



# Z-scheme photocatalyst systems employing Rh- and Ir-doped metal oxide materials for water splitting under visible light irradiation†

Akihiko Kudo,<sup>a</sup> <sup>\*ab</sup> Shunya Yoshino,<sup>a</sup> Taichi Tsuchiya,<sup>a</sup> Yuhei Udagawa,<sup>a</sup> Yukihiko Takahashi,<sup>a</sup> Masaharu Yamaguchi,<sup>a</sup> Ikue Ogasawara,<sup>a</sup> Hiroe Matsumoto<sup>a</sup> and Akihide Iwase <sup>ab</sup>

Received 27th November 2018, Accepted 24th January 2019

DOI: 10.1039/c8fd00209f

Various types of Z-scheme systems for water splitting under visible light irradiation were successfully developed by employing Rh- and Ir-doped metal oxide powdered materials with relatively narrow energy gaps (EG): BaTa<sub>2</sub>O<sub>6</sub>:Ir,La (EG: 1.9–2.0 eV), NaTaO<sub>3</sub>:Ir,La (EG: 2.1–2.3 eV), SrTiO<sub>3</sub>:Ir (EG: 1.6–1.8 eV), NaNbO<sub>3</sub>:Rh,Ba (EG: 2.5 eV) and TiO<sub>2</sub>:Rh,Sb (EG: 2.1 eV), with conventional SrTiO<sub>3</sub>:Rh (an H<sub>2</sub>-evolving photocatalyst) or BiVO<sub>4</sub> (an O<sub>2</sub>-evolving photocatalyst), and suitable electron mediators. The Z-scheme systems were classified into three groups depending on the combination of H<sub>2</sub>- and O<sub>2</sub>-evolving photocatalysts and electron mediator. The Z-scheme systems combining BaTa<sub>2</sub>O<sub>6</sub>:Ir,La with BiVO<sub>4</sub>, and NaTaO<sub>3</sub>:Ir,La with BiVO<sub>4</sub> were active when a [Co(bpy)<sub>3</sub>]<sup>3+/2+</sup> redox couple was used rather than an Fe<sup>3+/2+</sup> one. The combination of SrTiO<sub>3</sub>:Ir with SrTiO<sub>3</sub>:Rh gave an activity when the [Co(bpy)<sub>3</sub>]<sup>3+/2+</sup> and Fe<sup>3+/2+</sup> redox couple ionic mediators were used. The Z-scheme systems combining NaNbO<sub>3</sub>:Rh,Ba and TiO<sub>2</sub>:Rh,Sb with SrTiO<sub>3</sub>:Rh showed activities by using the [Co(bpy)<sub>3</sub>]<sup>3+/2+</sup> and Fe<sup>3+/2+</sup> redox couples and also *via* interparticle electron transfer by just contact with/without reduced graphene oxide (RGO). These suitable combinations can be explained based on the impurity levels of doped Rh<sup>3+</sup> and Ir<sup>3+</sup> toward the redox potentials of the ionic mediators for the Z-scheme systems employing ionic mediators, and *p*-/*n*-type and onset potentials of the photocurrent in the photoelectrochemical properties of those photocatalyst materials for the Z-scheme systems working *via* interparticle electron transfer.

<sup>a</sup>Department of Applied Chemistry, Faculty of Science, Tokyo University of Science, 1-3 Kagurazaka, Shinjuku-ku, Tokyo 162-8601, Japan. E-mail: a-kudo@rs.kagu.tus.ac.jp

<sup>b</sup>Photocatalysis International Research Center, Research Institute for Science and Technology, Tokyo University of Science, 2641, Yamazaki, Noda-shi, Chiba-ken 278-8510, Japan

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c8fd00209f



# Introduction

Artificial photosynthesis has attracted attention from the view point of solar energy conversion to storable chemical energy. Because solar water splitting is one of the representative reactions, photocatalytic water splitting has extensively been studied.<sup>1–7</sup> Powder-based metal oxide materials are attractive for photocatalytic water splitting because the cost will be low<sup>8–10</sup> and stability will be high compared with other materials such as chalcogenides. Although the efficiency of powder-based photocatalyst systems is behind that of systems of photovoltaic + electrolysis at the present stage, there are some advantages to powder-based photocatalyst systems.

It is crucial to demonstrate a solar water splitting system employing powder-based oxide photocatalysts. The present stage of research of artificial synthesis is along this topic. A reactor system including a gas separation system of evolved hydrogen from oxygen has been studied in addition to the development of photocatalyst materials aimed towards the practical use of solar water splitting.<sup>10</sup> The separation system can be achieved by the use of a suitable separation membrane. Moreover, the safety issue of the co-evolved H<sub>2</sub> and O<sub>2</sub> has also been examined *via* the use of a suitable gas transportation tube. However, even if an excellent reactor system with a gas separation membrane system is established, which photocatalyst is employed for it is still a key issue as high efficient photocatalysts for real solar water splitting into H<sub>2</sub> and O<sub>2</sub> without any sacrificial electron donors and acceptors have not yet been developed. It is essential to develop an efficient photocatalyst for demonstrating a solar water splitting system taking gas separation and safety issues into account.

High efficiency of a photocatalyst can be brought about by a high quantum yield and a response to light with a long wavelength. From this viewpoint, it is vital to develop photocatalysts that can utilize up to 600–700 nm of the solar spectrum. Although the use of noble metals might prevent the photocatalyst materials from practical use, the usage would be allowed if the amount is small and the materials can be recycled.

There are single particulate and Z-scheme systems in powder-based photocatalysts.<sup>2,11</sup> Z-scheme photocatalyst systems have the advantage that photocatalysts active for either H<sub>2</sub> or O<sub>2</sub> evolution can be employed. This means that a Z-scheme system can be constructed with various combinations of H<sub>2</sub>- and O<sub>2</sub>-evolving photocatalysts.<sup>12–15</sup> SrTiO<sub>3</sub>:Rh (an H<sub>2</sub>-evolving photocatalyst) and BiVO<sub>4</sub> (an O<sub>2</sub>-evolving photocatalyst) are representative metal oxide materials for the Z-scheme system.<sup>1,2,16</sup> It should be stressed that photocatalyst sheets prepared by a particle transfer method using SrTiO<sub>3</sub>:Rh,La and BiVO<sub>4</sub>:Mo powders demonstrate a quantum efficiency of 30% at 420 nm and a solar to hydrogen energy conversion efficiency of 1%.<sup>17–19</sup> This result suggests that the Z-scheme system is a promising photocatalyst system for practical solar water splitting. H<sub>2</sub> evolution separated from O<sub>2</sub> evolution is possible in the Z-scheme system if a suitable reactor is designed.<sup>8</sup> However, this photocatalyst sheet consisting of the SrTiO<sub>3</sub>:Rh,La and BiVO<sub>4</sub>:Mo powders responds up to 520 nm because the energy gap (EG) of SrTiO<sub>3</sub>:Rh,La and the band gap (BG) of BiVO<sub>4</sub>:Mo are 2.3 eV and 2.4 eV, respectively. So, it is a key issue to develop Z-scheme photocatalyst systems consisting of metal oxide photocatalysts with a response at longer wavelengths than 520 nm.



Transition metal doping is one strategy to make wide band gap photocatalysts responsive to visible light.<sup>1,2</sup> Rh and Ir are effective dopants in this strategy. We have reported that SrTiO<sub>3</sub>:Rh (EG: 2.3 eV),<sup>20</sup> SrTiO<sub>3</sub>:Rh,Sb (EG: 2.2–2.4 eV),<sup>21</sup> SrTiO<sub>3</sub>:Ir (EG: 1.6–1.8 eV),<sup>20,22</sup> BaTa<sub>2</sub>O<sub>6</sub>:Ir,La (EG: 1.9–2.0 eV)<sup>23</sup> and NaTaO<sub>3</sub>:Ir,La (EG: 2.1–2.3 eV)<sup>24</sup> are active for sacrificial H<sub>2</sub> evolution in the presence of electron donors such as methanol, while SrTiO<sub>3</sub>:Rh,Sb,<sup>21</sup> SrTiO<sub>3</sub>:Ir<sup>20</sup> and TiO<sub>2</sub>:Rh,Sb (EG: 2.1 eV)<sup>25</sup> are active for sacrificial O<sub>2</sub> evolution in the presence of electron acceptors such as Ag<sup>+</sup>. In these photocatalyst materials, the visible light responses are due to electronic transition from the impurity levels formed by the dopants to the conduction bands of the host materials. These energy and band gaps are close to, or narrower than, those of SrTiO<sub>3</sub>:Rh (EG: 2.3 eV)<sup>20</sup> and BiVO<sub>4</sub> (BG: 2.4 eV).<sup>26,27</sup> So, it is attractive to utilize these Rh- and Ir-doped metal oxide photocatalysts for the construction of Z-scheme systems.

