

Cite this: *Chem. Sci.*, 2018, 9, 8035

All publication charges for this article have been paid for by the Royal Society of Chemistry

Selective and catalytic carbon dioxide and heteroallene activation mediated by cerium N-heterocyclic carbene complexes†

Polly L. Arnold, ^{id}*^a Ryan W. F. Kerr, ^{id}^{ab} Catherine Weetman, ^{id}^a Scott R. Docherty, ^{id}^a Julia Rieb, ^{ac} Faye L. Cruickshank, ^a Kai Wang, ^a Christian Jandl, ^{ac} Max W. McMullon, ^{ac} Alexander Pöthig, ^{id}^c Fritz E. Kühn ^{id}*^c and Andrew D. Smith ^{id}*^b

A series of rare earth complexes of the form $\text{Ln}(\text{L}^{\text{R}})_3$ supported by bidentate *ortho*-aryloxoide–NHC ligands are reported ($\text{L}^{\text{R}} = 2\text{-O-}3,5\text{-}^t\text{Bu}_2\text{-C}_6\text{H}_2(1\text{-C}(\text{N}(\text{CH}_2)\text{N}(\text{R})))$); $\text{R} = ^i\text{Pr}, ^t\text{Bu}, \text{Mes}$; $\text{Ln} = \text{Ce}, \text{Sm}, \text{Eu}$). The cerium complexes cleanly and quantitatively insert carbon dioxide exclusively into all three cerium carbene bonds, forming $\text{Ce}(\text{L}^{\text{R}}\cdot\text{CO}_2)_3$. The insertion is reversible only for the mesityl-substituted complex $\text{Ce}(\text{L}^{\text{Mes}})_3$. Analysis of the capacity of $\text{Ce}(\text{L}^{\text{R}})_3$ to insert a range of heteroallenes that are isoelectronic with CO_2 reveals the solvent and ligand size dependence of the selectivity. This is important because only the complexes capable of reversible CO_2 -insertion are competent catalysts for catalytic conversions of CO_2 . Preliminary studies show that only $\text{Ce}(\text{L}^{\text{Mes}}\cdot\text{CO}_2)_3$ catalyses the formation of propylene carbonate from propylene oxide under 1 atm of CO_2 pressure. The mono-ligand complexes can be isolated from reactions using $\text{LiCe}(\text{N}^i\text{Pr}_2)_4$ as a starting material; LiBr adducts $[\text{Ce}(\text{L}^{\text{R}})(\text{N}^i\text{Pr}_2)\text{Br}\cdot\text{LiBr}(\text{THF})_2]$ ($\text{R} = \text{Me}, ^i\text{Pr}$) are reported, along with a hexanuclear N-heterocyclic dicarbene $[\text{Li}_2\text{Ce}_3(\text{OArC}^{\text{Me}}\text{-H})_3(\text{N}^i\text{Pr}_2)_5(\text{THF})_2]$ by-product. The analogous *para*-aryloxoide–NHC proligand ($\rho\text{-L}^{\text{Mes}} = 4\text{-O-}2,6\text{-}^t\text{Bu}_2\text{-C}_6\text{H}_2(1\text{-C}(\text{N}(\text{CH}_2)\text{N}(\text{Mes})))$) has been made for comparison, but the rare earth tris-ligand complexes $\text{Ln}(\rho\text{-L}^{\text{Mes}})_3(\text{THF})_2$ ($\text{Ln} = \text{Y}, \text{Ce}$) are too reactive for straightforward Lewis pair separated chemistry to be usefully carried out.

Received 26th July 2018

Accepted 5th September 2018

DOI: 10.1039/c8sc03312a

rsc.li/chemical-science

Carbon dioxide can be a useful and renewable C1 building block in the fine and bulk chemical industries due to its natural abundance and reactivity,^{1,2} and can provide carboxylic acids, esters and (cyclic) carbonates.³ Isoelectronic isocyanates and isothiocyanates are also valuable electrophilic elementary reagents used in polymerization and cyclisations,^{4–6} and thus chemistry which utilizes heteroallenes is of great interest.

Lewis basic N-heterocyclic carbenes (NHCs) are known to react with carbon dioxide, isocyanates and isothiocyanates as nucleophiles to form imidazolium carboxylates,^{7,8} imidazolium amidates⁹ and imidazolium carbimidothioates¹⁰ respectively **A** (Chart 1). While imidazolium carboxylates can successfully catalyse carbamate formation,¹¹ NHCs react as organocatalysts with isocyanates to form cyclic ureas **B** through an azolium amidate intermediate.⁹

Since the first reported isolation of lanthanide–NHC complexes in 1994,^{12,13} it has been shown that Lewis acidic rare-earth cations form hemilabile bonds with soft σ -donating NHCs.^{14,15} Between 2006 and 2010, Shen and co-workers published syntheses of aryloxoide–NHC lanthanide complexes, however no subsequent reactivity was reported.^{16–19} In 2014, we reported the activation of carbon dioxide **C** and carbon disulfide **D** using a scandium alkoxide–NHC complex, achieving frustrated Lewis pair (FLP) like reactivity which resulted in metal-ligand scrambling to form a polymeric $-(\text{Sc-NHC-CO}_2)_n$ containing network owing to the flexible alkoxide tether.²⁰

Cerium, the most abundant lanthanide has a relatively low toxicity; its trichloride is six times less toxic by ingestion than that of iron,^{21,22} and it has many applications in heterogeneous catalysis.²³ Previously we showed that cerium-silylamido NHC complexes $[\text{Ce}(\text{L})(\text{N}\{\text{SiMe}_3\}_2)_2]$ ($\text{L} =$ bidentate alkoxy-tethered NHC ligand) react with CO_2 to form an insoluble mixture while the uranium analogue $[\text{U}(\text{L})(\text{N}\{\text{SiMe}_3\}_2)_2]$ yields an equivalent of isocyanate.²⁴ In the latter instance it was not possible to isolate any intermediate that confirmed whether the NHC group was definitively involved in the CO_2 activation.²⁵ Recently, Suresh reported the first mononuclear *N*-carboxylate imidazolium lanthanide compounds, suggesting their potential

^aEaStCHEM School of Chemistry, University of Edinburgh, The King's Buildings, Edinburgh, EH9 3FJ, UK. E-mail: Polly.Arnold@ed.ac.uk^bEaStCHEM School of Chemistry, University of St. Andrews, North Haugh, St. Andrews, KY16 9ST, UK. E-mail: ads10@st-andrews.ac.uk^cMolecular Catalysis, Faculty of Chemistry and Catalysis Research Center, Technical University Munich, Lichtenbergstr. 4, 85748 Garching bei München, Germany

† Electronic supplementary information (ESI) available. CCDC 1856101–1856106, 1868204–1868209. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c8sc03312a





use as single-molecule magnets.²⁶ Here, we demonstrate that cerium (and other rare earth) complexes with aryloxide-tethered NHC ligands can successfully form homoleptic cerium imidazolium carboxylate complexes from CO₂ insertion into the Ce–C carbene bonds. We show how to control the reversibility for the

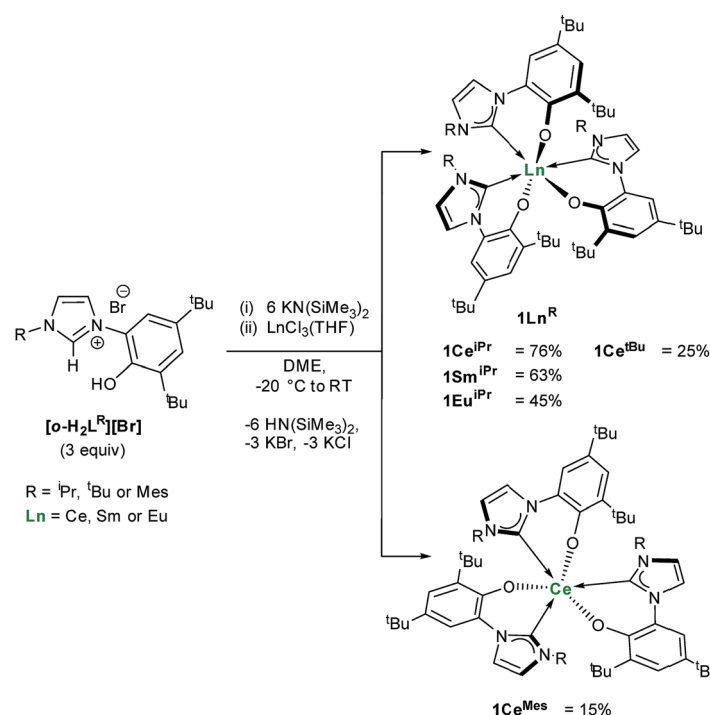
first time, and use this, and the extent of insertion of CO₂ or isoelectronic heteroallenes (isocyanates and isothiocyanates) by changing the ligand steric and electronic properties, and by solvent effects. This is important as we show that only the complexes capable of reversible CO₂-insertion are competent catalysts for the synthesis of cyclic carbonates from CO₂ and epoxides.

Results and discussion

ortho-Aryloxide Ln–NHC complex synthesis

One objective for synthesizing lanthanide aryloxide tethered–NHC complexes is to combine valuable hemilability within a rigid framework for selective reactivity and we envisioned that varying coordination environments arising from respective alkyl and aryl substituents could give distinctive chemistry. A suspension of an *ortho*-aryloxide NHC proligand,^{27,28} [*o*-H₂L^R][Br] where L^R = 2-O-3,5-^tBu₂-C₆H₂(1-C{N(CH₂)₂N(R)}) and R = ⁱPr, ^tBu and Mes were treated with 6 equivalents of KN(SiMe₃)₂ and LnCl₃(THF)_{*n*} (Ln = Ce, Sm, Eu) in DME to afford bright yellow solutions with colourless precipitates of KCl and KBr (Scheme 1). After work-up, **1Ln^R** (Ln(L^R)₃) can be afforded in moderate to good yields (15–76%), while over 8 g of **1Ce^{iPr}** can be isolated in a single reaction.

