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Exploring affordable cocatalysts with high performance for boosting charge separation and CO_2 activation is an effective strategy to reinforce CO_2 photoreduction efficiency. Herein, well-defined Co_9S_8 cages are exploited as a nonprecious promoter for visible-light CO_2 reduction. The Co_9S_8 cages are prepared *via* a multistep strategy with ZIF-67 particles as the precursor and fully characterized by physicochemical techniques. The hollow Co_9S_8 cocatalyst with a high surface area and profuse catalytically active centers is

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Special Professor in 2021. His longstanding interests are the developments of photocatalytic and photoelectrochemical devices for energy and environmental applications, such as water splitting, CO_2 reduction, organosynthesis and others. He has published more than 45 peer-reviewed publications, which have garnered more than 2000 citations and an H-index of 21.

discovered to accelerate separation and transfer of light-induced charges, and strengthen concentration and activation of CO_2 molecules. In a hybrid photosensitized system, these Co_9S_8 cages efficiently promote the deoxygenative reduction of CO_2 to generate CO, with a high yield rate of 35 μ mol h⁻¹ (i.e., 35 mmol h⁻¹ g⁻¹). Besides, this cocatalyst is also of high stability for the CO_2 photoreduction reaction. Density functional theory (DFT) calculations reveal that the $Ru(bpy)_3^{2+}$ photosensitizer is strongly absorbed on the Co_9S_8 (311) surface through forming four Co-C bonds, which can serve as the "bridges" to ensure quick electron transfer from the excited photosensitiser to the active Co_9S_8 cocatalyst, thus promoting the separation of photoexcited charges for ehannced CO_2 reduction performance.

Solar-driven CO_2 reduction, an ideal strategy to afford low-carbon fuels (*i.e.*, CO and CH_4) with renewable energy, is long being intensively studied,¹⁻⁶ given its promise in addressing the issues of energy and environment. However, as limited by the extreme chemical inertness of linear CO_2 molecules, the easy recombination feature of photoexcited charges and the intricacy of the multi-electron/proton engaged reaction,⁷⁻¹⁰ photocatalytic CO_2 reduction currently suffers from low efficiency coupled with moderate selectivity that is far behind the criteria of practical utilization.

During the past few decades, a great deal of strategies have been proposed to improve CO₂ photoreduction efficiency with inspiring progress realized. ^{11–16} Therein, cocatalyst engineering is verified to be a preferred approach. This is because cocatalysts can enable capture and activation of CO₂ molecules, accelerate separation and transfer kinetics of charge carriers, and provide catalytically active sites to selectively operate the redox reactions. ^{17,18} The noble metal nanoparticles are the classic CO₂ reduction cocatalysts with high catalytic performance, ^{19–22} and alternatively, the transition metal ions/complexes of cobalt and nickel can serve as efficient promoters for CO₂ photoreduction in cooperation with specific light-sensitizers. ^{23–30} Comparatively, considering the requirements of large-scale execution (*i.e.*, low price, high abundance, easy

Nanoscale Communication

separation/retrieval), exploring heterogeneous cocatalysts composed of cost-affordable elements for CO2 photoreduction is highly desirable.

Besides the chemical composition, designing catalytic materials with favored architectures is also of vital importance for photocatalytic CO₂ reduction.³¹ As a group of multifunctional catalysts for energy storage and conversion, 32,33 hollow structures are actively employed for CO2 photoreduction, 34-40 owing to their inherent advantages, such as a shortened perpendicular path for charge transfer to prevent their recombination, plentiful surface-active sites to boost CO2 capture and redox catalysis, and strengthened interior reflection/scattering to enhance incident light utilization. 41-43

With all the aforementioned concerns in mind, in this work, the well-defined Co₉S₈ cages are applied as a cocatalyst coupled with Ru(bpy)₃²⁺ (abbreviated as Ru, bpy = 2'2-bipyridine) as a photosensitizer for CO2 reduction under visible light. The dodecahedral Co₉S₈ cage is created via a multi-step method as depicted schematically in Fig. 1. Starting with Co₂ and 2-methylimidazole as the raw materials, the zeolitic imidazolate framework (ZIF)-67 dodecahedron is first prepared through a reported precipitation method with some adjustments (step I).44,45 Next, the ZIF-67 particle is converted to a CoS_x cage by a hydrothermal sulfidation reaction (step II). Finally, the Co₉S₈ cage is harvested after annealing the CoS_x intermediate under a N₂ atmosphere at 550 °C (step III).