There are several types of Z-scheme system employing different electron mediators such as Fe<sup>3+/2+</sup> and [Co(bpy)<sub>3</sub>]<sup>3+/2+</sup> redox couples,<sup>16,28,29</sup> and reduced graphene oxide (RGO)<sup>30–32</sup> as shown in Fig. 1(a) and (b). Moreover, some Z-scheme systems work even without electron mediators, as shown in Fig. 1(c).<sup>33,34</sup> In this case, electron transfer proceeds through an interface contacted between the particles of the H<sub>2</sub>- and O<sub>2</sub>-evolving photocatalysts. So, it is important to examine the electron mediator in the Z-scheme system.

In the present paper, sacrificial H<sub>2</sub> and O<sub>2</sub> evolutions over Rh- and Ir-doped metal oxide photocatalysts under visible light irradiation were examined first to see the relationship between the photocatalytic properties and their band structures. Then, these H<sub>2</sub>- and O<sub>2</sub>-evolving photocatalysts were employed for various types of Z-scheme system for water splitting into H<sub>2</sub> and O<sub>2</sub> in stoichiometric amounts under visible light irradiation, without any sacrificial reagents. The photocatalytic performances for the sacrificial H<sub>2</sub> and O<sub>2</sub> evolutions and the Z-schematic water splitting were discussed based on the band structures and

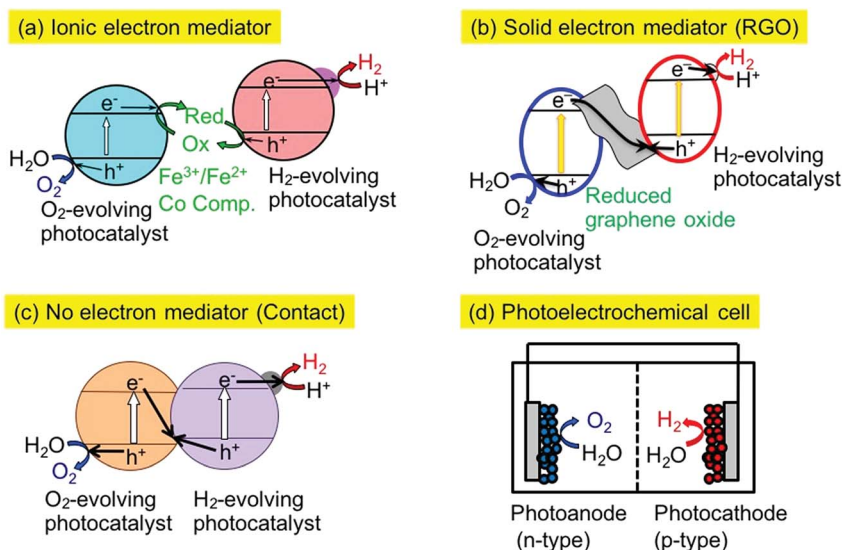


Fig. 1 Various types of powder material-based Z-scheme systems for water splitting.



photoelectrochemical properties of the photocatalyst materials such as p/n-types and the onset potentials of photocurrents.

## Experimental

### Preparation of Rh- or Ir-doped metal oxide photocatalysts

SrTiO<sub>3</sub>:Rh(1%),<sup>29</sup> SrTiO<sub>3</sub>:Ir(0.2%),<sup>22</sup> TiO<sub>2</sub>:Rh(x%),Sb(2x%) ( $x = 0.5$  or  $1.3$ ),<sup>25</sup> NaTaO<sub>3</sub>:Ir(1%),La(2%),<sup>24</sup> BaTa<sub>2</sub>O<sub>6</sub>:Ir(1%),La(2%)<sup>23</sup> and BiVO<sub>4</sub><sup>26,27</sup> were prepared by a solid-state reaction, a borate-flux method, and a liquid–solid reaction according to previous reports. In addition to them, NaNbO<sub>3</sub>:Rh(x%),Ba(y%) ( $x, y = (1.2, 1.44)$  or  $(1.0, 2.0)$ ) was newly prepared by a solid-state reaction. The starting materials, Na<sub>2</sub>CO<sub>3</sub> (Kanto Chemical; 99.5 or 99.8%), Nb<sub>2</sub>O<sub>5</sub> (Kanto Chemical; 99.99% or Kojundo Chemical; 99.99%), Rh<sub>2</sub>O<sub>3</sub> (Wako Chemical; 98%), and BaCO<sub>3</sub> (Kanto Chemical; 99%), were mixed at a molar ratio of Na/Nb/Rh/Ba = 1.05–1.05y : 1 – x : x : y. An excess of sodium was added in the starting materials to compensate for volatilization. The starting materials mixture was calcined at 1173 K for 1 h, and then 1423–1473 K for 10 h once or twice. The excess sodium was washed out with water after the calcination. The obtained powders had nonspecific shapes with aggregations, judging from the SEM images (Jeol; JSM-6700F) (Fig. S1†). The obtained samples were identified using X-ray diffraction (Rigaku; MiniFlex, Cu K $\alpha$ ). Diffuse reflectance spectra were obtained by a UV-vis-NIR spectrometer (JASCO, V-570) equipped with an integrator sphere and were converted to absorbance measurements *via* the Kubelka–Munk method.

### Preparation of an RGO–metal oxide composite

An RGO-incorporated O<sub>2</sub>-evolving photocatalyst was prepared by photocatalytic reduction of graphene oxide (GO) on the photocatalyst according to previous reports.<sup>30–32</sup> GO prepared by the Hummers' method<sup>35</sup> and the O<sub>2</sub>-evolving photocatalyst were dispersed in an aqueous methanol (Kanto Chemical; 99.8%) solution (50 vol%). The suspension was irradiated with visible light from a 300 W Xe lamp (PerkinElmer; CERMAX PE300BF) with a long pass filter (HOYA; L42) under a N<sub>2</sub> atmosphere with a pressure of 1 atm to obtain the RGO–photocatalyst composite. The methanol was carefully removed by washing with water. The RGO–photocatalyst composite was collected by filtration and was dried at room temperature in air.

### Sacrificial H<sub>2</sub> and O<sub>2</sub> evolutions (half reactions of water splitting)

H<sub>2</sub> and O<sub>2</sub> evolutions from aqueous solutions containing the sacrificial reagents CH<sub>3</sub>OH (Kanto Chemical; 99.8%) and AgNO<sub>3</sub> (Kojima Chemical; 99.9% or Toyo Chemical; 99.9%) that were half reactions of water splitting were examined using a top-irradiation reaction cell with a Pyrex window and a 300 W Xe lamp (PerkinElmer; CERMAX PE300BF). NaTaO<sub>3</sub>:Ir,La, BaTa<sub>2</sub>O<sub>6</sub>:Ir,La, NaNbO<sub>3</sub>:Rh,Ba, and TiO<sub>2</sub>:Rh,Sb were used as prepared, whereas SrTiO<sub>3</sub>:Ir without a cocatalyst was reduced under 1 atm of H<sub>2</sub> at 473 K for 2 h as a pretreatment for sacrificial O<sub>2</sub> evolution. The photocatalyst powders (0.1–0.3 g) were suspended in aqueous solutions (120–150 mL) and irradiated with visible light. For the H<sub>2</sub> evolution, Pt (0.3 wt%) cocatalyst, working as an H<sub>2</sub> evolution site, was loaded on photocatalysts by photodeposition from an aqueous methanol solution containing



$\text{H}_2\text{PtCl}_6$  (Tanaka Kikinzoku; 37.55% as Pt). The wavelength of the irradiation light was controlled to visible light using long-pass filters (HOYA; L42 and Y44). The amounts of evolved  $\text{H}_2$  and  $\text{O}_2$  were determined using an online gas chromatograph (Shimadzu; GC-8A, MS-5A column, TCD, Ar carrier).

### Z-schematic water splitting

Z-schematic water splitting was conducted using a gas-closed system with a top-irradiation cell with a Pyrex window.  $\text{H}_2$ -evolving photocatalyst and  $\text{O}_2$ -evolving photocatalyst powders (0.05 or 0.1 g, respectively) were suspended in 120 mL of water. For the interparticle Z-scheme systems without an electron mediator and with an RGO solid-state electron mediator, water not containing any ionic mediators was used. For the Z-scheme system with ionic mediator, an aqueous solution containing  $[\text{Co}(\text{bpy})_3]\text{SO}_4$  or  $\text{FeCl}_3$  as a mediator was used. The pH was adjusted with  $\text{H}_2\text{SO}_4$  in each of the Z-scheme systems with and without electron mediator, if necessary. Ru (0.7 wt%) cocatalyst, functioning as an  $\text{H}_2$  evolution site, was loaded on  $\text{SrTiO}_3\text{:Rh}$  (an  $\text{H}_2$ -evolving photocatalyst) by photodeposition from an aqueous methanol solution containing  $\text{RuCl}_3 \cdot n\text{H}_2\text{O}$  (Tanaka Kikinzoku; 36% as Ru in  $\text{RuCl}_3 \cdot n\text{H}_2\text{O}$ ). Pt (0.3–1 wt%) was loaded on the  $\text{NaTaO}_3\text{:Ir,L a}$  and  $\text{BaTa}_2\text{O}_6\text{:Ir,L a}$  ( $\text{H}_2$ -evolving photocatalysts), and  $\text{SrTiO}_3\text{:Ir}$  (an  $\text{O}_2$ -evolving photocatalyst) by an impregnation method. The photocatalyst powders and an aqueous  $\text{H}_2\text{PtCl}_6$  solution were placed in a porcelain crucible and dried on a hot plate. The  $\text{H}_2\text{PtCl}_6$ -impregnated  $\text{NaTaO}_3\text{:Ir,L a}$  and  $\text{BaTa}_2\text{O}_6\text{:Ir,L a}$  powders were calcined at 673 K for 2 h in air, whereas the  $\text{SrTiO}_3\text{:Ir}$  was not. The Pt-loaded  $\text{NaTaO}_3\text{:Ir,L a}$  and Pt-loaded  $\text{BaTa}_2\text{O}_6\text{:Ir,L a}$  were subsequently reduced at 673 K for 1 h under 1 atm of  $\text{H}_2$  as a pretreatment, while Pt-loaded  $\text{SrTiO}_3\text{:Ir}$  was reduced at 573 K for 1 h. The light source and GC setup were the same as those for the sacrificial  $\text{H}_2$  and  $\text{O}_2$  evolutions.

