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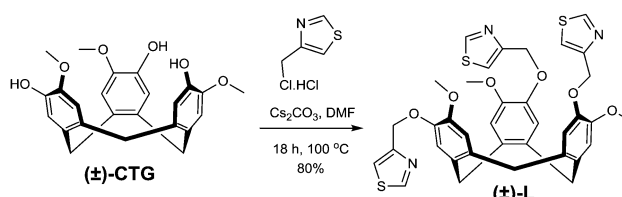
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M₁₂L₈ metallo-supramolecular cube with cyclotriguaiacylene-type ligand: spontaneous resolution of cube and its constituent host ligand†

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The racemic ligand (±)-tris-(4-methylthiazolyl)cyclotriguaiacylene forms a homochiral crystalline Ag₁₂L₈ cube with spontaneous resolution. The ligand itself likewise crystallises in a chirally pure fashion in two clathrate complexes. Ag₁₂L₈ is the first example of a cyclotriguaiacylene-type coordination cube and a rare example of a M₁₂L₈-type metallo-cube.

Metallo-cages, also referred to as coordination cages, are three-dimensional, hollow architectures that often closely resemble Platonic or Archimedean solids. Investigation of their self-assembly, and applications as nano-scale host assemblies is a highly active area.¹ One approach to a cube assembly is linking together eight corner units in an orthogonal manner. Most examples of such edge-directed coordination cage cubes have been M₈L₁₂ cages where the metal supplies the corner piece.² An alternative design is a M₁₂L₈ assembly where the metal represents the cube edge and the corner piece is a ligand, though examples are much rarer.³ Cyclotrimeratrylene (CTV) and its analogues have a bowl-shape tribenzo[*a,d,g*]-cyclononatriene core with the arene faces approximately orthogonal to one another which suggests that CTV-type ligands would be an ideal corner piece for an edge-directed cage. Indeed this concept has been demonstrated by Warmuth who reported an assembly where eight aldehyde-functionalised cyclotriguaiacylene (CTG) analogues were linked into a cube through reversible imine bond formation with twelve diaminobenzene units.⁴ Despite this organic precedent, directly analogous metallo-supramolecular examples with CTV or CTG-type ligands are not known. These ligand types do form an array of metallo-cages,^{5–8} but the closest examples to a cube are M₆L₈ stella octangula structures of octahedral symmetry,⁷ and a topologically complex self-entangled Pd₄L₄ cube.⁸ Likewise building tectons based on a conformationally-locked tribenzotriquinacene-core (akin to a



Scheme 1 Synthetic route to (±)-tris-(4-methylthiazolyl)cyclotriguaiacylene.

bridge-head CTV) can be assembled into cube structures,⁹ but the only 3D metallo-cage examples of these ligands are M₃L₂ assemblies with trigonal bipyramidal shape.¹⁰ We report here the first example of a M₁₂L₈ cube assembled from a CTG-type ligand, noting CTG-ligands are chiral and this cube displays both chiral self-sorting in its assembly and in its crystallisation into two conglomerate clathrate complexes.

The racemic ligand (±)-tris-(4-methylthiazolyl)cyclotriguaiacylene L was synthesised in 80% yield through reaction of 4-(chloromethyl)thiazole hydrochloride with cyclotriguaiacylene in the presence of base, Scheme 1. The ¹H NMR spectrum is characteristic of the C₃-symmetric bowl-shaped cavitanid with the bridging *endo* and *exo* methylene protons of the tribenzo[*a,d,g*]-cyclononatriene core appearing as doublets at 4.71 and 3.49 ppm respectively. The structures of two clathrate complexes of L, namely L·4(H₂O) and L·3(CH₃NO₂), were determined by single crystal X-ray techniques.‡ Each complex crystallises in a chiral trigonal/hexagonal space group with an asymmetric unit composed of solvent molecules and one third of an L ligand, with the ligand having crystallographic three-fold symmetry. For both complexes, the thiazole-methyl groups of the ligand extending outwards in-plane with the CTG-arene faces, at C_{thiazole}-CH₂-O-C_{aryl} torsion angle *ca.* 179°, however the rotation of the thiazole is distinct for the two complexes, Fig. 1.

