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## Hierarchical 3D porous carbon with facilely accessible Fe–N<sub>4</sub> single-atom sites for Zn–air batteries†

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Here, we report a highly efficient ORR electrocatalyst with Fe–N<sub>4</sub> active sites uniformly dispersed on a three-dimensional (3D) interconnected porous nitrogen-doped carbon network synthesized by pyrolyzing SiO<sub>2</sub>@ZIF-8 composites loaded with iron salts. The as-prepared single-atom Fe 3D-ordered mesoporous carbon (SA-Fe-3DOMC) possesses a high specific surface area of 1357.8 m<sup>2</sup> g<sup>-1</sup> and a high Fe loading of 0.84 wt% as well. Benefiting from these favourable structural properties, SA-Fe-3DOMC exhibits a superior ORR half-wave potential ( $E_{1/2}$ ) of 0.901 V and negligible activity loss (only 3 mV) after 10 000 cycles in alkaline media, surpassing the state-of-the-art Pt/C electrocatalyst. Particularly, an integrated zinc–air battery with SA-Fe-3DOMC as the air electrode shows a remarkable peak power density (140 mW cm<sup>-2</sup>) and a high specific capacity (786.6 mA h g<sup>-1</sup>), demonstrating great potential for practical application.

Zinc–air batteries (ZABs) have received significant attention as promising next-generation batteries because of their high power density, high specific energy, low cost, environmental friendliness and good safety.<sup>1–3</sup> However, the oxygen reduction reaction (ORR) at the cathode is subject to sluggish kinetics and usually requires high overpotentials.<sup>4</sup> In order to speed up the ORR and make ZABs more energy-efficient, suitable electrocatalysts are always needed.<sup>5–7</sup> Although precious metals and their alloys remain as outstanding candidates for catalyzing the ORR, their practical applications have been significantly hampered by the prohibitive cost and scarcity of precious

metals.<sup>8–12</sup> Therefore, it is important to develop high-performance inexpensive catalysts to replace precious metals.<sup>13,14</sup>

To date, Fe single-atom catalysts (SACs) dispersed as Fe–N<sub>4</sub> active sites on a porous carbon matrix have been regarded as one of the most efficient ORR electrocatalysts owing to their high activity, durable stability and maximum atom utilization.<sup>15–17</sup> In general, Fe SACs can be obtained by direct pyrolysis of mixtures of N- and C-containing precursors and Fe salts, followed by acid washing.<sup>18,19</sup> However, iron species are prone to agglomerate during the pyrolysis process, forming iron-based nanoparticles or clusters that are difficult to remove, thus blocking the formation of Fe–N<sub>4</sub> active sites.<sup>20</sup> Moreover, inaccessible pores will be generated owing to the random mixing of the N–C precursors and Fe salts. This leads to some buried Fe–N<sub>4</sub> active sites that cannot be exposed at the three-phase interface, thus reducing the utilization of Fe SACs.<sup>21–24</sup> Thus, an ideal Fe SAC should possess highly accessible Fe–N<sub>4</sub> active sites and a porous network that can facilitate the transport of electrons, oxygen and water.<sup>25</sup> A three dimensional (3D) ordered porous architecture with open channels would be suitable to maximize the accessibility of active sites, but suffers from the limitation of a low specific surface area and cannot support more Fe–N<sub>4</sub> active sites.<sup>26,27</sup> Considering these issues collectively, a 3D-ordered large-pore structure with suitable micropores and mesopores will be highly desirable, which facilitates the exposure of active sites and mass transfer.<sup>28–30</sup>

Herein, we have designed 3D-ordered porous nitrogen-doped carbon with accessible Fe–N<sub>4</sub> active sites as an efficient electrocatalyst for the ORR. Remarkably, the as-prepared SA-Fe-3DOMC possessed a high specific surface area of 1357.8 m<sup>2</sup> g<sup>-1</sup> and a high Fe loading (0.84 wt%). The electrochemical results demonstrated that SA-Fe-3DOMC exhibited excellent ORR catalytic activity and stability, with  $E_{1/2}$  increased to 0.901 V while the current almost remained similar after 10 000 potential cycles from 0.6 V to 1 V in alkaline media. Furthermore, an integrated ZAB with SA-Fe-3DOMC as the air electrode showed a remarkable peak power density (140 mW cm<sup>-2</sup>) and a high

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specific capacity ( $786.6 \text{ mA h g}^{-1}$ ), demonstrating great potential for practical application.