The Co_oS₈ cocatalyst is revealed by diverse photo-/electrochemical tests to augment CO2 activation and promote separation of light-excited charges. In a tandem photochemical system, the Co₉S₈ cocatalyst enables the deoxygenative reduction of CO2 by visible light, with a CO formation rate of 35 μ mol h⁻¹, under the cooperation of Ru(bpy)₃²⁺ as a photosensitizer. Indeed, the reduction ability of the Co₉S₈ catalyst in photocatalysis has been demonstrated by hydrogen evolution reactions; 46-50 however, its talent for CO₂ photoreduction has been seldom exploited so far.

Production of the ZIF-67 precursor was confirmed by powder X-ray diffraction (XRD) and energy-dispersive X-ray (EDX) characterization (Fig. S1, ESI†). The field emission scanning electron microscopy (FESEM) images show that the as-synthesized ZIF-67 particles present a morphology of nearly monodisperse dodecahedra with high uniformity (Fig. 2a and b).

After the sulfidation treatment, the ZIF-67 precursor was entirely transformed into an amorphous cobalt sulfide (CoS_x) product with a Co/S ratio of about 1:3.42 (Fig. S2, ESI†). The



Fig. 1 Schematic illustration of the preparation processes of dodecahedral Co₉S₈ cages: (I) self-assembly, (II) sulfidation, and (III) annealing in N_2 .

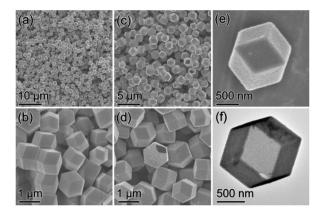


Fig. 2 (a and b) FESEM images of ZIF-67 dodecahedra. (c-e) FESEM images and (f) TEM image of CoS_x cages.

FESEM images indicate that the CoS_x intermediate can inherit the dodecahedral shape of the parental ZIF-67 particles perfectly (Fig. 2c). From the magnified FESEM images (Fig. 2d and e), it is found that the CoS_x particles have a relatively rougher surface than the ZIF-67 precursor, and the observed clear cavity of the broken CoSx dodecahedra implies their hollow structure. The empty interior of the well-defined CoS_x cages is then confirmed visually by the TEM image (Fig. 2f). The possible formation process of CoS_x cages is clarified as follows. During the hydrothermal sulfidation reaction, the thioacetamide is decomposed to release the sulfide ions to react with the cobalt ions on the ZIF-67 particle surface, yielding an outermost layer of CoS_x. Further formation of the CoS_x shell continues via the reactions between inward diffusing sulfide ions and outward diffusing metal ions, finally producing the polyhedral CoS_x cages. 46,51

The CoS_x material was treated by annealing in N_2 to attain the Co₉S₈ product. To monitor the conversion procedure, thermogravimetry (TG) analysis was conducted. As shown in

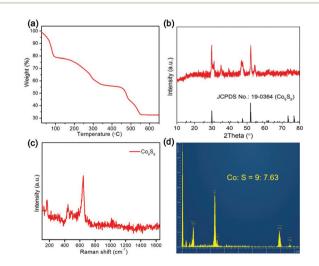


Fig. 3 (a) TG curve of CoS_x cages in in N_2 . (b) XRD pattern, (c) Raman spectrum and (d) EDX spectrum of Co₉S₈ cages.

Communication Nanoscale

Fig. 3a, the result of TG indicates that the CoS_x intermediate undergoes gradual weight loss during the thermal treatment due to sulphur volatilization, and finally forms a stable product at 550 °C, which is confirmed to be the Co₉S₈ phase (JCPDS card no.: 19-0364) by XRD analysis (Fig. 3b). No impurity XRD peaks are discerned, signifying the high phase purity of the Co₉S₈ material. The Raman spectrum of the Co₉S₈ sample presents a distinct vibrational peak at around 640 cm⁻¹ together with a set of bands below 750 cm⁻¹ (Fig. 3c), indexing to the characteristic Raman peaks of the Co₉S₈ phase.^{52,53} The EDX spectrum shows that the Co₉S₈ product is only composed of Co and S elements with a molar ratio of about 9:7.63 (Fig. 3d), which is close to the theoretical value.

