

Cite this: *RSC Adv.*, 2017, 7, 52632Received 18th September 2017
Accepted 2nd November 2017

DOI: 10.1039/c7ra10358a

rsc.li/rsc-advances

Hollow polymer particles: a review

Ros Azlinawati Ramli *

Herein, the basic principles, such as the definitions, classifications, and properties, of hollow polymer particles (HPPs) are critically investigated. HPP preparation methods, such as template, self-assembly, and acid/alkali swelling methods, have been briefly introduced. Moreover, advanced methods to prepare HPPs have been highlighted. In addition, recent studies and applications of HPPs have been briefly described. HPPs are polymeric spheres with size in the range from nanometer to micron and a single pore inside the particles. This review is focused on the classification of HPPs based on the core-shell, single particles, and stimuli-responsiveness. A single pore structure is the most desirable characteristic that allows this class of polymers to encapsulate large guest molecules and thus to be used as a drug delivery system.

1. Introduction

Hollow polymer particles (HPPs) are polymeric spheres with size in the range from nanometer to micron (1–1000 nm) and a single pore inside the particle; on the other hand, polymers with many pores are called porous polymers.^{1,2} Interest in HPPs has increased steadily over the past few decades because HPPs can readily be used to upload drugs and active species and can efficiently sequester specific target molecules into the hollow interiors by chemical modification.³ Furthermore, they have a variety of applications, such as they are used as coatings,^{8–13}

drug delivery systems,^{14–17} catalysts,^{18–20} and photonic crystals,^{21–23} in chemistry, biotechnology, and materials science.^{4–7} Core-shell HPPs are the most widely studied polymer particles in this class of polymers. The combination of inorganic-organic materials holds great promise for novel core-shell hybrid micro- and nanostructured HPPs.^{24,25}

HPPs are usually prepared in a spherical form using different methods to obtain the hollow part. HPPs can be synthesized in a liquid or solid form in the range of tens of nanometers to several millimeters.^{26,27} A variety of monomers, including methyl methacrylate (MMA), butyl acrylate (BA), butyl methacrylate (BMA), and vinyl acetate (VAc), have been used to synthesize HPPs. Moreover, one of the earliest processes for making hollow polymer particles was developed by Kowalski and co-workers in 1984 at Rohm and Haas company using an acid/alkali swelling method.^{28,29} The common methods that are used to form hollow structures of HPPs are template,^{1,30–32} self-assembly,^{33,34} and acid/alkali swelling,³⁵ with different synthesis techniques such as suspension, dispersion, emulsion, and precipitation polymerization. Advanced methods for the preparation of hollow polymers that have been reported recently include urea treatment,³⁶ an *in situ* reduction method,^{37,38} inverse mini-emulsion periphery RAFT polymerization (IMEPP),³⁹ and coaxial electrohydrodynamic and atomization (CEHDA).⁴⁰ Different types of materials, such as monomers,⁴¹ block copolymers,⁴² copolymers,⁴³ and inorganic nanoparticles,⁴⁴ have been used to synthesize HPPs. Advanced techniques, such as scanning electron microscopy, transmission electron microscopy, atomic force microscopy, and dynamic light scattering have been used to study the morphology system of HPPs.

2. Classifications

Hollow polymer particles (HPPs) can be classified depending on their morphology into core-shell and single-particles; in addition, depending on their state phenomenon, they can be

Material Technology Program, Faculty of Industrial Sciences & Technology, Universiti Malaysia Pahang (UMP), Lebuhraya Tun Razak, 26300 Gambang, Kuantan, Pahang, Malaysia. E-mail: azlinawati@ump.edu.my; Fax: +06-095492766; Tel: +06-0172949907



Ros Azlinawati Ramli received her BSc. (Hons) from Universiti Kebangsaan Malaysia (UKM) in 2004, and obtained her MSc. (2012) and Ph.D (2014) in Polymer Engineering from Universiti Teknologi Malaysia (UTM). She was awarded her PhD scholarship from the Ministry of Higher Education Malaysia (MOHE) and Best Student Doctor of Philosophy (Polymer Engineering) from UTM. Previously, she was

a Senior Chemist for Synthomer R&D Asia. Currently, she is a senior lecturer at Material Technology Program, Universiti Malaysia Pahang (UMP). Her research interests include emulsion polymers, polymer composites, microgels for biomedical applications, and hydrogels for agricultural applications.

classified into stimuli-responsive, dispersion, and non-dispersion. Moreover, core-shell HPPs can be classified into solid shell and porous shell, as shown in Fig. 1. This section of this study emphasizes on core-shell, single-particle, and stimuli-responsive HPPs.

2.1 Core-shell HPPs

Core-shell HPPs can be divided into solid shell and porous shell HPPs. The hollow part can be the core or in the shell of the particles. Hollow spheres with complex core-shell structures have attracted significant interest due to their superior properties such as tunable diameter, cavity size, large surface area, low density, and precise functionality.^{41,45} Efforts have been made to fabricate core-shell and hollow particles in the past 10–15 years.⁴⁶ The simplest method to fabricate core-shell HPPs is to coat a sacrificial core particle with a shell, followed by removal of the sacrificial core.^{21,47,48} The core can be removed by dissolution, thermolysis, or chemical etching to create a hollow interior.^{1,49}

2.1.1 Solid-shell HPPs. Solid-shell HPPs can be divided into single shell, multiple shell, and single hole in the shell HPPs. Solid-single shell HPPs are widely prepared as compared to other types. However, considering the complexity of some applications, solid-multiple shell HPPs have been developed to meet the increasing application requirements, *e.g.* carriers for multiple drug delivery with controlled release profiles or promising cellular organelle mimics.^{46,50,51} The solid-single hole in the shells of HPPs provided new opportunities in applications involved in the uptake of target species and in the encapsulation of drugs.³ Solid-shell HPPs are typically prepared by acid/alkali swelling methods using a multistep approach.^{13,35,52–54} During the entire polymerization process, the feeding rate of the monomer mixture, the emulsifier content, the cross-linking agent content, and the ratio of the monomers during the core and shell polymer preparation are the main factors that affect the formation of specific morphology and dispersity of the HPPs.⁵²

Yuan and co-workers prepared HPPs using seeded emulsion polymerization and alkali swelling methods. MMA-MAA-DVB was prepared as core particles, and styrene, acrylonitrile, and

divinylbenzene (DVB) were then synthesized onto the cores, followed by alkalization treatment to obtain hollow particles. They prepared hollow particles at different shell thicknesses with different weight ratios of the core monomers to shell monomers. The shell thickness increased with a decrease in the core/shell ratio. When the core/shell ratio was 1 : 8, the hollow emulsion particles obtained were monodispersed and maintained a more preferable balance between the rigidity of the shell polymers and the hollow percentages.⁵² Tan *et al.* developed novel core-double-shell particles with poly(methyl methacrylate-*co*-butyl acrylate) P(MMA-*co*-BA) as the core, poly(methyl methacrylate-*co*-butyl acrylate-*co*-methacrylic acid) P(MMA-*co*-BA-*co*-MAA) as the inner shells, and poly(styrene-*co*-methyl methacrylate) P(St-*co*-MMA) as the outer shells *via* a soap-free emulsion polymerization and acid-alkali osmotic swelling process.⁵⁵ Mandal *et al.* prepared HPPs by surface-confined living radical polymerization on silica templates. Using a silica microsphere as a sacrificial core, hollow poly(benzyl methacrylate) (PBzMA) microspheres were produced by core dissolution. Shell thicknesses were measured from the scanning electron microscopy (SEM) images, which are provided in Table 1. The data reveal that the shell thickness increases with the increasing polymerization time.⁵⁶

Xu and Asher prepared 150–700 nm HPP particles with 100–500 nm hollow cores *via* dispersion polymerization and HF etching. Monodisperse HPPs with outside diameters between 200 and 700 nm and shell thicknesses between 20 and 90 nm were synthesized by varying the silica core size, while maintaining the monomer/silica weight ratio at 0.9 and keeping all the other conditions constant.²² Minami *et al.* prepared HPPs with a single hole in the shell by the self-assembling of phase-separated polymer (SaPSeP) method. The formation of the hole in the shell is due to the emulsifier adsorbed at the interface of the droplet that affects the self-assembly of PDVB molecules at the interface. Fig. 2 shows the SEM images of PS/PDVB particles in the absence (a and c) and presence (b and d) of sodium dodecyl sulfate (SDS) and transmission electron microscopy (TEM) (c and d) images of their ultrathin cross-sections. A single hole was observed in the shell of the hollow polymer particles prepared in the presence of SDS (Fig. 2b).⁵⁷

Guan *et al.* prepared HPPs through direct polymerization and cross-linking reactions occurring on the surface of carboxyl-capping polystyrene (PS) colloids. The formation mechanism of holes that is based on the concomitant microphase separation and symmetrical volume shrinkage of shell materials provides



Fig. 1 Classification of hollow polymer particles (HPPs) depending on different parameters.

Table 1 Shell thickness and molecular weight of the hollow poly(benzyl methacrylate) microspheres prepared by varying the polymerization time. Reproduced from ref. 56 with permission from the American Chemical Society

Polymerization time (h)	Shell thickness (nm)	Molecular weight (M_n)	PDI (M_w/M_n)
3.5	175–225	9150	1.56
6.5	350–400	13 450	1.37
14.0	550–600	26 500	1.26





Fig. 2 SEM images (a) and (b) of PS/PDVB particles and TEM images (c) and (d) of ultrathin cross-sections of PS/PDVB particles prepared by seeded polymerizations in the absence (a) and (c) and presence (b) and (d) of SDS. Reproduced from ref. 57 with permission from the American Chemical Society.

a reproducible approach for creating holes in polymer shells. To obtain a better understanding of the formation mechanism of the hole, they obtained the reaction products at different polymerization stages. Core-shell microspheres with a single hole covered with a layer of an ultrathin film were observed after 12 h, and core-shell microspheres with open holes were observed after 20 h.³ Fig. 3 shows the schematic for the formation process of holes in the shells of PS involving: (1) the formation of a prepolymer layer, (2) the microphase separation of two polymer shell layers, (3) the development of a small interior void in the shell, (4) formation of a film-covered hole, and (5) formation of an open hole in the polymer shells.



Fig. 3 Schematic for the formation process of holes in the shells of polystyrene (PS). Reproduced from ref. 3 with permission from John Wiley and Sons.

