



Cite this: *CrystEngComm*, 2022, 24, 4481

Received 30th March 2022,
Accepted 9th June 2022

DOI: 10.1039/d2ce00439a

rsc.li/crystengcomm

Recent progress of the single crystal growth of homologous $(\text{InGaO}_3)_m(\text{ZnO})_n$

Naoki Kase,^{*a} Noboru Kimizuka^b and Nobuaki Miyakawa ^{*a}

Compounds composed of In–Ga–Zn–O (IGZO) have characteristic homologous crystal structures represented by the following general formula: $(\text{InGaO}_3)_m(\text{ZnO})_n$. Such compounds have been applied in many fields since the discovery that amorphous oxide thin films have high electron mobility. On the other hand, very few studies have been conducted on single crystals or the bulk behavior of IGZO compounds as growing large crystals is difficult. Consequently, the physical properties have not been clarified. In this review, we summarize the historical background of IGZO regarding the crystal structure and crystal growth, and introduce our successful method for growing large single crystals of InGaZnO_4 . We also introduce the micro-crystals of $(\text{InGaO}_3)_2(\text{ZnO})_1$.

1 Introduction

Renewable energy production from sources such as solar and wind power has attracted considerable attention. The use of metal oxide semiconductors that reduce thermal energy is expected to be an important step in such energy production. Transparent amorphous-oxide semiconductors from the In–Ga–Zn–O (IGZO) series have been proposed as candidates for the active channel in transparent thin-film transistors (TFTs)^{1,2} and are currently employed in various display devices.^{3–5} IGZO-based field-effect transistors (FETs) can be used to produce devices with extremely low power consumption compared with modern Si-based devices.^{1,2,6–8} The oxide semiconductors with the nearly zero off-state current, could be one of the trump cards to solve energy problems, as they can achieve devices with the extremely lower power consumption compared to cutting-edge Si-based devices.

In 2003, Nomura *et al.* reported the fabrication of transparent FETs with a single crystalline layer of $\text{InGaO}_3(\text{ZnO})_5$ as the channel.¹ In 2004, they demonstrated the promising potential of amorphous metal oxides as transparent TFT semiconductors.² When post-transition metals are used, the direct overlap of the isotropic s-orbitals is large and not significantly affected even in an amorphous structure. Amorphous-IGZO (a-IGZO) thin films have emerged as a particularly promising structure with Hall mobilities of up to $40 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ and practical field effect mobilities of $10 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$.² These mobilities are the result of an

omnidirectional carrier transport path, which is dominated by the s-orbitals of the cation, and these values are an order of magnitude larger than those recorded for amorphous Si (a-Si). Furthermore, the feature of the device fabricated on the polyethylene terephthalate sheet is flexible and can be bent. The characteristics are stable even after repeated bending of the transparent TFT sheet.

In 2009, a new layered crystalline structure type was discovered, known as *c*-axis-aligned crystalline IGZO (CAAC-IGZO),^{9,10} in which the elemental arrangement is strongly oriented along the *c*-axis of the three crystal axes. The extremely low off-state current (I_{off}) produced in FETs fabricated by using CAAC-IGZO make the design of electronic devices with significantly low power consumption, called metal-oxide-semiconductor field-effect transistors (MOSFETs), possible. MOSFETs developed by using CAAC-IGZO can exhibit an extremely low off-state current per micrometer of $10^{-24} \text{ A } \mu\text{m}^{-1}$ (ref. 11–14) and high mobilities of $30\text{--}70 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ when the composition ratio is changed. The off-state current is considerably smaller than that of a-Si ($10^{-12} \text{ A } \mu\text{m}^{-1}$).

The crystal phase of the In–Ga–Zn–O series was first recognized by Kimizuka *et al.* in 1985.¹⁵ They reported phase relationships and solubility ranges by using powder X-ray diffraction (XRD) data of the homologous $\text{InGaO}_3(\text{ZnO})_n$ compounds ($n = 1\text{--}13$).^{15–20} The structure consisted of stacked metal oxide layers along the *c*-axis of a trigonal (even values of *n*) or hexagonal (odd values of *n*) unit cell. Several homologous compounds of the oxides of $\text{ARO}_3(\text{MO})_n$ ($\text{M} = \text{Mg, Mn, Fe, Co, Cu, Zn, and Cd}$; $\text{A} = \text{Sc, Y, In, Ho, Er, Tm, Yb, and Lu}$; $\text{R} = \text{Al, Fe, and Ga}$) were characterized.^{21–25}

Two types of homologous compounds in the IGZO-series are reported: $\text{InGaO}_3(\text{ZnO})_n$ and $(\text{InGaO}_3)_m\text{ZnO}$. Both

^a Department of Applied Physics, Tokyo University of Science, 6-3-1 Niijuku, Katsushika-ku, Tokyo 125-8585, Japan. E-mail: n-kase@rs.tus.ac.jp, miyakawa@rs.tus.ac.jp

^b Departamento de Investigación en Polímeros y Materiales, Universidad de Sonora, Rosales y Luis Encinas s/n, Hermosillo, Sonora, CP 83000, Mexico



homologous structures are layered and have a common structural unit. The crystal structure of the $\text{InGaO}_3(\text{ZnO})_n$ homologous series consists of (i) edge sharing InO_6 octahedra, (ii) wurtzite-type ZnO_4 tetrahedra, and (iii) trigonal bipyramidal GaO_5 . These compounds crystallize in space group $R\bar{3}m$ for $n = \text{odd}$ and $P6_3/mmc$ for $n = \text{even}$.²⁶ In the case of $n = \text{odd}$, the number of InO_6 layers is 3, and in the case of $n = \text{even}$, the number is 2. In contrast, for the $(\text{InGaO}_3)_m\text{ZnO}$ series, only $m = 1$ and 2 are reported.¹⁵ This type of homologous series has also been observed in compounds of $(\text{YbFeO}_3)_m\text{FeO}$,^{28–30} $(\text{InGaO}_3)_m\text{CuO}$,³¹ and $\text{In}_2\text{S}_3(\text{ZnS})_1$.³² Although the InO_6 octahedral layers in the $\text{InGaO}_3(\text{ZnO})_n$ system are equidistance along the c -axis, these are not seen in the $(\text{InGaO}_3)_2\text{ZnO}$ system. This is owing to the $(\text{InGaO}_3)_2\text{ZnO}$ conformation having insufficient space for equal-sized wurtzite slabs between the octahedral layers.

Owing to recent progress in the measurement accuracy, certain physical properties can be determined on small single crystals.^{33–37} However, since the measurement methods are severely limited, a single crystal of sufficient size is required to evaluate bulk properties. In the past decade or so, studies have been conducted on the successful growth of small single crystals of the IGZO series, which are suitable for detailed crystal structure analysis. A large single crystal is required to perform sufficient measurements for elucidating the mechanism of the unusual transport phenomena of the IGZO series. In recent years, we have succeeded in growing large single crystals of IGZO compounds, and substantial progress has been made toward examining their physical properties. In this review, we introduce the recent progress in the growth of single crystals of compounds in the IGZO series.

2 Single-crystalline and c -axis-aligned crystal thin film

Several types of the IGZO thin films have been known so far. Hosono *et al.* reported that both amorphous- and single-crystalline IGZO thin films exhibit high-electron mobility in TFTs.^{1,2} Yamazaki *et al.* reported that CAAC-IGZO thin films can be used to produce highly reliable FETs.^{9–14} Here, we introduce the growth method and detailed structure of the single-crystalline IGZO and CAAC-IGZO thin films.

2.1 Single-crystalline film

Using a conventional vapor-phase growth technique to fabricate single-crystalline thin films of metal oxides with a complex layered structure and composition is difficult. To solve this problem, a practical method for growing the single-crystalline thin films of the layered complex oxides was developed, called the “Reactive Solid-Phase Epitaxy (R-SPE) technique”.^{38–40}

Single-crystalline films of $\text{InGaO}_3(\text{ZnO})_n$ were grown by R-SPE as follows. Pulsed laser deposition (PLD) was used to deposit a 2 nm-thick ZnO epitaxial layer at 973 K (700 °C) on a (111) single-crystal yttria-stabilized zirconia (YSZ) substrate

(the template), followed by a 120 nm-thick $\text{InGaO}_3(\text{ZnO})_5$ layer at room temperature. Next, the bilayer film was annealed in an appropriate atmosphere, to obtain a single-crystalline film. The fabricated bilayer was covered with a YSZ plate to suppress evaporation of the film components and was then subjected to thermal annealing at 1673 K (1400 °C) for 30 min, resulting in the growth of its single-crystalline phase.

Fig. 1(B and C) show cross-sectional high-resolution transmission electron microscopy (HRTEM) images of the $\text{InGaO}_3(\text{ZnO})_5$ film. The layered lattice structure with periodic stacking of InO_2 -layers and $\text{GaO}(\text{ZnO})_5$ blocks is clearly observed and is also found in the electron diffraction image. At the film–substrate interface there is less of this ordered stacking and observation of field-emission scanning electron microscopy (FESEM) images revealed no defect structure such as grain boundaries or dislocation over the entire area. It was concluded from these results that the epitaxial ZnO thin layer plays an essential role in controlling the crystallographic

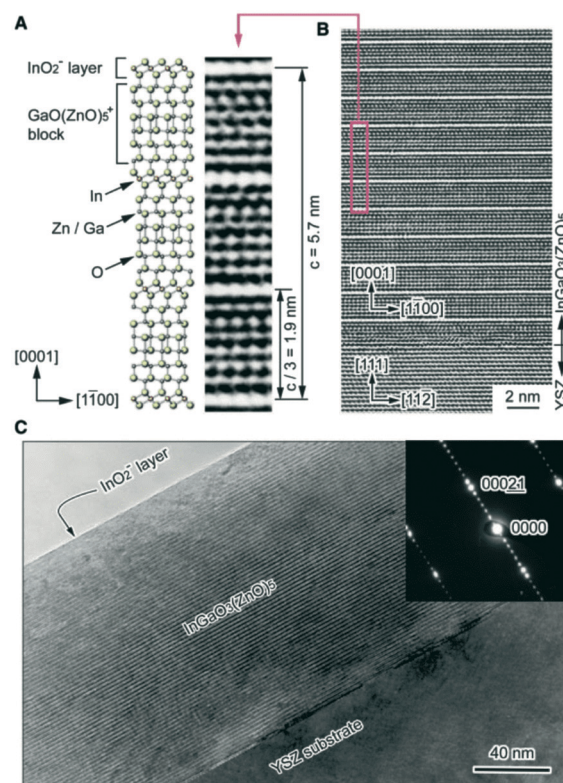


Fig. 1 (A) Crystal structure of $\text{InGaO}_3(\text{ZnO})_5$. The InO_2 layer (In^{3+} ion is located at an octahedral site coordinated by oxygens) and the $\text{GaO}(\text{ZnO})_5$ blocks (Ga^{3+} and Zn^{2+} ions share trigonal-bipyramidal and tetrahedral sites) are alternately stacked along the [0001] direction at a period of 1.9 nm (A HRTEM lattice image is shown for comparison). (B and C) Cross-sectional HRTEM images of a $\text{InGaO}_3(\text{ZnO})_5$ thin film grown on YSZ(111) by reactive solid-phase epitaxy. Periodic stacking of the InO_2 layer and the $\text{GaO}(\text{ZnO})_5$ block is clearly visible and confirmed in the electron diffraction image. A single-crystalline film is formed over the entire observation area. The topmost layer of the film is the InO_2 -layer. Panels (A–C) are reprinted with permission from ref. 1. © 2019 American Physical Society.



orientation, while the thickness ratio between the two layers controls the film composition.

Single-crystal thin films of the $\text{InGaO}_3(\text{ZnO})_n$ series with $n = 1\text{--}35$ can also be fabricated by the R-SPE method.⁴¹ The thicknesses of the resultant $\text{InGaO}_3(\text{ZnO})_n$ films were 90–250 nm. As shown in Fig. 2(b), atomically flat terraced and stepped surfaces were observed in the topographic atomic force microscopy (AFM) image of the $\text{InGaO}_3(\text{ZnO})_5$ film grown on a (111) YSZ substrate. Fig. 2(c and d) show a summary of the cross-sectional high angle annular dark field (HAADF) images of the $\text{InGaO}_3(\text{ZnO})_{15}$ single-crystalline film grown on a (111) YSZ substrate. A superlattice structure composed of InO-layers and $\text{GaO}(\text{ZnO})_n$ blocks was periodically observed.