The morphology and microstructures of the Co₉S₈ product were researched by FESEM and TEM. As can be seen from the FESEM images (Fig. 4a and b), the Co₉S₈ particles well preserve the pristine polyhedral construction of the CoS_x intermediate without perceptible agglomeration or fractures. The careful FESEM analysis discloses that the surface of Co₉S₈ particles experiences slight shrinkage after the thermal treatment (Fig. 4c). The structural features of Co₉S₈ cages revealed by TEM are consistent with the results of FESEM (Fig. 4d and e). The crystal features of Co₉S₈ cages were then checked by highresolution TEM (HRTEM) analysis. As shown in Fig. 4f, the HRTEM image displays strong lattice fringes with an interlayer d-spacing of 0.28 nm determined by the corresponding inverse fast Fourier transformation (IFFT) image and line scans (Fig. 4g and h), which is assigned to the (222) crystal plane of the cubic Co₉S₈ phase. The Co₉S₈ product is verified to be a polycrystalline material, as revealed by the selected area elec-

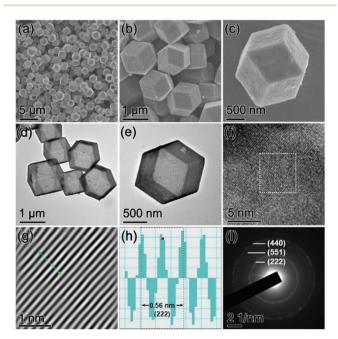


Fig. 4 (a-c) FESEM images, (d and e) TEM images, (f) HRTEM image, (g and h) the corresponding inverse fast Fourier transformation (IFFT) image and line scans of the dotted square region in (f), and (i) SAED pattern of Co₉S₈ cages.

tron diffraction (SAED) pattern (Fig. 4i), in which the clear diffraction fringes are attributed to the interplanar spacings of the cubic Co₉S₈ phase.

The surface chemical states of elements of the Co₉S₈ material were studied by X-ray photoelectron spectroscopy (XPS). As shown in Fig. 5a, the high-resolution Co 2p spectrum is fitted into two sets of spin-orbit doublets and satellite (Sat.) peaks. The first doublet located at 778.3 and 779.4 eV and the second doublet with binding energies of 793.4 and 795.5 eV are attributed to Co 2p_{3/2} and Co 2p_{1/2}, respectively. The presence of these doublets indicates the coexistence of Co3+ and Co₂⁺, which is in agreement with the results of previous work.^{54–56} In the XPS spectrum of S 2p (Fig. 5b), the fitted peaks with binding energies of about 162.1 and 163.1 eV are assigned to the Co-S species, while the other two peaks centered at 163.7 and 164.9 eV are attributable to the S-O bonds, 46 implying the surface oxidation of the Co₉S₈ sample that is generally detected during XPS measurements.

The textural properties of Co₉S₈ cages were examined by N₂ sorption measurements. As displayed in Fig. 5c, classified type-II N₂ adsorption-desorption isotherms together with a type H3 hysteresis loop are observed, suggesting the formation of mesopores in the material, which can be further confirmed by the corresponding pore size distribution plot (inset, Fig. 5c). Generation of the porous structure of Co_9S_8 cages is positive to facilitate mass transportation for heterogeneous photocatalytic applications. The Co₉S₈ material is measured to have a high specific Brunauer-Emmett-Teller (BET) surface area of about 137 m² g⁻¹, which should be mainly contributed by the well-defined hollow configuration. Such a high surface area will enable the capture and concentration of CO2 molecules, as reflected by the CO₂ adsorption isotherm, which gives a maximum CO₂ uptake of about 13.5 cm³ g⁻¹ at 0 °C and 1 atm (Fig. 5d). These textural features of Co₉S₈ cages are ben-

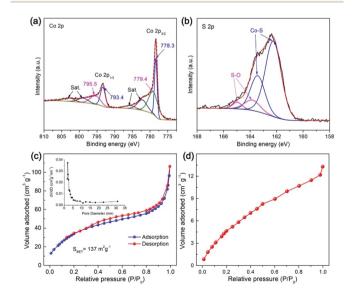


Fig. 5 (a) Co 2p and (b) S 2p XPS spectrum, (c) N₂ adsorption-desorption isotherms, and (d) CO₂ adsorption isotherm of Co₉S₈ cages.