### Photoelectrochemical measurements

A squeegee method was used to prepare the  $\text{SrTiO}_3\text{:Rh}(1\%)$  photoelectrode and a drop-casting method was used for the  $\text{NaTaO}_3\text{:Ir}(1\%),\text{L a}(2\%)$  and  $\text{BaTa}_2\text{O}_6\text{:Ir}(1\%),\text{L a}(2\%)$  with and without  $\text{H}_2$ -reduction;  $\text{SrTiO}_3\text{:Ir}(0.2\%)$  with  $\text{H}_2$ -reduction;  $\text{NaNbO}_3\text{:Rh}(1\%),\text{B a}(2\%)$  and  $\text{TiO}_2\text{:Rh}(0.5\%),\text{S b}(1\%)$  without  $\text{H}_2$ -reduction, and  $\text{BiVO}_4$  photoelectrodes using powdered photocatalyst materials. For the  $\text{SrTiO}_3\text{:Rh}(1\%)$  photoelectrode, a paste consisting of 20 mg of  $\text{SrTiO}_3\text{:Rh}(1\%)$  photocatalyst powder, 20  $\mu\text{L}$  of acetylacetone (Kanto Chemical; 99.5%) and 40  $\mu\text{L}$  of distilled water was coated on an indium tin oxide transparent electrode (ITO).<sup>36</sup> For the other photoelectrodes, the photocatalyst powders were dispersed in ethanol (1–2  $\text{mg mL}^{-1}$ ) by sonication. The suspensions were drop-cast onto a fluorine-doped tin oxide transparent electrode (FTO) to obtain 1–2  $\text{mg cm}^{-2}$  of photocatalyst on the FTO. The  $\text{H}_2$ -reduced  $\text{SrTiO}_3\text{:Ir}$ -loaded FTO substrate was not calcined, whereas the other photocatalyst-loaded ITO and FTO substrates were calcined at 573–673 K for 2 h in air. The photoelectrochemical properties were evaluated with a three-electrode system consisting of working, Ag/AgCl reference, and Pt counter electrodes with a potentiostat (Hokuto Denko; HZ-series or HSV-110) using a conventional H-type cell with a Nafion membrane. The electrolyte was 0.1  $\text{mol L}^{-1}$   $\text{K}_2\text{SO}_4$ , 0.025  $\text{mol L}^{-1}$   $\text{KH}_2\text{PO}_4$  + 0.025  $\text{mol L}^{-1}$   $\text{Na}_2\text{HPO}_4$  pH buffer



was added, if necessary. A 300 W Xe lamp (PerkinElmer; CERMAX PE300BF) with a long-pass filter (HOYA; L42) was employed as a light source.

## Results

### Photocatalytic activities for sacrificial H<sub>2</sub> and O<sub>2</sub> evolutions over Rh- and Ir-doped metal oxide materials and their band structures

Sacrificial H<sub>2</sub> and O<sub>2</sub> evolutions of half reactions were carried out as test reactions of water splitting over Rh- and Ir-doped metal oxide photocatalysts using a sacrificial electron donor and acceptor to see the ability for photocatalytic H<sub>2</sub> or O<sub>2</sub> evolution, as shown in Table 1, prior to conducting water splitting.

The energy gaps were determined from the diffuse reflectance spectra and wavelength dependence of the photocatalytic activities as shown in Fig. 2 and 3. In general, trivalent and tetravalent Rh and Ir species are doped at the Ti<sup>4+</sup>, Nb<sup>5+</sup> and Ta<sup>5+</sup> sites in metal oxide materials.<sup>20–25,37,38</sup> Among the species, Rh<sup>3+</sup> and Ir<sup>3+</sup> contribute to the visible light response for metal oxide photocatalysts.<sup>20,24,37,38</sup> Sb<sup>5+</sup> was codoped with Rh<sup>3+</sup> at the Ti<sup>4+</sup> and Nb<sup>5+</sup> sites for charge compensation to enhance the formation of Rh<sup>3+</sup> and Ir<sup>3+</sup> and suppress the formation of Rh<sup>4+</sup> and Ir<sup>4+</sup> as efficient recombination centers between photogenerated electrons and holes, while Ba<sup>2+</sup> and La<sup>3+</sup> were replaced at the alkali and alkaline earth metal sites for the same purpose. SrTiO<sub>3</sub>:Ir was reduced with H<sub>2</sub> at 473 K to form the dopant Ir<sup>3+</sup>.

In the case of the Ir-doped photocatalysts, SrTiO<sub>3</sub>:Ir was active for both the sacrificial H<sub>2</sub> and O<sub>2</sub> evolutions,<sup>20,22</sup> whereas BaTa<sub>2</sub>O<sub>6</sub>:Ir,La and NaTaO<sub>3</sub>:Ir,La were active only for the sacrificial H<sub>2</sub> evolution.<sup>23,24</sup> BaTa<sub>2</sub>O<sub>6</sub>:Ir,La and NaTaO<sub>3</sub>:Ir,La were not active for sacrificial O<sub>2</sub> evolution even if they were reduced with H<sub>2</sub> at 673 K, as SrTiO<sub>3</sub>:Ir was, to reduce the Ir<sup>4+</sup> species. The Rh-doped photocatalysts

Table 1 Sacrificial H<sub>2</sub> and O<sub>2</sub> evolutions under visible light irradiation over Rh- or Ir-doped metal oxide photocatalysts<sup>a</sup>

Photocatalyst	Energy gap/eV	Incident light/nm	Activity/ μmol h <sup>-1</sup>		Ref.
			H <sub>2</sub> <sup>b</sup>	O <sub>2</sub> <sup>c</sup>	
Pt/NaTaO <sub>3</sub> :Ir(1%),La(2%)	2.1–2.3	λ > 420	3.2	—	24
NaTaO <sub>3</sub> :Ir(1%),La(2%)	2.1–2.3	λ > 420	—	0	24
Pt/BaTa <sub>2</sub> O <sub>6</sub> :Ir(1%),La(2%)	1.9–2.0	λ > 420	4.6	—	23
BaTa <sub>2</sub> O <sub>6</sub> :Ir(1%),La(2%)	1.9–2.0	λ > 420	—	0	23
Pt/SrTiO <sub>3</sub> :Ir(0.2%)	1.6–1.8	λ > 440	8.6	—	20
SrTiO <sub>3</sub> :Ir(0.2%) <sup>d</sup>	1.6–1.8	λ > 420	—	4.3	20
Pt/NaNbO <sub>3</sub> :Rh(1%),Ba(2%)	2.5	λ > 420	0.2	—	This work
NaNbO <sub>3</sub> :Rh(1%),Ba(2%)	2.5	λ > 420	—	5.3	This work
Pt/TiO <sub>2</sub> :Rh(0.5%),Sb(1%)	2.1	λ > 440	0	—	This work
TiO <sub>2</sub> :Rh(0.5%),Sb(1%)	2.1	λ > 440	—	7.5	25

<sup>a</sup> Photocatalyst: 0.1–0.3 g; light source: 300 W Xe lamp with long-pass filters (λ > 420 nm or λ > 440 nm); reaction cell: top-irradiation cell with a Pyrex window. <sup>b</sup> Cocatalyst: Pt (0.3 wt%, photodeposition); reactant solution: 10 vol% aqueous methanol solution. <sup>c</sup> Cocatalyst: none; reactant solution: 0.02–0.05 mol L<sup>-1</sup> aqueous AgNO<sub>3</sub> solution (120–150 mL).

<sup>d</sup> Treatment: H<sub>2</sub>-reduction at 473 K for 2 h.



