



Cite this: *RSC Adv.*, 2019, 9, 17273

Mesoporous semiconductors combined with up-conversion nanoparticles for enhanced photodynamic therapy under near infrared light†

Fan Yang,^a Jun Liu,^a Xue Jiang,^a Weiwei Wu,^b Zhenni Wang,^b Qi Zeng^a and Ruichan Lv^{id}*^a

Photodynamic therapy (PDT) is a promising and effective method for tumor therapy that relies on the reactive oxygen species (ROS) produced by photosensitizers at specific wavelengths to inhibit tumor cells. Inorganic semiconductive materials are potential photosensitizers that can excellently absorb ultraviolet light to produce ROS to kill cancer cells. However, this strategy is still limited in terms of practical applications due to the weak penetration of ultraviolet light through biological tissue, as well as the hypoxic tumor microenvironment, largely decreasing ROS generation. In this research, novel PDT agents made with mesoporous lanthanide-semiconductor composites are developed to obtain a remarkable amount of generated ROS under near-infrared (NIR) laser irradiation. Due to the larger size (about 120 nm) of the up-conversion material (UCM) used as the substrate, coated with different amounts of semiconductors with mesoporous morphologies, this platform could emit higher blue emission under a 980 nm laser. Meanwhile, both of the semiconductors (SnO₂ and TiO₂) used have wide absorbance bands in the ultraviolet region, and the ultraviolet fluorescence emitted from the UCM core under NIR laser excitation can be used as the energy donor. Electron transfer processes in SnO₂ and TiO₂ are generated *via* the above platforms and produce ROS through photochemical action. Furthermore, the coated semiconductors are mesoporous with larger surface areas (about 302 m² g⁻¹) and various channels; this is beneficial to obtain enough oxygen to generate more ROS under a hypoxic environment. The PDT efficiency of a typical NaYF₄@SnO₂ sample is studied using a DPBF detector, *in vitro* MTT assays, and *in vivo* tumor inhibition experiments, revealing that this lanthanide-semiconductor platform could be potentially used as a PDT agent under NIR excitation.

Received 26th April 2019
 Accepted 20th May 2019

DOI: 10.1039/c9ra03116b

rsc.li/rsc-advances

Introduction

In recent years, compared with traditional radiotherapy and chemotherapy, photodynamic therapy (PDT) has attracted many researchers in the field of cancer treatment because of its few side effects, obvious efficacy, strong targeting, relatively low cost, and non-invasive mode of treatment.^{1–3} The main principle of PDT is to activate photosensitizers under radiation of a suitable wavelength to induce photodynamic effects leading to the production of reactive oxygen species

(ROS), which cause necrosis and apoptosis in diseased cells.⁴ Compared with traditional organic photosensitizers (*e.g.*, Ce6 and ZnPc), inorganic semiconductors (*e.g.*, TiO₂, SnO₂, MnO₂, *etc.*) with better biocompatibility can be stably maintained in the body longer with no obvious toxicity.^{5–8} Inorganic semiconductors can be utilized as photocatalysts and photosensitizer agents simultaneously.^{9,10} When a semiconductor is excited by energy from photons, electrons from the valence band can be excited to the conduction band, creating electron–hole pairs. Then, if they do not recombine, the electrons and holes can reduce or oxidize substances in the electrolyte solution. For example, in aqueous solution, holes oxidize water molecules to generate hydroxyl radicals ($\cdot\text{OH}$), and electrons reduce oxygen (O₂) to produce superoxide anions (O₂⁻) or hydrogen peroxide (H₂O₂). O₂⁻ is a single electron reduction product of oxygen, H₂O₂ is a two-electron reduction product of oxygen, and $\cdot\text{OH}$ is a three-electron reduction product of oxygen; they are all ROS.^{10,11} However, it is still difficult to extend current photodynamic therapy options to the NIR excitation range due to the limitations of existing photosensitizers; also, biological

^aEngineering Research Center of Molecular and Neuro Imaging, Ministry of Education, School of Life Science and Technology, Xidian University, Xi'an, Shaanxi 710071, China. E-mail: rchl@xidian.edu.cn

^bSchool of Advanced Materials and Nanotechnology, Xidian University, Xi'an, Shaanxi 710071, China

† Electronic supplementary information (ESI) available: SEM images and TEM images of the NaYF₄:Yb,Tm precursor; the particle size distribution of NaYF₄:Yb,Tm; the average shell thicknesses of three NYF@Sn samples; the average shell thicknesses of four NYF@Ti samples; N₂ adsorption/desorption isotherms and the pore size distribution of the NYF@Ti₂ composite; and a schematic diagram of NIR-induced ROS. See DOI: 10.1039/c9ra03116b



tissue absorbs visible ultraviolet light more strongly, which causes the excitation light to penetrate the tissue to a depth of only 1–2 mm. These factors easily lead to the failure of PDT toward cancer cells.^{12,13}

Studies of the optical window of biological tissue have shown that incident light in the near-infrared (NIR) range (700–1000 nm) has the greatest penetration into the tissue and it can penetrate the tissue by several centimeters.^{14–20} It would be very promising to transfer the excitation light wavelength during the PDT process into the NIR region. To overcome the drawbacks of the weak tissue penetration of ultraviolet light in the PDT field, combining the photosensitizer with an upconversion material has been proposed.^{21–27} Lanthanide-doped upconversion materials (UCMs) can absorb long-wavelength incident light and convert it into short-wavelength emission, such as converting near-infrared light into ultraviolet-visible light; for this purpose, the surface modification of upconversion materials has been widely used in biomedicine.^{28–36} Moreover, if a semiconductor is combined with an UCM, the excitation laser wavelength can be adjusted to the near-infrared region, and the obtained upconversion visible emission can be used for real-time image-guided diagnosis.^{37–41}