The ¹H NMR spectra of all four lanthanide complexes bearing alkyl R groups (ⁱPr or ^tBu) contain a complex set of paramagnetic resonances indicating C₁ symmetry and a unique environment for each ligand. In agreement with the ¹H spectrum of **1Ce^{iPr}** the ¹³C{¹H} NMR spectrum is also complicated, containing three carbene chemical shifts (δ = 174.8 ppm, 187.8 ppm, 192.3 ppm), slightly broadened compared to the rest of the spectrum (average fwhm 12 Hz), and shifted compared to



Scheme 1 Synthesis of homoleptic lanthanide(III) complexes **1Ln^R**.



similar diamagnetic lanthanide NHC complexes (normal region $\approx 200\text{--}238$ ppm for Y^{III} and Ce^{IV}).^{14,17,19,29,30} However in contrast, spectra of 1Ce^{Mes} contain a single set of paramagnetically shifted resonances indicating C_3 symmetry in solution on the ^1H NMR spectroscopic timescale and ^{13}C NMR spectroscopy of 1Ce^{Mes} displays a single carbene resonance at 184.2 ppm.

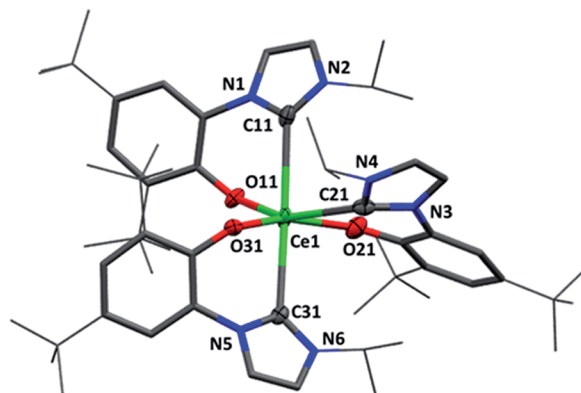
These ligand orientation differences are rationalized by consideration that three planar mesityl groups pack more easily than the aryloxy/*tert*-butyl groups would, and that the *tert*-butyl/*iso*-propyl steric repulsions are less prescriptive. The C_3 -symmetric complex would also be favoured if π -stacking between the mesityl substituent and an adjacent imidazolin-2-ylidene ring is possible. This high degree of steric crowding is used to rationalise the failed synthesis of related but bulkier diisopropylphenyl containing aryloxy-carbene ligands. Reactions aimed at targeting the mono- and bis-alkoxy-NHC analogues using this synthetic method yielded only the tris-ligand complex and unreacted LnCl_3 while the targeted synthesis of a $\text{Sm}(\text{II})$ analogue results in spontaneous oxidation and isolation of $\text{Sm}(\text{III})$ compound, 1Sm^{iPr} (see ESI†).

Single crystal X-ray analyses show that the *fac*- and *mer*-isomers are retained in the solid state for 1Ce^{Mes} and 1Ce^{iPr} respectively (see Fig. 1). The coordination geometry of cerium in each is a pseudo-octahedral geometry defined by average C–Ce–C bond angles ($172.93(18)^\circ/88.41(18)^\circ$ and $102.33(9)^\circ$) and OAr–CeOAr bond angles ($154.05(15)^\circ/102.44(16)^\circ$ and $94.82(9)^\circ$). The average Ce–C bond distances of 1Ce^{iPr} and 1Ce^{Mes} are 2.742(6) Å and 2.814(3) Å with the former within the regular range of a lanthanide–carbene bond. To the best of our knowledge the latter is the longest aryloxy tethered metal–carbene bond and amongst the longest lanthanide–carbene bonds known, consistent with the proposed high degree of hemilability. For 1Ce^{Mes} there is a conceivable offset aromatic donor–acceptor interaction between the electron deficient imidazolin-2-ylidene and the electron rich mesityl with an average centroid distance of 4.36 Å, within the upper limits of face-centred π -stacking.^{31,32}

Reactivity of 1Ce^{R} complexes

Exposure of a solution of 1Ce^{R} to an atmosphere of carbon dioxide results in the instant and quantitative formation of 2Ce^{R} ($\text{Ln}(\text{L}^{\text{R}}\cdot\text{CO}_2)_3$) as observed by the precipitation of a beige solid (hexanes reaction solvent) or monitoring by ^1H NMR spectroscopy (benzene reaction solvent), Scheme 2. As anticipated for a complex with a hemilabile metal–NHC bond, the CO_2 exclusively inserts into the three Ce–C bonds, and pleasingly, and in contrast to the complexes with more flexible, bidentate alkoxide–NHCs, the rest of the molecule remains relatively unperturbed, with no evidence of ligand redistribution between metal centres. Samples of 2Ce^{iPr} and 2Ce^{tBu} held at elevated temperatures under dynamic vacuum (100°C , 10^{-3} mbar) show no loss of CO_2 . However, a sample of 2Ce^{Mes} shows some loss of CO_2 under dynamic vacuum (25 to 100°C , 10^{-3} mbar), that is fully reversible. Solution phase analysis of the material formed shows it to be a complicated mixture that could be oligomeric, but the material is quantitatively converted back to 2Ce^{Mes} upon re-exposure to an atmosphere of CO_2 .

1Ce^{iPr} :



1Ce^{Mes} :

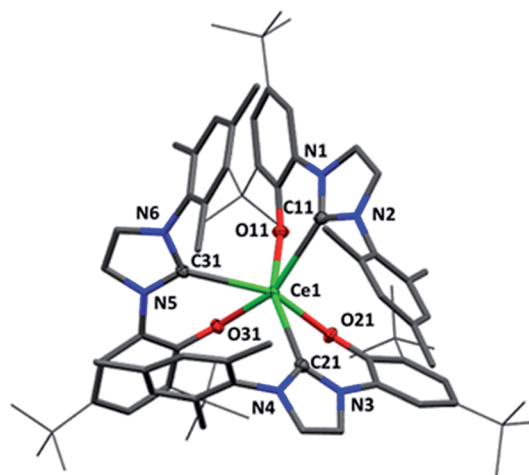
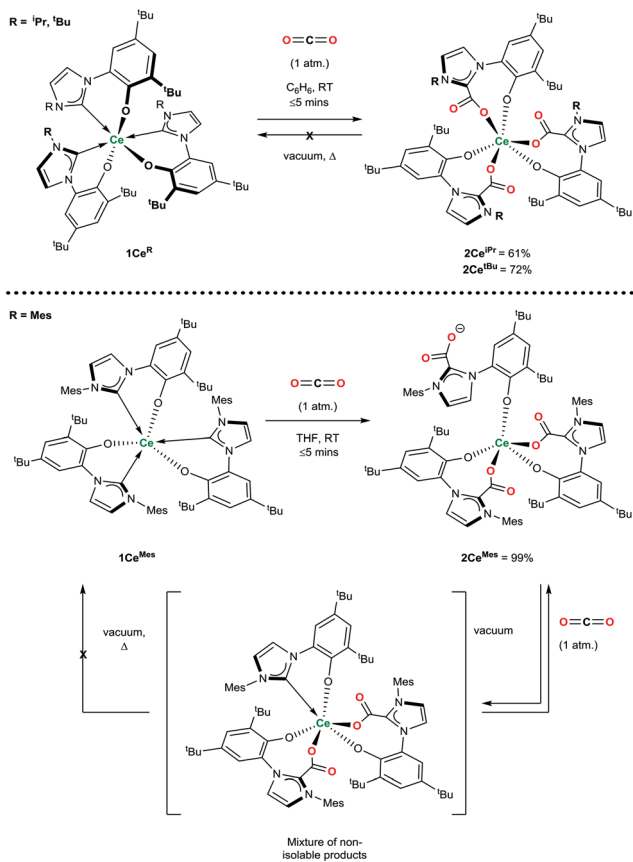


Fig. 1 Molecular structures of 1Ce^{iPr} (upper) and 1Ce^{Mes} (lower) with Ce, O and $\text{C}_{\text{carbene}}$ shown at 50% ellipsoid probability, framework and peripheral carbon atoms drawn capped stick and wireframe respectively, and H and lattice solvent omitted for clarity. Selected distances (Å) and angles ($^\circ$) for 1Ce^{iPr} : Ce1–C11 2.747(6), Ce1–C21 2.694(6), Ce1–C31 2.785(7), Ce1–O11 2.349(4), Ce1–O21 2.277(4), Ce1–O31 2.283(5), C11–Ce1–C21 88.28(18), C11–Ce1–C31 172.93(18), C21–Ce1–C31 88.53(18), O11–Ce–O21 154.05(15), O11–Ce–O31 107.42(15), O21–Ce–O31 97.45(16), C11–Ce1–O11 68.68(16), C21–Ce1–O21 69.09(18), C31–Ce1–O31 69.66(17); for 1Ce^{Mes} : Ce1–C11 2.823(3), Ce1–C21 2.814(3), Ce1–C31 2.806(3), Ce1–O11 2.266(2), Ce1–O21 2.264(2), Ce1–O31 2.251(2), C11–Ce1–C21 105.49(9), C11–Ce1–C31 100.01(9), C21–Ce1–C31 101.48(9), O11–Ce1–O21 93.66(8), O11–Ce1–O31 96.58(8), O21–Ce1–O31 94.23(8), C11–Ce1–O11 65.85(9), O21–Ce1–O21 66.82(8), O31–Ce1–O31 66.25(9).

