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How do you (dis)solve a problem like methylene chloride?†

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Chemical regulation in the European Union (EU) and the United States of America (USA) has restricted the use of some historically important solvents, essentially banning certain uses. The most high profile regulatory action thus far has been a 'Final Risk Management Rule' prohibiting all consumer uses, and many commercial uses, of methylene chloride (dichloromethane, DCM) by the United States Environmental Protection Agency (EPA) through the Toxic Substances Control Act (TSCA). The unique properties of chlorinated solvents makes direct substitution difficult or impossible for most uses, and creative solutions are needed. The replacement of methylene chloride in synthesis, extraction, and chromatography with green solvents will be discussed as a way of using regulatory intervention as the catalyst for innovation and positive change.

Sustainability spotlight

New restrictions on how methylene chloride can be used means that there is a need for alternative solvents, and a danger that the replacements introduce new risks and unforeseen consequences. This work establishes how and where methylene chloride is being used at present, and what solvents offer a viable alternative. Solvent substitution in favour of less hazardous solvents contributes to the aspirations of Good Health and Well-being (SDG3) and Climate Action (SDG13). Chemical processes that utilise safer solvents to comply with regulation promote Industry, Innovation and Infrastructure (SDG9) and permit Responsible Consumption and Production (SDG12).

Introduction

Methylene chloride, also known as dichloromethane (DCM), is one of the most useful and popular organic solvents. Methylene chloride was discovered in 1840 by Henri Victor Regnault when they chlorinated methyl chloride. Commercial production of methylene chloride was first performed by the reaction of chloroform with zinc and hydrochloric acid.¹ This was replaced by the chlorination of methane at 500 °C which results in a mixture of chloromethanes (methyl chloride, methylene chloride, chloroform, carbon tetrachloride). As methylene chloride gained popularity in the mid twentieth century, earlier manufacturing techniques were surpassed by the contemporary method of producing methyl chloride from methanol and hydrogen chloride, which is subsequently chlorinated (Fig. 1).² The main uses of methylene chloride which established it as a mainstay solvent were as a paint remover, for decaffeination, the production of polycarbonates, and as a blowing agent for polyurethane and the subsequent cleaning of machinery.² Over time, methylene chloride was replaced in some of these

applications due to health concerns, a trend culminating in recent regulatory action.

We can gain some insight into modern methylene chloride use in the USA from the Environmental Protection Agency (EPA) Toxics Release Inventory (data for 2020 to 2023 is shown in Fig. 2).³ The data up to 2022 shows emissions of methylene chloride topping 1 million litres each year. The 2023 data set records 800 000 litres of methylene chloride emissions, a notable drop, which could indicate the start of a concerted effort to reduce methylene chloride use. The industrial sector responsible for the most methylene chloride emissions is



Fig. 1 (a) Four stage synthesis of methylene chloride from methane via methanol. (b) COSMO surface of methylene chloride indicating positive charge in blue, neutral charge in green and negative charge in yellow.

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Fig. 2 Toxics release inventory data for the annual quantities of environmental releases of methylene chloride (USA, 2020–2023). Left (a): separated by industry. Right (b): separated by fate, where 'other' includes sewage treatment, storage, and untracked emissions.

chemical manufacturing. Cleaning machinery is also a source of methylene chloride emissions. Some emissions are attributed to hazardous waste processors (*i.e.* off site emissions). Each year, the vast majority of methylene chloride emissions are to the air, either as intentional emissions or as fugitive emissions.

The reasons why chlorinated solvents are highly valued have been covered elsewhere.⁴ In brief, methylene chloride has the benefits of halogen bonding,⁵ providing strong interactions with typical dipolar substrates for high solubility, while at the same time having a low boiling point (40 °C) for easy removal. This combination of properties is extremely convenient for laboratory and manufacturing uses, but as explained next, comes at the cost of human health and environmental damage.