In L·4(H₂O) there are two types of water molecules with one sited on a three-fold rotation axis. A network of hydrogen bonds forms between the thiazole groups and waters of crystallisation (Fig. 1a) such that a tetrameric cluster of water molecules connects to three L ligands, and *vice versa* to form a 2D network

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Fig. 1 Clathrate complexes of L. (a) 2D hydrogen bonded network from crystal structure of L·4(H₂O) viewed down c axis. Water shown as spheres, N...O and O...O separations are 2.895 and 2.918 Å respectively. (b) Packing diagram of L·3(CH₃NO₂) with CH₃NO₂ shown in ball-and-stick.

with 6³ topology. In the overall lattice layers of networks pack in an ABAB arrangement with a rotated bowl-in-bowl stacking of the cavitand ligands (see ESI[†]).

Bowl-in-bowl stacking of ligands also occurs in complex L·3(CH₃NO₂), however here the ligands are perfectly aligned with a separation of 4.4 Å, Fig. 1b. The CH₃NO₂ guest molecules occupy lattice positions within the structure. Both clathrate complexes exhibit spontaneous resolution as all L ligands within each structure are the same enantiomer. Spontaneous resolution on crystallisation is known as conglomerate formation.

Vapour diffusion of diethyl ether into a dimethylformamide (DMF) solution of L and AgX (X = ReO₄⁻, BF₄⁻) gave crystals of the complexes [Ag₁₂(L)₈]·12ReO₄·n(DMF) **1** and [Ag₁₂(L)₈]·12BF₄·n(DMF) **2**.[‡] The structures of complexes **1** and **2** were determined by single crystal crystallography and are isostructural, being

solved in the chiral cubic space group *F*432.[‡] The asymmetric unit of complex **1** comprises one third of ligand L, a Ag(I) cation on a 2-fold axis and fragments of ReO₄⁻ anions, one on a 3-fold axis and the other, disordered and with low occupancy, sited on a 4-fold axis.[§]

As for the ligand structures, ligand L in **1** has crystallographic 3-fold symmetry, however with a quite different conformation where the thiazole-methyl groups are folded inwards above the cavitand bowl, at C_{thiazole}-CH₂-O-C_{aryl} torsion angle -83.6°. As expected for a thiazole, only the nitrogen heterocyclic atoms act as donors to the metal, with each thiazole binding to a separate Ag(I) cation. The Ag(I) cations have near linear coordination to thiazole groups of two L ligands at Ag-N distance 2.151(15) Å and N-Ag-N angle 175.9(8)°. Each L1 ligand thus bridges between three metals and each metal links two L1 ligands. These linkages result in a discrete [Ag₁₂L₈]¹²⁺ cubic cage with the L1 ligands at the corners, Fig. 2a. The Ag(I) cations are located at the internal rather than external edge of the cube due to the conformation adopted by the thiazole-methyl groups. The S-heteroatoms of each thiazole moiety are located at the faces of the cube, and the space-filling view, Fig. 2b, shows small windows into the cubes which are bounded by these S atoms (S...S distances across diagonal of window is ca. 6 Å). Space-filling also shows that Ag(I) centres are well within the cube, and thus protected from additional coordination *exo* to the cage. The Ag(I) cations are arranged in a cuboctahedron with Ag...Ag separation of 8.71 Å. The overall size of the cube is ca. 27 Å across the body diagonal measured between centres of the -(CH₂)₃- cavitand ligands.