The synthetic procedure of the SA-Fe-3DOMC catalysts is illustrated in Scheme 1. Briefly, silica nanospheres with a size of 180–200 nm were first modified with  $\text{Fe}^{3+}$  ( $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ ) and  $\text{Zn}^{2+}$  ( $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ ) (Fig. S1a, ESI†). The subsequent *in situ* growth of ZIF-8 embedded the silica nanospheres into the MOF precursors (Fig. S1b, ESI†).<sup>26</sup> After carbonization at  $900^\circ\text{C}$  and HF etching, SA-Fe-3DOMC was obtained, showing a honeycomb-like structure and uniform pores with a diameter ranging from 180 to 200 nm (Fig. 1a, b, S1c and d, ESI†), matching well with the size of silica templates. Consistently, a well-defined interconnected macroporous structure with a 10 nm-thick wall was identified by transmission electron microscopy (TEM) (Fig. 1c). Moreover, Fig. 1d reveals that no nanoparticles were detected in the high-angle annular dark field-scanning transmission electron microscope (HAADF-STEM) observation. Elemental mapping images clearly demonstrated the homogeneous distribution of C, N, O and Fe in SA-Fe-3DOMC (Fig. 1e), indicating that iron atoms exist in the carbon matrix as atomically dispersed active sites. Moreover, the HAADF-STEM image of SA-Fe-3DOMC exhibited that abundant and isolated Fe atoms with brighter spots were indeed atomically dispersed on the porous carbon matrix (Fig. 1f), which was a piece of strong evidence for the successful synthesis of single atoms.<sup>27,28</sup>

The XRD patterns of N-C, SA-Fe-3DOMC and Fe-N-C-50 showed almost the same diffraction peaks centered at around  $24.7^\circ$  and  $43.6^\circ$  (Fig. 2a), which are assigned to the (002) and (101) lattice planes of graphitic carbon, respectively.<sup>31,32</sup> Remarkably, no peaks belonging to Fe and Zn species were observed, confirming the absence of crystalline Fe and Zn particles in SA-Fe-3DOMC. The Fe loading amount of SA-Fe-3DOMC was measured to be 0.84 wt% by inductively coupled plasma-mass spectrometry (ICP-MS). The Raman spectra of the three samples exhibited two prominent peaks at  $1340 \text{ cm}^{-1}$  and  $1590 \text{ cm}^{-1}$ , which are ascribed to the D-band ( $\text{sp}^3$  disordered carbon) and G-band ( $\text{sp}^2$  graphitic carbon), respectively



Scheme 1 Schematic illustration of the synthesis of SA-Fe-3DOMC.



Fig. 1 (a and b) SEM images of SA-Fe-3DOMC. (c) TEM image of SA-Fe-3DOMC. (d) HAADF-STEM image of SA-Fe-3DOMC and (e) corresponding elemental images. (f) AC-HAADF-STEM image of SA-Fe-3DOMC.

(Fig. 2b).<sup>33</sup> The intensity ratio  $I_D/I_G$  values for N-C, SA-Fe-3DOMC and Fe-N-C-50 were calculated to be 0.96, 0.93 and 0.95, indicating that SA-Fe-3DOMC possessed a much higher graphitization degree than the other two counterparts.

The pore structures of the as-synthesized catalysts were further characterized by  $\text{N}_2$  adsorption-desorption measurements. All catalysts presented a characteristic type IV isotherm, suggesting the presence of abundant mesopores in the samples. Furthermore, a rapid adsorption of  $\text{N}_2$  at relatively low pressure ( $P/P_0 < 0.1$ ) was observed, indicating that micropores can be speculated in the sample (Fig. 2c).<sup>34,35</sup> With the increased number of Fe salts, the specific surface area dramatically decreased from  $1681.6 \text{ m}^2 \text{ g}^{-1}$  for N-C to  $1357.8 \text{ m}^2 \text{ g}^{-1}$  for SA-Fe-3DOMC and to  $629.7 \text{ m}^2 \text{ g}^{-1}$  for Fe-N-C-50. Moreover, the