Hazards and regulatory action

Since April 2024, the United States Environmental Protection Agency (EPA) has issued an intervention *via* the Toxic Substances Control Act (TSCA) to reduce the risk to human health posed by methylene chloride.⁶ It was found that methylene chloride poses an unacceptable risk of cancer,⁷ acute neurotoxicity, and chronic liver toxicity due to exposure arising from many of its uses.⁸ The final risk management rule affects methylene chloride in the USA in the following ways.⁹ All consumer uses of methylene chloride have been banned since May 2025. The following uses of methylene chloride are permitted within industrial and commercial operations, as well as research, government, and academic institutions, but a 'Workplace Chemical Protection Program' (WCPP) is required: manufacturing of methylene chloride, import of methylene chloride, uses of methylene chloride as a reactant, incorporation into a mixture (*e.g.* formulation), repackaging (*e.g.* re-bottling) of methylene chloride containing products, recycling of methylene chloride, its use as a laboratory chemical (*i.e.* for research and teaching purposes), as a paint or coating remover for aircraft and spacecraft, as a bonding agent for solvent welding, as a 'processing aid' – the example given in the legislation is of a heat transfer fluid but it can also be understood to

include solvent uses,⁹ in plastics and rubber manufacturing, in closed loop processes where the methylene chloride is reclaimed, and methylene chloride disposal (waste management). A WCPP consists of monitoring the airborne methylene chloride concentration and compliance with exposure limits, as well as requirements for protective equipment and training and an exposure control plan.¹⁰ Furthermore, the 8 hours time-weighted average exposure limit has been lowered from 25 to 2 ppm, and the short-term exposure limit has been lowered from 125 to 16 ppm.¹¹ Commercial uses of methylene chloride not previously listed or able to comply to a WCPP are banned after April 2026. There are some important exceptions where the TSCA ruling does not apply, including the pharmaceutical industry (accounting for about a third of methylene chloride use) and food processing (*e.g.* decaffeination) which are covered by other legislation.

The number of complete TSCA final risk management rules is now five, also covering carbon tetrachloride, perchloroethylene, trichloroethylene, and asbestos.¹² It is not a coincidence that four of the five finalised risk management rules target hazardous chlorinated solvents. Risk management rules are in the pipeline for a variety of other solvents, including 1,2-dichloroethane, which is on the REACH authorisation list in Europe.¹³ Methylene chloride is also subject to EU restrictions, specifically forbidding its use in paint removers.¹⁴

The health hazards, environmental impact, and safety (physical) hazards of solvents can be ranked using the Safe and Sustainable by Design (SSbD) approach pioneered in the EU (Fig. 3).¹⁵ Compared to other chlorinated solvents, the hazards of methylene chloride are less severe, which contributed to its status as a safer option compared to chloroform or 1,2-dichloroethane for example in early solvent selection guides.¹⁶ This option is no longer tenable for applications not permitted under the recent TSCA risk evaluation of course. Methylene chloride is suspected of causing cancer, as designated with the hazard statement H351. Methylene chloride is also an irritant. 1,2-Dichloroethane and trichloroethylene definitively cause cancer (H350) and are regarded as more hazardous in this respect. The





Fig. 3 Radar charts: hazard severity levels in health, environment and physical hazard categories as defined by the EU safe and sustainable by design guidelines. Hazard severity is categorised in four levels, from inside to outside on each radar chart: no applicable hazards, minor hazards, concerning hazards, most severe hazards. Blocks: underneath each radar plot are the greenness scores from the CHEM21 solvent selection guide, in order of safety, health, and environment. Hazard codes and further information are given in the ESI.†

chronic organ damage (H372) of chloroform and carbon tetrachloride is also highly concerning and warrants a greater health risk score than methylene chloride in the SSbD framework. Regarding physical hazards, methylene chloride is not flammable and generally stable, and has no formal hazard statements regarding environmental toxicity. Fig. 3 also includes potential replacements for methylene chloride, all of which have lower health risks but tend to be flammable or potentially explosive in the case of ethers. The CHEM21 solvent selection guide ranks ethers poorly due to these safety concerns.¹⁷ The transition from non-flammable chlorinated solvents to flammable hydrocarbon and oxygenated solvents will require adjustments to safety procedures.