Most recently, there have been reports on the application of TiO₂ combined upconversion nanocomposites to PDT.^{42,43} Also, in our previous work, we have proved that ROS could be generated due to energy transfer from the upconversion blue emission of a lanthanide-based compound to a TiO₂-based photocatalyst that was responsive to blue emission.⁴⁴ However, up to now, there are several problems to be resolved relating to the structures of the energy donor UCM and energy acceptor photocatalyst: (1) commonly used upconversion substrates are too small (less than 50 nm), resulting in unstable combination and small amounts of attached TiO₂; (2) the structures of the photocatalysts used are always simple, and the surface areas of the bulk morphologies are not high enough to generate a sufficient amount of ROS; and (3) when inorganic PDT agents are used in hypoxic tumor cells, the PDT efficiency is further limited because of a lack of oxygen.^{45–50} Due to these three main reasons, PDT efficiency is not satisfactory under simple excitation pump power.

In this work, we used a slightly larger up-conversion material (~120 nm diameter) as the substrate, coated with different amounts of semiconductors with mesoporous morphologies. Due to the larger surface area with various channels, NaYF₄@SnO₂ and NaYF₄@TiO₂ have good photodynamic efficiency. Both of the semiconductors (SnO₂ and TiO₂) have wide absorbance bands in the ultraviolet region, and the ultraviolet fluorescence emitted from the NaYF₄ core under NIR laser excitation can be used as the energy donor. Electron transfer processes in SnO₂ and TiO₂ are generated in the above platforms and produce ROS, undergoing photochemical action. The PDT efficiency of a typical NaYF₄@SnO₂ sample is studied using DPBF, *in vitro* MTT assays, and *in vivo* tumor inhibition experiments.

Experimental section

Materials

All chemical reagents were used as received without further purification. Yttrium oxide (Y₂O₃, 99.99%), ytterbium oxide

(Yb₂O₃, 99.9%), thulium oxide (Tm₂O₃, 99.99%), ethylenediaminetetraacetic acid disodium salt dihydrate (EDTA), sodium fluoride (NaF), poly(acrylic acid) (PAA), sodium tetrafluoroborate (K₂SnO₃), urea (CON₂H₄), diethylenetriamine (DETA), and titanium isopropoxide, were purchased from the Shanghai Aladdin Bio-Chem Technology Co., Ltd., Shanghai, China. Isopropanol was obtained from the China National Pharmaceutical Group Corporation.

Synthesis of NaYF₄:Yb³⁺,Tm³⁺ upconversion nanoparticles (denoted as NYF)

The synthesis of the precursor was carried out according to a previous work.⁵¹ Typically, 0.8 mmol of EDTA was dissolved in 15 mL of deionized water, and 4 mL of 0.2 M RE(NO₃)₃ mixture solution (Y : Yb : Tm = 70 : 30 : 0.5) was added to the above solution under stirring at room temperature for 1 h. Then 12 mL of aqueous NaF (0.8 M) was added to the above solution and it was stirred for another hour. Then the mixed solution was transferred into a 50 mL Teflon-lined autoclave and this was heated at 180 °C for 2 h. The autoclave was cooled to room temperature naturally, and the obtained precipitates were collected *via* centrifugation, washed with deionized water and ethanol three times, and then dried in air at 60 °C for 12 h.

Synthesis of NaYF₄:Yb³⁺,Tm³⁺@SnO₂ core/shell nanoparticles (denoted as NYF@Sn)

This process of coating with SnO₂ of different thicknesses was carried out with slight changes to the published literature.⁵² Typically, 40 mg of NYF UCNPs and 0.08 g of PAA were put into 25 mL of ultrapure water and, after ultrasonication for 20 min, the mixed solution was stirred at 50 °C for 6 hours. Then, the above solution was centrifuged and the obtained precipitate was put into ultrapure water (UPW)/ethanol solution ($V_{UPW} : V_{ethanol} = 15.625 : 9.625$ mL). A certain amount of urea (40 mg) and K₂SnO₃ (10 mg) was added and the mixture was stirred for 10 min. Finally, the mixture was transferred into a 50 mL Teflon-lined autoclave and kept at 170 °C for 12 h. The obtained NYF@Sn was centrifuged and washed with deionized water and ethanol three times before drying at 60 °C for 720 min. The thickness of SnO₂ could be adjusted by adding different amounts of K₂SnO₃ – 5 mg, 10 mg, and 20 mg – and the corresponding core/shell nanoparticles are denoted as NYF@Sn1, NYF@Sn2, and NYF@Sn3, respectively.

