



Cite this: *RSC Adv.*, 2019, 9, 14782

# Adsorption and anticorrosive behavior of aromatic epoxy monomers on carbon steel corrosion in acidic solution: computational studies and sustained experimental studies†

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Herein, the synthesis, characterization and corrosion inhibition effectiveness of two aromatic epoxy monomers (AEMs) namely, 2-(oxiran-2-yl-methoxy)-*N,N*-bis(oxiran-2-yl-methyl)aniline (AEM1) and *N,N*-bis(oxiran-2-ylmethyl)-2-((oxiran-2-ylmethyl) thio)aniline (AEM2), in carbon steel corrosive dissolution in 1 M HCl solution is investigated using computational and experimental techniques. AEM1 and AEM2 were characterized using FT-IR, <sup>1</sup>H NMR and <sup>13</sup>C NMR spectroscopy techniques. Electrochemical results demonstrated that AEMs act as reasonably good corrosion inhibitors for carbon steel in 1 M HCl medium and their effectiveness followed the sequence: AEM2 (95.4%) > AEM1 (94.3%). A PDP study showed that AEMs act as mixed-type inhibitors with slight anodic predominance. Adsorption of the AEMs obeyed the Langmuir isotherm model. Interactions between AEMs and the metallic surface was further studied using DFT and MD simulations that give several computational parameters such as *I*, *A*, *E*<sub>HOMO</sub>, *E*<sub>LUMO</sub>, Δ*E*, δ, χ, ρ, σ, η, Δ*N* and *E*<sub>ads</sub>. The experimental and computational results were in good agreement and well complimented each other.

Received 5th March 2019

Accepted 17th April 2019

DOI: 10.1039/c9ra01672d

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## 1. Introduction

Carbon and mild steel alloys have great importance in a variety of industrial applications because of their excellent mechanical properties and low cost. However, they are highly unstable in aggressive environments mostly during acidic cleaning processes.<sup>1–5</sup> A review of the literature shows that the presence

of organic compounds having heteroatoms (such as P, S, O and N) and homo- and hetero-atomic multiple bonds such as  $\text{>C}\equiv\text{N}$ ,  $\text{>C}=\text{C}<$ ,  $\text{>N}=\text{O}$ ,  $\text{>N}=\text{N}<$ ,  $\text{>C}=\text{S}$ ,  $\text{>C}=\text{O}$ ,  $\text{>C}=\text{N}$  etc. minimize the probability and rate of corrosive dissolution. The non-bonding and  $\pi$ -electrons of heteroatoms and multiple bonds offer co-ordination bonding with d-orbitals of the metals thereby they adsorb and form protective hydrophobic films.<sup>6–10</sup> Epoxy monomers are widely used in several industrial domains such as high performance adhesives, coatings, construction, architecture, automotive and aeronautical due to their good adhesion ability to the metal surfaces, excellent chemical resistance and good durability.<sup>11,12</sup> Epoxy monomers are well established oxygen containing three membered heterocyclic compounds.<sup>13,14</sup> Recently, use of corrosion inhibitors based on some aromatic epoxy monomers (AEMs) has gained particular importance because of their relatively high protectiveness, the facile route of synthesis and high purity.<sup>15,16</sup> We herein, synthesized two aromatic epoxy monomers (AEMs) as tri-functional molecules designated as AEM1 and AEM2 and tested their anticorrosive properties on the corrosion of carbon steel in 1 M HCl media. Novelty, of the study is lying behind the facts that these AEMs have never be tested as corrosion inhibitors for steel alloys and they offer reasonably high inhibition efficiency at relatively low concentration. Obviously, AEMs are associated with extensive conjugation in the form of heteroatoms (N, S, and O) and  $\pi$ -electrons by way of which they can

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† Electronic supplementary information (ESI) available. See DOI: 10.1039/c9ra01672d



adsorb effectually and obstruct corrosion thereafter. Both the tested AEMs have similar molecular structure and they differ from each other with respect to the nature of heteroatoms. Therefore, present study is mainly focused and designed to demonstrate the effect of heteroatoms and their electronegativity on the corrosion inhibition of carbon steel in acidic medium. Both inhibitors were synthesized using commercially available cheap and non-toxic chemicals and solvent (isopropyl alcohol/water) in high yield. The inhibition effect of AEMs was studied using experimental and computational methods using PDP, EIS, DFT and MD methods. Computational and experimental outcomes showed that inhibitor containing relatively less electronegative sulfur atom (AEM2) acts as superior corrosion inhibitor as compared to the inhibitor having more electronegative oxygen atom (AEM1). The results derived from various techniques were in good agreement.

## 2. Experimental

### 2.1. Materials

Carbon steel employed in our previous reports has been used in the present study.<sup>14,16</sup> Test solution (1 M HCl) was prepared by dilution of 37% analytical grade HCl. The aromatic epoxy monomers were synthesized by the practice described in the literature.<sup>17–20</sup> Schematic diagram for the preparation of the aromatic epoxy monomers (AEMs) is shown in Fig. 1. In brief, a mixture of an aromatic epoxy monomers compound ( $10^{-2}$  mol) and ethanol (10 mL) was added in condenser fitted two-necked RB flask. In the above solution, 2.5 mL of epichlorohydrin was added at once. After stirring the resulting mixture at 70 °C for 4 h, mixture was allowed to cool down to 40 °C. After that, 3 mL of  $\text{NEt}_3$  was added in the mixture and stirred again for 3 h. The excess epichlorohydrin and the residual solvent were removed using a rotary evaporator in order to obtain the desired viscose solution of resin.

Both aromatic epoxy monomers were characterized using  $^1\text{H}$  NMR,  $^{13}\text{C}$  NMR and FT-IR spectroscopic characterization techniques is shown in Fig. SI 1 and SI 2.† AEM1: yield 92%,  $^1\text{H}$ -NMR (DMSO- $d_6$ , 300 MHz):  $\delta$  ppm = 2.7 2, 3.8–2, 6.3 (dd, 8H,  $\text{CH}_2$  epoxy cycle) (A,B), 2, 7.7 (m, 2H, CH epoxy cycle) (X), 2.81 (m,

1H, CH epoxy cycle) (X), 3, 3.6–3, 6.0 (dd, 4H,  $\text{N-CH}_2$ ) (C,D), 3, 9.5–4, 2.0 (dd, 2H,  $\text{O-CH}_2$ ) (C,D), 6, 6.5–6, 9.2 (s, 4H aromatic ring) (Ar);  $^{13}\text{C}$ -NMR:  $\delta$  ppm = 44, 2–45, 6 (s,  $\text{CH}_2$  epoxy cycle) (A), 51, 1–51, 9 (s, CH epoxy cycle) (X), 62 (s,  $\text{N-CH}_2$ ) (C), 69, 4 (s,  $\text{O-CH}_2$ ) (C), 113–122 (s, CH aromatic) (Ar), 138, 7; 145, 1 (s, C–C aromatic) (Ar); FTIR-ATR ( $\text{cm}^{-1}$ ): 3257 ( $\text{O-H}$  and secondary  $\text{N-H}$ ), 2967, 2925, 2877 (C–H vibration), 1600, 1500, 1450 (C=C aromatic), 1295, 1370 (C–N vibration), 1240, 1216, 1089 (C–O), 920, 835 (epoxy groups). AEM2: yield 88%,  $^1\text{H}$ -NMR (DMSO- $d_6$ , 300 MHz):  $\delta$  ppm = 2, 3.8–2, 6.3 (dd, 8H,  $\text{CH}_2$  epoxy cycle) (A,B), 2, 7.7 (m, 2H, CH epoxy cycle) (X), 3, 0.4 (m, 1H, CH epoxy cycle) (X), 3, 3.6–3, 6.1 (dd, 4H,  $\text{N-CH}_2$ ) (C,D), 2, 7.7–3, 0.2 (dd, 2H,  $\text{O-CH}_2$ ) (C,D), 6.65–7, 2.1 (s, 4H aromatic ring) (Ar);  $^{13}\text{C}$ -NMR:  $\delta$  ppm = 45, 6–46, 8 (s,  $\text{CH}_2$  epoxy cycle) (A), 51, 1–53, 4 (s, CH epoxy cycle) (X), 61, 3 (s,  $\text{N-CH}_2$ ) (C), 40, 8 (s,  $\text{S-CH}_2$ ) (C), 115–134, 6 (s, CH aromatic) (Ar), 117, 9; 135 (s, C–C aromatic) (Ar); FTIR-ATR ( $\text{cm}^{-1}$ ): 3257 (NH/OH residual), 2450 (S–H residual), 2967, 2925, 2877 (C–H vibration), 1600, 1500, 1450 (C=C aromatic), 1000–1400 (C–N, C–O, C–S vibration), 950, 835 (epoxy groups)).