Despite the inherent hazards of methylene chloride being less severe than other chlorinated solvents, ultimately the risk of using methylene chloride also depends on exposure. Chronic toxicity (*e.g.* cancer) requires repeated or prolonged exposure to become a relevant mode of action. That said, there is a lot of evidence for the acute toxicity of methylene chloride, including death, that ultimately resulted in the regulatory action being taken. Since 1980, 85 deaths have occurred directly due to the use of methylene chloride in the USA.¹⁸ Seventy four of those

fatalities occurred at work, with the use of paint strippers in enclosed spaces the main culprit.¹⁹ Exposure to methylene chloride vapor causes heart attacks, which can occur after only brief accidental exposure, for example from leaks.²⁰ There have also been reports of necrosis after accidental injection of methylene chloride in laboratory accidents.²¹ Less typically, physical risks can occur from the use of methylene chloride. The combination of methylene chloride and sodium azide can form the explosive diazidomethane.²² It is generally best practice to avoid using halogenated solvents with azides for this reason.

Chlorinated solvents are known to deplete stratospheric ozone, the most notorious of which is carbon tetrachloride, earning itself a place amongst the substances banned under the conditions of the Montreal Protocol.²³ Methylene chloride also depletes ozone,²⁴ and worldwide emissions have been recorded as increasing,²⁵ posing a greater threat to the environment. Therefore, evidence suggests methylene chloride poses greater risks to people and the planet than might be expected from hazard statements alone.

Uses of methylene chloride

Formulation

It is perhaps paint remover products that have been the most high profile use of methylene chloride and the main driver behind legislative action. Use of paint stripper products containing methylene chloride was banned in the EU in 2012,²⁶ and banned for consumers in the USA in 2019 (prior to the more recent ruling introducing wider restrictions on methylene chloride use).²⁷

Alternative solvents have been suggested to reformulate paint removers. A mixture of methanol and DMSO has been suggested, primarily on the basis of environmental impact rather than technical performance.²⁸ Milescu *et al.* tested dihydrolevoglucosenone (brand name Cyrene) for its ability to remove paint from porous substrates such as brick with some success compared to other neat solvents.²⁹ However, neither of these solutions is a formulated product. Contemporary, commercially available paint removal products contain a variety of ingredients, often alkanes, alkoxy alcohols, and a base (hydroxide or amine, see Table S2 of the ESI[†]). These products are corrosive but do not present a chronic toxicity hazard because of the absence of methylene chloride.

Synthetic chemistry

Although there is concern over the future of methylene chloride in industrial chemistry, in those research labs shielded from the brunt of regulations, the use of methylene chloride has not wavered. A survey of some representative journals shows that reported reactions in methylene chloride are just as popular as they were in 2020 (Fig. 4). *ACS Sustainable Chemistry and Engineering* has the lowest instances of methylene chloride being used as a solvent at just 1 reported reaction in every 10 original research articles. *RSC Sustainability* in 2023 also had an average of 0.1 reactions using methylene chloride as a solvent per





Fig. 4 (a) Scifinder reports of reactions using methylene chloride as a solvent divided by the number of research articles published that year from 2020 to 2023 in seven journals. (b) Solvent occurrences in the 79 articles published in OPRD in the first quarter of 2025. Multiple uses in a single article are counted as one entry. Percentage frequency given for methylene chloride (dichloromethane, DCM) and all more commonly used solvents for comparison.

article. *Green Chemistry* reports a reaction in methylene chloride an average of 1–2 times per article. Journals more likely to contain substrate screenings of a synthetic methodology or catalogues of novel compounds have more instances of methylene chloride use per article, and likely much more is used in extractions and chromatography. A prime example is the *Journal of Medicinal Chemistry*, featuring an average of 12 reactions in methylene chloride per article.