2.2 *c*-Axis-aligned crystalline film

In this section, we introduce the *c*-axis-aligned crystalline (CAAC) thin film. CAAC-IGZO was first discovered in 2009.^{9,10} The characteristic structure of CAAC-IGZO is a *c*-axis orientation perpendicular to the surface of the entire film but with no order in the *ab*-plane direction. The benefit of this structure is that the CAAC-IGZO films can be synthesized at a

lower temperature than single-crystal thin films and stable production is possible.

Films with thicknesses 100–300 nm were fabricated on quartz substrates by the sputtering method.^{9,10,42–44} A polycrystalline substance with In, Ga, and Zn at an atomic ratio of 1 : 1 : 1 was used as the target. The conditions for dc-magnetron sputtering were as follows: the deposition power density was 0.87 W cm^{-2} and the deposition pressure was 0.4 Pa with a $\text{O}_2/(\text{Ar} + \text{O}_2)$ gas flow rate of 0.33. The power density in ac-magnetron sputtering was 1.66 W cm^{-2} and the pressure was 0.6 Pa with $\text{O}_2/(\text{Ar} + \text{O}_2)$ gas flow rate of 0.5. The deposition temperature necessary for the deposition of CAAC-IGZO was 423 K (150 °C). The temperature is lower than that of R-SPE method.

Fig. 3(a and b) show the in-plane and out-of-plane XRD spectra of the CAAC-IGZO film. The results show that the out-of-plane XRD spectrum of the fabricated IGZO film has a peak corresponding to a crystal at $\sim 31^\circ$. The peak corresponds to the (009) plane of a InGaZnO_4 crystal, which indicates that the fabricated IGZO film differed from an amorphous structure. There was no evidence of symmetry in the (110) plane detected in the in-plane XRD images. Thus, it was concluded that there was no preferential orientation in the *a*-axis and *b*-axis directions, although the *c*-axis was aligned.

Fig. 3(c and d) show the results of the selected area electron diffraction (SAED) measurements with an electron beam (beam diameter; 300 nm) positioned perpendicularly, then parallel to the substrate. In the images where the electron beam was incident perpendicularly, rings are clearly observed. The rings were ascribed to overlapped diffractions

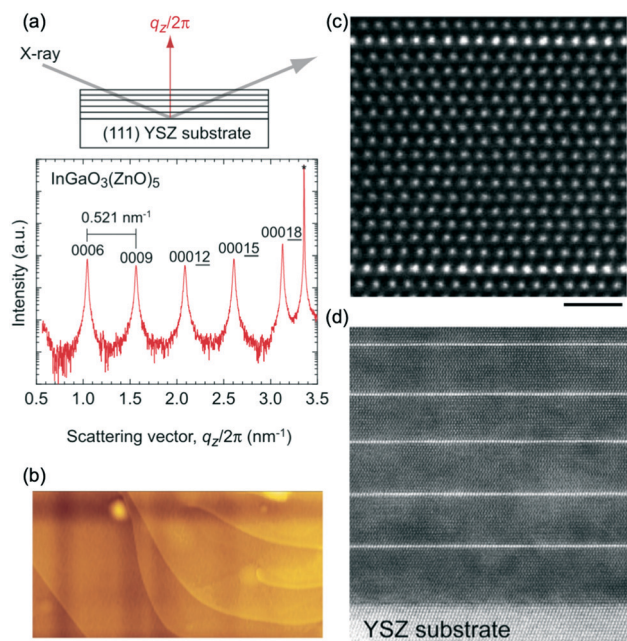


Fig. 2 X-ray diffraction (XRD) data of the superlattices. (a) Out-of-plane XRD pattern of the resultant $\text{InGaO}_3(\text{ZnO})_5$ film on a (111) YSZ substrate. The peak interval (0.521 nm^{-1}) corresponds to the distance between two adjacent InO-layers (1.92 nm). (b) Topographic atomic force microscopy (AFM) images of $\text{InGaO}_3(\text{ZnO})_5$ films on a (111) YSZ substrate (the scale bar is 500 nm). (c) Atomic arrangement of the superlattices. HAADF images of the resultant $\text{InGaO}_3(\text{ZnO})_{15}$ film grown on a (111) YSZ single crystal substrate. Brighter lines of one-atomic-layer thickness are periodically present, corresponding to InO-layers. Atomic structure of the layered structure (scale bar = 1 nm). (d) The bottom part (scale bar = 5 nm). Panels (a–d) are reprinted with permission from ref. 41. © 2022 John Wiley and Sons, Inc.

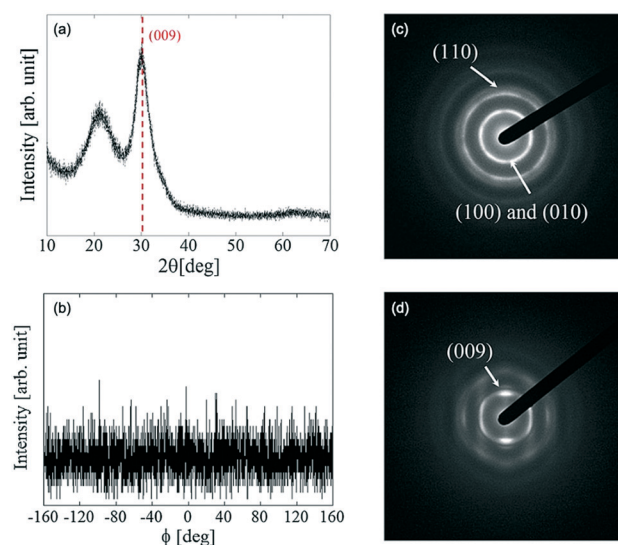


Fig. 3 XRD measurement results for *c*-axis-aligned crystalline In-Ga-Zn-O: (a) out-of-plane XRD spectra and (b) in-plane (ϕ -scan) XRD spectra ($2\theta = 56^\circ$). (c) SAED diffraction pattern of CAAC-IGZO in the perpendicular direction, and (d) in the parallel direction. Panels (a–d) are reprinted with permission from ref. 10. © 2014 Society for Information Display.



Highlight

of the (010) and (100) planes of the InGaZnO_4 . In addition, a second ring originating from the (110) plane was also observed. By contrast, when the electron beam was parallel to the substrate, spots attributed to the (009) plane were observed. These results revealed that the c -axis of the CAAC-IGZO is aligned perpendicularly to the thin film.

3 Small single crystal

The growth of small single crystals of $\text{InGaO}_3(\text{ZnO})_n$ ($n = 1-4$) was reported by Mader *et al.*, although the crystal size of $n = 3$ was not sufficiently large to perform single crystal analysis. These small crystals were unsuitable for measurements of physical properties but were sufficient for determining the detailed crystal structure. In this section, the explanation is divided into two situations, (1) where n is odd ($n = 1, 3; R\bar{3}m$) and (2) where n is even ($n = 2, 4; P6_3/mmc$).

3.1 Small single crystal of IGZO-11 and IGZO-13

The growth of a small single crystal of IGZO-11 has been reported by two groups. The methods were using high-pressure synthesis⁴⁵ and slow-cooling in a sealed Pt crucible.⁴⁶ These single crystals of the $\text{InGaO}_3(\text{ZnO})_n$ (or IGZO-1 n) are part of the homologous series where n is odd ($n = 1, 3; R\bar{3}m$).

3.1.1 Crystal growth of IGZO-11 (slow cooling). Single crystals of IGZO-11 were synthesized in a sealed Pt-tube at elevated temperatures under atmospheric pressure without flux.⁴⁶ Powdered In_2O_3 (99.99%), Ga_2O_3 (99.99%) and ZnO (99.99%) were preheated to 1073 K (800 °C) in air for one day prior to mixing. Stoichiometric masses of the starting compounds were weighed and mixed under ethanol in an agate mortar for 25 min. The mixture was sealed in platinum crucibles within 1.5 hours to allow heating to 1823 K (1550 °C). After 6 days at 1823 K, slow cooling from 1803 K (1530 °C) to 1723 K (1450 °C) over an 80 min timeframe was followed by cooling to room temperature over several hours.

Fig. 4(a) shows the optical micrograph image of a crystal of IGZO-11. Bluish-gray transparent crystals were obtained as the product of the synthesis in addition to a small amount of colorless powder. The crystal was very small, but this was the largest available in the IGZO series at that time. A HAADF image at the $[1\bar{1}00]$ orientation is shown in Fig. 4(b) with the structure model as overlay. The octahedral layers containing In appear with a strong contrast and have a distance from the (0003) lattice fringes of $d = 0.868$ nm, which agrees with the values from single crystal refinement data. The $(2\bar{1}\bar{1}0)$ prism planes with spacing of 0.164 nm are well resolved. The image IGZO-11 at atomic resolution in the electron microscope supports a perfect periodic stacking of atomic layers.

Convergent beam electron diffraction (CBED) was useful to determine the space group, because the whole crystal contributes to the diffraction pattern. The CBED patterns along the c -axis show a three-fold symmetry for the entire pattern (Fig. 4(d)) as well as for the central CBED disc

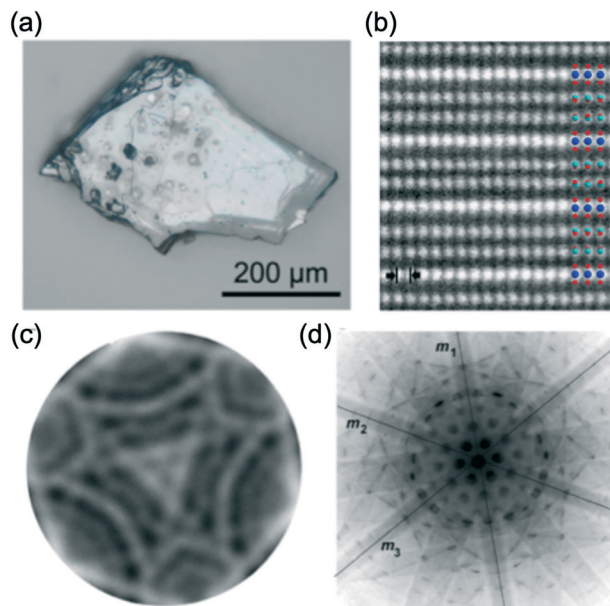


Fig. 4 (a) Optical micrograph image of a single crystal of $\text{InGaO}_3(\text{ZnO})_1$. (b) HAADF image of InGaZnO_4 in $[1\bar{1}00]$ orientation with structure model. The blue, red, and light green circles represent In, Zn/Ga, and O atoms, respectively. Cation columns of In (bright) and Zn/Ga (medium bright) are well resolved indicating the cation layer stacking in InGaZnO_4 . A spacing of 0.16495 nm of the $(2\bar{1}\bar{1}0)$ prism planes is shown. (c) Central CBED disc and (d) whole pattern results suggest $3m$ symmetry. Panels (a-d) are reprinted with permission from ref. 46. © 2014 Elsevier.

(Fig. 4(c)). Analysis of the results suggested a trigonal crystal system with $3m$ symmetry, although the results of CBED yield two possible space groups: $R\bar{3}m$ and $R\bar{3}m$.⁴⁷ However, the six-fold symmetry of the diffraction pattern of the zero order Laue zone (ZOLZ) yields only the space group $R\bar{3}m$.

The cell parameters at room temperature were determined by refining powder XRD data ($a = 0.3284(3)$ nm and $c = 2.6037(3)$ nm).⁴⁶ The transparent crystals had a bluish-gray color and exhibited an excess of Ga expressed by the cation ratio of $\text{In}:\text{Ga}:\text{Zn} = 30.2(9):39(1):31(1)$ at% as determined by energy dispersive X-ray spectroscopy (EDXS).⁴⁶ This ratio indicated that the composition of the single crystal did not correspond to the ideal value, and that the quality of the crystal was poor. In addition, the aliovalent substitution of Zn by Ga was observed for the first time in the $\text{InGaO}_3(\text{ZnO})_n$ compounds. The chemical composition of the crystals was determined as $(\text{In}_{0.9}\text{Ga}_{0.1})\text{Ga}_{1.06}\text{Zn}_{0.91}\square_{0.03}\text{O}_4$ which is indicative of aliovalent substitution of Zn^{2+} by Ga^{3+} leading to vacancies in the cation sublattice due to the synthesis conditions. While the isovalent substitution of In^{3+} by Ga^{3+} is known, this is the first report of an aliovalent substitution of Zn^{2+} by Ga^{3+} in a member of the homologous series $\text{InGaO}_3(\text{ZnO})_n$. This resulted in the formula $(\text{In}_{0.9}\text{Ga}_{0.1})\text{Ga}_{1.06}\text{Zn}_{0.91}\square_{0.03}\text{O}_4$ where 3% of the cations in 4 + 1 coordination are vacant due to aliovalent substitution. Thus, the depletion of Zn-ions was also found for the first time in a member of the homologous series $\text{InGaO}_3(\text{ZnO})_n$.