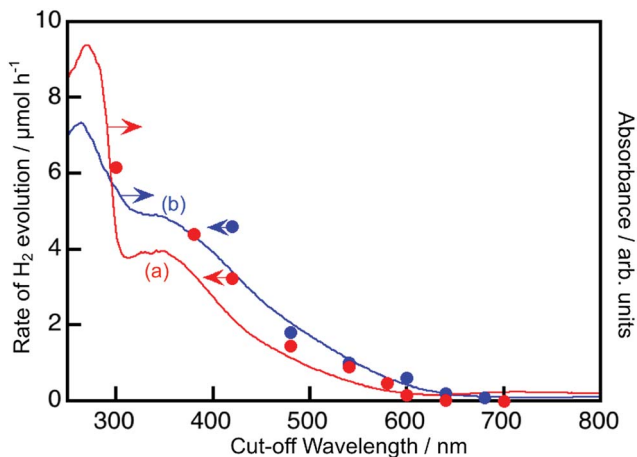


Fig. 2 The wavelength dependence of H<sub>2</sub> evolution from a 10 vol% aqueous methanol solution (closed circles) and diffuse reflectance spectra (solid line) of (a) NaTaO<sub>3</sub>:Ir(1%),La(2%) and (b) BaTa<sub>2</sub>O<sub>6</sub>:Ir(1%),La(2%). Photocatalyst: 0.1 g, cocatalyst: Pt (photo-deposition) for H<sub>2</sub> evolution, reactant solution: 120 mL, light source: 300 W Xe lamp with long-pass filters, and reaction cell: top-irradiation cell with a Pyrex window. Samples of diffuse reflectance spectra were reduced at 473 K.

TiO<sub>2</sub>:Rh,Sb<sup>25</sup> and NaNbO<sub>3</sub>:Rh,Ba were active for sacrificial O<sub>2</sub> evolution using Ag<sup>+</sup> as an electron acceptor. NaNbO<sub>3</sub>:Rh,Ba showed very low activity for the sacrificial H<sub>2</sub> evolution. These properties will be discussed based on the band structure in the Discussion section. These results of the sacrificial H<sub>2</sub> and O<sub>2</sub> evolutions suggest that NaTaO<sub>3</sub>:Ir,La and BaTa<sub>2</sub>O<sub>6</sub>:Ir,La can be used as H<sub>2</sub>-evolving

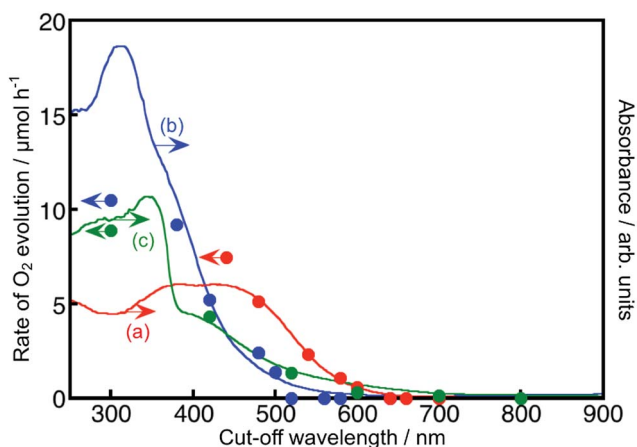


Fig. 3 The wavelength dependence of O<sub>2</sub> evolution from a 0.05 mol L<sup>-1</sup> aqueous silver nitrate solution (closed circles) and diffuse reflectance spectra (solid line) of (a) TiO<sub>2</sub>:Rh(1.3%),Sb(2.6%), (b) NaNbO<sub>3</sub>:Rh(1.2%),Ba(1.44%) and (c) SrTiO<sub>3</sub>:Ir(0.2%) with H<sub>2</sub> reduction at 473 K. Photocatalyst: 0.1–0.3 g, reactant solution: 150 mL, light source: 300 W Xe lamp with long-pass filters, and reaction cell: top-irradiation cell with a Pyrex window.



photocatalysts for the construction of Z-scheme systems, while SrTiO<sub>3</sub>:Ir, NaNbO<sub>3</sub>:Rh,Ba and TiO<sub>2</sub>:Rh,Sb are expected to be employed as O<sub>2</sub>-evolving photocatalysts.

Fig. 2 and 3 show the diffuse reflectance spectra and wavelength dependence of the photocatalytic H<sub>2</sub> and O<sub>2</sub> evolutions of the Rh- and Ir-doped metal oxide photocatalysts in the presence of sacrificial reagents. The wavelengths were controlled with long-pass filters. It is vital to see the wavelength dependency of the photocatalytic activity because it is not guaranteed that photocatalysts with visible light absorption bands always give activities under visible light irradiation. The onsets of the wavelength dependence agreed with those of the diffuse reflection spectra. The onset wavelengths for the H<sub>2</sub> evolutions were 640 and 600 nm for BaTa<sub>2</sub>O<sub>6</sub>:Ir,La and NaTaO<sub>3</sub>:Ir,La, respectively. These onset wavelengths were longer than the 540 nm of SrTiO<sub>3</sub>:Rh (a conventional H<sub>2</sub>-evolving photocatalyst). The onset wavelengths for O<sub>2</sub> evolution were 600, 500 and 700 nm for TiO<sub>2</sub>:Rh,Sb, NaNbO<sub>3</sub>:Rh,Ba and SrTiO<sub>3</sub>:Ir, respectively. It is noteworthy that TiO<sub>2</sub>:Rh,Sb and SrTiO<sub>3</sub>:Ir responded at longer wavelengths than the BiVO<sub>4</sub> (BG: 2.4 eV) (a conventional O<sub>2</sub>-evolving photocatalyst).

Fig. 4 shows the band structures of the Rh- and Ir-doped metal oxide photocatalysts. The impurity levels of Rh<sup>3+</sup> and Ir<sup>3+</sup> were estimated from the energy gaps determined by diffuse reflection spectra supposing that the valence band consisting of O 2p located at +3 V vs. NHE at a pH of 0.<sup>39</sup> The absorption bands in the visible light region shown in Fig. 2 and 3 are due to electronic transition from the impurity levels consisting of Rh<sup>3+</sup> and Ir<sup>3+</sup> to the conduction bands of the host materials. The impurity levels formed with electron-filled orbitals of Ir<sup>3+</sup> were around 1.0–1.2 V for NaTaO<sub>3</sub> and BaTa<sub>2</sub>O<sub>6</sub>, while Ir<sup>3+</sup> in SrTiO<sub>3</sub> formed an impurity level around 1.4–1.6 V that was deeper than those in the cases of the tantalates. The energy levels formed with electron-filled orbitals of Rh<sup>3+</sup> located around 2.0–2.1 V that were similar to those of SrTiO<sub>3</sub>:Rh and SrTiO<sub>3</sub>:Rh,Sb.<sup>20,21</sup> The reason Rh<sup>3+</sup> forms a deeper impurity level than Ir<sup>3+</sup> is due to Ir<sup>4+</sup> being more stable than Rh<sup>4+</sup> in metal oxides. Therefore, electronic transition from the Ir<sup>3+</sup> impurity level to a conduction band is easier than that from Rh<sup>3+</sup>, resulting in the formation of the shallow impurity level by Ir<sup>3+</sup>.

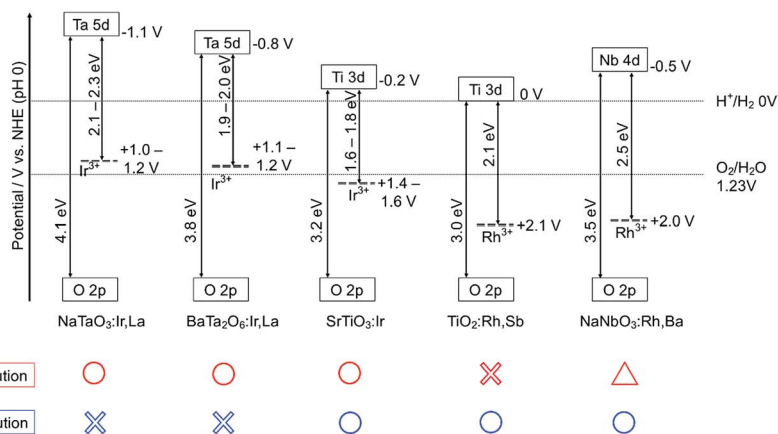


Fig. 4 The band structures of Rh- or Ir-doped metal oxide photocatalysts at pH 0.





