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Synthesis and characterization of a core–shell $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ photo-catalyst with enhanced photocatalytic activity under visible light irradiation†

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Novel core–shell structured ellipsoid-like $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ composites, with different amounts of $g\text{-C}_3\text{N}_4$, have been successfully prepared by a simple hydrothermal–chemisorption method. Their performance as photocatalysts was systematically evaluated during RhB degradation under visible light irradiation. The composite with 7 wt% $g\text{-C}_3\text{N}_4$ was found to be 7 times more efficient as a photocatalyst than pristine BiVO_4 . Its core–shell structure and activity were also found to be highly stable after it was used for 5 times in RhB degradation. The new composites were examined by various characterization techniques. The core–shell structure enhanced the contact area between the BiVO_4 core and $g\text{-C}_3\text{N}_4$ nano-sheet shell, which provided more active sites and strengthened the chemical band interaction. The thin $g\text{-C}_3\text{N}_4$ nano-sheets reduced the charge carrier transfer distance, which further suppressed the recombination of the photo-induced electron–hole pairs and therefore enhanced the photocatalytic activity of the composites. A reaction mechanism of the photocatalytic RhB degradation was proposed and discussed in detail.

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1. Introduction

Photocatalytic technology has attracted the interest of many researchers since Fujishima and Honda¹ decomposed water into H_2 on a TiO_2 electrode in 1972. It^{2–4} can serve as a solution to the global energy crisis and environmental pollution. Semiconductor photo-catalysts,⁵ such as metal oxides, sulfides, and ternary materials with appropriate band gaps and high activity as well as stability, have attracted worldwide attention. However, some of them are only active when illuminated with ultra-violet (UV) light,^{6,7} which only makes up 4% of solar-light, thus limiting their applications in reality. In addition, some photo-generated charge carriers go through recombination instead of leading to reactions, resulting in low photocatalytic activity.⁸ At present, the critical research area in photo-catalysis

is to develop a visible-light responsive, highly efficient, and chemically stable photo-catalyst.

Among the numerous semiconductor photo-catalysts examined so far, Bi-based compounds such as BiOX ($X = \text{Cl}, \text{Br}, \text{I}$),^{9–13} BiWO_6 ,¹⁴ and BiVO_4 have received the most attention. Bismuth vanadate (BiVO_4) was identified as one of the most promising photo-catalysts under visible light irradiation because of its nontoxicity, narrow bandgap energy ($E_g = 2.40$ eV), and chemical stability.¹⁵ Pure BiVO_4 had different morphologies and crystal forms.^{16,17} However, all of them are in fact relatively poor in absorbing visible lights. There have been many attempts by researchers to achieve better performance of pollution degradation with BiVO_4 . BiVO_4 with double crystal phases was obtained for more reactive species during photo-degradation,¹⁸ in addition to studies on different morphologies.¹⁹ Several reports demonstrate that $\text{Cu}_2\text{O}/\text{BiVO}_4$ p–n junction photo-catalyst,²⁰ $\text{Ag}_3\text{PO}_4/\text{BiVO}_4$,²¹ BiVO_4/rGO ,²² core–shell $\text{BiVO}_4@/\text{Bi}_2\text{O}_3$ (ref. 23) and $\text{BiVO}_4@/\text{MoS}_2$,²⁴ achieved better stability and higher photocatalytic degradation efficiency. However, according to previous studies, the increase of photo-catalysis activity was limited and the morphology of composites was rarely changed.

$g\text{-C}_3\text{N}_4$ (graphene-like carbon nitride) has become one of the most promising novel photocatalytic materials. It is a π -conjugate 2D layered material that is non-toxic, metal-free, has a desirable band gap position as well as excellent thermal and chemistry stability, and a simple product synthesis.^{25,26} It was first studied to split water for hydrogen production by Wang

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*et al.*²⁷ Unfortunately, pure g-C₃N₄ has a high recombination rate of photo-excited charges, resulting in limited photo-degradation efficiency. Many reports have combined g-C₃N₄ with other suitable semiconductors to produce novel composite photo-catalysts. TiO₂/g-C₃N₄,^{28,29} CdS/g-C₃N₄,^{30,31} AgX/g-C₃N₄ (X = Br, I),³² g-C₃N₄/AgVO₃,³³ CeO₂/g-C₃N₄ (ref. 34) and g-C₃N₄/Bi-based composites^{35–37} have been studied and all had higher photocatalytic efficiency both in pollution degradation and hydrogen production. g-C₃N₄ nano-sheets³⁸ had better performance when compared with g-C₃N₄ polymer. With the protection of g-C₃N₄ nano-sheets, novel core-shell Ag₃PO₄@g-C₃N₄,³⁹ CdS@g-C₃N₄ nanowire, ZnO/g-C₃N₄ (ref. 40) photo-catalyst showed enhanced photocatalytic activity.

Several reports have shown various combinations of BiVO₄ and g-C₃N₄ including BiVO₄/g-C₃N₄ Z-scheme photo-catalyst,⁴¹ sulfur-doped g-C₃N₄/BiVO₄,⁴² and g-C₃N₄/BiVO₄.^{43–45} These compositions had impressive activity in water oxidation, pollution degradation, and CO₂ reduction. Many attempts have been made to produce composite photo-catalysts with improved degradation efficiency including Z-scheme photo-catalysts, heterojunctions, and core-shell structures.⁴⁶ As far as we know, there are no reports on the synthesis and application of core-shell BiVO₄@g-C₃N₄ for pollution degradation. Herein, we coated thin g-C₃N₄ nano-sheets onto ellipsoid BiVO₄ by simple stirring-chemisorption while varying the wt% of g-C₃N₄ to prepare different BiVO₄@g-C₃N₄ (x wt%). The photocatalytic activity is assessed by decomposing rhodamine B under visible light. The influence of g-C₃N₄ content is investigated and discussed. A possible photo-degradation mechanism is investigated based on the results of characterization and activity evaluation.

2. Experimental

All chemicals used in the experiments are analytical grade and were used without further purification.

2.1. Synthesis of g-C₃N₄ and g-C₃N₄ nano-sheets

g-C₃N₄ was synthesized through thermal condensation.⁴⁷ 5 g of melamine were put into an aluminum crucible with a cover on and subsequently heated in a muffle furnace at a rate of 5 °C min⁻¹ until the temperature reached 550 °C where it was maintained for 4 hours. After cooling down to the room temperature, the light-yellow g-C₃N₄ products were ground and collected.

To prepare the g-C₃N₄ nano-sheets, 0.2 g g-C₃N₄ powder was ultra-sonicated in 1000 mL distilled water for 24 hours followed by centrifugation to remove the un-exfoliated g-C₃N₄, resulting in g-C₃N₄ nano-sheets dispersed solution. The concentration of the g-C₃N₄ nano-sheet suspension is about 0.07 g L⁻¹.

2.2. Synthesis of BiVO₄

BiVO₄ was synthesized through a hydrothermal method.⁴⁸ In this process, bismuth nitrate (Bi(NO₃)₃·5H₂O) and ammonium vanadate (NH₄VO₃) were used as the bismuth and vanadate sources. 1.98 g Bi(NO₃)₃·5H₂O were dissolved in 40 mL ethylene

glycol, forming a homogeneous milk-white solution under magnetic stirring. Then, 0.98 g of NH₄VO₃ were dissolved in 40 mL hot water, forming an orange-yellow transparent solution. The above NH₄VO₃ solution was added into the Bi(NO₃)₃ solution dropwise under vigorous stirring and incubated for 30 min. The resulting solution was then transferred to 100 mL Teflon-lined stainless steel autoclave and maintained at 100 °C for 12 h. After cooling the reactor down to room temperature, the yellow precipitates were collected and washed three times with distilled water and ethyl alcohol. Finally, the products were dried at 80 °C in electric oven for 6 h and a yellowish BiVO₄ powder was obtained.

