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Integration of nanosized ZIF-8 particles onto mesoporous TiO₂ nanobeads for enhanced photocatalytic activity†

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Hexavalent chromium, or Cr(vi), is a highly toxic contaminant in industrial wastewater that needs to be treated before being released. A Zeolitic Imidazolate Framework ZIF-8 was assembled on the surface of mesoporous TiO₂ beads to achieve a TiO₂/ZIF-8 heterostructure for photocatalytic reduction of Cr(vi). The TiO₂/ZIF-8 nanobeads exhibited remarkable photocatalytic activity of Cr(vi) reduction relative to that of pristine TiO₂ nanobeads attributed to the strong Cr(vi) adsorbing properties of ZIF-8 and improved charge transfer efficiency of TiO₂/ZIF-8. This work illustrates a strategy for enhanced photocatalysis by modifying the metal oxide catalyst with high surface area Zeolitic Imidazolate Framework to improve adsorption and charge transfer efficiency.

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Introduction

Hexavalent chromium, or Cr(vi), is one of the most toxic and dangerous contaminants in industrial wastewater due to its high toxicity and extreme solubility.^{1–3} Compared to Cr(vi), Cr(III) is less toxic and can be readily precipitated in neutral or alkaline solution. The preferred treatment of Cr(vi) in wastewater is simultaneous detoxification of Cr(vi) to Cr(III) and adsorption of the latter from water.⁴ Among different detoxifying treatments (such as chemical precipitation, ion exchange, membrane separation, adsorption and photocatalysis), photocatalysis has been proven to be a useful approach in reducing Cr(vi) to Cr(III).^{5–7} So far, most reported photocatalysts have been based on semiconductors such as TiO₂,^{5–7} CdS,⁸ SnS₂,⁹ and Ag₂S.¹⁰ It is known that reactant adsorption or permeation plays an important role in quantum efficiency of the photocatalysis process.^{11–13} Increasing the surface area of photocatalysts has been proven to be an effective way to provide more reactive sites and better adsorption properties and thus improve the photocatalytic activity.¹¹ However, the limited surface area of traditional semiconductor photocatalysts restricts their photocatalytic activity.

Metal–organic frameworks (MOFs) have demonstrated wide applications in molecular recognition, gas separation, catalysis, and drug delivery due to their high surface area, large pore size, tunability and well-defined nanometer-scale cavity, and chemical tailor-ability.¹⁴ More specifically, MOFs have been

considered as promising candidate materials for photocatalytic reaction, such as water splitting,^{15,16} CO₂ reduction,^{17–19} photo-degradation of organic pollutants,^{20–22} photocatalytic oxidation of alcohol²³ and photocatalytic reduction of Cr(vi).^{24–26} For instance, Li and his coworkers synthesized an amine-functionalized titanium MOFs (NH₂-MIL-125(Ti)) photocatalyst for CO₂ photoreduction in acetonitrile with triethanolamine as sacrificial agent under visible light.¹⁷ Mahata *et al.* used different MOFs based on Ni, Co, and Zn as photocatalysts to degrade organic dyes.²⁷ Furthermore, Fe(III)-based MOFs exhibited good photocatalytic activity for Cr(vi) reduction under visible-light.^{24,26} Although many kinds of MOFs have been proven to be novel photocatalysts, to date, compared with the traditional semiconductor photocatalysts, the photocatalytic efficiency of MOFs photocatalyst is still low due to the low efficiency in exciton generation and charge separation, which limit its practical application. On the other hand, the stability of MOFs during the photocatalytic reaction is an important issue that needed to be further consideration.

Incorporation the high surface area of MOFs and the superiority of MOFs in adsorption together with the high activity of semiconductor photocatalysts provide a promising strategy in photocatalytic application. Recently, semiconductor nano-structures/MOF composites have been intensively studied as photocatalysts (*e.g.* C₃N₄/Co-ZIF-9,²⁸ Cd_{0.2}Zn_{0.8}S@UiO-66-NH₂,²⁹ TiO₂@MIL-53,³⁰ TiO₂/ZIF-8,³¹ Cu₃(BTC)₂@TiO₂,³² ZnO@ZIF-8,^{33,34} Fe₃O₄@MIL-101,³⁵ UiO-66/-C₃N₄,³⁶ ZIF-8/Zn₂GeO₄,³⁷ CPO-27-Mg/TiO₂³⁸). It has been demonstrated that the combination of semiconductor and MOF showed performance exceed that of the individual component due to their synergistic effect. Besides increasing the surface area of the composites, charge transfer can occur between photoexcited inorganic semiconductors and MOFs, which substantially suppresses

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electron-hole recombination in the semiconductor and supplies long-life time electrons for photocatalytic reaction.^{21,32} Despite the rapidly growing interest in semiconductor nanostructures/MOF composites, the studies on these MOF-based hybrid photocatalysts are still scarce and in the infancy.