Synthesis of NaYF₄:Yb³⁺,Tm³⁺@TiO₂ core/shell nanoparticles (denoted as NYF@Ti)

The process of coating with TiO₂ with different morphologies was carried out with slight changes to the published literature.⁵³ Typically, 25 mg of NYF UCNPs was dissolved in 21 mL of isopropanol *via* sonication, followed by the addition of 0.02 mL of DETA. After the solution was stirred for 10 min, 1 mL of titanium isopropoxide was added and stirring was continued for 10 min. Then, the above solution was transferred into a 50 mL Teflon-lined autoclave and maintained at 200 °C for 24 h. After the autoclave cooled to room temperature naturally, the products were separated *via* centrifugation, washed with ethanol



three times, and then dried at 60 °C for 12 h. Finally, the obtained NYF@Ti was calcined in a muffle furnace at 350 °C for 2 h at a heating rate of 1 °C min⁻¹ to make highly crystalline phase anatase TiO₂. The thickness of TiO₂ could be adjusted by adding different amounts of titanium isopropoxide – 0.1 mL, 0.2 mL, and 0.6 mL – and the corresponding core/shell nanoparticles are denoted as NYF@Ti1, NYF@Ti2, and NYF@Ti3, respectively.

Singlet oxygen detection *via* DPBF

A DPBF probe was used to detect the production of reactive oxygen species. In a dark environment, 2 mL of NYF@Sn/NYF@Ti aqueous solution (0.32 mg mL⁻¹) was prepared in a 10 mm cuvette, and then 36 μL of DPBF ethanol solution (1 M) was added. Then, absorbance spectra of the solution were sequentially detected using a UV spectrophotometer after irradiation with a 980 nm NIR laser for different time periods (0 min, 2 min, 4 min, 6 min, and 8 min).

In vitro cytotoxicity of NYF@Sn and NYF@Ti

MTT assays were carried out to detect the cytotoxicity of NYF@Sn and NYF@Ti towards cells. SGC 7901 cells were incubated overnight in a 96-well plate at 5000 per well to obtain a monolayer, then different concentrations of samples (0, 31.25, 62.5, 125, 250 and 500 μg mL⁻¹) were added before incubation at 37 °C for 24 hours under a 5% CO₂ atmosphere. Thereafter, each well was washed several times with PBS, and then 200 μL of medium and 20 μL of MTT solution (5 mg mL⁻¹) were added. After the 96-well plate was further incubated for 4 hours at 37 °C, the solution was discarded and 100 μL of DMSO was added. Finally, the plate was shaken for 20 min, and then the absorbance of the cells was recorded at a wavelength of 490 nm using a microplate reader.

In vitro anticancer efficiency using MTT assays

After SGC 7901 cells were cultured in 96-well plates for 12 hours to form a monolayer, NYF@Sn2 and NYF@Ti2 at a certain concentration (500 μg mL⁻¹) were added and incubated with the cells for 4 hours. Then each well was washed three times using PBS, discarding the mixture, followed by the addition of 100 μL of fresh medium to every well and irradiation with a 980 nm laser (1 W cm⁻²) for different time periods (1 min, 2 min and 5 min). Simultaneously, untreated cells were used as controls for cells treated with laser irradiation. Afterwards, the therapeutic effects of the semiconductor photodynamic materials were confirmed *via* detecting the cell viability using MTT assays. The treated and untreated cells were stained with Calcein AM and PI for 15 minutes at 37 °C and 5% CO₂, and background fluorescence interference was removed *via* washing several times using PBS; then fluorescence images were obtained *via* a Leica DMI 4000 B instrument.

In vivo anticancer efficiency

All experiments were conducted in accordance with internationally accepted laboratory animal use and care guidelines and

the protocol was approved by the School of Pharmacy Ethics Committee. Mice were acclimatized for one week before the study, and the mice were provided with water and food pellets until the end of the experimental period. The animal study procedures were approved and followed by the Xi'an Jiaotong University Animal Care and Use Committee. 4T1 cells dissolved in the medium were injected subcutaneously into the right hind limbs of BALB/c mice (4–5 weeks; 18–22 g). When tumors grew to ~30 mm³, a mouse tumor model was established and the *in vivo* experiments could be started. The mice were divided into two groups, the control group and the treatment group, with three mice in each group. No measures were taken for the control group. For the treatment group, NYF@Sn2 dissolved in PBS was injected into the mice tumors, and the tumors were irradiated with a 980 nm NIR laser for ten minutes. Laser treatment was conducted every two days and for a total of four times. Body weight and tumor size data from the two groups of mice were recorded every two days during the eight-day treatment cycle.

Results and discussion

Structure and morphology

The synthesis of the cubic-phase NaYF₄:Yb³⁺,Tm³⁺ nanoparticles is facile and precipitation is uniform. During the synthesis process, the solubility of NYF is not very good and it will precipitate in the aqueous solution for a long time. Therefore, PAA is used to modify the surface functionality of NYF, leading to better NYF hydrophilicity, which is advantageous for coating the semiconductors on the surface of the precursor. SEM images and TEM images are shown in Fig. S1;† the synthesized NYF is relatively uniform in size (about 120 nm) and has good dispersibility (Fig. S2)†.