Fig. 2 (a) XRD spectra, (b) Raman spectra, (c)  $\text{N}_2$  adsorption/desorption isotherms and (d) pore size distributions of N-C, SA-Fe-3DOMC and Fe-N-C-50. (e) The N 1s high resolution spectrum of SA-Fe-3DOMC. (f) Fe 2p high-resolution spectra. (g) Fe K-edge XANES spectra and (h) Fourier transforms of the Fe K-edge EXAFS oscillations of Fe foil, FeO,  $\text{Fe}_2\text{O}_3$ , SA-Fe-3DOMC, and FePc. (i) FT-EXAFS curve fitting of SA-Fe-3DOMC in  $R$  space.

pore distribution revealed two sharp peaks at 0.565 nm and 2 to 4 nm (Fig. 2d), implying that a large number of micropores and mesopores were formed,<sup>36</sup> owing to the evaporation of Zn species and the catalytic effect of Fe salts, respectively.<sup>37,38</sup> Notably, the higher specific surface area and abundant pores facilitate the exposure of more Fe-N<sub>x</sub> active sites, resulting in better ORR catalytic performance.

The elemental chemical state of the obtained catalyst surface was analysed by X-ray photoelectron spectroscopy (XPS). The XPS survey spectrum of SA-Fe-3DOMC presented showed the presence of C (89.28 at%), N (4.63 at%), O (5.04 at%), and Fe (1.05 at%) (Fig. S2a, ESI<sup>†</sup>). The C 1s spectrum of SA-Fe-3DOMC was resolved into C-C (~284.7 eV), C-N (~286.2 eV) and C-O (~289.0 eV), suggesting that heteroatom N was successfully doped into the porous carbon network (Fig. S3, ESI<sup>†</sup>).<sup>39</sup> The high-resolution N1s spectra of N-C, SA-Fe-3DOMC and Fe-N-C-50 can be deconvoluted into pyridinic-N (~398.5 eV), pyrrolic-N (~400.1 eV), graphitic-N (~401.1 eV) and oxidized-N (~403.9 eV). In addition, a peak at 399.6 eV assigned to Fe-N<sub>x</sub> species was observed in both SA-Fe-3DOMC and Fe-N-C-50 (Fig. 2e, S2b and c, ESI<sup>†</sup>).<sup>28</sup> Furthermore, SA-Fe-3DOMC possessed the highest proportion of pyridinic N and graphitic N compared to the other two counterparts (Fig. S4, ESI<sup>†</sup>), which is beneficial for regulating the properties of the Fe-N-C catalyst.<sup>39</sup> The Fe 2p high-resolution spectra of SA-Fe-3DOMC and Fe-N-C-50 can be convoluted into Fe 2p<sub>3/2</sub> (710.0 eV) and Fe 2p<sub>1/2</sub> (723.7 eV), with peaks around 725.5 eV and 716.1 eV classified as the satellite peaks of Fe, while the peak at 713.2 eV could be assigned to Fe-N<sub>x</sub> species (Fig. 2f and S2d, ESI<sup>†</sup>).<sup>40,41</sup> Linear fitting curve of Fe foil, FeO, Fe<sub>2</sub>O<sub>3</sub> and SA-Fe-3DOMC were derived from corresponding Fe K-edge X-ray absorption near edge structure (XANES) position. The iron K-edge XANES position of SA-Fe-3DOMC was located close to that of FeO (Fig. 2g and S9, ESI<sup>†</sup>), indicating that the valence state of Fe in SA-Fe-3DOMC was close to +2.<sup>42,43</sup> The Fourier-transform (FT) of *k*<sup>3</sup>-weighted extended X-ray absorption fine structure (EXAFS) revealed a predominant peak at 1.50 Å in *R* space for SA-Fe-3DOMC (Fig. 2h), ascribed to the Fe-N(O) coordination. No Fe-Fe coordination peak (at 2.2 Å) was observed for SA-Fe-3DOMC, further demonstrating the absence of metallic Fe.<sup>44</sup> The EXAFS fitting results were used to indicate the coordination environments of Fe in SA-Fe-3DOMC (Fig. 2i and Table S1, ESI<sup>†</sup>), which revealed that the Fe atoms in SA-Fe-3DOMC were coordinated with four N atoms, indicating a high probability that the active sites in SA-Fe-3DOMC could be Fe-N<sub>4</sub>.<sup>45</sup>