In this paper, we report the design and fabrication of TiO₂/ZIF-8 hybrid photocatalyst. ZIF-8 was chosen as metal-organic framework material owing to its large surface area, excellent thermal and chemical stability.³⁹ The new hybrid TiO₂/ZIF-8 nanobeads are produced by decorating the surface of as-synthesized mesoporous TiO₂ spheres with ZIF-8 nanoparticles. The hybrid TiO₂/ZIF-8 nanobeads were characterized by their morphology, absorption capacity, surface area and photocatalytic reduction of Cr(IV). Compared to TiO₂ beads with similar sizes, the catalytic activity of TiO₂/ZIF-8 in reducing Cr(VI) to Cr(III) under 300 mW cm⁻² full-spectrum light irradiation at room temperature is significantly enhanced. It was found that the enhanced photocatalytic properties of TiO₂/ZIF-8 spheres were based on strong Cr(VI) adsorbing properties of ZIF-8 as well as more efficient charge transfer. The stability and long term performance of the new catalyst was investigated by means of its transient photocurrent response to periodic irradiation and repeated photocatalytic reduction experiment of synthetic Cr(IV) waste solution using the recovered photocatalyst. To the best of our knowledge, this work represents the design of hybrid TiO₂/ZIF-8 nanosphere and its use for Cr(IV) photoreduction. We hope that the current work could inspire growing interest on the fabrication of other high-performance semiconductor/MOFs composite by taking the advantage of MOFs.

Experimental

Preparation of the mesoporous TiO₂ beads

TiO₂ colloids were prepared with a modified strategy as previously reported.⁴⁰ Typically, 0.2 mL of tetrabutyl titanate (TBT) was added to 10 mL of ethylene glycol under stirring. After magnetically stirred for 24 h at room temperature, the solution then poured into 100 mL acetone (containing 0.3% water) under vigorous stirring for about 10 min and then kept at rest for 1 h. The white precipitate was separated by centrifugation and washed several times with ethanol and deionized water to remove excess ethylene glycol from the surfaces of the particles. The obtained powder mixed with 10 mL deionized water and then transferred to a stainless Teflon-lined autoclave of 25 mL inner volume. The hydrothermal synthesis was performed under an auto-generated pressure at 180 °C for 24 h in an electric oven, followed by cooling naturally to room temperature. The white powder was collected by centrifugation and washed thoroughly with alcohol for several times, then dried at 60 °C for 12 h.

Synthesis of ZIF-8 nanocrystals

ZIF-8 nanocrystals were prepared according to our previously reported work.³⁷ Typically, 0.587 g of Zn(NO₃)₂·6H₂O and 1.298 g of 2-methylimidazole (Hmim) were dissolved in 40 mL of methanol (MeOH) separately. The Zn²⁺ solution was added to

the Hmim solution under stirring. The mixture was stirred at room temperature for 2 h. The ZIF-8 nanocrystals were separated from the milky colloidal dispersion by centrifugation and washed with anhydrous ethanol for three times, and dried at 60 °C.

Preparation of TiO₂/ZIF-8 heterostructure

Typically, 0.4 g as-synthesized mesoporous TiO₂ beads and 0.587 g Zn(NO₃)₂·6H₂O were added to 40 mL of MeOH and stirred for 1 h. 1.298 g Hmim was dissolved in 40 mL of MeOH to get clear solution, then poured into the former solution under stirring and further stirred for another 1 h at room temperature (theoretical mass ratio 1 : 2 of ZIF-8 to TiO₂). The obtained precipitate was collected by repeated centrifugation and washing with water and alcohol for several times before drying at 60 °C for 12 h.

Characterization

X-ray diffraction (XRD) measurements were performed on a Philips PW1140/90 diffractometer with CuKα radiation (25 mA and 40 kV) at a scan rate of 2° per min with a step size of 0.02°. The morphology of the samples was observed by scanning electron microscopy (FEI Nova NanoSEM 450) and transmission electron microscopy (JEOL 3010, Japan). The specific surface area of the samples and nitrogen adsorption-desorption isotherms were measured by nitrogen sorption at 77 K on surface area and porosity analyzer (Micromeritics ASAP 2020) and calculated by the BET method. Prior to BET surface area measurement, all the samples were degassed at 300 °C for 4 h. Thermal gravimetric (TG) measurement was taken on a thermal instrument (Shimadzu Corp. Japan) with a heating rate of 10 °C min⁻¹. 6 mg samples were filled into alumina crucibles and heated in a flow of air with a ramp rate of 5 °C min⁻¹ from room temperature up to 700 °C. UV visible light adsorption spectra was obtained using a UV-Vis spectrophotometer (UV-2600 Shimadzu, Japan).