2.3. Synthesis of BiVO₄@g-C₃N₄ core-shell composite photo-catalyst

BiVO₄@g-C₃N₄ core-shell composite photo-catalysts were synthesized by vigorous-stirring chemisorption (Fig. 1). 0.2 g BiVO₄ were added into measured amounts of g-C₃N₄ nano-sheets suspension (0.07 g L⁻¹) followed by ultra-sonication for 60 minutes and stirring for 24 h. Various weight ratios of g-C₃N₄ to BiVO₄@g-C₃N₄ (x wt%) (x = 1, 3, 5, 7, 10) were prepared by tuning the volume of g-C₃N₄ nano-sheets between 40 mL, 90 mL, 150 mL, 200 mL, and 300 mL, respectively.

2.4. Characterization

The crystal structures and phase data for the prepared BiVO₄@g-C₃N₄ composite samples were detected by X-ray diffraction (XRD) using a Rigaku D/MAX2500 PC diffractometer with Cu-Kα radiation and an operating voltage and current of 40 kV and 100 mA, respectively. The morphology of the core-shell BiVO₄@g-C₃N₄ samples was investigated by scanning electron microscope (SEM) (Hitachi, S-4800) and transmission electron microscopy (TEM) (JEOL Ltd., JEM-2010). Energy dispersive X-ray spectroscopy (EDX) displayed the distribution of the composed elements. UV-vis light (UV-vis) diffuse reflectance spectra were recorded on a UV-vis spectrometer (Puxi, UV1901) with BaSO₄ used as the reference, based on the data, the band-gap energy of the samples could be concluded. A similar trend occurs when calculating the band-gap energy (E_g) according to the follow formula:

$$(\alpha h\nu)^n = A(h\nu - E_g)$$



Fig. 1 The schematic illustration for BiVO₄@g-C₃N₄ synthesis process.



where α , h , ν , E_g and A respectively represent the absorption coefficient, Planck's constant, light frequency, band gap and a constant. The index, n , depends on the material and has a value of $n = 2$ for a direct-gap semiconductor and $n = 0.5$ for an indirect-gap semiconductor. The Fourier transform infrared spectra (FTIR) of the samples were recorded using an IR Vertex70 FTIR spectrometer from 4000–400 cm^{-1} . The chemical composition and the chemical states of the photo-catalysts were analyzed by X-ray photoelectron spectroscopy (XPS) using a XSAM 800 apparatus. Photoluminescence (PL) was performed on a fluorescence and phosphorescence lifetime and steady-state spectrometer (Edinburgh instruments FLSP 920, LED 360 nm). Electrochemical and photo-electrochemical measurements were performed in a constructed three electrode quartz cell system containing 0.5 M Na_2SO_4 electrolyte solution. Pt sheet was used as a counter electrode, $\text{Hg}/\text{Hg}_2\text{Cl}_2/\text{sat}$ were used as reference electrodes, and a thin film of obtained samples on indium-tin-oxide (ITO) was used as the working electrode for investigation. The photo-electrochemical experimental results were recorded with a CHI 660B electrochemical system.

2.5. Photocatalytic activity evaluation

The photocatalytic reaction of $\text{BiVO}_4/\text{g-C}_3\text{N}_4$ composite photo-catalyst under visible light was carried out by photodegradation of RhB. RhB is a very stable dye and has been used widely in evaluating the performance of numerous photo-catalysts under visible light irradiation. The photocatalytic reactions were carried out in quartz tubes (Nanjing, Xujiang), with a 3 cm distance from the light source. Visible light was supplied by a 400 W halogen lamp with cut-off filters for $\lambda \geq 420$ nm. 100 mg of the catalyst powder was dispersed into 50 mL of 10 mg L^{-1} RhB solution. The temperature of the reaction mixture was maintained at 25 ± 2 °C using a recirculating cooling water system. The suspension was stirred for 30 min in the dark to make ensure that the absorption–desorption equilibrium between the photo-catalyst and the RhB had been reached, at which point the light was turned on. 5 mL of the sample were taken out at constant time intervals during the reaction process over the course of 60 minutes. The samples were then centrifuged for 6 minutes at 10 000 rpm to remove the photo-catalyst powder. The supernatant was analyzed by UV-vis (TU-1901, Puxi) spectroscopy and the absorbance (A) was used to determine the concentration of RhB according to the Beer–Lambert law. The degradation efficiency of RhB was calculated by the following formula:

$$\text{Degradation (\%)} = \frac{C_0 - C}{C_0} \times 100\%$$

where C_0 represents the original concentration of RhB and C represents the concentration of RhB at time “ t ”.

3. Result and discussion

3.1. XRD

XRD analysis was applied to evaluate the crystal phase of BiVO_4 and $\text{BiVO}_4/\text{g-C}_3\text{N}_4$ composites. Fig. 2 shows the XRD patterns

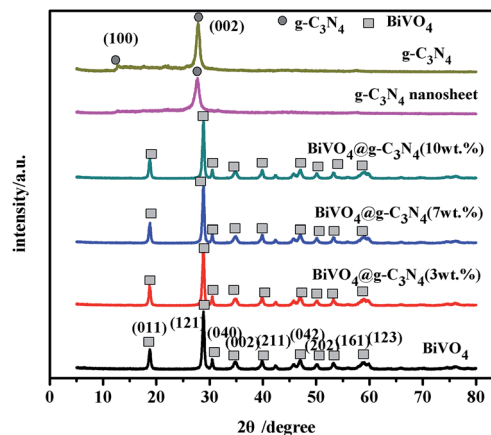


Fig. 2 XRD patterns of $\text{g-C}_3\text{N}_4$, BiVO_4 and $\text{BiVO}_4/\text{g-C}_3\text{N}_4$ with different $\text{g-C}_3\text{N}_4$ coated amounts.

of $\text{g-C}_3\text{N}_4$, BiVO_4 , and $\text{BiVO}_4/\text{g-C}_3\text{N}_4$ with different coated amounts of $\text{g-C}_3\text{N}_4$ nano-sheets. For the pure $\text{g-C}_3\text{N}_4$, it can be clearly seen that the peaks at 12.8° and 27.8° correspond to the (100) and (002) diffraction planes (JCPDS 87-1526), respectively, which are characteristic of graphitic materials and fit well with reported literatures.⁴⁹ Where the (100) diffraction peak at 12.8° represents the in-plane structural packing motif of tri-s-triazine units, and the peak at 27.8° is indexed as (002) peak of the stacking of conjugated aromatic system, and the distance of interlayer spacing is 0.320 nm.⁵⁰ The XRD pattern of $\text{g-C}_3\text{N}_4$ nano-sheet obtained by ultra-sonicated shows only one obvious peak of (002), and the (002) diffraction peak of $\text{g-C}_3\text{N}_4$ nano-sheet is wider and weaker, indicating the stacking of interlayer structure is broken and the exfoliated $\text{g-C}_3\text{N}_4$ nano-sheet is obtained.^{39,51} As for BiVO_4 , the observed peaks at 18.67° , 28.82° , 30.55° , 35.22° , 39.78° , 47.31° , 50.31° , 53.31° and 58.53° correspond to the (011), (121), (040), (002), (211), (042), (202), (161) and (123) crystal planes of m- BiVO_4 (JCPDS 14-0688).^{17,48,52,53} The characteristic diffraction peaks were sharp, proving the better purity and crystallinity of BiVO_4 . However, for the $\text{BiVO}_4/\text{g-C}_3\text{N}_4$ (x wt%) ($x < 10$) samples, only diffraction peaks of BiVO_4 were observed with no evidence of the existence of $\text{g-C}_3\text{N}_4$. This is likely due to small amounts of $\text{g-C}_3\text{N}_4$ nano-sheets being coated as well as the low diffraction intensity²² of $\text{g-C}_3\text{N}_4$ in the composites. The existence of $\text{g-C}_3\text{N}_4$ can be shown through EDS, FT-IR and XPS, which is discussed later.