The electrocatalytic activity of NC, SA-Fe-3DOMC and Fe-N-C-50 was first evaluated by cyclic voltammetry (CV). Apparently, the CV curves of all samples showed obvious oxygen reduction peaks under O<sub>2</sub>-saturated conditions (Fig. 3a, S5a and b, ESI<sup>†</sup>), and a more positive oxygen reduction peak of SA-Fe-3DOMC was observed, indicating a superior catalytic activity. Moreover, linear sweep voltammetry (LSV) curves recorded by using a rotating disk electrode (RDE) at 1600 rpm were used to investigate the ORR activities in O<sub>2</sub>-saturated 0.1 M KOH solution. The SA-Fe-3DOMC catalyst exhibited the highest ORR catalytic activity with an onset potential (*E*<sub>onset</sub>) of 0.994 V, a *E*<sub>1/2</sub> of 0.901 V and a diffusion-limited current densities of 5.696 mA



Fig. 3 (a) CV curves in O<sub>2</sub>- and N<sub>2</sub>-saturated 0.1 M KOH solution. (b) LSV of Pt/C, N-C, SA-Fe-3DOMC, and Fe-N-C-50 in O<sub>2</sub>-saturated 0.1 M KOH at 1600 rpm. (c) RDE curves of SA-Fe-3DOMC from 400 rpm to 2400 rpm (inset: K-L plots at 0.4 V, 0.5 V and 0.6 V). (d) Comparison of kinetic current density (*j<sub>k</sub>*) and *E*<sub>1/2</sub> for Pt/C, N-C, SA-Fe-3DOMC, and Fe-N-C-50. (e) Chronoamperometric curves of SA-Fe-3DOMC and Pt/C before and after the addition of 2% (v/v) methanol in O<sub>2</sub>-saturated 0.1 M KOH. (f) ORR polarization curves of SA-Fe-3DOMC before and after 10 000 cycles.

cm<sup>-2</sup>, which were much higher than those of commercial Pt/C (0.946 V, 0.855 V, and 5.228 mA cm<sup>-2</sup>), Fe-N-C-50 (0.955 V, 0.868 V, and 5.608 mA cm<sup>-2</sup>) and NC (0.876 V, 0.781 V, and 4.982 mA cm<sup>-2</sup>) (Fig. 3b). To further explore the ORR pathway of SA-Fe-3DOMC, LSV curves were recorded at different rotation speeds (from 400 to 2400 rpm). As depicted in Fig. 3c, the Koutecky-Levich (K-L) plots of SA-Fe-3DOMC showed good linearity and a similar slope, demonstrating that the ORR takes place on the SA-Fe-3DOMC catalyst in a similar way at different potentials. The average electron transfer number of SA-Fe-3DOMC calculated from the K-L plot at 0.4–0.6 V was about 3.7, reflecting an approximate 4e<sup>-</sup> oxygen reduction pathway (inset of Fig. 3c). RRDE measurements demonstrated very low H<sub>2</sub>O<sub>2</sub> yields (3.4–5.1%) and electron transfer numbers in the range of 3.7–3.8 for the SA-Fe-3DOMC catalyst in the potential range from 0.2 to 0.8 V (Fig. S6a and b, ESI<sup>†</sup>). Fig. 3d further shows the kinetic current density (*j<sub>k</sub>*) of N-C, SA-Fe-3DOMC, Fe-N-C-50 and commercial Pt/C. Remarkably, the *j<sub>k</sub>* of SA-Fe-3DOMC reached as high as 18.70 mA cm<sup>-2</sup> at 0.85 V, and it was almost 2.8-fold that of commercial Pt/C (6.63 mA cm<sup>-2</sup>).