## Z-schematic systems for photocatalytic water splitting employing Rh- and Ir-doped metal oxide materials

The combination of SrTiO<sub>3</sub>:Rh and BiVO<sub>4</sub> photocatalysts can be a benchmark of a Z-scheme photocatalyst system. The Z-schematic water splitting proceeds using Fe<sup>3+/2+</sup> and [Co(bpy)<sub>3</sub>]<sup>3+/2+</sup> redox couples as ionic mediators (Fig. 1(a))<sup>16,28,29</sup> and also *via* interparticle electron transfer between SrTiO<sub>3</sub>:Rh and BiVO<sub>4</sub> particles with and without RGO (Fig. 1(b) and (c)).<sup>30,33,34</sup> In general, water splitting *via* Z-schematic interparticle electron transfer by contact between the particles of the H<sub>2</sub>- and O<sub>2</sub>-evolving photocatalysts with and without RGO can be achieved (Fig. 1(b) and (c)) when the H<sub>2</sub>- and O<sub>2</sub>-evolving photocatalysts satisfy the following two requirements; (i) H<sub>2</sub>- and O<sub>2</sub>-evolving photocatalysts possess p- and n-type semiconductor properties, respectively, (ii) there is a certain electrode potential at which the cathodic photocurrent of the p-type semiconductor overlaps with the anodic photocurrent of the n-type semiconductor, being similar to water splitting using a photoelectrochemical cell working with no applied external bias as shown in Fig. 1(d).<sup>31,32</sup> Moreover, H<sub>2</sub>- and O<sub>2</sub>-evolving photocatalysts must have contact with each other for interparticle electron transfer. In contrast to this, a Z-scheme system employing an ionic electron mediator could work regardless of the p- or n-type properties of the H<sub>2</sub>- or O<sub>2</sub>-evolving photocatalysts, if the photocatalysts have potentials for the reduction or oxidation of redox couple ionic mediators and the adsorption abilities for the redox couples. The SrTiO<sub>3</sub>:Rh and BiVO<sub>4</sub> photocatalysts satisfy these factors resulting in all of the Z-scheme systems showing activities for water splitting into H<sub>2</sub> and O<sub>2</sub> in stoichiometric amounts without any sacrificial reagents (Fig. 1(a)–(d)).

Various types of Z-scheme photocatalyst systems employing Ir- and Rh-doped photocatalysts for water splitting into H<sub>2</sub> and O<sub>2</sub> under visible light irradiation are shown in Table 2. Here, the absolute evaluation of the performance of photocatalyst systems for water splitting is how much H<sub>2</sub> and O<sub>2</sub> is obtained under certain experimental conditions. In this sense, the activities of the different Z-scheme systems in Table 2 are comparable with each other, but not with those of Table 1, because of the almost identical experimental conditions for the water splitting, even if the kinetics would be different among the systems.

For the construction of Z-scheme systems, BiVO<sub>4</sub> (an O<sub>2</sub>-evolving photocatalyst) was combined with BaTa<sub>2</sub>O<sub>6</sub>:Ir,La and NaTaO<sub>3</sub>:Ir,La (H<sub>2</sub>-evolving photocatalysts), while SrTiO<sub>3</sub>:Rh (an H<sub>2</sub>-evolving photocatalyst) was combined with TiO<sub>2</sub>:Rh,Sb, NaNbO<sub>3</sub>:Rh,Ba and SrTiO<sub>3</sub>:Ir (O<sub>2</sub>-evolving photocatalysts) as suggested by their H<sub>2</sub> and O<sub>2</sub> evolution abilities. In addition to the suspension system, the photoelectrochemical properties of photocatalyst powders immobilized on a conducting substrate as shown in Fig. 1(d) were examined using a Pt counter electrode without any sacrificial reagents to see the cathodic or anodic photocurrent and the onset potentials of semiconductor properties, not as a water splitting device, as shown in Fig. 5, in order to consider the potential overlap for Z-schematic water splitting *via* interparticle electron transfer with and without RGO. It has been reported that SrTiO<sub>3</sub>:Rh<sup>36,40</sup> and BiVO<sub>4</sub><sup>32,41–44</sup> function as a photocathode and photoanode, respectively. NaTaO<sub>3</sub>:Ir,La showed only an anodic photocurrent. Although BaTa<sub>2</sub>O<sub>6</sub>:Ir,La and SrTiO<sub>3</sub>:Ir showed cathodic and anodic photocurrents, the photocurrents were poor and the onset potentials shifted with the sweeping direction of the CV curves. This result implies that





Table 2 Z-schematic water splitting under visible light irradiation using Rh- or Ir-doped metal oxide photocatalysts<sup>a</sup>

H <sub>2</sub> photocat.	O <sub>2</sub> photocat.	Mediator	Initial pH	Activity/ $\mu\text{mol h}^{-1}$		Ref.
				H <sub>2</sub>	O <sub>2</sub>	
Pt/NaNbO <sub>3</sub> :Ir(1%),La(2%) <sup>b</sup>	BiVO <sub>4</sub>	None <sup>e</sup>	4.2	Trace	Trace	This work
Pt/NbTaO <sub>3</sub> :Ir(1%),La(2%) <sup>b</sup>	BiVO <sub>4</sub>	RGO <sup>e</sup>	4.2	Trace	Trace	This work
Pt/NbTaO <sub>3</sub> :Ir(1%),La(2%) <sup>b</sup>	BiVO <sub>4</sub>	Fe <sup>3+/2+</sup> <sup>f</sup>	2.4	Trace	Trace	This work
Pt/NbTaO <sub>3</sub> :Ir(1%),La(2%) <sup>b</sup>	BiVO <sub>4</sub>	[Co(bpy) <sub>3</sub> ] <sup>3+/2+</sup> <sup>g</sup>	4.2	0.8	0.3	This work
Pt/BaTa <sub>2</sub> O <sub>6</sub> :Ir(1%),La(2%) <sup>b</sup>	BiVO <sub>4</sub>	None <sup>e</sup>	4.2	0.9	0.4	This work
Pt/BaTa <sub>2</sub> O <sub>6</sub> :Ir(1%),La(2%) <sup>b</sup>	BiVO <sub>4</sub>	RGO <sup>e</sup>	4.2	0.9	0.4	This work
Pt/BaTa <sub>2</sub> O <sub>6</sub> :Ir(1%),La(2%) <sup>b</sup>	BiVO <sub>4</sub>	Fe <sup>3+/2+</sup> <sup>f</sup>	2.4	0.1	0.4	This work
Pt/BaTa <sub>2</sub> O <sub>6</sub> :Ir(1%),La(2%) <sup>b</sup>	BiVO <sub>4</sub>	[Co(bpy) <sub>3</sub> ] <sup>3+/2+</sup> <sup>g</sup>	4.2	5.9	2.1	22
Ru/SrTiO <sub>3</sub> :Rh(1%) <sup>c</sup>	Pt/SrTiO <sub>3</sub> :Ir(0.2%) <sup>d</sup>	None <sup>e</sup>	3.5	0.4	Trace	This work
Ru/SrTiO <sub>3</sub> :Rh(1%) <sup>c</sup>	Pt/SrTiO <sub>3</sub> :Ir(0.2%) <sup>d</sup>	RGO <sup>e</sup>	3.5	0.8	Trace	This work
Ru/SrTiO <sub>3</sub> :Rh(1%) <sup>c</sup>	Pt/SrTiO <sub>3</sub> :Ir(0.2%) <sup>d</sup>	Fe <sup>3+/2+</sup> <sup>f</sup>	2.4	2.5	1.0	This work
Ru/SrTiO <sub>3</sub> :Rh(1%) <sup>c</sup>	Pt/SrTiO <sub>3</sub> :Ir(0.2%) <sup>d</sup>	[Co(bpy) <sub>3</sub> ] <sup>3+/2+</sup> <sup>g</sup>	7.8	5.1	2.4	This work
Ru/SrTiO <sub>3</sub> :Rh(1%) <sup>c</sup>	NbNbO <sub>3</sub> :Rh(1%),Ba(2%)	None <sup>e</sup>	2.4	2.1	1.1	This work
Ru/SrTiO <sub>3</sub> :Rh(1%) <sup>c</sup>	NbNbO <sub>3</sub> :Rh(1%),Ba(2%)	Fe <sup>3+/2+</sup> <sup>f</sup>	2.4	5.7	2.3	This work
Ru/SrTiO <sub>3</sub> :Rh(1%) <sup>c</sup>	NbNbO <sub>3</sub> :Rh(1%),Ba(2%)	[Co(bpy) <sub>3</sub> ] <sup>3+/2+</sup> <sup>g</sup>	3.8	4.1	1.7	This work
Ru/SrTiO <sub>3</sub> :Rh(1%) <sup>c</sup>	TiO <sub>2</sub> :Rh(0.5%),Sb(1%)	None <sup>e</sup>	3.5	5.4	2.4	This work
Ru/SrTiO <sub>3</sub> :Rh(1%) <sup>c</sup>	TiO <sub>2</sub> :Rh(0.5%),Sb(1%)	Fe <sup>3+/2+</sup> <sup>f</sup>	2.4	12	5.3	This work
Ru/SrTiO <sub>3</sub> :Rh(1%) <sup>c</sup>	TiO <sub>2</sub> :Rh(0.5%),Sb(1%)	[Co(bpy) <sub>3</sub> ] <sup>3+/2+</sup> <sup>g</sup>	3.5	28	13	This work

<sup>a</sup> Photocatalyst: 0.05 or 0.1 g each; light source: 300 W Xe lamp with a long-pass filter ( $\lambda > 420$  nm); reaction cell: top-irradiation cell with a Pyrex window. <sup>b</sup> Cocatalyst: Pt (0.3 wt%), impregnation at 673 K for 2 h and subsequent H<sub>2</sub>-reduction at 673 K for 1 h). <sup>c</sup> Cocatalyst: Ru (0.7 wt%, photodeposition). <sup>d</sup> Cocatalyst: Pt (1 wt%), impregnation without calcination and subsequent H<sub>2</sub>-reduction at 573 K for 1 h). <sup>e</sup> Reactant solution: H<sub>2</sub>SO<sub>4</sub> solution. <sup>f</sup> Reactant solution: 2 mmol L<sup>-1</sup> FeCl<sub>3</sub> solution. <sup>g</sup> Reactant solution: 0.02 mmol L<sup>-1</sup>. <sup>h</sup> Reactant solution: 0.5 mmol L<sup>-1</sup> [Co(bpy)<sub>3</sub>]SO<sub>4</sub> solution (120 mL).