3.2. SEM & TEM analysis

The morphology and structure of the as-synthesized photo-catalysts were characterized by both scanning electron microscopy (SEM) and transmission electron microscopy (TEM). As shown in Fig. 3(a), bulk $\text{g-C}_3\text{N}_4$ had a layered structure, and the average size was about 2–3 μm , while the $\text{g-C}_3\text{N}_4$ nano-sheets treated by ultrasonication were much looser (Fig. 3(b)). The BiVO_4 precursor was a regular ellipsoid about 500 nm in diameter and 1–2 μm in length (Fig. 3(c)). For the $\text{BiVO}_4/\text{g-C}_3\text{N}_4$ (7 wt%) composites seen in Fig. 3(d) and (e), the thinner $\text{g-C}_3\text{N}_4$ nano-sheets can be seen coated on the ellipsoid-like



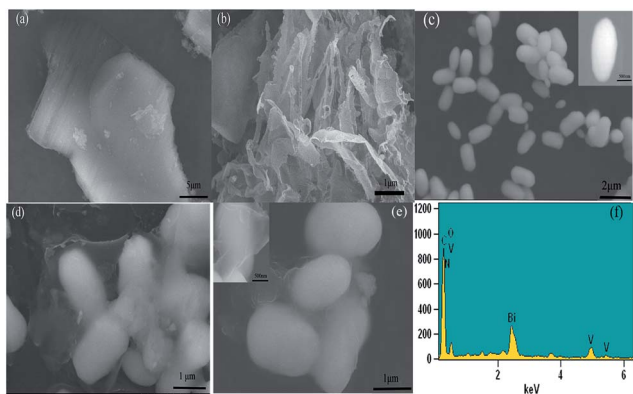


Fig. 3 SEM images of (a) $g\text{-C}_3\text{N}_4$; (b) $g\text{-C}_3\text{N}_4$ nano-sheets; (c) BiVO_4 ; (d) and (e) $\text{BiVO}_4@g\text{-C}_3\text{N}_4$; (f) the EDS for $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (7 wt%).

BiVO_4 , the core-shell structure enlarges the contact area between the BiVO_4 and $g\text{-C}_3\text{N}_4$ nano-sheet, which benefits the transfer of photo-excited charge and enlarges the light absorption areas. More details about the core-shell structure will be discussed in the TEM images. The EDS analysis (shown in Fig. 3(f)) proves the existence of the elements Bi, V, O, C and N, which supports the existence of $g\text{-C}_3\text{N}_4$.

TEM characterization of the prepared samples was performed in order to further study the details of the morphology and microstructure of the pure $g\text{-C}_3\text{N}_4$, the BiVO_4 and the $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (Fig. 4). The $g\text{-C}_3\text{N}_4$ exfoliated by ultrasound had a thin nano-sheet morphology, just like the thin film with the soft edges, the wrinkled layer with some stacking layers represents the soft nanosheets (Fig. 4(a)). Fig. 4(b) shows that the pure BiVO_4 was ellipsoid-like with tiny bulges on its rough surface, the average size is 500 nm in diameter. The TEM images of the $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ samples are shown in Fig. 4(c) and



Fig. 4 TEM images of the prepared samples (a) $g\text{-C}_3\text{N}_4$ nano-sheets; (b) the pristine BiVO_4 ; (c) $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (the inner figure showed the SAED pattern of BiVO_4); (d) the structure of core-shell $\text{BiVO}_4@g\text{-C}_3\text{N}_4$.

(d). The dark black area surrounded by a gray edge represents the BiVO_4 core protected by the $g\text{-C}_3\text{N}_4$ nano-sheet shell. The thickness of $g\text{-C}_3\text{N}_4$ shell was about 7 nm (light-yellow lined areas). The obtained core-shell structure strengthens the tight contact between BiVO_4 and $g\text{-C}_3\text{N}_4$ nano-sheets, differs from the point contact of the loaded composites, the surface contact of core@shell structure enlarges the contact area, more contact areas does help to strengthen the synergetic effect between the π -conjugated $g\text{-C}_3\text{N}_4$ and inner BiVO_4 . This not only benefits the area of light absorption, but also improves the chemical band interaction of BiVO_4 and $g\text{-C}_3\text{N}_4$, enhancing the separation efficiency of photo-induced charge, with the higher photo-degradation efficiency, this will be studied in the following discussion. The SAED pattern of BiVO_4 core indicates high crystallinity of BiVO_4 and is in agreement with the XRD analysis.

3.3. FT-IR spectra analysis

FT-IR spectra were used to characterize the special chemical bands of the prepared samples. Fig. 5 shows the FT-IR spectra of $g\text{-C}_3\text{N}_4$, BiVO_4 and $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (7 wt%). For the pure BiVO_4 , the peak at 698 cm^{-1} represents the Bi-O bond while the 733 cm^{-1} stretching peak corresponds to $\nu_3(\text{VO}_4)$.⁵⁴ Both of them are typical peaks of BiVO_4 .⁵⁵ The peaks at around 1400 cm^{-1} can be

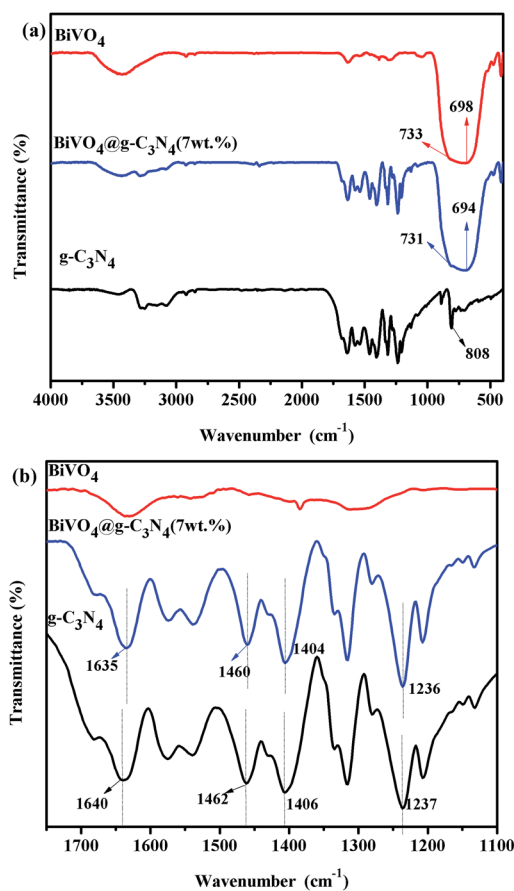


Fig. 5 (a) FTIR spectra of the prepared $g\text{-C}_3\text{N}_4$, BiVO_4 and $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ photo-catalysts; (b) partial enlarged details of FTIR spectra for $g\text{-C}_3\text{N}_4$, BiVO_4 and $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ between $1100\text{--}1750\text{ cm}^{-1}$.