Further evidence of the defect structure was found by considering the shortened a -axis (the c -axis does not change).^{16,19} Clearly, 6% of the Zn sites were substituted with the smaller Ga^{3+} ions ($r(\text{Ga}^{3+}) = 0.047$ nm vs. $r(\text{Zn}^{2+}) = 0.06$ nm in fourfold coordination) and with 3% vacant sites, there was clearly an effect on the lattice parameters when compared to InGaZnO_4 . If the oxygen ions were already in contact along the c -axis, there would not be an effect observed. However, if smaller ions and vacant sites were introduced into the wurtzite-type layer in InGaZnO_4 the oxygen ion distance in the basal plane would be reduced. This results in a reduction in the length of the a -axis, which is exactly as observed in the XRD data.

3.1.2 Detail crystal structure of IGZO-11. Fig. 5 shows the crystal structure of IGZO-11. The InO_6 octahedra are slightly compressed in the direction of the c -axis yielding $\text{O}(1)$ – In – $\text{O}(1)$ -angles of 98.3° .⁴⁶ The deviation from the ideal 90° angle was also observed for members of the homologous series $\text{InGaO}_3(\text{ZnO})_n$ with $n = 2, 3, 4$.^{48,49} The InO_6 octahedra are connected *via* edges to layers perpendicular to the c -axis as was found in the CdI_2 -type structure with the composition InO_2 . No evidence was found from thermal displacement parameters or by electron density difference maps²⁶ for a displacement of In atoms out of the center of symmetry, as was observed for Lu in LuFeZnO_4 .

Two layers of Zn and Ga ions sharing the same site with the composition ZnGaO_2^+ were found between the InO_6 octahedral layers. Although the Zn/Ga mix site can be expressed as two layers of tetrahedra (trigonal bipyramid with

off-center cations), in this article, the Zn/Ga mixed site is referred to as the trigonal bipyramid as shown in Fig. 5. Free refinement modeling made it difficult to determine the occupation factor of the isoelectronic Ga^{3+} and Zn^{2+} ions, so the factor was set to 0.5 to maintain charge neutrality. Zn and Ga showed trigonal bipyramidal $4 + 1$ coordination with O ions; however, the cations were significantly displaced from the equatorial plane. Hence, InGaZnO_4 (= $\text{InGaO}_3(\text{ZnO})_1$) as the first member of the $\text{InGaO}_3(\text{ZnO})_n$ ($n = \text{integer}$) family has the same structural characteristics as other known members of homologous compounds.

3.1.3 Small single crystal of IGZO-13. A crystal of $\text{InGaO}_3(\text{ZnO})_3$ (or IGZO-13) was grown by the K_2MoO_4 -flux method, but the size was not large enough to perform single crystal analysis at this stage.⁴⁸ Hence, the scanning electron microscopy (SEM) image is not attached in the article and the composition ratio of the sample was not determined. The crystal structure was determined by powder X-ray diffraction data using the fine yellow powder produced. The details of the growth method are described in a later section.

Structural data of $\text{LuFeO}_3(\text{ZnO})_3$ were used as a starting model for the Rietveld refinement.²⁶ The analysis revealed that IGZO-13 crystallized as the trigonal space group ($R\bar{3}m$; no. 166) with parameters ($a = 0.32871(9)$ nm; $c = 4.1589(1)$ nm, $Z = 3$, reflection intensity-based reliability factor $R_{\text{Bragg}} = 0.0506$).⁴⁸ The structural model was validated by bond valence calculations of the Ga^{3+} distribution.²⁷ These

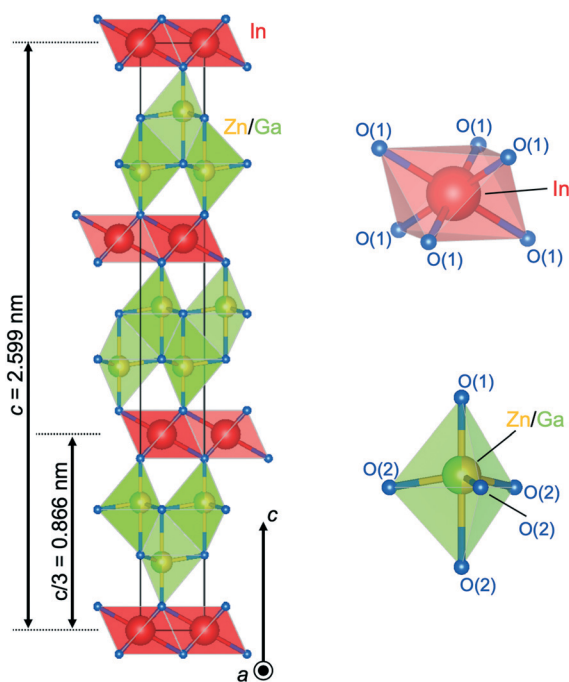


Fig. 5 Crystal structure of $\text{InGaO}_3(\text{ZnO})_1$. Coordination of InO_6 is octahedra and Zn/GaO_5 is trigonal bipyramid. Because Zn and Ga cannot be separated, the Zn/Ga mixed site is referred to as the trigonal bipyramid in this paper.

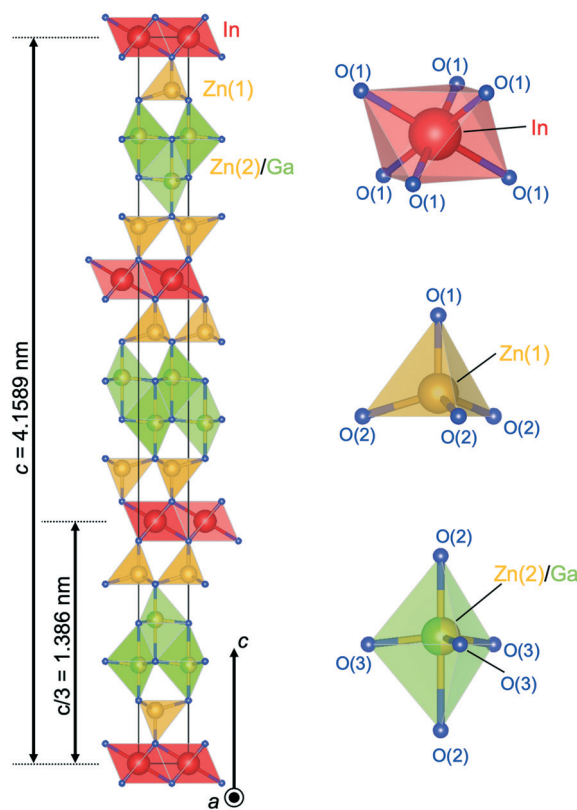


Fig. 6 Crystal structure of $\text{InGaO}_3(\text{ZnO})_3$. Coordination of InO_6 is octahedra, $\text{Zn}(1)\text{O}_4$ is tetrahedra, and $\text{Zn}(2)/\text{GaO}_5$ is trigonal bipyramid.



calculations indicated a model with Ga^{3+} only in the trigonal bipyramid at the inversion plane within the wurtzite region. Thus, these trigonal bipyramids are statistically occupied by alternating Zn and Ga, while the tetrahedra linked to the layers of InO_6 octahedra are occupied by Zn only.

Fig. 6 shows the crystal structure of IGZO-13. The inversion plane in the center of the ZnO_4 region lies between two $(\text{Zn}/\text{Ga})\text{O}_5$ trigonal bipyramids. In the GaO_5 trigonal bipyramid, the difference in the distances from Ga to the apical O(2) ions was a result of the position of Ga in the equatorial triangle of the trigonal bipyramid, which is slightly off-center in the c -axis direction. As distance increased from the InO_6 octahedral layer, the Zn^{2+} ion in the ZnO_4 tetrahedron shifted along the c -axis towards the center of the trigonal bipyramid. The distance between the cation and the apical oxygen increased from 0.18981 nm for $\text{Zn}(1)\text{-O}(1)$ to 0.20395 for $(\text{Zn}(2)/\text{Ga}(2))\text{-O}(2)$ and the distance to the basal oxygen decreased.

3.2 Small single crystal of IGZO-12 and IGZO-14

In this section, we introduce the small single crystals of the $\text{InGaO}_3(\text{ZnO})_n$ homologous series where n is even ($n = 2, 4$). One key feature of the analysis of the crystals is that the Zn and Ga sites are perfectly separated into ZnO_4 tetrahedra and GaO_5 trigonal bipyramids.

3.2.1 Crystal growth method. Single crystals of the homologous series $\text{InGaO}_3(\text{ZnO})_2$ (or IGZO-12) and $\text{InGaO}_3(\text{ZnO})_4$ (or IGZO-14) were synthesized by a K_2MoO_4 -flux method in sealed platinum tubes.^{48–52} In_2O_3 (99.999%, Alfa Aesar), Ga_2O_3 (99.999%, Alfa Aesar), and ZnO (99.999%, Alfa Aesar) were used as purchased, without further purification. Calculated stoichiometric amounts of the starting compounds were weighed and thoroughly mixed in an agate mortar and dispersed in acetone. After drying, the mixtures of the raw materials were pressed into pellets. The pellets and K_2MoO_4 (98%, Aldrich) flux were sealed in a platinum crucible. K_2MoO_4 melts at 1192 K (919 °C) and is soluble in water, both of which were important during crystal growth. The ratio between the mixture and the flux was 1:20 by weight. The platinum crucible was heated in a sealed platinum tube to 1623 K (1350 °C) for 5 days then cooled at a rate of 0.1 K min^{-1} to 873 K (600 °C). The single crystals obtained were removed from the K_2MoO_4 -flux by washing with distilled water at 343 K (70 °C). The IGZO-12 crystals were colorless and the IGZO-14 were yellow.

Fig. 7(a and d) show the SEM images of single crystals of IGZO-12 and -14. The crystals are very small (less than 0.1 mm) and are smaller than the IGZO-11 crystals grown by the slow-cooling method.⁴⁶ The shape of the single crystal indicates a hexagonal lattice. The chemical composition was determined by EDXS to be the ratio of $\text{In}:\text{Ga}:\text{Zn} = 25.55:24.38:50.07$ for IGZO-12 (ref. 48) and $\text{In}:\text{Ga}:\text{Zn} = 17.83:16.37:65.80$ for IGZO-14 (ref. 49) in at% with an error less than ± 1 each. Both ratios are in close agreement with the ideal values. Based on these results, it appears that the

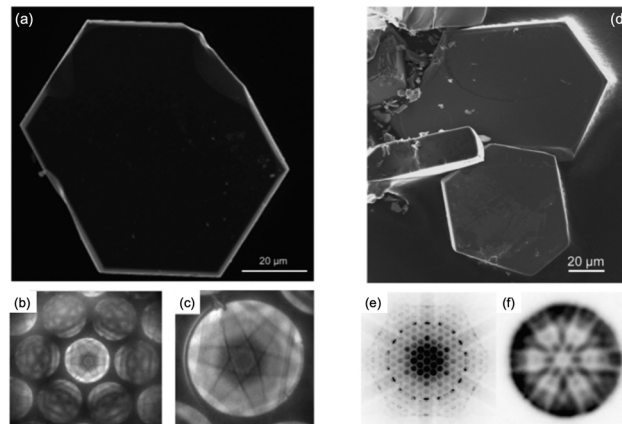


Fig. 7 (a) SEM image of single crystals of $\text{InGaO}_3(\text{ZnO})_2$. (b and c) CBED pattern in the $[0001]$ zone axis of $\text{InGaO}_3(\text{ZnO})_2$. Whole pattern (b) and bright field CBED disc (c) both exhibit hexagonal symmetry. (d) SEM image of single crystals of $\text{InGaO}_3(\text{ZnO})_4$. (e and f) CBED pattern in the $[0001]$ zone axis of $\text{InGaO}_3(\text{ZnO})_4$. $6mm$ whole pattern symmetry and $6mm$ bright field symmetry. Panels (a–f) are reprinted with permission from ref. 48 and 49. © 2009 and 2010 John Wiley and Sons, Inc.

crystals with $n = 2, 4$ are easier to grow than those with $n = 1, 3$.

The CBED patterns in the $[0001]$ zone axis (Fig. 7(b and c)) exhibit $6mm$ whole-pattern symmetry as well as $6mm$ symmetry in the bright field disc (Fig. 7(e and f)). According to Morniroli and Steeds, two hexagonal space groups of $P6_3mc$ and $P6_3/mmc$ are still possible candidates.⁴⁷ The number of independent mirror planes decides the space group, *i.e.*, $2mm$ or $2m$ point group symmetry, respectively. As a result of higher order Laue zone reflections⁵³ being examined in the electron diffraction patterns, it was concluded that $\text{InGaO}_3(\text{ZnO})_2$ and $\text{InGaO}_3(\text{ZnO})_4$ crystallize in the space group $P6_3/mmc$ (no. 194).