Nanoscale Communication

eficial to afford rich exposed catalytically active sites and boost mass transport for heterogeneous CO₂ fixation reactions.

We then evaluated the performance of Co₉S₈ cages by the visible-light CO2 reduction reaction conducted in a classic hybrid system, ^{2,25,54,57-61} engaging Ru(bpy)₃²⁺ as the light-harvester and H₂O/acetonitrile/TEOA mixture as the reaction medium. Fig. 6a presents the CO2 reduction activity of the Co₉S₈ cocatalyst as a function of reaction time. As can be seen, the system manifests a high photocatalytic performance in the initial first 1 h of the reaction, showing a CO2-to-CO conversion rate of 35 µmol h⁻¹, coupled with a H₂-releasing rate of 12 µmol h⁻¹. In addition to CO and H₂, no detectable formation of hydrocarbon products is observed in the liquid phase. On further continuing the reaction, generation of the products increases progressively, but with a decreased rate, which should result from degradation of the ruthenium photosensitizer. 57,62,63 The total yield of CO reaches 65 µmol after photoreaction for 7 h, corresponding to a catalytic turnover number (TON) of about 54 relative to the cocatalyst.

The working mechanism of the CO₂ reduction system was inspected by controlling the reaction conditions. As shown in Fig. 6b, compared to that of the normal reaction (column 1),

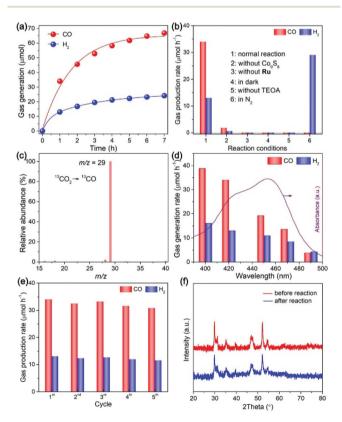


Fig. 6 (a) CO_2 photoreduction activity of the Co_9S_8 cocatalyst as a function of reaction time. (b) Performance of the CO_2 reduction system under various conditions. (c) Mass spectrum of CO_3 yielded from the $^{13}CO_2$ isotope reaction. (d) CO_3 and CO_3 generation by visible light with different wavelengths. (e) Stability tests of the Co_9S_8 cocatalyst. (f) XRD patterns of the Co_9S_8 cocatalyst before and after CO_2 reduction reactions.

the evolution of CO/H_2 reduces significantly once the Co_9S_8 cocatalyst is omitted from the system, highlighting its critical role in promoting the CO_2 reduction reaction. No product is detected without the introduction of a photosensitizer (column 3) or visible light irradiation (column 4), which indicates that the CO_2 reduction reaction is a visible-light-sensitized process. It is revealed that TEOA is essential to achieve the CO_2 reduction photocatalysis, as the reaction would be completely terminated without its participation (column 5), which matches well with the results of reported studies in analogous photochemical systems. $^2, ^2, ^2, ^5, ^7, ^6, ^2, ^64$ When employing N_2 as the gas feedstock to substitute CO_2 for running the reaction (column 6), the only product generated is H_2 , suggesting that the CO product should be derived from the CO_2 gas.

To directly track the origin of CO, we performed a 13 C-labelled reaction with 13 CO₂ as the reactant and analyzed the produced CO by mass spectroscopy (MS). As indicated in Fig. 6c, the MS spectrum manifests a predominant peak with a m/z value of 29, which is attributed distinctly to 13 CO. This observation firmly validates the source of CO generation, that is, the CO₂ feedstock.

