Thus, we will not go into detail, but rather highlight two general results. First, the long-range structuring decreases, and the order beyond the first shell is less pronounced in polarisable models.<sup>63,199,200,202,206</sup> Thus, screening is much more efficient if the ions are polarisable.<sup>63</sup> Second, the dynamics are significantly accelerated, and predicted transport properties are in much better agreement with experimental values.<sup>196,206,207</sup>

The tendency of polarisability to enhance the dynamics in ionic liquids should be kept in mind during the design process, *i.e.* when an ionic liquid is conceptually composed from functional groups and atoms of a given polarisability.<sup>208–210</sup> However, the experimental possibilities are very limited, at least for targeted modification, since volume and polarisability are closely related.<sup>55,211,212</sup> Examples are the aforementioned studies of the series  $[C_4C_1im][XF_6]$  ( $X = P, As$  or  $Sb$ ), where a substantial increase in dynamics has been observed for the heavier central elements.<sup>106,107</sup> Although the effect is very clear, it cannot be attributed to either size or polarisability without the help of theoretical methods, since both change together.



Experimentally more accessible than the effects of polarisability are those of an asymmetric charge distribution, *i.e.* those related to the presence of a permanent 'dipole moment'. In principle, the dipole moment of a molecule can be obtained from its wavefunction, and such calculations are indeed common for ionic liquids.<sup>170,176,213,214</sup> The unperturbed dipole moment  $\vec{\mu}$  of a molecule as the sum of its nuclear  $\vec{\mu}_N$  and electronic  $\vec{\mu}_e^0$  contributions is given in eqn (1) and (2).<sup>215,216</sup>

$$\vec{\mu}_N = \sum_i^N z_i \vec{r}_i \quad (1)$$

$$\vec{\mu}_e^0 = -\langle \psi | \vec{r} | \psi \rangle = -\int_{\mathbb{R}^3} \vec{r} \rho(\vec{r}) d\vec{r} \quad (2)$$

The sum runs over the nuclei  $N$ , where  $z_i$  are the nuclear charges. The wavefunction  $\psi$  is usually obtained as a Slater determinant of Gaussian basis functions, so  $\langle \psi | \vec{r} | \psi \rangle$  can be evaluated analytically. The last identity arises as the position operator can be separated since it does not act on the wavefunction, thus allowing the latter to be interpreted as probability (electron)  $\rho(\vec{r})$  density according to Born's rule. However for molecular ions, the dipole moment becomes ill defined, as only the non-vanishing multipole moment is independent of the choice of coordinate system.<sup>195,215</sup> Hence, we will use the 'charge arm'  $L_c = |Ql_q|$  as defined by Kobrak *et al.* for this discussion.<sup>49</sup> The charge arm is proportional to the total charge  $Q$  of the ion and the distance  $l_q$  between centre of charge  $\vec{R}_{cq}$  and centre of mass  $\vec{R}_{cm}$ , (3)–(5).<sup>49</sup>

$$\vec{l}_q = \vec{R}_{cq} - \vec{R}_{cm} \quad (3)$$

$$\vec{R}_{cm} = \frac{1}{M} \sum_i^N m_i \vec{r}_i \quad (4)$$

$$\vec{R}_{cq} = \frac{1}{Q} \sum_i^N q_i \vec{r}_i \quad (5)$$

$M$  is the mass of the ion,  $m_i$  is the mass of atom  $i$ . The similarity of  $\vec{R}_{cq}$  and  $\vec{\mu}$  is evident;  $\vec{R}_{cq} = \vec{\mu}/Q$ . The situation for a pair of spherical ions is shown in Fig. 15. The Coulombic attraction (or repulsion, for like charges) acts on the centre of charge. Thus, if  $\vec{R}_{cq}$  and  $\vec{R}_{cm}$  don't coincide, the Coulomb force translates to a torque  $\tau$  on the ion. The torque is proportional to  $L_c$ , but the angular acceleration  $\alpha$  also depends on the mass – or, more precisely, on the moment of inertia  $I$ , (6).<sup>¶</sup> Thus, the librational coupling of an ion with asymmetric charge distribution varies with the length of the charge arm  $L_c$ , the mass of the ion  $M$ , and the radius of gyration  $R_g$ . Li *et al.* use the term 'charge lever moment' for  $Z$ , to distinguish it from the charge arm  $L_c$ .<sup>194</sup>

$$\alpha = \frac{\tau}{I} \propto \frac{L_c}{MR_g^2} = Z \quad (6)$$

The simplicity of the model in Fig. 15 comes with limitations that have to be kept in mind. First, higher order multipole interactions are neglected.<sup>195</sup> Second, it relies on spherical, rigid

¶ Equations are simplified for the case of a spherical ion, for a force perpendicular to  $\vec{l}_q$ . For a more detailed discussion, see ref. 194.

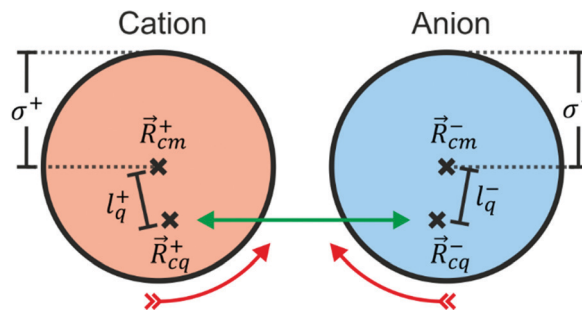


Fig. 15 Schematic representation of the construction of a model system with displaced charge.

ions.<sup>194</sup> Nevertheless, the charge arm concept captures important features of ionic liquids, and at least a qualitative correlation is found with experimental viscosities of ionic liquids based on rigid ions.<sup>49,194</sup>

Molecular dynamics simulations allow the direct implementation of the simple model system shown in Fig. 15. These targeted modifications are similar to those discussed in the section 'size and mass of ions', *i.e.* they are usually realised using Lennard-Jones or hard sphere fluid models, with the only difference that the (point) charge is displaced from the (point) mass by the distance  $l_q$ . Charge displacement models consistently show a number of features:

- With increasing charge displacement, initially the viscosity is found to decrease, and conductivity and translational diffusion increase.<sup>111,217</sup>
- Rotational diffusion slows down with increasing charge displacement, indicative of directional bonding.<sup>217,218</sup>
- For large  $l_q/\sigma$ , ion pairing to form a dipolar liquid occurs.<sup>111,217,219</sup>
- Thus, for too large  $l_q/\sigma$ , the dynamics slow down again – fluidity, conductivity, diffusion decrease.<sup>217</sup>
- Charge displacement in general lowers melting points; see the discussion of Fig. 18.<sup>219–221</sup>

The effect of charge displacement on dynamic properties is shown in Fig. 16 (black curves).<sup>111</sup> There is an initial drop in viscosity for increasing  $l_q$ , accompanied by an increase in diffusion. However, if the charge displacement becomes too large – *i.e.* larger than 50% of the ion radius in this case – the viscosity shows an abrupt increase. This increase can be attributed to the onset of ion pairing, which is also evident in the convergent diffusion coefficients of cation and anion.<sup>111</sup> For comparison, the red curves show the behaviour of charge centred model systems with the same charge separation. It is evident that without the asymmetry, a monotonous decrease in the dynamics is found, *i.e.* viscosities increase and diffusion coefficients decrease.<sup>111</sup>

Li and Kobrak demonstrated that the general trends observed in the simulation of Lennard-Jones and hard sphere

|| Thus, for the charge centred system,  $\vec{l}_q^- = \vec{l}_q^+ = 0$ . The distance between  $\vec{R}_{cq}$  (given by  $\sigma^+ + \sigma^-$ ) for the charge centred system is set equal to the minimum possible charge separation in the asymmetric charge system (given by  $\sigma^+ + \sigma^- - \vec{l}_q^+$ ).