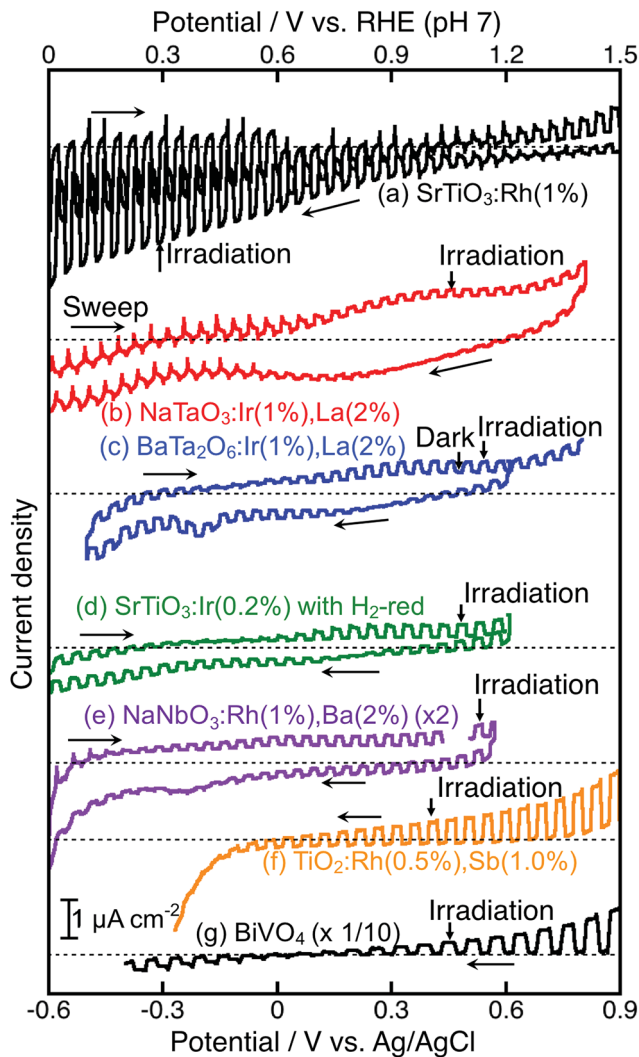


Fig. 5 Current vs. potential curves of (a)  $\text{SrTiO}_3\text{:Rh}(1\%)$ , (b)  $\text{NaTaO}_3\text{:Ir}(1\%),\text{La}(2\%)$ , (c)  $\text{BaTa}_2\text{O}_6\text{:Ir}(1\%),\text{La}(2\%)$ , (d)  $\text{SrTiO}_3\text{:Ir}(0.2\%)$  with  $\text{H}_2$  reduction, (e)  $\text{NaNbO}_3\text{:Rh}(1\%),\text{Ba}(2\%)$  (x2), (f)  $\text{TiO}_2\text{:Rh}(0.5\%),\text{Sb}(1.0\%)$ , and (g)  $\text{BiVO}_4$  (x 1/10). Electrolyte:  $0.1 \text{ mol L}^{-1} \text{ K}_2\text{SO}_4$  aqueous solution (pH 7, phosphate buffer was added if necessary), light source: 300 W Xe lamp ( $\lambda > 420 \text{ nm}$ ).

those photocurrents might not be due to  $\text{H}_2$  and  $\text{O}_2$  evolutions, but possibly due to redox reactions of the doped Ir species. In contrast,  $\text{TiO}_2\text{:Rh,Sb}$ ,<sup>45</sup> and  $\text{NaNbO}_3\text{:Rh,Ba}$  gave clear anodic photocurrents indicating an n-type semiconductor character. The onset potential of  $\text{NaNbO}_3\text{:Rh,Ba}$  was more negative than that of  $\text{TiO}_2\text{:Rh,Sb}$ , whereas the anodic photocurrent of  $\text{NaNbO}_3\text{:Rh,Ba}$  was much smaller than that of  $\text{TiO}_2\text{:Rh,Sb}$ . These anodic photocurrents overlapped with the cathodic photocurrent of  $\text{SrTiO}_3\text{:Rh}$  at a certain electrode potential.

The Z-scheme systems were classified into three groups depending on the combination of  $\text{H}_2$ - and  $\text{O}_2$ -evolving photocatalysts and an electron mediator;



being active for only a  $[\text{Co}(\text{bpy})_3]^{3+/2+}$  redox couple, active for  $[\text{Co}(\text{bpy})_3]^{3+/2+}$  and  $\text{Fe}^{3+/2+}$  redox couple ionic mediators, active not only for  $[\text{Co}(\text{bpy})_3]^{3+/2+}$  and  $\text{Fe}^{3+/2+}$  redox couples but also *via* interparticle electron transfer with and without RGO.  $\text{NaTaO}_3:\text{Ir},\text{La} + \text{BiVO}_4$  and  $\text{BaTa}_2\text{O}_6:\text{Ir},\text{La} + \text{BiVO}_4$  were active when not  $\text{Fe}^{3+/2+}$  but a  $[\text{Co}(\text{bpy})_3]^{3+/2+}$  redox couple was used. Although the  $\text{BaTa}_2\text{O}_6:\text{Ir},\text{La} + \text{BiVO}_4$  showed activities *via* interparticle electron transfer at a pH of 4.2 with and without RGO, the activities were smaller than that with the ionic electron mediator  $[\text{Co}(\text{bpy})_3]^{3+/2+}$ .  $\text{SrTiO}_3:\text{Rh} + \text{SrTiO}_3:\text{Ir}$  was active when the  $[\text{Co}(\text{bpy})_3]^{3+/2+}$  and  $\text{Fe}^{3+/2+}$  redox couple ionic mediators were used, whereas it was not active *via* interparticle electron transfer with and without RGO at pH 3.5.  $\text{SrTiO}_3:\text{Rh} + \text{NaNbO}_3:\text{Rh},\text{Ba}$  and  $\text{SrTiO}_3:\text{Rh} + \text{TiO}_2:\text{Rh},\text{Sb}$  were active for the  $[\text{Co}(\text{bpy})_3]^{3+/2+}$  and  $\text{Fe}^{3+/2+}$  redox couples and *via* interparticle electron transfer without RGO. These Z-scheme systems with RGO using  $\text{NaNbO}_3:\text{Rh},\text{Ba}$  and  $\text{TiO}_2:\text{Rh},\text{Sb}$  were not examined in detail, because the GOs were not suitably photoreduced on the  $\text{NaNbO}_3:\text{Rh},\text{Ba}$  and  $\text{TiO}_2:\text{Rh},\text{Sb}$  due to their poor reducing activities, as expected from their poor  $\text{H}_2$  evolution activities as shown in Table 1.  $\text{SrTiO}_3:\text{Rh} + \text{TiO}_2:\text{Rh},\text{Sb}$  showed the best performances among the Z-scheme systems in Table 2.

## Discussion

$\text{H}_2$  and  $\text{O}_2$  evolution abilities (shown in Table 1) can be considered based on the band structure, as shown in Fig. 4. Photogenerated holes in the  $\text{Ir}^{3+}$  levels in  $\text{NaTaO}_3:\text{Ir},\text{La}$  and  $\text{BaTa}_2\text{O}_6:\text{Ir},\text{La}$  have no potentials for water oxidation. In contrast, the  $\text{Ir}^{3+}$  level of  $\text{SrTiO}_3:\text{Ir}$  possesses the water oxidation potential, though the driving force is not so large. Conduction bands consisting of Ta 5d orbitals in  $\text{BaTa}_2\text{O}_6:\text{Ir},\text{La}$  and  $\text{NaTaO}_3:\text{Ir},\text{La}$  possess thermodynamically enough potentials for water reduction to form  $\text{H}_2$ . On the other hand, the photogenerated holes in the  $\text{Rh}^{3+}$  levels in  $\text{NaNbO}_3:\text{Rh},\text{Ba}$  and  $\text{TiO}_2:\text{Rh},\text{Sb}$  have enough potential for water oxidation to form  $\text{O}_2$ . These energy levels were consistent with the  $\text{H}_2$  and  $\text{O}_2$  evolution abilities, as shown in Table 1. Of course, other kinetic factors of the active sites would exist for the  $\text{H}_2$  and  $\text{O}_2$  evolution abilities in addition to the energy levels of the thermodynamic factor.