3.2.2 Detailed crystal structure of IGZO-12. The crystal structure of IGZO-12 was identified by single-crystal XRD measurements. Although the basic units of the crystal structure are common to $n = \text{odd}$, the arrangement is different leading to a different space group ($P6_3/mmc$; no. 194). Based on single crystal structure refinement with the parameters ($a = 0.32909(2)$ nm; $c = 2.2485(2)$ nm; $Z = 2$; conventional residual factor $R_1 = 0.0237$; weighted residual factor $wR_2 = 0.0711$),⁴⁸ the compound was found to consist of an alternate stacking of the InO_6 octahedra, ZnO_4 tetrahedra, and GaO_5 trigonal bipyramids.

Fig. 8 shows the crystal structure of IGZO-12. The inversion plane was built by Ga^{3+} ions in trigonal bipyramidal coordination in the wurtzite-type block at half distance between the layers of InO_6 octahedra. The GaO_5 polyhedra share corners between each other and with neighboring ZnO_4 units. The site occupation of the isoelectronic Zn^{2+} and Ga^{3+} , which are indistinguishable by XRD was confirmed by valence sum calculations.²⁷ These calculations showed reasonable values only for Zn at the tetrahedral site and Ga in the trigonal bipyramidal interstices. Thus, the structure of



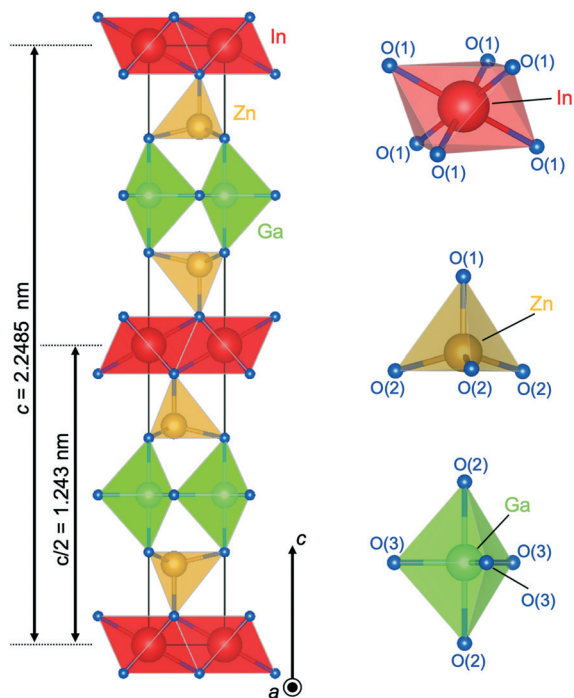


Fig. 8 Crystal structure of $\text{InGaO}_3(\text{ZnO})_2$. Coordination of InO_6 octahedra, ZnO_4 tetrahedra, and GaO_5 trigonal bipyramid.

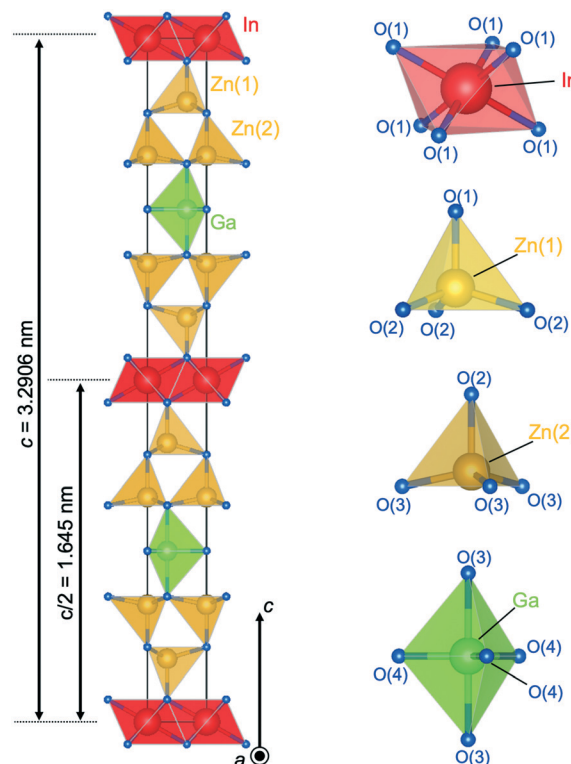


Fig. 9 Crystal structure of $\text{InGaO}_3(\text{ZnO})_4$. Coordination of InO_6 octahedra, $\text{Zn}(1)\text{O}_4$ tetrahedra, $\text{Zn}(2)\text{O}_4$ tetrahedra, and GaO_5 trigonal bipyramid.

IGZO-12 from single-crystal data clearly shows an ordering of the R cation in $\text{AR}_3(\text{ZnO})_n$ with Ga^{3+} which occupies the inversion boundary.

One of the differences between $n = \text{odd}$ and even in $\text{InGaO}_3(\text{ZnO})_n$ structures is the Ga position, which was found to be present in the trigonal bipyramid. Although the Ga position is off-center in the c -axis direction in IGZO-11 and -13, the GaO_5 trigonal bipyramid has the same distance between Ga and the apical O(2) ions of 0.21379(2) nm in IGZO-12, because the position of the Ga atom is in the equatorial triangle of the trigonal bipyramid. The Zn^{2+} ion is shifted off the center of the ZnO_4 tetrahedron along c towards the trigonal bipyramid which increased the angle O(2)–Zn–O(2) with respect to O(1)–Zn–O(2). This change in the Zn position was found to be the same regardless of the value of n .

3.2.3 Detailed crystal structure of IGZO-14. The crystal structure of IGZO-14 was determined from single-crystal XRD data. The compound crystallized in a hexagonal space group ($P6_3/mmc$; no. 194). Single crystal structure refinement with XRD data ($a = 0.32850(2)$ nm; $c = 3.2906(3)$ nm, $Z = 2$, $R_1 = 0.0685$, $wR_2 = 0.1760$)⁴⁹ revealed that the crystal structure consisted of the same units of IGZO-12.

Fig. 9 shows the crystal structure of IGZO-14. The structure has $\text{Zn}(2)\text{O}$ layers inserted between $\text{Zn}(1)\text{O}$ layers and GaO_5 layers in the IGZO-12 structure. At both sides, the InO_6 layers are connected by corners with ZnO_4 tetrahedra. The position of GaO_5 in IGZO-14 is different, and it is located under the $\text{Zn}(1)$ atom instead of under the In atom as in IGZO-12. This inversion plane is built by Ga^{3+} ions filling

trigonal bipyramidal sites in the wurtzite-type blocks at half distance between the layers of InO_6 octahedra. The position of the inversion is located at the center of the trigonal bipyramidal GaO_5 , in the equivalent position to IGZO-12.

As mentioned before, the cations are shifted from the tetrahedron center towards the basal plane with increasing distance from the InO_6 octahedral layer. The tendency was most clearly found in the IGZO-14 structure. The Zn^{2+} ion is shifted off the center of the ZnO_4 tetrahedron along c -axis towards the oxygen triangle on the basal plane. The distance from Zn in the tetrahedron to the apex oxygen increased with increasing distance from the InO_6 octahedral layer. The $\text{Zn}(1)\text{–O}(1)$ distance was found to be 0.1955(9) nm, and $\text{Zn}(2)\text{–O}(2)$ was 0.2038(9) nm. The $\text{Zn}(2)$ site, which is farther from InO_2 , had a longer distance from the apex oxygen, which indicated that $\text{Zn}(2)$ is closer to the basal plane. The GaO_5 at an equal distance from the InO_6 octahedra is in the equatorial triangle of the trigonal bipyramid without the Ga being shifting from the center as in other homologues.

The displacement parameters of O(4) showed that the oxygen ion in the equatorial plane of the GaO_5 unit are elongated along c -axis. A comparison between the components with $n = 2$ and 4 showed an increasing displacement parameter of O(4). Although the result indicated a split position, attempts to refine this for the O(4) site showed no reasonable results. Similar results were not obtained for the single crystals of the homologous



InGaO₃(ZnO)_n series. The reason for this significantly elongated displacement parameter along *c*-axis of the O(4) remains to be clarified.

The site occupation of Zn and Ga was determined by bond valence sum calculations,²⁷ which is the same method as IGZO-12. The tetrahedra in the wurtzite-type region are increasingly deformed with decreasing distance to the InO₆ octahedral layers. However, calculations taking into consideration the statistical distribution and replacement of Ga(1), Zn(1), and Zn(2) yielded less valid results. The calculation of the bond valence sum showed reasonable results only with Zn in tetrahedral sites and Ga in trigonal bipyramidal interstices. The crystal parameters of the InGaO₃(ZnO)_n (*n* = 1–4) single crystals are summarized in Table 1.

3.2.4 Detailed crystal structure of InFeO₃(ZnO)₄. Although the perfect ordering of Ga and Zn has been reported in IGZO-12 and IGZO-14 by structural analysis and calculation, the mixed ratio of Ga/Zn was not directly observed because isoelectronic Ga³⁺ and Zn²⁺ could not be distinguished by XRD measurements. To investigate the distribution of R³⁺ and M²⁺, the single crystal of InFeO₃(ZnO)₄ was grown because XRD measurements can distinguish between Fe³⁺ and Zn²⁺.⁵⁴ Fig. 10(a) shows SEM image of the single crystals of InFeO₃(ZnO)₄ grown by the K₂MoO₄ flux method with the same conditions as IGZO-12 and IGZO-14.

The structure of InFeO₃(ZnO)₄ was calculated as being in space group *P*6₃/*m**m**c*; no. 194 (*a* = 0.3211(6) nm; *c* = 3.3032(6) nm, *Z* = 2, *R*₁ = 0.0368, *wR*₂ = 0.0928)⁵⁴ from XRD data. Alternating stacks of edge-sharing InO₆ octahedra and 5 layers of corner-sharing (Zn/Fe)O₄ tetrahedra were found. The composition was measured by EDXS and resulted in a ratio In:Fe:Zn equal to 18.3:16:65.7.⁵⁴ The ratio indicated that the Fe value was slightly low, and that the ratio of IGZO-14 was closer to the ideal value. As shown in Fig. 10(b), a HRTEM image in *<2110>* orientation is observed in InFeO₃(ZnO)₄. The In layers are seen as parallel contrasting dark lines with a spacing of *c*/2 = 1.65 nm and 5 lines generated from the 5 layers of Zn/Fe cations are also easily observed.

Refining the site occupation factors of the cations in (Zn/Fe)O₄ tetrahedra resulted in a clear and reasonable gradient for the distribution of Fe. The site occupation factor refinement results for the highest Fe amount (34%) showed Fe on the trigonal bipyramidal position Zn(4)/Fe(4) and a

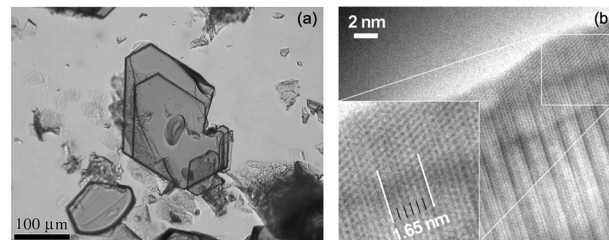


Fig. 10 (a) Optical image of single crystals of InFeO₃(ZnO)₄. (b) HRTEM images of InFeO₃(ZnO)₄ imaged along *<2110>* zone axis. Dark spot lines have the In-layer spacing of 1.65 nm. Panels (a and b) are reprinted with permission from ref. 54. © 2014 John Wiley and Sons, Inc.

gradual decrease from 23.2% with Zn(3)/Fe(3) to 9% with Zn(2)/Fe(2) towards the InO₆ layers (the Zn(*x*)/Fe(*x*) sites are surrounded by the same *x* of the basal plane O(*x*) ion as shown in Fig. 9). Since the distribution of Fe was observed, it is desirable to consider a method that can directly observe the ratio of Ga/Zn ratio even in IGZO.

4 Single crystal growth by an optical floating zone method

In 2019, our group succeeded in growing large and high-quality single crystals of IGZO-11 using a novel and reproducible method.⁵⁵ Using this process, we also succeeded in growing very small single crystals of IGZO-21.⁴⁶ In this section, we introduce this growth method and the resulting bulk single crystals of IGZO-11 and tiny crystal of IGZO-21.