### 2.2. Electrochemical measurements

Before performing the experiments, the steel specimens were polished with grit sand papers using different grades (180 to 1000), rinsed with alcohol and distilled water. Stock solution of the inhibitors (AEM1 and AEM2) was prepared by dissolving them in small amount (nearly 2 mL) of isopropyl alcohol followed by the addition of tested electrolyte. The electrochemical behavior of an aromatic epoxy monomers (AEMs) was studied using Potentiostat (BioLogic SP-200) instrument as designated in our prior reports.<sup>11,12</sup> In order to get more accuracy and reproducibility of experimental data, the electrochemical studies were performed triply at each tested concentration of the AEM1 and AEM2 and mean values are reported.

### 2.3. Computational details

For DFT study, both AEM1 and AEM2 were geometrically optimized using B3LYP/6-31+G(d,p) level in both gas phase and in aqueous solution as described elsewhere. All calculations were performed using the G09 suite program.<sup>21–28</sup> To validate the

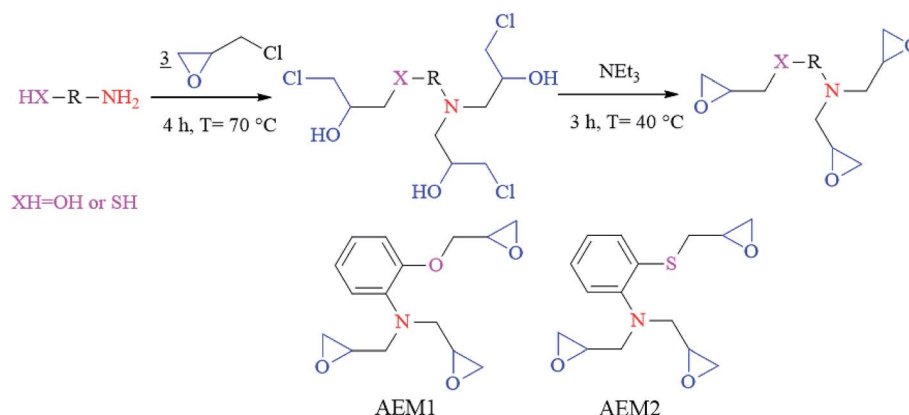


Fig. 1 Schematic outline for the synthesis of two aromatic epoxy monomers AEM1 and AEM2.



experimental results and to gain a deeper insight into the reactivity of the studied AEM1 and AEM2 compounds, the calculated quantum global parameters, which were extracted based on the values of the highest occupied and lowest unoccupied molecular orbitals ( $E_{\text{HOMO}}$  and  $E_{\text{LUMO}}$ ) were used. Different DFT parameters are calculated as follows:<sup>29–43</sup>

The ionization potential:

$$I = -E_{\text{HOMO}}, \quad (1)$$

The electron affinity:

$$A = -E_{\text{LUMO}}, \quad (2)$$

The energy difference between  $E_{\text{LUMO}}$  and  $E_{\text{HOMO}}$ ,

$$\Delta E = (E_{\text{HOMO}} - E_{\text{LUMO}}), \quad (3)$$

The electronegativity:

$$\chi = (I + A)/2, \quad (4)$$

The global hardness:

$$(\eta = (I - A)/2) \text{ and softness } (\sigma = 1/\eta), \quad (5)$$

The fraction of electrons transferred:

$$\Delta N = \frac{\varphi_{\text{Fe}} - \chi_{\text{inh}}}{2(\eta_{\text{Fe}} + \eta_{\text{inh}})}, \quad (6)$$

The initial molecule–metal interaction energy:

$$\Delta\psi = \frac{(\chi_{\text{Fe}} - \chi_{\text{inh}})^2}{2(\eta_{\text{Fe}} + \eta_{\text{inh}})}, \quad (7)$$

The electron transfer electron back-donation:

$$\Delta E_{\text{back-donation}} = -\frac{\eta}{4} \quad (8)$$

where  $\varphi_{\text{Fe}}$ ,  $\chi_{\text{Fe}}$  and  $\eta_{\text{Fe}}$  are the work function, electronegativity and hardness of the iron, respectively. In this study, three different plans namely, (100), (110) and (111) were considered for the interaction with AEM1 and AEM2. The work function values of (100), (110) and (111) are 3.91, 4.82 and 3.88 eV, respectively. The electronegativity ( $\chi_{\text{Fe}}$ ) and hardness ( $\eta_{\text{Fe}}$ ) of iron are selected as 7.0 and 0 eV, respectively.

#### 2.4. Molecular dynamics simulation

The interaction between AEM1 and AEM2 in their neutral as well as in their protonated (cationic) forms and metal surface was studied employing the molecular dynamics (MD) simulations. MD simulations were carried out using Forcite module of Materials Studio 8.0 program developed by BIOVIA Inc. MD simulations were done in a simulation box of 1.98 nm × 1.98 nm × 4.01 nm size with periodic boundary conditions. The box consisted of a lower Fe slab and an upper solvent layer (containing 500 water molecules and one inhibitor molecule of AEM1 and AEM2). For the iron substrate, Fe (110) was selected as explored surface because it has a density packed structure and is the most stable. The Fe(110) surface was modelled with a six-layer slab model. In this model, there were 64 iron atoms (8 × 8) in each layer of unit cell. Particulars of similar simulations have also been documented in our aforementioned publications.<sup>44–48</sup>

### 3. Results and discussion

#### 3.1. Electrochemical measurements

**3.1.1. Open circuit potential (OCP).** Fig. 2 represents the OCP versus time (for 1800 s) curves for dissolution of metal in acidic medium with and without AEM1 and AEM2. Formation of straight lines showed the establishment of steady state potential and dissolution of surface oxide layers of  $\text{Fe}_2\text{O}_3$  and

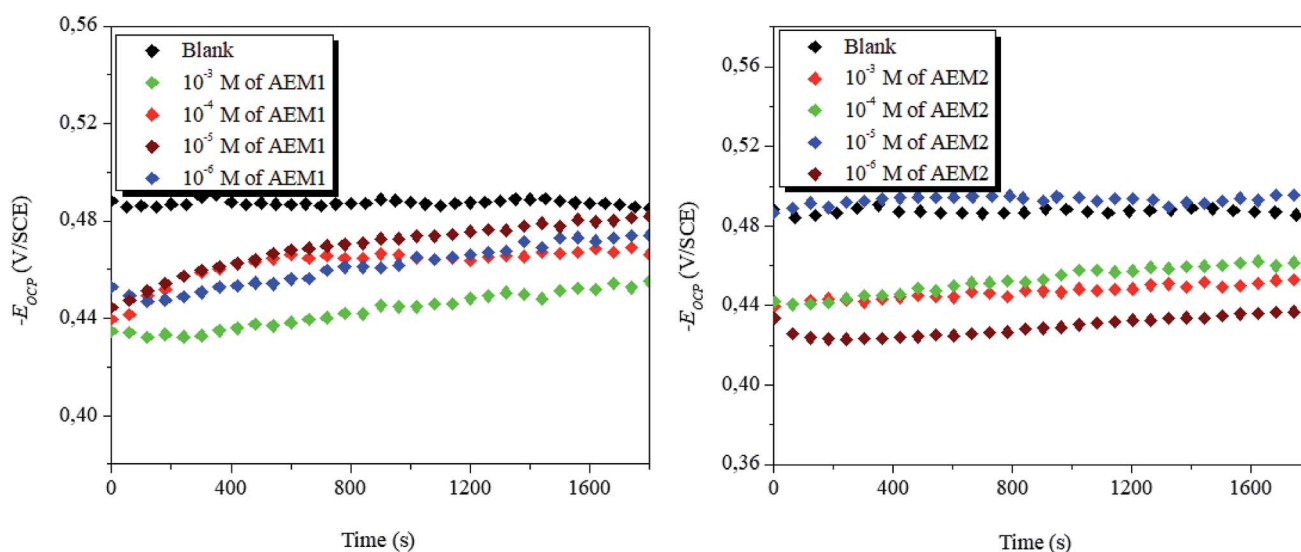


Fig. 2 Variation of  $E_{\text{OCP}}$  vs. with time for carbon steel corrosion in 1 M HCl with and without AEM1 and AEM2 molecules at 298 K.