Photoelectrochemical measurement

To prepare the photoelectrodes, 10 mg TiO₂ or TiO₂/ZIF-8 was added into 1 mL of ethanol. The mixture was stirred for 1 h to ensure that the particles were uniformly dispersed in the solution. 10 μL of this prepared solution was dropped onto the FTO glass substrate (exposed area of 1.0 × 1.0 cm²), and then dried under vacuum condition for 1 h at 60 °C. This step was repeated five times to ensure a uniform coverage of photocatalysts on ITO. Photocurrent measurements were performed in a three-electrode cell using an electrochemical analyzer (CHI-630D, Shanghai Chenhua). The electrolyte was 0.1 M Na₂SO₄ aqueous solution (pH = 7.0). The TiO₂ or TiO₂/ZIF-8 photoanodes were used as working electrodes; a Pt wire served as a counter electrode, a Ag/AgCl electrode was used as the reference electrodes. The electrodes were irradiated through the conducting glass by a 300 W Xe lamp. Cyclic voltammetry was performed with a scan rate of 30 mV s⁻¹.



Photocatalytic measurement

All the photocatalysts were activated in 200 °C for 4 h before use. The photocatalytic activities of photocatalysts were evaluated by photocatalytic reduction of Cr(vi) under full-light irradiation of a 300 W Xe arc lamp with a cooling water filter. Potassium dichromate ($K_2Cr_2O_7$) was selected as the Cr(vi) compound. The photocatalytic reduction of Cr(vi) was performed at room temperature in a quartz reactor containing 20 mg photocatalyst and 40 mL of Cr(vi) solution (20 mg L⁻¹ based on Cr in a dilute $K_2Cr_2O_7$ solution, pH = 7). Nitrogen was then purged through the system, followed by the introduction of 5 mg hole scavenger (ammonium oxalate). As we known, Cr(vi) is often discharged together with hazardous organics from industrial wastewater, these hazardous organics can be used as hole scavenger.

Prior to irradiation, the suspension was stirred for 40 min in dark to reach adsorption equilibrium and then was exposed to light irradiation. The concentrations of Cr(vi) was measured by UV-Vis maximum absorbance at 365 nm. The photocatalytic efficiency was determined by the following equation, photocatalytic efficiency = C/C_0 , where C is the Cr(vi) concentration at time of measurement and C_0 is the initial Cr(vi) concentration.

Results and discussion

The XRD patterns of the as-synthesized samples are shown in Fig. 1. The as-synthesized TiO₂ precursor was amorphous in structure with essentially no diffraction peaks. After hydrothermal synthesis, all the diffraction peaks of the sample can be assigned to the anatase phase of TiO₂ (JCPDS no. 21-1272) with lattice constants of $a = b = 3.785 \text{ \AA}$ and $c = 9.514 \text{ \AA}$, $\alpha = \beta = \gamma = 90^\circ$. The broadening of the diffraction peaks indicates the crystallite size of this TiO₂ sample is small. The crystallite size D was calculated using the Scherrer equation:

$$D = 0.9\lambda/\beta \cos \theta$$

where, λ is wavelength of X-ray (1.5405 Å), β corresponds to full width at half maxima (FWHM), θ is the diffraction angle. The estimated size of this TiO₂ sample was about 18 nm. The diffraction peaks of the as-synthesized ZIF-8 (Fig. 1c) fits well with ZIF-8 phase with a cubic space group ($I\bar{4}3m$)⁴¹ by comparing with the simulated ZIF-8 and published pattern,¹⁴ which indicates that the product is pure-phase ZIF-8. After growing of ZIF-8 particles on the TiO₂ beads, characteristic diffraction peaks of ZIF-8 was observed ($2\theta = 7.4^\circ$ and 16.6°). Both distinct characteristic peaks of ZIF-8 and anatase TiO₂ were observed, confirming the purity of the product.

The FE-SEM images taken at different magnification show that the prepared TiO₂ colloids precursor is entirely composed of uniform, spherical crystallites with a diameter of 300 ± 50 nm. The surfaces of these beads are very smooth without obvious granular features. After hydrothermal treatment, the surface of TiO₂ colloids particles were roughened as a result of crystallization. The TiO₂ beads are monodisperse beads with a diameter of 250 ± 50 nm. The high magnification SEM image (Fig. 2d) demonstrates that these TiO₂ beads are composed of



Fig. 1 Powder X-ray diffraction patterns of (a) TiO₂ colloids, (b) TiO₂ beads, (c) pure ZIF-8, and (d) TiO₂/ZIF-8 spheres.

uniform nanocrystals with a size of about 20 ± 5 nm, consistent with the XRD result.