The eight L ligands within each [Ag₁₂L₈]¹²⁺ cube are the same ligand enantiomer, hence the cube displays homochiral recognition or chiral self-sorting. Homochiral self-sorting of a racemic mixture of enantiomeric ligands during metal-directed self-assembly processes has been previously reported, both for helicate systems,¹¹ and also for cage-like assemblies,^{6-8,12} including with CTG-type ligands.⁶⁻⁸ Complex **1** displays a further level of spontaneous chiral resolution as it crystallises as a conglomerate where each crystal of **1** contains only one enantiomer of the chiral [Ag₁₂L₈]¹²⁺ cube. The bulk sample contains crystals of both enantiomers, and coincidentally the selected crystal studied for the isostructural complex **2** was of the opposite enantiomer to that selected for **1** (see ESI[†]). Spontaneous resolution is not common for metallo-supramolecular systems. Resolution of metallo-supramolecular species with achiral ligands and helical chirality at the metal has been reported for a M₃L₃ helicate¹³ and recently Rissanen and co-workers reported the first example for a 3D metal-linked coordination cage.¹⁴ A handful of examples have been reported for simpler CTG-type systems, namely a metallated hemi-cryptophane,¹⁵ a CTG-type coordination polymer,¹⁶ a trinuclear Re(I)₃L complex¹⁷ as well as organic cage cryptophanes.¹⁸

There are extensive π-π stacking interactions between [Ag₁₂(L)₈]¹²⁺ cubes in the crystal lattice, with each arene of the tribenzo[*a,d,g*]-cyclononatriene ligand core forming a face-to-face interaction at centroid separation 3.81 Å. The cubes form a cubic close packed array, Fig. 2c, with π-stacking along the cube edges such that each cube interacts with twelve others.





Fig. 2 From the crystal structure of $[\text{Ag}_{12}(\text{L})_8] \cdot 12\text{ReO}_4 \cdot n(\text{DMF})$ **1** showing (a) the $[\text{Ag}_{12}(\text{L})_8]^{12+}$ cube; (b) space-filling view; (c) partial packing of cubes through π - π stacking interactions into a cubic close packed array.

There are two crystallographically distinct ReO_4^- positions in the crystal lattice of complex **1**, one is fully occupied on a 3-fold axis and is external to the $[\text{Ag}_{12}\text{L}_8]^{12+}$ cube, whilst the other has less than full occupancy, and is located on a 4-fold axis (hence is symmetry-disordered) and located inside the cube, aligned with the facial windows. The very high symmetry of the structure is reflected in the Russian Doll arrangement of prisms that can be identified, with the internal ReO_4^- anions forming an octahedron which is internal to the cuboctahedron of $\text{Ag}(\text{I})$ sites, itself part of the $[\text{Ag}_{12}\text{L}_8]^{12+}$ cube, which is surrounded by a rhombicuboctahedron of external ReO_4^- anions, Fig. 3 and Fig. S19 (ESI[†]).

Thermogravimetric analysis of complexes **1** and **2** showed a gradual weight loss to *ca.* 190–200 °C of *ca.* 7–10% for **1** and **2**

respectively (see ESI[†]), consistent with the loss of approximately 10 DMF solvent molecules from inside the $[\text{Ag}_{12}\text{L}_8]^{12+}$ cubes from **1** and 12 DMF molecules from **2** which are solvation levels easily accommodated within the cube, with an estimated internal volume of *ca.* 5000 Å³ (excluding anions) using a 1.2 Å probe.¹⁹

Crystals of the cube can be re-dissolved in coordinating solvents such as DMSO but this leads to disassociation of the complex. The ¹H NMR spectrum of L and AgBF_4 in *d*₇-DMF was very broad (see Fig. S9, ESI[†]), and ESI-MS of a DMF solution of complex **1** showed only *m/z* peaks corresponding to $[\text{AgL}]^+$ and $[\text{AgL}_2]^+$ species (Fig. S10, ESI[†]).

In summary, the racemic ligand (±)-tris-(4-methylthiazolyl)-cyclotriguiacylene crystallises in a chiral fashion as a hydrate, and $\text{Ag}(\text{I})$ complexes of the same ligand form a Ag_{12}L_8 cube structure that shows dual levels of chiral sorting, with both chiral self-sorting of the cubes and spontaneous resolution of the chiral cubes on crystallisation for both ReO_4^- and BF_4^- salts.