^1H NMR spectroscopic analysis reveals that the *N*-alkyl functionalised 2Ce^{R} complexes have C_3 symmetry, *i.e.* a *fac*-conformation of the three bidentate ligands. The ^{13}C NMR spectra contain diagnostic CO_2 carbon resonances for 2Ce^{iPr} and 2Ce^{tBu} ($\delta = 173.1$ ppm and 173.5 ppm respectively) at significantly higher frequency than known organic NHC· CO_2 compounds (≈ 20 ppm)^{7,8} as might be anticipated from proximity to the paramagnetic metal center. The FTIR spectrum of 2Ce^{iPr} shows a characteristic absorption at 1666 cm^{-1} (typical range $\sim 1630\text{--}1690$).^{8,33,34} The conversion of 1Ce^{Mes} to 2Ce^{Mes}





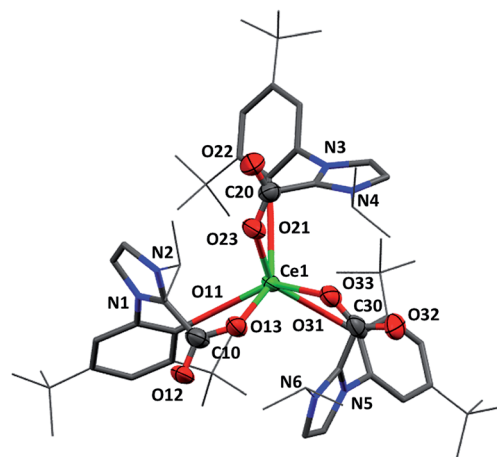
Scheme 2 Reactivity of 1Ce^{R} with CO_2 that forms the triply $\text{CO}_2\text{-NHC}$ inserted adducts irreversibly (2Ce^{iPr} and 2Ce^{tBu}) and reversibly (2Ce^{Mes}).

results in a lowering of symmetry from C_3 to C_1 according to room temperature solution spectroscopies. The ^1H NMR spectrum shows three broadened sets of paramagnetic ligand resonances, and two C–O stretches observable in the FTIR spectrum (1678 and 1715 cm^{-1}). We suggest that due to steric hindrance of three mesityl groups that one of the imidazolium carboxylate units is non-coordinating in solution.

Single crystal X-ray analysis confirms that CO_2 insertion products 2Ce^{iPr} and 2Ce^{tBu} have a pseudo-trigonal prismatic geometry with C_3 -symmetric *fac*-arrangement described by the average $\text{O}^{\text{Ar}}\text{-Ce-O}^{\text{Ar}}$ bond angles ($97.03(10)^\circ$ and $95.94(8)^\circ$ respectively) and $\text{O}^{\text{CO}}\text{-Ce-O}^{\text{CO}}$ bond angles ($77.77(10)^\circ$ and $80.48(8)^\circ$) (Fig. 2). The average Ce– O^{CO} bond length is within the regular bond length range at ($2.472(6)$ Å and $2.473(2)$ Å respectively) suggesting a strong degree of stabilisation despite an increase of metal chelate ring size from 6 to 8.

The substrate scope was further explored with carbon disulfide and other isoelectronic (hetero)allenes shown in Scheme 3. Interestingly, treatment of a benzene solution of 1Ce^{iPr} with excess carbon disulfide at temperatures up to 80°C shows no reaction. This differs from the alkoxide-tethered carbene complex **D** for which the product arising from the insertion of CS_2 into two (of the three) M–C bonds was characterized.²⁰ The higher reactivity of CO_2 compared to CS_2 in

2Ce^{iPr} :



2Ce^{tBu} :

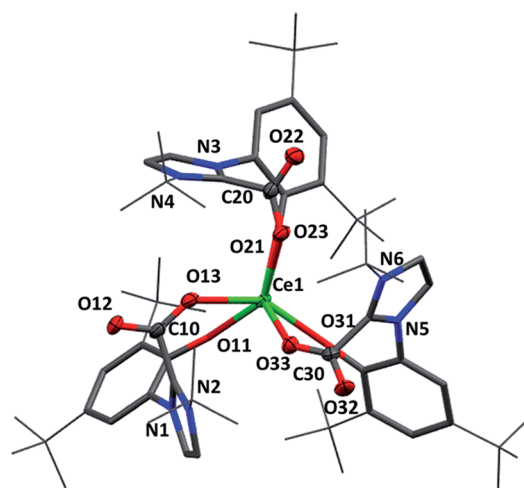
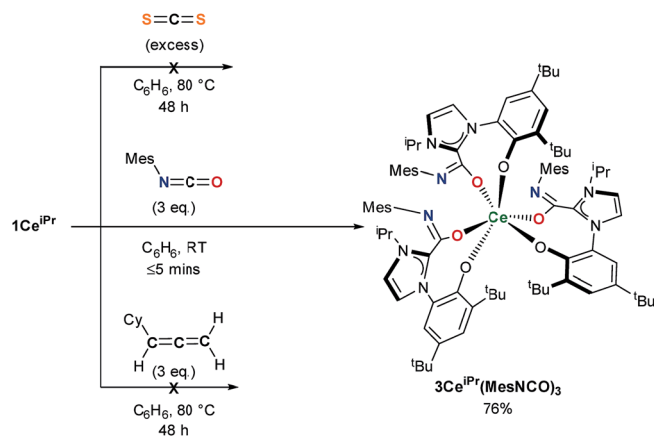


Fig. 2 Molecular structures of 2Ce^{iPr} (upper) and 2Ce^{tBu} (lower) with carboxylate and Ce atoms shown at 50% ellipsoid probability, framework and peripheral carbon atoms drawn capped stick and wireframe respectively, and H and lattice solvent omitted for clarity. Selected distances (Å) and angles ($^\circ$) for 2Ce^{iPr} : Ce1–O11 2.790(5), Ce1–O21 2.274(5), Ce1–O31 2.256(6), Ce1–O13 2.468(6), Ce1–O23 2.482(6), Ce1–O33 2.466(6), O11–Ce1–O21 97.18(19), O11–Ce1–O31 99.3(2), O21–Ce1–O31 94.8(2), O13–Ce1–O23 77.02(19), O13–Ce1–O33 76.68(19), O23–Ce1–O33 79.6(2), O11–Ce1–O13 75.72(19), O21–Ce1–O23 75.7(2), O31–Ce1–O33 75.9(2). For 2Ce^{tBu} : Ce1–O11 2.261(2), Ce1–O21 2.268(2), Ce1–O31 2.274(2), Ce1–O13 2.473(2), Ce1–O23 2.477(2), Ce1–O33 2.470(2), O11–Ce1–O21 95.38(8), O11–Ce1–O31 95.25(8), O21–Ce1–O31 97.20(8), O13–Ce1–O23 76.76(8), O13–Ce1–O33 80.89(8), O23–Ce1–O33 80.80(8), O11–Ce1–O13 75.50(8), O21–Ce1–O23 74.77(2), O31–Ce1–O33 75.03(8).

this system is reasonable considering the stronger affinity of Ce for oxygen, and the lower dipole moment in the latter reagent.

Treatment of a benzene or THF solution of 1Ce^{iPr} with three equivalents of mesityl isocyanate (MesNCO) immediately results in the insertion of isocyanate into all three Ce–NHC bonds, affording a pale-yellow solution from which the tris-azoliumamidate $3\text{Ce}^{\text{iPr}}(\text{MesNCO})_3$ can be readily isolated as colourless microcrystalline powder in 76% yield, Scheme 3. No dimer or trimer isocyanate products were observed as





Scheme 3 Treatment of 1Ce^{iPr} with reagents isoelectronic to CO_2 .

a comparison to “free” NHC isocyanate chemistry.⁹ As could be expected, the non-polar and more sterically hindered cyclohexallene shows no reactivity with 1Ce^{iPr} .

In the reaction of 1Ce^{iPr} with three equivalents of *tert*-butyl isocyanate (${}^t\text{BuNCO}$) in benzene or THF, two molecules of isocyanate insert to form $3\text{Ce}^{\text{iPr}}({}^t\text{BuNCO})_2$, however in DME solution, three molecules insert to form $3\text{Ce}^{\text{iPr}}({}^t\text{BuNCO})_3$ as a 3 : 1 mixture of the *fac*- and *mer*-isomers observable by ${}^1\text{H}$ NMR spectroscopy, Scheme 4. We suggest that in the former two solvents, the steric bulk of the *tert*-butyl groups restricts access to the third equivalent, but the stronger, bidentate donor solvent DME increases the lability of the NHC groups, enabling three insertions to occur. If 1Ce^{iPr} is treated with 3 equivs of *tert*-

$3\text{Ce}^{\text{iPr}}({}^t\text{BuNCS})_2$:

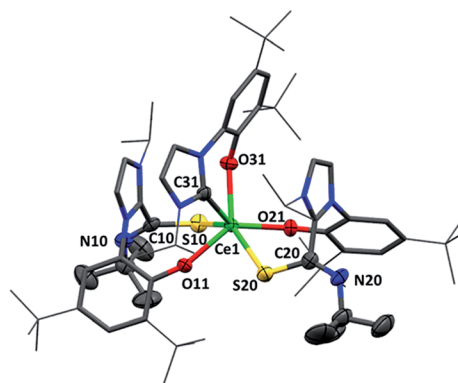
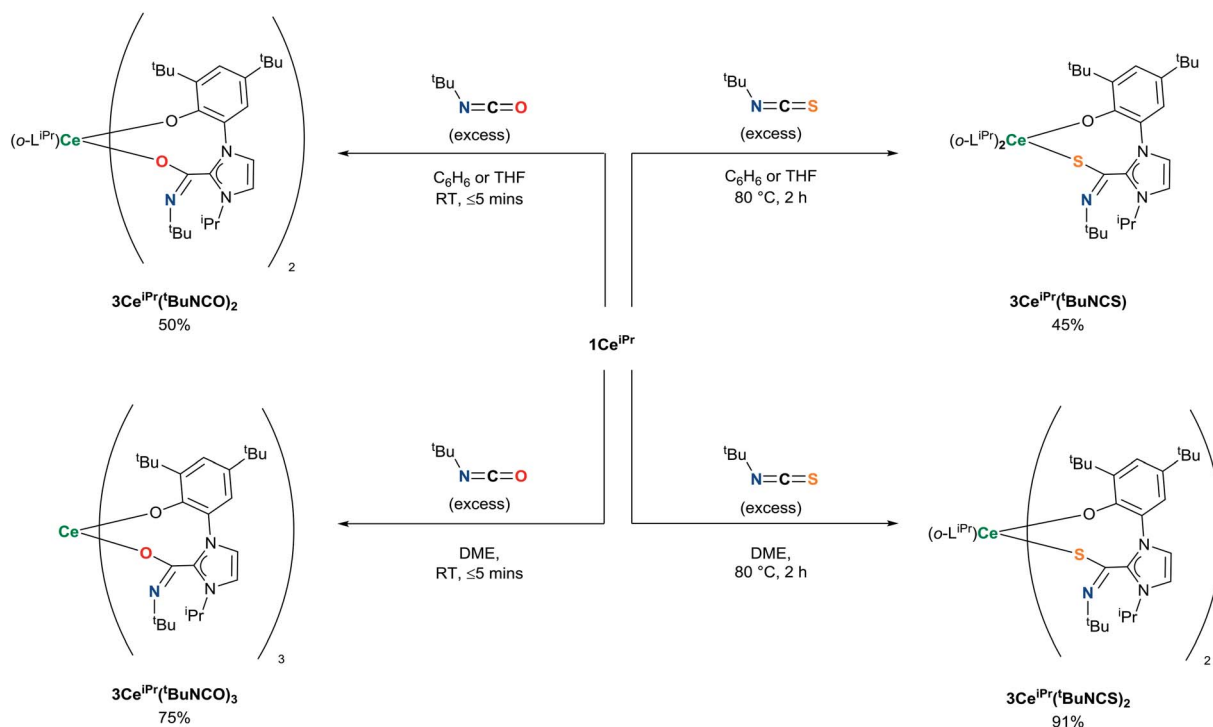