After *in situ* growth of ZIF-8 on TiO₂ beads, a layer of ZIF-8 polyhedrons were grown on the surface of the TiO₂ beads (Fig. 3a). The EDS results also confirm the presence of C, N, O, Ti and Zn in the TiO₂/ZIF-8 hybrid spheres (Fig. S1†). A high magnification FE-SEM image of the sample reveals that there are many polyhedrons on the surface of TiO₂ beads, which indicates that the ZIF-8 particles were successfully decorated onto the beads to form TiO₂/ZIF-8 hybrid nanospheres (Fig. 3b). The ZIF-8 polyhedrons on the TiO₂ beads are 40–70 nm in size, which are similar with those ZIF-8 crystals synthesized in the absence of TiO₂ beads (Fig. S2†). Discrete ZIF-8 polyhedrons of approximately 50 nm were decorated on the surface of TiO₂ beads, as shown in the TEM images of TiO₂@ZIF-8 nanobeads (Fig. 3c and d). By means of UV-vis spectroscopy (Fig. S3†), it was observed that the spectrum absorption edge of TiO₂/ZIF-8 was slightly blue-shifted compared to the TiO₂.



Fig. 2 FE-SEM images of TiO₂ colloids (a), (b), and TiO₂ beads (c), (d).





Fig. 3 (a) and (b) FE-SEM images, and (c) and (d) TEM images of the $\text{TiO}_2/\text{ZIF-8}$ hybrid beads.

Fig. 4 shows the nitrogen adsorption–desorption isotherms of TiO_2 beads, ZIF-8 and $\text{TiO}_2/\text{ZIF-8}$ hybrid nanospheres. The TiO_2 beads with surface decorated with ZIF-8 nanoparticles exhibit higher BET area due to the high surface area of ZIF-8 as compared to TiO_2 beads. The specific BET surface area of TiO_2 beads, ZIF-8 and $\text{TiO}_2/\text{ZIF-8}$ are about $250 \text{ m}^2 \text{ g}^{-1}$, $1058 \text{ m}^2 \text{ g}^{-1}$ and $397 \text{ m}^2 \text{ g}^{-1}$, respectively. ZIF-8 (Fig. 4a) shows a typical type I nitrogen adsorption–desorption isotherm,⁴² which fit well with the microporous frameworks of ZIF-8.³⁷ The obvious hysteresis loop at high relative pressure (0.8–0.9) indicates the existence of textural macroporosity formed by packing of ZIF-8 crystals, consistent with some previous literature.^{34,37,43} On the other hand, TiO_2 beads exhibits a type IV isotherm with an H_3 hysteresis loop according to the Brunauer–Deming–Deming–Teller (BDDT) classification, indicating a mesoporous characteristic.^{44,45} The initial adsorption of $\text{TiO}_2/\text{ZIF-8}$ is higher than that of TiO_2 (Fig. 4b and c), indicating that micropores exist in



Fig. 4 N_2 adsorption–desorption isotherms of (a) ZIF-8, (b) $\text{TiO}_2/\text{ZIF-8}$, (c) TiO_2 beads.

$\text{TiO}_2/\text{ZIF-8}$ hybrid nanospheres,⁴² which can be attributed to the frameworks of ZIF-8 particles. The adsorption of $\text{TiO}_2/\text{ZIF-8}$ at high pressure corresponds to composite spherical morphology of the sample, resulting from mesoporous TiO_2 beads and microporous ZIF-8 particles.

To determine the content of ZIF-8 in $\text{TiO}_2/\text{ZIF-8}$ composites, TG-DTA analysis of these three samples was carried out in air atmosphere, as shown in Fig. 5. The TG of ZIF-8 exhibits a 75% total mass loss up to ca. 700°C in two steps. For the bare TiO_2 beads, TG analysis shows only 1.7% mass loss is observed up to 700°C , which can be attributed to the desorption of H_2O



Fig. 5 TG (black curve) and DTA (blue curve) analysis of (a) ZIF-8, (b) $\text{TiO}_2/\text{ZIF-8}$, (c) TiO_2 beads.



adsorbed on the surfaces of the beads. The weight loss of $\text{TiO}_2/\text{ZIF-8}$ at the beginning stage ($\sim 1.5\%$) is the vaporization of adsorbed water or methanol in the sample. The obvious weight loss occurred in the range of $250\text{--}580\text{ }^\circ\text{C}$ owing to rapid decomposition of ZIF-8 molecules. When the temperature reaches $600\text{ }^\circ\text{C}$, ZIF-8 molecules are transformed to ZnO completely. A total mass loss of $\sim 14.18\%$ in the measured temperature range means that the prepared sample contains around 20 wt% ZIF-8 and 80 wt% TiO_2 on the basis of the weight loss of pure ZIF-8 in air ($\sim 75\%$). The actual amount of ZIF-8 grown on TiO_2 is smaller than the theoretical value (theoretical ZIF-8/ TiO_2 mass ratio of 1 : 2). This can be explained by the fact that apart from ZIF-8 growth on TiO_2 , ZIF-8 also crystallizes in the solution, consuming some of its precursors.