4.1 Large single crystal of IGZO-11

The IGZO-11 crystals were successfully grown for the first time by the OFZ (optical floating zone) method in applying a gas pressure using a feed rod with a Zn-rich composition.⁵⁵ A stable liquid phase was visually apparent at a pressure of ~0.9 MPa. In addition, it was found that IGZO-11 melts at about 1923 K (1650 °C).

4.1.1 Optical floating zone method of IGZO-11. The OFZ method is a crucible-free crystal growth method. Fig. 11 shows a scheme of the OFZ method. Ellipsoidal mirrors focused the light from halogen lamps onto a vertically held feed rod. A molten zone was kept between two vertical solid rods by its own surface tension. After stabilization, the

Table 1 Structural parameters for small single crystals of homologous InGaO₃(ZnO)_n

	IGZO-11 (ref. 45)	IGZO-11 (ref. 46)	IGZO-12 (ref. 48)	IGZO-13 (ref. 48)	IGZO-14 (ref. 49)
Crystal group	<i>R</i> 3̄ <i>m</i> (no. 166, <i>Z</i> = 3)	<i>R</i> 3̄ <i>m</i> (no. 166, <i>Z</i> = 3)	<i>P</i> 6 ₃ / <i>m</i> <i>m</i> <i>c</i> (no. 194, <i>Z</i> = 2)	<i>R</i> 3̄ <i>m</i> (no. 166, <i>Z</i> = 3)	<i>P</i> 6 ₃ / <i>m</i> <i>m</i> <i>c</i> (no. 194, <i>Z</i> = 2)
Lattice constants (nm)	<i>a</i> = 0.32990(2) <i>c</i> = 2.6101(3)	<i>a</i> = 0.3275(1) <i>c</i> = 2.599(1)	<i>a</i> = 0.32909(2) <i>c</i> = 2.2485(2)	<i>a</i> = 0.328971(9) <i>c</i> = 4.1589(1)	<i>a</i> = 0.32850(2) <i>c</i> = 3.2906(3)
Cell volume <i>V</i> (nm ³)	0.24601(3)	0.2414(3)	0.21089(3)	0.389788	0.30752(4)
Crystal color	Colorless	Bluish-gray	Colorless	Yellow	Yellow
Growth method	High pressure (2 GPa)	Slow cooling	K ₂ MoO ₄ flux	K ₂ MoO ₄ flux	K ₂ MoO ₄ flux
Crystal size (mm)	Not shown	~0.2	~0.2	Not shown	~0.2
In:Ga:Zn ratio	Not shown	30.2:39:31	25.6:24.4:50.1	Not shown	17.8:16.4:65.8



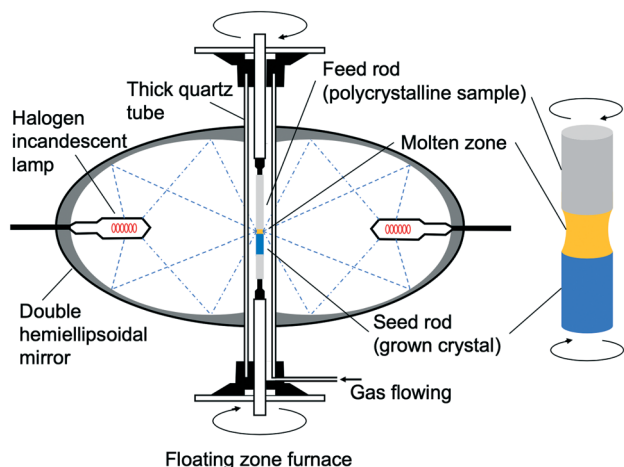


Fig. 11 Schematic image of a floating zone furnace and crystal growth.

molten zone was moved along the sample to grow the single crystal. A single crystal was grown by dipping a seed crystal into one end of the zone and moving the molten zone toward the feedstock.

The feed rods were prepared by a standard solid-phase reaction method using high purity reactants. Powders of In_2O_3 (99.99%, Furuuchi Chem.), Ga_2O_3 (99.99%, Furuuchi

Chem.) and ZnO (99.99%, Kojundo Chem. Lab.) were preheated at 1523 K (1250 °C) in air for 24 hours. Using a molar ratio of $1:1:2+x$ ($0 < x < 0.2$) of $\text{In}_2\text{O}_3:\text{Ga}_2\text{O}_3:\text{ZnO}$, the powders were weighed and thoroughly mixed in an agate mortar with ethanol. After drying, the mixtures were further ground, calcined at 1523 K (1250 °C) for 48 hours, then the calcined mixtures were thoroughly re-ground. The re-ground powders were densely packed in a rubber tube with a diameter of 6 mm and then evacuated using a rotary pump to improve both density and uniformity. The tube was then hydrostatically pressed at a pressure of 18 MPa. The polycrystalline rod was sintered at 1523 K (1250 °C) for 24 hours in air and was also used as the support rod (Fig. 12(a)). Single crystals of In–Ga–Zn–O were grown at a growth rate of 1.9 mm h^{-1} and gap speed of 2.0 mm h^{-1} in constantly flowing dry-air with a pressure of 0.1–0.9 MPa. The upper and lower polycrystalline rods were rotated in opposite directions at a rate of 16 and 14 rpm, respectively. The rod-rotation speed may have been too slow, as the sample that was grown had a large variation in electrical conductivity.

In the growing process, a vaporized material (mainly ZnO) was found in the molten zone and deposited on the quartz tube located in the infrared radiation furnace. Suppression of the ZnO vaporization was found to be important in attaining a stable liquid phase in the molten zone and thus precipitating IGZO-11 as the primary phase. Thus, the two methods that were followed were: (i) synthesis was carried out in a pressurized atmosphere of $P = 0.1\text{--}0.9 \text{ MPa}$ under dry-air flow, and (ii) continuous supply of ZnO was provided using a feed rod with a Zn-rich composition during crystal growth.

The optimal growth conditions were found to be when the crystal growth was performed under pressures of 0.1, 0.6 and 0.9 MPa for feed rods with a stoichiometric ratio ($\text{In}:\text{Ga}:\text{Zn} = 1:1:1$) or pressures of 0.8 and 0.9 MPa for feed rods with a Zn-rich starting mixture ($\text{In}:\text{Ga}:\text{Zn} = 1:1:1.1$). The presence of IGZO-21 impurities was detected by the powder XRD patterns post synthesis. We found that the XRD peak-intensity of IGZO-21 decreased with increasing gas pressure and/or using a feed rod with a Zn-rich composition. This demonstrated that applying gas pressures of $\leq 0.9 \text{ MPa}$ were very effective at reducing ZnO vaporization, but that these conditions did not suppress it completely. Thus, excess ZnO in the ratio was necessary to compensate for the loss of the ZnO due to vaporization.

4.1.2 Bulk single crystal of IGZO-11. Fig. 12(b) shows an image of the IGZO-11 crystal bar as it grew. The crystal growth mainly occurred along the ab -plane. The color of the crystal was bluish, and the crystal was easily cleaved. Fig. 12(c) shows a plate-shaped single crystal which was mechanically cleaved from a rod-shaped crystal. The cleavage surface represents the ab -plane (perpendicular to the c -axis), because all the peaks obtained from XRD analysis were able to be assigned to the $(00l)$ reflections of IGZO-11.

As the evaporation of ZnO was observed during the crystal growth, it was likely that the crystals have Zn defects. In the

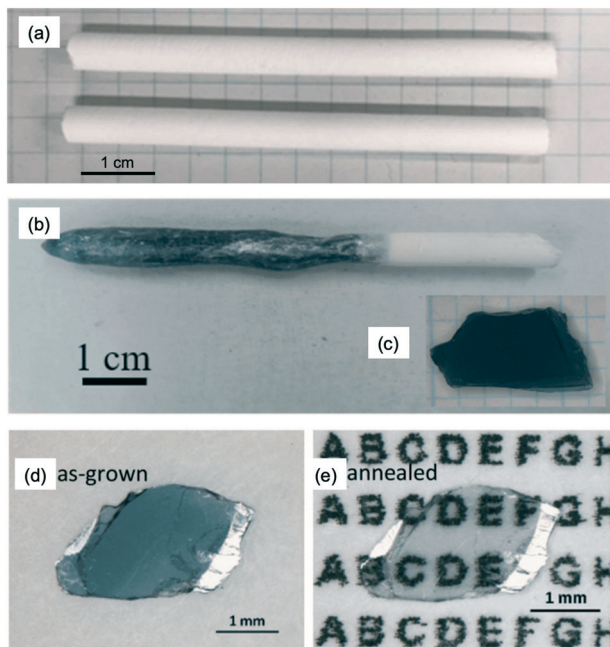


Fig. 12 (a) Sintered polycrystalline rod of $\text{InGaO}_3(\text{ZnO})_1$. (b) Rod-shaped $\text{InGaO}_3(\text{ZnO})_1$ crystal grown by the OFZ method under a pressure of 0.9 MPa. (c) Plate-shaped single crystal, that was cleaved mechanically from the rod-shaped crystal, where the cleavage plane was perpendicular to the c -axis. (d) Image of the crystal without modification. (e) Image of the post-annealed crystal (1273 K (1000 °C) for 72 hours under a flow O_2 gas at 0.1 MPa). Panels (b–d) are reprinted with permission from ref. 55. © 2019 Royal Society of Chemistry.



Highlight

sample grown at a low-gas pressure where the vaporization of ZnO was not suppressed, IGZO-21 crystals were obtained as primary products together with IGZO-11 crystals containing many Zn defects. On the other hand, samples prepared in combination with Zn-rich solutions under a pressure of $P > 0.8$ MPa had only a small amount of Zn vacancies. In addition, there were no IGZO-21 impurities in the crystals grown by under these optimum conditions.

X-ray fluorescence (XRF) spectra showed the defected areas and allowed the composition ratio of the crystals to be measured. The XRF spectra of the crystal shows the ratio of In:Ga:Zn = 0.95:1.00:0.94 in at%.⁵⁵ The result indicated that the amount of cation defects is very small, which is consistent with the lattice constant, and indicates that a high quality crystal was produced. In addition, the chemical composition ratio was nearer to the stoichiometric ratio than that obtained from slow cooling in the Pt crucible,⁴⁶ which indicates that the crystal is of higher quality. The crystal parameters of the large single crystal of IGZO-11 are summarized in Table 2.

Single crystals with a size of 3–5 mm were annealed at 1273 K (1000 °C) for 72 hours under an oxygen pressure of 0.1 MPa. The heat treatment brought about a change in color of the crystals from blue to colorless. Fig. 12(d and e) show images of a single crystal before and after the annealing process. This result suggested that the color change may be caused by oxygen filling oxygen-deficient sites during the O₂ annealing process. This color-change also indicates that the oxygen vacancies caused a change in the valence state of the three cations (In, Ga, Zn); however, which specific cation is affected is not yet clear.

To investigate the optical properties and determine the origin of the color-change, transmittance spectral data were collected, as shown in Fig. 13(a). An almost linear decrease was observed from 3.27 to 3.02 eV when n increased from 1 to 9,⁵⁶ which was consistent with the decrease also observed for $n = 1$ to $n = 3$.¹⁶ Furthermore, when UV-visible spectra were recorded from one of the bluish single crystals, the spectra showed a substantial increase in remission in the UV region. This corresponded to the optical band gap, which was determined to be $E_g = 3.17$ eV by Tauc plot.^{57–59} This optical gap is slightly smaller than the band gap $E_g = 3.27$ eV of InGaO₃(ZnO) that was determined by our group.⁵⁵ The

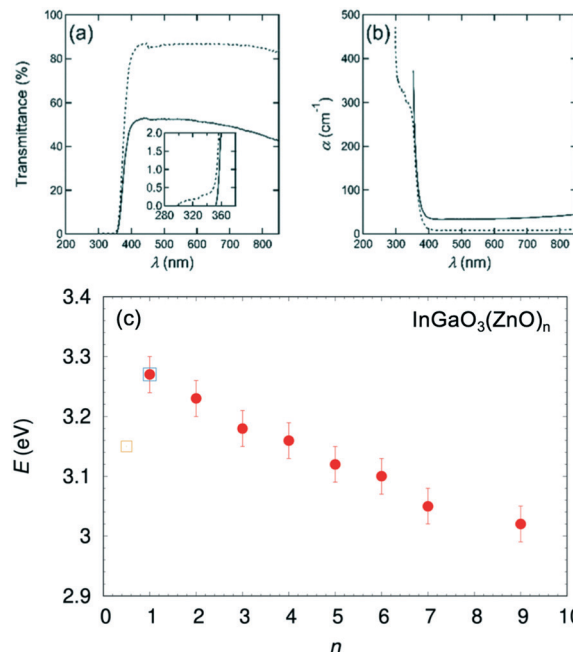


Fig. 13 (a) The transmission spectra of single crystal IGZO-11 before (solid line) and after (dotted line) the O₂ annealing at 1273 K (1000 °C) in 72 hours. The inset is an enlargement of the region of shorter incident photon wavelength. The sample thickness is 0.195 mm. (b) The adsorption coefficient α vs. wavelength of the incident photon λ for an unmodified, single crystal. (c) Evolution of the band gap in the series InGaO₃(ZnO)_{*n*} estimated using a Tauc plot. Circle symbol represents the gap values estimated from the polycrystalline sample⁵⁶ and square symbols^{55,68} estimated from single crystals. Panels (a and b) are reprinted with permission from ref. 55. © 2019 Royal Society of Chemistry.

energy gap value is plotted in Fig. 13(c), as InGaO₃(ZnO)_{0.5}. The value is a large deviation from that found in InGaO₃(ZnO)_{*n*}.