Fig. 3 Molecular structure of $3\text{Ce}^{\text{iPr}}({}^t\text{BuNCS})_2$ with selected C and non-C/H atoms shown at 50% ellipsoid probability, framework and peripheral carbon atoms drawn capped stick and wireframe respectively, and H and lattice solvent omitted for clarity. Selected average distances (Å) and angles ($^\circ$) for $3\text{Ce}^{\text{iPr}}({}^t\text{BuNCS})_2$: Ce1–S10 2.996(12), Ce1–S20 3.048(12), Ce1–C31 2.716(4), Ce1–O11 2.280(3), Ce1–O21 2.273(3), Ce1–O31 2.277(3), S10–Ce1–S20 143.91(3), S10–Ce1–C31 126.96(9), S20–Ce1–C31 80.70(7), S10–Ce1–O11 78.71(8), S20–Ce1–O21 78.79(7), C31–Ce1–O31 69.93(11).

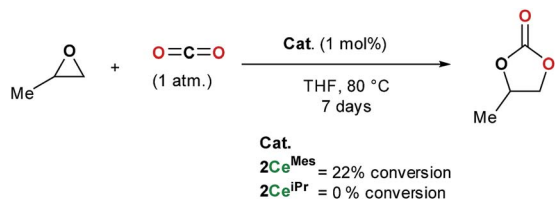
butyl isothiocyanate (${}^t\text{BuNCS}$) at 80 °C in benzene or THF, a single equivalent of isothiocyanate inserts to form $3\text{Ce}^{\text{iPr}}({}^t\text{BuNCS})$ while in DME two equivalents of isothiocyanate insert to form $3\text{Ce}^{\text{iPr}}({}^t\text{BuNCS})_2$.

Single crystals of $3\text{Ce}^{\text{iPr}}({}^t\text{BuNCS})_2$ were grown by slow diffusion of heptane into a toluene solution. An X-ray diffraction study, Fig. 3, reveals a pseudo-trigonal prismatic molecular



Scheme 4 Differences in reactivity of 1Ce^{iPr} with sterically hindered isocyanates and isothiocyanates depending on solvent, and ligand substituents to afford $3\text{Ce}^{\text{iPr}}({}^t\text{BuNCO})_2$, $3\text{Ce}^{\text{iPr}}({}^t\text{BuNCO})_3$, $3\text{Ce}^{\text{iPr}}({}^t\text{BuNCS})$, and $3\text{Ce}^{\text{iPr}}({}^t\text{BuNCS})_2$.





Scheme 5 Catalytic formation of propylene carbonate from propylene oxide in an atmosphere of carbon dioxide using 2Ce^{Mes} and 2Ce^{iPr}.

geometry at the metal center in the solid state. The Ce–S bond lengths average at 3.022 Å, and the Ce–C bond (2.716 Å) is only a little shorter than the average Ce–C bond length in the parent compound 1Ce^{iPr} (2.742 Å). The obtuse S–Ce–S bond angle (143.91°) and chelate angle of each bidentate ligand is within the expected range; S–Ce–O^{Ar} (78.75° avg.) and C–Ce–O^{Ar} (69.93°).

Catalytic applications of 2Ce^R complexes

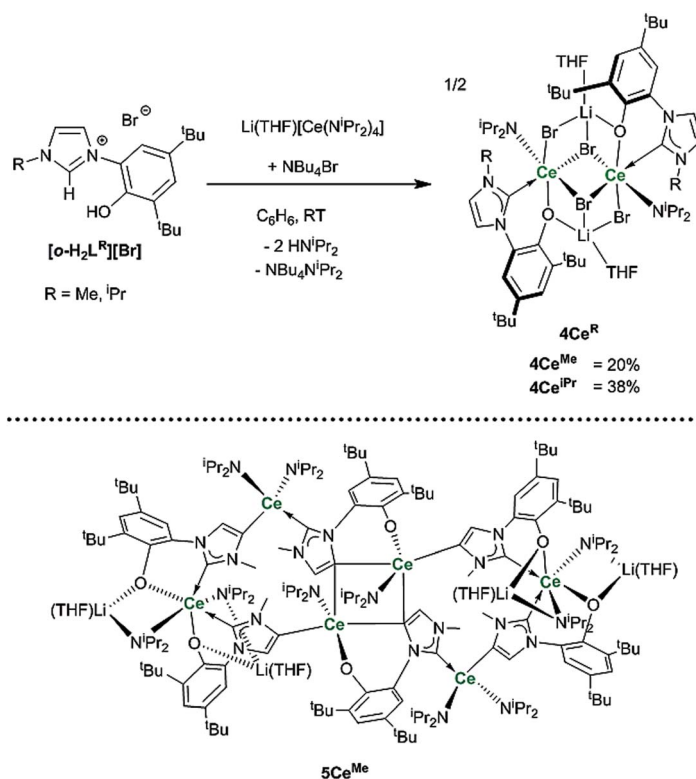
The formation of cyclic carbonates from epoxides and carbon dioxide was chosen for a preliminary study of the catalytic activity of the tris(ligand) CO₂ adducts 2Ce^{iPr} and 2Ce^{Mes}. Both free base NHCs and imidazolium carboxylates can be used as catalysts for the formation of cyclic carbonates from epoxides and carbon dioxide under high temperatures and pressures (up to 120 °C and 20 atm), while rare earth initiators are known to

function at lower temperatures and/or pressures, a co-catalyst is usually required.^{11,35} Scheme 5 shows how under an atmosphere of carbon dioxide, 1 mol% of 2Ce^{Mes} catalyses the conversion of propylene oxide to propylene carbonate with 22% conversion at 80 °C in THF over 7 days, a much higher activity than the imidazolium carboxylates alone. On the other hand, the more compact 2Ce^{iPr} shows no reactivity. The solid-state structures show a higher steric congestion in the L^{Mes} adduct 1Ce^{Mes}, and IR and NMR spectroscopies confirm different ligand solution environments for 2Ce^{Mes}, suggesting both the Ce–C_{carbene} and Ce–C_{CO₂} interactions are weaker and more labile for the Mes system. We propose that the catalysis requires a combination of Lewis base type NHC–CO₂ activation, and Lewis acid type Ce-epoxide activation.

Synthesis of the heteroleptic substituted NHC analogues

To target reactions with single equivalents of CO₂, reactions designed to make complexes containing a single NHC ligand were carried out. The reactions of the ligands [o-H₂L^R][Br], R = Me, ⁱPr and equimolar amounts of Li(THF)[CeN(ⁱPr₂)₄] only afford clean material in low yields and significant decomposition can be observed. Adding an additional bromide source improves the yield of the mono-NHC–Ce complexes 4Ce^{Me} and 4Ce^{iPr} (Ce₂Br₄L^R(ⁱPr₂N)₂Li₂(THF)₂) to a moderate level (20% and 38% respectively, see Scheme 6).

Crystallographic analysis reveals a dimeric structure still containing unreacted base and lithium ions (see ESI[†]). A



Scheme 6 Reactions to target mono–NHC Ln complexes that afford 4Ce^{Me}, 4Ce^{iPr} and the hexanuclear 5Ce^{Me} that is the by-product isolated as single X-ray quality crystals for R = Me.



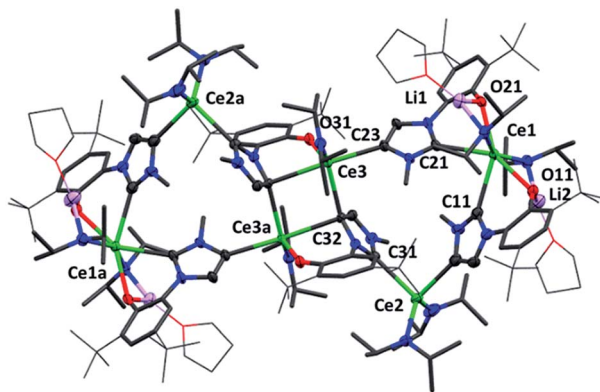


Fig. 4 Molecular structure of 5Ce^{Me} with imidazolium, Ce, O and Li atoms shown at 50% ellipsoid probability, framework atoms drawn wireframe, and H and lattice solvent omitted for clarity. Selected distances (Å) and angles ($^{\circ}$) for 5Ce^{Me} : Ce1–C11 2.743(5), Ce1–C21 2.728(5), Ce2–C13 2.667(7), Ce2–Ce31 2.710(5), Ce3–C23 2.651(6), Ce3–C32 2.854(4), Ce3–C32a 2.685(6), Ce1–O11 2.517(4), Ce1–C21 2.462(4), Ce3–O31 2.232(3), C11–Ce1–O11 65.1(2), C11–Ce1–C21 82.1(2), C21–Ce1–O21 65.9(1), C13–Ce2–C31 94.4(2), Ce3–C32–Ce3a 98.5(1), C32–Ce3–C32a 81.5(1), C32–Ce3–O31 71.1(1), C32–Ce3–C23a 163.8(2).