Let us consider the reason why the Z-scheme systems shown in Table 2 are classified into three groups, based on the band structure and photoelectrochemical properties. Fig. 6 shows energy diagrams for the Z-schemes employing  $\text{Fe}^{3+/2+}$  and  $[\text{Co}(\text{bpy})_3]^{3+/2+}$  redox couples at pH 2.4 and 4.0, respectively. It is assumed that the band levels shift with  $-0.059 \text{ V pH}^{-1}$  because those materials are metal oxides.

$\text{NaTaO}_3:\text{Ir},\text{La} + \text{BiVO}_4$  and  $\text{BaTa}_2\text{O}_6:\text{Ir},\text{La} + \text{BiVO}_4$  were active with a  $[\text{Co}(\text{bpy})_3]^{3+/2+}$  redox couple at pH 4.2, but not with  $\text{Fe}^{3+/2+}$  at pH 2.4. Photogenerated holes in the  $\text{Ir}^{3+}$  levels in  $\text{NaTaO}_3:\text{Ir},\text{La}$  and  $\text{BaTa}_2\text{O}_6:\text{Ir},\text{La}$  (an  $\text{H}_2$ -evolving photocatalyst) possess small driving forces for oxidation of  $\text{Fe}^{2+}$  (an electron mediator) at pH 2.4 as shown in Fig. 6(a), whereas they possess enough potential for oxidation of  $[\text{Co}(\text{bpy})_3]^{2+}$  at pH 4.2 as shown in Fig. 6(b). The activities from the interparticle electron transfer were very small or negligible because of poor photoresponse and poor overlaps of the potentials giving cathodic photocurrents of  $\text{NaTaO}_3:\text{Ir},\text{La}$  and  $\text{BaTa}_2\text{O}_6:\text{Ir},\text{La}$  and an anodic photocurrent of  $\text{BiVO}_4$  as shown in Fig. 5, and also may be due to poor contact between  $\text{NaTaO}_3:\text{Ir},\text{La}$  or  $\text{BaTa}_2\text{O}_6:\text{Ir},\text{La}$  (an  $\text{H}_2$ -evolving photocatalyst) and  $\text{BiVO}_4$  (an  $\text{O}_2$ -



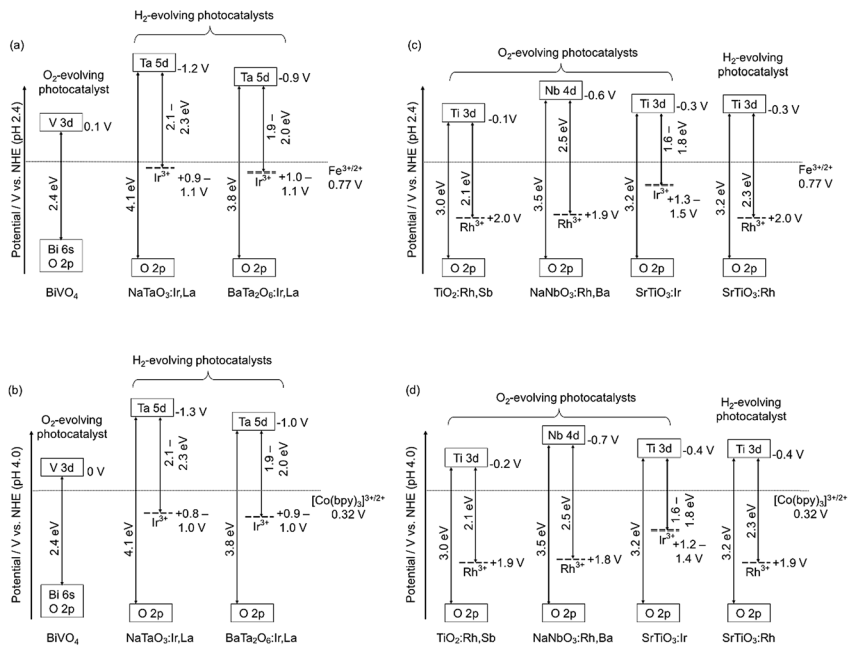


Fig. 6 The band structures of Rh- or Ir-doped metal oxide photocatalysts, and redox potentials of Fe and Co-complex ionic electron mediators at pH 2.4 and 4.0. (a) and (b) represent Z-scheme systems of Ir-doped metal oxides as H<sub>2</sub>-evolving photocatalysts combined with BiVO<sub>4</sub>. (c) and (d) represent Z-scheme systems of Rh-doped metal oxides and SrTiO<sub>3</sub>:Ir as O<sub>2</sub>-evolving photocatalysts combined with SrTiO<sub>3</sub>:Rh.

evolving photocatalyst) in the suspension at pH 4.2. In contrast to them, SrTiO<sub>3</sub>:Rh + SrTiO<sub>3</sub>:Ir was active when the [Co(bpy)<sub>3</sub>]<sup>3+/2+</sup> and Fe<sup>3+/2+</sup> redox couple ionic mediators were used, because the conduction band of SrTiO<sub>3</sub>:Ir (an O<sub>2</sub>-evolving photocatalyst) has enough potential for the reduction of Fe<sup>3+</sup> and [Co(bpy)<sub>3</sub>]<sup>3+</sup> (Fig. 6(c) and (d)). However, because the anodic photocurrent of SrTiO<sub>3</sub>:Ir hardly overlapped with the cathodic photocurrent of SrTiO<sub>3</sub>:Rh, the activities by interparticle electron transfer with and without RGO were negligible. The conduction bands of TiO<sub>2</sub>:Rh,Sb and NaNbO<sub>3</sub>:Rh,Ba (O<sub>2</sub>-evolving photocatalysts) also possessed enough potential for the reduction of Fe<sup>3+</sup> and [Co(bpy)<sub>3</sub>]<sup>3+</sup> as well as SrTiO<sub>3</sub>:Ir (Fig. 6(c) and (d)). Moreover, in these cases, the anodic photocurrents of the TiO<sub>2</sub>:Rh,Sb and NaNbO<sub>3</sub>:Rh,Ba photoelectrodes overlapped enough with the cathodic photocurrent of SrTiO<sub>3</sub>:Rh. So, it is reasonable that SrTiO<sub>3</sub>:Rh + NaNbO<sub>3</sub>:Rh,Ba and SrTiO<sub>3</sub>:Rh + TiO<sub>2</sub>:Rh,Sb were active not only for the [Co(bpy)<sub>3</sub>]<sup>3+/2+</sup> and Fe<sup>3+/2+</sup> redox couples but also *via* interparticle electron transfer with and without RGO. The activity of SrTiO<sub>3</sub>:Rh + TiO<sub>2</sub>:Rh,Sb was higher than that of SrTiO<sub>3</sub>:Rh + NaNbO<sub>3</sub>:Rh,Ba for all of the types of Z-scheme system. In the cases with the use of ionic electron mediators, it is probably due to the higher activity for O<sub>2</sub> evolution and the narrower energy gaps of TiO<sub>2</sub>:Rh,Sb than those of NaNbO<sub>3</sub>:Rh,Ba, as shown in Table 1. The reason why SrTiO<sub>3</sub>:Rh + TiO<sub>2</sub>:Rh,Sb showed a higher activity than SrTiO<sub>3</sub>:Rh + NaNbO<sub>3</sub>:Rh,Ba *via* interparticle electron transfer is that TiO<sub>2</sub>:Rh,Sb gave much larger anodic photocurrents than NaNbO<sub>3</sub>:Rh,Ba, as shown in Fig. 5.



## Conclusions

We have successfully developed Z-scheme photocatalyst systems for water splitting under visible light irradiation employing Rh- and Ir-doped metal oxide photocatalysts with longer wavelength responses than conventional SrTiO<sub>3</sub>:Rh and BiVO<sub>4</sub>. The impurity levels of doped Ir<sup>3+</sup> and Rh<sup>3+</sup> that contributed to the visible light responses for the photocatalytic reactions were determined from diffuse reflectance spectra and supposing +3.0 V vs. NHE of a valence band of the oxide photocatalyst materials. The impurity levels of Ir<sup>3+</sup> in NaTaO<sub>3</sub>:Ir,La and BaTa<sub>2</sub>O<sub>6</sub>:Ir,La have sufficient potentials for the oxidation of the electron mediator [Co(bpy)<sub>3</sub>]<sup>2+</sup>, but insufficient potentials for water oxidation to form O<sub>2</sub> and oxidation of the electron mediator Fe<sup>2+</sup>. The conduction bands of the NaTaO<sub>3</sub>:Ir,La and BaTa<sub>2</sub>O<sub>6</sub>:Ir,La have enough potential for water reduction to form H<sub>2</sub>. Therefore, the NaTaO<sub>3</sub>:Ir,La and BaTa<sub>2</sub>O<sub>6</sub>:Ir,La could be used as H<sub>2</sub>-evolving photocatalysts only when a [Co(bpy)<sub>3</sub>]<sup>3+/2+</sup> redox couple was used. The impurity levels of Ir<sup>3+</sup> in SrTiO<sub>3</sub>:Ir and Rh<sup>3+</sup> in TiO<sub>2</sub>:Rh,Sb and NaNbO<sub>3</sub>:Rh,Ba have potentials for water oxidation and their conduction bands possess the potentials for reduction of the electron mediators Fe<sup>3+</sup> and [Co(bpy)<sub>3</sub>]<sup>3+</sup>. This property means that SrTiO<sub>3</sub>:Ir, TiO<sub>2</sub>:Rh,Sb and NaNbO<sub>3</sub>:Rh,Ba could be employed as O<sub>2</sub>-evolving photocatalysts for the construction of Z-scheme systems employing ionic electron mediators. Moreover, photoelectrochemical measurements using photocatalyst powders immobilized on conducting substrate revealed that the n-type characters and relatively negative onset potentials of the TiO<sub>2</sub>:Rh,Sb and NaNbO<sub>3</sub>:Rh,Ba photoanodes enabled the Z-scheme systems to work by interparticle electron transfer. Although the efficiencies of the present Z-scheme systems are low at the present stage, these will be improved by interfacial controls from a kinetic point of view in basic research. These results and discussion will contribute to the design of a highly active photocatalyst system for water splitting into H<sub>2</sub> and O<sub>2</sub>, aiming for the demonstration of actual solar water splitting using a suitable reactor.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