2.1.2 Porous-shell HPPs. HPPs with pores in their shells have received significant attention owing to their low density and high specific surface area; these materials are very useful in surface-related applications such as in sensors⁵⁸ electro-rheological applications,⁵⁹ solar cells,⁶⁰ mimicking photosynthesis,⁶¹ catalysts,⁶² and optical applications.⁶³ The porous hollow structure also plays an important role in the superhigh capacity of drugs.⁶⁴ A single Au nanoparticle encapsulated in the porous shell of HPPs remarkably improved the catalytic efficiency and recyclability for high performance heterogeneous catalysts.⁶⁵ Template synthesis is one of the most used strategies for the preparation of porous-shell HPPs.^{66–69} This method usually involves tedious treatment, such as rinsing, calcination, or chemical etching, to remove the surfactants or templates, where the structure of the shell may be damaged.⁷⁰ Accordingly, several novel methods have been introduced such as the self-assembly of emulsion particles at the monomer droplet interface,⁷¹ Pickering emulsion polymerization,⁷⁰ and a facile swelling–evaporation process.⁶⁵ Nanoparticles, such as ferric oxide (Fe_3O_4) NPs, can be added to the porous shells to transform HPPs to magnetic hollow porous molecularly imprinted polymers (HPMIPs).⁷²

He *et al.* prepared HPPs with a porous shell by self-assembling of sulfonated polystyrene (PS) emulsion particles at the monomer droplets interface. The swelling of the PS emulsion particles by the oil phase provided a driving force to develop the hollow core. The increase in the polymerization rate led to a thicker shell formation. After polymerization, the self-disengagement of the adsorbed PS particles from the polymeric shell resulted in the porous shell. Fig. 4 shows the SEM (A) and TEM (B) images of PVAc. The TEM images show that the microspheres have hollow core/porous-shell structures, and the pores are distributed compactly and evenly in the shell, which can be seen from the contrast in the bright and dark fields. The porous shell was very thin, with an approximately 35–172 nm thickness for all the system.⁷³

Chen *et al.* prepared poly(acrylic acid) (PAA) nanogels with a hollow core-porous shell structure by the direct polymerization of an acrylic acid monomer in the presence of hydroxypropyl cellulose (HPC) and a cross-linking agent, *N,N*-



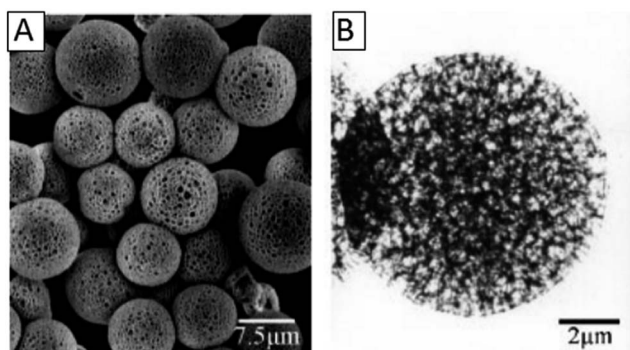


Fig. 4 SEM (A) and TEM (B) images of the hollow core/porous shell of PVAc. Reproduced from ref. 73 with permission from John Wiley and Sons.

methylenebisacrylamide (MBAAm), followed by removal of HPC at a medium pH ($\text{pH} > 6.0$) from the generated HPC-PAA nanoparticles.⁶⁴ Guan *et al.* prepared HPPs with a mesoporous shell by Pickering emulsion polymerization as a potential colloidal collector for organic contaminants in water. Isopentyl acetate (PEA) was included to generate a hydrophobic core that could be easily removed during the drying process, resulting in hollow particles. The hollow microspheres were reinforced with silica nanoparticles; this resulted in an amphiphilic surface and nanotunnels in the shell.⁷⁰ Han *et al.* developed HPPs with a single Aurum (Au) nanoparticle encapsulated in the porous shell with superior catalytic efficiency and recyclability. At first, a Au-poly(*o*-methoxyaniline) (POMA) core-shell hybrid was synthesized, followed by swelling treatment in a suitable swelling solvent. Finally, evaporation of the swollen Au-POMA core-shell hybrids resulted in a hollow nanostructure formation.⁶⁵

2.2 Single-particle HPPs

All the aforementioned HPPs have a closed-shell structure; this means that the encapsulated materials need to diffuse through the shell to reach the interior, which is a slow and ineffective process. Accordingly, single-particle HPPs with one hole on the surface were developed to easily and quickly load a variety of functional materials (including molecular species, proteins,

and nanoparticles) into the interiors. The size of the hollow spheres and holes can be controlled by the temperature, amount of solvent, and the initiator to monomer ratio. The special characteristic of single-particle HPPs is that the hole can be closed *via* a thermal or solvent-induced annealing process after loading the materials.^{74,75} This class of HPPs exhibits promising applications in drug encapsulation and controlled release systems.

Jeong *et al.* prepared hollow particles with engineered holes on their surfaces (microscale fish bowls) using two different routes. The first route required the use of pre-synthesized latex beads and was mainly limited to a few amorphous polymers such as PS and PMMA. The second route involved the use of an emulsion derived from a polymer solution. This route was more generic and could be directly applied to both amorphous and semicrystalline polymers such as polycaprolactone (PCL) and poly (*L*-lactide) (PLLA).⁷⁴ Fig. 5 shows the SEM images of microscale fish bowls prepared from both routes.

Im *et al.* prepared another class of macroporous capsule-polymer shells with controllable holes on their surfaces. The existence of holes (224 nm in radius) on the surfaces of the PS hollow particles allowed for direct loading of the functional materials. After loading, the holes could be closed by annealing the system at a temperature slightly above the glass-transition temperature (T_g) of PS. Fig. 6 shows an SEM image demonstrating the possibility to load the bowl-shaped PS particles with solid PS beads of 260 nm diameter.⁷⁵

Han *et al.* prepared HPPs of poly(*o*-toluidine) (POT) in an aqueous solution by a droplet template without surfactant stabilization. By simply changing the concentration of *o*-toluidine, the size of the resulting hollow spheres ranged from nanometers to micrometers. Temperature was found to have a significant effect on the size of the hollow spheres. The size of the hole on the surface of each hollow microsphere could be controlled by changing the molar ratio of the initiator to monomer. The droplets formed by the monomer in an aqueous solution were proposed to act as a template during the polymerization process. Fig. 7 shows the TEM images of the POT synthesized at different initiator to monomer ratios (a and b) and at different temperatures (c and d). The results show that the size of the hole on the surface of each hollow

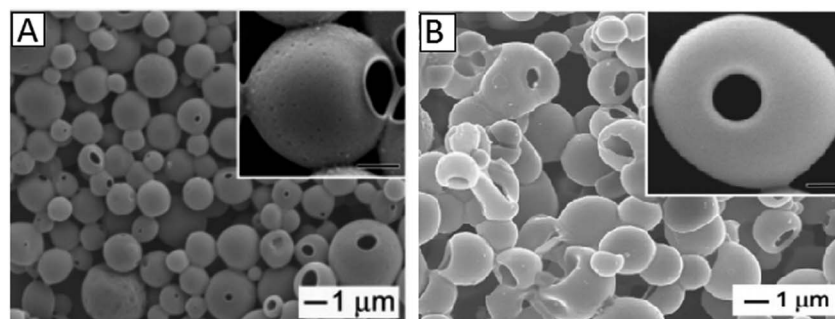


Fig. 5 (A) SEM images of microscale fish bowls derived from PS bead swelling by styrene and (B) fish bowls prepared from PMMA emulsions containing different amounts of toluene. Reproduced from ref. 74 with permission from the American Chemical Society.



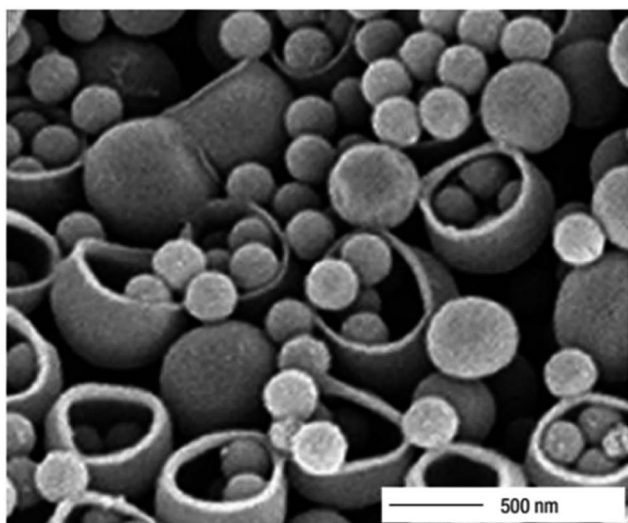


Fig. 6 Demonstration of the possibility to load bowl-shaped PS particles with solid PS beads of 260 nm diameter. Reproduced from ref. 75 with permission from the Nature Publishing Group.

microsphere increased with the decreasing [APS]/[*o*-toluidine] ratio. Temperature is another important factor that influences the sizes of the POT hollow microspheres. When the polymerization reaction temperature was maintained above 15 °C, irregular nanoparticles as well as the hollow microspheres were observed. If the temperature is between 0 and 10 °C, the size of the hollow spheres slightly changes, whereas the size decreases significantly at lower temperatures (<0 °C).⁷⁶

2.3 Stimuli-responsive HPPs

Stimuli-responsive HPPs are smart materials that are capable of responding to environmental stimuli such as temperature, pH, light, and magnetic field, which allows for the controlled release and site-specific targeting of the loaded drugs modulated by environmental stimuli.^{77,78} The best technique to prepare stimuli-responsive HPPs is polymerization from templates. This technique allows the employment of a variety of functional monomers, which contributes to better stimuli-responsive performance than other techniques.⁷⁹ The functionality of stimuli-responsive HPPs can be increased by introducing more than one shell layer and the addition of inorganic/magnetic particles, such as silica, aurum, and ferric oxide, to the particles. The magnetic-targeting stimuli-responsive HPPs show unique performance as intelligent carriers for the site-specific controlled delivery of drugs or genes. The most used monomer to prepare stimuli-responsive HPPs is *N*-isopropylacrylamide (NIPAM), which exhibits a volume-phase transition temperature (VPTT) close to the temperature of the human body.⁸⁰ Upon a slight change of temperature, poly(*N*-isopropylacrylamide) (PNIPAM) chains undergo macromolecular transition from a hydrophilic to a hydrophobic structure.⁸¹

Li *et al.* prepared silica-polymer double-shell HPPs with independent pH and temperature sensitivity *via* a sol-gel process. The poly(methacrylic acid) (PMAA) inner shell and the PNIPAM outer shell could respond independently to pH and temperature of the environment, respectively. Monodispersed double-walled hollow PMAA-PNIPAM microspheres were obtained by selective removal of the inorganic silica core and interlayer by HF treatment for 48 h.⁸² Du *et al.* designed double-shell HPPs with independent pH-responsive cross-linked PMAA



Fig. 7 SEM images of poly(*o*-toluidine) (POT) hollow microspheres synthesized at different [APS]/[*o*-toluidine] ratios of: (a) 0.8 : 1 and (b) 0.6 : 1. (c) and (d) SEM images of POT hollow microspheres synthesized at different temperatures: (c) 25 °C and (d) −10 °C. The insets show the corresponding TEMs image of the POT hollow structures. Reproduced from ref. 76 with permission from John Wiley and Sons.





Fig. 8 Schematic of the synthesis process for doubly temperature-responsive core-double-shell microgels and hollow spheres. Reproduced from ref. 51 with permission from John Wiley and Sons.

inner shells, temperature-responsive cross-linked PNIPAM outer shells, and movable superparamagnetic cores from $\text{Fe}_3\text{O}_4/\text{SiO}_2/\text{PMAA}/\text{SiO}_2/\text{PNIPAM}$ penta-layer onion-shaped core-shell microspheres based on polymerization on template approaches.⁷⁹ Pan *et al.* synthesized magnetic/hollow double-shell-imprinted polymers (MH-MIPs) by Pickering emulsion polymerization. In this preparation process, ATP particles were used as stabilizers for oil-in-water Pickering emulsions, and a few hydrophilic Fe_3O_4 nanoparticles in the water phase were used as magnetic separation carriers. Then, the interfacial imprinted shells of double-shelled MIPs were fabricated in the oil phase, and ATP particles were attached to the surfaces of the interfacial imprinted shells to form permeable outer shells.⁸³ Macková *et al.* designed hollow particles with a triple functionality by inverse emulsion polymerization. These hybrid particles have potential in magnetic and biomedical applications for temperature-driven preparative separations of biological media.⁸⁴ Dubbert *et al.* prepared hollow double-shell microgels consisting of two thermoresponsive polymer shells made of PNIPAM and poly(*N*-isopropylmethacrylamide) (PNIPMAM). The hollow particles were derived from silica-PNIPAM-PNIPMAM core-shell-shell (CSS) particles by dissolution of the silica core. In the first synthesis step, the PNIPAM shell was attached to a functionalized silica core; this resulted in particles with a single shell. In the second step, a shell made of PNIPMAM was added. This led to the formation of doubly temperature-sensitive CSS particles, as shown in Fig. 8.⁵¹

3. Preparation methods of HPPs

Hollow polymer particles (HPPs) are typically prepared by template, self-assembly, and acid/alkali swelling using different synthesis techniques such as suspension, dispersion, emulsion, and precipitation polymerization. One of the earliest processes for making hollow emulsion particles was developed by Rohm and Haas. Their process involved emulsion polymerization and alkali swelling methods.⁵⁶ Among the various approaches

reported, template-directed synthesis is the most widely used method for the formation of polymer hollow particles.^{21,85,86} However, solid-shell HPPs are mostly prepared by acid/alkali swelling method.^{13,35,52–54} Table 2 lists the advantages and disadvantages of methods used for the preparation of hollow polymer particles (HPPs).