Two facile hydrothermal methods are used for the synthesis of NYF coated with SnO₂ and TiO₂ shells. Both the formed NYF@Sn and NYF@Ti preserve their original spherical structure; porous SnO₂ and TiO₂ shells are coated on NYF with adjustable morphologies. For NYF@Sn, we can control the thickness of the SnO₂ shell by changing the amount of K₂SnO₃ added during the reaction. In Fig. 1A–C, NYF@Sn samples with SnO₂ shells of increasing thickness (from 5 nm to 15 nm) are presented, corresponding to the addition of 5 mg, 10 mg, and 20 mg of K₂SnO₃, respectively (Fig. S3)†. For NYF@Ti (Fig. 2), we synthesized three samples of NYF@Ti with different thicknesses (from 8 nm to 103 nm) by changing the amount of isopropyl titanate added (0.1 mL, 0.2 mL, 0.6 mL, and 1 mL) (Fig. S4)†. NYF@Ti1 with a thinner TiO₂ shell has a shell thickness of ~8 nm, and the coating is uniform. When coated with a thicker TiO₂ shell, NYF@Ti3 remains uniform with a thickness of ~103 nm. Further typical morphology analysis is presented in Fig. 3. As shown, NYF@Sn2 with a thickness of 10 nm is mono-dispersed (Fig. 3A), and the porous SnO₂ is clearly visible (Fig. 3B). Element mapping imaging and energy dispersive spectroscopy (EDS) data demonstrate the presence of Na, Y, Yb, Tm, F, O and Sn elements, as shown in Fig. 3C, confirming that SnO₂ is successfully coated on the NYF core and the elements are dispersed uniformly. The surfaces of the core-





Fig. 1 SEM images and TEM images of NYF@Sn synthesized from different amounts of K_2SnO_3 : (A1–A3) 5 mg; (B1–B3) 10 mg; and (C1–C3) 20 mg. Here, the amount of added NYF is 25 mg.

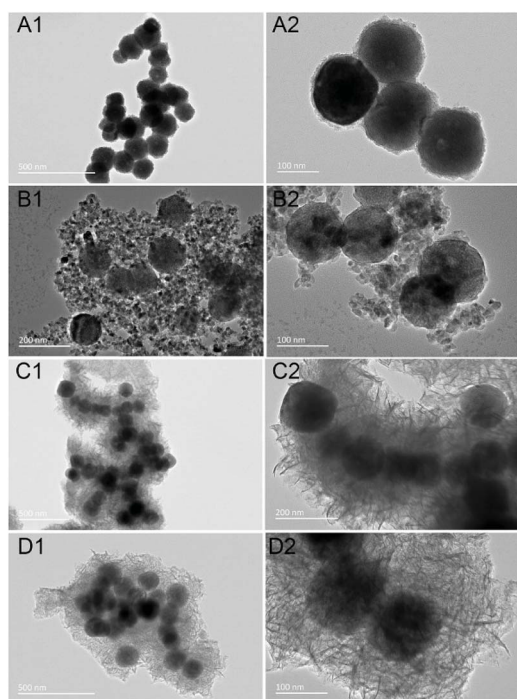


Fig. 2 TEM images of NYF@Ti synthesized with different amounts of isopropyl titanate: (A1 and A2) 100 μ L; (B1 and B2) 200 μ L; (C1 and C2) 600 μ L; and (D1 and D2) 1000 μ L. Here, the amount of added NYF is 25 mg.

shell nanoparticles become rough due to the porous shell, which increases the specific surface area and is more favorable for absorbing incident light for upconversion; the shell can also achieve better photocatalysis.

Fig. 4A shows XRD patterns of the as-synthesized precursor, and NYF@Sn2 and NYF@Ti2 samples. As shown, the peaks of the precursor correspond well with JCPDS 06-0342. After coating with TiO_2 or SnO_2 separately, the typical peaks of TiO_2 or SnO_2 occur correspondingly, further indicating the successful



Fig. 3 (A) An SEM image, (B) a TEM image, (C) element mapping data, and (D) EDS results from NYF@Sn2.

combination of the semiconductors. The N_2 adsorption/desorption isotherm of a typical NYF@Sn2 composite was further detected and is shown in Fig. 4B. It is clear that the as-synthesized NYF@Sn2 shows a H1 hysteresis loop, indicating typical mesopores and channels. The BET surface area, pore volume, and average pore width are calculated to be $84 \text{ m}^2 \text{ g}^{-1}$, $0.23 \text{ cm}^3 \text{ g}^{-1}$, and 17 nm, respectively (Fig. 4C), while the detected BET surface area, pore volume, and average pore width of NYF@Ti2 are $185 \text{ m}^2 \text{ g}^{-1}$, $0.24 \text{ cm}^3 \text{ g}^{-1}$ and 19 nm, respectively (Fig. S5[†]). These large BET surface areas and pore sizes are beneficial for absorbing oxygen in water and tissue, which is helpful for improving the PDT efficiency.

Optical properties and ROS detection

Upconversion luminescence spectra of NYF, NYF@Sn2 and NYF@Ti2 under laser irradiation at 980 nm are shown in Fig. 5A and C. The NYF UCNPs show typical Tm^{3+} photoluminescence emission characteristics. In the ultraviolet region, two emission peaks appear at 347 nm and 362 nm, induced by the radioactive transitions of $^1I_6 \rightarrow ^3F_4$ and $^1D_2 \rightarrow ^3H_6$. Two blue emission peaks are located at 452 nm and 475 nm, derived from the electronic transitions of $^1D_2 \rightarrow ^3F_4$ and $^1G_4 \rightarrow ^3H_6$, respectively.⁵⁴ A red emission peak at 646 nm is derived from the $^1G_4 \rightarrow ^3F_4$ transitions of Tm^{3+} ions. After coating with the SnO_2 shell (or TiO_2 shell), almost all of the peaks are weakened, while the peak at 646 nm is unchanged. This phenomenon has two main causes: firstly, the SnO_2 shell (or TiO_2 shell) attenuated a portion of the NIR light, resulting in a decrease in the intensity of the excitation light reaching the NYF core; secondly, the Tm^{3+} ion transitions became weaker because SnO_2 (or TiO_2) absorbed





Fig. 4 (A) XRD patterns of as-synthesized samples. (B) The N_2 adsorption/desorption isotherm and (C) pore size distribution of the NYF@Sn2 composite.