This work was supported by JSPS KAKENHI Grant Numbers 17H06440 and 17H06433 in Scientific Research on Innovative Areas "Innovations for Light-Energy Conversion (I<sup>4</sup>LEC)".

## References

- 1 A. Kudo, H. Kato and I. Tsuji, *Chem. Lett.*, 2004, **33**, 1534.
- 2 A. Kudo and Y. Miseki, *Chem. Soc. Rev.*, 2009, **38**, 253.
- 3 F. E. Osterloh, *Chem. Mater.*, 2008, **20**, 35.
- 4 Y. Inoue, *Energy Environ. Sci.*, 2009, **2**, 364.
- 5 R. Abe, *J. Photochem. Photobiol., C*, 2010, **11**, 179.
- 6 T. Hisatomi, J. Kubota and K. Domen, *Chem. Soc. Rev.*, 2014, **43**, 7520.



- 7 K. Maeda and K. Domen, *Bull. Chem. Soc. Jpn.*, 2016, **89**, 627.
- 8 B. A. Pinaud, J. D. Benck, L. C. Seitz, A. J. Forman, Z. Chen, T. G. Deutsch, B. D. James, K. N. Baum, G. N. Baum, S. Ardo, H. Wang, E. Miller and T. F. Jaramillo, *Energy Environ. Sci.*, 2013, **6**, 1983.
- 9 D. M. Fabian, S. Hu, N. Singh, F. A. Houle, T. Hisatomi, K. Domen, F. E. Osterloh and S. Ardo, *Energy Environ. Sci.*, 2015, **8**, 2825.
- 10 T. Yamada and K. Domen, *Chem. Eng.*, 2018, **2**, 36.
- 11 A. Kudo, in "Photocatalysis", *Contemporary Catalysis; Science, Technology and Applications*, ed. P. Kamer, D. Vogt and J. Thybaut, Royal Society of Chemistry, Cambridge, 2017, ch. 3.7, pp. 326–343.
- 12 A. J. Bard, *J. Photochem.*, 1979, **10**, 59.
- 13 A. Kudo, *MRS Bull.*, 2011, **36**, 32.
- 14 K. Maeda, *ACS Catal.*, 2013, **3**, 1486.
- 15 Y. Wang, H. Suzuki, J. Xie, O. Tomita, D. J. Martin, M. Higashi, D. Kong, R. Abe and J. Tang, *Chem. Rev.*, 2018, **118**, 5201.
- 16 H. Kato, M. Hori, R. Konta, Y. Shimodaira and A. Kudo, *Chem. Lett.*, 2004, **33**, 1348.
- 17 Q. Wang, T. Hisatomi, Q. Jia, H. Tokudome, M. Zhong, C. Wang, Z. Pan, T. Takata, M. Nakabayashi, N. Shibata, Y. Li, I. D. Sharp, A. Kudo, T. Yamada and K. Domen, *Nat. Mater.*, 2016, **15**, 611.
- 18 Q. Wang, T. Hisatomi, Y. Suzuki, Z. Pan, J. Seo, M. Katayama, T. Minegishi, H. Nishiyama, T. Takata, K. Seki, A. Kudo, T. Yamada and K. Domen, *J. Am. Chem. Soc.*, 2017, **139**, 1675.
- 19 Q. Wang, T. Hisatomi, M. Katayama, T. Takata, T. Minegishi, A. Kudo, T. Yamada and K. Domen, *Faraday Discuss.*, 2017, **197**, 491.
- 20 R. Konta, T. Ishii, H. Kato and A. Kudo, *J. Phys. Chem. B*, 2004, **108**, 8992.
- 21 R. Niishiro, S. Tanaka and A. Kudo, *Appl. Catal., B*, 2014, **150**, 187.
- 22 S. Suzuki, H. Matsumoto, A. Iwase and A. Kudo, *Chem. Commun.*, 2018, **54**, 10606.
- 23 A. Iwase and A. Kudo, *Chem. Commun.*, 2017, **53**, 6156.
- 24 A. Iwase, K. Saito and A. Kudo, *Bull. Chem. Soc. Jpn.*, 2009, **82**, 514.
- 25 R. Niishiro, R. Konta, H. Kato, W. J. Chun, K. Asakura and A. Kudo, *J. Phys. Chem. C*, 2007, **111**, 17420.
- 26 A. Kudo, K. Omori and H. Kato, *J. Am. Chem. Soc.*, 1999, **121**, 11459.
- 27 A. Iwase, H. Kato and A. Kudo, *J. Sol. Energy Eng.*, 2010, **132**, 21106.
- 28 Y. Sasaki, A. Iwase, H. Kato and A. Kudo, *J. Catal.*, 2008, **259**, 133.
- 29 Y. Sasaki, H. Kato and A. Kudo, *J. Am. Chem. Soc.*, 2013, **135**, 5441.
- 30 A. Iwase, Y. H. Ng, Y. Ishiguro, A. Kudo and R. Amal, *J. Am. Chem. Soc.*, 2011, **133**, 11054.
- 31 K. Iwashina, A. Iwase, Y. H. Ng, R. Amal and A. Kudo, *J. Am. Chem. Soc.*, 2015, **137**, 604.
- 32 A. Iwase, S. Yoshino, T. Takayama, Y. H. Ng, R. Amal and A. Kudo, *J. Am. Chem. Soc.*, 2016, **138**, 10260.
- 33 Y. Sasaki, H. Nemoto, K. Saito and A. Kudo, *J. Phys. Chem. C*, 2009, **113**, 17536.
- 34 Q. Jia, A. Iwase and A. Kudo, *Chem. Sci.*, 2014, **5**, 1513.
- 35 W. S. Hummers and R. E. Offeman, *J. Am. Chem. Soc.*, 1958, **80**, 1339.
- 36 K. Iwashina and A. Kudo, *J. Am. Chem. Soc.*, 2011, **133**, 13272.
- 37 S. Kawasaki, K. Akagi, K. Nakatsuji, S. Yamamoto, I. Matsuda, Y. Harada, J. Yoshinobu, F. Komori, R. Takahashi, M. Lippmaa, C. Sakai, H. Niwa, M. Oshima, K. Iwashina and A. Kudo, *J. Phys. Chem. C*, 2012, **116**, 24445.



- 38 S. Kawasaki, R. Takahashi, K. Akagi, J. Yoshinobu, F. Komori, K. Horiba, H. Kumigashira, K. Iwashina, A. Kudo and M. Lippmaa, *J. Phys. Chem. C*, 2014, **118**, 20222.
- 39 D. E. Scaife, *Sol. Energy*, 1980, **25**, 41.
- 40 S. Kawasaki, K. Nakatsuji, J. Yoshinobu, F. Komori, R. Takahashi, M. Lippmaa, K. Mase and A. Kudo, *Appl. Phys. Lett.*, 2012, **101**, 033910.
- 41 K. Sayama, A. Nomura, Z. Zou, R. Abe, Y. Abe and H. Arakawa, *Chem. Commun.*, 2003, 2908.
- 42 K. Sayama, A. Nomura, T. Arai, T. Sugita, R. Abe, M. Yanagida, T. Oi, Y. Iwasaki, Y. Abe and H. Sugihara, *J. Phys. Chem. B*, 2006, **110**, 2908; F. F. Adbi, L. Han, A. H. M. Smets, M. Zeman, B. Dam and R. van de Krol, *Nat. Commun.*, 2013, **4**, 2195.
- 43 Y. Park, K. J. McDonald and K.-S. Choi, *Chem. Soc. Rev.*, 2013, **42**, 2321.
- 44 Q. Jia, K. Iwashina and A. Kudo, *Proc. Natl. Acad. Sci. U. S. A.*, 2012, **109**, 11564.
- 45 M. Yamaguchi, R. Niishiro, Q. Jia, Y. Kuang, K. Kitamura, A. Iwase, T. Minegishi, T. Yamada, K. Domen and A. Kudo, to be submitted.