Fig. 16 Diffusion and viscosity as a function of charge separation. The blue shaded area indicates the  $l_q^+/\sigma^+$  range for the  $[C_4C_{1im}]$  cation. The charge centred systems where charge separation is achieved by decreasing the size are shown in red. Reprinted (adapted) from H. V. Spohr and G. N. Patey, *J. Chem. Phys.*, 2009, **130**, 104506, with the permission of AIP Publishing.<sup>217</sup>

models also apply to a more complex model based on the  $[PF_6]^-$  anion.<sup>222</sup> The authors systematically varied the charges of a generic classical force field in two ways, Fig. 17. First, they shifted a small proportion of the charge – 0.10e per fluorine atom, *i.e.*  $\approx 30\%$  of the native charge – from one side of the anion to the other. Second, they placed the charges on dummy atoms between phosphorus and fluorine, leaving fluorine uncharged. Thus, the charges were buried/more localised in comparison to the native force field. For the combination of the two targeted modifications, the amount of charge shifted from one half of the dummy atoms to the others was increased to 0.30e to achieve the same charge lever moment for both asymmetric cases. The phosphorus atom was left unchanged, and an identical structure with reversed charges was used as counterion.

Two dynamical properties are of central interest in the model system in Fig. 17, the translational and rotational diffusion. Both translation and rotation are significantly accelerated for the accessible charge compared to the case where the charge is buried on the dummy atoms.<sup>222</sup> However, opposite trends in translational and rotational dynamics were reported for the asymmetric models. The translational diffusion coefficients increased significantly for the asymmetric charge models – by a factor of 12 for the accessible charge, and by a factor of 52 for the buried charge.<sup>222</sup> In contrast, the rotational relaxation time almost doubled for the asymmetric charge models compared with their symmetric counterpart.<sup>222</sup>

These results indicate that the beneficial effect of an asymmetric charge distribution arises from the coupling between rotational (usually fast in liquids) and translational (usually slower in liquids) modes. It can be expected that this mechanism of structural relaxation is particularly important for relatively rigid, close to spherical ions. The relative charge arm  $l_q/\sigma$  of the  $[PF_6]^-$  model with artificial charge asymmetry is in the range where a viscosity decrease is anticipated, *i.e.* 0.35 (using the distance between P and F as  $\sigma$ ) to 0.18 (additionally including the fluorine radius from the Lennard-Jones potential<sup>223</sup>), Fig. 16.

Charge asymmetry is an important design aspect of ionic liquids, at least for those with rigid ions. Borate based ionic liquids provide an excellent platform for targeted modifications of this kind. The reason is that borate anions with different (mixed) combinations of the four substituents are synthetically accessible, such as cyano(fluoro)borates or cyano(chloro)borates.<sup>192,224–228</sup> Table 2 shows targeted modifications involving a systematic exchange of the nitrile groups in the tetracyanoborate anion with either fluorine or chlorine. It is evident that the ionic liquids with charge asymmetric anions (*i.e.* with large charge arm) have qualitatively lower viscosities and melting points. The physical properties of the cyano(chloro)borates differ less than those of the cyano(fluoro)borates, in line with the smaller charge asymmetry. However, this is not sufficient to determine causation, since other effects might be relevant. For example, nitrile groups are relatively large and prone to dipolar interactions, which distinguishes these from the halides.<sup>229,230</sup> Furthermore, the anions all have different masses and volumes, and the resulting ionic liquids differ in their densities. Thus, molecular dynamics simulations like those presented in Fig. 17 would be valuable to test the charge asymmetry hypothesis. The approach is straightforward: starting from a native force field for  $[BF_4]^-$ ,  $[BCl_4]^-$  and  $[B(CN)_4]^-$ , the atomic charges are rearranged to give a nonzero charge arm of reasonable magnitude. The relative changes in predicted properties – such as the viscosity – can then be compared to the relative changes found in Table 2.

Apart from very general statements, we have so far not considered the effect of targeted modifications on thermal transitions. The reason for that is that thermal properties depend on a multitude of subtle effects, and the obtained results also vary with experimental parameters such as cooling rate or thermal history.<sup>231</sup> Even simple spherical ion models predict a variety of thermal transitions,<sup>220,221</sup> which we will discuss here briefly. The thermal transitions reported by Lindenberg and Patey show a pronounced dependency both on the size difference and the charge displacement of the ions, Fig. 18.<sup>221</sup>

In Fig. 18, the cation size (both absolute and in relation to the anion) increases from left to right, the shading of the blocks corresponds to the magnitude of charge displacement  $d$ . Systems with small or no charge displacement tend to crystallise into high-melting CsCl (similar ion sizes) or NaCl (different ion sizes) structures. Introducing displaced charges into the systems with similar size leads to glassy states upon cooling with an underlying orientationally ordered crystal structure.<sup>220,221</sup> The melting points are reduced considerably by several hundred



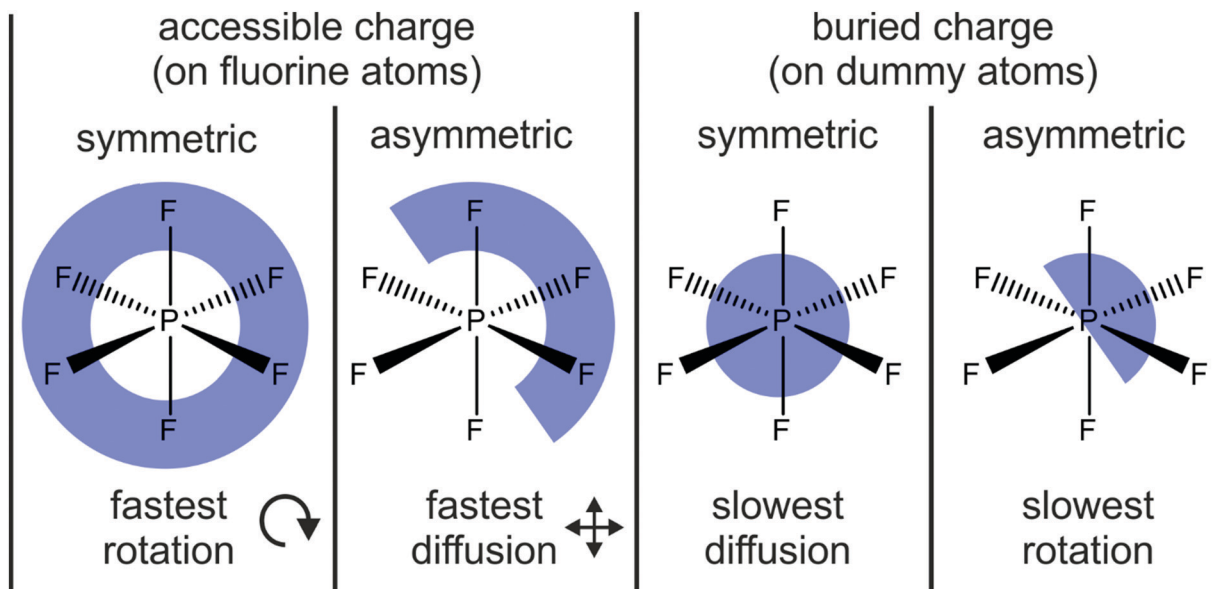


Fig. 17 Four model systems with accessible vs. buried and symmetric vs. asymmetric charge distribution as targeted modification are used to demonstrate the impact on transport properties. Translational diffusion is fastest with accessible and asymmetric charge distribution, whereas rotational diffusion is fastest with accessible and symmetric charge distribution.<sup>222</sup>

**Table 2** Properties of ionic liquids with different borate anions. The cation is always  $[\text{C}_2\text{C}_1\text{im}]^+$ . Viscosities are at 20 °C, physicochemical data taken from ref. 226, see also ref. 224, 225 and 228. For volume and charge arm computed at the full MP2/cc-pVTZ//B3LYP-GD3BJ/6-311+G(d,p) level of theory, see ESI

Anion	Viscosity/mPa s	Melting point/°C	Volume/Å <sup>3</sup>	$l_q/\sigma$
$[\text{B}(\text{CN})_4]^-$	22.2	8	152	0
$[\text{BF}(\text{CN})_3]^-$	12.6	-7.8	132	0.085
$[\text{BF}_2(\text{CN})_2]^-$	11.1	-16	112	0.114
$[\text{BF}_3(\text{CN})]^-$	14.7	-28	91	0.116
$[\text{BF}_4]^-$	34	14	69	0
$[\text{B}(\text{CN})_4]^-$	22.2	8	152	0
$[\text{BCl}(\text{CN})_3]^-$	18.6	—	149	0.004
$[\text{BCl}_2(\text{CN})_2]^-$	19.6	-4	147	0.015
$[\text{BCl}_3(\text{CN})]^-$	21.1	8	144	0.030
$[\text{BCl}_4]^-$	36.6	15	141	0