The photocatalysis performance of $\text{TiO}_2/\text{ZIF-8}$ in $\text{Cr}(\text{vi})$ aqueous solutions is shown in Fig. 6a, as compared with that of TiO_2 and ZIF-8 nanoparticles. Under light irradiation, TiO_2 in the $\text{TiO}_2/\text{ZIF-8}$ hybrid nanospheres was excited. The electrons from the excited TiO_2 can transfer to the ZIF-8, and reduce the $\text{Cr}(\text{vi})$ to $\text{Cr}(\text{iii})$. The control $\text{Cr}(\text{vi})$ reduction experiment performed in the absence of the photocatalyst showed no obvious decrease of $\text{Cr}(\text{vi})$ concentration, proving that the $\text{Cr}(\text{vi})$

reduction reaction is driven by light with the photocatalyst. Although several MOFs have recently been shown to have good photocatalytic activity under light illumination,^{17,46} there was no remarkable change of $\text{Cr}(\text{vi})$ concentration detected for 60 min irradiation using pure ZIF-8 as photocatalyst. Only $\sim 10\%$ of $\text{Cr}(\text{vi})$ is removed after 60 min irradiation, which is due to the cooperation of weak photocatalytic activity and adsorption of ZIF-8. The pure TiO_2 beads showed photocatalytic reduction of $\text{Cr}(\text{vi})$ of $\sim 80\%$ after 60 min irradiation. After incorporation of ZIF-8 nanoparticles, the photocatalytic activity of $\text{TiO}_2/\text{ZIF-8}$ hybrid spheres towards $\text{Cr}(\text{vi})$ is remarkably enhanced. For the photocatalytic reduction of $\text{Cr}(\text{vi})$, around 99% of $\text{Cr}(\text{vi})$ was removed after 60 min irradiation. The removal efficiency is better than that of TiO_2 and ZIF-8, exhibiting enhanced photocatalysis properties. First, the high BET surface and the porous structure of ZIF-8 are beneficial to absorption and permeation of $\text{Cr}(\text{vi})$ species. Second, in neutral conditions ($\text{pH} = 7$), ZIF-8 particles exhibit positive charge⁴⁷ and the main species of $\text{Cr}(\text{vi})$ in neutral aqueous solution is CrO_4^{2-} anion.^{34,48} Therefore, CrO_4^{2-} could be effectively adsorbed around the $\text{TiO}_2/\text{ZIF-8}$ photocatalyst by electrostatic interaction, which lead to better photocatalytic activity of $\text{Cr}(\text{vi})$ on $\text{TiO}_2/\text{ZIF-8}$. Time-dependent evolution of the photocatalytic removal of $\text{Cr}(\text{vi})$ by $\text{TiO}_2/\text{ZIF-8}$ is characterized by optical absorption measurements, as shown in Fig. 6b. One can see the absorption peak at 365 nm ascribed to $\text{Cr}(\text{vi})$ ion^{34,49} decreases rapidly and remarkably with the increase of time and almost disappears after 60 min, which demonstrates the photocatalytic reduction of $\text{Cr}(\text{vi})$ over $\text{TiO}_2/\text{ZIF-8}$. The final concentration of the Cr ions remaining in the solution after 60 min of UV light irradiation drops to 0.06 mg L^{-1} from the initial 20 mg L^{-1} , and the solution becomes colorless, indicating that the $\text{TiO}_2/\text{ZIF-8}$ can almost thoroughly remove $\text{Cr}(\text{vi})$ ions from the wastewater.