In addition, the methanol crossover effects on SA-Fe-3DOMC and Pt/C catalysts were measured in O<sub>2</sub>-saturated 0.1 M KOH at 0.6 V. With the addition of 2% (v/v) methanol, a sharp decay in current density was immediately detected for Pt/C (Fig. 3e). In contrast, there was no noticeable variation in current density for SA-Fe-3DOMC, showing a better tolerance to methanol

crossover. The ORR stability of SA-Fe-3DOMC and commercial Pt/C was further investigated by an accelerated durability test (ADT). As illustrated in Fig. 3f, the  $E_{1/2}$  of SA-Fe-3DOMC displayed a negligible negative shift of only 3 mV after 10 000 potential cycles from 0.6 V to 1 V. In comparison, the commercial Pt/C showed a severe loss of 12 mV (Fig. S7, ESI†). As shown in Fig. S8,† the morphology and corresponding elemental distribution of SA-Fe-3DOMC before and after 10 000 cycles remain almost unchanged, further confirming the robust properties of SA-Fe-3DOMC. These results suggested that the SA-Fe-3DOMC catalyst has superior electrocatalytic activity and durability in alkaline medium, making it very attractive as a cost-effective electrocatalyst alternative to replace Pt-based counterparts for the ORR.

To further reveal the potential of the SA-Fe-3DOMC electrocatalyst for practical applications, a liquid ZAB was assembled, where carbon paper loaded with SA-Fe-3DOMC and a gas diffusion layer was used as the air electrode (Fig. 4a). Such a battery afforded a stable open-circuit voltage (1.48 V), which was more positive than that of the Pt/C-based battery (1.37 V) (Fig. 4b), confirming outstanding catalytic performance. As shown in Fig. 4c, the battery with SA-Fe-3DOMC exhibited a peak power density of  $140 \text{ mW cm}^{-2}$ , which was superior to that with Pt/C ( $124.5 \text{ mW cm}^{-2}$ ). A typical constant current discharge test also was performed at different current densities (Fig. 4d). Apparently, the voltage gap between SA-Fe-3DOMC and commercial Pt/C widened with the increase of current density, suggesting that the well-designed interconnected 3D nanoporous structure facilitated the mass transfer, especially at higher current densities.<sup>46</sup> When normalized to the consumed Zn during a long-term constant current discharge process at  $10 \text{ mA cm}^{-2}$ , the SA-Fe-3DOMC battery exhibited a much higher specific capacity with a value of  $786.6 \text{ mA h g}^{-1}$ , which means that the utilization of catalysts is equal to 95.9% since the theoretical capacity is  $820 \text{ mA h g}^{-1}$  (Fig. 4e). Moreover, the SA-Fe-3DOMC-based battery can provide electrical energy for 8 light-emitting diodes (LEDs) (Fig. 4f), demonstrating great potential for practical applications.



Fig. 4 (a) Schematic of ZABs using SA-Fe-3DOMC as the cathode. (b) Open-circuit plots of Pt/C and SA-Fe-3DOMC. (c) Discharge polarization curves and corresponding power density curves of ZABs respectively with Pt/C and SA-Fe-3DOMC as the cathodes. (d) Discharge curves of ZABs at different current densities. (e) Specific capacities of ZABs using SA-Fe-3DOMC and Pt/C as the cathode at  $10 \text{ mA cm}^{-2}$ . (f) A demonstration of an LED screen, powered by two ZAB sets.

In summary, we designed and synthesized a SA-Fe-3DOMC electrocatalyst with high ORR catalytic activity, using a  $\text{SiO}_2$ -embedded Zn-MOF assisted strategy. The achieved SA-Fe-3DOMC showed an excellent  $E_{1/2}$  of 0.901 V and a high durability of only 3 mV decrease after 10 000 potential cycles, surpassing the state-of-the-art Pt/C in alkaline medium. For ZAB applications, the as-prepared SA-Fe-3DOMC electrocatalyst also exhibited an outstanding peak power density of  $140 \text{ mW cm}^{-2}$  and a superior specific capacity of  $786.6 \text{ mA h g}^{-1}$ . Collectively, the hierarchical pores generated by  $\text{SiO}_2$  and Zn species as porogens allowed the facile access of the single-atom active sites, providing effective strategies not only for seeking alternatives to Pt/C catalysts but also to develop more catalysts for various energy conversion devices, such as  $\text{CO}_2$  reduction, water splitting, and  $\text{N}_2$  fixation.

## Conflicts of interest

There are no conflicts to declare.

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