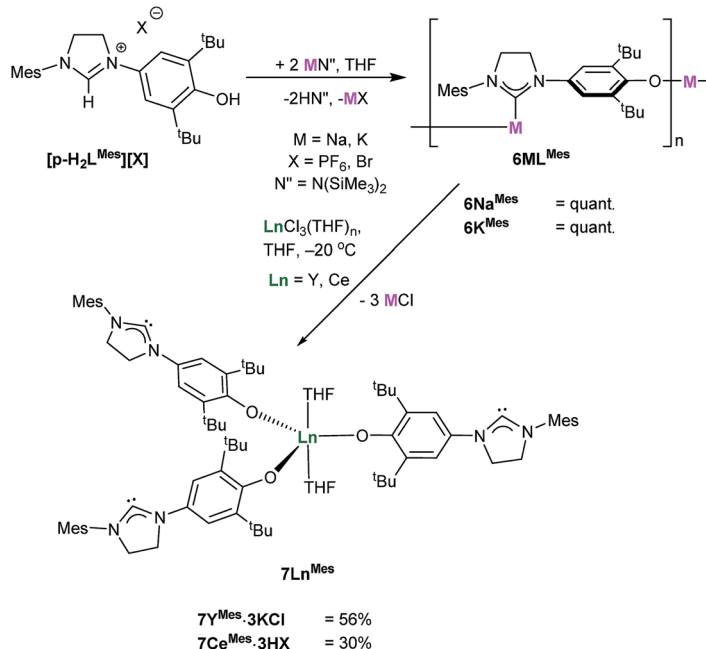
complicated bis(ligand) Li_4Ce_6 cluster 5Ce^{Me} , in which each ligand has been deprotonated at the NHC backbone (in the 4-position) yielding a dianionic OC ligand that bridges two cerium cations, is isolated in low yield as orange crystals that are suitable for single crystal diffraction studies (Fig. 4 and ESI †). Syntheses to target 4 or 5 in the absence of an additional bromide source, or from cerium bromide starting materials, yield only complicated mixtures of compounds in our hands.

Synthesis of the *para*-aryloxy substituted NHC analogues

The analogous complexes of the *para*-substituted aryloxy ligand $p\text{-L}^{\text{R}}$ separate the Lewis acid and Lewis base centers, and thus offer a potential insight into the importance of the adjacent Ln centre and the nucleophilic NHC in the combined activation of CO_2 and the other unsaturated substrates. A modification of Wang's proligand synthesis using saturated-backbone imidazoline analogues allows access to the *para*-functionalized proligand in 15% yield.²⁷ Treatment of this *N*-mesityl functionalized proligand $[p\text{-H}_2\text{L}^{\text{Mes}}][\text{X}]$, ($p\text{-L}^{\text{Mes}} = 4\text{-O-2,6-}t\text{Bu}_2\text{-C}_6\text{H}_2(1\text{-C}\{\text{N}(\text{CH}_2)_2\text{-NMe}\})$), $\text{X} = \text{PF}_6, \text{Br}$) with either $\text{MN}(\text{SiMe}_3)_2$ ($\text{M} = \text{Na}$ or K) in THF at room temperature affords the group 1 NHC salts 6M^{Mes} $[(\text{M}(p\text{-L}^{\text{Mes}}))]_n$, ($\text{M} = \text{Na}, \text{K}$) in quantitative yield, Scheme 7. The solid-state structures of both are polymeric, according to single crystal X-ray data, with 6Na^{Mes} displaying repeating $\text{C}[\text{Na}(\mu\text{-ArO})_2\text{Na}]\text{C}$ diamond units, while 6K^{Mes} displays a perpendicular ArO-K-C arrangement, see ESI † .

Salt 6M^{Mes} can be treated with YCl_3 or CeCl_3 at $-20\text{ }^{\circ}\text{C}$ to synthesise 7Ln^{Mes} ($\text{Ln}(p\text{-L}^{\text{Mes}})_3(\text{THF})_2$ where $\text{Ln} = \text{Y}, \text{Ce}$) in 56% and 30% yield as yellow powders. Due to their high reactivity all the compounds start to degrade rapidly making further analysis difficult, and the complexes are best stored in their protonated form, *i.e.* $[\text{Ln}(p\text{-HL}^{\text{Mes}})]_3(\text{THF})_2\text{X}_3$.

Analysis of 7Y^{Mes} by ^{13}C NMR spectroscopy reveals a characteristic carbene signal (δ 238.2 ppm) is observed as a singlet indicating that the carbene does not bind to yttrium in solution. These complexes were found to be extremely air sensitive, were only ever isolated as KCl and HCl salts and became highly insoluble in a range of solvents so were not pursued further (see ESI †).



Scheme 7 Reaction to target the *para*-ligand adducts 7Ln^{Mes} ($\text{Ln} = \text{Y}, \text{Ce}$).



Conclusions

The tris(*ortho*-aryloxy-NHC) rare earth complexes $\text{Ln}(\text{L}^{\text{R}})_3$ are readily isolated and are the thermodynamic sink in this system. Insertion of CO_2 or a range of isoelectronic (hetero)allenes into the labile cerium carbene bond in $\text{Ce}(\text{L}^{\text{R}})_3$ shows a dependence on solvent and N-R group on L^{R} that enables control of the degree of insertion. The CO_2 -insertion products form cleanly at ambient pressure, but only reversibly for the bulky mesityl substituted $\text{Ce}(\text{L}^{\text{Mes}})_3$. The reversibility of the CO_2 insertion appears to be crucial for further reactivity as only $\text{Ce}(\text{L}^{\text{Mes}} \cdot \text{CO}_2)_3$ is an active catalyst for the conversion of propylene oxide to propylene carbonate. Although yields in these preliminary tests using low temperatures and one atmosphere of CO_2 are low, the catalyst is more active than a monodentate NHC, and when the ligands are better fit to the metal in $\text{Ce}(\text{L}^{\text{iPr} \cdot \text{CO}_2})_3$, the complexes are inactive. We propose that the catalysis requires a combination of Lewis base type NHC- CO_2 activation, and Lewis acid type Ce-epoxide activation.

Although the tris-ligand complexes are the thermodynamic sink in the system, the mono-ligand complexes can be isolated from reactions using $\text{LiCe}(\text{N}^{\text{iPr}}\text{Pr}_2)_4$ as a starting material; LiBr adducts $[\text{Ce}(\text{L}^{\text{R}})(\text{N}^{\text{iPr}}\text{Pr}_2)\text{Br} \cdot \text{LiBr}(\text{THF})_2]$ ($\text{R} = \text{Me}, \text{iPr}$) are reported, along with a hexanuclear N-heterocyclic dicarbene complex $[\text{Li}_2\text{Ce}_3(\text{OArC}^{\text{Me}}\text{-H})_3(\text{N}^{\text{iPr}}\text{Pr}_2)_5(\text{THF})_2]_2$ which is formed as a by-product. The analogous *para*-aryloxy-NHC proligand $p\text{-L}^{\text{Mes}}$ has been made for comparison. The group 1 salts $[\text{Na}(p\text{-L}^{\text{Mes}})]_n$ and $[\text{K}(p\text{-L}^{\text{Mes}})]_n$ form two different types of infinite coordination polymers through metal carbene-bonds. Synthesis of the analogous lanthanide *para*-aryloxy NHC complexes $\text{Ln}(p\text{-L}^{\text{Mes}})_3(\text{THF})_2$ ($\text{Ln} = \text{Y}, \text{Ce}$) is possible but they are all highly reactive leading to rapid degradation. Therefore straightforward Lewis pair separated chemistry cannot usefully be carried out.

Further work is underway to use the C_3 -symmetric tris(*ortho*-aryloxy-NHC)- CO_2 adducts in asymmetric catalysis and to expand the scope of the CO_2 functionalisation.

Experimental

General details

All manipulations were carried out under a dry, oxygen-free atmosphere of nitrogen using standard Schlenk and glovebox technique. All gases were supplied by BOC gases UK. All glassware items, cannulae and Fisherbrand 1.2 μm retention glass microfibre filters were dried in a 170 $^\circ\text{C}$ oven overnight before use. Benzene and DME were distilled from potassium and stored over 4 \AA molecular sieves. Hexane, heptane, THF, and toluene were degassed and purified by passage through activated 4 \AA molecular sieves or activated alumina towers and stored over 4 \AA molecular sieves. Deuterated solvents, benzene- d_6 , THF- d_8 and pyridine- d_5 were dried over potassium, vacuum-transferred, and freeze-pump-thaw degassed prior to use. ^1H and ^{13}C NMR spectra were recorded on Bruker AVA400, AVA500, or PRO500 spectrometers at 300 K. Chemical shifts are reported in parts per million, δ , referenced to residual proton resonances, and calibrated against external TMS. Infrared spectra were recorded on a Perkin Elmer Spectrum 65 FT-IR

spectrometer as nujol mulls between KBr disks. Mass spectra were acquired using a Solarix FT-ICR (12 T) (Bruker UK Ltd) equipped with a Bruker APPI source. Samples were prepared as ca. 1 mM toluene solutions of the substrate. Elemental analyses were carried out at London Metropolitan University, London, UK.

$\text{NaN}(\text{SiMe}_3)_2$, $\text{KN}(\text{SiMe}_3)_2$,³⁶ and the $[\text{o-H}_2\text{L}^{\text{R}}][\text{Br}]$ ^{27,28} proligands were prepared according to the literature procedures. $\text{YCl}_3(\text{H}_2\text{O})_n$ and $\text{LnCl}_3(\text{H}_2\text{O})_n$ were purchased and stirred overnight with TMSCl (40 equiv.) in THF before vacuum drying for several hours.