attributed to the absorption of atmospheric carbon dioxide, while the peak at 1650 cm^{-1} is characteristic of residual water on the surface.⁵³ The FT-IR spectra of $g\text{-C}_3\text{N}_4$ shows the molecular structure of the compound: there is a peak at 808 cm^{-1} that is attributed to the C–N hetero-cycle, there are bands at 1237 cm^{-1} , 1315 cm^{-1} , 1406 cm^{-1} representing the aromatic C–N stretching peaks, and a peak at 1640 cm^{-1} representing the C=N stretching. These peaks are consistent with those previously reported in literature.²⁵ The FT-IR spectra of core-shell $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ shows both the typical stretching peaks of BiVO_4 and $g\text{-C}_3\text{N}_4$. The bands of BiVO_4 appear at 694 cm^{-1} and 731 cm^{-1} and are slightly red shifted when compared with pure BiVO_4 . Likewise, the peaks of $g\text{-C}_3\text{N}_4$ appear at their typical positions but with a slight red shift, as seen in the enlarged portion of the FT-IR spectra in Fig. 5(b). The band strengths are weakened in the composite photocatalyst, which indicates the increased stretching of the C–N, C=N, Bi–O and V–O bonds and a more widely π -conjugated system appears between the $g\text{-C}_3\text{N}_4$ nano-sheets and BiVO_4 core.^{40,56} The core-shell structure induced modified interactions between chemical bands, a covalent bond between $g\text{-C}_3\text{N}_4$ and BiVO_4 was formed, which in turn resulted in an improvement in the transfer and separation efficiency of photo-induced charges and increased photocatalytic degradation efficiency.

3.4. UV-vis analysis

UV-vis (ultraviolet-visible) absorption spectra were acquired in order to study the optical properties of the prepared samples and to further prove their photo-catalytic activity under visible light. The UV-vis of $g\text{-C}_3\text{N}_4$, BiVO_4 and $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ are shown in Fig. 6(a). For pure $g\text{-C}_3\text{N}_4$, the edge of the absorption spectra is at about 470 nm and shows strong absorption both in the ultraviolet and partial visible regions. The UV-vis absorption spectra of the $g\text{-C}_3\text{N}_4$ nano-sheets shows a blue shift of about $10\text{--}20\text{ nm}$ in the absorption edge. As for pure BiVO_4 , a stronger and wider absorbency is seen in visible light range. For the $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ composites photo-catalyst, the absorption edge is red shifted significantly compared to $g\text{-C}_3\text{N}_4$, and wider light absorption is seen in the visible region. This indicates that $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ may be able to absorb more light to generate electron-hole pairs. However, there was no obvious change in absorbance range between the BiVO_4 and the composite photo-catalysts with differing $g\text{-C}_3\text{N}_4$ coating amounts. $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ is a direct-gap semiconductor and uses a value of $n = 2$. According to the transformed Kubelka–Munk function, the band-gap energy can be estimated by the intercept of the tangent to the plot of $(\alpha h\nu)^2$ vs. $(h\nu)$. The results obtained from these plots are shown in Fig. 6(b). The band-gap energy of $g\text{-C}_3\text{N}_4$ was found to be 2.78 eV , whereas $g\text{-C}_3\text{N}_4$ nano-sheet has a larger energy of 2.82 eV which might the effect of quantum confinement between the tiny nano-sheets and the larger light absorption area of the nano-sheet. The calculated band-gap energies of BiVO_4 and $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (3, 5, 7, 10 wt%) are both 2.45 eV , because of the little change of the edge of the absorption spectra among the samples. The effects of core@shell structure on composites photo-catalyst activity will be discussed further.



Fig. 6 (a) UV-vis diffuse reflectance spectra of prepared photo-catalysts; (b) plot of $(\alpha h\nu)^2$ vs. energy ($h\nu$) and band-gap energy of the $g\text{-C}_3\text{N}_4$, $g\text{-C}_3\text{N}_4$ nano-sheet, BiVO_4 , $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ samples.

3.5. XPS spectra analysis

X-ray photoelectron spectroscopy (XPS) was performed to detect the surface chemical composition and chemical states of the prepared BiVO_4 and $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (7 wt%) samples. Fig. 7(a) shows the survey scan XPS spectrum of the two samples. Bi 4f, V 2p, O 1s, as well as C 1s, N 1s for the composite photo-catalyst are seen, proving that the photo-catalyst included all of the elemental components, which is in agreement with the EDS analysis results. In the core-shell $\text{BiVO}_4@g\text{-C}_3\text{N}_4$, the XPS spectra showed C 1s (Fig. 7(b)), and the binding energies of 284.8 eV and 288.05 eV are representative of the adsorbed carbon on the surface and the C atom of sp^2 -hybridized N–C=N bond,⁵⁷ respectively. Three peaks in the N 1s binding energy area are clearly seen (Fig. 7(c)). The peak at 397.43 eV is due to C=N–C bonds, and the other two peaks at about 398.92 eV and 401.11 eV attributed to the N atoms of N–(C)₃ bond and H–N–C bond,²⁵ respectively. For the composite photo-catalyst, the binding energies of these C 1s and N 1s peaks are typical in $g\text{-C}_3\text{N}_4$. In order to further study the interaction between the chemical states and core-shell structure, the XPS spectra of Bi, V and O were obtained and are depicted in Fig. 7(d)–(f). For bismuth, the typical peaks at 158.78 eV and 164.08 eV were the binding energies of Bi 4f_{7/2} and Bi 4f_{5/2}, respectively, which indicates the existence of Bi^{3+} in BiVO_4 . As for $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (7 wt%), the binding energy of Bi^{3+} was found to be slightly





Fig. 7 (a) All XPS spectra of BiVO_4 and $\text{BiVO}_4@g\text{-C}_3\text{N}_4$; XPS spectra of $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (b) C 1s; (c) N 1s; XPS spectra of BiVO_4 and $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (d) Bi 4f; (e) V 2p; (f) O 1s.

increased (158.81 eV and 164.25 eV). For V 2p orbital, the binding energy peaks at 523.8 eV and 516.43 eV of V $2p_{1/2}$ and V $2p_{3/2}$ represent the V^{5+} in BiVO_4 . Similarity, the peaks of V $2p_{1/2}$ and V $2p_{3/2}$ for $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (7 wt%) shift to higher binding energy at 524.16 eV and 516.75 eV (Fig. 7(e)). The XPS spectra of O 1s show three main peaks: the peak at 529.42 eV can be assigned to the lattice oxygen in BiVO_4 , while the peaks at 530.14 eV and 532.25 eV are signals from hydroxyl groups and absorbed water on the surface of the sample. In addition, three forms of oxygen are found in the composite, the binding energies of O 1s of $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (7 wt%) are larger than those found in pure BiVO_4 . The interaction of shell $g\text{-C}_3\text{N}_4$ nano-sheets and core BiVO_4 is confirmed by these increased binding energies,⁴⁰ the inner interaction between shell and core induces to the shift of Bi 4f, V 2p and O 1s orbitals.⁵⁸ The positive shift of binding energy means the increase of the electron density around the Bi, V, O, indicating a stronger chemical band of core-shell photo-catalyst,⁴⁵ not the simple physical mixture of BiVO_4 and $g\text{-C}_3\text{N}_4$. The π -conjugated shell $g\text{-C}_3\text{N}_4$ nano-sheets strengthens the chemical bonds of BiVO_4 . The synergetic effect between the core@shell structure forms a wider π -conjugated system, the chemical interaction is strong enough to promote better separation of photo-excited electron-

hole pairs, also the larger contact area does help to transfer the separated charge, effectively increasing the photo-catalysis degradation efficiency. This results support the conclusion of the FT-IR analysis.