3.1 Template method

The template method is known to be one of the most effective approaches to achieve hollow nanostructures.⁸⁷ The method can be divided into four categories: (1) conventional hard-templating synthesis, (2) sacrificial-templating synthesis, (3) soft-templating synthesis, and (4) template-free methods. The most commonly employed hard templates are monodisperse silica particles and polymer latex colloids. The sacrificial template is the template itself, which is involved as a reactant in the synthetic process of the shell material (or its intermediate). The most commonly used soft (liquid or gaseous) templates are emulsion droplets, surfactants and other supramolecular micelles, polymer aggregates/vesicles, and gas bubbles.⁸⁸ The template-free method/self-templated method is based on novel mechanisms, such as Ostwald ripening (a spontaneous process, first described by Wilhelm Ostwald in 1896).⁸⁹ Templates are selected, and their diameters can be controlled easily to predetermine the sizes of HPPs;⁹⁰ to ensure that the hollow particles retain integrity after the template synthesis, a cross-linking strategy is employed.⁹¹ Polymer nanoparticles, such as emulsion and polymeric micelles, have been increasingly used to fabricate HPPs because it is easy to control the size, morphology, and surface functionality of these nanoparticles. Emulsion particles are employed to fabricate hollow particles with larger diameters in the range from the submicron to micrometer, whereas polymeric micelles are used for smaller hollow particles with diameters less than 100 nm.^{92,93}

MAH/VAc (maleic anhydride/vinyl acetate) (PMV) particles have been used as hard templates for the synthesis of maleic anhydride/divinylbenzene (MAH/DVB) hollow polymer particles.



Table 2 Advantages and disadvantages of methods used for the preparation of hollow polymer particles (HPPs)

Method	Advantages	Disadvantages
(A) Template method		
Hard-templates	Void size and shape of the particles are determined by the size and shape of the template Hard-templating method in principle can be used to prepare any kind of HPPs and therefore is considered as the most popular method Easy to control size, morphology, and surface functionality of the HPPs	The templates are removed completely without any contribution to the composition of the final products In some cases, toxic etchants or solvents are required to remove the template core, which seriously limits its application
Soft-templates	In most soft-templating approaches, only the coating process is needed, because it is unnecessary to remove the soft templates	Coating process usually requires an additional surface modification procedure, which could potentially cause low reproducibility and high cost
Self-templates	Self-templating strategy is preferred in practical applications because of the significantly reduced production cost, simplified synthesis procedures, and the ease of scaling up	Self-templating is only suitable for the synthesis of hollow structures with specific compositions, limiting its application field Practical application of the soft-templating method has been mainly limited by the size and uniformity because it produces materials mainly in the micrometer range with high polydispersity
(B) Self-assembly		
	Produces HPPs in nanometer scale with high spherical particles	Shell cross-linking must be conducted at very low concentration of micelles solution; this restricts the self-assembly method for industrial application Electrostatic interaction as a driving force is only suitable for layer-by-layer (LbL) assembly of polymer electrolytes Micelle formation of block copolymer involves UV cross-linking of the shell-layer and subsequent tedious ozonolysis of the core Compositions and structures of the block copolymers are very important for the preparations of polymer vesicles as seen in the limited successful examples
(C) Acid/alkali swelling		
	Formation of hollow particles with a definite morphology (particles containing one central hollow)	Complicated process controlled by a number of thermodynamic and kinetic factors

The approaches consist of three major steps, as shown in Fig. 9: the synthesis of template particles based on MAH/VAc; the formation of core/shell particles *via* precipitation polymerization using MAH/DVB as comonomers and PMV as templates; and the removal of the core by dissolving it with an organic solvent, *i.e.*, *n*-butyl acetate. The shell thickness of the hollow spheres and the volume of their hollow core could be controlled simply by varying the size of the template, the monomer concentration, and the ratio of the monomer to template.⁹⁴

In another study reported by Li *et al.*, poly(divinylbenzene-*co*-acrylic acid) (P(DVB-*co*-AA)) microspheres were used as templates to synthesize hollow microspheres with movable P(DVB-*co*-AA) cores. Tri-layered-structure microspheres were prepared by three-stage distillation precipitation polymerization. The hollow P(DVB-*co*-AA) microspheres with a 48 nm shell thickness were formed by selective dissolution of the noncross-linked PAA segment.⁹⁵ In another study, polypyrrole-chitosan (PPy-CS) hollow nanospheres with movable Ag nanoparticles (Ag@PPy-CS) cores were prepared using a photoreduction method. This method does not require the preparation of polymer layers as compared to the template-assisted approach. In this method, hollow nanospheres were prepared first before the introduction of the core material.⁴

Boyer *et al.* prepared hollow polymer nanocapsules of styrene (St) and maleic anhydride (MAH), (St-*alt*-MAH), exploiting gold

nanoparticles (GNP) as sacrificial templates for polymers obtained by RAFT polymerization, as illustrated in Fig. 10. Block polymers (made by RAFT polymerization) were assembled onto the GNP surfaces using a grafting process. The polymers were then cross-linked using a small diamine molecule. Finally, the gold cores were removed using aqua regia; this resulted in the formation of a stable, biocompatible, low cytotoxic, and anti-fouling hollow polymer.⁹⁶

McKelvey and co-workers reported a cationic equilibrium vesicles templating method for producing hollow polymeric spheres of styrene and divinylbenzene. The cationic equilibrium vesicles were formed by cetyltrimethylammonium tosylate (CTAT) and sodium dodecylbenzenesulfonate (SDBS) or cetyltrimethylammonium bromide (CTAB) and sodium octyl sulfate (SOS). The hydrophobic monomers were added to swell the bilayers of a vesicle solution and subsequently polymerized. The goal was to capture or template the vesicle structure to form hollow polymer spheres. The particles had an average radius of *ca.* 60 nm and a membrane shell less than 10 nm thick.⁹⁷ In another study, a surfactant bilayer of dimethyldioctadecylammonium chloride (DODAC) was used as a vesicle template to allow the preparation of poly(butyl methacrylate) (PBMA) hollow spheres with dimensions ranging from several nanometers up to several hundred micrometers. After extraction of the surfactant matrix, the polymer particles contracted



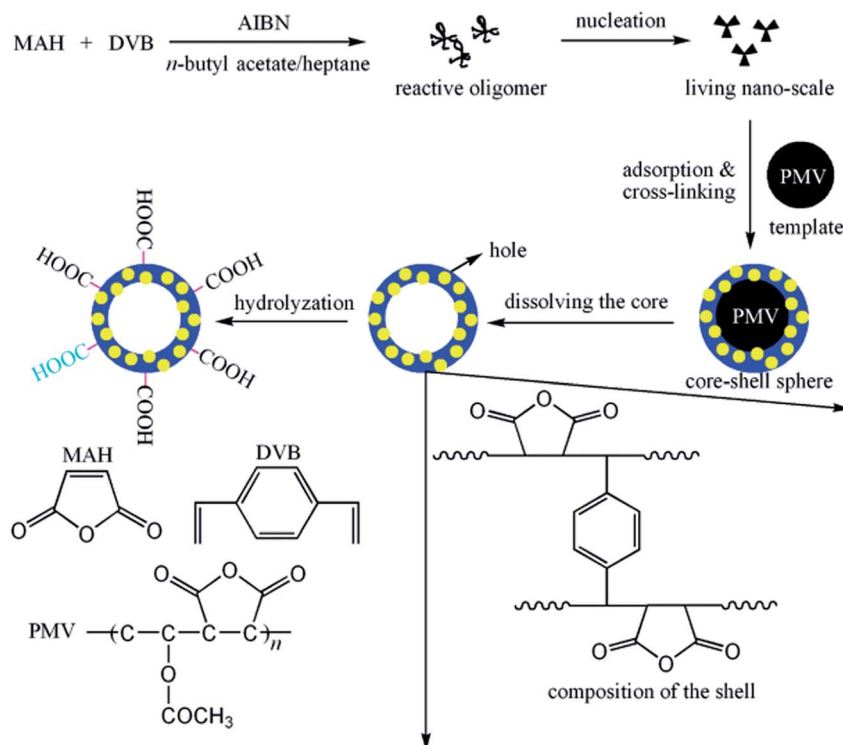


Fig. 9 Schematic of the strategy employed in the preparation of hollow spheres. Reproduced from ref. 94 with permission from the American Chemical Society.

considerably; thus, this led to the formation of hollow spheres.⁹⁸ Wang *et al.* prepared hollow ferrocenyl coordination polymer microspheres with micropores in the shells by Ostwald ripening. The formation of hollow microspheres with microporous shells resulted from the mass transfer and crystallization of small crystallites from the Ostwald ripening process. The detailed formation mechanism of hollow microspheres is depicted in Fig. 11 involving: (a) dissolution of the small particles; (b) diffusion and redeposition of the dissolved species; and (c) dissolution and outmigration of the inner crystallites.⁶⁶

Gao *et al.* prepared polymer nanocapsules embedded with noble metal nanoparticles by a template-free method, as shown in Fig. 12. Compared with other methods for fabricating nanocapsules, this method is performed in an aqueous solution without any organic solvent or template core. By this technique,

it is not necessary to prepare any uniform spherical templates, regular amphiphilic polymers, complex dendrimers, or hyper-branched polymers, and repeated absorption and washing processes are also not required.⁹⁹

3.2 Self-assembly

The self-assembly approach involves a careful selection of the experimental parameters, such as the copolymer composition, block length, ionic content, solvent, and polymer concentration, to control the morphology.¹⁰⁰ HPPs have been prepared *via* the self-assembly of amphiphilic block copolymers, rod-coil block copolymers, polymer pairs, and polymer-peptide conjugates.⁴⁸ Owing to their amphiphilic nature and molecular geometry, lipid molecules can aggregate in a dilute aqueous solution into spherically closed bilayer structures, the so-called vesicles or liposomes. It is quite reasonable that the hollow



Fig. 10 Synthesis of nanocapsules using gold nanoparticles as sacrificial templates: (1) nanoparticle assembly; (2) cross-linking; and (3) GNP core removal. Reproduced from ref. 96 with permission from John Wiley and Sons.