It is worthy to note that $\text{TiO}_2/\text{ZIF-8}$ was stable during the whole photocatalytic reduction process. The XRD patterns before and after photocatalysis are almost same, as shown in Fig. S5. † In addition, no obvious morphology change was observed after photocatalysis process (Fig. S6†). The stability and reusability of the photocatalysts are very important for photocatalytic application. To further evaluate the long term performance of the $\text{TiO}_2/\text{ZIF-8}$ photocatalysts, photocatalytic reduction of $\text{Cr}(\text{vi})$ by irradiation using the recycled photocatalysts was conducted. After each reaction, the used photocatalyst was recovered by filtration, washed with water and ethanol, and then dried under vacuum. Fig. 7b shows the transient photocurrent responses of TiO_2 beads and $\text{TiO}_2/\text{ZIF-8}$ under intermittent light illumination. The photocurrent density of TiO_2 beads, ZIF-8 and $\text{TiO}_2/\text{ZIF-8}$ are about 4.3 and $3.4\text{ }\mu\text{A cm}^{-2}$, respectively. The incorporation of ZIF-8 particles on TiO_2 enhance the photocurrent significantly. This indicates that the separation efficiency of photoinduced electron-hole (e^- - h^+) pairs and the lifetime of the photogenerated charge carriers are improved. For semiconductor/MOF photocatalysts, Jiang proved that the photogenerated electrons can be effectively transferred from the semiconductor to the MOF, which not only facilitates charge separation in the semiconductor but supplies energetic electrons to molecules adsorbed on the MOF.^{21,32} So,



Fig. 6 (a) Reduction profiles of photocatalytic reduction of $\text{Cr}(\text{vi})$ over various photocatalysts; reaction condition: 20 mg photocatalyst, 40 mL of 20 mg L^{-1} $\text{Cr}(\text{vi})$, $\text{pH} = 7$; (b) absorption spectra of the $\text{Cr}(\text{vi})$ solution after photocatalytic reaction for different times over $\text{TiO}_2/\text{ZIF-8}$.



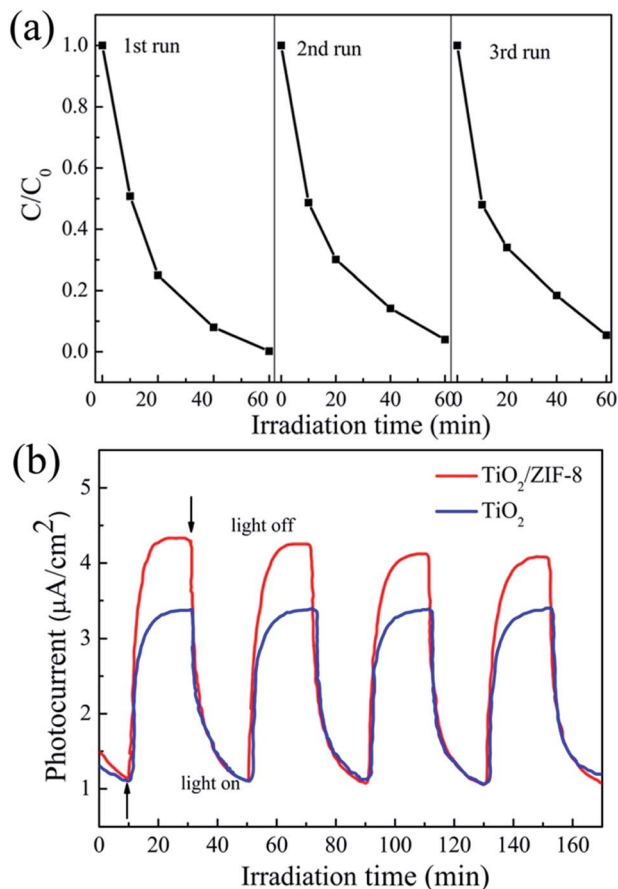


Fig. 7 (a) Recycling test on $\text{TiO}_2/\text{ZIF-8}$ for photocatalytic reduction of $\text{Cr}(\text{VI})$. (b) Transient photocurrent response of $\text{TiO}_2/\text{ZIF-8}$ and TiO_2 in $0.1 \text{ M Na}_2\text{SO}_4$ aqueous solution.

the higher photocatalytic activity of the $\text{TiO}_2/\text{ZIF-8}$ toward reduction of $\text{Cr}(\text{VI})$ relative to that of the TiO_2 sample should be attributed to synergistic combination of ZIF-8 and TiO_2 nanobeads. Firstly, ZIF-8 can effectively adsorb the dissolved $\text{Cr}(\text{VI})$ in the water solution; secondly, the charge transfer of $\text{TiO}_2/\text{ZIF-8}$ is more efficient than TiO_2 alone. Both the improvement lead to higher photocurrent density of $\text{TiO}_2/\text{ZIF-8}$ nanobeads.