$\text{Fe}_3\text{O}_4$ .<sup>49</sup> It is also observable that at most of the tested concentrations, OCP *versus* time curves are shifted towards more positive (anodic) directions. This implies that AEM1 and AEM2 are acted as predominantly anodic type inhibitors with predominant anticorrosive effect on anodic reaction.<sup>49</sup>

**3.1.2. PDP study.** The anodic and cathodic polarizations curves with and without AEMs are shown in Fig. 3 and PDP parameters are presented in Table 1. The surface coverage values at different concentrations of the AEMs were derived by dividing the percentage of efficiency by 100. Extrapolation of the linear segments of the PDP curves gives the values of  $i_{\text{corr}}$  through which inhibition efficiency ( $\eta_{\text{PDP}}\%$ ) of AEMs was calculated as follows:<sup>14,16</sup>

$$\eta_{\text{PDP}}\% = \left(1 - \frac{i_{\text{corr}}}{i_{\text{corr}}^0}\right) \times 100 \quad (9)$$

whereas,  $i_{\text{corr}}^0$  and  $i_{\text{corr}}$  represent the corrosion current densities without and with the AEMs correspondingly.

From the results it can be seen that  $E_{\text{corr}}$  values for inhibited cases did not showed any significant change as compared to the  $E_{\text{corr}}$  value of uninhibited case which implies that both AEM1 and AEM2 acted as mixed type corrosion inhibitors. Observation of the results showed that presence of the AEMs affected both anodic as well as cathodic reactions without affecting the common characteristics of PDP curves.<sup>50</sup> This observation showed that AEM1 and AEM2 inhibit metallic corrosion by blocking the active sites without affecting mechanism of anodic and cathodic reactions. Further, it can also be seen that at most of the studied concentrations the anodic Tafel slope values ( $\beta_a$ ) showed slightly higher shift as compared to the values of cathodic Tafel slope values ( $\beta_c$ ) which suggested that both AEM1 and AEM2 have marked effect on anodic dissolution as compared to the cathodic hydrogen evolution reaction. It can also be noted that values of  $E_{\text{corr}}$  are shifted towards more positive *i.e.* anodic direction with the presence the different

concentration of AEMs. On this basis, it can be assumed that AEMs are acted as mixed type inhibitors with anodic predominance. Blocking of the active sites of metals can be resulted due to the adsorption of the AEM1 and AEM2 can using their several heteroatoms (N, S and O) that offer non-bonding electrons and aromatic rings that offer  $\pi$ -electrons for metal-inhibitor interactions. Apart from the metal-inhibitors (AEMs) interactions, presence of heteroatoms in the form of polar functional group(s), enhances the solubility of the investigated inhibitor molecules.<sup>51</sup>

**3.1.3. EIS study.** The Nyquist plots for metal dissolution in 1 M HCl with and without AEM1 and AEM2 are presented in Fig. 4. It is easy to notice that all these Nyquist plots curves present a single capacitive loop in the frequency range of 100 kHz to 10 mHz, which is usually related to charge transfer phenomenon.<sup>52</sup> The imperfection in the semicircle is resulted because of the frequency dispersion that resulted mainly due to the rough and inhomogeneous metallic surfaces.<sup>53</sup> Careful observation of the Nyquist plots showed that diameters of the plots are increasing on increasing the concentration of AEM1 and AEM2 and maximum increase was observed at  $10^{-3}$  M concentration. This observation suggested that resistance for the metallic corrosion in acidic medium of 1 M HCl is increasing on increasing the concentration of AEM1 and AEM2. The fitted Nyquist and Bode plots and equivalent circuit used for the evaluation of EIS parameters are shown in Fig. 5. The CPE is defined as the below eqn (10):<sup>54,55</sup>

$$Z_{\text{CPE}} = Y_0^{-1}(i\omega)^{-n} \quad (10)$$

In the above equation, all symbols have their usual meaning.<sup>45</sup> The capacitance values ( $C_{\text{dl}}$ ) is calculated using eqn (11):<sup>56,57</sup>

$$C_{\text{dl}} = (Y_0 R_{\text{ct}}^{1-n})^{1/n} \quad (11)$$

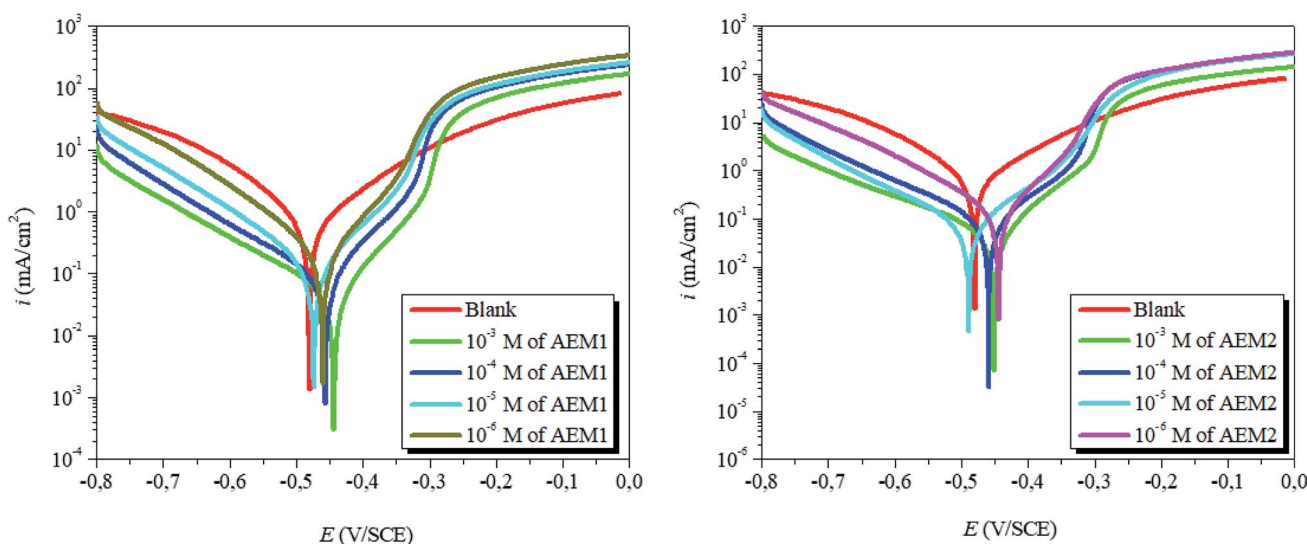


Fig. 3 PDP curves for carbon steel corrosion in 1 M HCl solution in the absence and presence of different concentrations of aromatic epoxy monomers AEM1 and AEM2 at 298 K.



**Table 1** PDP parameters ( $\pm$ SD) for carbon steel corrosion in 1 M HCl solution in the absence and presence of different concentrations of AEM1 and AEM2

Inh	C (M)	$E_{\text{corr}}$ (mV)	$i_{\text{corr}}$ ( $\mu\text{A cm}^{-2}$ )	$\beta_a$ (mV dec $^{-1}$ )	$-\beta_c$ (mV dec $^{-1}$ )	$\eta\%$	$\theta$
Blank	—	-473.80	916.6( $\pm$ 1.78)	163.6( $\pm$ 1.10)	155.0( $\pm$ 1.33)	—	—
AEM1	$10^{-3}$	-443.94	058.96( $\pm$ 1.56)	097.8( $\pm$ 1.08)	180.1( $\pm$ 1.12)	93.6	0.936
	$10^{-4}$	-457.68	097.75( $\pm$ 1.72)	102.2( $\pm$ 1.11)	183.0( $\pm$ 1.17)	89.3	0.893
	$10^{-5}$	-474.14	179.68( $\pm$ 1.85)	116.0( $\pm$ 1.13)	155.9( $\pm$ 1.07)	80.1	0.801
	$10^{-6}$	-456.89	251.78( $\pm$ 1.74)	081.3( $\pm$ 1.07)	141.1( $\pm$ 1.35)	72.5	0.725
AEM2	$10^{-3}$	-451.26	032.88( $\pm$ 1.55)	73.90( $\pm$ 1.18)	166.1( $\pm$ 1.25)	96.4	0.964
	$10^{-4}$	-459.41	094.77( $\pm$ 1.65)	102.7( $\pm$ 1.12)	169.2( $\pm$ 1.14)	89.6	0.896
	$10^{-5}$	-489.01	107.53( $\pm$ 1.16)	134.9( $\pm$ 1.20)	172.1( $\pm$ 1.37)	88.2	0.882
	$10^{-6}$	-444.30	254.31( $\pm$ 1.62)	134.8( $\pm$ 1.16)	171.4( $\pm$ 1.18)	72.2	0.722