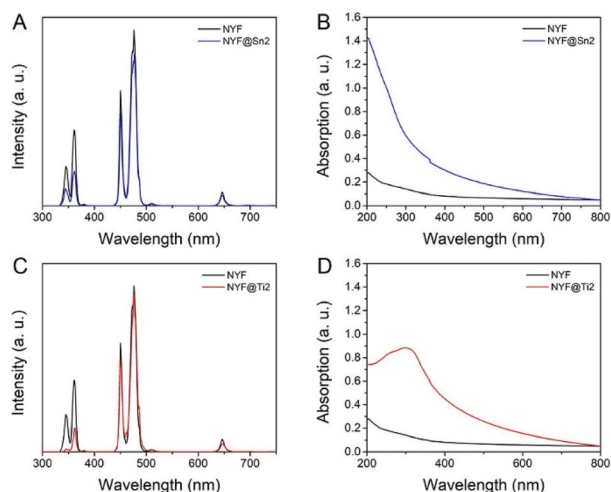


Fig. 5 Upconversion luminescence spectra of the NYF precursor and (A) NYF@Sn2 or (C) NYF@Ti2 under 980 nm excitation. UV-vis absorbance spectra of NYF solution and (B) NYF@Sn2 solution or (D) NYF@Ti2 solution.

a part of the ultraviolet light emitted by Tm^{3+} . Therefore, the two peaks in the ultraviolet region decreased more. UV-vis absorbance spectra of NYF, NYF@Sn2 and NYF@Ti2 are measured

and compared in Fig. 5B and D. NYF has almost no absorbance in the range of 200 nm to 800 nm. However, NYF@Sn2 has a strong and sharp absorbance edge from 400 nm to 200 nm, which corresponds to the absorbance characteristics of SnO_2 , and this edge covers the $^1I_6 \rightarrow ^3F_4$ and $^1D_2 \rightarrow ^3H_6$ emission peaks of NYF, confirming the theory that SnO_2 absorbed the emitted light of NYF. Similarly, NYF@Ti2 has a broad absorbance peak in the ultraviolet region, which also overlaps with the emission peaks of Tm^{3+} ions at 374 nm and 362 nm and corresponds to the intrinsic absorbance of TiO_2 .

The chemical probe 1,3-diphenylisobenzofuran (DPBF) is used to detect the active oxygen produced by the photodynamic action of NYF@Sn2 and NYF@Ti2 under excitation from a 980 nm NIR laser. ROS can react irreversibly with DPBF, resulting in a decrease in the characteristic absorbance peak of DPBF at 420 nm. Under NIR irradiation, the UV/blue light emitted from the NYF core is absorbed by the SnO_2 (or TiO_2) shell. Then, electrons in the semiconductor valence band are excited to the conduction band and holes are generated in the valence band. A schematic diagram of this NIR-induced PDT process can be seen in Scheme S1†. These electron-hole pairs are transferred to the surface of SnO_2 and react with substances in the solvent, and the ROS generated in this process reduce the characteristic peak of DPBF. It can be seen in Fig. 6 that the characteristic absorbance peak at 420 nm of DPBF solution mixed with NYF@Sn or NYF@Ti decreases significantly after different periods of NIR laser irradiation. As the NIR irradiation time increases, the characteristic peak drops more. After 8 min of NIR irradiation, NYF@Sn2 produces the most ROS, reducing the absorbance peak of DPBF by 86% (Fig. 6C), while the singlet oxygen generated in the greatest amount by NYF@Ti2 decreases the characteristic peak of DPBF by 49% (Fig. 6E).

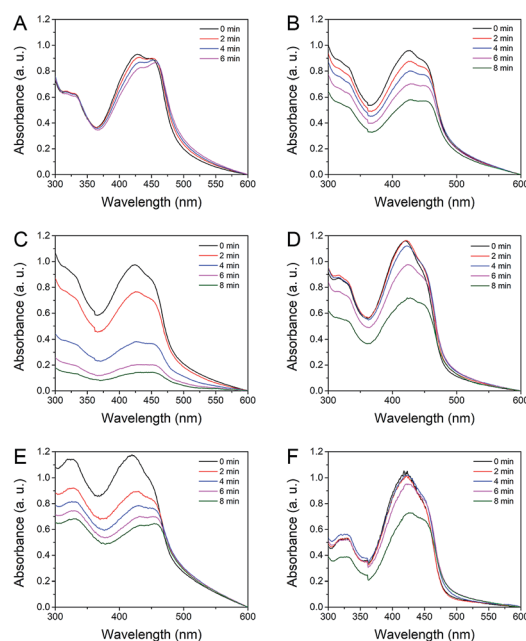


Fig. 6 Absorbance spectra of DPBF solution treated with (A) the NYF precursor, (B) NYF@Sn1, (C) NYF@Sn2, (D) NYF@Sn3, (E) NYF@Ti2, and (F) NYF@Ti3 at different irradiation time points.