Conclusions

In conclusion, $\text{TiO}_2/\text{ZIF-8}$ hybrid nanospheres have been successfully prepared by depositing nanosized ZIF-8 particles on the surfaces of mesoporous TiO_2 nanobeads. As a result, the obtained heterostructures show higher BET area due to the high surface area of ZIF-8 as compared to TiO_2 beads. TG-DTA analysis result indicate that the prepared sample contains around 20 wt% ZIF-8 and 80 wt% TiO_2 . The $\text{TiO}_2/\text{ZIF-8}$ hybrid nanospheres exhibited enhanced photocatalytic activity of $\text{Cr}(\text{VI})$ reduction than pristine TiO_2 beads, which can be attributed to the strong $\text{Cr}(\text{VI})$ adsorption property of ZIF-8 as well as more efficient charge transfer compared to pristine TiO_2 beads. It is anticipated that the strategy for implementing MOF structures in photocatalyst shown here can be applied to other catalytic systems for enhanced photocatalytic conversion efficiency.

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Notes and references

- 1 S. Rengaraj, S. Venkataraj, J.-W. Yeon, Y. Kim, X. Z. Li and G. K. H. Pang, *Appl. Catal., B*, 2007, **77**, 157–165.
- 2 D. Mohan and C. U. Pittman Jr, *J. Hazard. Mater.*, 2006, **137**, 762–811.
- 3 V. K. Gupta, A. Rastogi and A. Nayak, *J. Colloid Interface Sci.*, 2010, **342**, 135–141.
- 4 D. Shao, X. Wang and Q. Fan, *Microporous Mesoporous Mater.*, 2009, **117**, 243–248.
- 5 F. D. Kuanping Gong, Z. Xia, M. Durstock and L. Dai, *Science*, 2009, **323**, 760–764.
- 6 P. Mohapatra, S. K. Samantaray and K. Parida, *J. Photochem. Photobiol., A*, 2005, **170**, 189–194.
- 7 L. Wang, S. Z. Kang, X. Li, L. Qin, H. Yan and J. Mu, *Mater. Chem. Phys.*, 2016, **178**, 190–195.
- 8 X. Liu, L. Pan, T. Lv, G. Zhu, Z. Sun and C. Sun, *Chem. Commun.*, 2011, **47**, 11984–11986.
- 9 Y. C. Zhang, J. Li, M. Zhang and D. D. Dionysiou, *Environ. Sci. Technol.*, 2011, **45**, 9324–9331.
- 10 W. Yang, L. Zhang, Y. Hu, Y. Zhong, H. B. Wu and X. W. Lou, *Angew. Chem., Int. Ed.*, 2012, **51**, 11501–11504.
- 11 H. Tong, S. Ouyang, Y. Bi, N. Umezawa, M. Oshikiri and J. Ye, *Adv. Mater.*, 2012, **24**, 229–251.
- 12 X. Chang, T. Wang and J. Gong, *Energy Environ. Sci.*, 2016, **9**, 2177–2196.
- 13 W. Tu, Y. Zhou and Z. Zou, *Adv. Mater.*, 2014, **26**, 4607–4626.
- 14 K. S. Park, Z. Ni, A. P. Cote, J. Y. Choi, R. Huang, F. J. Uribe-Romo, H. K. Chae, M. O’Keeffe and O. M. Yaghi, *Proc. Natl. Acad. Sci. U. S. A.*, 2006, **103**, 10186–10191.
- 15 G. Huang, Y. Chen and H. Jiang, *Acta Chim. Sin.*, 2016, **74**, 113–129.
- 16 J. H. Cavka, S. Jakobsen, U. Olsbye, N. Guillou, C. Lamberti, S. Bordiga and K. P. Lillerud, *J. Am. Chem. Soc.*, 2008, **130**, 13850–13851.
- 17 Y. H. Fu, D. R. Sun, Y. J. Chen, R. K. Huang, Z. X. Ding, X. Z. Fu and Z. H. Li, *Angew. Chem., Int. Ed.*, 2012, **51**, 3364–3367.
- 18 H. Q. Xu, J. Hu, D. Wang, Z. Li, Q. Zhang, Y. Luo, S. H. Yu and H. L. Jiang, *J. Am. Chem. Soc.*, 2015, **137**, 13440–13443.
- 19 A. Dhakshinamoorthy, A. M. Asiri and H. Garcia, *Angew. Chem., Int. Ed.*, 2016, **55**, 5414–5445.
- 20 K. G. Laurier, F. Vermoortele, R. Ameloot, D. E. De Vos, J. Hofkens and M. B. Roeffaers, *J. Am. Chem. Soc.*, 2013, **135**, 14488–14491.
- 21 Y. Li, H. Xu, S. Ouyang and J. Ye, *Phys. Chem. Chem. Phys.*, 2016, **18**, 7563–7572.
- 22 D. Wang, M. Wang and Z. Li, *ACS Catal.*, 2015, **5**, 6852–6857.