The inhibition efficiency  $\eta_{\text{EIS}}\%$  is calculated using the values of polarization resistance  $R_p$  using following relationship:

$$\eta_{\text{EIS}}(\%) = \frac{R_p - R_p^0}{R_p} \times 100 \quad (12)$$

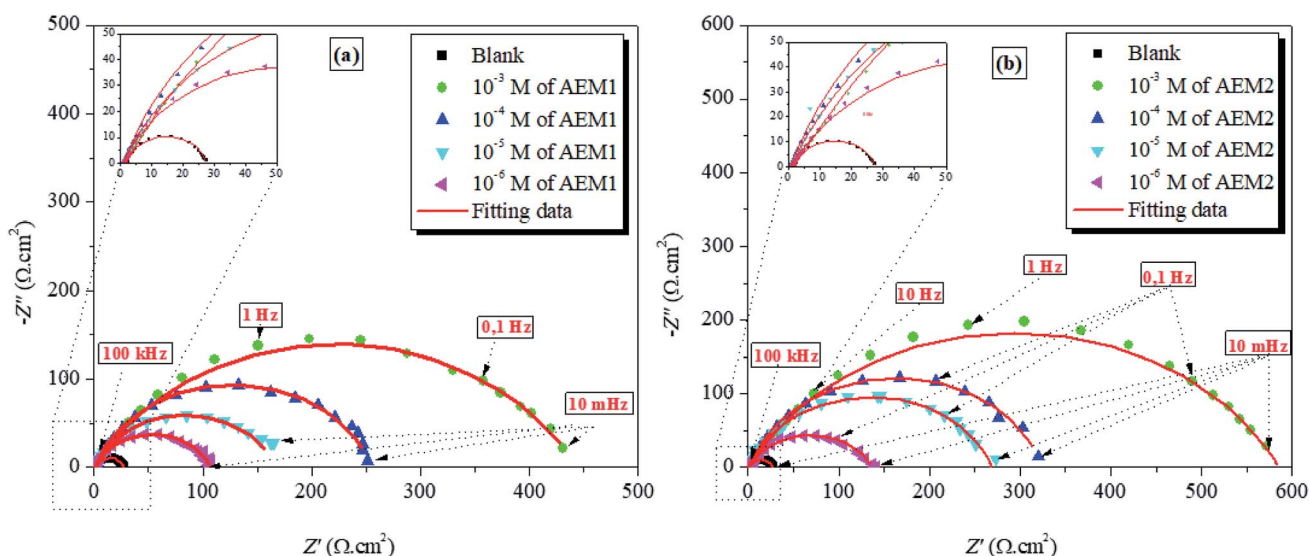
where,  $R_p^0$  and  $R_p$  are the polarization resistances without and with AEMs, respectively. As can be seen from Table 2, the  $R_p$  values are increasing on increasing the AEMs concentrations. It can also be seen that increase in  $R_p$  value at tested concentrations of AEMs are relatively greater of AEM2 as compared to the AEM1 which showed that AEM2 is better corrosion inhibitor as compared to the AEM1.<sup>58</sup> The high inhibition efficiency of the AEM2 is might be resulted due to the presence of less electro-negative sulfur in its molecular structure atom that generally offers strong bonding with the metallic surface. AEM1 and AEM2 showed that highest inhibition efficiencies of 94.3% and 95.4% respectively. Moreover, the  $C_{\text{dl}}$  values diminish on enhancing the concentration for the AEMs which is resulted from the increase in the thickness of the protective layer and/or the decrease in local dielectric constant.<sup>59</sup> Bode modulus plots with and without aromatic epoxy monomers (AEMs) are given in

Fig. 6. Analyze of the figure displays that the magnitude of slope values for the inhabited (by AEMs) curves are greater than that of the uninhabited curve. Among the investigated AEMs, AEM2 (Fig. 6b) showed higher slope value than that of AEM1 (Fig. 6a) which implies that AEM2 has higher effectiveness towards the interaction with metal surface as compared to the AEM1.

The corresponding Bode plots for the corrosion of metal in acidic medium are given in Fig. 6c and d, that show a single maxima in the absence and presence of all concentrations of AEM1 and AEM2 which is a common characteristics of carbon steel corrosion in acidic medium.<sup>60</sup> Increase in the phase angle values of inhibited (by AEMs) Bode plots is resulted due to increase in the metallic surface smoothness due to the adsorption of AEM1 and AEM2 at the interface of metal and electrolyte.<sup>60</sup>

### 3.2. Adsorption studies

In the present study, adsorption behavior of AEM1 and AEM2 is investigated using several commonly used adsorption isotherm models among which most frequently used Langmuir



**Fig. 4** Nyquist (a and b) diagrams for carbon steel in 1 M HCl solution in the absence and presence of different concentrations of aromatic epoxy monomers AEM1 and AEM2 at 298 K.



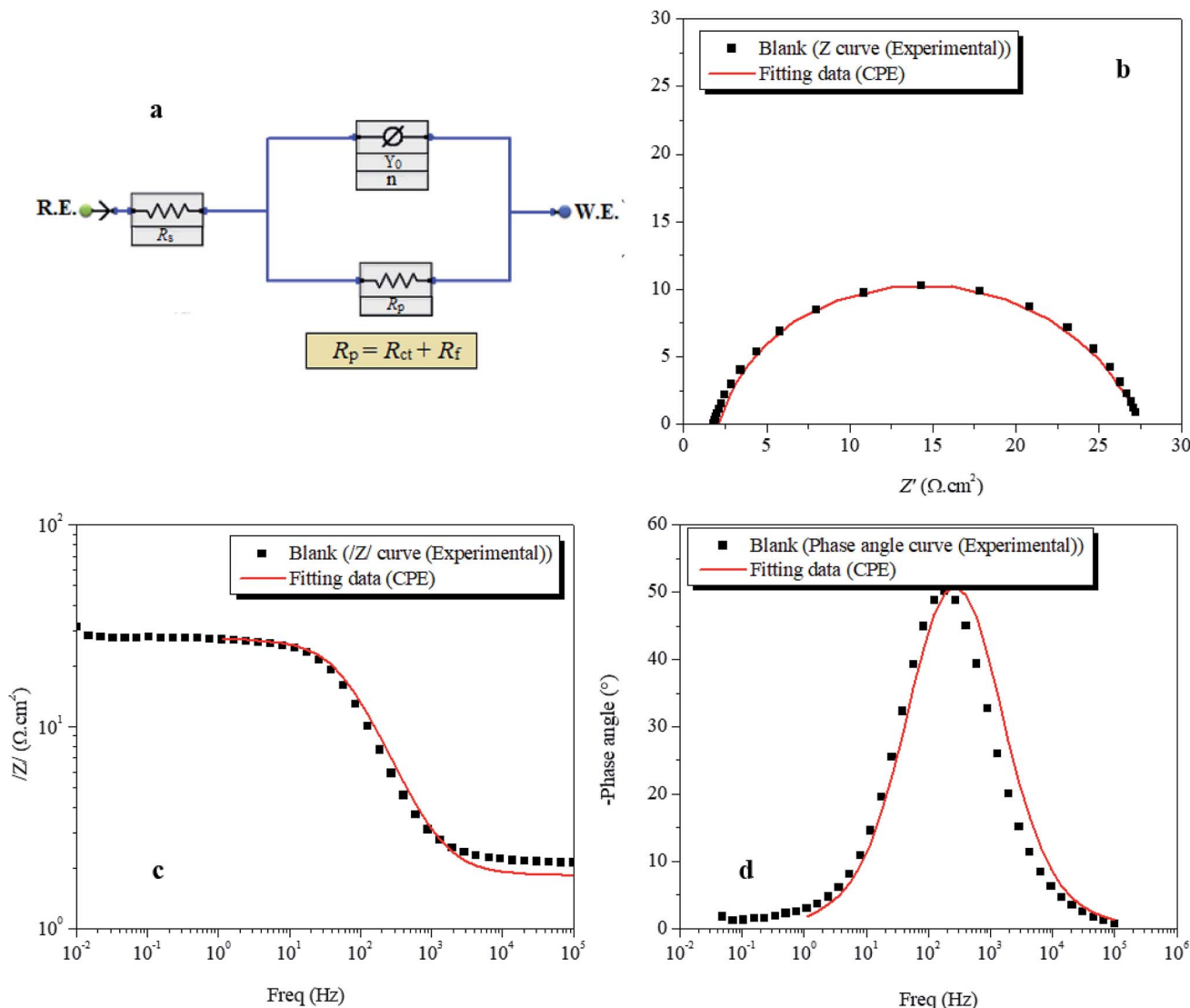


Fig. 5 Equivalent circuit model used to fit the EIS data.