Taking a more detailed look at articles published in *Organic Process Research and Development* (OPRD) between January and March 2025, which regularly covers the kilogram synthesis of fine chemicals, the choice of solvents for reactions, work up (extraction or trituration or solvent swap) and recrystallisation, and column chromatography was surveyed. Methylene chloride accounts for 10–15% of reported solvent use and it was the third most common choice of solvent in each category (if alkanes are grouped together). Methanol and tetrahydrofuran (THF) are the most frequently used options for reaction solvents (closely followed by methylene chloride), while ethyl acetate and alkanes (with *n*-heptane the most popular) are the most common post-reaction solvents (again, methylene chloride is third). A previous survey of OPRD covering the years 1997 to 2012 found between 17% and 64% of the papers examined featured methylene chloride with a mean annual occurrence of 38%.³⁰ The equivalent figure for 2025 (first quarter) is 32%. Overall, the use of methylene chloride appears to fluctuate without a strong trend but it is always significant. We can infer the popularity of methylene chloride in small scale exploratory reactions is still transferring onto larger scale and pilot plant chemistry.

Chromatography in safer solvents

When utilised, flash chromatography is the greatest source of waste in a chemical process, and the quantity of solvent required is the main contributor. Ethyl acetate combined with alkanes is a common option, but methanol–methylene chloride is considered superior for the separation of polar substances. A visual guide for greener chromatography, showing proportions of alternative solvents has been designed by Taygerly *et al.* (Fig. 5).³¹ Heptane or methyl *tert*-butyl ether (MTBE) modified with 3 : 1 ethyl acetate–ethanol is a broadly useful alternative. Note that MTBE scores poorly in the CHEM21 solvent selection guide due to its flammability and resistivity.¹⁷ This presents a new challenge because methylene chloride is non-flammable and so risk assessments must account for new safety hazards. All the recommended solvent systems contain flammable solvents and pose some health hazards. The practitioner must be aware of these hazards, especially because purification solvents are required in greater quantities than reaction solvents (and must not be overlooked in risk assessments). Another seminal study by MacMilan *et al.* concluded that cyclopentyl methyl ether (CPME)–methanol has the potential to replace methylene chloride–methanol in chromatography. To avoid toxic methanol, an isopropanol (IPA)–ethyl acetate eluent can also be used.³² Heptane in combination with methyl acetate or ethyl acetate has also been supported for the separation of analgesics.³³

The chromatographic separation of acetaminophen and caffeine is a common system to study new solvents. Interpreting thin layer chromatography (TLC) retardation factors (R_f) with a statistical thermodynamic approach to intermolecular interactions has suggested methyl pivalate, *tert*-butyl acetate, *sec*-butyl acetate and ethyl isobutyrate are suitable





Fig. 5 A chart for substituting mixtures of methylene chloride (DCM) and methanol (MeOH) for chromatography with alternative solvent systems including ethyl acetate (EtOAc), ethanol (EtOH), 2-propanol (IPA), and methyl *tert*-butyl ether (MTBE). Trace vertically from a DCM-MeOH system to an appropriate alternative. Hazard pictograms included for solvents and solvent mixtures.

replacements for methylene chloride in combination with methanol.³⁴ Similarly, Sharma *et al.* used TLC to identify methyl acetate and ethyl acetate as replacements for methylene chloride.³⁵ Instead of adapting conventional preparatory chromatography, alternative purification techniques can be adapted. Peterson *et al.* suggest investigating ion exchange resins, reverse-phase chromatography, or supercritical fluid chromatography.³⁶

Replacing methylene chloride in extractions

Aside from column chromatography, the greatest use of methylene chloride within a synthetic process will often be the work up. This is where water soluble by-products and other contamination is separated from a solution of the primary product. Other types of extractions are also used in the analysis of environmental samples, restoration, and food manufacturing (*e.g.* oil extraction or decaffeination). Ethyl acetate is also a common extraction solvent and sometimes interchangeable with methylene chloride. For instance, lake sediment extraction for environmental monitoring purposes has been achieved with ethyl acetate, albeit with some loss of recovery, particularly of fatty acids, compared to methylene chloride.³⁷ There are more reasons why methylene chloride might be preferred over ethyl acetate or other alternatives. The solubility of water in methylene chloride is an order of magnitude lower than water in ethyl acetate,³⁸ limiting cross contamination and the possibility of emulsions forming. The high density of methylene chloride is advantageous when removing it from water by separating

funnel and this is particularly true at larger manufacturing scales. Ethyl acetate is also flammable, whereas methylene chloride is not. Methylene chloride has a lower boiling point for easier removal too, although this can translate to more fugitive emissions.