3.6. Photocatalytic activity evaluation

Rhodamine B was employed in order to investigate the photocatalytic degradation rate of the $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ samples under visible light radiation ($\lambda \geq 420$ nm). A blank experiment without any photocatalyst powder was carried out in order to quantify background degradation of RhB. The results of all experiments under irradiation for 60 min are shown in Fig. 8(a), including the blank and dark experiments as well as the RhB photo-degradation by $g\text{-C}_3\text{N}_4$, BiVO_4 , and $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (x wt%). The highest degradation activity is seen when the composite contains 7% $g\text{-C}_3\text{N}_4$. An increase in activity is seen as the $g\text{-C}_3\text{N}_4$ concentration increases up to 7%, but a decrease is seen beyond this concentration at 10%. The $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (7 wt%) degraded about 90% of the RhB, a result that is 9 times greater than pure BiVO_4 (9%) and 1.3 times greater than pure $g\text{-C}_3\text{N}_4$ nano-sheets. When the $g\text{-C}_3\text{N}_4$ content was increased to 10%, the degradation efficiency of $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (10 wt%) decreased to 80%. The kinetic constants for the different compositions of $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ were calculated (Fig. 8(b)) by assuming that the degradation rate fit the pseudo-first-order reaction model ($\ln(C_0/C_t) = kt$). When the $g\text{-C}_3\text{N}_4$ amounts were 1%, 3%, 5%, 7%, 10%, the values of k were 0.00435 min^{-1} , 0.00761 min^{-1} , 0.00996 min^{-1} , 0.03971 min^{-1} and 0.02661 min^{-1} , respectively. The coated amount of $g\text{-C}_3\text{N}_4$ nano-sheets had a great effect on the photocatalytic activity. When the $g\text{-C}_3\text{N}_4$ nano-sheet content was less than 7 wt%, the contact area of the core-shell structure was not large enough to promote photo-degradation. However, excessive coating with $g\text{-C}_3\text{N}_4$ nano-sheet might lead to resistance for the separation of electron-hole pairs and transfer. With the optimal coating amount of 7%, the larger light absorption area and π -conjugation promoted the best photocatalytic degradation efficiency, which agrees with the FT-IR and XPS results shown previously.

To further prove the effectiveness of the composite photocatalyst, a comparison of $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (7 wt%) and mechanically mixed $\text{BiVO}_4/g\text{-C}_3\text{N}_4$ (7 wt%) was performed and the results are shown in Fig. 8(c). The M- $\text{BiVO}_4/g\text{-C}_3\text{N}_4$ (7 wt%) had a similar photocatalytic performance as pure $g\text{-C}_3\text{N}_4$ nano-sheets, and a lower performance than $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (7 wt%). This result shows that the core-shell structure plays an important role in promoting photocatalytic activity. Fig. 8(d) shows the whole spectra scan of RhB during the photocatalytic degradation progress using $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (7 wt%), the descending peaks of RhB for other photocatalyst samples are shown in the ESI (Fig. S1†), the inner line of Fig. 8(d) is the wavelength shift of the RhB spectra under the photo irradiation of Fig. 8(d). As the symbols in red shows, the major absorption peaks of the RhB showed a blue shift from 553 nm to 497 nm gradually, the blue shift of 56 nm indicating that the RhB was mineralized by de-ethylation. De-ethylation of RhB (N,N,N',N' -tetra-ethylated rhodamine molecule, $\lambda_{\text{max}} = 554$ nm) has the





Fig. 8 (a) The comparison of photocatalytic activity for $g\text{-C}_3\text{N}_4$, BiVO_4 and $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (x wt%) (x = 1, 3, 5, 7, 10); (b) the comparison of first-order kinetic constants of different photo-catalysts; (c) the comparison of photocatalytic activity for $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (7 wt%) and M- $\text{BiVO}_4/g\text{-C}_3\text{N}_4$ (7 wt%); (d) the whole spectral scanning of RhB during the photo-catalytic degradation for $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (7 wt%).

characteristic peak shifted to the blue region, rhodamine, $\lambda_{\text{max}} = 497 \text{ nm}$, the conjugated structure of stable RhB was destroyed^{59,60} photocatalytically.

3.7. Trapping experiment & repeating experiment

It is well known that photo-excited electrons and holes have strong reduction and oxidation properties. This allows them to react with absorbed O_2 on the catalyst surface and generate reactive species during the photo-degradation process, such as $\cdot\text{O}_2^-$, h^+ and $\cdot\text{OH}$. In order to investigate the photocatalytic mechanism of $\text{BiVO}_4@g\text{-C}_3\text{N}_4$, a trapping experiment was designed to detect and quantify these reactive species. Benzoquinone (BQ) was added to the reaction system as a scavenger of $\cdot\text{O}_2^-$, isopropanol (IPA) was added to trap the $\cdot\text{OH}$ radicals, and the ethylene diamine tetra-acetic acid disodium salt (EDTA-2Na) was used as a h^+ scavenger. The trapping experiments were maintained under the same reaction conditions as the photocatalytic activity evaluation. As seen in Fig. 9(a), the presence of IPA had little effect on the degradation rate of RhB, indicating that the $\cdot\text{OH}$ radical was not the main reactive species. The degradation rate was greatly suppressed with the addition of EDTA-2Na, indicating that h^+ was an important species in the degradation of RhB. O_2^- also played an assistant role in RhB degradation. These results provide the framework for a possible photo-degradation mechanism.

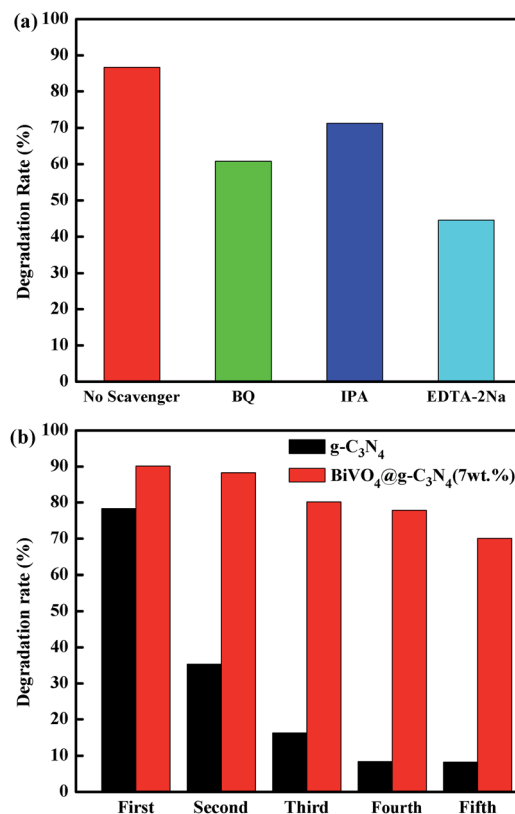


Fig. 9 (a) The comparison of degradation rate with addition of different scavengers; (b) five cycle running of photo-degradation RhB over $g\text{-C}_3\text{N}_4$ and $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (7 wt%).