In vitro and *in vivo* photodynamic therapy effects

Based on the morphology and PDT efficiency, we chose NYF@Sn2 and NYF@Ti2 as the final PDT agents. Before the prepared PDT agents were subjected to biological applications, their biocompatibilities were firstly assessed. The cytotoxicity of NYF@Sn2 and NYF@Ti2 towards cells was determined using the MTT method. Fig. 7A illustrates cell viability data after incubation with NYF@Sn and NYF@Ti (different concentrations) at 37 °C. It can be seen that the cell survival rate is maintained at 85% overall after incubation with NYF@Sn, and the cell survival rate after incubation with NYF@Ti is 90% on average. This shows that both samples have good biocompatibility and can be applied to photodynamic therapy.

In vitro anticancer efficiency assays under 980 nm NIR irradiation were further carried out. Compared with the control group, with an increase in the 980 nm laser irradiation time (1, 2 and 5 min), the cell viability is significantly reduced as a result of the photodynamic effects of the samples. When cells are incubated with NYF@Sn2 under 980 nm laser irradiation, the cell viability decreases to 80% after 6 min, while under the influence of NYF@Ti2, the cell viability decreases more before eventually decreasing to 70% (Fig. 7B). This indicates that the ROS produced by the photodynamic effects of NYF@Sn2 and NYF@Ti2 under 980 nm laser irradiation inhibit cell growth. Furthermore, microscopy images of cells incubated with NYF@Sn2 without and with 980 nm laser irradiation for 10 min are shown in Fig. 7C and D.

To test the *in vivo* photodynamic therapy effects of NYF@Sn, we established an animal model of 4T1 tumor cells. The tumor-bearing mice were divided into two groups: a blank group was used as a control and the treatment group was treated with NYF@Sn under a NIR laser. Mouse body weight (Fig. 8A) and tumor size (Fig. 8B) changes were recorded every two days during the treatment cycle. Eight days later, we found that the tumors of the control mice grew to 120 mm³, while the tumors of the treated mice were reduced to 10 mm³. This shows that

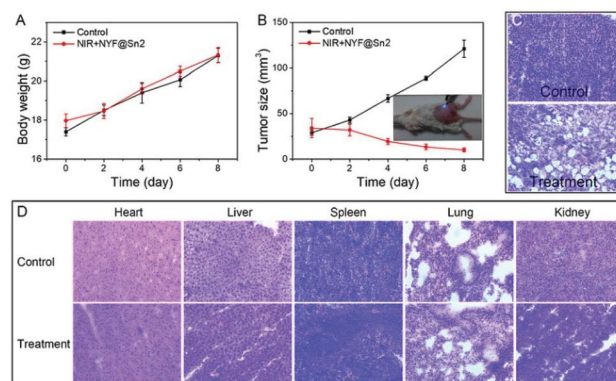


Fig. 8 (A) Body weight and (B) tumor size changes in mice from control and treatment groups over 8 days. H&E stained images of (C) a tumor and (D) heart, liver, spleen, lung, and kidney samples from the control group and NYF@Sn treatment group.

photodynamic therapy with NYF@Sn worked. Fig. 8C shows H&E stained tissue sections of tumor and major organ samples from the blank and treated mice group. As can be seen from the tumor sections, many necrotic and apoptotic cells appear in the tumor tissue after treatment compared with the control group. In the organs of heart, liver, spleen, lung and kidney, normal cells are found in both mouse groups and no abnormal necrosis is detected (Fig. 8D).

Conclusions

In summary, we synthesized a series of lanthanide-semiconductor PDT agents in the form of NaYF₄:Yb,Tm@SnO₂ and NaYF₄:Yb,Tm@TiO₂, in which SnO₂ and TiO₂ act as photosensitizers. For the two core-shell composites, both the porous SnO₂ and TiO₂ shells are relatively uniformly coated on the surface of NYF. NYF@Sn and NYF@Ti have good photodynamic characteristics under 980 nm laser irradiation; the produced ROS significantly weakened the characteristic absorbance peak of DPBF, the survival rates of cells incubated with the photosensitizers were reduced after NIR irradiation, and the growth of living tumors was inhibited after periodic light therapy. This combination of upconversion materials and semiconducting photosensitizers, using near-infrared light with its deeper penetration of biological tissue as the excitation light and semiconductors with higher energy conversion efficiencies as photosensitizers, can be better applied to future photodynamic therapy research.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by the Natural Science Foundation of China (NSFC 81801744), the National Key R&D Program of

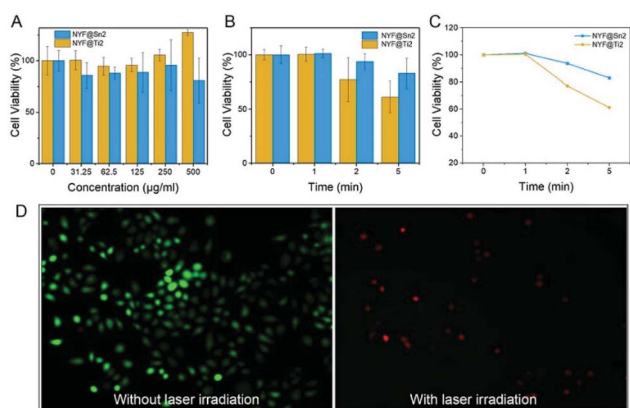


Fig. 7 The viability of cells incubated with NYF@Sn2 and NYF@Ti2 at different concentrations (A) without laser irradiation and (B) following different irradiation exposure times, obtained using MTT assays. AM/PI stained images of cells (C) before and (D) after PDT treatment under 980 nm laser irradiation.