4.1.3 Transport properties of the bulk crystal IGZO-11.

One of the most important reasons to grow bulk single crystals of IGZO compounds is that they make it possible to measure anisotropy in the electrical properties. Given the layered structure of IGZO-*mn*, anisotropic properties are expected because non-relativistic first-principles calculations suggest the two-dimensional electronic structure.⁶⁰ However, conductivity along the *c*-axis (σ_c) has not yet been reported, because a large single crystal required for bulk measurements was not available. Fig. 14(a) shows the representative *T*-dependence of σ_{ab} and σ_c of unmodified IGZO-11 single crystals, where both σ_{ab} and σ_c were measured using the same fragment of crystal and then compared. The magnitude of σ_c at 300 K was in the range 0.2–50 S cm⁻¹, which is considerably smaller than the σ_{ab} (300 K) of the unmodified single crystals. A comparison σ_c and σ_{ab} in the same crystal gives a ratio of $\sigma_{ab}/\sigma_c > 40$, as shown in Fig. 14(a). Moreover, the ratio increases for crystals with lower σ_{ab} , which corresponds to a lower carrier concentration (Fig. 14(c)). The In–In distance along the *c*-axis is much longer than the distance along the *ab* plane. Thus, we suggest that a reduction in the overlap of the In-5s orbital in the

Table 2 Structural parameters for single crystals of homologous (InGaO₃)_{*m*}ZnO using an optical floating zone method

	IGZO-11 (ref. 55)	IGZO-21 (ref. 68)
Crystal group	$R\bar{3}m$ (no. 166, $Z = 3$)	$P6_3/mmc$ (no. 166, $Z = 4$)
Lattice constants (nm)	$a = 0.32990(4)$ $c = 2.6018(3)$	$a = 0.33137(1)$ $c = 2.9523(1)$
Cell volume V (nm ³)	0.24523	0.280748
Crystal color	Bluish	Bluish
Growth method	Floating zone	Floating zone
Crystal size (mm)	~30	~0.5
In:Ga:Zn ratio	0.95:1.00:0.94	19.9:59.9:20.2
Air pressure (MPa)	0.9	0.1



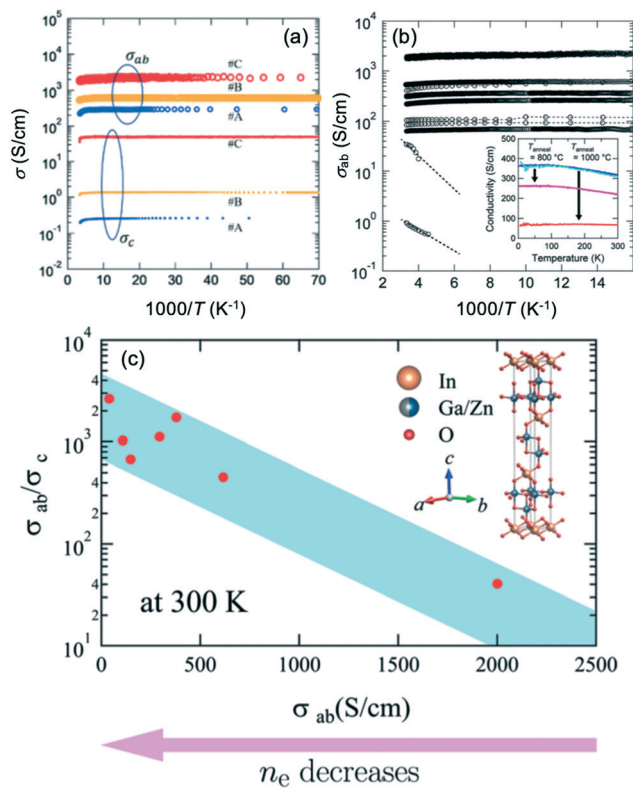


Fig. 14 (a) Comparison of electrical conductivities for in-plane (σ_{ab}) and out-of-plane (σ_c) unmodified IGZO-11 single crystals. (b) Arrhenius plots of the in-plane electrical conductivity $\sigma_{ab}(T)$, inset of (b) shows that the annealing temperature effect of the σ_{ab} of IGZO-11, where the T_{anneal} for purple and red circles are 1073 K (800 °C) and 1273 K (1000 °C), respectively, and the annealing time for both samples was fixed at 10 hours. (c) The anisotropic ratio σ_{ab}/σ_c for single crystals at 300 K. Panels (a-c) are reprinted with permission from ref. 55. © 2019 Royal Society of Chemistry.

c-axis direction is possibly the origin of the anisotropic conductivity of the IGZO-11 crystal.

Fig. 15(a and b) show the Arrhenius plots of the carrier density $n_e(T)$ and the Hall mobility $\mu_e(T)$ for bulk IGZO-11 single crystals. The increase in the Hall mobility with increasing carrier density is the opposite of the standard behavior of single crystal semiconductors such as $\beta\text{-Ga}_2\text{O}_3$.^{61–63} Confirmation of these anomalous properties in bulk single crystals is important to understand whether this phenomenon is intrinsic in IGZO systems. Fig. 15(c) shows $\mu_e(300\text{ K})$ of bulk single crystals of IGZO-11 plotted as a function of $n_e(300\text{ K})$ including a-IGZO,^{2,64} sc-IGZO-1*n*,^{39,65} poly-crystalline IGZO (pc-IGZO).⁶

The relationship between $\mu_e(300\text{ K})$ and $n_e(300\text{ K})$ is typical for the thermally activated conductivity observed in amorphous semiconductors, our bulk single crystals also exhibited the same behavior, which is unusual. However, the Hall mobility of the bulk single crystals was larger than those of the other reported sample,^{2,6,39,64,65} when $n_e(300\text{ K})$ exceeds 10^{19} cm^{-3} . The reason for this difference was attributed to the decrease in relaxation time due to disorders such as grain boundaries and anti-site defects in the InO_2

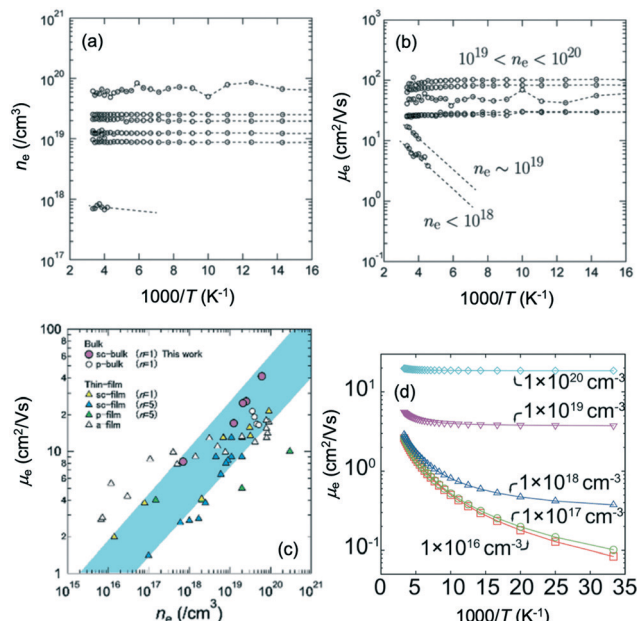


Fig. 15 Arrhenius plots of (a) the carrier density $n_e(T)$ and (b) the Hall mobility $\mu_e(T)$ for bulk IGZO-11 single crystals. (c) The carrier density dependence of the Hall mobility at 300 K for bulk IGZO-11 single crystals (purple circles). Hall mobilities $\mu_e(n_e)$ of bulk IGZO-11 single crystals are compared to those of a-IGZO, sc-IGZO-1*n*, pc-IGZO-11. (d) Theoretical calculation of the temperature dependence of Hall mobility at various carrier densities. Panels (a-c) are reprinted with permission from ref. 55. © 2019 Royal Society of Chemistry. Panel (d) is reprinted with permission from ref. 66. © 2013 AIP Publishing LLC.

conducting layer. Based on these characteristics, we believe that it is worth investigating possible device applications using IGZO bulk single crystals.

Fig. 15(d) shows the calculated Arrhenius plots of the Hall mobility $\mu_e(T)$.⁶⁶ In the model, cation-disorder scattering, ionized impurity, and polar optical phonon scattering^{66,67} were considered. The calculated mobility values accurately reproduced the temperature dependence of the mobility reported in both conventional thin film and amorphous systems. Relaxation time of cation-disorder scattering increased rapidly at low electron energies and this was seen as the main reason the mobility increased with increasing carrier density. However, there was a difference of an order of magnitude in our measurement results, which indicated that more accurate parameters are required in future studies.

4.2 Small single crystal of IGZO-21

As mentioned previously, we discovered a $(\text{InGaO}_3)_2\text{ZnO}$ phase (or $\text{InGaO}_3(\text{ZnO})_{0.5}$, IGZO-21) in the process of growing the IGZO-11 crystal by the OFZ method. The crystal structure is another homologues type, and has a structure in which *m* is changed in $(\text{InGaO}_3)_m\text{ZnO}$. The homologues compounds with only *m* = 1 and 2 have been reported.

4.2.1 Optical floating zone method of IGZO-21. Single crystals of IGZO-21 were grown by employing the OFZ method.⁶⁸ The procedure for growing crystals was almost



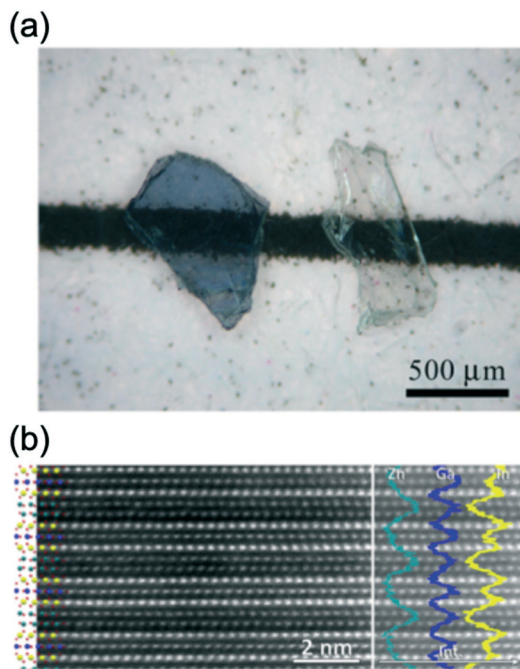


Fig. 16 (a) Single crystals of $\text{In}_{1-x}\text{Ga}_{1+x}\text{O}_3(\text{ZnO})_{0.5}$ before (left) and after annealing (right). (b) HAADF image with crystal structure of $(\text{InGaO}_3)_2\text{ZnO}$ overlaid (left) and extracted EDX line-scan profiles for In-L (yellow), Ga-K (blue) and Zn-K (cyan) as overlay (right). Panels (a and b) are reprinted with permission from ref. 68. © 2020 Elsevier.