Fig. 11 Proposed formation process of hollow iron-based ferrocenyl coordination polymer microspheres (Fe-Fc-HCPS) by the Ostwald ripening mechanism, together with the corresponding TEM images. Reproduced from ref. 66 with permission from the American Chemical Society.

sphere morphology of these aggregates should render them suitable as precursors for the preparation of more stable nanocapsules. This can be realized using different concepts.⁶ Amphiphilic diblock copolymers possess similar properties as surfactants. They can undergo a self-assembly process to obtain diverse morphologies.⁴⁶ The self-assembly method produces HPPs in the nanometer scale with high spherical particles. This method involves shell cross-linking reactions and then removal of their core to form stable HPPs. The shell cross-linking must be conducted at a very low concentration of micelle solution to prevent the polymer micelles from aggregating; this restricts the use of the self-assembly method in industrial applications.¹⁰¹

Polymer aggregates exhibit a higher stability and durability as compared to small-molecule aggregates due to their mechanical and physical properties. Therefore, polymer self-assembly has attracted significant attention because of its potential applications in many fields, such as biomedicine, biomaterials, and microelectronics, and in the synthesis of photoelectric materials, catalysts, etc.¹⁰⁰

Jang and Ha reported the preparation of P(MMA)/cross-linked PS core/shell nanospheres *via* emulsion polymerization. The self-assembly of the triblock copolymers poly(oxyethylene)-poly(oxypropylene)-poly(oxyethylene) was used to form hollow polystyrene nanospheres. PS hollow nanospheres were obtained



Fig. 12 Proposed schematic for the formation of nanocapsules. Reproduced from ref. 99 with permission from the Royal Society of Chemistry.



by a polymer-core etching technique using methylene chloride. Fig. 13 shows the schematic of polystyrene hollow nanosphere fabrication.³³

In another study reported by He and co-workers, cage-like polymer microspheres with hollow core/porous-shell structures were prepared by the self-assembly of PS particles at the emulsion droplet interface. The swelling of the PS particles by the oil phase provides a driving force to develop hollow cores. The PS particles also serve as a porogen that will disengage automatically during polymerization, leaving behind a uniform porous shell.^{73,102} In another contribution, He and co-workers prepared a polystyrene/melamine-formaldehyde (PS/MF) hollow polymer based on a similar approach. The reserving time of emulsion after the addition of the MF prepolymer will affect the morphology of the polymer layer. It was found that many pores appeared on the particle surface with the shortening of the reserving time and finally formed a porous structure.⁷¹

3.3 Acid/alkali swelling

Acid/alkali swelling/osmotic swelling is the earliest method and dominant for obtaining submicron-sized HPPs. This method was invented by Kowalski and colleagues and commercialized by Rohm & Haas Company.¹⁰³ The typical concept is to prepare a core/shell polymer with a soft carboxyl core and rigid shell, followed by neutralizing the carboxyl group to generate a polyelectrolyte inside and then expanding the core by an osmotic swelling mechanism.³⁵ This method was subsequently improved by Vanderhoff and co-workers¹⁰⁴ and other researchers.^{105,106} The advantage of these methods comprises the formation of hollow particles with a definite morphology (particles containing one central hollow) as compared to other techniques that provide particles containing several hollows separated by polymer walls.¹³ A large number of emulsion polymerization techniques are available to produce hollow

particles by acid/alkali swelling. The unsaturated acid monomer plays an important role in the formation of the hollow structure.⁴⁹ At first, seed latex with carboxylic acid groups is prepared; then, a hard permeable polymer shell is formed on the outer surfaces of the carboxylic core polymer particles; after this, at around the glass-transition temperature (T_g) of the shell, particles containing water and polyelectrolyte are prepared by neutralizing the carboxylic acid groups of the core polymer particles using a volatile alkali.¹⁰¹

Yuan and co-workers prepared monodispersed HPPs with a core/shell monomer ratio of 1 : 8 and 1% DVB cross-linker in the shell. At first, MMA-BA-MAA copolymer emulsions were prepared as seeds; after this, the seed emulsion polymerization of MMA-MAA-DVB was carried out to prepare carboxylated core particles. Then, the hydrophobic shell of St-ACN-DVB was synthesized onto the core. Finally, the HPP particles were obtained by alkalization treatment of the core-shell latex particles using NaOH.⁵² Based on a similar approach, Zhang and co-workers prepared HPPs with an MAA/BA/MMA core and BA/St shell, followed by a stepwise alkalization treatment with ammonia (NH₄OH). The results showed that the optimum content of the cross-linker diallyl maleate (DAM) in the shell was 0.5–1.0 wt%. Monodispersed HPPs were obtained at a core/shell monomer ratio of 1 : 7.⁵³ This concept was utilized by Lee and co-workers to produce P(NIPAM-co-MAA)/Fe₃O₄ thermo-sensitive magnetic HPPs.^{54,107}

3.4 Advanced techniques to prepare hollow polymer particles

Various physical and chemical methods have been reported for the preparation of HPPs. Preparation of HPPs by emulsion polymerization *via* the encapsulation of a hydrocarbon non-solvent is a complicated process controlled by a number of thermodynamic and kinetic factors.¹⁰⁸ Electrostatic interaction,



Fig. 13 Schematic for polystyrene hollow nanosphere fabrication. Reproduced from ref. 33 with permission from the American Chemical Society.



as a driving force, is only suitable for the layer-by-layer (LbL) assembly of polymer electrolytes. Micelle formation of a block copolymer involved UV cross-linking of the shell layer and a subsequent tedious ozonolysis of the core.¹⁰⁹ However, although the abovementioned pioneering works are very interesting, the preparation processes seem to be time-consuming; for example, multistep processes are needed for the synthesis of core-shell composite particles, and then, removal of the core particles by selective dissolution or by calcination is required. Therefore, development of facile and feasible methods to prepare HPPs continues to be a great challenge to material scientists. The fabrication of monodisperse HPPs with size in the range from the micrometers to nanometers and a tailored structural shape and surface properties remains a great challenge in material science.¹¹⁰ Accordingly, in this review, advanced techniques such as electrohydrodynamic atomization (CEHDA), inverse mini-emulsion periphery RAFT polymerization (IMEPP) process, urea treatment, and a four-step approach for the preparation of HPPs have been reported.

Chang *et al.* prepared a hollow microsphere with a single hole in its shell by a CEHDA technique. Fig. 14 shows the experimental setup of the CEHDA technique consisting of a pair of concentric needles. The inner stainless steel needles with diameters in the range between 300 and 150 μm were supplied with liquid perfluorohexane (PFH), and outer needles with diameters in the range between 1100 and 685 μm were supplied with a polymethylsilsequioxane (PMSQ) solution. The mean diameter of the hollow microsphere could be varied in the range from 310 to 1000 nm by adjusting the flow rate, polymer concentration, or applied voltage. The shell thickness could be controlled by the polymer concentration. Compared with other methods, this strategy has two advantages. First, it enables the preparation of microspheres with a single surface hole at ambient temperature and pressure. Second, this green process does not require the use of surfactants or other additives and does not entail a template removal step.^{40,111}

Utama *et al.* synthesized HPPs with a hydrophilic liquid core using a one-pot approach *via* a novel IMEPP process. An inverse mini-emulsion was obtained by dissolving the stabilizer, monomer, cross-linker, and initiator in the organic phase and the lipophile in the aqueous phase. The combined mixture was then ultrasonicated, followed by polymerization in sealed glass ampoules under vacuum at 60 $^{\circ}\text{C}$. The size and stability of the droplets (prior to polymerization) were monitored using varying amounts of dispersed phase, stabilizer, lipophile, and the ultrasonication period.¹¹² Further developments by Utama *et al.* showed that the shell thickness could be conveniently tuned by adjusting the stoichiometry and monomer conversion.³⁹ Wang *et al.* prepared porous and hollow polymer particles by urea treatment. In this study, core/shell particles were synthesized by multistep seed polymerization using methyl methacrylate (MMA), acrylic acid (AA), divinylbenzene (DVB), and styrene (St), and then, a certain amount of ethanol and urea were added to the particles. The advantage of this approach as compared to that of the conventional methods is the creation of polymers with mesoporous shells. Fig. 15 shows the morphology of the particles (A) without urea treatment, and the particle has a smooth surface, and (B) after urea treatment, the particle displays a rough surface. The porous shells were created by carbon dioxide produced by the decomposition of urea and was released out from the particles.³⁶

Shi *et al.* reported a novel route to prepare HPPs decorated with gold nanoparticles (AuNPs) in the shell. The approach needed no additional reducing agent in the formation process of the AuNPs. As shown in Fig. 16, the method mainly involves four steps: (1) pretreatment of the surface of silica particles with polyvinylpyrrolidone (PVP), and PVP-treated silica particles are labeled as SiO_2/PVP ; (2) seeded polymerization of ethylene glycol dimethacrylate (EGDMA) and methacrylic acid (MAA) on the SiO_2/PVP particles, and the resultant particles are labeled as $\text{SiO}_2/\text{PVP}/\text{PEM}$, where PEM represents the poly(EGDMA-co-MAA) copolymer; (3) selective etching of SiO_2 from $\text{SiO}_2/\text{PVP}/\text{PEM}$ particles using an aqueous HF solution to generate PVP/PEM hollow particles; and (4) AuNP formation in the shell of



Fig. 14 Experimental setup for the electrohydrodynamic atomization (CEHDA) technique, consisting of a pair of concentric needles. Reproduced from ref. 111 with permission from the American Chemical Society.





Fig. 15 TEM images of the multistage polymer particle (A) without and (B) with urea treatment. Reproduced from ref. 36 with permission from Springer.



Fig. 16 Schematic for the synthesis process of PVP/PEM-Au hybrid hollow particles. Reproduced from ref. 38 with permission from Springer.

the PVP/PEM hollow particles to yield PVP/PEM-Au hybrid hollow particles.³⁸

4. Properties

The features, such as encapsulation property, surface functionality, low density, high specific area, good flow ability, optical properties, surface permeability, and opacity,^{35,101,113,114} of HPPs have been actively investigated due to their possible role in technological applications in various areas. The functionalization of HPPs could be achieved by encapsulating guest species into the hollow core, which endowed them with diverse properties.¹¹⁵ Consequently, incorporation of functional groups (*e.g.*, $-\text{COOH}$) comprising biological units by surface modification could specifically deliver the loaded guest molecules to target areas and ideally suited for *in vivo* applications.¹¹⁶ Shi *et al.* developed a charged hollow core/shell lipid-polymer-lipid hybrid nanostructure constructed with three functional features. An outer PEG-lipid layer was incorporated into the system and could enable the particles to escape immunological recognition, and improve the particle stability during circulation and slow polymer degradation and drug release. The middle poly(lactic-co-glycolic acid) (PLGA) forms a barrier through the sustained release of encapsulated small interfering RNA (siRNA) and additionally serves to carry water-insoluble drugs. Moreover, the inner positively charged hollow core, composed of cationic lipids, could encapsulate siRNA much more efficiently than PLGA alone.¹¹⁷

Nanoparticles, such as tin, gold, Fe_3O_4 , or silica NPs, could be incorporated into the interior and exterior of HPPs.⁴ The inclusion of stimuli-responsive properties into these microstructures has allowed their unique application in biomaterials, such as in anticancer drug delivery systems, because of their

increased functionality in comparison with that of conventional inert structures.¹¹⁸ Kim *et al.* prepared monodispersed HPPs embedded with magnetic Fe_3O_4 nanoparticles. The resultant magnetic hollow polymer showed excellent colloidal stability. The interior void size could be controlled by adjusting the particle size of the core polymer, and the shell thickness could be adjusted by varying the weight ratio of the iron oxide nanoparticles to the styrene monomer. These magnetic polystyrene microspheres were superparamagnetic and highly magnetized as much as 4.69 emu g^{-1} microsphere.¹¹⁹ Shi *et al.* prepared HPPs decorated with gold nanoparticles in the shell. By introducing the PVP polymer into the shell of the hollow particles, AuNPs were easily formed within the polymer shell just in the presence of a gold precursor under heating. This method will open up the possibility for the generation of a series of HPPs that can incorporate various noble metal nanoparticles (Au, Ag, Pt, *etc.*) into the shell.³⁸ Lee *et al.* synthesized P(NIPAM-MAA)/ Fe_3O_4 thermosensitive magnetic HPPs. They found that an increase in the amount of Fe_3O_4 particles would increase the saturation magnetic flux density of the magnetic HPPs.⁵⁴