- 23 W. G. Tian, C. Xiao, R. V. Maligal-Ganesh, X. Li and W. Huang, *Chem. Eng. Sci.*, 2015, **124**, 45–51.
- 24 L. Shi, T. Wang, H. Zhang, K. Chang, X. Meng, H. Liu and J. Ye, *Adv. Sci.*, 2015, **2**, 1500006.
- 25 R. Liang, L. Shen, F. Jing, W. Wu, Q. Na, L. Rui and W. Ling, *Appl. Catal., B*, 2015, **162**, 245–251.
- 26 R. Liang, F. Jing, L. Shen, N. Qin and L. Wu, *J. Hazard. Mater.*, 2015, **287**, 364–372.
- 27 P. Mahata, G. Madras and S. Natarajan, *J. Phys. Chem. B*, 2006, **110**, 13759–13768.
- 28 S. Wang, J. Lin and X. Wang, *Phys. Chem. Chem. Phys.*, 2014, **16**, 14656–14660.
- 29 Y. Su, Z. Zhang, H. Liu and Y. Wang, *Appl. Catal., B*, 2017, **200**, 448–457.
- 30 N. Chang, D. Y. He, Y. X. Li, Z. W. Tang and Y. F. Huang, *RSC Adv.*, 2016, **6**, 71481–71484.
- 31 X. Zeng, L. Huang, C. Wang, J. Wang, J. Li and X. Luo, *ACS Appl. Mater. Interfaces*, 2016, **8**, 20274–20282.
- 32 R. Li, J. Hu, M. Deng, H. Wang, X. Wang, Y. Hu, H. L. Jiang, J. Jiang, Q. Zhang, Y. Xie and Y. Xiong, *Adv. Mater.*, 2014, **26**, 4783–4788.
- 33 W. Zhan, Q. Kuang, J. Zhou, X. Kong, Z. Xie and L. Zheng, *J. Am. Chem. Soc.*, 2013, **135**, 1926–1933.
- 34 X. Wang, J. Liu, S. Leong, X. Lin, J. Wei, B. Kong, Y. Xu, Z. X. Low, J. Yao and H. Wang, *ACS Appl. Mater. Interfaces*, 2016, **8**, 9080–9087.
- 35 X. Yue, W. Guo, X. Li, H. Zhou and R. Wang, *Environ. Sci. Pollut. Res.*, 2016, **23**, 1–9.
- 36 R. Wang, L. Gu, J. Zhou, X. Liu, F. Teng, C. Li, Y. Shen and Y. Yuan, *Adv. Mater. Interfaces*, 2015, **2**, 1500037.
- 37 Q. Liu, Z. X. Low, L. Li, A. Razmjou, K. Wang, J. Yao and H. Wang, *J. Mater. Chem. A*, 2013, **1**, 11563–11569.
- 38 M. Wang, D. Wang and Z. Li, *Appl. Catal., B*, 2016, **183**, 47–52.
- 39 X. C. Huang, Y. Y. Lin, J. P. Zhang and X. M. Chen, *Angew. Chem., Int. Ed.*, 2006, **45**, 1557–1559.
- 40 X. Jiang, T. Herricks and Y. Xia, *Adv. Mater.*, 2003, **15**, 1205–1209.
- 41 U. P. N. Tran, K. K. A. Le and N. T. S. Phan, *ACS Catal.*, 2011, **1**, 120–127.
- 42 M. Kruk and M. Jaroniec, *Chem. Mater.*, 2001, **13**, 3169–3183.
- 43 M. He, J. Yao, Q. Liu, K. Wang, F. Chen and H. Wang, *Microporous Mesoporous Mater.*, 2014, **184**, 55–60.
- 44 S. C. Yan, S. X. Ouyang, J. Gao, M. Yang, J. Y. Feng, X. X. Fan, L. J. Wan, Z. S. Li, J. H. Ye, Y. Zhou and Z. G. Zou, *Angew. Chem., Int. Ed.*, 2010, **49**, 6400–6404.
- 45 X. W. Lou, D. Deng, J. Y. Lee and L. A. Archer, *J. Mater. Chem.*, 2008, **18**, 4397–4401.
- 46 C. Wang, Z. Xie, K. E. deKrafft and W. Lin, *J. Am. Chem. Soc.*, 2011, **133**, 13445–13454.
- 47 N. A. Khan, B. K. Jung, Z. Hasan and S. H. Jhung, *J. Hazard. Mater.*, 2015, **282**, 194–200.
- 48 X. Wang, *J. Mater. Chem.*, 2010, **20**, 8582–8590.
- 49 J. Yu, S. Zhuang, X. Xu, W. Zhu, B. Feng and J. Hu, *J. Mater. Chem. A*, 2014, **3**, 1199–1207.