Kelvin even for relatively small charge displacements. It is worth noting that the authors observed pre-melting transitions relevant to solid state physics in the region of large charge displacements and disparate ion sizes: first, plastic crystals with ion pairs rotating together; second, fast ion conductors where the larger cations remain on their lattice positions, while the anions are mobile.<sup>221</sup>

Naturally, for ionic liquids, liquefaction temperatures will depend on a multitude of factors beyond the dipole moment. However, as a general trend, a lowering of liquefaction temperatures with increasing dipole moments has been observed experimentally.<sup>232,233</sup>

## Hydrogen bonding

In the previous section we have collected sound evidence that weak directional bonding can fluidize ionic liquids – both with respect to melting points and transport properties.<sup>111,232</sup>



Fig. 18 Melting point trends with different charge displacements ' $d$ ' and cation sizes.<sup>221</sup> The charge displacement  $d$ , does not translate directly to  $l_q$  because the charge centre also is assigned a small mass and because only 2/3 of the charge are displaced. Thus,  $d = 26$  corresponds to  $l_q \sim 1.7$  Å, or a relative charge arm of 0.67 (cation size 100%) to 0.22 (cation size 300%). The anion radius is held constant, the cation radius is increased from 2.5 to 7.5 Å, units of  $d$  are 0.1 Å. Reprinted from E. K. Lindenberg and G. N. Patey, *J. Chem. Phys.*, 2015, **143**, 024508, with the permission of AIP Publishing.



This observation is directly related to the potential effects of hydrogen bonding in ionic liquids. Indeed, hydrogen bonding is recognised as a concept of central importance in this field, including hydrogen bonds in aprotic ionic liquids as well as the 'conventional' hydrogen bonds present in protic ionic liquids.<sup>48,58,234–236</sup> In this section we will focus on how targeted modifications can be exploited to understand and design hydrogen bonding, using the well-known imidazolium cation as an example.

In imidazolium cations, any ring atom bonded to a hydrogen atom can serve as hydrogen bond donor. In general, the strongest hydrogen bonds are formed by N1–H or N3–H groups,<sup>237</sup> followed by C2–H, and finally the much weaker C4–H and C5–H groups.<sup>234,238</sup> The numbering of ring atoms is shown in Fig. 19a. While hydrogen bonds are formed by the alkyl chains,<sup>234,239–241</sup> these will not be discussed and we will focus on hydrogen bonds at the C2 position. A systematic experimental study of both melting points and – where possible – transport properties of imidazolium salts with all 20 possible methylation patterns would be desirable. We refer the interested reader to the work by Fumino, Ludwig, *et al.*;<sup>242–245</sup> structures and available melting points are given in the ESI.†<sup>242,243,246</sup>

In many cases of practical importance, hydrogen bonding is a dominant interaction between the solvent (ionic liquid) and solute. Prototypical examples are the dissolution of cellulose and the Diels–Alder reaction in ionic liquids as solvents.<sup>247–253</sup> Both experimental and theoretical studies of solute–solvent interactions for practical applications rely on having a reference solvent with different hydrogen bonding properties, without significantly altering other properties.

In the spirit of targeted modifications, the hydrogen atom may simply be replaced by a methyl group. Bonhôte *et al.* used this concept in their seminal work on the [NTf<sub>2</sub>]<sup>−</sup> anion, Fig. 19.<sup>161</sup> The strongest hydrogen bond in this set of ionic liquids is formed at the C2 position, whereas C4 and C5 hydrogen bonds are much weaker. Surprisingly at the time of the discovery, an unusually high viscosity was observed for the ionic liquid 1-ethyl-2,3-dimethylimidazolium bis(trifluoromethanesulfonyl) imide [C<sub>2</sub>C<sub>1</sub>C<sub>1</sub>im][NTf<sub>2</sub>] (Fig. 19d), where the C2 hydrogen bond is disrupted by methylation.<sup>161</sup> Other imidazolium ionic liquids with different anions show the same behaviour, which sparked a debate on the nature of the increase in viscosity relevant in the context of this perspective.<sup>254–258</sup> Critically, the cation in Fig. 19d differs from the other cations in Fig. 19 not only by the absence of hydrogen bonding, but also in other parameters – in particular size and conformational flexibility.<sup>254</sup>

The effect of size and flexibility on viscosity can be seen by comparing the viscosities of the [NTf<sub>2</sub>]<sup>−</sup> based ionic liquids in Fig. 19a–c, which vary from 44 to 25 mPa s. The ionic liquid with C2 methylated cation (Fig. 19d) clearly exceeds this range, with a viscosity of 88 mPa s. However, this modification alone is not sufficient to identify the cause of the difference – it can be argued that the methyl group in C2 position impedes rotation of the ethyl side chain around the N–C bond, thus reducing conformational flexibility.<sup>254</sup> This second hypothesis can be tested by comparing the properties of ionic liquids with the

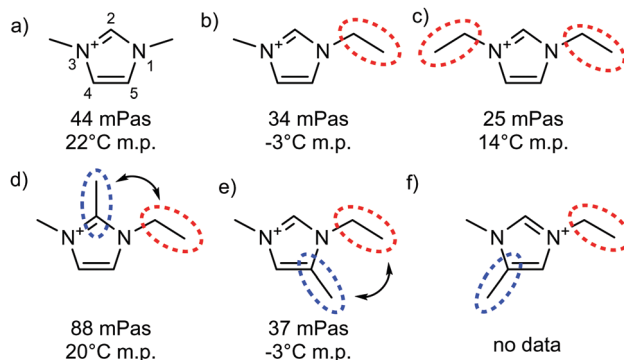


Fig. 19 Viscosities at 20 °C and melting points taken from ref. 161, all systems with [NTf<sub>2</sub>]<sup>−</sup> as anion. Red dotted circles highlight flexible groups, blue dotted circles highlight hydrogen bonding sites which are blocked by methylation, double-headed arrows indicate situations where conformational flexibility is potentially reduced by steric clash.

cations shown in Fig. 19d–f, all having the same mass, but different substitution patterns. If restricted rotation was indeed the sole cause of the higher viscosity, but not hydrogen bonding, then the C2 and C5 methylated analogues would be expected to show the same, or very similar viscosities, whereas the C4 methylated cation (Fig. 19f) would be more fluid than these two. In contrast, the C2 methylated imidazolium ionic liquid (Fig. 19d) is more than twice as viscous as the C5 methylated counterpart (Fig. 19e). Thus, hydrogen bonding at C2 can be identified as the dominating factor in this case. The comparison with the C4 methylated imidazolium ionic liquid would allow for more insight into the interplay with conformational flexibility. However, the required data are not yet available.