The CO_2 photoreduction reactions were also initiated with light irradiation of different wavelengths through applying specific long-pass cut-off filters. As revealed in Fig. 6d, the formation of CO/H_2 diminishes gradually under light irradiation of longer wavelengths, because of the decline in the incident photons when extending the wavelength of the light filter. These findings indicate that the CO_2 reduction is motivated by light excitation of the photosensitizer.

To examine the stability of the Co_9S_8 cocatalyst, after CO_2 photoreduction reactions, it was separated, washed, and readded into fresh reaction mixtures for repeated operations. No apparent decrease in CO/H_2 production is detected during the stability tests (Fig. 6e), pointing to its high activity stability. The results of XRD, FTIR, Raman, and XPS tests for the Co_9S_8 cocatalyst before and after photoreaction expose its high stabilities in the crystal, chemical and surface structures in the photocatalytic CO_2 reduction system (Fig. 6f, and Fig. S3, ESI†).

In order to gain an understanding of the high performance of the CO_2 photoreduction system mediated by $\mathrm{Co}_9\mathrm{S}_8$ cages, we carried out photo-/electro-chemical measurements. To demonstrate the function of $\mathrm{Co}_9\mathrm{S}_8$ in the CO_2 reduction reaction, linear sweep voltammetry (LSV) was conducted. As shown in Fig. 7a, the $\mathrm{Co}_9\mathrm{S}_8$ catalyst delivers a more positive onset potential and a higher cathodic current in CO_2 -saturated solution than those under N_2 -saturated conditions. These outcomes verify that the $\mathrm{Co}_9\mathrm{S}_8$ cages can activate CO_2 molecules and drive their reduction reaction.

UV-vis diffuse reflectance spectra (DRS) indicate that the Co_9S_8 cocatalyst and $\text{Ru}(\text{bpy})_3^{2^+}$ photosensitizer show strong visible light absorption (Fig. 7b), while their hybrid possesses an optical harvesting ability. By using UV-vis DRS and the Tauc curve of Co_9S_8 (Fig. S4, ESI†), its bandgap energy is measured to be 0.98 eV. ^{37,47} To define the conduction band bottom (CBM) and valence band maximum (VBM) positions, the flat

Communication Nanoscale

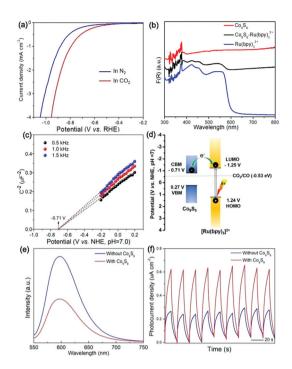


Fig. 7 (a) LSV curve of the Co₉S₈ cocatalyst under N₂- and CO₂-saturated solutions. (b) DRS spectra of Co₉S₈, Ru(bpy)₃²⁺ and Ru(bpy)₃²⁺/ Co₉S₈ hybrid. (c) Mott-Schottky plots. (d) Schematic illustration of energy levels and electron transfer from the ruthenium photosensitizer to the Co₉S₈ cocatalyst. (e) PL spectra and (f) transient photocurrent generation of reaction systems with and without the Co₉S₈ cocatalyst.

band potential of Co₉S₈ was estimated by electrochemical Mott-Schottky plots. As shown in Fig. 7c, the derived CBM potential of the Co₉S₈ cocatalyst is about -0.71 V (vs. NHE, pH = 7.0), which integrated with the bandgap energy determines its VBM position at about 0.27 V (vs. NHE, pH = 7.0). Regarding the band structure of Co₉S₈, the lowest unoccupied molecular orbital (LUMO) and highest occupied molecular orbital (HOMO) levels of Ru(bpy)32+, and the redox potential for reducing CO₂ to CO (Fig. 7d), ^{26,62,66} the excited electrons of the photosensitizer are encouraged thermo-dynamically to travel to the cocatalyst for driving the CO₂ reduction reaction.