China Grant (No. 2017YFA0205202), and the Fundamental Research Funds for the Central Universities.

References

- Z. Zhou, J. Song, L. Nie and X. Chen, *Chem. Soc. Rev.*, 2016, **45**, 6597–6626.
- A. Bansal, F. Yang, T. Xi, Y. Zhang and J. S. Ho, *Proc. Natl. Acad. Sci. U. S. A.*, 2018, **115**, 1469–1474.
- D. Dolmans, D. Fukumura and R. K. Jain, *Nat. Rev. Cancer*, 2003, **3**, 380–387.
- G. Zheng, J. Chen, K. Stefflova, M. Jarvi, H. Li and B. C. Wilson, *Proc. Natl. Acad. Sci. U. S. A.*, 2007, **104**, 8989–8994.
- M. Feng, R. Lv, L. Xiao, B. Hu, S. Zhu, F. He, P. Yang and J. Tian, *Inorg. Chem.*, 2018, **57**, 14594–14602.
- R. Lv, P. Yang, B. Hu, J. Xu, W. Shang and J. Tian, *ACS Nano*, 2017, **11**, 1064–1072.
- N. M. Idris, M. K. Gnanasammandhan, J. Zhang, P. C. Ho, R. Mahendran and Y. Zhang, *Nat. Med.*, 2012, **18**, 1580–1586.
- L.-S. Lin, J. Song, L. Song, K. Ke, Y. Liu, Z. Zhou, Z. Shen, J. Li, Z. Yang, W. Tang, G. Niu, H.-H. Yang and X. Chen, *Angew. Chem., Int. Ed.*, 2018, **57**, 4902–4906.
- C. Zhang, K. Zhao, W. Bu, D. Ni, Y. Liu, J. Feng and J. Shi, *Angew. Chem., Int. Ed.*, 2015, **54**, 1770–1774.
- E. Ju, K. Dong, Z. Chen, Z. Liu, C. Liu, Y. Huang, Z. Wang, F. Pu, J. Ren and X. Qu, *Angew. Chem., Int. Ed.*, 2016, **55**, 11467–11471.
- C. Seidl, J. Ungelenk, E. Zittel, T. Bergfeldt, J. P. Sleeman, U. Schepers and C. Feldmann, *ACS Nano*, 2016, **10**, 3149–3157.
- J. Hu, Y. a. Tang, A. H. Elmenoufy, H. Xu, Z. Cheng and X. Yang, *Small*, 2015, **11**, 5860–5887.
- I. S. Turan, D. Yildiz, A. Turksoy, G. Gunaydin and E. U. Akkaya, *Angew. Chem., Int. Ed.*, 2016, **55**, 2875–2878.
- J. Xu, A. Gulzar, P. Yang, H. Bi, D. Yang, S. Gai, F. He, J. Lin, B. Xing and D. Jin, *Coord. Chem. Rev.*, 2019, **381**, 104–134.
- X. Ai, L. Lyu, Y. Zhang, Y. Tang, J. Mu, F. Liu, Y. Zhou, Z. Zuo, G. Liu and B. Xing, *Angew. Chem., Int. Ed.*, 2017, **56**, 3031–3035.
- M.-H. Chan and R.-S. Liu, *Nanoscale*, 2017, **9**, 18153–18168.
- X. Wang, S. S. Huang, W. D. W. Heston, H. Guo, B.-C. Wang and J. P. Basilion, *Mol. Cancer Ther.*, 2014, **13**, 2595–2606.
- X. Yue, Q. Zhang and Z. Dai, *Adv. Drug Delivery Rev.*, 2017, **115**, 155–170.
- W. Zheng, P. Huang, Z. Gong, D. Tu, J. Xu, Q. Zou, R. Li, W. You, J.-C. G. Bunzli and X. Chen, *Nat. Commun.*, 2018, **9**, 3462.
- Q. Gong, R. Zou, J. Xing, L. Xiang, R. Zhang and A. Wu, *Adv. Sci.*, 2018, **5**, 1700664.
- J. Zhou, Q. Liu, W. Feng, Y. Sun and F. Li, *Chem. Rev.*, 2015, **115**, 395–465.
- K. Liu, X. Liu, Q. Zeng, Y. Zhang, L. Tu, T. Liu, X. Kong, Y. Wang, F. Cao, S. A. G. Lambrechts, M. C. G. Aalders and H. Zhang, *ACS Nano*, 2012, **6**, 4054–4062.
- C. Chen, F. Wang, S. Wen, Q. P. Su, M. C. L. Wu, Y. Liu, B. Wang, D. Li, X. Shan, M. Kianinia, I. Aharonovich, M. Toth, S. P. Jackson, P. Xi and D. Jin, *Nat. Commun.*, 2018, **9**, 3290.
- J. Shen, G. Chen, T. Y. Ohulchanskyy, S. J. Kesseli, S. Buchholz, Z. Li, P. N. Prasad and G. Han, *Small*, 2013, **9**, 3213–3217.
- M. Haase and H. Schaefer, *Angew. Chem., Int. Ed.*, 2011, **50**, 5808–5829.
- X. Liu, C.-H. Yan and J. A. Capobianco, *Chem. Soc. Rev.*, 2015, **44**, 1299–1301.
- R. Lv, M. Feng and W. J. Parak, *Nanomaterials*, 2018, **8**, 989.
- G. Tian, X. Zheng, X. Zhang, W. Yin, J. Yu, D. Wang, Z. Zhang, X. Yang, Z. Gu and Y. Zhao, *Biomaterials*, 2015, **40**, 107–116.
- J. V. Garcia, J. Yang, D. Shen, C. Yao, X. Li, R. Wang, G. D. Stucky, D. Zhao, P. C. Ford and F. Zhang, *Small*, 2012, **8**, 3800–3805.
- Y. Wang, S. Song, J. Liu, D. Liu and H. Zhang, *Angew. Chem., Int. Ed.*, 2015, **54**, 536–540.
- L.-D. Sun, Y.-F. Wang and C.-H. Yan, *Acc. Chem. Res.*, 2014, **47**, 1001–1009.
- A. Escudero, A. I. Becerro, C. Carrillo-Carrion, N. O. Nunez, M. V. Zyuzin, M. Laguna, D. Gonzalez-Mancebo, M. Ocana and W. J. Parak, *Nanophotonics*, 2017, **6**, 881–921.
- H. Wen, H. Zhu, X. Chen, T. F. Hung, B. Wang, G. Zhu, S. F. Yu and F. Wang, *Angew. Chem., Int. Ed. Engl.*, 2013, **52**, 13419–13423.
- W. Fan, W. Bu and J. Shi, *Adv. Mater.*, 2016, **28**, 3987–4011.
- J. Wang, R. Deng, M. A. MacDonald, B. Chen, J. Yuan, F. Wang, D. Chi, T. S. A. Hor, P. Zhang, G. Liu, Y. Han and X. Liu, *Nat. Mater.*, 2014, **13**, 157–162.
- H. H. Gorris and O. S. Wolfbeis, *Angew. Chem., Int. Ed.*, 2013, **52**, 3584–3600.
- R. Bardhan, S. Lal, A. Joshi and N. J. Halas, *Acc. Chem. Res.*, 2011, **44**, 936–946.
- D. Wang, D. Wang, A. Kuzmin, A. Pliss, W. Shao, J. Xia, J. Qu and P. N. Prasad, *Adv. Opt. Mater.*, 2018, **6**, 1701142.
- Y. Yang, Q. Shao, R. Deng, C. Wang, X. Teng, K. Cheng, Z. Cheng, L. Huang, Z. Liu, X. Liu and B. Xing, *Angew. Chem., Int. Ed.*, 2012, **51**, 3125–3129.
- S. Zeng, Z. Yi, W. Lu, C. Qian, H. Wang, L. Rao, T. Zeng, H. Liu, H. Liu, B. Fei and J. Hao, *Adv. Funct. Mater.*, 2014, **24**, 4051–4059.
- R. Lv, M. Feng, L. Xiao, J. A. Damasco, J. Tian and P. N. Prasad, *ACS Appl. Bio Mater.*, 2018, **1**, 1505–1511.
- Y. Liu, Y. Liu, W. Bu, C. Cheng, C. Zuo, Q. Xiao, Y. Sun, D. Ni, C. Zhang, J. Liu and J. Shi, *Angew. Chem., Int. Ed.*, 2015, **54**, 8105–8109.
- S. S. Lucky, N. M. Idris, Z. Li, K. Huang, K. C. Soo and Y. Zhang, *ACS Nano*, 2015, **9**, 191–205.
- R. Lv, C. Zhong, R. Li, P. Yang, F. He, S. Gai, Z. Hou, G. Yang and J. Lin, *Chem. Mater.*, 2015, **27**, 1751–1763.
- R. Song, M. Zhang, Y. Liu, Z. Cui, H. Zhang, Z. Tang, X. Chen, H. Wu, Z. Yao, M. He and W. Bu, *Biomaterials*, 2018, **175**, 123–133.
- Z. Yin, D. Chen, J. Zou, J. Shao, H. Tang, H. Xu, W. Si and X. Dong, *ChemistrySelect*, 2018, **3**, 4366–4373.