Opacity or the hiding power is an important optical property of white paint coating and results from the multiple scattering of incident light due to interactions with pigment particles.¹²⁰ HPPs with excellent optical properties, such as opacity and whiteness, are utilized as an opacifying agent or a white pigment for industrial coatings. The needed structure and properties of this type of HPPs can be obtained by changing the composition of the shell, the size of particles, and the holes or kinds of internal substances.⁵³ Huang *et al.* prepared novel HPPs with both opaque and self-film forming properties. The HPPs of four-layered polymers consisted of a high carboxyl-containing soft core, a transition layer, a rigid supporting layer, and an outermost film-forming layer. The optimized



recipe with a BA/St mass ratio of 2/1, DVB content of 1 wt%, and core/film forming layer mass ratio of 1/6 was utilized in the film forming layer preparation, and the self-film forming hollow latex with the best opaque property was obtained. Accordingly, this kind of hollow polymers might have prospective applications in various fields such as in waterborne coatings, inks, paints, and biomedicine. Borthakur *et al.* prepared a series of core-shell particles containing a copolymer of P(*n*-BA-*co*-MAA-*co*-EGDMA) as the core and a copolymer of P(St-*co*-MMA) as the shell. In their studies, newly synthesized core-shell particles were used as a pigment in the emulsion paint along with titanium dioxide (TiO₂). Particles with a St : MMA ratio of 60 : 40 by percentage showed the best result in these studies. It was found that the optical properties of the paint did not change with an up to 17% reduction of TiO₂ in the pigment volume. The gloss, rock hardness, and washability of the emulsion paint also remained unchanged.⁹

Furthermore, creation of a single hole in polymer shells offers a large number of binding sites generated in the proximity of the interior surface. The target species can easily diffuse into the interior sites of hollow spheres through the open holes; this leads to a higher capacity and faster kinetics for the uptake of target species. Hollow core-porous-shells have promising application as high-loading drug carriers. This is due to the large cavity inside the HPPs and the numerous mesopores present in the shell.⁶⁴ A permeable shell can allow small molecules to penetrate the core for effective catalytic reactions and drug-loading capacity.^{64,65} Several interesting materials have been reported such as inorganic silica, in which both macropores and micro- or mesopores are incorporated into the structure.⁷¹ Guan *et al.* reported HPPs with a mesoporous shell containing small silica particles on the surface. The surface roughness of the hollow particles has technical importance in controlling the surface-associated adsorption process. The resultant hollow particle had a hydrophobic cavity and an amphiphilic surface, thus serving as a good colloidal collector (high loading) for hydrophobic contaminants in water. The particles were also found to be good particle emulsifiers probably because of the amphiphilicity of the silicate and polymeric moieties on the surface.⁷⁰

5. Recent study on hollow polymer particles

Hollow polymer particles (HPPs) have attracted wide research interest from fundamental science to prospective applications. The unique properties of HPPs have attracted attention from scientists, who are re-investigating the older systems to develop a new system using advanced techniques and methods. To simplify the complex processes of fabrication of HPPs, a one-step method was proposed by Yang *et al.*¹²¹ Other facile methods to prepare HPPs have been developed by a number of researchers recently.^{122–124} In this section, the recent studies have been reviewed.

Nakamura *et al.* prepared poly(ionic liquid) (PIL) particles with a single-hollow structure by suspension polymerization from monomer droplets. The obtained PIL hollow particle

shells had switchable properties, which could be changed from hydrophobic to hydrophilic by anion exchange using a LiBr/ethanol solution. These properties are useful for preserving encapsulated materials in the hollow part by changing the hydrophobicity of the shell particle.¹²⁵ Qin *et al.* designed and synthesized a novel type of ammonium-functionalized HPPs (HPP-NH₃⁺). Due to the presence of dual functional groups (ammonium and carboxyl groups), HPP-NH₃⁺ showed a significant pH-dependent equilibrium adsorption capacity, which increased dramatically from 59 mg g⁻¹ to 449 mg g⁻¹ as the solution pH was decreased from 9 to 2. Because of its high adsorption capacity, pH-sensitivity, easy regeneration, and good reusability, the HPP-NH₃⁺ has great potential for application in the field of water treatment, controlled drug release, and pH-responsive delivery.¹²⁶ Kitayama *et al.* studied an efficient pathway for the preparation of HPPs. In this approach, site-specific a posteriori photocrosslinking of homogeneous spherical polymer particles and subsequent removal of the particle core – the self-templating strategy – has been developed as an efficient pathway for hollow particle formation. This facile route *via* a self-templating strategy has great potential to be used as an alternative route because this route has high mass productivity and high simplicity as a result of the non-use of additional sacrificial template particles and highly toxic solvents.¹²² Zhang *et al.* prepared a novel amphiphilic diblock copolymer, PHE-MAPOSS-*b*-P(DMAEMA-*co*-CMA), *via* reversible addition-fragmentation chain transfer (RAFT) polymerization. The hollow polymeric capsules were obtained *via* etching the POSS core in the solution of hydrofluoric acid (HF). The tetraphenylporphyrin tetrasulfonic acid hydrate (TPPS)-loaded polymeric capsules demonstrated efficient photodynamic therapy (PDT) efficacy and a low dark toxicity towards the MCF-7 cells. Thus, the TPPS-loaded polymeric capsules presented potential application in PDT.¹²⁷ Nakayama *et al.* fabricated HPPs from the solvent treatment of aqueous bubble-stabilized polystyrene (PS) emulsion particles. The aqueous bubbles stabilized with PS particles could be transformed into polymeric hollow microspheres containing air by exposure to a solvent. The polymeric hollow microspheres maintained their shapes even after complete water evaporation; this confirmed that a rigid polymeric film was successfully formed.¹²⁸ Wichaita *et al.* prepared natural rubber (NR)-based hollow latex (HL) particles by a seeded emulsion polymerization of methyl methacrylate/divinyl benzene/acrylic acid (MMA/DVB/AA) monomers on the NR seed. A void was subsequently formed at an MMA/DVB molar ratio of 2.7/1 without the necessity of any residual core removal. The void cavity was enlarged when the monomer to seed (M/S) weight ratio was increased to 4/1. The large void cavity (155 ± 37 nm) of these NR-based HL particles makes them particularly suited as delivery vehicle systems.¹²⁹ Yang *et al.* developed an innovative one-step synthesis of hollow polymer particles by microsuspension polymerization of styrene (St) and methyl acrylate (MMA) with Mg(OH)₂ as the dispersant. Due to difference in the reactive activity of St and MMA, hollow polymer particles were generated at an appropriate pH and polymerization temperature.¹³⁰ Okubo *et al.* proposed an innovative easy method to synthesize hydrophobic hollow PS



particles by emulsion polymerization with an polyoxyethylene nonylphenyl ether nonionic emulsifier (Emulgen 910). The vol% (relative to the particles) of water was increased to 46% by heat treatment at 90 °C for 24 h, which was based on further water absorption and resulted in spherical hollow particles.¹³¹ Other contributions relating to HPPs have been described by a number of researchers.^{132–138}

6. Applications

Hollow polymer particles (HPPs) have attracted extensive attention in recent years due to their potential applications in the fields of coating, papermaking, and cosmetics because of their excellent characteristics of light scattering and low density;¹³⁹ moreover, they have potential applications in catalysis, biochemistry, and biomedicine due to their capacity of accommodating objects in their hollow structure.^{3,140} The application of HPPs as fillers that decrease the density of polymeric composites is a well-known technique. Another and more complicated area of HPPs is their use as white pigments in coatings and paint materials. This approach almost eliminates the possibility for using many of the conventional microspheres.^{13,141} Stimuli-responsive HPPs are very promising candidates as drug delivery systems because of their high stability for prolonged circulation in the blood stream and their ability to encapsulate bioactive drugs.¹⁴² In particular, their application provides an interesting opportunity for drug delivery in which the delivery system becomes an active participant, rather than a passive carrier, in the optimization of disease therapy.¹⁴³ When used as drug carriers, stimuli-responsive HPPs have shown a much larger drug-loading capacity than ordinary HPPs as stimuli-responsive HPPs are easier to load with drugs, particularly biomacromolecular drugs, or to release the entrapped drugs on demand at the desired target because the pores of their shells can be gated upon phase transition triggered by external stimuli.¹⁴³ However,

porous-shell HPPs are also well known for a superhigh capacity of drugs due to their porous shell structures.

Yang *et al.* prepared monodisperse pH-sensitive hollow P(MBAAm-co-MAA) microspheres with a diameter below 200 nm and a shell thickness in the range of 8–26 nm. The *in vitro* loading and release behavior studied demonstrated that the hollow P(MBAAm-co-MAA) microspheres toward the DOX drug possessed a good loading capacity (as high as 357 mg mg⁻¹) with a high encapsulation efficiency (as high as 83%). The loading and release behavior of DOX drug with P(MBAAm-co-MAA) HPPs as a reservoir was highly dependent on the pH values in the environment, which could be well interpreted by the difference in strength of the hydrogen-bonding interaction between the host P(MBAAm-co-MAA) HPPs and the guest DOX drug under different pH conditions. Furthermore, these HPPs were improved to enhance the effect of targeted delivery by adding Fe₃O₄ nanoparticles to P(MBAAm-co-MAA) HPPs. These results imply that P(MBAAm-co-MAA) hollow microspheres may have potential application as a reservoir for controlled release. The drug loading of the dual-targeting hollow P(MBAAm-co-MAA) microspheres toward the DOX drug possessed a good loading capacity (as high as 176 mg mg⁻¹) and a high encapsulation efficiency (as high as 61%) in the case of an initial DOX concentration of 230 mg mL⁻¹, whereas the releasing behavior with dual-targeting hollow microspheres as reservoirs was highly dependent on the pH values in the environment. All these results may offer a suitable way for the preparation of multi-functionalized drug carriers for tumor combination therapy.¹⁴⁰ Chiang *et al.* prepared polymeric hollow nanogels designed as carriers for intracellular triggered drug release.¹⁴⁴ Another contribution reported by Chiang *et al.* involved the use of superparamagnetic hollow nanogels as a potential guidable vehicle system for stimuli-mediated magnetic resonance imaging (MRI) and multiple cancer therapeutics. The system works by magnetic transport guidance toward the target and subsequent exposure to an alternating magnetic field. Then, the system exhibited accelerated drug release in response to both