General procedure 1 – synthesis of 1Ln^{R}

To a suspension of $[\text{o-H}_2\text{L}^{\text{R}}][\text{Br}]$ (3 equiv.) in DME (0.1 M) $\text{KN}(\text{SiMe}_3)_2$ (6 equiv.) was added and the resulting mixture was stirred for 5 min at -20 $^\circ\text{C}$ while and warmed to room temperature. $\text{LnCl}_3(\text{THF})_n$ (1 eq.) was added, and the resulting mixture was stirred at room temperature for 2 h. Volatiles were removed under reduced pressure, the crude product was extracted three times with hexane and the combined filtrates were concentrated to saturation and cooled to -20 $^\circ\text{C}$ overnight. The resulting suspension was filtered and the solid collected and dried under vacuum to give the title compound which was stored at -20 $^\circ\text{C}$ under a nitrogen atmosphere.

1Ce^{iPr} . Using general procedure 1 – 3-(3,5-di-*tert*-butyl-2-hydroxyphenyl)-1-isopropyl-1H-imidazol-3-ium bromide $[\text{o-H}_2\text{L}^{\text{iPr}}][\text{Br}]$ (11.82 g, 30 mmol), $\text{KN}(\text{SiMe}_3)_2$ (11.97 g, 60 mmol), $\text{CeCl}_3(\text{THF})_{1.15}$ (3.29 g, 10 mmol) and DME (100 mL) gave after recrystallization the title compound 1Ce^{iPr} as a yellow solid (8.17 g, 7.6 mmol, 76%). X-ray quality crystals were grown from a concentrated hexane solution at -20 $^\circ\text{C}$ over 1 week. ^1H NMR (500 MHz, C_6D_6) δ_{H} : -9.76 (3H, s, $\text{CH}(\text{CH}_3)$), -6.89 (3H, s, $\text{CH}(\text{CH}_3)$), -4.37 (3H, s, $\text{CH}(\text{CH}_3)$), -3.38 (9H, s, $\text{C}(\text{CH}_3)_3$), -1.62 (3H, s, $\text{CH}(\text{CH}_3)$), 0.47 (3H, s, $\text{CH}(\text{CH}_3)$), 0.77 (s, 1H, CH), 1.51 (9H, s, $\text{C}(\text{CH}_3)_3$), 1.57 (9H, s, $\text{C}(\text{CH}_3)_3$), 1.75 (1H, s, CH), 1.76 (1H, s, CH), 2.10 (9H, s, $\text{C}(\text{CH}_3)_3$), 2.31 (9H, s, $\text{C}(\text{CH}_3)_3$), 3.37 – 3.32 (9H, m, $\text{C}(\text{CH}_3)_3$), 3.43 (3H, s, $\text{CH}(\text{CH}_3)$), 5.97 (1H, s, CH), 6.70 (1H, s, CH), 7.08 (1H, app d, J 2.7, CH), 7.08 (1H, app d, J 2.7, CH), 7.62 (1H, s, CH), 7.70 (1H, app d, J 2.7, CH), 8.91 (1H, s, CH), 9.75 (1H, s, CH), 10.18 (2H, m, $2 \times \text{CH}$), 10.39 (1H, s, CH), 11.01 (1H, s, CH), 11.22 (1H, s, CH). $^{13}\text{C}\{^1\text{H}\}$ NMR (126 MHz, C_6D_6) δ_{C} : 14.4, 21.7, 22.3, 23.1, 24.6, 24.7, 30.1, 30.1, 31.9, 32.0, 32.5, 33.4, 33.9, 34.2, 35.4, 36.2, 36.5, 39.4, 41.0, 41.8, 46.4, 51.4, 114.4, 115.6, 118.9, 119.3, 119.5, 121.3, 122.2, 122.6, 122.7, 123.8, 124.2, 129.7, 131.4, 138.8, 140.3, 140.5, 141.9, 146.2, 147.4, 147.7, 148.6, 154.0, 155.8, 162.9, 174.8, 187.8, 192.2. Elemental analysis $\text{C}_{60}\text{H}_{87}\text{CeO}_3\text{N}_6$: C 66.70%, H 8.12%, N 7.78% calculated. C 66.72%, H 8.13%, N 7.78% found; $\text{APPI}^+ \text{C}_{60}\text{H}_{87}\text{CeN}_6\text{O}_3^+ [\text{M}]^+$ requires 1079.5894, found 1079.5711 (-17.0 ppm).

1Ce^{tBu} . Using general procedure 1 – 1-(*tert*-butyl)-3-(3,5-di-*tert*-butyl-2-hydroxyphenyl)-1H-imidazol-3-ium bromide $[\text{o-H}_2\text{L}^{\text{tBu}}][\text{Br}]$ (306 mg, 0.75 mmol), $\text{KN}(\text{SiMe}_3)_2$ (300 mg, 1.5 mmol), $\text{CeCl}_3(\text{THF})_{1.15}$ (80 mg, 0.25 mmol) and DME (2.5 mL) gave after recrystallization title compound 1Ce^{tBu} as a yellow solid (70 mg, 0.0625 mmol, 25%). ^1H NMR (400 MHz, C_6D_6) δ_{H} : -18.43 (9H, s, $\text{C}(\text{CH}_3)_3$), -9.47 (9H, s, $\text{C}(\text{CH}_3)_3$), -4.21 (9H, s, $\text{C}(\text{CH}_3)_3$), -3.27



(9H, s, C(CH₃)₃), -1.25 (9H, s, C(CH₃)₃), 1.18 (9H, s, C(CH₃)₃), 2.41 (9H, s, C(CH₃)₃), 4.02 (9H, s, C(CH₃)₃), 4.88 (1H, s, ArH), 6.23 (1H, s, ArH), 7.45 (1H, s, ArH), 8.66 (1H, s, ArH), 9.30–9.38 (10H, m, ArH + C(CH₃)₃), 9.63 (1H, s), 9.68 (1H, s), 11.97 (1H, s), 16.43 (1H, s), 16.84 (1H, s). ¹³C{¹H} NMR (126 MHz, C₆D₆) δ_C: 23.9, 25.3, 27.0, 29.0, 29.1, 30.1, 33.8, 34.2, 35.5, 35.7, 36.7, 36.9, 37.9, 38.9, 48.2, 51.8, 53.0, 57.4, 113.5, 115.1, 117.1, 118.1, 118.8, 120.6, 121.4, 123.0, 123.7, 124.3, 125.5, 130.3, 132.5, 133.4, 134.5, 137.9, 140.6, 142.4, 144.7, 147.0, 149.9, 157.3, 160.8, 163.1, 171.1, 206.2, 213.0. APPI⁺ C₆₃H₉₃CeN₆O₃⁺ [M]⁺ requires 1121.6364, found 1121.6333 (-2.7 ppm). After several attempts, this compound did not give satisfactory elemental analysis results, presumably because of its thermal sensitivity.

1Ce^{Mes}. Using general procedure 1 - 3-(3,5-di-*tert*-butyl-2-hydroxyphenyl)-1-mesityl-1*H*-imidazol-3-ium bromide [**o**-H₂L^{Mes}][Br] (353 mg, 0.75 mmol), KN(SiMe₃)₂ (300 mg, 1.5 mmol), CeCl₃(THF)_{1.15} (80 mg, 0.25 mmol) and DME (2.5 mL) gave after extraction and recrystallization in benzene title compound **1Ce^{Mes}** as a yellow solid (42 mg, 0.037 mmol, 15%). X-ray quality crystals were grown from a concentrated benzene solution over 1 week at room temperature. ¹H NMR (400 MHz, C₆D₆) δ_H: -8.50 (9H, ArCH₃), -3.80 (27H, s, C(CH₃)₃), 1.51 (9H, s, ArCH₃), 2.13 (3H, s, ArH), 2.80 (27H, s, C(CH₃)₃), 7.92 (9H, s, ArCH₃), 8.33 (3H, s, ArH), 8.53 (3H, s, ArH), 9.07 (3H, s, ArH), 11.54 (3H, s, ArH), 12.01 (3H, s, ArH). ¹³C{¹H} NMR (126 MHz, C₆D₆) δ_C: 20.1 (ArCH₃), 21.8 (ArCH₃), 25.4 (ArC(CH₃)₃), 32.7 (ArCH₃), 34.9 (ArC(CH₃)₃), 36.0 (ArC(CH₃)₃), 122.4 (ArC), 123.2 (ImC), 123.8 (ArC), 124.8 (ArC), 125.4 (ArC), 129.5 (ArC), 130.5 (ArC), 130.5 (ArC), 133.0 (ImC), 135.1 (ArC), 135.7 (ArC), 138.3 (ArC), 139.8 (ArC), 147.9 (ArC), 148.3 (ArC), 184.2 (NCN). Elemental analysis C₇₈H₈₈CeO₃N₆: C 71.58%, H 7.62%, N 6.42% calculated. C 71.43%, H 7.76%, N 6.31% found; APPI⁺ C₇₈H₈₈CeN₆O₃⁺ [M]⁺ requires 1307.6833, found 1307.6810 (-1.7 ppm).

1Sm^{iPr}. Using general procedure 1 - 3-(3,5-di-*tert*-butyl-2-hydroxyphenyl)-1-isopropyl-1*H*-imidazol-3-ium bromide [**o**-H₂L^{iPr}][Br] (296 mg, 0.75 mmol), KN(SiMe₃)₂ (300 mg, 1.5 mmol), SmCl₃(THF)₂ (100 mg, 0.1575 mmol) and DME (2.5 mL) gave after recrystallization title compound **1Sm^{iPr}** as a yellow solid (171 mg, 7.6 mmol, 63%). ¹H NMR (400 MHz, C₆D₆) δ_H: -9.17(-9.07) (1H, m, CH(CH₃)), -4.24 (3H, app d, *J* 5.8, CH(CH₃)), -2.42 (3H, app d, *J* 5.5, CH(CH₃)), -1.13 (9H, s, C(CH₃)), -0.85(-0.73) (6H, m, 2xCH(CH₃)), -0.63(-0.59) (1H, m, CH(CH₃)), 1.08 (9H, s, C(CH₃)), 1.65-1.70 (3H, m, CH(CH₃)), 1.81 (10H, s, C(CH₃)), 1.85 (10H, s, C(CH₃)), 2.02 (10H, s, C(CH₃)), 2.39 (10H, s, C(CH₃)), 3.24 (3H, app d, *J* 5.5, CH(CH₃)), 4.99 (1H, app p, *J* 6.8, CH(CH₃)), 5.54 (1H, app d, *J* 1.8, ArH), 6.05 (1H, app d, *J* 1.7, ArH), 7.26 (1H, app d, *J* 1.7, ArH), 7.90 (1H, app d, *J* 2.5, ImH), 8.06 (1H, app d, *J* 1.7, ArH), 8.14 (1H, app d, *J* 2.4, ImH), 8.15 (1H, app d, *J* 2.4, ImH), 8.32 (1H, app d, *J* 2.6, ImH), 8.43 (1H, app d, *J* 1.8, ArH), 8.46 (1H, app d, *J* 2.6, ImH), 8.53 (1H, app d, *J* 1.7, ArH), 8.94 (1H, app d, *J* 2.5, ImH). Elemental analysis C₆₀H₈₇SmO₃N₆: C 66.07%, H 8.04%, N 7.70% calculated. C 60.10%, H 8.33%, N 7.54% found. APPI⁺ C₆₀H₈₇SmN₆O₃⁺ [M]⁺ requires 1091.6037, found 1091.6076 (+3.6 ppm).