adsorption isotherm gave the best fit.<sup>64</sup> The linear form of Langmuir adsorption isotherm is represented in eqn (13):

$$K_{\text{ads}} C_{\text{inh}} = \frac{\theta}{1 - \theta} \quad (13)$$

In which  $C_{\text{inh}}$  is the concentration of the compound,  $\theta$  is surface coverage degree and  $K_{\text{ads}}$  is the equilibrium constant. The association between  $C_{\text{inh}}$  and  $C_{\text{inh}}/\theta$  yielded straight lines with intercepts of  $K_{\text{ads}}$  as shown in Fig. SI 3.† From the  $K_{\text{ads}}$  values, the  $\Delta G_{\text{ads}}$  values were calculated using eqn (14) and (15):<sup>62</sup>

Table 2 EIS parameters ( $\pm$ SD) for carbon steel corrosion in 1 M HCl solution in the absence and presence of different concentrations of AEM1 and AEM2

Inh	C (M)	$R_s$ ( $\Omega \text{ cm}^2$ )	$R_p$ ( $\Omega \text{ cm}^2$ )	$n$	$C_{\text{dl}}$ ( $\mu\text{F cm}^{-2}$ )	$\eta\%$	$\theta$	$\chi^2$
Blank	—	1.81( $\pm$ 0.03)	25( $\pm$ 1.52)	0.682( $\pm$ 0.01)	485.8( $\pm$ 0.004)	—	—	0.056
AEM1	$10^{-3}$	1.44( $\pm$ 0.07)	441.3( $\pm$ 1.85)	0.842( $\pm$ 0.01)	103.4( $\pm$ 0.008)	94.3	0.943	0.028
	$10^{-4}$	0.68( $\pm$ 0.12)	253.7( $\pm$ 1.75)	0.785( $\pm$ 0.02)	178.6( $\pm$ 0.006)	90.0	0.900	0.095
	$10^{-5}$	1.13( $\pm$ 0.05)	160.3( $\pm$ 1.45)	0.768( $\pm$ 0.03)	216.2( $\pm$ 0.003)	84.4	0.844	0.044
	$10^{-6}$	0.98( $\pm$ 0.13)	106.4( $\pm$ 0.96)	0.762( $\pm$ 0.01)	346.3( $\pm$ 0.009)	76.5	0.765	0.186
AEM2	$10^{-3}$	2.12( $\pm$ 0.07)	584.3( $\pm$ 1.62)	0.881( $\pm$ 0.02)	102.1( $\pm$ 0.001)	95.4	0.954	0.225
	$10^{-4}$	0.70( $\pm$ 0.11)	324.4( $\pm$ 1.65)	0.855( $\pm$ 0.01)	173.8( $\pm$ 0.005)	92.3	0.923	0.455
	$10^{-5}$	1.24( $\pm$ 0.05)	268.2( $\pm$ 1.80)	0.840( $\pm$ 0.02)	213.7( $\pm$ 0.003)	90.7	0.907	0.081
	$10^{-6}$	0.45( $\pm$ 0.08)	136.5( $\pm$ 0.57)	0.798( $\pm$ 0.01)	316.6( $\pm$ 0.007)	81.7	0.817	0.257



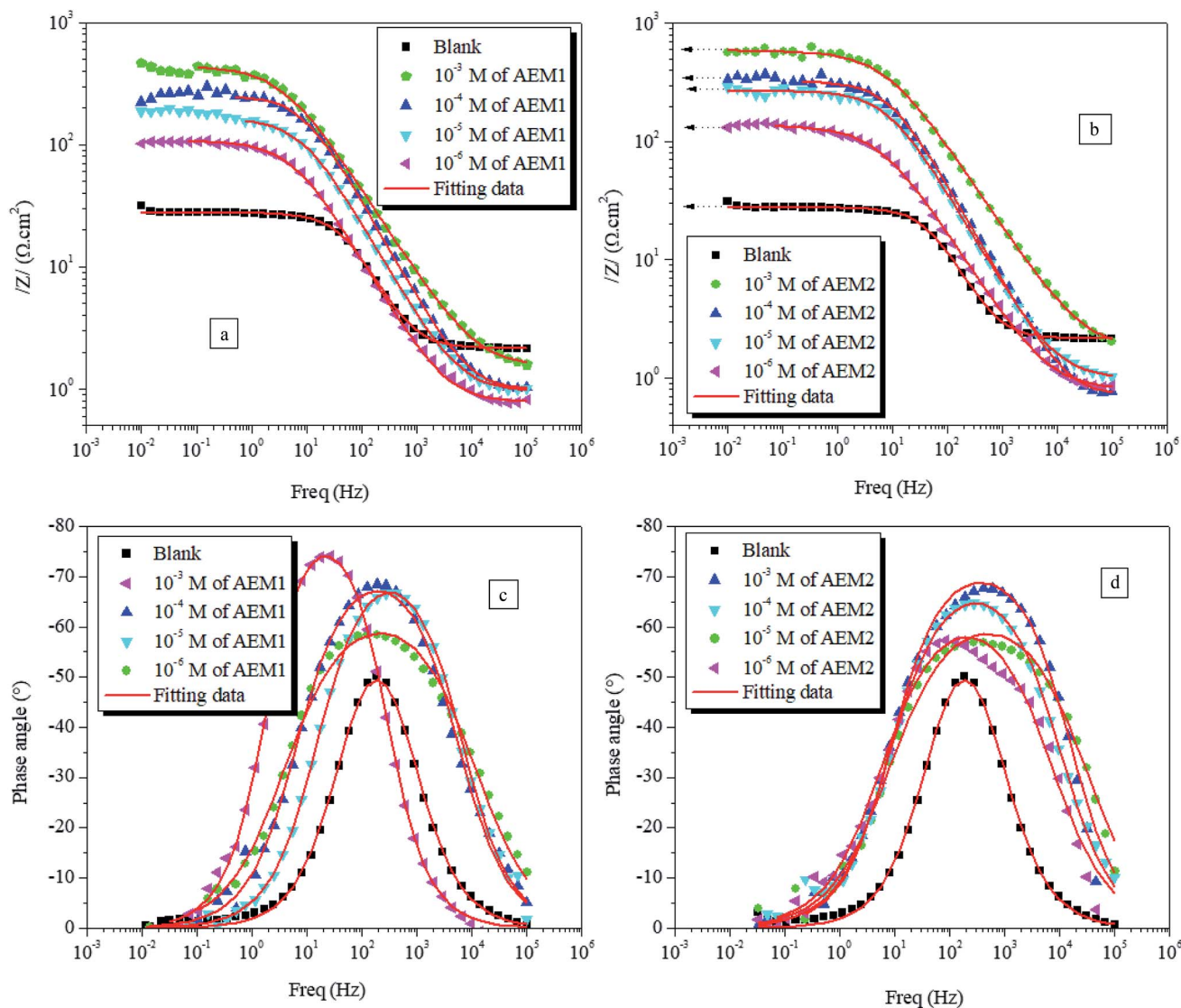


Fig. 6 Bode frequency (a and b) and phase angle (c and d) plots for carbon steel corrosion in the absence and presence of different concentrations of AEM1 and AEM2 at 298 K.

$$K_{\text{ads}} = \frac{1}{55.5} \exp\left(\frac{-\Delta G_{\text{ads}}}{RT}\right) \quad (14)$$

$$\Delta G_{\text{ads}} = -RT \ln(55.5K_{\text{ads}}) \quad (15)$$

where,  $R$  represents universal gas constant, value 55.55 is the molar concentration of water (M) and  $T$  is an absolute temperature. Table SI 1<sup>†</sup> provides the values of  $K_{\text{ads}}$  and  $\Delta G_{\text{ads}}$  obtained for the adsorption of the AEMs on the carbon steel surface. Generally, high negative value of  $\Delta G_{\text{ads}}$  and high positive high value of  $K_{\text{ads}}$  are associated with high inhibitor effectiveness. It is well known that value of  $\Delta G_{\text{ads}}$  either equal to  $-40 \text{ kJ mol}^{-1}$  or more negative is consistent with chemisorption.<sup>61</sup>  $\Delta G_{\text{ads}}$  value in the range of/or more positive than  $-20 \text{ kJ mol}^{-1}$  is associated with a physisorption (electrostatic) interaction.<sup>61</sup> The calculated  $\Delta G_{\text{ads}}$  values for AEM1 and AEM2 are varied between  $-42.03$  and  $-44.20 \text{ kJ mol}^{-1}$  which indicates that both investigated

inhibitors adsorb on metallic surface mostly *via* chemisorption mechanism.<sup>63-65</sup>