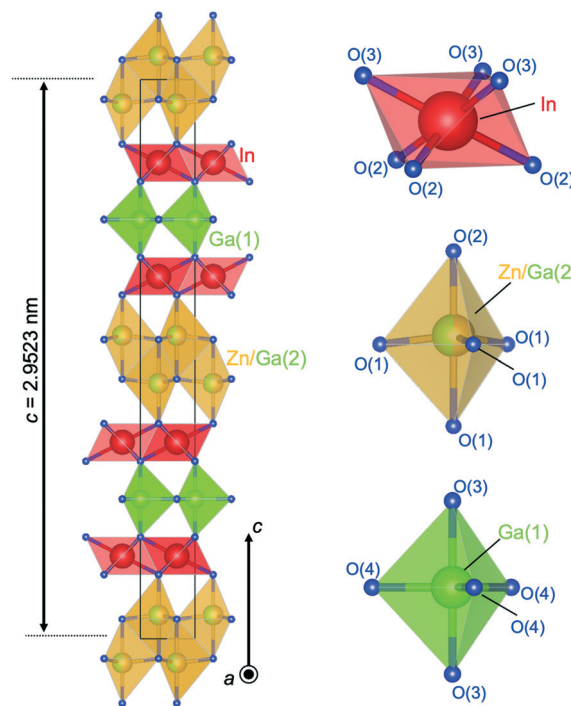


Fig. 17 Crystal structure of $(\text{InGaO}_3)_2\text{ZnO}$. Coordination of InO_6 is octahedra, Zn/Ga(2)O_5 trigonal bipyramids (two Zn/Ga(2)O_4 tetrahedra), and Ga(1)O_5 trigonal bipyramid.

the same as mentioned earlier. The feed rods were prepared by a standard solid-phase reaction method using high purity reactants. The powders were weighed with a stoichiometric ratio ($\text{In}:\text{Ga}:\text{Zn} = 1:1:1$). The crystal bar was grown in constant gas flow of dry air in a pressure region of 0.1 MPa where Zn easily evaporates due to insufficient pressurization. We note that the conditions were not optimal for the crystal growth of IGZO-21 and further optimization is required.

4.2.2 Crystal parameters and chemical composition ratio of IGZO-21. As shown in Fig. 16(a), the dark rod was crushed into small crystals, which were bluish in color. The size of the crystals was about $0.5 \times 0.2 \text{ mm}^2$. These bluish crystals became colorless after annealing at 1273 K (1000 °C) for 10 hours in air. Chemical composition ratios of several crystals of $\text{In}_{1-x}\text{Ga}_{1+x}\text{O}_3(\text{ZnO})_{0.5}$ were confirmed by EDX spectroscopy. The measured spectra of 27 crystals show different values depending on the composition of $\text{In}_{1-x}\text{Ga}_{1+x}\text{O}_3(\text{ZnO})_{0.5}$ with $0.20 \leq x \leq 0.82$. The averaged value was estimated to a cation ratio of $\text{In}:\text{Ga}:\text{Zn} = 19.87:59.88:20.24$ in at% with an error of ± 1.4 at%.⁶⁸ Thus, the cation ratio was assigned to the chemical formula $\text{In}_{1-x}\text{Ga}_{1+x}\text{O}_3(\text{ZnO})_{0.5}$. Consequently, in crystals of the averaged composition, the octahedral site is occupied with In and Ga in a ratio of $\text{In}:\text{Ga} = 0.5:0.5$ resulting in octahedral $(\text{In}_{0.5}\text{Ga}_{0.5})\text{O}_6$ and trigonal bipyramidal $(\text{Ga}_{1.0})\text{O}_5$ arrangements. The wide range in the chemical ratio at each crystal may have originated from the method of growth. In the OFZ method, the temperature gradient from the center axis of the rod to the surface led to

evaporation of volatile oxides preferentially at the surface where there was also a loss of In_2O_3 .

4.2.3 Detailed crystal structure of IGZO-21. Analyses of the XRD and electron diffraction spectra indicated a hexagonal crystal system and the space group $P6_3/mmc$ (no. 194). Structural refinement from single crystal XRD data ($a = 0.33137(1) \text{ nm}$; $c = 2.9523(1) \text{ nm}$, $R_1 = 0.0211$, $wR_2 = 0.0538$) indicated an isotopic structure to that of $\text{YbFeO}_3(\text{FeO})_{0.5}$.⁶⁹ Fig. 17 shows the crystal structure of IGZO-21. Although the InO_6 octahedral layers in the $\text{InGaO}_3(\text{ZnO})_n$ family are equidistant along the c -axis, they were not found to be in $(\text{InGaO}_3)_2\text{ZnO}$. This is because the composition refuses permission to exist equal-sized wurtzite slabs between the octahedral layers. The InO_6 octahedral layers are alternately stacked with single layers of nearly ideal GaO_5 trigonal bipyramids and two layers of tetrahedra (trigonal bipyramid with off-center cations). The cation in the tetrahedra is shifted from the center position towards the basal plane of O(3) triangle and thus, the tetrahedra tend towards GaO_5 trigonal bipyramids having distances to the apical atoms of 0.1955 nm and 0.2328 nm, respectively. The crystal parameters of the IGZO-21 single crystal are listed in Table 2.

Distinguish Ga^{3+} and Zn^{2+} by using XRD is difficult. ⁷¹Ga NMR and EXAFS revealed that Ga also occupied the octahedral sites found for $\text{InGaO}_3(\text{ZnO})_1$.⁷⁰ Mixed occupation of Ga and Zn in the two layers of tetrahedra was also demonstrated. The distribution of the cations on the octahedral, trigonal-bipyramidal and tetrahedral sites was made in concordance with the results of HRTEM. Fig. 16(b)



shows a HAADF image with well-resolved cation columns in the structure model derived from single crystal XRD (left) and extracted EDX line-scan profiles (right) (In–L (yellow), Ga–K (blue) and Zn–K (cyan) as overlay). The layer stacking of these blocks is seen in Fig. 16(b). The GaO₃ trigonal bipyramid is sandwiched between InO₆ layers. Two Zn/Ga(2)O₄ tetrahedra were also found to be sandwiched between InO₆ layers.

A refinement of the single crystal data with a mixed occupation of the octahedral site by In and Ga led to a slight improvement of the *R*-values and to an occupation of the octahedra with 91%–In and 9%–Ga. All distances and angles in (InGaO₃)₂ZnO agreed with previously reported distances and angles from XRD data for InGaO₃(ZnO)_{*n*} compounds with *n* = 1, 2, 3 and 4 (ref. 46, 48, 49 and 68) and the only exception was the Ga(2) coordination. The Ga atom in the wurtzite slabs shows a close to ideal trigonal bipyramid and is in the center of the trigonal O(4) plane with distances of 0.191317(6) nm to the O(4) plane and 0.1999(3) nm to the apical O(3) site. The site symmetry of Ga in the wurtzite slabs in InGaO₃(ZnO)_{*n*} with *n* = 2, 4 also gave ideal angles for a trigonal bipyramid. However, the distances to the apical oxygen positions were elongated to 0.215 nm, while the Ga–O distances of 0.191 nm in the trigonal plane are the similar to those in (InGaO₃)₂ZnO.

To determine the cation distribution at the individual polyhedral layers of (InGaO₃)₂ZnO, high resolution SEM and TEM studies were conducted as shown in Fig. 16(b). Because of the dependence of intensity on the atomic number *Z'*, the In cations (*Z'* = 49) are the brightest HAADF signal. However, the layers of Ga between the two layers of InO₆ appear more intensely than the pair of Ga/Zn layers. This could be explained by “crosstalk” with the adjacent In atoms and/or by the shape of the electron probe. EDX analysis and the findings of Narendranath *et al.*,⁷⁰ meant that we could not exclude the idea that Ga partially occupies the octahedral layers. Additionally, atomic resolution imaging in the electron microscope suggested a perfect periodic stacking of the atomic layers.

5 Conclusion

We introduced the recent advances in growing single crystals of the homologous (InGaO₃)_{*m*}(ZnO)_{*n*} series. The detailed crystal structures were determined for (InGaO₃)_{*m*}(ZnO)_{*n*} (*n* = 1–4). Large single crystals facilitate experiments that cannot be conducted by using polycrystalline or thin films. Our group declared that the OFZ method in applying air pressure is an efficient tool for growing large single crystals of IGZO. We noted that excess ZnO is also essential, because the gas pressure is effective but not suppress ZnO vaporization completely.

Furthermore, understanding the anisotropy of the electrical conductivity of InGaO₃(ZnO)_{*n*} is critical for clarifying the mechanism of the conductivity of the IGZO series. Large single crystals facilitate experimentation for

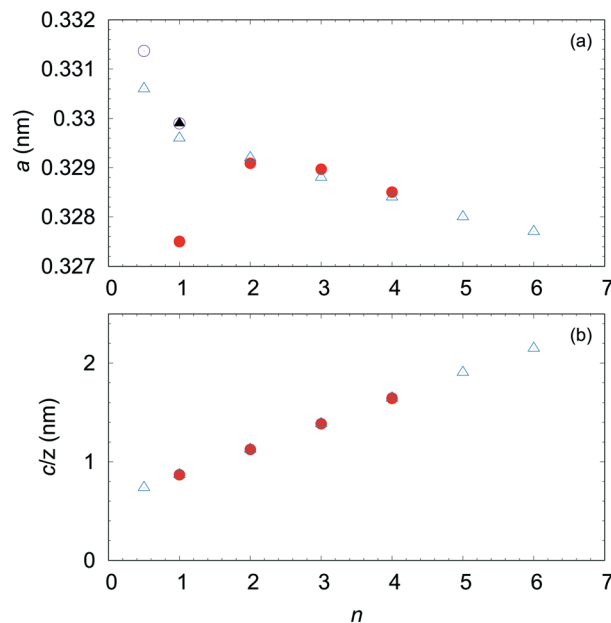


Fig. 18 (a) Lattice parameter *a* and (b) layer width *c/z* of homologous compounds InGaO₃(ZnO)_{*n*} as function of *n*. Open triangle symbols represent lattice constants estimated from the polycrystalline sample.²³ Solid triangles represent small single crystal obtained from high pressure (2 GPa).⁴⁵ Open circle symbols represent estimates from single crystals grown by the OFZ method.⁵⁵ Solid circles represent estimates from small single crystals.^{46,48,49}

determining the transport properties due to the structures of these compounds. Fig. 18 shows the length of the *a*-axis and *c*-axis/*Z* of InGaO₃(ZnO)_{*n*}. Because *Z* is the number of formula units in the structural unit cell, *c/z* shows the In–In distance along the *c*-axis in InGaO₃(ZnO)_{*n*} (*n* ≥ 1). The length of the *c*-axis in the InGaO₃(ZnO)_{*n*} series is expressed by *c/z* = 6.060 + 2.598 × *n*.¹⁷ The length determined for the single crystals was in good agreement with that of the polycrystalline sample. When In-5s mainly contributes to the electrical conductivity owing to an increasing distance of In–In layers with increasing *n*, we can predict that the conductivity along the *c*-axis becomes worse. In addition, IGZO-21 may have less anisotropy than IGZO-11 because it has a short average In–In distance. The relationship between the *c*-axis conductivity and the number of ZnO layers *n* remains to be clarified in experiments. To demonstrate the anisotropy, large single crystal growth of InGaO₃(ZnO)_{*n*} with larger *n* is necessary, and we are currently conducting these experiments.

The analysis of neutron diffraction results could lead to an understanding of the distribution of Zn/Ga in the single crystals. As distinguishing between Zn²⁺ and Ga³⁺ by performing XRD is impossible, an experiment to directly observe the distribution should be conducted instead. With the large single crystals we grew, the crystal structure can be analyzed using neutron diffraction, to directly observe the distribution. Although the Ga and Zn sites of *n* = 2, 4 are completely separated in the analysis using XRD, more studies are required to confirm this result.