It would be desirable to have the data shown in Fig. 19 for anions such as methanesulfonate or acetate, which are known to engage in stronger hydrogen bonding than [NTf<sub>2</sub>]<sup>−</sup>. This data is not available at the moment, especially not for the more exotic cations in Fig. 19c–f. In addition, melting points can cause practical problems. An example is [C<sub>2</sub>C<sub>1</sub>C<sub>1</sub>im][OAc] with a melting point around 90 °C, whereas [C<sub>2</sub>C<sub>1</sub>im][OAc] is a room temperature ionic liquid.<sup>259</sup> Such large differences pose a challenge when viscosities are to be compared. Furthermore, decomposition becomes a relevant issue for anions which are good hydrogen bond acceptors due to their inherent basicity.<sup>259</sup>

Methylation as a targeted modification can be used with theoretical methods to not only identify that hydrogen bonding in C2 position fluidises ionic liquids, but also to understand why. Several groups investigated the barrier associated with anions moving between the (equivalent) positions on top and below of the imidazolium ring *via* the in plane orientation on the C2–H site, where hydrogen bonding can occur for the non-methylated analogues.<sup>255,260,261</sup> Fig. 20a shows the potential energy surface for the position of the [NTf<sub>2</sub>]<sup>−</sup> anion around the two cations 1,2,3-trimethylimidazolium and 1,3-dimethylimidazolium reported by Rodrigues *et al.*<sup>255</sup> The C–H···N angles of 45° and 360 – 45 = 315° correspond to the [NTf<sub>2</sub>]<sup>−</sup> positions above and below the imidazolium ring plane; which is the preferred geometry in the gas phase. Transition between these two



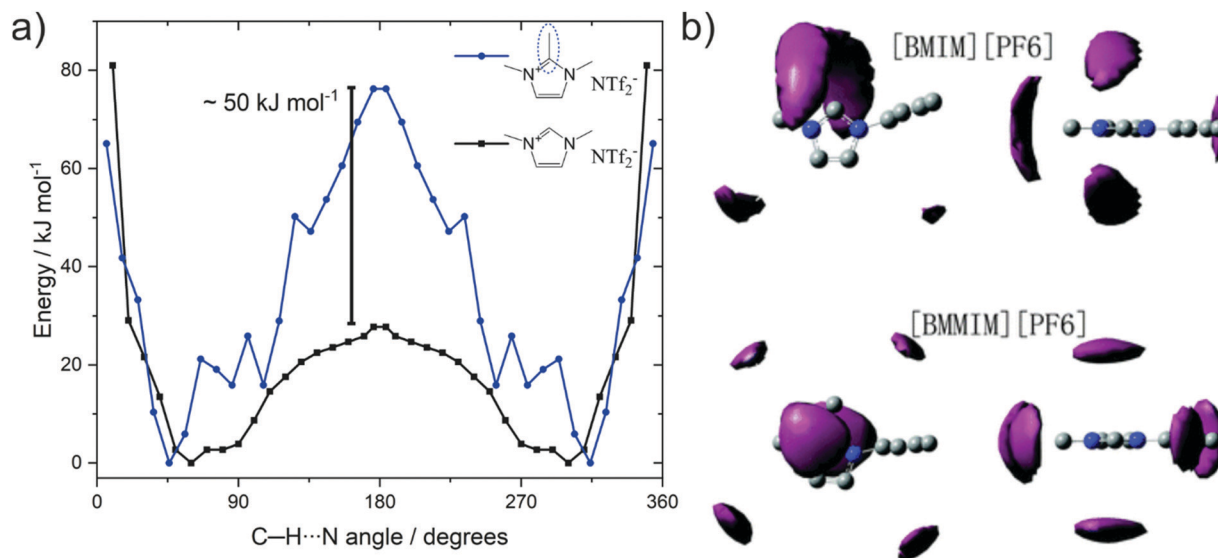


Fig. 20 (a) Energy as a function of the angle between C2 carbon, acidic hydrogen, and imide-N of the  $[\text{NTf}_2]^-$ .<sup>255</sup> (b) Spatial distribution function of  $[\text{PF}_6]^-$  around  $[\text{C}_4\text{C}_1\text{im}]^+$  (top) and  $[\text{C}_2\text{C}_1\text{im}]^+$  (bottom). Reproduced from ref. 257 with permission from the PCCP Owner Societies.

minima occurs *via* the in-plane geometry at  $180^\circ$ , which is associated with a much higher barrier for the methylated cation, *i.e.* the blue curve in Fig. 20a.

The significantly increased barrier for the C2 methylated analogue decreases the entropy, as fewer positions are available for the anion relative to the cation. The resulting localisation of the anion for C2 methylated imidazolium cations is visible in spatial distribution functions obtained from bulk molecular dynamics simulations, Fig. 20b.<sup>257,260</sup> Thus, the loss in conformational entropy of the side chains is overcompensated by the gain in configurational or vibrational (in the sense of an intermolecular vibration mode) entropy.<sup>255,257,260</sup> However, we would like to emphasise that this entropy argument does not conflict with hydrogen bonding as cause of the higher fluidity. Although hydrogen bonding is an additional interaction – which would be expected to decrease fluidity – it comes to act upon an arrangement of ions which is high in energy, *i.e.* the in-plane position of the anion. Thus, this additional interaction will flatten the potential energy surface, rendering more minima accessible. In analogy to conformational flexibility, additional pathways become available for structural relaxation, lowering the viscosity and accelerating the dynamics.

The anions in the examples shown in Fig. 20 are generally considered to be weakly coordinating, even though aprotic ionic liquids such as  $[\text{C}_2\text{C}_1\text{im}][\text{NTf}_2]$  show rich hydrogen bonding.<sup>262,263</sup> Nevertheless, the higher transition barrier upon methylation shown in Fig. 20a is also found for anions that are good hydrogen bond acceptors, such as halides.<sup>260,261</sup> Hydrogen bonding facilitates the position change of halides. The structure with a  $\text{C-H}\cdots\text{Cl}$  angle of  $180^\circ$  in  $[\text{C}_1\text{C}_1\text{im}]\text{Cl}$  is a minimum, separated by less than  $10 \text{ kJ mol}^{-1}$  from the structures where the chloride is situated above and below the ring.<sup>260</sup> In the C2-methylated  $[\text{C}_1\text{C}_1\text{im}]\text{Cl}$ , the  $180^\circ$  structure is a transition state rather than a minimum, and the energy barrier increases to above  $40 \text{ kJ mol}^{-1}$ .<sup>260</sup>

The interpretation of hydrogen bonding above is not without critics. Several groups question the relevance of hydrogen bonding in imidazolium cations, or at least point out the deviations from ‘conventional’ hydrogen bonds.<sup>264–267</sup> Some of the discrepancies are rooted in more or less strict definitions of hydrogen bonding and the overlap between different interpretations. The last point is illustrated with the targeted modifications of the  $[\text{C}_4\text{C}_1\text{im}]^+$  cation shown in Fig. 21. As discussed above, methylation of the C2 position leads to an increase in viscosity, Fig. 21b. However, a similar increase in viscosity can be achieved with a branched side chain, Fig. 21c. The cation shown in Fig. 21c is interesting as the increase in viscosity compared to ionic liquids with linear side chains significantly exceeds what one would expect given the small reduction in conformational flexibility, Fig. 21d.<sup>268</sup> Using molecular dynamics simulations, Zhang *et al.* could show that the curious properties of this particular ionic liquid are rooted in the preferred orientation of the side chain in space, altering the distribution of the anion around the cation as well as the barriers which separate preferred positions.<sup>269</sup> Thus, the increases in viscosity for both the C2 methylated and the branched imidazolium ionic liquids can be explained with a change in configurational entropy. However, the deeper-rooted causes are different: loss in hydrogen bonding upon methylation of the C2 position, and steric exclusion of the anion by the branched side chain. The choice of a model and the level of coarse graining are to some degree subjective; this arbitrariness must be kept in mind when conflicting conclusions are encountered.