1Eu^{iPr}. Using general procedure 1 - 3-(3,5-di-*tert*-butyl-2-hydroxyphenyl)-1-isopropyl-1*H*-imidazol-3-ium bromide [**o**-H₂L^{iPr}][Br] (296 mg, 0.75 mmol), KN(SiMe₃)₂ (300 mg, 1.5

mmol), EuCl₃(THF)_{2.5} (110 mg, 0.1575 mmol) and DME (2.5 mL) gave after recrystallization title compound **1Eu^{iPr}** as an orange-red solid (121 mg, 0.11 mmol, 45%). ¹H NMR (400 MHz, C₆D₆) δ_H: -21.12 (1H, s), -14.06 (9H, s, C(CH₃)₃), -11.59 (3H, s, CH(CH₃)), -6.76 (9H, s, C(CH₃)₃), -5.94 (1H, s, CH), -5.90 (1H, s, CH), -5.38 (1H, s, CH), -2.60 (1H, s, CH), -1.63 (9H, s, C(CH₃)₃), -1.48 (9H, s, C(CH₃)₃), -1.44 (9H, s, C(CH₃)₃), -0.72 (1H, s, CH), -0.64 (1H, s, CH), -1.77 (3H, s, CH(CH₃)), 3.26 (1H, s, CH), 4.76 (3H, s, CH(CH₃)), 6.07 (1H, s, CH), 6.20 (1H, s, CH), 7.20 (1H, s, CH), 7.39 (1H, s, CH), 11.89 (9H, s, C(CH₃)₃), 15.05 (3H, s, CH(CH₃)), 15.85 (1H, s, CH), 17.88 (1H, s, CH), 24.72 (3H, s, CH(CH₃)), 33.23 (3H, s, CH(CH₃)), 49.20 (1H, s, CH), 96.66 (1H, s, CH). Elemental analysis C₆₀H₈₇EuO₃N₆: C 65.97%, H 8.03%, N 7.69% calculated. C 66.00%, H 8.01%, N 7.67% found; APPI⁺ C₆₀H₈₇EuN₆O₃⁺ [M]⁺ requires 1092.6052, found 1092.6095 (+3.9 ppm).

General procedure 2 - synthesis of 2Ce^R

A solution of **1Ce^R** (3 equiv.) in benzene, toluene, hexane or THF (0.5 M) was freeze-pump-thaw degassed 3 times and exposed to an atmosphere of dry CO₂ in a Teflon-valved ampoule. The solvent was removed under reduced pressure, and the crude product was extracted with toluene and concentrated to saturation and cooled to -30 °C overnight. The resulting suspension was filtered and dried under vacuum to yield the title compound which was stored at -20 °C under a nitrogen atmosphere.

2Ce^{iPr}. Using general procedure 2 - **1Ce^{iPr}** (3.0 g, 2.78 mmol) in toluene (50 mL) was charged with an atmosphere of CO₂ and after recrystallization gave the title product **2Ce^{iPr}** as a colourless solid (2.05 g, 1.69 mmol, 61%). Colourless crystals suitable for X-ray diffraction were grown from slow diffusion of hexanes into a concentrated THF solution. ¹H NMR (400 MHz, C₆D₆) δ_H: 0.91 (27H, s, C(CH₃)₃), 1.27-1.31 (9H, m, CH(CH₃)_a(CH₃)_b), 1.72-1.76 (9H, m, (CH(CH₃)_a(CH₃)_b), 2.52 (27H, s, C(CH₃)₃), 4.46-4.50 (3H, m, ArH), 4.62-4.66 (3H, m, ArH), 5.97 (3H, d, *J* 2.6, ImH), 7.59 (3H, d, *J* 2.6, ImH), 8.42 (3H, m, CH(CH₃)_a(CH₃)_b). ¹³C{¹H} NMR (126 MHz, C₆D₆) δ_C: 21.4 (CH(CH₃)_a(CH₃)_b), 24.6 (CH(CH₃)_a(CH₃)_b), 30.1 (C(CH₃)₃), 31.5 (C(CH₃)₃), 33.6 (C(CH₃)₃), 36.7 (C(CH₃)₃), 53.1 (CH(CH₃)_a(CH₃)_b), 112.4 (NCN), 119.1 (ArC), 119.5 (ArC), 123.6 (ImC), 125.2 (ImC), 134.6 (ArC), 139.4 (ArC), 143.2 (ArC), 154.9 (ArC), 173.1 (OCO). ν_{max} (nujol mull): 1666. Elemental analysis C₆₃H₈₇CeO₃N₆: C 62.25%, H 7.46%, N 6.91% calculated. C 62.36%, H 7.58%, N 7.08% found.

2Ce^{tBu}. Using general procedure 2 - **1Ce^{tBu}** (32 mg, 0.022 mmol) and THF (1 mL) was charged with an atmosphere of CO₂ and after recrystallization gave the title product **2Ce^{tBu}** as a colourless solid (20 mg, 16 mmol, 72%). ¹H NMR (400 MHz, C₆D₆) δ_H: 0.87 (27H, s, C(CH₃)₃), 2.60 (27H, s, C(CH₃)₃), 2.94 (27H, s, C(CH₃)₃), 4.58-4.62 (3H, m, ArH), 4.84-4.88 (3H, m, ArH), 5.91 (3H, d, *J* 2.5, ImH), 7.69 (3H, d, *J* 2.5, ImH). ¹³C{¹H} NMR (126 MHz, C₆D₆) δ_C: 30.0 (C(CH₃)₃), 30.1 (C(CH₃)₃), 31.1 (C(CH₃)₃), 33.2 (C(CH₃)₃), 36.4 (C(CH₃)₃), 62.5 (NC(CH₃)₃), 113.9 (ArC), 118.7 (ArC), 120.4 (ArC), 122.9 (ImC), 125.5 (ImC), 133.9 (ArC), 141.8 (ArC), 143.6 (ArC), 160.8 (ArC(2)OCe), 173.5 (OCO). Elemental analysis C₆₆H₉₃CeO₃N₆: C 62.25%, H 7.46%, N 6.91% calculated. C 62.36%, H 7.58%, N 7.08% found.



2Ce^{Mes}. Using a modification of general procedure 2 – **1Ce^{Mes}** (25 mg, 0.022 mmol) and THF (1 mL) was charged with an atmosphere of CO₂ and the resulting solution was left to slow evaporate to give the title product **2Ce^{Mes}** as a bright yellow solid (31 mg, 0.022 mmol, 99%). ¹H NMR (400 MHz, *d*₈-THF) δ_H: –9.96 (3H, br. s, CH₃), –6.09 (9H, br. s, C(CH₃)₃), –5.17 (9H, br. s, C(CH₃)₃), –2.95 (9H, br. s, C(CH₃)₃), –2.11 (3H, br. s, CH₃), 0.82 (9H, br. s, C(CH₃)₃), 0.97 (3H, br. s, CH₃), 1.13 (9H, br. s, C(CH₃)₃), 2.64 (12H, app. br. s, C(CH₃)₃ + CH₃), 3.05 (1H, br. s, CH), 3.51 (1H, br. s, CH), 4.46 (1H, br. s, CH), 4.74 (6H, app. br. s, 2 × CH₃), 5.61 (1H, br. s, CH), 5.72 (1H, br. s, CH), 6.06 (1H, br. s, CH), 6.28 (3H, br. s, CH₃), 7.04 (1H, br. s, CH), 7.51 (1H, br. s, CH), 8.36 (1H, br. s, CH), 8.61 (1H, br. s, CH), 9.10 (3H, br. s), 10.61 (1H, br. s, CH), 12.01 (1H, br. s, CH), 12.23 (1H, br. s, CH), 12.59 (1H, br. s, CH), 13.00–3.60 (4H, m, CH₃ + CH). Three CH resonances could not be located. ¹³C{¹H} NMR (126 MHz, *d*₈-THF) δ_C: 7.3, 13.3, 16.7, 17.4, 19.5, 19.7, 20.0, 20.3, 21.1, 21.5, 22.8, 23.4, 28.2, 29.9, 30.6, 32.0, 32.5, 33.8, 34.2, 34.5, 36.3, 36.5, 113.9, 119.3, 119.5, 120.6, 121.0, 121.7, 122.4, 123.1, 123.2, 124.8, 126.3, 126.6, 127.4, 127.7, 128.1, 129.5, 130.1, 130.9, 131.8, 132.2, 132.6, 133.2, 134.3, 134.7, 135.3, 135.7, 135.9, 136.7, 137.5, 138.5, 139.3, 140.1, 141.0, 141.0, 141.5, 141.8, 142.4, 143.4, 145.7, 148.0, 160.0, 163.3, 164.7, 169.6, 170.9, 175.4, 180.4, 200.2. ν_{max} (nujol mull): 1678, 1716. Elemental analysis C₈₁H₉₉CeO₉N₆: C 67.52% H 6.93% N 5.83% calculated. C 67.21% H 7.25% N 5.66% found.