### 3.3. Effect of temperature

The inhibition performance of AEMs at their optimum concentrations was tested at different temperatures (298, 308, 318 and 328 K) using PDP measurements and results are presented in Fig. SI 4 and Table SI 2.<sup>†</sup> It can be seen that increase in temperature causes significant increase in the  $i_{\text{corr}}$  without affecting the common features of polarization curves. This is due to the desorption of adsorbed AEMs molecules from the metallic surface at elevated temperatures.<sup>45</sup> The effect of temperature can be demonstrated using Arrhenius and transition state equations:<sup>66-68</sup>

$$i_{\text{corr}} = A \exp\left(-\frac{E_a}{RT}\right) \quad (16)$$



$$i_{\text{corr}} = \frac{RT}{hN} \exp\left(\frac{\Delta S_a}{R}\right) \exp\left(-\frac{\Delta H_a}{RT}\right) \quad (17)$$

where,  $A$  is the pre-exponential factor,  $\Delta S_a$  is the entropy of activation,  $\Delta H_a$  is the enthalpy,  $N$  is the Avogadro number,  $h$  is the Planck constant. Results (Fig. SI 5†) showed that  $E_a$  values are higher in the presence of AEM1 and AEM2 as compared to in their absence which implies that both AEM1 and AEM2 adsorb through physisorption mechanism that takes place during the initial stage of the metal–inhibitors interactions.<sup>69</sup> Positive values of  $\Delta H_a$  suggested that interaction of AEM1 and SEM2 with metallic surface is an endothermic process whereas large negative values of  $\Delta S_a$  suggested the creation of active complex and dissociation of the complex is the rate determining step.<sup>70–72</sup>

### 3.4. Comparative efficiency

Literature study reveals that few of the aromatic epoxy monomers based organic inhibitors have been evaluated for their anti-corrosive behavior for metals and alloys in different electrolytic media.<sup>14–16,71,72</sup> Inhibition effectiveness of some common representatives of the similar inhibitors is presented in Table 3. It can be seen that these types of compounds are acted as good corrosion inhibitors and their molecules structures contain N, O, S and/or P heteroatoms along with the aromatic ring(s).

### 3.5. Theoretical studies

**3.5.1. Quantum chemical calculation.** Fig. 7 and 8 show the optimized structures, the frontier molecular orbitals and electrostatic potential maps of the neutral and the protonated species of the studied compound in aqueous solution. According to FMO theory, the chemical reaction mainly occurred on HOMO and LUMO sites of the molecules. The values of the HOMO and LUMO energies give additional understanding related to the ability of electron donation and acceptance to and from the metal surface. In this sense, the electron dissemination of HOMO and LUMO is a challenging way to forecast the reactivity and, therefore, the capability of the inhibitor compounds to adsorb onto the metal surface. Depending on FMO theory, higher value of  $E_{\text{HOMO}}$  signifies that the inhibitor has a higher ability to transfer its electrons to suitable acceptor molecule whereas converse it true for  $E_{\text{LUMO}}$ .<sup>73–76</sup> Therefore, a lower  $E_{\text{LUMO}}$  value implies that inhibitor is more plausible to accept electrons.<sup>77</sup> The energy gap ( $\Delta E$ ) is a very useful quantum parameter, which measures the chemical reactivity and complexation with the metal surfaces.<sup>78</sup> It was

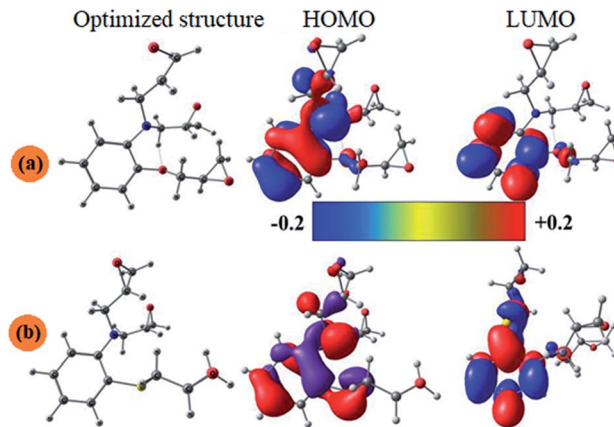


Fig. 7 The optimized molecular structures, HOMO and LUMO of the neutral inhibitor molecules AEM1 (a) and AEM2 (b) using B3LYP using functional and 6-31+G(d,p) basis set.

found that the lower value of the energy gap is generally correlated with high chemical reactivity. The lower energy gap inhibitor, the higher inhibition efficiency one.<sup>75,79,80</sup>

Table 4 displays the computed results of total energies, dipole moments,  $E_{\text{HOMO}}$  and  $E_{\text{LUMO}}$ , and the estimated quantum parameters of the protonated and no protonated structures of the investigated inhibitors in both gas phase and in aqueous medium. In section of Table 4 indicates that the  $E_{\text{HOMO}}$  of AEM2 in its protonated form, in both media, is higher than that of AEM1, which leads us to expect that the protection effectiveness of AEM2 is superior to AEM1. This finding is healthy interrelated with the obtained experimental results (AEM2 > AEM1). Whereas, the order is reserved when the non-protonated species are considered. It is also found that the trend in  $E_{\text{LUMO}}$  values of the protonated species in aqueous phase follows the order: AEM2 < AEM1, which is also well correlated with the experimental findings. On the other hand, there is no correspondence among the experimental results and computed ones when the non-protonated species were taken into account. Based on these results, thereafter, our attention will be focused on the on the results obtained for the protonated species in the aqueous which is similar to the corrosion medium used in the experimental study.

As indicated from Table 4, the values of the energy gap between  $E_{\text{LUMO}}$  (ability to accept) and  $E_{\text{HOMO}}$  (ability to donate) follow the order AEM2 (5.271 eV) < AEM1 (5.463 eV). This finding reflects that AEM2 has the higher inhibition efficiency than AEM1, because lower energy gap causes the improvement

Table 3 Inhibition efficiency comparisons for some traditional corrosion inhibitors

Type of corrosion inhibitor	Type of solution	$C$ (M)	Type of substrate	$\eta_{\text{Max}}$ (%)	Ref.
Hexa(3-methoxy propan-1,2-diol)cyclotri-phosphazene	3% NaCl	$10^{-3}$	Carbon steel	99.0	14
Triglycidyl ether of triethoxytriazine	1 M HCl	$10^{-3}$	Carbon steel	88.0	15
Tetraglycidyl-1,2-aminobenzamide	1 M HCl	$10^{-3}$	Carbon steel	96.0	16
<i>S,S'</i> -Diglycidyl <i>O,O'</i> -dicarbonodithioate of bisphenol A	1 M HCl	$10^{-3}$	Carbon steel	96.0	71
Octaglycidylether tetra-aniline <i>para</i> methylene dianiline	0.5 M H <sub>2</sub> SO <sub>4</sub>	$10^{-3}$	Mild steel	98.0	72



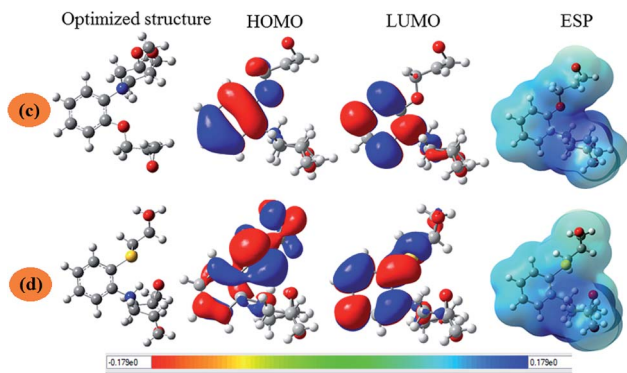


Fig. 8 The optimized molecular structures, HOMO and LUMO of the protonated inhibitor molecules AEM1-H<sup>+</sup> (c) and AEM2-H<sup>+</sup> (d) using B3LYP using functional and 6-31+G(d,p) basis set.

on molecular reactivity, which facilitates reactivity. It is also found that the global hardness value of AEM1 (2.733 eV) inhibitor is greater than of the AEM2 molecule (2.636 eV), therefore, the trend of the protection effectiveness of the investigated inhibitors as follows: AEM2 > AEM1.