As a high-performance photo-catalyst, it is important to evaluate the stability of prepared core-shell $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ composite under visible light irradiation. In order to better study the stability and recyclability of the composite photo-catalyst, a repeating experiment of photo-degradation of RhB was carried out for five cycles. Fig. 9(b) shows the experimental degradation rate of the sample after five cycles. The pure $g\text{-C}_3\text{N}_4$ sample showed a great deal of instability, with the photo-degradation efficiency decreasing from 78% to 8%. Conversely, the $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (7 wt%) photocatalyst was quite stable, and the degradation rate only decreased to 75% after five recycling runs. The core-shell $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ composite photocatalyst demonstrated pretty good stability and reusability. The core-shell structure promoted chemical band interaction and stability of the composite photo-catalyst while enlarging the reaction contact area.

3.8. Photoluminescence (PL) spectra & photocurrent & EIS analysis

Photoluminescence emission spectra were obtained in order to further study the separation efficiency of electron-hole pairs. Higher intensities in the emission spectra of the sample indicate quicker recombination of photo-generated charges, thus leading to reduced photo-catalytic activity. The PL spectra of $g\text{-C}_3\text{N}_4$, BiVO_4 and $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ under excitation at 320 nm are shown in Fig. 10(a). It can be clearly seen that the pure BiVO_4 produced strong emission peaks around 520 nm.⁶¹ The relative intensity of BiVO_4 was much greater than that of $\text{BiVO}_4@g\text{-C}_3\text{N}_4$, and the PL intensity of $g\text{-C}_3\text{N}_4$ was also greater, indicating that coating with $g\text{-C}_3\text{N}_4$ nano-sheets played an important role in suppressing the recombination of electrons and holes. The core-shell structure resulted in an improvement in the separation efficiency of photo-generated charges. Fig. 10(b) shows a magnified view of the PL spectra of $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ (x wt%) ($x = 3, 5, 7, 10$). With increasing $g\text{-C}_3\text{N}_4$ content, the intensities of PL decrease until the content reaches 7%, after which the intensity increases. This may be due to an excess in $g\text{-C}_3\text{N}_4$ nano-sheets on the surface resulting in a gradual disappearance of the π -conjugated synergetic effect. These results are in agreement with the photo-catalytic activity results.

In order to provide more evidence of photoelectron and hole separation, a transient photocurrent response experiment was performed under visible light irradiation every 30 seconds. A higher photocurrent response value indicates lower recombination and more efficient separation of electron-hole pairs.⁶² As the results in Fig. 11 show, the semiconductor electrodes modified by the BiVO_4 , $g\text{-C}_3\text{N}_4$ and $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ samples show quick photocurrent response under the light irradiation, and decrease to zero immediately as soon as the light off. The $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ composite conducts the most photocurrent when compared with both pure BiVO_4 and $g\text{-C}_3\text{N}_4$. Namely, more electron-hole pairs can be separated efficiently over the $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ sample, and the longer lifetime of the photo-excited charge carrier than that of pure BiVO_4 , $g\text{-C}_3\text{N}_4$. In this case, the core-shell structure resulted in a larger contact area which contributed to a higher photocurrent response and



Fig. 10 (a) Photoluminescence spectra of as-prepared $g\text{-C}_3\text{N}_4$, BiVO_4 , $\text{BiVO}_4@g\text{-C}_3\text{N}_4$; (b) partial enlarged PL spectra of $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ with different amounts of $g\text{-C}_3\text{N}_4$ nano-sheets.



Fig. 11 Transient photocurrent response for $g\text{-C}_3\text{N}_4$, BiVO_4 and $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ photo-catalysts.

inhibited the recombination rate of photo-excited reactive electrons and holes.

Electrochemical impedance spectroscopy (EIS) of pure $g\text{-C}_3\text{N}_4$, BiVO_4 and $\text{BiVO}_4@g\text{-C}_3\text{N}_4$ was performed in order to investigate the charge transfer resistance and assess the separation efficiency of electron-hole pairs, which is a crucial factor for photocatalytic degradation. In the Nyquist diagram of EIS,



a smaller radius is evidence of smaller charge transfer resistance and a faster charge transfer rate.⁶³ Fig. 12 shows that the arc radius of BiVO₄@g-C₃N₄ (7 wt%) is the smallest among the three samples. The core-shell structure formed between BiVO₄ and the g-C₃N₄ nano-sheets changed the chemical bands and altered the resistance on the electrode surface, resulting in higher separation efficiency and a faster transfer rate of photo-generated charges. As the widely conjugated system developed, just as the FT-IR and XPS analysis, the interaction between g-C₃N₄ nano-sheets shell and BiVO₄ core can provide an effective path for the transfer of photo-induced charges. These data also support the results from the photocatalytic activity, PL, and photocurrent results.

3.9. The proposed mechanism of photocatalytic degradation

Based on the analysis and characterization of the experimental results above, a possible photocatalytic mechanism for core-shell BiVO₄@g-C₃N₄ is proposed and discussed below. The layered g-C₃N₄ nano-sheets coated on the ellipsoid-like BiVO₄ provide a larger contact area and interface between the core and shell, resulting in strengthening of the π -conjugated system and promotion of the separation of photo-generated electron-hole pairs. This enables a higher photo-degradation efficiency under light irradiation. Proper band gap position of the photocatalyst is one of the key factors to promote separation of electron-hole pairs and is a result of the composite's properties. According to the data from the UV-vis experiments, the band-gap energy of BiVO₄ is 2.46 eV (Fig. 5(b)). The band gap edge positions of the conduction band and valence band can be calculated by the formulas:

$$E_{CB} = X - E^e - 0.5E_g, E_{VB} = E_{CB} + E_g$$

where X represents the absolute electronegativity of the semiconductor and is obtained from the geometric mean of the constituent atoms, E^e is the energy of free electrons on the hydrogen scale (about 4.5 eV), and E_g is the band-gap energy of the semiconductor.⁶⁴ The E_{CB} and E_{VB} values were 0.31 eV and 2.77 eV for BiVO₄, respectively. For g-C₃N₄, the conduction band and valence band were found at -1.13 eV and 1.57 eV, which

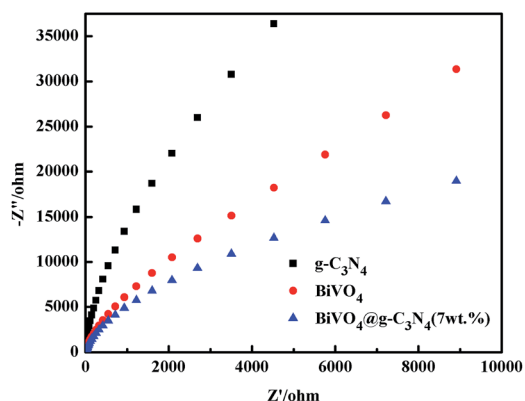


Fig. 12 Electrochemical impedance spectroscopy of g-C₃N₄, BiVO₄ and BiVO₄@g-C₃N₄.