To monitor the separation and transport of light-excited charges in the photochemical system, room temperature photoluminescence (PL) was implemented. As revealed in Fig. 7e, the system involving Co₉S₈ reveals a quenched PL emission compared to the system free of the cocatalyst, suggesting the prohibited recombination of light-stimulated charges of the former. 31,40 Simultaneously, the time-resolved PL (TRPL) spectra (Fig. S5, ESI†) reveal that the average emission lifetime of the reaction mixture with the cocatalyst (399 ns) is much shorter than that of the mixture without its presence (450 ns), indicating that quick electron transfer exists between the photosensitizer and the cocatalyst. Consistently, the photocurrent generation of the Co₉S₈-mediated system is higher than that of the system without its existence (Fig. 7f), which mirrors the boosted transport of charge carriers induced by the metal

sulfide cocatalyst. 26,47,67 All the above results show that the Co₉S₈ cocatalyst can expedite separation and transfer of lightexcited charges and promote activation of CO2 molecules, thus reinforcing the CO₂ photoreduction efficiency.

Density functional theory (DFT) calculations were carried out to gain some insights into the interaction between the Ru $(bpy)_3^{2+}$ photosensitizer and Co_9S_8 cocatalyst. The results prove that the Ru(bpy)₃²⁺ complex can be strongly adsorbed on the Co_9S_8 (311) surface, in which the carbon atoms of Ru(bpy)₃²⁺ and the surface Co atoms of Co₉S₈ form four Co-C bonds with a length of about 2.1 Å (Fig. 8). The chemical bonds between the photosensitizer and the cocatalyst may serve as the "bridges" to guarantee quick electron migration from the excited photosensitiser to the active Co₉S₈ cocatalyst to run the CO2 reduction reaction.

Finally, we have proposed a probable mechanism of the CO₂ photoreduction reaction catalyzed by the caged Co₉S₈ cocatalyst. Being stimulated by visible light illumination, the ground state photosensitizer Ru(bpy)₃²⁺ moves to the excited state of $Ru(bpy)_3^{2+*}$. Such an excited state will react with the electron donor TEOA through a reductive quenching process, leading to the formation of the photosensitizer in a reduced state of $Ru(bpy)_3^{+}$. ^{26,64} The excited electrons of the $Ru(bpy)_3^{+}$ species will delocalize and move to the porous Co₉S₈ cocatalyst to run the CO₂-to-CO conversion reaction, 60,62,65 during which the Co₉S₈ cages not only strengthen the adsorption and activation of CO₂ molecules but also push the separation and migration of light-generated charges to support the reaction.

In summary, well-defined dodecahedral Co₉S₈ cages have been fabricated and applied as a new cocatalyst for CO₂ photoreduction with visible light. Owing to the high surface area and plentiful catalytically reactive locations, the hollow C₉S₈ cocatalyst effectively boosts CO2 activation and prevents charge recombination. In a classic light-sensitized system, such a non-noble-metal cocatalyst can powerfully promote the conver-

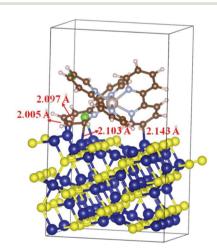


Fig. 8 The structure of Ru(bpy)₃Cl₂ adsorbed on the Co₉S₈ (311) surface with an adsorption energy of -4.5 eV. The Co, S, Ru, N, C, H, and Cl atoms are denoted by blue, yellow, pink, gray, brown, white, and green, respectively.

Nanoscale Communication

sion of CO₂ to CO with high efficiency and good stability. Inspired by this contribution, future studies about the sulfide semiconductors with narrow bandgaps and the hybrids of Co₉S₈/semiconductor for photocatalytic CO₂ reduction may be anticipated.

Author contributions

Xiahui Lin: conceptualization, methodology, feasibility analysis, validation, writing-original draft, and writing-review and editing. Zidong Xie: methodology, feasibility analysis, and validation. Bo Su: methodology, validation, and writing-review and editing. Mei Zheng: theoretical calculation. Wenxin Dai: feasibility analysis. Yidong Hou: writing-review and editing. Zhengxin Ding: writing-review and editing. Wei Lin: theoretical calculation and supervision. Yuanxing Fang: writing-review and editing and supervision. Sibo Wang: validation, writingreview and editing, and supervision.

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

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