Another example of a ‘conflicting’ interpretation is the free volume model. This model explains the high viscosities of C2 methylated imidazolium ionic liquids with their decreased free volume compared to non-methylated ionic liquids.<sup>270</sup> In other words, the ions are more densely packed in the case of C2 methylated cations. There clearly is some conceptual overlap





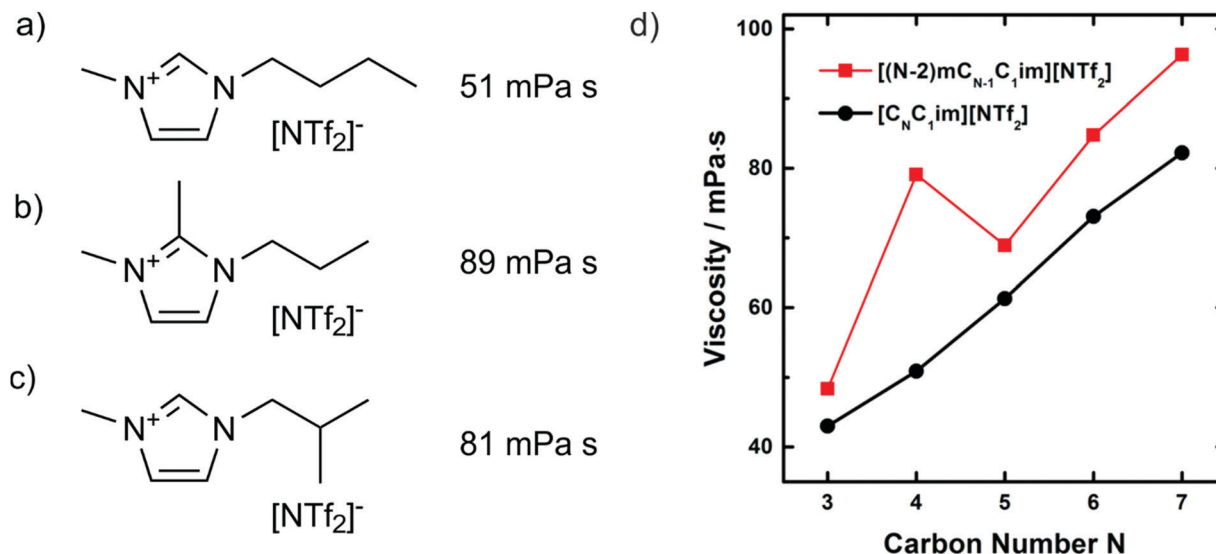


Fig. 21 Viscosities at 25 °C of (a)<sup>268</sup> [C<sub>4</sub>C<sub>1</sub>im][NTf<sub>2</sub>] and (b)<sup>271,379</sup> and (c)<sup>268</sup> targeted modifications with the same molecular weight. Slightly lower values for (a) (46 mPa s) and (b) (84 mPa s) were reported in.<sup>380</sup> The dynamic viscosity as a function of the number of carbons in the side chain is shown in (d) for both linear (black curve) and branched (red curve) cations.<sup>268</sup> Reprinted with permission from L. Xue, E. Gurung, G. Tamas, Y. P. Koh, M. Shadeck, S. L. Simon, M. Maroncelli and E. L. Quitevis, *J. Chem. Eng. Data*, 2016, **61**, 1078–1091. Copyright 2016 American Chemical Society.

with the spatial distribution interpretation we used above. If the anions are more localised and less mobile due to high barriers, then the free volume in the bulk liquid will be lower, and *vice versa*. However, the hypothesis of a standalone ‘free volume model’ can be tested with relative ease by considering isodensity conditions. The density of [C<sub>3</sub>C<sub>1</sub>im][NTf<sub>2</sub>], Fig. 21b, is 1.4579 g mL<sup>-1</sup> at 25 °C and atmospheric pressure, under which conditions it has a viscosity of 89 mPa s.<sup>270</sup> Under the same conditions, the density of [C<sub>4</sub>C<sub>1</sub>im][NTf<sub>2</sub>], Fig. 21a, is lower at 1.4366 g cm<sup>-3</sup>, indicating a less dense packing and thus a larger free volume, under these conditions it has a viscosity of 51 mPa s.<sup>271</sup> Isodensity conditions are reached at approximately 31 MPa, where [C<sub>4</sub>C<sub>1</sub>im][NTf<sub>2</sub>] reaches a density of 1.4579 g mL<sup>-1</sup> and a viscosity of 73 mPa s, see ESI.†<sup>271</sup> Thus, it may be necessary but it is not sufficient to consider free volume and isodensity conditions to explain the high viscosity of the C2 methylated analogue.

The value of methylation in understanding the behaviour of imidazolium ionic liquids is apparent. However, changes in size and shape of the cation also contribute to the changes in macroscopic properties. Instead, imidazolium based ionic liquids can be compared with ionic liquids based on other azoles which are not significantly different in size or mass, Fig. 22. Few studies exist on triazolium and diazolum (other than 1,3-disubstituted imidazolium) ionic liquids, usually focusing on niche uses as catalysts or using highly functionalised cations.<sup>272–278</sup> Some groups make use of the higher stability towards base of 1,2,3-triazolium cations which is a result of the missing acidic hydrogen atom in 2-position.<sup>279–281</sup> It would be beneficial to have data with stronger hydrogen bond acceptors than [NTf<sub>2</sub>]<sup>-</sup>, the available thermal properties indicate that very low melting points can be obtained even with ions such as [OMs]<sup>-</sup> or [BF<sub>4</sub>]<sup>-</sup>.<sup>282</sup> For example, the triazolium analogue of [C<sub>4</sub>C<sub>1</sub>im]

[OMs] with the triazolium nitrogen in 4-position is reported to be a liquid, however we could not find viscosity data.<sup>278</sup>

The structures shown in Fig. 22 are useful targeted modifications to understand hydrogen bonding in ionic liquids. As a first approximation, the protic diazolum ionic liquids in Fig. 22a will form stronger hydrogen bonds than their aprotic imidazolium counterparts. Conversely, the triazolium cations in Fig. 22b are missing one hydrogen bonding site. Indeed, higher viscosities compared to [C<sub>4</sub>C<sub>1</sub>im][NTf<sub>2</sub>] have been reported for the triazolium counterparts, Fig. 22b, in agreement with the methylation approach.<sup>283–285</sup> However, it is surprising that the increase in viscosity of ionic liquids is higher when the

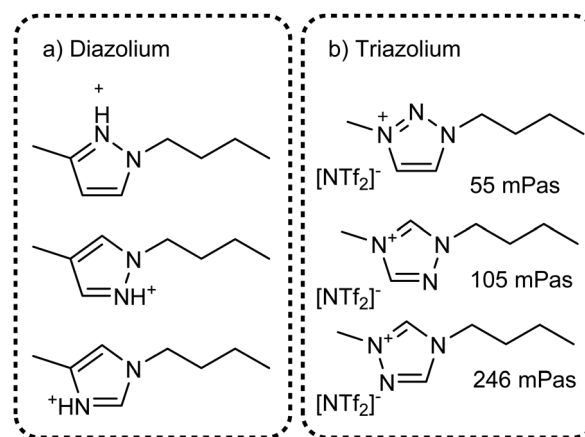


Fig. 22 Targeted modifications of the [C<sub>4</sub>C<sub>1</sub>im]<sup>+</sup> cation: (a) diazolum based cations and (b) triazolium cations. Some properties for ionic liquids with triazolium [NTf<sub>2</sub>]<sup>-</sup> ionic liquids have been measured, viscosities at 25 °C are given in (b). Values are taken from: 2-triaz viscosity,<sup>285</sup> 4-triaz viscosity,<sup>283</sup> 5-triaz viscosity,<sup>284</sup> for the sake of simplicity we follow the numbering of the 1-butyl-3-methylimidazolium base structure.



hydrogen bonding site in 4 or 5 position is omitted. In addition, the viscosity at room temperature of the 1,2,3-triazolium analogue is almost identical to that of  $[C_4C_1im][NTf_2]$ , Fig. 22a. Thus, the simple picture of missing hydrogen bond donors does not correlate well with the situation found for imidazolium cations. As a second approximation, changes in the molecular electrostatic potential<sup>286</sup> and the acidity of the different sites<sup>287–289</sup> for each specific azole have to be considered.