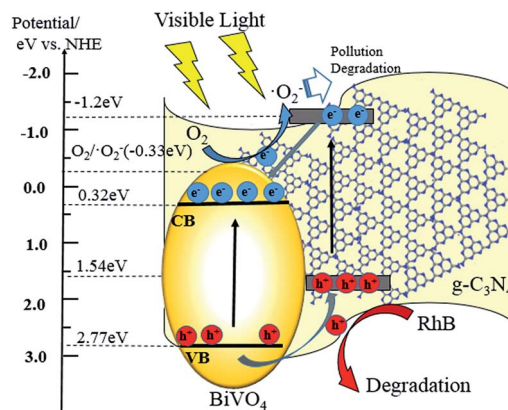


Fig. 13 The proposed mechanism for photo-degradation process.

agrees with previously reported literature values.⁴⁵ A proposed mechanism that considers the relative band gap edge positions is illustrated in Fig. 13. Under the irradiation of visible light, both BiVO₄ and g-C₃N₄ are excited and easily induce the generation of electron-hole pairs. Due to the well-matched band structure and large interfacial contact area of the core-shell structure, the photo-excited electrons of g-C₃N₄ are easily transferred to the CB of BiVO₄. Simultaneously, the holes in the VB of BiVO₄ easily migrate to the VB of g-C₃N₄ shell, where they are able to oxidize the absorbed RhB molecules into micro molecules. As the CB edge potential of g-C₃N₄ is more negative than the reduction potential of O₂/·O₂⁻ (-0.33 eV),⁶⁵ during the excited electrons migrated from conduction band of g-C₃N₄ to BiVO₄, only part of the electrons are able to reduce the absorbed oxygen into superoxide radicals (·O₂⁻), some of the electrons transported to the CB of BiVO₄. As was described in the trapping experiment, the h⁺ are the main reactive species and the ·O₂⁻ are an assistant species. The core-shell structure of the composite photo-catalyst facilitates efficient separation of photo-induced electron-hole pairs, and the layered shell provides more active sites due to an enlarged contact area. The wider π -conjugated system of the core-shell structure, the thin shell material reduces the distance for migration, allowing for more photo-excited charges to be separated and transferred.

4. Conclusions

Ellipsoid-like BiVO₄ and core-shell structured BiVO₄@g-C₃N₄ were successfully synthesized through a hydrothermal and ultra-sonicated-chemisorption method. Ellipsoid-like BiVO₄ was prepared through the addition of ethylene glycol. The performance of the composite photo-catalyst was accessed by evaluating its ability to degrade RhB pollution in water. The highest photo-degradation efficiency occurred when the content of g-C₃N₄ nano-sheet was 7% and was 9 times greater than that of pure BiVO₄. In addition, the core-shell BiVO₄@g-C₃N₄ showed high stability after five cycles of light irradiation. The stable core-shell structure of the composite photo-catalyst not only enlarged the contact area, but also strengthened the interaction of chemical bands, effectively enhancing the π -



conjugation. This resulted in efficient separation of the photo-excited electron-hole pairs and a high rate of photo-degradation of RhB (90% degradation in 60 min). This promising novel core-shell composite photo-catalyst provides an exciting area for further application and study.