The  $\Delta N$  values were derived using eqn (6). It was previously reported that the aptitude of a compound to transference its electrons to or from the iron surface can be indicated by measuring the  $\Delta N$  value. It was found that for  $\Delta N > 0$  case, the electron is transferred from inhibitor to metal and, *vice versa*, if  $\Delta N < 0$ .<sup>81–83,85</sup> Our results (Table 4) show that, the positive values of  $\Delta N$  on all iron planes (100), (110) and (111) are more positive for AEM2 (1.679, 2.879 and 1.640) than AEM1 (1.609, 2.852, 1.568). These results suggest the higher ability of AEM2 to give electrons to the iron surface than AEM1. It is also found that the donation of electrons from the inhibitor to the 110 plan is the

most positive among all cases. Sastri and Perumareddi<sup>84</sup> was calculated  $\Delta\psi$  using eqn (7).  $\Delta\psi$  is also considered as an important property it is found that the trend in  $\Delta\psi$  follow the order AEM2 (1.379 eV) > AEM1 (1.282 eV) (Table 4). This indicates that AEM2 molecule has an inhibition efficiency greater than AEM1 molecule, which also agrees with the experimental results.

The  $\Delta E_{\text{back-donation}}$  ( $\Delta E_{\text{b-d}}$ ) parameter was calculated using eqn (8).  $\Delta E_{\text{b-d}}$  suggests that back-donation or retro-donation progression *i.e.* electron transfer from metal surface to inhibitor (AEMs) molecules.<sup>85</sup> Eqn (8) implies that when the hardness ( $\eta$ ) of the inhibitor is higher than zero or  $\Delta E_{\text{b-d}}$  is less than zero. The process is energetically favored. Our results displayed in Table 4 indicate that  $\Delta E_{\text{b-d}} < 0$ , therefore, the charge transfer to an inhibitor followed by back-donation from the inhibitor is dynamically favorable. The calculated  $\Delta E_{\text{b-d}}$  values follow the order AEM2 (−0.659 eV) > AEM1 (−0.683 eV), which also correlates with the experimental results.

Finally, the computed dipole moment of the investigated species in all proposed cases are also gathered in Table 4. Xianghong, *et al.*<sup>86</sup> reported that the high value of  $\mu$  possibly increments the interactions between adsorbate and adsorbent. Results in Table 4 reveal that the dipole moments of the protonated AEM2 inhibitor in aqueous medium is 1.13 Debye and it is higher than that of the AEM1 molecule. Therefore, the inhibition efficiency trend may follow the order AEM2 > AEM1. Another important quantum index is the proton affinity (PA).<sup>87</sup> Results presented in Table 4 showed that the proton affinity of AEM2-H<sup>+</sup> (Fig. 6b) is higher than that of the AEM1-H<sup>+</sup> (Fig. 6a) which establishes that effectiveness of the AEMs obeyed the order: AEM2-H<sup>+</sup> > AEM1-H<sup>+</sup>. This finding is in good agreement with the reported experimental data. Results of *N*-protonated species in

Table 4 Quantum chemical parameters for the neutral and protonated forms of AEM1 and AEM2 in both gas phase and aqueous solution, calculated using B3LYP/6-311++G(d,p)

	Neutral form				Protonated forms (opened epoxy rings)				Protonated forms (nitrogen atom)			
	Gas phase		Aqueous solution		Gas phase		Aqueous solution		Gas phase		Aqueous solution	
	AEM1	AEM2	AEM1	AEM2	AEM1	AEM2	AEM1	AEM2	AEM1-H <sup>+</sup>	AEM2-H <sup>+</sup>	AEM1-H <sup>+</sup>	AEM2-H <sup>+</sup>
<i>E</i>	−938.57	−1261.55	−938.59	−1261.56	−1167.99	−1490.95	−1168.01	−1490.99	−938.95	−1261.93	−939.03	−1262.01
$\mu$	2.81	4.18	4.27	5.31	5.47	4.08	5.19	6.32	4.60	6.19	5.85	8.69
<i>E</i> <sub>HOMO</sub>	−5.76	−6.156	−5.872	−6.054	−5.958	−5.329	−5.989	−5.867	−10.324	−10.048	−7.054	−7.232
<i>E</i> <sub>LUMO</sub>	−0.474	−1.047	−0.555	−0.943	−0.401	−0.345	−0.524	−0.596	−4.617	−4.625	−1.319	−1.595
$\Delta E$	5.286	5.108	5.318	5.111	5.557	4.984	5.465	5.271	5.707	5.424	5.735	5.637
<i>I</i>	5.76	6.156	5.872	6.054	5.958	5.329	5.989	5.867	10.324	10.048	7.054	7.232
<i>A</i>	0.474	1.047	0.555	0.943	0.401	0.345	0.524	0.596	4.617	4.625	1.319	1.595
$\chi$	3.117	3.601	3.213	3.498	3.179	2.837	3.257	3.232	7.470	7.337	4.186	4.414
$\eta$	2.643	2.554	2.659	2.556	2.778	2.492	2.733	2.636	2.854	2.712	2.868	2.818
$\sigma$	0.378	0.392	0.376	0.391	0.36	0.401	0.366	0.379	0.350	0.369	0.349	0.355
$\Delta N_{100}$	0.15	0.06	0.131	0.081	1.572	1.767	1.609	1.679	−0.624	−0.632	−0.048	−0.089
$\Delta N_{110}$	0.322	0.239	0.302	0.259	2.836	2.901	2.852	2.879	−0.464	−0.464	0.111	0.072
$\Delta N_{111}$	0.144	0.055	0.125	0.075	1.53	1.73	1.568	1.64	−0.629	−0.637	−0.053	−0.095
$\Delta\psi$	2.852	2.262	2.697	2.399	2.628	3.477	2.563	2.693	0.039	0.021	1.380	1.186
$\Delta E_{\text{b-d}}$	−0.661	−0.639	−0.665	−0.639	−0.695	−0.623	−0.683	−0.659	−0.713	−0.678	−0.717	−0.705
<sup>a</sup> PA	202.1	205.4	−40.6	−39.8								

<sup>a</sup> PA =  $E(\text{protonated}) - (E(\text{neutral}) + E(\text{H}^+))$  (in kcal mol<sup>−1</sup>);  $E(\text{H}^+)$  in gas phase = 1.48 kcal mol<sup>−1</sup> and in aqueous solution = −275.12 kcal mol<sup>−1</sup>.



aqueous solution have the same inhibition efficiency order as in the case of the neutral and the opened epoxy rings.

**3.5.2. Molecular dynamics simulation.** The nature of metal–AEMs interactions and orientations of AEMs on metallic surface was further studied using MD simulation method. The equilibrium configurations (top and side view) of the inhibitors in neutral and protonated forms adsorbed on the Fe(110) surface are shown in Fig. 9. It can be seen that the neutral inhibitors adsorb on the metal surface through the nitrogen (N), oxygen (O), sulfur (S) atoms as well as the aromatic rings. The adsorption energy ( $E_{\text{ads}}$ ) for AEMs adsorption can be calculated as follow (18):

$$E_{\text{ads}} = E_{\text{total}} - (E_{\text{surf+water}} + E_{\text{inh+water}}) + E_{\text{water}} \quad (18)$$

where,  $E_{\text{total}}$  is denote the total energy related to metal–AEMs interactions, which include iron crystal, the adsorbed inhibitor molecule and solution;  $E_{\text{surf+water}}$  and  $E_{\text{inh+water}}$  are the potential energies of the system without the inhibitor and the system without the iron crystal, respectively;  $E_{\text{water}}$  is the potential energy of the water molecules. The obtained  $E_{\text{ads}}$  values are  $-634.0$ ,  $-263.1$ ,  $-689.2$ , and  $-346.5$   $\text{kJ mol}^{-1}$  for AEM1, AEM1- $\text{H}^+$ , AEM2, and AEM2- $\text{H}^+$ , respectively. Negative values of  $E_{\text{ads}}$  indicate the spontaneous behaviour of metal–AEMs interactions. The more negative value of  $E_{\text{ads}}$  for AEM2 as compared to the AEM1 indicates that later case has relatively strong probability of metal–inhibitor interactions as compared to former one.<sup>88</sup> It can also be seen that  $E_{\text{ads}}$  values are higher for neutral form of the AEMs as compared to the protonated form of the AEMs which implies that interactions between AEMs and metallic surface mostly followed the chemisorption mechanism. Effect of the temperature for AEM2 has also been demonstrated using MD simulations and result is shown in Fig. 10. The calculated values of  $E_{\text{ads}}$  are  $-689.2$ ,  $-665.4$ ,  $-636.7$ , and  $-611.9$   $\text{kJ mol}^{-1}$  at 298 K, 308 K, 318 K and 328 K,

respectively. From the results it can be observe that increase in the temperature causes significant decrease in the values of  $E_{\text{ads}}$  that can be resulted due to increase in the kinetic energy of the inhibitor molecules at elevated temperatures.