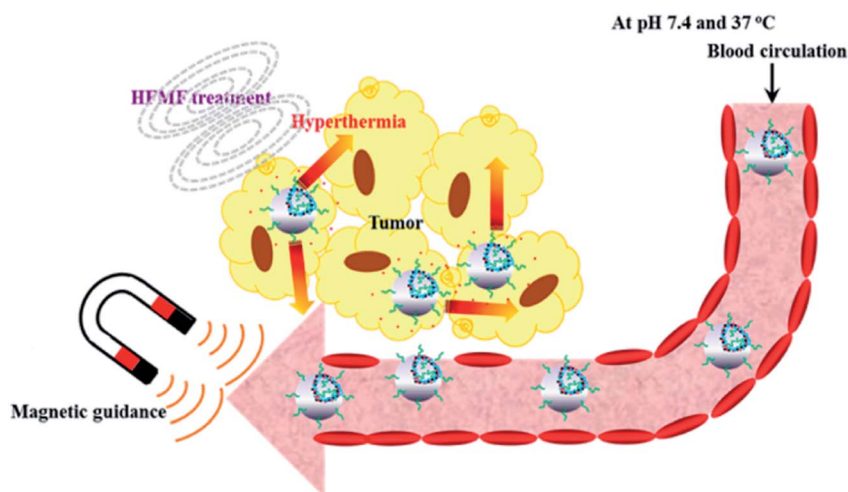


Fig. 17 Schematic hollow nanogels serving as a multifunctional anticancer theranostic platform. Reproduced from ref. 145 with permission from the American Chemical Society.



pH reduction and temperature increase as a result of a disruption of the interactions between the drug molecules and nanogels structures. Fig. 17 shows a schematic of the hollow nanogels serving as a multifunctional anticancer theranostic platform.¹⁴⁵

Ke *et al.* reported a novel smart system based on PLGA hollow microspheres (HMs) that could deliver an anticancer drug into tumor cells and quickly release the drug in an acidic organelle such as a lysosome.¹⁴⁶ Nuasen *et al.* prepared hollow latex (HL)

particles functionalized with chitosan for the removal of formaldehyde from indoor air. Chitosan and polyethyleneimine (PEI)-functionalized HL particles were conveniently fabricated by coating (P(MMA/DVB/AA)) HPPs with 5 wt% chitosan or 14 wt% PEI. Besides providing opacity and whiteness, the multilayer HL-chitosan particles could effectively remove indoor air pollutants, *i.e.*, formaldehyde gas, and, hence, would be useful in special coating applications.¹⁰

Table 3 Polymers that form HPPs and classification, techniques, and the methods

Polymer	Classification	Preparation		Ref.
		Technique	Method	
MMA/MAA/DVB core, St/ACN/DVB shell	Solid-shell	Seeded emulsion	Acid/alkali swelling	52
P(St-co-MMA-co-BA-co-MAA)	Solid-shell	Soap-free emulsion	Acid/alkali swelling	55
PBzMA	Solid-shell	LRP	Core dissolution	56
PS	Solid-shell	Dispersion	HF etching	21
PS/DVB	Solid-shell	Seeded dispersion	SaPSeP	57
P(AA-EGDMA)	Solid-shell	2-Steps polymerization	Core dissolution	3
PVAc, PMMA, PBA, PBMA	Porous-shell	Emulsion	Self-assembly	73 and 102
PAA	Porous-shell	Suspension	Core-removal	64
PEA core, silica shell	Porous-shell	Pickering emulsion	Solvent evaporation	70
Au-POMA	Porous-shell	Precipitation	Solvent evaporation	65
PS, PMMA	Single particles	Pre-synthesis, emulsion	Solvent evaporation	74
PS	Single particles	3-Steps approach	Solvent evaporation	75
POT	Single particles	Chemical polymerization	Droplet template	76
PMAA-PNIPAM	Stimuli responsive	Sol-gel process	Core removal/HF treatment	82
Fe ₃ O ₄ /SiO ₂ /PMAA/SiO ₂ /PNIPAM	Stimuli responsive	Precipitation	Template	79
MH-MIPs	Stimuli responsive	Pickering emulsion	Soft-templates	83
P(NIPAAm-co-MBAAm-co-GA)	Stimuli responsive	Inverse emulsion	Phase separation	84
Silica-PNIPAM-PNIPAM	Stimuli responsive	2-Steps approach	Core dissolution	51
P(MAH/DVB)	Solid-shell	Distillation precipitation	Hard templates	94
P(DVB-co-AA)	Solid-shell	Precipitation	Hard templates	95
PPy-CS	Solid-shell	Photoreduction	Sacrificial templates	4
St- <i>alt</i> -MA	Solid-shell	RAFT	Sacrificial templates	96
St, DVB	Solid-shell	Vesicle polymerization	Vesicle templates	97 and 98
Fe-Fc-HCPS	Porous-shell	Ostwald ripening	Self-templates	66
AG-VP	Single-particles	FRP	Template-free	99
PS	Solid-shell	Emulsion	Self-assembly/Core etching	33
PS	Porous-shell	Emulsion	Self-assembly/Emulsion droplet	73 and 102
PS/MF	Porous-shell	Emulsion	Self-assembly/Emulsion droplet	71
MMA-MAA-DVB	Solid-shell	Emulsion	Acid/alkali swelling	52
MAA/BA/MMA core, BA/St shell	Solid-shell	Emulsion	Acid/alkali swelling	53
P(NIPAAm-MAA)/Fe ₃ O ₄	Solid-shell	Emulsion	Acid/alkali swelling	54 and 107
PMSQ	Solid-shell	CEHDA		40 and 111
PMMA	Single particles	Inverse mini-emulsion	IMEPP	39 and 112
P(MMA/AA/St)	Porous-shell	Emulsion	Urea treatment	36
PVP/PEM-Au	Solid-shell	Emulsion	HF etching	38
Lipid-polymer-lipid	Solid-shell	Emulsion	Self-assembly	117
PS-Fe ₃ O ₄	Solid-shell	Emulsion	Acid/alkali swelling	119
St/BA	Solid-shell	Emulsion	Acid/alkali swelling	35
P(BA-co-MAA-co-EGDMA) core, P(St-co-MMA) shell	Solid-shell	Emulsion	Acid/alkali swelling	9
PIL	Solid-shell	Suspension	SaPSeP	125
PHEMA- <i>b</i> -P(DMAEMA-co-CMA)	Solid-shell	RAFT polymerization	HF etching	127
PS	Single particles	Emulsion	Phase separated	128
NR-based P(MMA/DVB/AA)	Solid-shell	Emulsion	Core removal	129
St/MMA	Single particles	Microsuspension	Phase separation	130
P(MBAAm-co-MAA)	Solid-shell	Distillation precipitation	Dissolution of core	140
P(NIPAM-co-AA)	Stimuli responsive	FRP	Dissolution of core	144
P(MMA/DVB/AA)/CS@PEI	Solid-shell	Emulsion	Phase separated	10



7. Conclusions and outlook

This review presents a general overview of hollow polymer particles (HPPs). The formation of single pore structures in the particles endows a high potential to encapsulate large guest molecules. Core-shell polymer particles with hollow core structures are the most developed HPPs due to their interesting morphologies, which lead to them achieve superior properties. These advanced morphologies have accelerated their exploration in a growing list of applications such as in drug delivery, cosmetics, inks, pigments, chemical reagents, the protection of biologically active species, and the removal of pollutants. The highest loading capacity of functional materials can be achieved by porous-shell structures, and the fastest loading capacity can be achieved by creating a single hole on the polymer particle surface. Magnetic-targeting stimuli-responsive HPPs have received significant attention from scientists due to their capability to work as intelligent carriers. Template-directed synthesis is the most widely used method for the formation of HPPs. However, with solid-shell HPPs, the acid/alkali swelling method is mostly used to obtain the hollow core. Pioneering works show that the preparation processes is time-consuming; therefore, development of a facile method with high-quality (*e.g.*, nonagglomerated, uniform, controlled size) HPPs is still a great challenge to material scientists. The developed methods also need to be environmentally friendly, clean, and versatile to meet the demands of modern technologies. The best example is the opaque polymers developed by Rohm and Haas (bought by Dow chemical in 2009) for filler and coating applications. However, the development of opaque polymers took 30 years before they could be commercialized. For applications, especially biomedical applications, careful studies should be conducted to reveal the toxicity of the HPPs to the human body.

A number of interesting developments have been focused in principle; however, more studies are necessary to optimize the experiments, to improve results reproducibility, and to demonstrate applications testing. Scaling-up this product to produce industrial-scale quantities for applications is expected to face critical challenges, such as uniformity, size, stability, synthesis scale, and production cost. To meet the requirements for industrial applications, future research needs to overcome the critical challenges. These aspects of HPPs synthesis are still in an early stage; this offers exciting opportunities for newcomers to this area. Table 3 summarizes the polymers that form HPPs and classifications and the preparation methods of HPPs.

Conflicts of interest

There are no conflicts to declare.

Abbreviations

AA	Acrylic acid
AAm	Acrylamide
ACN	Acrylonitrile

Ag	Argentum
AG	Alginic acid
Au	Aurum
4-AP	4-Aminophenol
4-NP	4-Nitrophenol
VP	4-Vinyl pyridine
BA	Butyl acrylate
BMA	Butyl methacrylate
CMA	Coumarin methacrylate
CS	Chitosan
CSS	Core-shell-shell
CTAB	Cetyltrimethylammonium bromide
CTAT	Cetyltrimethylammonium tosylate
CEHDA	Coaxial electrohydrodynamic atomization
DAM	Diallyl maleate
DMAEMA	2-(Dimethylamino)ethyl methacrylate
DODAC	Dimethyldioctadecylammonium chloride
DOX	Doxorubicin
DVB	Divinylbenzene
EGDMA	Ethyleneglycol dimethacrylate
EHDA	Electrohydrodynamic atomization
Fe	Ferrocenyl
Fe	Ferum
Fe ₃ O ₄	Ferric oxide
FRP	Free radical polymerization
GA	Glycidyl acrylate
GNP	Gold nanoparticles
HCPS	Hollow coordination polymer microspheres
HEMAPOSS	Polyhedral oligomeric silsesquioxane
HF	Hydrofluoric acid
HFMF	High-frequency magnetic field
HMs	Hollow microspheres
HPC	Hydroxypropyl cellulose
HP	Hollow polymer
HPMIPs	Hollow porous molecularly imprinted polymers
HPPs	Hollow polymer particles
IMEPP	Inverse mini-emulsion periphery RAFT polymerization
LBL	Layer-by-layer
LiBr	Lithium bromide
LRP	Living radical polymerization
MAA	Methacrylic acid
MAH	Maleic anhydride
MBAAm	<i>N,N</i> -methylenebisacrylamide
MC	Methylene chloride
MCF-7	Michigan Cancer Foundation-7
MIP	Molecularly imprinted polymers
MA	Methyl methacrylate
MF	Melamine-formaldehyde
Mg(OH) ₂	Magnesium hydroxide
<i>M_n</i>	Number average molecular weight
<i>M_w</i>	Weight average molecular weight
MRI	Magnetic resonance imaging
NaBH ₄	Sodium borohydride
NaOH	Sodium hydroxide
NH ₃ ⁺	Ammonium
NIPAM	<i>N</i> -isopropylacrylamide
NPs	Nanoparticles
PBA	Poly(butyl acrylate)



PBMA	Poly(butyl methacrylate)
PBzMA	Poly(benzyl methacrylate)
PCL	Polycaprolactone
PDT	Photodynamic therapy
PEA	Isopentyl acetate
PEG	Polyethylene glycol
PEI	Polyethyleneimine
PEM	P(EGDMA-co-MAA)
PFH	Liquid perfluorohexane
PIL	Poly(ionic liquid)
PLGA	Poly(lactic-co-glycolic acid)
PLLA	Poly(L-lactide)
PMMA	Poly(methyl methacrylate)
PMSQ	Polymethylsilsesquioxane
PNIPAM	Poly(<i>N</i> -isopropylacrylamide)
PNIPMAM	Poly(<i>N</i> -isopropylmethacrylamide)
POT	Poly(<i>o</i> -toluidine)
POMA	Poly(<i>o</i> -methoxyaniline)
PMV	MAH/VAc particles
PPy	Polypyrrole
PS	Poly(styrene)
Pt	Platinum
PVA	Poly(vinyl acetate)
PVP	Polyvinylpyrrolidone
RAFT	Reversible addition-fragmentation chain transfer
SaPSeP	Self-assembling of phase-separated polymer
SDBS	Sodium dodecylbenzenesulfonate
SDS	Sodium dodecyl sulfate
SEM	Scanning electron microscope
SiO ₂	Silicone dioxide
siRNA	Small interfering RNA
St	Styrene
SOS	Sodium octyl sulfate
TEM	Transmission electron microscope
TiO ₂	Titanium dioxide
TPPS	Tetraphenylporphyrin tetrasulfonic acid hydrate
TPM	3-(Trimethoxysilyl)propyl methacrylate
T _g	Glass-transition temperature
VAc	Vinyl acetate

Acknowledgements

The author would like to thank Universiti Malaysia Pahang for the financial support under University Grant (No. RDU1703286).