The insight promised by the targeted modifications justifies the effort, even if the resulting ionic liquids themselves are too expensive or otherwise unsuited for practical applications. The ultimate goal is the transition from understanding to design. Examples of how such a transition can be achieved have been put forward by Ludwig and Paschek.<sup>290</sup> The authors suggest adding substituents of different electronegativity to the 4 and/or 5 positions of  $[C_2C_1im][NTf_2]$  to change physicochemical properties of the bulk ionic liquid by changing the strength of hydrogen bonds.<sup>245,290</sup>

One of the least disruptive targeted modifications is the use of isotopologues. Consequently, deuteration is commonly used as a probe, for example in NMR or neutron scattering studies.<sup>166,238,239,291–294</sup> However, it must be kept in mind that the free energy surface might be affected by the mass of the nuclei. Although the effects are in general small, one needs to be aware and wary of these to correctly interpret results. This is important for the exchange of protium with deuterium, and the effect of this exchange on hydrogen bonds. Deuteration changes the zero point energy and the energy difference between vibrational eigenstates, even if the potential provided by the internal electronic energy does not change.<sup>295–298</sup> However, this does not directly translate into one isotope forming stronger hydrogen bonds than the other.<sup>296,299–301</sup> In general, at lower temperatures deuterium bonds are more stable than hydrogen bonds, and the trend reverses at higher temperatures where entropic effects become more important.<sup>297,300</sup> Yoshimura *et al.* confirmed the pronounced temperature dependence of deuterium bonds and the qualitative difference to protium bonds for  $D_2O$  and  $H_2O$  in the ionic liquid  $[C_4C_1im][BF_4]$ .<sup>302</sup> In some cases the results are counterintuitive – for example, Remsing *et al.* provided evidence that the  $Cl \cdots D$  bond is stronger than the  $Cl \cdots H$  bond in  $[C_4C_1im]Cl$ , despite the measurement being at 90 °C.<sup>238</sup>

Differences in hydrogen bonding due to deuteration can affect both the solid and liquid state. For the solid, small changes in geometry upon deuteration can accumulate, leading eventually to polymorphism or differences in melting point.<sup>303–305</sup> Abe *et al.* observed significant changes in the crystallisation temperature for wet diethyl(2-methoxyethyl)(methyl) ammonium tetrafluoroborate  $[N221(2O1)][BF_4]$  for different deuterium contents.<sup>304,306</sup>  $[N221(2O1)][BF_4]$  with 1 mol%  $H_2O$  was observed to crystallise at  $-14$  °C, whereas  $[N221(2O1)][BF_4]$  with 1 mol%  $D_2O$  crystallised at  $-40$  °C.<sup>304</sup> The water content in these measurements was rather low. 1 mol% is approximately equivalent to 0.1 mass% for this ionic liquid, once again showing the importance of impurities.<sup>307</sup> Isotope effects of the deuterium can also be observed in the liquid state. The miscibility of

$[C_4C_1im][BF_4]$  with protiated *vs.* deuterated butanols shows that this is relevant beyond the subtleties of spectroscopic methods.<sup>308</sup> For all four butanols, the UCST\*\* is lowered by 4–7 °C upon deuteration, *i.e.* the  $d1$ -butanols are more miscible.<sup>308</sup> Finally, we note that a change in conformer distribution has been reported for several  $[C_4C_1im]$  based ionic liquids.<sup>309,310</sup> However, further work is required to demonstrate that the reported preference for the gauche conformer is not an artefact of the higher intensity in the corresponding Raman band for the C2 deuterated cation.

Targeted modifications to investigate the effects of hydrogen bonding are not limited to the polar core. In some cases, the periphery of a cation is even more accessible experimentally than the core, since a variety of different alkylating agents are readily available. Thus, a functional group capable of hydrogen bonding can be formally introduced into a cation in an ionic liquid. Targeted modifications have been used to show that this can lead to pronounced cation–cation interactions and the formation of cation clusters.<sup>92,311–315</sup> In addition, the cation–anion interaction can be disrupted if the hydrogen bonding site competes with the polar core.<sup>311,316</sup> Systematic studies of hydrogen bonding and cationic clustering led to a detailed understanding of the lubricity in closely related ionic liquids, Fig. 23.<sup>317</sup> Li *et al.* observed an approximately linear relationship between lateral and normal force in an atomic force microscopy experiment for  $[C_4py][NTf_2]$ , Fig. 23a.<sup>317</sup> In contrast, the hydroxyl functionalised analogue  $[HOC_4py][NTf_2]$  showed a discontinuous behaviour in the same setup, Fig. 23b. The authors rationalised this observation with the presence of an ordered cation layer stabilised by hydrogen bonding, Fig. 23c. Notably, the authors also successfully used other targeted modifications to disrupt the bilayer, such as an anion with increased hydrogen bonding tendency.<sup>317</sup> Thorough understanding of the surface behaviour of ionic liquids is of vital importance for several practical applications such as lubrication<sup>318–320</sup> or catalysis with supported ionic liquid phase (SILP).<sup>321–323</sup>

## Fluorinated *vs.* alkyl side chains and nanosegregation

There are two design parameters relevant to this section, the length of the side chain and the degree of fluorination. However, this leads to a problem since both design parameters are coupled to other properties. The replacement of hydrogen with fluorine can be considered a targeted modification if one is aware of the manifold of changes that come with it. We will thus present only a few select examples of fluorinated *vs.* protonated side chains and use this section to discuss the limits of the ‘targeted modifications’ approach. First, we will discuss small fluorine containing groups, followed by longer fluorinated side chains that induce fluorophilic behaviour.

\*\* “Upper critical solution temperature”, above this temperature ionic liquid and butanol become completely miscible across the whole composition range.



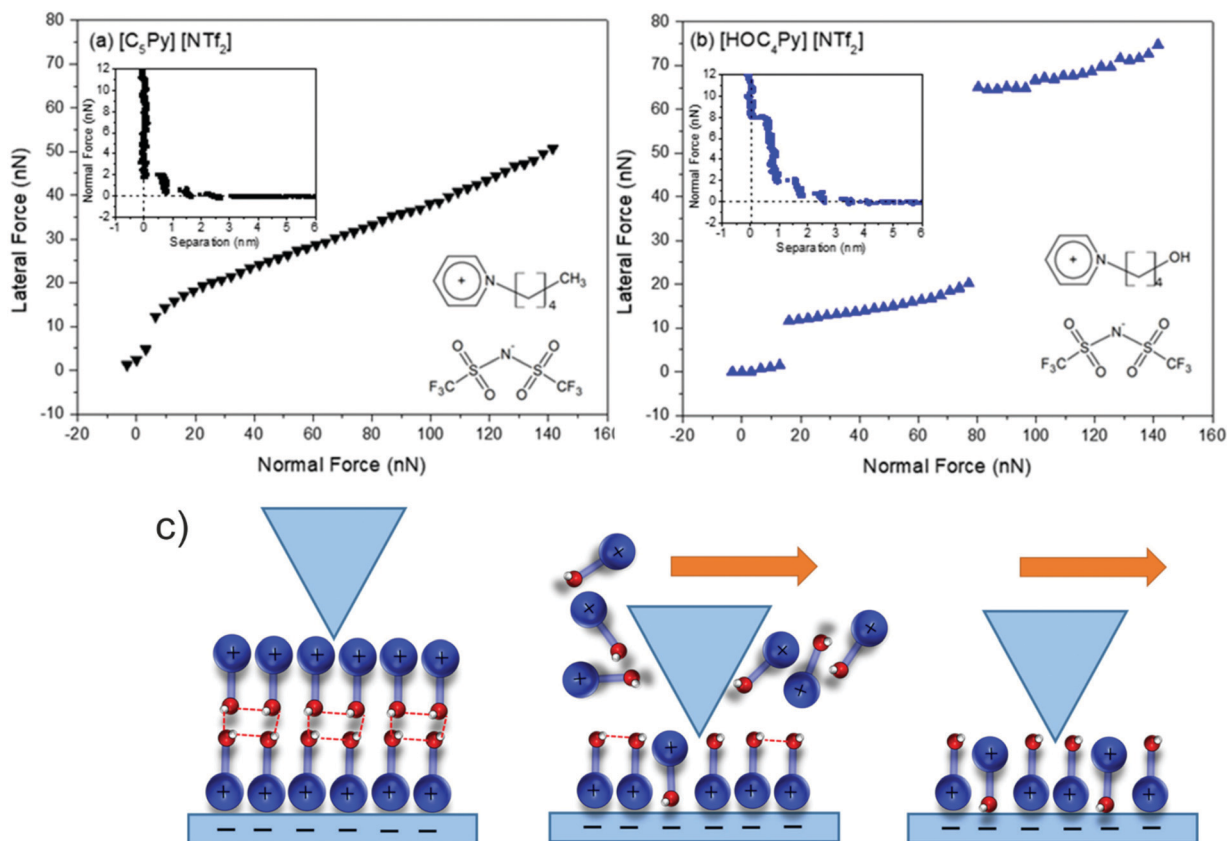


Fig. 23 The comparison between ionic liquids with alkyl- and hydroxyalkyl pyridinium cations revealed their interesting lubrication behaviour. Lateral vs. normal force shows a linear behaviour for the alkyl pyridinium cation (a), whereas distinct steps are observed for the hydroxyalkyl pyridinium cation (b). The authors provide a plausible rationalisation in terms of bilayer formation (c).<sup>317</sup> Reprinted (adapted) with permission from H. Li, T. Niemann, R. Ludwig and R. Atkin, *J. Phys. Chem. Lett.*, 2020, 3905–3910. Copyright 2020 American Chemical Society.