Ideally, the need for extraction solvents should be completely eliminated. In one innovation, glass powder coated with an absorbent silicone polymer is added to water to remove organic compounds, which then can be filtered and added as the input for column chromatography.³⁹ Obviously the high solvent demand of the column is not solved, but significant solvent savings were reported all the same. Similarly, a polyurethane foam sponge does the same job in place of a traditional solvent extraction.⁴⁰

Where a solvent is still required, Cayot *et al.* used an acetone-cyclopentane or ethyl acetate-ethanol mixture for food extractions in the food processing industry.⁴¹ Caffeine extraction is the most well known example of where methylene chloride has traditionally been used in the food industry, for which several alternative solvents are known, including ethyl lactate,⁴² and supercritical carbon dioxide.⁴³ Note there are regulations in the USA and elsewhere controlling the use of solvents for food applications. Rather than banning solvents, regulation for the food sector tends to describe which solvents you can use from restrictive lists. Methylene chloride is in fact permitted for the decaffeination of coffee if the solvent residue is present in less than 10 ppm as described in USA regulation,⁴⁴ or 2 ppm in the equivalent EU law.⁴⁵ This approach to regulation does not encourage innovation, because novel, safer solvents suitable for food extraction will not be permitted without revisions to the law.



Methylene chloride-free synthesis

Much of the published solvent substitution activities are focused on organic synthesis, especially within the pharmaceutical industry.⁴⁶ The EPA risk management rule for methylene chloride does not extend to the pharmaceutical sector, which is outside the jurisdiction of TSCA, and is an industry that routinely handles very hazardous substances with high safety standards (thereby minimising risk to acceptable levels). Nevertheless, the impressive efforts made by the pharmaceutical industry to reduce methylene chloride use shows regulation is not always required as a driver for greener practices.^{47–49}

Beyond the pharmaceutical sector, it is valuable to know what uses of methylene chloride will be impacted most by regulation. Of the 1 627 689 uses of methylene chloride recorded in the SciFinder database between 2020–2025 (as of March 2025), 8 reaction categories stand out as the most important (Fig. 6). Almost a quarter of the recently recorded uses of methylene chloride can be attributed to amide formation, usually between a carboxylic acid and an amine using a coupling agent (distribution of methylene chloride reactions is shown on the x-axis of Fig. 7). The equivalent process of making esters with alcohols is the third most prevalent reaction type. The hydrolysis of amides (ranked 2nd) and esters (8th) are also common reactions, specifically the cleavage of the *tert*-butyloxycarbonyl (boc) protecting group and *tert*-butyl esters respectively with trifluoroacetic acid (TFA). The prominence of amide formation and boc deprotection as methylene chloride uses are indicative of the essential role of this



Fig. 7 Reaction survey results represented by the prevalence of each reaction (from 1.6 million hits where methylene chloride (DCM) is the reaction solvent) on the x-axis plotted against how often DCM is used in the generic reactions of the accompanying schemes in Fig. 6 (y-axis), all according to SciFinder for the period January 2020 to March 2025.

chemistry in contemporary bioactive molecule research, where large numbers of small molecule amides and peptides are produced for activity screening.