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Notes and references

‡ Crystal data. L·4(H₂O): C₃₆H₄₁N₃O₁₀S₃, *M_r* = 771.9, hexagonal, *a* = *b* = 14.5223(19) Å, *c* = 10.2701(12) Å, *V* = 1875.8(4) Å³, space group *P6₃*, *Z* = 2, λ = 1.54184 Å, θ_{max} = 73.42°, 158 parameters, 1 restraint, *R*₁ = 0.0886 (for 1394 data *I* > 2σ(*I*)), *wR*₂ = 0.2598 (all 2239 data), Flack parameter = -0.07(7). CCDC 1470477. L·3(CH₃NO₂): C₃₉H₄₂N₆O₁₂S₃, *M_r* = 882.97, trigonal (hexagonal), *a* = *b* = 28.138(3) Å, *c* = 4.4445(5) Å, *V* = 3047.4(5) Å³, space group *R3*, *Z* = 3, λ = 1.54184 Å, θ_{max} = 73.82°, 183 parameters, 1 restraint, *R*₁ = 0.0719 (for 1635 data *I* > 2σ(*I*)), *wR*₂ = 1873 (all 2150 data), Flack parameter = 0.08(5). CCDC 1478562. Complex **1**: C₂₈₈H₂₆₄Ag₁₂N₂₄O₉₆Re₁₂S₂₄, *M_r* = 9895.52, cubic, *a* = *b* = *c* = 35.7237(4) Å, *V* = 45590(9) Å³, space group *F432*, *Z* = 4, λ = 1.54184 Å, θ_{max} = 73.74°, 168 parameters, 4 restraints, *R*₁ = 0.0847 (for 2667 data *I* > 2σ(*I*)), *wR*₂ = 0.2886 (all 3794 data), Flack parameter = 0.03(4). CCDC 1470478. Complex **2**: C₂₈₈H₂₆₄Ag₁₂B₁₂F₄₈N₂₄O₄₈S₂₄, *M_r* = 7934.84,

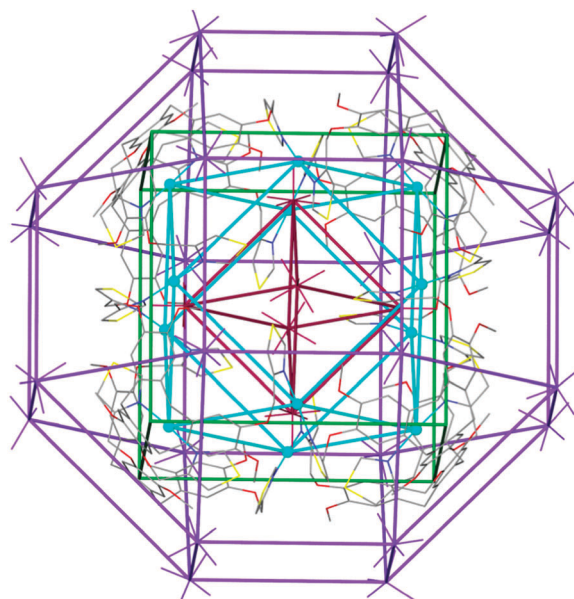


Fig. 3 Matryoshka (Russian Doll) arrangement of prisms-in-prisms of complex **1** with octahedron of *endo*- ReO_4 sites shown in dark red; cuboctahedron of $\text{Ag}(\text{I})$ sites in light blue; cube of ligand positions in green; and rhombicuboctahedron of *exo*- ReO_4 sites in purple.



cubic, $a = b = c = 35.4548(7) \text{ \AA}$, $V = 44568(1) \text{ \AA}^3$, space group $F432$, $Z = 4$, $\lambda = 1.54184 \text{ \AA}$, $\theta_{\text{max}} = 73.64^\circ$, 152 parameters, 42 restraints, $R_1 = 0.0718$ (for 1639 data $I > 2\sigma(I)$), $wR_2 = 0.2256$ (all 2572 data), Flack parameter = 0.15(3). Squeeze was employed for complex 2.¹⁸ CCDC 1470479.

§ Large crystals of complexes 1 and 2 can be obtained repeatedly in good yield in ca. 1 week from crude L which contains a slight HCl contamination. We have previously noted such contamination in a similar ligand synthesis.²⁰ Use of highly pure L only gives a small number of very small crystals of 1 or 2 after a month alongside crystals of the ligand. EDX measurements on individual single crystals do not show Cl with Re, Ag and S the only heavy elements.