The trifluoromethyl group is a common motif in ionic liquids, with a methyl group being the protonated analogue, Fig. 24. Trifluoromethanesulfonate [OTf]<sup>−</sup> and methanesulfonate [OMs]<sup>−</sup> anions are commonly used as anions in ionic liquid studies.<sup>324–328</sup> Interestingly, the motivation for the use of the [OTf]<sup>−</sup>/[OMs]<sup>−</sup> pair in these studies is indeed usually targeted modification – however to investigate the influence of hydrogen bonding rather than fluorination. Due to the electron withdrawing effect of fluorine substitution, the [OTf]<sup>−</sup> anion has a lower basicity than [OMs]<sup>−</sup>, which results in weaker hydrogen bonds.<sup>326–328</sup> The desire for weakly coordinating anions (together with high chemical and thermal stability) is at the root of the widespread use of highly fluorinated anions in ionic liquids, which is why understanding the effects of fluorination is so important.<sup>103,104,329,330</sup> Studies that compare [OMs]<sup>−</sup> and [OTf]<sup>−</sup> might thus also include an anion like trifluoroacetate [TFA]<sup>−</sup>, which also shows a more pronounced tendency to form hydrogen bonds than [OMs]<sup>−</sup>.<sup>327</sup> Finally, we have mentioned earlier that alkyl side chains (in cations) can act as hydrogen bond donors, which is naturally not the case for perfluoroalkyl chains.<sup>234,239–241</sup> In contrast, fluorine atoms can serve as weak hydrogen bond acceptors in some cases.<sup>331–336</sup> However, this is generally believed not to be relevant for ionic liquids.

The effect of fluorination on the properties of [NTf<sub>2</sub>]<sup>−</sup> ionic liquids has been studied by comparison with the non-fluorinated analogue [NMs<sub>2</sub>]<sup>−</sup>, Fig. 24b.<sup>103</sup> Similar to [OMs]<sup>−</sup> vs. [OTf]<sup>−</sup>, Pringle *et al.* report higher viscosities for the corresponding ionic liquids with the non-fluorinated anion as a result of stronger interactions.<sup>103</sup> Furthermore, the higher basicity of the non-fluorinated anion leads to reduced thermal stability of corresponding imidazolium and pyrrolidinium ionic liquids.<sup>103,330</sup> These effects and their rationalisation can be understood intuitively. Much more subtle and not immediately evident are the effects of fluorination on conformational flexibility. The [NMs<sub>2</sub>]<sup>−</sup> anion can be considered a rigid anion for practical purposes, thus the well-known plasticising properties of the flexible [NTf<sub>2</sub>]<sup>−</sup> anion also depend on its degree of fluorination.<sup>105</sup> This finding poses an interesting starting point for further systematic investigations of partly fluorinated anions intermediate between [NTf<sub>2</sub>]<sup>−</sup> and [NMs<sub>2</sub>]<sup>−</sup>.<sup>337</sup>

Size and mass are inevitably linked to fluorination. As a rule of thumb, a trifluoromethyl group is approximately equivalent to an isopropyl group in terms of size.<sup>331</sup> The effects of size and mass have been discussed earlier. In principle, classical molecular dynamics simulations can differentiate between the different effects. For example, it is straightforward to selectively change a force field to produce an artificial ‘lighter fluorine’





Fig. 24 Fluorinated (top row) and non-fluorinated (bottom row) anions and their typically observed behaviours. Fluorinated anions such as (a) [OTf] or [NTf<sub>2</sub>] are often used for their beneficial physicochemical properties.

(or 'smaller fluorine') without changing other characteristics such as the charge distribution. The mixing and solvation behaviour of fluorocarbons is partly determined by the volumes of the involved species, hence such a study would be relevant for practical applications where ionic liquids are to be used as solvents.<sup>338</sup> The steric demand of fluorine in these systems has three implications:

(1) The size of the ion itself increases (relative to the counterion), see the implications in the section 'size and mass of ions'.

(2) The importance of conformational flexibility for ions such as [NTf<sub>2</sub>]<sup>-</sup> is believed to be due to the relatively large CF<sub>3</sub> groups, see the section 'conformations and conformational flexibility'.<sup>163,165</sup>

(3) The polarisable carbon atoms are sterically shielded by the less polarisable fluorine atoms, reducing dispersion interactions.<sup>339</sup>

These points are also valid for longer alkyl and fluoroalkyl side chains. In addition, it is well established that alkyl and (per)fluoroalkyl side chains of sufficient length lead to the formation of apolar and fluorous domains, coexistent with the polar domains formed by the charged parts of the ions.<sup>340–353</sup> Hollóczy *et al.* investigated this behaviour with the help of molecular dynamics simulations.<sup>354</sup> The authors simulated mixtures of [C<sub>8</sub>C<sub>1</sub>im]Br and the fluorinated counterpart [(R<sub>6</sub>C<sub>2</sub>)C<sub>1</sub>im]Br, Fig. 25. All seven compositions ranging from pure [C<sub>8</sub>C<sub>1</sub>im]Br over a 1:1 mixture to pure [(R<sub>6</sub>C<sub>2</sub>)C<sub>1</sub>im]Br showed well-defined polar, apolar, and fluorous domains.<sup>354</sup> The [(R<sub>6</sub>C<sub>2</sub>)C<sub>1</sub>im]<sup>+</sup> cation is a targeted modification of the [C<sub>8</sub>C<sub>1</sub>im]<sup>+</sup> cation, which allowed the authors to study the influence of fluorination. Future studies might focus on the different contributions of the manifold changes that come with the introduction of fluorination, isolating the effects of mass, size, charge distribution, *etc.*

To this point, the aggregation of alkyl and fluoroalkyl side chains in separate domains has been treated as an empirical result. Nanosegregation in ionic liquids with extended side chains is not simply the result of the (fluoro)alkyl chains being excluded from the polar parts; otherwise only two domains

would form rather than the three domains visible in Fig. 25. The conundrum can be solved by considering the interactions favoured by alkyl and perfluoroalkyl chains, respectively.

Long alkyl side chains in ionic liquids are believed to actively aggregate due to dispersion interactions, as opposed to being merely expelled from the charged region.<sup>355–359</sup> Firaha *et al.* used *ab initio* simulations to directly investigate the effects of dispersion interactions.<sup>358</sup> To this end, the authors compared results from *ab initio* molecular dynamics simulations using DFT functionals with and without dispersion correction. Thus, even basic physical concepts such as dispersion interactions can be probed with targeted modifications, in this case different functionals. Surprisingly, the authors observed a decrease in alkyl chain aggregation when dispersion was included.<sup>358</sup> Although the interaction between alkyl chains did indeed increase, it was found that the interactions between polar parts benefited even more.