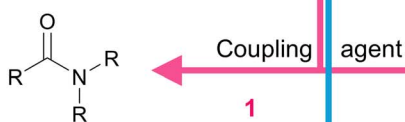
Ester hydrolysis



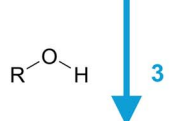
Amide/carbamate hydrolysis



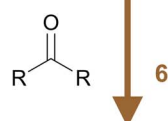
Ether hydrolysis



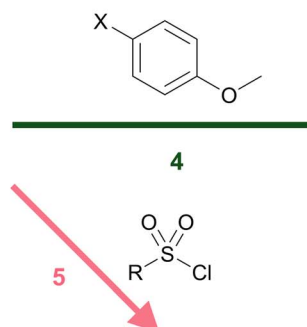
Amidation



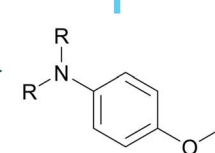
Esterification



Reductive amination



Sulphonamidation



Aromatic nucleophilic substitution

Fig. 6 Common reactions performed in methylene chloride (numbered 1 to 8). Numbers are the popularity ranking of each reaction according to SciFinder for the period between January 2020 and March 2025. The schemes represent generic versions of the most common transformations found within each category.



Fourth on the list of common reactions is nucleophilic substitution by amines, consisting of aromatic nucleophilic substitution (40 000 examples) and nucleophilic substitution (21 000 examples). The other top uses of methylene chloride as a reaction solvent are the production of sulphonamides, reductive amination, and ether cleavage. Nucleophilic addition to a carbonyl accounts for almost half the reported uses of methylene chloride as a reaction solvent. Carbonyl addition is very solvent dependent, the kinetics of which are inversely proportional to the hydrogen bond accepting ability of the solvent.⁵⁰ This makes the poorly hydrogen bonding chlorinated solvents and aromatic solvents very appealing despite their hazards.

Fig. 7 (y -axis) also shows the dependency of the generic reaction types on methylene chloride (*i.e.* the percentage of reactions in each category performed in methylene chloride). This is based on the reaction schemes in Fig. 6, which represent the most common example in each reaction category. It is important to consider this, because it demonstrates that although there are many more examples of methylene chloride being used in amidations than other reactions, that reaction category is so large that actually only a minority of direct amidations using coupling agents require methylene chloride. This could indicate that the use of methylene chloride in amidations is easily eliminated by following other common protocols in other solvents. Four of the 8 reactions are predominantly performed in methylene chloride: direct esterification by coupling agent, sulphomamidations with sulphonyl chlorides, conversion of aryl methyl ethers to phenols, and TFA hydrolysis of *tert*-butyl esters. Aromatic nucleophilic substitution on the other hand is rarely practiced in methylene chloride and accounts for a modest amount of total methylene chloride use in synthesis. Notable solvent substitutions for the major reactions reliant on methylene chloride will now be summarised.

Amidation

Direct amide formation is necessary in a variety of fields, notably for synthetic peptides and other bioactive compounds.⁵¹ The reaction between a carboxylic acid and an amine can be achieved with a coupling agent. Analysing hexafluorophosphate azabenzotriazole tetramethyl uronium (HATU) mediated transformations,⁵² DMF is the most common solvent (74% of SciFinder examples since 2020), with methylene chloride accounting for the solvent choice in 18% of examples. However, because amide coupling is so prevalent in synthetic chemistry, it accounts for nearly a quarter of the reactions using methylene chloride in the SciFinder database and thus is worthy of attention.

The solvent Cyrene has been used successfully in HATU mediated coupling reactions to make small molecule amides including dipeptides.⁵³ Cyrene can also be used as a solvent in the reaction of acid chlorides and amines to make amides.⁵⁴ Both reports make specific mention of washing the Cyrene out of the crude product mixture with water.^{53,54} Unlike methylene chloride, Cyrene is miscible with water, but needs considerable washing to completely remove.