References

- J. Han, G. Song and R. Guo, *Chem. Mater.*, 2007, **19**, 973–975.
- D. Wu, F. Xu, B. Sun, R. Fu, H. He and K. Matyjaszewski, *Chem. Rev.*, 2012, **112**, 3959–4015.
- G. Guan, Z. Zhang, Z. Wang, B. Liu, D. Gao and C. Xie, *Adv. Mater.*, 2007, **19**, 2370–2374.
- D. Cheng, X. Zhou, H. Xia and H. S. O. Chan, *Chem. Mater.*, 2005, **17**, 3578–3581.
- D. G. Shchukin, G. B. Sukhorukov and H. Möhwald, *Angew. Chem., Int. Ed.*, 2003, **42**, 4472–4475.
- W. Meier, *Chem. Soc. Rev.*, 2000, **29**, 295–303.
- H. Bamnolker, B. Nitzan, S. Gura and S. Margel, *J. Mater. Sci. Lett.*, 1997, **16**, 1412–1415.
- A. Khan, B. Ray, J. Maiti and S. Dolui, *Pigm. Resin Technol.*, 2009, **38**, 159–164.
- L. J. Borthakur, T. Jana and S. Dolui, *J. Coat. Technol. Res.*, 2010, **7**, 765–772.
- S. Nuasaen and P. Tangboriboonrat, *Prog. Org. Coat.*, 2015, **79**, 83–89.
- A. Khan, B. Ray and S. Dolui, *Prog. Org. Coat.*, 2008, **62**, 65–70.
- R. F. Brown, C. Carr and M. E. Taylor, *Prog. Org. Coat.*, 1997, **30**, 185–194.
- V. Pavlyuchenko, O. Sorochinskaya, S. Ivanchev, V. Klubin, G. Kreichman, V. Budtov, M. Skrifvars, E. Halme and J. Koskinen, *J. Polym. Sci., Part A: Polym. Chem.*, 2001, **39**, 1435–1449.
- X. Yang, L. Chen, B. Han, X. Yang and H. Duan, *Polymer*, 2010, **51**, 2533–2539.
- X. Kang, D. Yang, P. A. Ma, Y. Dai, M. Shang, D. Geng, Z. Cheng and J. Lin, *Langmuir*, 2013, **29**, 1286–1294.
- C. Mangeney, S. Bousalem, C. Connan, M.-J. Vaulay, S. Bernard and M. M. Chehimi, *Langmuir*, 2006, **22**, 10163–10169.
- J.-H. Kim, R. D. Burnett and A. Gabriel, *J. Biomed. Nanotechnol.*, 2012, **8**, 432–438.
- E. Donath, G. B. Sukhorukov, F. Caruso, S. A. Davis and H. Möhwald, *Angew. Chem., Int. Ed.*, 1998, **37**, 2201–2205.
- S. Miao, C. Zhang, Z. Liu, B. Han, Y. Xie, S. Ding and Z. Yang, *J. Phys. Chem. C*, 2008, **112**, 774–780.
- X. Wang, H. Ji, X. Zhang, H. Zhang and X. Yang, *J. Mater. Sci.*, 2010, **45**, 3981–3989.
- X. Xu and S. A. Asher, *J. Am. Chem. Soc.*, 2004, **126**, 7940–7945.
- B. Debord, M. Alharbi, T. Bradley, C. Fourcade-Dutin, Y. Wang, L. Vincetti, F. Gérôme and F. Benabid, *Opt. Express*, 2013, **21**, 28597–28608.
- C. Markos, G. Antonopoulos and G. Kakarantzas, *IEEE Photonics Technol. Lett.*, 2013, **25**, 2003–2006.
- G. Li, Q. Shi, S. Yuan, K. Neoh, E. Kang and X. Yang, *Chem. Mater.*, 2010, **22**, 1309–1317.
- R. A. Ramli, W. A. Laftah and S. Hashim, *RSC Adv.*, 2013, **3**, 15543–15565.
- F. Caruso, R. A. Caruso and H. Möhwald, *Chem. Mater.*, 1999, **11**, 3309–3314.
- D. L. Wilcox, *Hollow and Solid Spheres and Microspheres: Science and Technology Associated with their Fabrication and Application: Symposium held November 3-December 1, 1994, Mater. Res. Soc., Boston, Massachusetts, USA*, 1995.
- A. Kowalski, M. Vogel and R. M. Blankenship, *US Pat.*, 4, 468, 498, 1984.
- R. M. Blankenship and A. Kowalski, *US. Pat.*, 4, 594, 363, 1986.
- Z. Zhang, J. Deng, J. Sui, L. Yu, M. Wan and Y. Wei, *Macromol. Chem. Phys.*, 2006, **207**, 763–769.
- J. Du and Y. Chen, *Angew. Chem., Int. Ed.*, 2004, **43**, 5084–5087.



- 32 O. Velev, K. Furusawa and K. Nagayama, *Langmuir*, 1996, **12**, 2374–2384.
- 33 J. Jang and H. Ha, *Langmuir*, 2002, **18**, 5613–5618.
- 34 M. S. Wong, J. N. Cha, K.-S. Choi, T. J. Deming and G. D. Stucky, *Nano Lett.*, 2002, **2**, 583–587.
- 35 X. Huang, L. Li, M. Xu, J. Xu, Y. Jiang and C. Kan, *J. Appl. Polym. Sci.*, 2015, **132**, 42541.
- 36 Z. Wang, Q. Hong, Y. Shi, J. Zhang, Y. Song, S. Wu and B. Yang, *Polym. Sci., Ser. B*, 2015, **57**, 600–607.
- 37 D. Hu, X. Zhang, C. Zhang, S. Ding and Z. Yang, *Polymer*, 2012, **53**, 3802–3806.
- 38 S. Shi, C. Chang, T. Wang, S. Ren, Y. Gao and N. Wang, *J. Polym. Res.*, 2015, **22**, 632.
- 39 R. H. Utama, M. H. Stenzel and P. B. Zetterlund, *Macromolecules*, 2013, **46**, 2118–2127.
- 40 M.-W. Chang, E. Stride and M. Edirisinghe, *Langmuir*, 2010, **26**, 5115–5121.
- 41 C. F. Lee, M. L. Hsu, C. H. Chu and T. Y. Wu, *J. Polym. Sci., Part A: Polym. Chem.*, 2014, **52**, 3441–3451.
- 42 M. Sasidharan and K. Nakashima, *Acc. Chem. Res.*, 2013, **47**, 157–167.
- 43 P. Panahian, M. Salami-Kalajahi and M. Salami Hosseini, *Ind. Eng. Chem. Res.*, 2014, **53**, 8079–8086.
- 44 M. Fujii, T. Shin, H. Watanabe and T. Takei, *Adv. Powder Technol.*, 2012, **23**, 562–565.
- 45 R. G. Chaudhuri and S. Paria, *Chem. Rev.*, 2012, **112**, 2373–2433.
- 46 G. L. Li, H. Möhwald and D. G. Shchukin, *Chem. Soc. Rev.*, 2013, **42**, 3628–3646.
- 47 F. Caruso, *Adv. Mater.*, 2001, **13**, 11–22.
- 48 G.-D. Fu, G. L. Li, K. Neoh and E. Kang, *Prog. Polym. Sci.*, 2011, **36**, 127–167.
- 49 W. Deng, W. Ji, Y. Jiang and C. Kan, *J. Appl. Polym. Sci.*, 2013, **127**, 651–658.
- 50 X. Huang and B. Voit, *Polym. Chem.*, 2013, **4**, 435–443.
- 51 J. Dubbert, K. Nothdurft, M. Karg and W. Richtering, *Macromol. Rapid Commun.*, 2015, **36**, 159–164.
- 52 C. D. Yuan, A. H. Miao, J. W. Cao, Y. S. Xu and T. Y. Cao, *J. Appl. Polym. Sci.*, 2005, **98**, 1505–1510.
- 53 Q. Zhang, Z. Yang, X. Zhan and F. Chen, *J. Appl. Polym. Sci.*, 2009, **113**, 207–215.
- 54 C. F. Lee, M. L. Lin, Y. C. Wang and W. Y. Chiu, *J. Polym. Sci., Part A: Polym. Chem.*, 2012, **50**, 2626–2634.
- 55 T. Tan, S. Wang, X. Li, C. Wang and Y. An, *Curr. Appl. Phys.*, 2009, **9**, 989–992.
- 56 T. K. Mandal, M. S. Fleming and D. R. Walt, *Chem. Mater.*, 2000, **12**, 3481–3487.
- 57 H. Minami, H. Kobayashi and M. Okubo, *Langmuir*, 2005, **21**, 5655–5658.
- 58 D. Chen, J. Deng, J. Liang, J. Xie, C. Hu and K. Huang, *Sens. Actuators, B*, 2013, **183**, 594–600.
- 59 W. L. Zhang, Y. D. Liu and H. J. Choi, *J. Mater. Chem.*, 2011, **21**, 6916–6921.
- 60 W. Lu, C. Wang, W. Yue and L. Chen, *Nanoscale*, 2011, **3**, 3631–3634.
- 61 J. Sun, J. Zhang, M. Zhang, M. Antonietti, X. Fu and X. Wang, *Nat. Commun.*, 2012, **3**, 1139.
- 62 R. Liu, Y. Guo, G. Odusote, F. Qu and R. D. Priestley, *ACS Appl. Mater. Interfaces*, 2013, **5**, 9167–9171.
- 63 X. Xia, D. Chao, X. Qi, Q. Xiong, Y. Zhang, J. Tu, H. Zhang and H. J. Fan, *Nano Lett.*, 2013, **13**, 4562–4568.
- 64 Y. Chen, X. Zheng, H. Qian, Z. Mao, D. Ding and X. Jiang, *ACS Appl. Mater. Interfaces*, 2010, **2**, 3532–3538.
- 65 J. Han, M. Wang, R. Chen, N. Han and R. Guo, *Chem. Commun.*, 2014, **50**, 8295–8298.
- 66 J. Huo, L. Wang, E. Irran, H. Yu, J. Gao, D. Fan, B. Li, J. Wang, W. Ding and A. M. Amin, *Angew. Chem., Int. Ed.*, 2010, **49**, 9237–9241.
- 67 J. Liu, Q. Yang, L. Zhang, H. Yang, J. Gao and C. Li, *Chem. Mater.*, 2008, **20**, 4268–4275.
- 68 S. W. Choi, Y. Zhang and Y. Xia, *Adv. Funct. Mater.*, 2009, **19**, 2943–2949.
- 69 T. Chen, P. J. Colver and S. A. Bon, *Adv. Mater.*, 2007, **19**, 2286–2289.
- 70 Y. Guan, X. Meng and D. Qiu, *Langmuir*, 2014, **30**, 3681–3686.
- 71 X. D. He, X. W. Ge, M. Z. Wang and Z. C. Zhang, *J. Colloid Interface Sci.*, 2006, **299**, 791–796.
- 72 H. Li, X. Hu, Y. Zhang, S. Shi, X. Jiang and X. Chen, *J. Chromatogr. A*, 2015, **1404**, 21–27.
- 73 X. He, X. Ge, H. Liu, M. Wang and Z. Zhang, *J. Polym. Sci., Part A: Polym. Chem.*, 2007, **45**, 933–941.
- 74 U. Jeong, S. H. Im, P. H. Camargo, J. H. Kim and Y. Xia, *Langmuir*, 2007, **23**, 10968–10975.
- 75 S. H. Im, U. Jeong and Y. Xia, *Nat. Mater.*, 2005, **4**, 671–675.
- 76 J. Han, G. Song and R. Guo, *Adv. Mater.*, 2006, **18**, 3140–3144.
- 77 L. Zha, B. Banik and F. Alexis, *Soft Matter*, 2011, **7**, 5908–5916.
- 78 R. A. Ramli, S. Hashim and W. A. Laftah, *J. Colloid Interface Sci.*, 2013, **391**, 86–94.
- 79 P. Du, T. Wang and P. Liu, *Colloids Surf., B*, 2013, **102**, 1–8.
- 80 M. Heskins and J. E. Guillet, *J. Macromol. Sci., Chem.*, 1968, **2**, 1441–1455.
- 81 L.-H. Wang, T. Wu, Z. Zhang and Y.-Z. You, *Macromolecules*, 2015, **49**, 362–366.
- 82 G. Li, C. Lei, C. Wang, K. Neoh, E. Kang and X. Yang, *Macromolecules*, 2008, **41**, 9487–9490.
- 83 J. Pan, L. Li, H. Hang, R. Wu, X. Dai, W. Shi and Y. Yan, *Langmuir*, 2013, **29**, 8170–8178.
- 84 H. Macková, D. Horák, E. Petrovský and J. Kovářová, *Colloid Polym. Sci.*, 2013, **291**, 205–213.
- 85 G. Réthoré and A. Pandit, *Small*, 2010, **6**, 488–498.
- 86 Y. Yang, Y. Chu, F. Yang and Y. Zhang, *Mater. Chem. Phys.*, 2005, **92**, 164–171.
- 87 C. Wang, J. Yan, X. Cui, D. Cong and H. Wang, *Colloids Surf., A*, 2010, **363**, 71–77.
- 88 X. W. D. Lou, L. A. Archer and Z. Yang, *Adv. Mater.*, 2008, **20**, 3987–4019.
- 89 W. Ostwald, *Lehrbuch der Allgemeinen Chemie*, Engelmann, Leipzig, Germany, vol. 2, 1896, part 1.
- 90 X. Wang, J. Feng, Y. Bai, Q. Zhang and Y. Yin, *Chem. Rev.*, 2016, **116**, 10983–11060.