Highly fluorinated compounds have properties which are distinctively different from their protonated counterparts. In their seminal work from 1994, Horváth and Rábai introduced the term 'fluorous' to address this behaviour.<sup>360,361</sup> In general, fluorous compounds and solvents are poorly miscible both with water and conventional organic solvents.<sup>338–340,362</sup> Fluorous biphasic catalysis makes use of this phenomenon by tagging a catalyst with a 'fluorous ponytail', which leads to a preferred solvation in fluorous solvents and facilitates catalyst recovery.<sup>340,363,364</sup>

The behaviour is commonly explained in two ways. First, fluorine atoms in perfluoroalkyl chains sterically shield the more polarisable carbon atoms, thus reducing dispersion interactions.<sup>331,339,365</sup> Second, long perfluoroalkyl chains tend to adopt a helical conformation,<sup>336,339,366</sup> thus leading to non-vanishing permanent dipole moments which interact with each other, forming a network of 'stratified dipole arrays'.<sup>367–369</sup> In reality, both explanations contribute to some degree to the macroscopic properties of fluorous compounds, and the relative importance of each will depend on the exact system. The energy differences between conformers in perfluoroalkanes are small,<sup>370,371</sup> and in some cases all-*trans* staggered



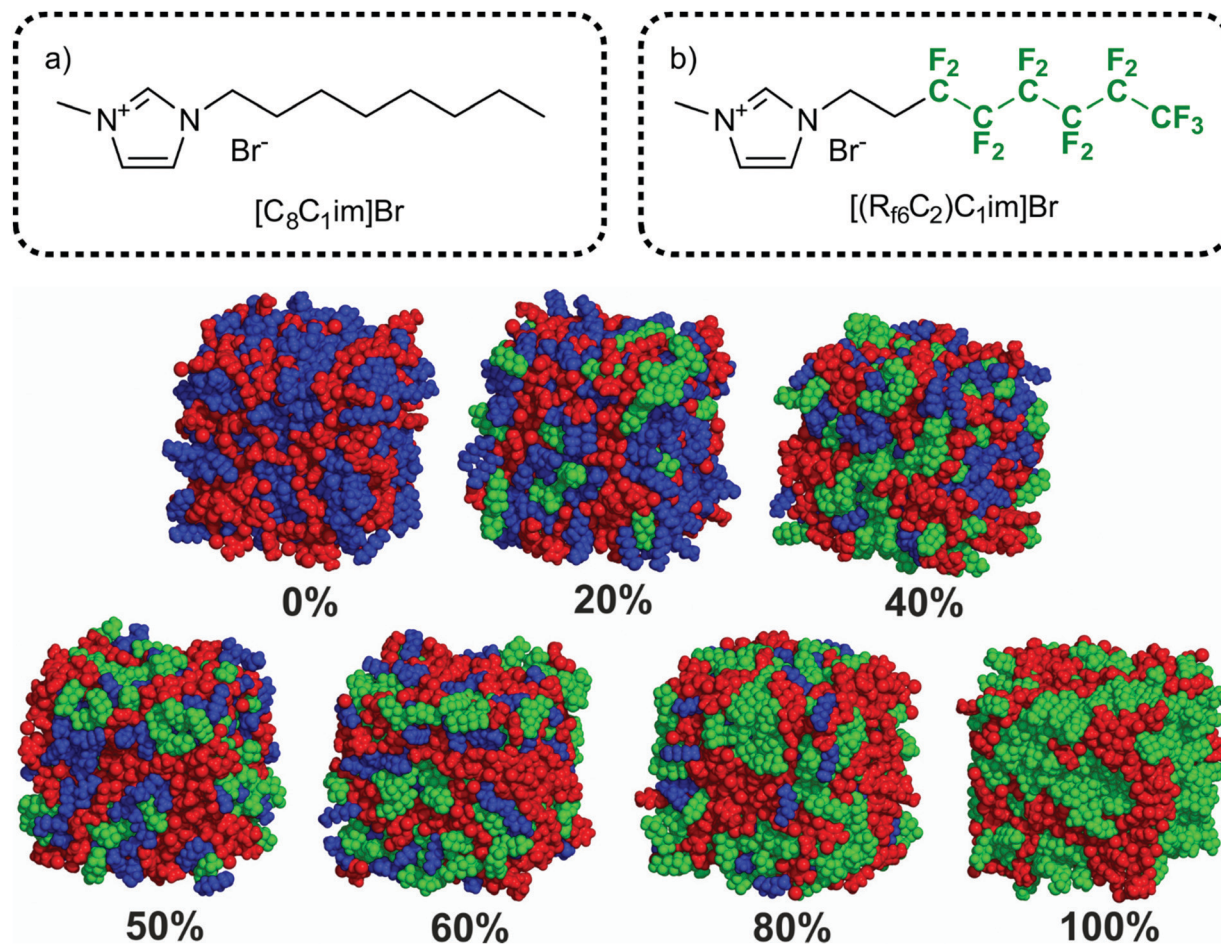


Fig. 25 Mixtures of (a)  $[C_8C_1im]Br$  and (b) the highly fluorinated analogue  $[(R_{f6}C_2)C_1im]Br$  lead to a triphasic structure, i.e. nanosegregation into apolar (blue), polar (red), and fluorous (green) domains. The three domains are clearly discernible in the different snapshots shown in (c) for an increasing percentage of the fluorous IL.<sup>354</sup> Reprinted with permission from M. Brehm, H. Weber, M. Thomas, O. Hollóczki and B. Kirchner, *ChemPhysChem*, 2015, **16**, 3271–3277. Copyright 2015 Wiley-VCH Verlag GmbH & Co. KGaA, Weinheim.

conformations are found, stabilised by numerous weak fluorine–fluorine interactions.<sup>336</sup>

Ionic liquids lead to a unique situation as fluorous and hydrocarbon moieties are forced together to meet electroneutrality conditions. It is an interesting question as to what extent this nanosegregation (literally demixing on a molecular scale) influences bulk properties of ionic liquids, including those with short side chains. Molecular mimics and their mixing behaviour provide a potential way forward to answer this question. Thus, fluorination and nanosegregation conflate all aspects covered in this perspective, which in turn shows that great care has to be exercised to rationalise bulk properties with molecular concepts.

## Summary and outlook

In this perspective, we present targeted modifications as an approach to understand how the attractive macroscopic properties of ionic liquids arise, and how they can be controlled. This approach seeks to overcome the limitations of structure–property relationships by further breaking down the structural changes

which inevitably lead to changes in more than one molecular property. An example is the side chain length (structure) in imidazolium ionic liquids and the effect on liquefaction temperatures (property). Extending the side chain leads to irregular behaviour of the liquefaction temperatures, since many molecular properties change at the same time.<sup>56</sup> Several underlying concepts such as size and mass, conformational flexibility, or dispersion interactions need to be deployed to rationalise the observed behaviour. We thus present targeted modifications as a ‘concept–property relationship’, where the object of study is modified in a way that allows one to address one underlying concept at a time.

We discuss several examples where targeted modifications could be used and have been used to enable the design of ionic liquids by understanding how their properties arise. In many cases, a synergistic combination of theory and experiment is required to provide a satisfactory answer. For example, localised charge and flexibility have both been suggested as cause for the faster dynamics in phosphonium ionic liquids compared to ammonium ionic liquids. Computational approaches can investigate these two contributions separately



and settle the dispute. In this example, flexibility is now thought to be the dominant contribution, whereas the localised charge slows down dynamics.<sup>139</sup> However, one point cannot be stressed enough: the existence of two models does not mean that only one of these can be right. We showed this for the free volume model vs. hydrogen bonding in imidazolium ionic liquids, where targeted modifications revealed that neither of the two provides a convincing, standalone rationale.

Despite the invaluable insight provided and promised by targeted modifications, making quantitative predictions remains an issue. This is due to the complexity of some of the relationships and the coupling between these. Thus, to achieve *a priori* design beyond what is covered by targeted modifications, statistical methods such as machine learning might be necessary. However, even then, knowledge of targeted modifications will be helpful. The ideas described in this perspective can be used to reduce or even inspire the parameter space for statistical methods, since targeted modifications can reveal which concepts are important and which aren't. Goal-oriented feature selection and machine learning are rapidly advancing fields which become more and more relevant for ionic liquids.<sup>372</sup> In this context, targeted modifications serve as a bridge between basic research and practical applications.

Several promising models aim to provide quantitative predictions based on microscopic structure, such as mode coupling theory or excess entropy scaling.<sup>181,373–377</sup> To some degree, these models suffer from a limited interpretability on a molecular level. Future work in this direction might use a fruitful combination of statistical mechanics and targeted modifications, for example to quantify the contribution of conformational flexibility to excess entropy and thus to macroscopic transport properties.

Further experimental and theoretical studies in this direction have considerable potential, for example using molecular mimics and diazolum/triazolum analogues. The successes of existing studies presented in this perspective show that there is merit to targeted modifications.

## Conflicts of interest

There are no conflicts to declare.

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