Solid phase synthesis is a vital technology for amidation but reliant on solvents that swell the resin support. Propylene

carbonate can be used as the solvent for solid phase peptide synthesis, as was demonstrated with the synthesis of non-peptide bradykinin.⁵⁵

Acetonitrile was demonstrated to give superior conversion and selectivity compared to DMF in the coupling of hindered carboxylic acids with weakly nucleophilic amines using *N*-acyl imidazolium intermediates in combination with tetramethylchloroformamidinium hexafluorophosphate (TCFH).⁵⁶

A large screening of acid-amine pairs and coupling agents has suggested that (1-cyano-2-ethoxy-2-oxoethylideneaminoxy) dimethylamino-morpholino-carbenium hexafluorophosphate (COMU) is actually a preferable coupling agent to HATU,⁵⁷ which works well in combination with dimethyl carbonate, ethyl acetate, and 2-methyltetrahydrofuran.⁵⁸ Another methodology uses a dithiocarbonate to replace typical coupling agents, which can be performed in ethyl acetate.⁵⁹ Other additions to the literature include the use of aqueous surfactants,⁶⁰ and deep eutectic solvents.⁶¹ There are also instances of solvent-free amidation: to this end ball milling of esters and amines is known,⁶² as is the iodine mediated Boc protection of amines with Boc₂O.⁶³

Due to the toxicity of amide coupling agents,⁶⁴ catalytic amidation methods are desirable, but have not superseded the established coupling agent protocols. These catalytic methods can be biocatalytic,⁶⁵ which are possible in a variety of solvents, or chemocatalytic,⁶⁶ which tend to rely on toluene and tetrahydrofuran (THF). Despite scoring slightly better in solvent selection guides than methylene chloride,¹⁷ toluene and THF both have chronic toxicity hazards.

Amide (and carbamate) hydrolysis

The deprotection of *tert*-butyloxycarbonyl (boc)-protected amines with trifluoroacetic acid is a major category of amide hydrolysis. 44% of recent examples are conducted in methylene chloride, and the only other popular solvent option is 1,4-dioxane (23%), which is known to cause cancer. Furthermore, trifluoroacetic acid is under scrutiny as a persistent environmental pollutant.⁶⁷ Therefore, safer solvents/reagents for amide hydrolysis can be considered a priority action area despite some other reactions having a stronger dependence on methylene chloride. Fortunately there are several protocols for the thermally-initiated removal of Boc groups, reported to occur in flow reactors in acetonitrile,⁶⁸ as well as in trifluoroethanol,⁶⁹ or water.^{70,71} Instead of using trifluoroacetic acid, alternative conditions have been developed, either solvent-free with hydrogen chloride,⁷² or using oxalyl chloride in methanol.⁷³

Esterification

Steglich esterifications between carboxylic acids and alcohols using a carbodiimide coupling agent are usually performed in methylene chloride. Solubility and stability plays an essential role in most reactions, but additionally chlorinated solvents are especially suited to a variety of esterification methods because they accelerate the reaction through entropic effects.^{50,74} Despite this, a solvent selection guide documents several viable alternative solvents.⁷⁵ The favoured reaction system appears to be 2-chloro-1-methylpyridinium iodide (Mukaiyama's reagent) in



dimethyl carbonate. 5-Nitro-4,6-dithiocyanatopyrimidine (NDTP) can also be used as a coupling reagent in excess alcohol (as solvent and reactant), or acetonitrile as the solvent.⁷⁶ 1-Propanephosphonic anhydride (T3P) and cyclopentanone have been found to promote the synthesis of thioesters from thiols and carboxylic acids.⁷⁷

Luis *et al.* have demonstrated the action of TCFH upon acyl imidazolium intermediates in acetonitrile as the solvent to produce esters and thioesters.⁷⁸ The method is mild (0–25 °C, ~3 hours) and applicable to poorly reactive substrates.

Aromatic nucleophilic substitution

The typical aromatic nucleophilic substitution of fluorine with an amine tends to be performed in dimethyl sulphoxide (DMSO, 33%) or *N,N*-dimethylformamide (DMF, 26%) rather than chlorinated solvents. Dipolar aprotic solvents have their own issues, with DMSO prone to decompose and transfer solutes through the skin, and DMF being reprotoxic. Alternatives to dipolar aprotic solvents have been discussed elsewhere.⁷⁹

Sulphonamidation

Two-thirds of recently reported sulphonamidations (from the sulphonyl chloride and an amine) are performed in methylene chloride, and although THF and DMF are also popular solvents they do not represent greener alternatives on account of their respective suspected carcinogenicity and reproductive toxicity. When it concerns the use of (sulphonic or carboxylic) acid chlorides, anhydrous solvents tend to be used to avoid hydrolysis of the reactant. However, it has been shown that a pH of 8 (or lower) is sufficient to suppress hydrolysis and permit facile sulphonamidation in water.⁸⁰ Alternative methods from nitroarenes instead of amines also work with water as the solvent.⁸¹ Conditions utilising deep eutectic mixtures based on choline chloride as solvents,⁸² or no solvent at all,⁸³ are also known.