Highlight

High-quality single crystals allow us to systematically evaluate the physical properties of the IGZO series, e.g., thermal properties (Seebeck coefficient, thermal conductivity and specific heat). The properties were measured using a polycrystalline sample and IGZO thin film along the *c*-axis,^{41,56,71,72} and not reported using the bulk single crystal. In addition, angle-resolved photoemission spectroscopy (ARPES) could be performed using our large single crystals. Using ARPES data, we may be able to determine both a detailed energy gap and the Fermi surface. These experiments are now in progress. As many different measurements can be performed on our large single crystals, rapid and significant progress is expected in the research on the IGZO series.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by JSPS KAKENHI Grant Number JP20K05306 and JP21K04909. A part of this work is also supported by Tokyo University's Science Grant for Young or Female Researchers, the Grants for Young Researcher from JGC-S Scholarship Foundation, Iketani Science and Technology Foundation, and Murata Science Foundation. Crystal structures schematics were produced using VESTA.⁷³

References

- 1 K. Nomura, H. Ohta, K. Ueda, T. Kamiya, M. Hirano and H. Hosono, *Science*, 2003, **300**, 1269.
- 2 K. Nomura, H. Ohta, A. Takagi, T. Kamiya, M. Hirano and H. Hosono, *Nature*, 2004, **432**, 488–492.
- 3 S. Amano, H. Harada, K. Akimoto, J. Sakata, T. Nishi, K. Moriya, K. Wakimoto, J. Koyama, S. Yamazaki, Y. Oikawa, T. Ikeyama and M. Sakakura, *Dig. Tech. Pap. – Soc. Inf. Disp. Int. Symp.*, 2010, **41**, 626.
- 4 T. Nishi, K. Moriya, S. Fukai, Y. Kubota, K. Akimoto, J. Sakata, K. Kusunoki, R. Arasawa, K. Wakimoto, J. Koyama, S. Yamazaki, Y. Oikawa, K. Okazaki and M. Sakakura, *Dig. Tech. Pap. – Soc. Inf. Disp. Int. Symp.*, 2010, **41**, 1685.
- 5 K. Toyotaka, K. Kusunoki, T. Nagata, Y. Hirakata, K. Wakimoto, J. Koyama, S. Yamazaki, R. Sato, K. Okazaki and M. Sakakura, *Jpn. J. Appl. Phys.*, 2011, **50**, 03CC09.
- 6 M. Orita, M. Takeuchi, H. Sakai and H. Tanji, *Jpn. J. Appl. Phys.*, 1995, **34**, L1550–L1552.
- 7 K. Kato, Y. Shionoiri, Y. Sekine, K. Furutani, T. Hatano, T. Aoki, M. Sasaki, H. Tomatsu, J. Koyama and S. Yamazaki, *Jpn. J. Appl. Phys.*, 2012, **51**, 021201.
- 8 N. Yamada, R. Ino, H. Tomura, Y. Kondo and Y. Ninomiya, *Adv. Electron. Mater.*, 2017, **3**, 1700298.
- 9 S. Yamazaki, J. Koyama, Y. Yamamoto and K. Okamoto, *Dig. Tech. Pap. – Soc. Inf. Disp. Int. Symp.*, 2012, **43**, 183–186.
- 10 S. Yamazaki, T. Hirohasi, M. Takahashi, S. Adachi, M. Tsubuku, J. Koezuka, K. Okazaki, Y. Kanzaki, H. Matsukizono, S. Kaneko, S. Mori and T. Matsuo, *J. Soc. Inf. Disp.*, 2014, **22**(1), 55–67.
- 11 K. Kato, Y. Shionoiri, Y. Sekine, K. Furutani, T. Hatano, T. Aoki, M. Sasaki, H. Tomatsu, J. Koyama and S. Yamazaki, *Jpn. J. Appl. Phys.*, 2012, **51**, 21201.
- 12 Y. Sekine, K. Furutani, Y. Shionoiri, K. Kato, J. Koyama and S. Yamazaki, *ECS Trans.*, 2011, **37**, 77–88.
- 13 M. Tsubuku, D. Matsubayashi, T. Takeuchi, R. Honda, T. Murakawa, S. Matsuda, Y. Yamamoto and S. Yamazaki, *ECS Trans.*, 2015, **67**, 17–22.
- 14 *Physics and Technology of Crystalline Oxide Semiconductor CAAC-IGZO: Application to Displays*, ed. S. Yamazaki and T. Tsutsui, Wiley, U.K., 2017.
- 15 N. Kimizuka and T. Mohri, *J. Solid State Chem.*, 1985, **60**, 382–384.
- 16 T. Moriga, D. R. Kammler and T. O. Mason, *J. Am. Ceram. Soc.*, 1999, **82**, 2705–2710.
- 17 N. Kimizuka, M. Isobe and M. Nakamura, *J. Solid State Chem.*, 1995, **116**, 170–178.
- 18 C. Li, Y. Bando, M. Nakamura, M. Onoda and N. Kimizuka, *J. Solid State Chem.*, 1998, **139**, 347–355.
- 19 M. Nakamura, N. Kimizuka and T. Mohri, *J. Solid State Chem.*, 1991, **93**, 298–315.
- 20 N. Kimizuka, T. Mohri, Y. Matsui and K. Siratori, *J. Solid State Chem.*, 1988, **74**, 98–109.
- 21 M. Nakamura, N. Kimizuka and T. Mohri, *J. Solid State Chem.*, 1989, **81**, 70.
- 22 J. Iida, S. Takekawa and N. Kimizuka, *J. Cryst. Growth*, 1990, **102**, 398–400.
- 23 M. Nakamura, N. Kimizuka and T. Mohri, *J. Solid State Chem.*, 1991, **93**, 298–315.
- 24 N. Kimizuka, T. Mohri and Y. Matsui, *J. Solid State Chem.*, 1988, **74**, 98–109.
- 25 N. Kimizuka and E. Takayama, *J. Solid State Chem.*, 1982, **41**, 166–173.
- 26 M. Isobe, N. Kimizuka, M. Nakamura and T. Mohri, *Acta Crystallogr., Sect. C: Cryst. Struct. Commun.*, 1994, **50**, 332–336.
- 27 S. Adams, *Acta Crystallogr., Sect. B: Struct. Sci.*, 2001, **57**(2001), 287, <http://kristall.unimki.gwdg.de/softbv/index.html>.
- 28 N. Kimizuka, A. Takenaka, Y. Sasada and T. Katsura, *Solid State Commun.*, 1974, **15**, 1199–1201.
- 29 K. Kato, I. Kawada, N. Kimizuka and T. Katsura, *Z. Kristallogr.*, 1976, **143**, 278–284.
- 30 R. Grajczyk and M. A. Subramanian, *Prog. Solid State Chem.*, 2015, **43**, 37–46.
- 31 N. Kimizuka and E. Takayama, *J. Solid State Chem.*, 1984, **53**, 217–226.
- 32 F. Lappe, A. Niggli, R. Nitsche and J. White, *Z. Kristallogr., Kristallgeom., Kristallphys., Kristallchem.*, 1962, **117**, 146.
- 33 S. Yonezawa, T. Kajikawa and Y. Maeno, *Phys. Rev. Lett.*, 2013, **110**, 077003.
- 34 S. Kittaka, Y. Aoki, Y. Shimura, T. Sakakibara, S. Seiro, C. Geibel, F. Steglich, H. Ikeda and K. Machida, *Phys. Rev. Lett.*, 2014, **112**, 067002.



- 35 Y. Matsumoto and S. Nakatsuji, *Rev. Sci. Instrum.*, 2018, **89**, 033908.
- 36 N. Kase, Y. Terui, T. Nakano and N. Takeda, *Phys. Rev. B*, 2017, **96**, 214506.
- 37 N. Kase, *J. Phys. Soc. Jpn.*, 2019, **88**, 041007.
- 38 H. Ohta, K. Nomura, M. Orita, M. Hirano, K. Ueda, T. Suzuki, Y. Ikuhara and H. Hosono, *Adv. Funct. Mater.*, 2003, **13**, 1398.
- 39 K. Nomura, T. Kamiya, H. Ohta, K. Ueda, M. Hirano and H. Hosono, *Appl. Phys. Lett.*, 2004, **85**, 1993.
- 40 H. Hosono, T. Kamiya and M. Hirano, *Bull. Chem. Soc. Jpn.*, 2006, **79**, 1–24.
- 41 H. J. Cho, Y. Wu, Y.-Q. Zhang, B. Feng, M. Mikami, W. Shin, Y. Ikuhara, Y.-M. Sheu, K. Saito and H. Ohta, *Adv. Mater. Interfaces*, 2021, **8**, 2001932.
- 42 S. Yamazaki, H. Suzawa, K. Inoue, K. Kato, T. Hirohashi, K. Okazaki and N. Kimizuka, *Jpn. J. Appl. Phys.*, 2014, **53**, 04ED18.
- 43 S. Yamazaki, S. Ohshita, M. Oota, H. Baba, T. Onuki, H. Kunitake, K. Ohshima, D. Shimada, H. Kimura, T. Murakawa, T. Atsumi and K. Kato, *Int. J. Ceram. Eng. Sci.*, 2019, **1**, 6–20.
- 44 D. M. Lynch, B. Zhu, B. D. A. Levin, D. A. Muller, D. G. Ast, R. G. Greene and M. O. Thompson, *Appl. Phys. Lett.*, 2014, **105**, 262103.
- 45 M. Nespolo, A. Sato, T. Osawa and H. Ohashi, *Cryst. Res. Technol.*, 2000, **35**, 151–165.
- 46 W. Assenmacher, G. Schnakenburg, Y. Michiue, Y. Kanke, N. Kimizuka and W. Mader, *J. Solid State Chem.*, 2014, **215**, 176–183.
- 47 J. P. Morniroli and J. W. Steeds, *Ultramicroscopy*, 1992, **45**, 219.
- 48 I. Keller, W. Assenmacher, G. Schnakenburg and W. Mader, *Z. Anorg. Allg. Chem.*, 2009, **635**, 2065–2071.
- 49 I. Keller and W. Mader, *Z. Anorg. Allg. Chem.*, 2010, **636**, 1045–1049.
- 50 S. Eichhorn and W. Mader, *J. Solid State Chem.*, 2016, **233**, 75–81.
- 51 S. Eichhorn, H. Schmid, W. Assenmacher and W. Mader, *J. Solid State Chem.*, 2017, **46**, 214–220.
- 52 J. Garling, W. Assenmacher, H. Schmid, P. Longo and W. Mader, *J. Solid State Chem.*, 2018, **258**, 809–817.
- 53 A. Magrez, J. P. Morniroli, M. Caldes, A. Marie, O. Joubert and L. Brohan, *J. Solid State Chem.*, 2003, **172**, 243.
- 54 S. Erichsen, W. Assenmacher, G. Schnakenburg and W. Mader, *Z. Anorg. Allg. Chem.*, 2014, **640**, 2817–2822.
- 55 Y. Tanaka, K. Wada, Y. Kobayashi, T. Fujii, S. J. Denholme, R. Sekine, N. Kase, N. Kimizuka and N. Miyakawa, *CrystEngComm*, 2019, **21**, 2985.
- 56 S. Préaud, C. Byl, F. Brisset and D. Berardan, *J. Am. Ceram. Soc.*, 2020, **103**, 3030–3038.
- 57 J. Tauc, R. Grigovici and A. Vancu, *Phys. Status Solidi*, 1966, **15**, 627.
- 58 B. D. Viezbicke, S. Patel, B. E. Davis and D. P. Birnie, *Phys. Status Solidi B*, 2015, **252**, 1700–1710.
- 59 T. Kamiya, K. Nomura and H. Hosono, *Phys. Status Solidi A*, 2009, **206**, 860–867.
- 60 M. Orita, H. Tanji, M. Mizuno, H. Adachi and I. Tanaka, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2000, **61**, 1811–1815.
- 61 H. H. Tippins, *Phys. Rev. A*, 1965, **140**, A316.
- 62 M. R. Lorenz, J. F. Woods and R. J. Gambino, *J. Phys. Chem. Solids*, 1967, **28**, 403.
- 63 N. Ueda, H. Hosono, R. Waseda and H. Kawazoe, *Appl. Phys. Lett.*, 1997, **70**, 3561.
- 64 A. Takagia, K. Nomura, H. Ohta, H. Yanagi, T. Kamiya, M. Hirano and H. Hosono, *Thin Solid Films*, 2005, **486**, 38–41.
- 65 T. Kamiya, K. Nomura and H. Hosono, *Sci. Technol. Adv. Mater.*, 2010, **11**, 044305.
- 66 Y. Kang, Y. Cho and S. Han, *Appl. Phys. Lett.*, 2013, **102**, 152104.
- 67 B. K. Ridley, *Quantum Process in Semiconductors*, Oxford University Press, New York, 3rd edn, 1993.
- 68 D. Lahr, W. Assenmacher, H. Schmid, N. Kimizuka, A. Kamai, Y. Tanaka, N. Miyakawa and W. Mader, *J. Solid State Chem.*, 2020, **288**, 121341.
- 69 K. Kato, I. Kawada, N. Kimizuka and T. Katsura, *Z. Kristallogr.*, 1975, **141**, 314–320.
- 70 S. B. Narendranath, A. K. Yadav, T. G. Ajithkumar, D. Bhattacharyya, S. N. Jha, K. K. Dey, T. Raja and R. N. Devi, *Dalton Trans.*, 2014, **43**, 2120–2126.
- 71 T. Yoshikawa, T. Yagi, N. Oka, J. Jia, Y. Yamashita, K. Hattori, Y. Seino, N. Taketoshi, T. Baba and Y. Shigesato, *Appl. Phys. Express*, 2013, **6**, 021101.
- 72 Y. Fujimoto, M. Uenuma, Y. Ishikawa and Y. Uraoka, *J. Electron. Mater.*, 2016, **45**, 1377–1381.
- 73 K. Momma and F. Izumi, *J. Appl. Crystallogr.*, 2011, **44**, 1272–1276.