- 1 For recent reviews, see: L. Chen, Q. Chen, M. Wu, F. Jiang and M. Hong, *Acc. Chem. Res.*, 2015, **48**, 201; S. Zarra, D. M. Wood, D. A. Roberts and J. R. Nitschke, *Chem. Soc. Rev.*, 2015, **44**, 419; J. J. Henkelis and M. J. Hardie, *Chem. Commun.*, 2015, **51**, 11929; M. Han, D. M. Engelhard and G. H. Clever, *Chem. Soc. Rev.*, 2014, **43**, 1848; T. R. Cook, Y.-R. Zheng and P. J. Stang, *Chem. Rev.*, 2013, **113**, 734; K. Harris and M. Fujita, *Chem. Commun.*, 2013, **49**, 6703; M. D. Ward, *Chem. Commun.*, 2009, 4487.
- 2 For example: C. Browne, S. Brenet, J. K. Clegg and J. R. Nitschke, *Angew. Chem., Int. Ed.*, 2013, **52**, 1944; I. S. Tidmarsh, T. B. Faust, H. Adams, L. P. Harding, L. Russo, W. Clegg and M. D. Ward, *J. Am. Chem. Soc.*, 2008, **130**, 15167; J.-Y. Xu, X. Qiao, H.-B. Song, S.-P. Yan, D.-Z. Liao, G. Song, Y. Journaux and J. Cano, *Chem. Commun.*, 2008, 6414; Y. Liu, V. Kravtsov, R. D. Walsh, P. Poddar, H. Srikanth and M. Eddaoudi, *Chem. Commun.*, 2004, 2807; Z. R. Bell, L. P. Harding and M. D. Ward, *Chem. Commun.*, 2003, 2432; S. Roche, C. Haslam, S. L. Heath and J. A. Thomas, *Chem. Commun.*, 1998, 1681.
- 3 M. Wang, C. Wang, X.-Q. Hao, X. Li, T. J. Vaughn, Y.-Y. Zhang, Y. Yu, Z.-Y. Li, M.-P. Song, H.-B. Yang and X. Li, *J. Am. Chem. Soc.*, 2014, **136**, 10499; C. Wang, X.-Q. Hao, M. Wang, C. Guo, B. Xu, E. N. Tan, Y.-Y. Zhang, Y. Yu, Z.-Y. Li, H.-B. Yang, M.-P. Song and X. Li, *Chem. Sci.*, 2014, **5**, 1221; B. F. Abrahams, S. J. Egan and R. Robson, *J. Am. Chem. Soc.*, 1999, **121**, 3535.
- 4 D. Xu and R. Warmuth, *J. Am. Chem. Soc.*, 2008, **130**, 7520.
- 5 J. J. Henkelis, C. J. Carruthers, S. E. Chambers, R. Clowes, A. I. Cooper, J. Fisher and M. J. Hardie, *J. Am. Chem. Soc.*, 2014, **136**, 14393; B. F. Abrahams, B. A. Boughton, N. J. Fitzgerald, J. L. Holmes and R. Robson, *Chem. Commun.*, 2011, **47**, 7404; T. K. Ronson, H. Nowell, A. Westcott and M. J. Hardie, *Chem. Commun.*, 2011, **47**, 176; B. F. Abrahams, N. J. Fitzgerald and R. Robson, *Angew. Chem., Int. Ed.*, 2010, **49**, 2896; A. Westcott, J. Fisher, L. P. Harding, P. Rizkallah and M. J. Hardie, *J. Am. Chem. Soc.*, 2008, **130**, 2950; Z. Zhong, A. Ikeda, S. Shinkai, S. Sakamoto and K. Yamaguchi, *Org. Lett.*, 2001, **3**, 1085.
- 6 A. Schaly, Y. Rousselin, J.-C. Chambron, E. Aubert and E. Espinosa, *Eur. J. Inorg. Chem.*, 2016, 832; J. J. Henkelis, T. K. Ronson, L. P. Harding and M. J. Hardie, *Chem. Commun.*, 2011, **47**, 6560; C. J. Sumby, J. Fisher, T. J. Prior and M. J. Hardie, *Chem. – Eur. J.*, 2006, **12**, 2945.