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Notes and references

- 1 A. Fujishima and K. Honda, *Nature*, 1972, **238**, 37–38.
- 2 H. Tong, S. Ouyang, Y. Bi, N. Umezawa, M. Oshikiri and J. Ye, *Adv. Mater.*, 2012, **24**, 229–251.
- 3 C. McCullagh, N. Skillen, M. Adams and P. K. J. Robertson, *J. Chem. Technol. Biotechnol.*, 2011, **86**, 1002–1017.
- 4 D. Chatterjee and S. Dasgupta, *J. Photochem. Photobiol., C*, 2005, **6**, 186–205.
- 5 X. Chen, Y. Dai and X. Wang, *J. Alloys Compd.*, 2015, **649**, 910–932.
- 6 M. Pelaez, N. T. Nolan, S. C. Pillai, M. K. Seery, P. Falaras, A. G. Kontos, P. S. M. Dunlop, J. W. J. Hamilton, J. A. Byrne, K. O'Shea, M. H. Entezari and D. D. Dionysiou, *Appl. Catal., B*, 2012, **125**, 331–349.
- 7 M. Ni, M. K. H. Leung, D. Y. C. Leung and K. Sumathy, *Renewable Sustainable Energy Rev.*, 2007, **11**, 401–425.
- 8 Z. Yi, J. Ye, N. Kikugawa, T. Kako, S. Ouyang, H. Stuart-Williams, H. Yang, J. Cao, W. Luo, Z. Li, Y. Liu and R. L. Withers, *Nat. Mater.*, 2010, **9**, 559–564.
- 9 X. Zhang, X. B. Wang, L. W. Wang, W. K. Wang, L. L. Long, W. W. Li and H. Q. Yu, *ACS Appl. Mater. Interfaces*, 2014, **6**, 7766–7772.
- 10 J. Xia, S. Yin, H. Li, H. Xu, Y. Yan and Q. Zhang, *Langmuir*, 2011, **27**, 1200–1206.
- 11 J. Zhang, F. Shi, J. Lin, D. Chen, J. Gao, Z. Huang, X. Ding and C. Tang, *Chem. Mater.*, 2008, **20**, 2937–2941.
- 12 H. Huang, X. Han, X. Li, S. Wang, P. K. Chu and Y. Zhang, *ACS Appl. Mater. Interfaces*, 2015, **7**, 482–492.
- 13 H. Huang, K. Xiao, Y. He, T. Zhang, F. Dong, X. Du and Y. Zhang, *Appl. Catal., B*, 2016, **199**, 75–86.
- 14 S. Guo, X. Li, H. Wang, F. Dong and Z. Wu, *J. Colloid Interface Sci.*, 2012, **369**, 373–380.
- 15 D. Sánchez-Martínez, D. B. Hernández-Uresti, L. M. Torres-Martínez and S. Mejía-Rosales, *Res. Chem. Intermed.*, 2015, **41**, 8839–8854.
- 16 B.-X. Lei, L.-L. Zeng, P. Zhang, Z.-F. Sun, W. Sun and X.-X. Zhang, *Adv. Powder Technol.*, 2014, **25**, 946–951.
- 17 T. Saison, N. Chemin, C. Chanéac, O. Durupthy, L. Mariey, F. Maugé, V. Brezová and J.-P. Jolivet, *J. Phys. Chem. C*, 2015, **119**, 12967–12977.
- 18 O. F. Lopes, K. T. G. Carvalho, A. E. Nogueira, W. Avansi and C. Ribeiro, *Appl. Catal., B*, 2016, **188**, 87–97.
- 19 S. Obregón, A. Caballero and G. Colón, *Appl. Catal., B*, 2012, **117–118**, 59–66.
- 20 W. Wang, X. Huang, S. Wu, Y. Zhou, L. Wang, H. Shi, Y. Liang and B. Zou, *Appl. Catal., B*, 2013, **134–135**, 293–301.
- 21 C. Li, P. Zhang, R. Lv, J. Lu, T. Wang, S. Wang, H. Wang and J. Gong, *Small*, 2013, **9**, 3951–3956.
- 22 S. Dong, Y. Cui, Y. Wang, Y. Li, L. Hu, J. Sun and J. Sun, *Chem. Eng. J.*, 2014, **249**, 102–110.
- 23 M. L. Guan, D. K. Ma, S. W. Hu, Y. J. Chen and S. M. Huang, *Inorg. Chem.*, 2011, **50**, 800–805.
- 24 W. Zhao, Y. Liu, Z. Wei, S. Yang, H. He and C. Sun, *Appl. Catal., B*, 2016, **185**, 242–252.
- 25 X.-S. Zhang, J.-Y. Hu and H. Jiang, *Chem. Eng. J.*, 2014, **256**, 230–237.
- 26 T. Zhu, Y. Song, H. Ji, Y. Xu, Y. Song, J. Xia, S. Yin, Y. Li, H. Xu, Q. Zhang and H. Li, *Chem. Eng. J.*, 2015, **271**, 96–105.
- 27 X. Wang, K. Maeda, A. Thomas, K. Takanebe, G. Xin, J. M. Carlsson, K. Domen and M. Antonietti, *Nat. Mater.*, 2009, **8**, 76–80.
- 28 H. Yan and H. Yang, *J. Alloys Compd.*, 2011, **509**, L26–L29.
- 29 K. Sridharan, E. Jang and T. J. Park, *Appl. Catal., B*, 2013, **142–143**, 718–728.
- 30 L. Ge, F. Zuo, J. Liu, Q. Ma, C. Wang, D. Sun, L. Bartels and P. Feng, *J. Phys. Chem. C*, 2012, **116**, 13708–13714.
- 31 X. Dai, M. Xie, S. Meng, X. Fu and S. Chen, *Appl. Catal., B*, 2014, **158–159**, 382–390.
- 32 H. Xu, J. Yan, Y. Xu, Y. Song, H. Li, J. Xia, C. Huang and H. Wan, *Appl. Catal., B*, 2013, **129**, 182–193.
- 33 H. Shi, C. Zhang and C. Zhou, *RSC Adv.*, 2015, **5**, 50146–50154.
- 34 N. Tian, H. Huang, C. Liu, F. Dong, T. Zhang, X. Du, S. Yua and Y. Zhang, *J. Mater. Chem. A*, 2015, **3**, 17120–17129.
- 35 L. Ge, C. Han and J. Liu, *Appl. Catal., B*, 2011, **108–109**, 100–107.
- 36 L. Liu, Y. Qi, J. Lu, S. Lin, W. An, J. Hu, Y. Liang and W. Cui, *RSC Adv.*, 2015, **9**, 99339–99346.
- 37 C. Liu, H. Huang, X. Du, T. Zhang, N. Tian, Y. Guo and Y. Zhang, *J. Phys. Chem. C*, 2015, **119**, 17156–17165.
- 38 P. Niu, L. Zhang, G. Liu and H.-M. Cheng, *Adv. Funct. Mater.*, 2012, **22**, 4763–4770.
- 39 L. Liu, Y. Qi, J. Lu, S. Lin, W. An, Y. Liang and W. Cui, *Appl. Catal., B*, 2016, **183**, 133–141.
- 40 S. Kumar, A. Baruah, S. Tonda, B. Kumar, V. Shanker and B. Sreedhar, *Nanoscale*, 2014, **6**, 4830–4842.
- 41 N. Tian, H. Huang, Y. He, Y. Guo, T. Zhang and Y. Zhang, *Dalton Trans.*, 2015, **44**, 4297–4307.
- 42 H. J. Kong, D. H. Won, J. Kim and S. I. Woo, *Chem. Mater.*, 2016, **28**, 1318–1324.
- 43 Y. Ji, J. Cao, L. Jiang, Y. Zhang and Z. Yi, *J. Alloys Compd.*, 2014, **590**, 9–14.
- 44 C. Li, S. Wang, T. Wang, Y. Wei, P. Zhang and J. Gong, *Small*, 2014, **10**, 2783–2790.
- 45 M. Ou, Q. Zhong and S. Zhang, *J. Sol-Gel Sci. Technol.*, 2014, **72**, 443–454.



- 46 Z. Li, M. Li, Z. Bian, Y. Kathiraser and S. Kawi, *Appl. Catal., B*, 2016, **188**, 324–341.
- 47 S. W. Hu, L. W. Yang, Y. Tian, X. L. Wei, J. W. Ding, J. X. Zhong and P. K. Chu, *Appl. Catal., B*, 2015, **163**, 611–622.
- 48 M. Ge, L. Liu, W. Chen and Z. Zhou, *CrystEngComm*, 2012, **14**, 1038–1044.
- 49 L. Ge, C. Han, J. Liu and Y. Li, *Appl. Catal., A*, 2011, **409–410**, 215–222.
- 50 G. Liao, S. Chen, X. Quan, H. Yu and H. Zhao, *J. Mater. Chem.*, 2012, **22**, 2721–2726.
- 51 X. Zhang, X. Xie, H. Wang, J. Zhang, B. Pan and Y. Xie, *J. Am. Chem. Soc.*, 2013, **135**, 18–21.
- 52 H. Li, K. Yu, X. Lei, B. Guo, H. Fu and Z. Zhu, *J. Phys. Chem. C*, 2015, **119**, 22681–22689.
- 53 J. Liu, H. Wang, S. Wang and H. Yan, *Mater. Sci. Eng., B*, 2003, **104**, 36–39.
- 54 M. Yang and X. Jin, *J. Wuhan Univ. Technol., Mater. Sci. Ed.*, 2015, **30**, 217–222.
- 55 M. Gotić, S. Musić, M. Ivanda, M. Šoufek and S. Popović, *J. Mol. Struct.*, 2005, **744–747**, 535–540.
- 56 Y. Wang, R. Shi, J. Lin and Y. Zhu, *Energy Environ. Sci.*, 2011, **4**, 2922.
- 57 S. C. Yan, Z. S. Li and Z. G. Zou, *Langmuir*, 2010, **26**, 3894–3901.
- 58 S. Wang, D. Li, C. Sun, S. Yang, Y. Guan and H. He, *Appl. Catal., B*, 2014, **144**, 885–892.
- 59 T. Wu, G. Liu and J. Zhao, *J. Phys. Chem. B*, 1998, **102**, 5845–5851.
- 60 M. Yan, Y. Wu, Y. Yan, X. Yan, F. Zhu, Y. Hua and W. Shi, *ACS Sustainable Chem. Eng.*, 2016, **4**, 757–766.
- 61 D. K. Ma, M. L. Guan, S. S. Liu, Y. Q. Zhang, C. W. Zhang, Y. X. He and S. M. Huang, *Dalton Trans.*, 2012, **41**, 5581–5586.
- 62 Y. Xu, H. Xu, L. Wang, J. Yan, H. Li, Y. Song, L. Huang and G. Cai, *Dalton Trans.*, 2013, **42**, 7604–7613.
- 63 J. Kim, S. Byun, A. J. Smith, J. Yu and J. Huang, *J. Phys. Chem. Lett.*, 2013, **4**, 1227–1232.
- 64 J. Zhang, F. Ren, M. Deng and Y. Wang, *Phys. Chem. Chem. Phys.*, 2015, **17**, 10218–10226.
- 65 J. Kim, C. W. Lee and W. Choi, *Environ. Sci. Technol.*, 2010, **44**, 6849–6854.