- 91 E. S. Read and S. P. Armes, *Chem. Commun.*, 2007, 3021–3035.
- 92 A. Khanal, Y. Inoue, M. Yada and K. Nakashima, *J. Am. Chem. Soc.*, 2007, **129**, 1534–1535.
- 93 D. E. Bergbreiter, *Angew. Chem., Int. Ed.*, 1999, **38**, 2870–2872.
- 94 J. Deng, Y. Yu, S. Dun and W. Yang, *J. Phys. Chem. B*, 2010, **114**, 2593–2601.
- 95 G. Li and X. Yang, *J. Phys. Chem. B*, 2007, **111**, 12781–12786.
- 96 C. Boyer, M. R. Whittaker, C. Nouvel and T. P. Davis, *Macromolecules*, 2010, **43**, 1792–1799.
- 97 C. McKelvey, E. Kaler, J. Zasadzinski, B. Coldren and H.-T. Jung, *Langmuir*, 2000, **16**, 8285–8290.
- 98 J. Hotz and W. Meier, *Langmuir*, 1998, **14**, 1031–1036.
- 99 Y. Gao, X. Ding, Z. Zheng, X. Cheng and Y. Peng, *Chem. Commun.*, 2007, 3720–3722.
- 100 Y. Mai and A. Eisenberg, *Chem. Soc. Rev.*, 2012, **41**, 5969–5985.
- 101 B. Wei, S. Wang, H. Song, H. Liu, J. Li and N. Liu, *Pet. Sci.*, 2009, **6**, 306–312.
- 102 X. He, X. Ge, H. Liu, M. Wang and Z. Zhang, *Chem. Mater.*, 2005, **17**, 5891–5892.
- 103 A. Kowalski and M. Vogel, *US. Pat.*, 4, 469, 825, 1984.
- 104 E. S. Daniels, E. D. Sudol and M. S. El-Aasser, *Polymer Latexes: Preparation, Characterization, and Applications*, ACS Publications, 1992.
- 105 D. I. Lee, M. R. Mulders, D. J. Nicholson and A. N. Leadbetter, *US. Pat.*, 5, 521, 253, 1996.
- 106 G. K. Chip and A. Rudin, *US. Pat.*, 4, 863, 973, 1989.
- 107 C. F. Lee, Y. C. Wang and W. Y. Chiu, *J. Polym. Sci., Part A: Polym. Chem.*, 2013, **51**, 2880–2891.
- 108 M.-K. Park, K. Onishi, J. Locklin, F. Caruso and R. C. Advincula, *Langmuir*, 2003, **19**, 8550–8554.
- 109 S. Stewart and G. Liu, *Chem. Mater.*, 1999, **11**, 1048–1054.
- 110 Z. Deng, M. Chen, S. Zhou, B. You and L. Wu, *Langmuir*, 2006, **22**, 6403–6407.
- 111 M.-W. Chang, E. Stride and M. Edirisinghe, *J. R. Soc., Interface*, 2010, **7**, S451–S460.
- 112 R. H. Utama, Y. Guo, P. B. Zetterlund and M. H. Stenzel, *Chem. Commun.*, 2012, **48**, 11103–11105.
- 113 G. Liu, H. Wang and X. Yang, *Polymer*, 2009, **50**, 2578–2586.
- 114 S. Yang and H. Liu, *J. Mater. Chem.*, 2006, **16**, 4480–4487.
- 115 S. Malathi and S. Balasubramanian, *J. Biomed. Nanotechnol.*, 2011, **7**, 150–151.
- 116 H. Liu, N. Finn and M. Yates, *Langmuir*, 2005, **21**, 379–385.
- 117 J. Shi, Z. Xiao, A. R. Votruba, C. Vilos and O. C. Farokhzad, *Angew. Chem.*, 2011, **123**, 7165–7169.
- 118 L. Zhang, R. Guo, M. Yang, X. Jiang and B. Liu, *Adv. Mater.*, 2007, **19**, 2988–2992.
- 119 N. S. Kim and J.-D. Kim, *J. Ind. Eng. Chem.*, 2012, **18**, 1721–1729.
- 120 C. J. McDonald and M. J. Devon, *Adv. Colloid Interface Sci.*, 2002, **99**, 181–213.
- 121 J. Shen, J. Xu, Y. Hu, J. Li and C. Kan, *Colloid Polym. Sci.*, 2017, **295**, 679–688.
- 122 Y. Kitayama, K. Yoshikawa and T. Takeuchi, *Langmuir*, 2016, **32**, 9245–9253.
- 123 S. H. Park, J. Kim, W.-E. Lee, D.-J. Byun and M. H. Kim, *Langmuir*, 2017, **33**, 2275–2282.
- 124 S.-J. Park, H.-S. Lim, Y. M. Lee and K.-D. Suh, *RSC Adv.*, 2015, **5**, 10081–10088.
- 125 R. Nakamura, M. Tokuda, T. Suzuki and H. Minami, *Langmuir*, 2016, **32**, 2331–2337.
- 126 Y. Qin, L. Wang, C. Zhao, D. Chen, Y. Ma and W. Yang, *ACS Appl. Mater. Interfaces*, 2016, **8**, 16690–16698.
- 127 Z. Zhang, Y. Xue, P. Zhang, A. H. Müller and W. Zhang, *Macromolecules*, 2016, **49**, 8440–8448.
- 128 S. Nakayama, K. Fukuhara, Y. Nakamura and S. Fujii, *Chem. Lett.*, 2015, **44**, 773–775.
- 129 W. Wichaita, D. Polpanich, T. Suteewong and P. Tangboriboonrat, *Polymer*, 2016, **99**, 324–331.
- 130 Z.-k. Yang, Z.-l. Wang, Z.-m. Mao, W.-s. Li, Y.-j. Zhou, X. Liu and M. Okubo, *Colloid Polym. Sci.*, 2017, **295**, 565–572.
- 131 M. Okubo, H. Kobayashi, C. Huang, E. Miyanaaga and T. Suzuki, *Langmuir*, 2017, **33**, 3468–3475.
- 132 H. Minami, A. Kojima and T. Suzuki, *Langmuir*, 2017, **33**, 1541–1546.
- 133 N. S. J. Williams, S. A. Wheeler and R. D. Bradley, *US Pat.*, 9,546,297, 2017.
- 134 Y. Zhang, S. Costeux and T. H. Kalantar, *US Pat.*, 9, 650, 500, 2017.
- 135 N. Sudjaipraparat, C. Kaewsaneha, S. Nuasaen and P. Tangboriboonrat, *Polymer*, 2017, **121**, 165–172.
- 136 W. Deng, H.-C. Guo, G.-A. Li and C.-Y. Kan, *Chin. Chem. Lett.*, 2017, **28**, 367–371.
- 137 M. Kano, N. Yanagisawa, Y. Takahashi and Y. Kondo, *J. Fluorine Chem.*, 2017, **197**, 34–41.
- 138 D. Chen, Y. Liu, Y. Qin, L. Wang, Y. Ma and W. Yang, *Chin. J. Chem.*, 2017, **35**, 596–604.
- 139 S. Freiberg and X. Zhu, *Int. J. Pharm.*, 2004, **282**, 1–18.
- 140 X. Yang, L. Chen, B. Huang, F. Bai and X. Yang, *Polymer*, 2009, **50**, 3556–3563.
- 141 V. Pavlyuchenko, N. Byrdina, S. Ivanchev, M. Skrifvars, E. Halme, H. Laamanen and J. Koskinen, *US. Pat.*, 6, 235, 810, 2001.
- 142 S. V. Vinogradov, *Nanomedicine*, 2010, **5**, 165–168.
- 143 Z. Xing, C. Wang, J. Yan, L. Zhang, L. Li and L. Zha, *Soft Matter*, 2011, **7**, 7992–7997.
- 144 W.-H. Chiang, V. T. Ho, W.-C. Huang, Y.-F. Huang, C.-S. Chern and H.-C. Chiu, *Langmuir*, 2012, **28**, 15056–15064.
- 145 W.-H. Chiang, V. T. Ho, H.-H. Chen, W.-C. Huang, Y.-F. Huang, S.-C. Lin, C.-S. Chern and H.-C. Chiu, *Langmuir*, 2013, **29**, 6434–6443.
- 146 C. J. Ke, T. Y. Su, H. L. Chen, H. L. Liu, W. L. Chiang, P. C. Chu, Y. Xia and H. W. Sung, *Angew. Chem.*, 2011, **123**, 8236–8239.