- 47 G. Lan, K. Ni, Z. Xu, S. S. Veroneau, Y. Song and W. Lin, *J. Am. Chem. Soc.*, 2018, **140**, 5670–5673.
- 48 L. Du, S. Liao, H. A. Khatib, J. F. Stoddart and J. I. Zink, *J. Am. Chem. Soc.*, 2009, **131**, 15136–15142.
- 49 Y. Liu, Y. Liu, W. Bu, Q. Xiao, Y. Sun, K. Zhao, W. Fan, J. Liu and J. Shi, *Biomaterials*, 2015, **49**, 1–8.
- 50 W. Xu, J. Ding, L. Li, C. Xiao, X. Zhuang and X. Chen, *Chem. Commun.*, 2015, **51**, 6812–6815.
- 51 Y. Sun, Y. Chen, L. Tian, Y. Yu, X. Kong, J. Zhao and H. Zhang, *Nanotechnology*, 2007, **18**, 447.
- 52 Q. Tian, W. Yao, W. Wu, J. Liu, Z. Wu, L. Liu, Z. Dai and C. Jiang, *ACS Sustainable Chem. Eng.*, 2017, **5**, 10889–10899.
- 53 Z. Xu, M. Quintanilla, F. Vetrone, A. O. Govorov, M. Chaker and D. Ma, *Adv. Funct. Mater.*, 2015, **25**, 2950–2960.
- 54 R. Lv, D. Yang, P. Yang, J. Xu, F. He, S. Gai, C. Li, Y. Dai, G. Yang and J. Lin, *Chem. Mater.*, 2016, **28**, 4724–4734.