- 7 J. J. Henkelis, J. Fisher, S. L. Warriner and M. J. Hardie, *Chem. – Eur. J.*, 2014, **20**, 4117; T. K. Ronson, J. Fisher, L. P. Harding and M. J. Hardie, *Angew. Chem., Int. Ed.*, 2007, **46**, 9086.
- 8 T. K. Ronson, J. Fisher, L. P. Harding, P. J. Rizkallah, J. E. Warren and M. J. Hardie, *Nat. Chem.*, 2009, **1**, 212.
- 9 S. Klotzbach and F. Beuerle, *Angew. Chem., Int. Ed.*, 2015, **54**, 10356; J. Strübe, B. Neumann, H.-G. Stammer and D. Kuck, *Chem. – Eur. J.*, 2009, **15**, 2256.
- 10 J. Wei, Z.-M. Li, X.-J. Jin, X.-J. Yao, X.-P. Cao, H.-F. Chow and D. Kuck, *Chem. – Asian J.*, 2015, **10**, 1150.
- 11 O. Gidron, M. Jirásek, N. Trapp, M.-O. Ebert, X. Zhang and F. Diederich, *J. Am. Chem. Soc.*, 2015, **137**, 12502; C. Gütz, R. Hovorka, N. Struch, J. Bunzen, G. Meyer-Eppler, Z.-W. Qu, S. Grimme, F. Topic, K. Rissanen, M. Cetina, M. Engeser and A. Lützen, *J. Am. Chem. Soc.*, 2014, **136**, 11830; M. Boiocchi and L. Fabbri, *Chem. Soc. Rev.*, 2014, **43**, 1835; M. Huntin, C. J. Cramer, L. Gagliardi, A. R. M. Shahi, G. Bernardinelli, R. Cerny and J. R. Nitschke, *J. Am. Chem. Soc.*, 2007, **129**, 8774; S. G. Telfer, T. Sato, R. Kuroda, J. Lefebvre and D. B. Leznoff, *Inorg. Chem.*, 2004, **43**, 421.
- 12 S. A. Boer and D. R. Turner, *Chem. Commun.*, 2015, **51**, 17375; L.-L. Yan, C.-H. Tan, G.-L. Zhang, L.-P. Zhou, J.-C. Bünzli and Q.-F. Sun, *J. Am. Chem. Soc.*, 2015, **137**, 8550; C. Gütz, R. Hovorka, G. Schnakenburg and A. Lützen, *Chem. – Eur. J.*, 2013, **19**, 10890; C. Maeda, T. Kamada and A. Osuka, *Coord. Chem. Rev.*, 2007, **251**, 2743.
- 13 R. Krämer, J.-M. Lehn, A. de Cian and J. Fischer, *Angew. Chem., Int. Ed. Engl.*, 1993, **32**, 703.
- 14 P. Bonakdarzadeh, F. Pan, E. Kalenius, O. Jurček and K. Rissanen, *Angew. Chem., Int. Ed.*, 2015, **54**, 14890.
- 15 Y. Makita, K. Sugimoto, K. Furuyoshi, K. Ikeda, T. Fujita, S.-I. Fujiwara and A. Ogawa, *Supramol. Chem.*, 2011, **23**, 269.
- 16 J. J. Henkelis, S. A. Barnett, L. P. Harding and M. J. Hardie, *Inorg. Chem.*, 2012, **51**, 10657.
- 17 F. L. Thorp-Greenwood, V. E. Pritchard, M. P. Coogan and M. J. Hardie, *Organometallics*, 2016, **35**, 1632.
- 18 M. A. Little, J. Donkin, J. Fisher, M. A. Halcrow, J. Loder and M. J. Hardie, *Angew. Chem., Int. Ed.*, 2012, **51**, 764; C. Givélet, J. Sun, D. Xu, T. J. Emge, A. Dhokte and R. Warmuth, *Chem. Commun.*, 2011, **47**, 4511.
- 19 A. L. Spek, *Acta Crystallogr., Sect. A: Found. Crystallogr.*, 1990, **A46**, C34.
- 20 T. K. Ronson and M. J. Hardie, *CrystEngComm*, 2008, **10**, 1731.