Reductive amination

The reductive amination of aldehydes has a very strong dependence on chlorinated solvents. Where sodium triacetoxymethylborohydride (STAB) is the reducing agent, 42% of recent reports are in methylene chloride and a further 25% are in 1,2-dichloroethane. The problem of chlorinated solvent use in the reductive amination of aldehydes has been comprehensively addressed by the work of McGonagle *et al.*⁸⁴ Ethyl acetate and dimethyl carbonate were found to be green, high performance alternative solvents that work on a wide range of substrates.

Ether hydrolysis

The use of methylene chloride in ether hydrolysis is tied to a specific methodology for the demethylation of anisoles (methoxybenzenes) with boron trifluoride dating back to the 1960s.⁸⁵ The mechanism requires the bromination of the methyl group. In the intervening years, alternative methods have been developed, such as the use of aluminium triiodide in acetonitrile,⁸⁶ or an oxidative route using 2-iodoxybenzoic acid (IBX) in a mixture of dimethyl carbonate and water.⁸⁷

Biocatalytic methods are also known to proceed in DMSO for guaiacol and its derivatives.⁸⁸

Ester hydrolysis

The hydrolysis of *tert*-butyl esters with TFA has an overwhelming dependence on methylene chloride (81% of recent examples). There are some examples of using water as a solvent for this transformation, including the synthesis of saccharide oligomers,⁸⁹ and radiolabelled medical imaging compounds.⁹⁰

Other notable reactions

Naturally methylene chloride is used in many other reactions outside the top 8 summarised above. Brominations are commonly conducted in chlorinated solvents, but ethyl acetate can be an acceptable substitute.⁹¹ Olefin metathesis is another reaction overwhelmingly reliant on chlorinated solvents, but dimethyl carbonate and ethyl acetate have been shown to be suitable for ring closing metathesis reactions in particular.⁹² A mechanistic study was used to investigate alternative solvents for nitrene transfer reactions, which found that dimethyl carbonate, isopropyl acetate and anisole may be good candidates.⁹³

Conclusions

We are at a transitional point in the history of methylene chloride. Once viewed as a vital and irreplaceable solvent, its commercial uses are soon to be relegated to small scale exploratory chemistry, pharmaceutical manufacturing and a few other permitted uses. The cost-benefit approach favoured by lawmakers has been criticised,⁹⁴ but the latest ruling on methylene chloride is just a continuation of a recent trend of international regulatory action against solvents with chronic toxicity hazards.

Chemists and chemical engineers need to avoid developing products and processes using methylene chloride to avoid being trapped into a reliance on substances subject to, or at risk of, regulatory interventions. The major uses of methylene chloride across formulations, separations, and reactions all have promising alternatives if practitioners are willing to adapt protocols and this transition to safer solvents should start in the teaching laboratory.^{95,96} Ethyl acetate and dimethyl carbonate recur as promising substitutes across different case studies with safer environmental health and safety profiles (see ESI Fig. S1†). To discover the optimum solvent in any application, it is essential to understand the role of the solvent through experiment, but it would be pertinent to first try either ethyl acetate or dimethyl carbonate in place of methylene chloride before undertaking more thorough redesigns of products or processes to accommodate other solvents. Many sectors have enthusiastically taken to the challenge of eliminating methylene chloride,⁹⁷ and this latest risk management rule will make others follow suit.

Data availability

The data supporting this article (*i.e.* literature survey results) have been included as part of the ESI.†



Conflicts of interest

There are no conflicts to declare.

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