### 3.6. Mechanism of corrosion inhibition

On the basis of combined computational modeling and experimental results a simple inhibition mechanism can be proposed for carbon steel in hydrochloric acid medium. AEMs molecules in HCl can protonate and exit in their protonated forms as demonstrated as follows:

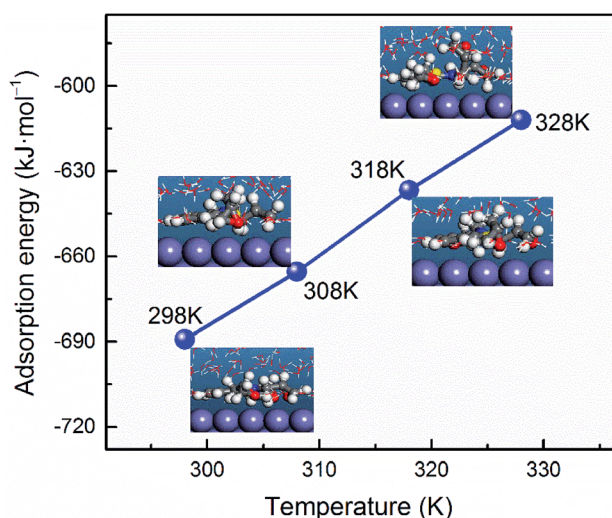
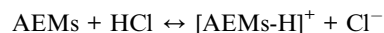


Fig. 10 Dependence of adsorption energy on the temperature for adsorbed AEM2 molecule on Fe(110).

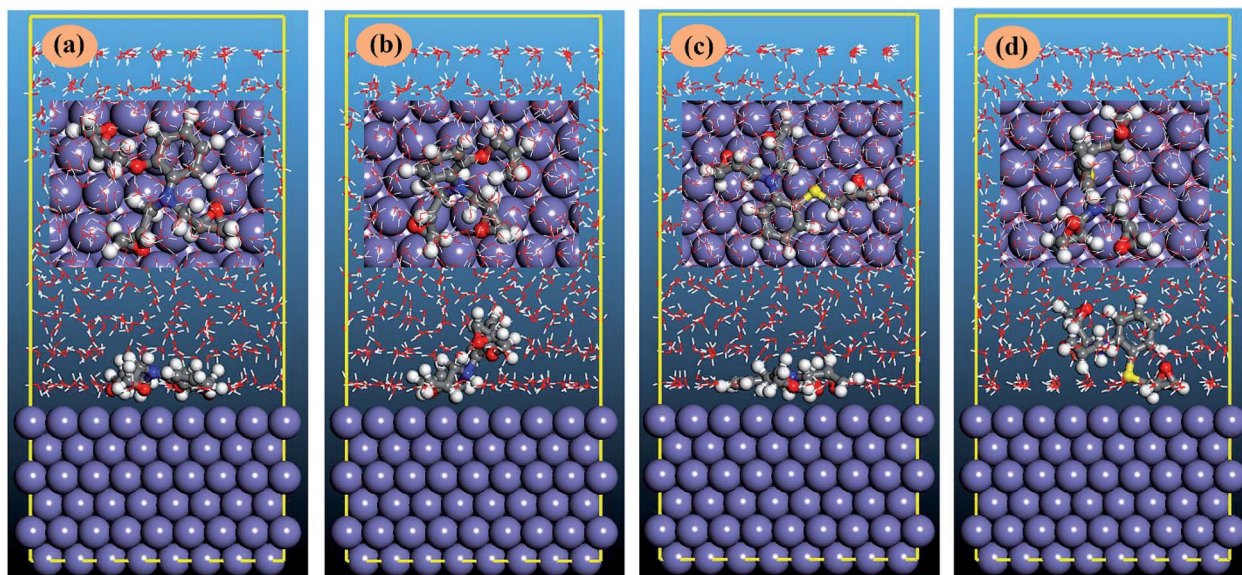


Fig. 9 The equilibrium configurations of (a) AEM1, (b) AEM1- $\text{H}^+$ , (c) AEM2, (d) AEM2- $\text{H}^+$  inhibitors adsorbed on Fe(110) surface.



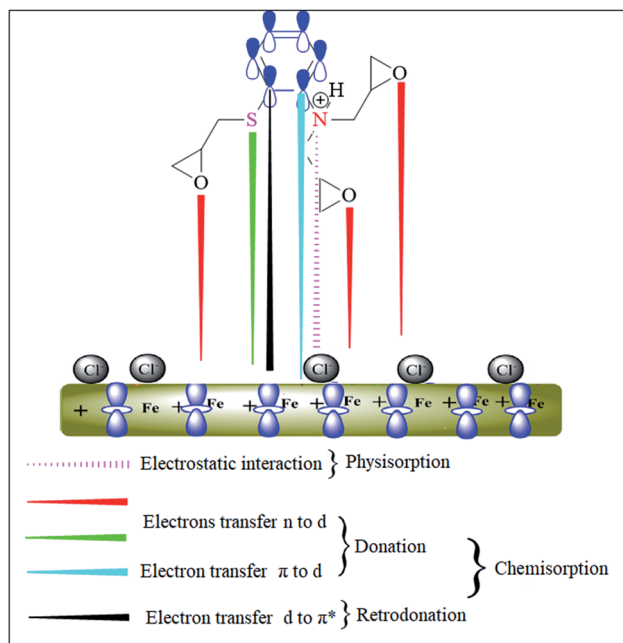


Fig. 11 Pictorial representation of adsorption of AEM2 on carbon steel surface in 1 M HCl solution.

Metallic surface becomes negatively charged due to adsorption of chloride ions that ultimately attracts the positively charged inhibitor molecules through electrostatic force of attractions.<sup>89,90</sup> However, AEMs can interact chemically through several heteroatoms such as N, O, S atoms and aromatic rings that offer strong bonding with the metallic surface. A possible mechanism of the adsorption of AEM on the steel surface is shown in Fig. 11.

## 4. Conclusions

From above experimental and theoretical studies, it can be concluded that:

(i) AEM1 and AEM2 acted as good inhibitors and their inhibition effectiveness increases with increasing their concentration.

(ii) The PDP results reveal that, the AEMs behave as mixed-type inhibitors with anodic predominance.

(iii) The EIS results suggested that, AEMs molecules protect the carbon steel from corrosion by the formation of a protective film at the metal–electrolyte interface.

(iv) Adsorption of the tested AEMs molecules obeyed Langmuir adsorption isotherm.

(v) Values of  $\Delta G_{\text{ads}}$  suggested that AEM1 and AEM2 mainly interact through chemisorption mechanism.

(vi) The chemisorption mechanism was further supported by MD simulations study.

(vii) Theoretical results derived for neutral and protonated forms of the AEMs in both gas phase and aqueous solution indicated that AEM2 acted as is more efficient inhibitor than AEM1.

(viii) The more negative value of  $E_{\text{ads}}$  for AEM2 as compared to the AEM1 indicated that AEM2 acted as is more efficient inhibitor than AEM1.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

C. Verma thankfully acknowledge the North-West University (Mafikeng Campus) South Africa for providing financial assistance under postdoctoral fellowship scheme. Nuha Wazzan gratefully acknowledges King Abdulaziz University's High Performance Computing Centre (Aziz Supercomputer) for assisting with the calculations for this study. This research was partially supported by the National Natural Science Foundation of China (21706195), the Science and Technology Program of Guizhou Province (QKHJC2016-1149), and the Guizhou Provincial Department of Education Foundation (QJHKYZ2016-105). We would like to thank the Laboratory of Metallurgical Analysis, Cetim Maroc Développement (Centre Technique des Industries de la Mécanique) and Quality Control Laboratory, Casablanca Aeronautics Group Figeac Aero. Aeronautical Technopole of Nouaceur, Mohammed V-Casablanca Airport, Morocco.

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