3Ce^{iPr}(MesNCO)₃. To a solution of **1Ce^{iPr}** (108 mg, 0.1 mmol) in C₆H₆ (2 mL), MesNCO (48 mg, 0.03 mmol) was added and stirred for 15 min. The reaction mixture was filtered and cooled to –30 °C and the title product was isolated as a colourless powder by filtration of the solvents and drying under vacuum (123 mg, 79%). ¹H NMR (500 MHz, C₆D₆) δ_H: –6.53 (9H, s, C(CH₃)₃), –5.17 (3H, s, CH₃), –4.52 (3H, s, CH₃), –3.89 (3H, s, CH₃), –0.21 (3H, s, CH₃), 1.05 (3H, s, CH₃), 1.10 (3H, s, CH₃), 1.37 (9H, s, C(CH₃)₃), 1.53 (6H, s, Mes(2,6)CH₃), 1.72 (2H, s, Mes(3,5)H), 1.99 (9H, s, C(CH₃)₃), 2.00 (6H, s, Mes(2,6)CH₃), 2.02 (6H, s, Mes(2,6)CH₃), 2.22 (3H, s, CH₃), 2.56 (9H, s, C(CH₃)₃), 2.58 (3H, s, CH₃), 2.64 (9H, s, C(CH₃)₃), 2.99 (1H, s, CH), 3.26 (1H, s, CH), 3.44 (1H, s, CH), 3.64 (2H, s, Mes(3,5)H), 4.34 (1H, s, CH), 4.87 (1H, s, CH), 5.86 (1H, s, CH), 6.52 (2H, s, Mes(3,5)H), 6.62 (1H, s, CH), 6.79 (3H, s, CH₃), 7.27 (1H, s, CH₃), 7.35 (1H, s, CH), 7.71 (1H, s, CH), 9.07 (1H, s, CH), 10.04 (1H, s, CH), 10.58 (9H, s, C(CH₃)₃), 10.64 (1H, s, CH), 10.85 (1H, s, CH), 12.62 (1H, s, CH). Elemental analysis C₉₀H₁₂₀CeN₉O₆: C 69.11%, H 7.73%, N 8.06% calculated. C 69.09%, H 8.11%, N 7.93% found; APPI⁺ C₉₀H₁₂₁CeN₉O₆⁺ [M + H]⁺ requires 1563.8494, found 1563.8419 (–4.8 ppm).

3Ce^{iPr}(^tBuNCO)₃. To a solution of **1Ce^{iPr}** (108 mg, 0.1 mmol) in DME (2 mL), ^tBuNCO (20 mg, 0.3 mmol) was added and stirred for 15 min. The reaction mixture was filtered into hexane (1 mL) and cooled to –30 °C and the title product was isolated as a colourless powder by filtration of the solvents and drying under vacuum (126 mg, 91%). ¹H NMR (500 MHz, C₆D₆) (*fac*)-**3Ce^{iPr}(^tBuNCO)₃** δ_H: –6.17 (27H, s), –5.18 (9H, s), –3.72 (27H, s), –3.48 (3H, s), 0.50 (9H, s), 5.52 (27H, s), 9.91 (3H, s), 12.68 (3H, s), 17.99 (3H, s), 21.06 (3H, s). (*mer*)-**3Ce^{iPr}(^tBuNCO)₃** δ_H:

–12.69 (1H, s), –12.11 (9H, s), –7.69 (3H, s), –6.97 (3H, s), –5.96 (9H, s), –5.39 (9H, s), –5.00 (3H, s), –4.86 (3H, s), –2.67 (1H, s), –2.26 (9H, s), –2.08 (1H, s), –1.45 (1H, s), –1.09 (9H, s), –0.95 (3H, s), 0.14 (9H, s), 1.82 (3H, s), 2.11 (1H, s), 5.11 (9H, s), 5.26 (9H, s), 5.36 (9H, s), 5.94 (1H, s), 8.73 (1H, s), 9.24 (3H, s), 9.91 (1H, s), 12.01 (1H, s), 12.25 (1H, s), 12.62 (1H, s), 16.82 (1H, s), 17.31 (1H, s), 19.71 (1H, s), 20.30 (1H, s). APPI⁺ C₇₅H₁₁₆CeN₉O₇⁺ [M + H₂O]⁺ requires 1394.8052, found 1394.8426 (+26.8 ppm). Elemental analysis C₇₇H₁₁₄CeN₉O₆: C 65.38%, H 8.34%, N 9.15% calculated. C 65.52%, H 8.45%, N 8.98% found.

3Ce^{iPr}(^tBuNCO)₂. To a solution of **1Ce^{iPr}** (108 mg, 0.1 mmol) in C₆H₆ or THF (2 mL), ^tBuNCO (20 mg, 0.3 mmol) was added and stirred for 15 min. The reaction mixture was filtered into hexane (1 mL) and cooled to –30 °C and the title product was isolated as a pale-yellow powder by filtration of the solvents and drying under vacuum (76 mg, 59%). ¹H NMR (500 MHz, C₆D₆) δ_H: –6.95 (9H, s), –4.75 (3H, s), –2.95 (1H, s), –2.62 (3H, s), –1.00 (9H, s), –0.1 (9H, s), 0.13 (3H, s), 0.19 (9H, s), 0.59–0.64 (3H, m), 0.66 (1H, s), 1.05 (1H, s), 1.17 (3H, s), 1.35 (1H, s), 1.44 (1H, s), 1.77 (9H, s), 2.31 (9H, s), 2.52 (9H, s), 3.42 (1H, s), 4.15 (1H, s), 5.96 (3H, s), 6.28 (1H, s), 7.03 (1H, s), 7.24 (1H, s), 9.06 (1H, s), 9.38 (9H, s), 9.88 (1H, s), 10.38 (1H, s), 10.50 (1H, s), 12.30 (1H, s). APPI⁺ C₇₀H₁₀₆CeN₈O₅⁺ [M + H]⁺ requires 1278.7341, found 1278.7213 (–10.0 ppm). Elemental analysis C₇₀H₁₀₅CeN₈O₅: C 65.75%, H 8.28%, N 8.75% calculated. C 65.50%, H 8.58%, N 8.64% found.

3Ce^{iPr}(^tBuNCS)₂. To a solution of **1Ce^{iPr}** (108 mg, 0.1 mmol) in DME (2 mL), ^tBuNCS (34 μL, 0.3 mmol) was added and stirred for 2 h at 80 °C. The reaction mixture was cooled to room temperature then evaporated to dryness. Colourless X-ray quality crystals were grown by diffusion of heptane into a toluene solution of the crude product, and isolated by decanting (89 mg, 68%). ¹H NMR (500 MHz, C₆D₆) δ_H: –9.13 (1H, s), –8.58 (3H, s), –8.29 (1H, s), –6.14 (3H, s), –4.01 (3H, s), –3.97 (9H, s), –3.53 (3H, s), –3.00 (9H, s), –2.28 (9H, s), –1.34 (1H, s), 0.23 (9H, s), 1.38 (1H, s), 1.68 (9H, s), 3.13 (1H, s), 3.24 (3H, s), 3.25 (9H, s), 4.78 (1H, s), 6.07 (1H, s), 8.29 (1H, s), 8.37 (1H, s), 9.15 (9H, s), 9.76 (1H, s), 9.90 (1H, s), 10.86 (1H, s), 13.99 (1H, s), 14.23 (3H, s), 17.32 (1H, s), 18.17 (9H, s), 52.70 (1H br. s). Elemental analysis C₇₀H₁₀₅CeN₈O₃S₂: C 64.14%, H 8.07%, N 8.55% calculated. C 64.17%, H 8.35%, N 8.24% found. APPI⁺ C₇₀H₁₀₆CeN₈O₃S₂⁺ [M + H]⁺ requires 1310.6884, found 1310.6816 (–5.2 ppm).

3Ce^{iPr}(^tBuNCS). To a solution of **1Ce^{iPr}** (108 mg, 0.1 mmol) in DME (2 mL), ^tBuNCS (34 μL, 0.3 mmol) was added and stirred for 2 h at 80 °C. The reaction mixture was cooled to room temperature and evaporated to dryness yielding a colourless powder (117 mg, 98%). ¹H NMR (500 MHz, C₆D₆) δ_H: –8.72 (3H, s), –8.44 (1H, s), –6.21 (3H, s), –4.08 (3H, s), –4.06 (9H, s), –3.74 (1H, s), –3.61 (3H, s), –2.35 (9H, s), –1.49 (1H, s), –0.10 (1H, s), 0.32 (1H, s), 1.50 (9H, s), 1.67 (9H, s), 1.98 (9H, s), 2.82 (1H, s), 3.25 (9H, s), 3.32 (3H, s), 3.51 (1H, s), 4.69 (1H, s), 6.08 (1H, s), 6.88 (1H, s), 8.34 (1H, s), 9.25–9.35 (9H, m), 9.82 (1H, s), 9.87 (1H, s), 14.04 (1H, s), 14.45 (3H, s), 17.40 (1H, s), 18.47 (9H, s). APPI⁺ C₆₅H₉₆CeN₇O₃S⁺ [M]⁺ requires 1194.6350, found 1194.6571 (+18.5 ppm). Elemental analysis C₆₅H₉₆CeN₇O₃S: C



65.29%, H 8.09%, N 8.20% calculated. C 65.42%, H 8.21%, N 7.59% found.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

We thank the EPSRC for funding through the Centre for Doctoral Training in Critical Resource Catalysis (CRITICAT, EP/L016419/1, R. W. F. K.), EP/J018139/1 and the UK Catalysis Hub (EP/K014714/1, P. L. A., C. W.), EP/M010554/1 (P. L. A.). This project has received funding from the European Research Council (ERC) under the European Union's Horizon 2020 research and innovation programme (grant agreement No. 740311, P. L. A.). K. W. thanks the China Scholarship Council (CSC) for a postgraduate fellowship. P. L. A., M. W. M., J. R. and F. E. K. thank the Technische Universität München – Institute for Advanced Study, funded by the German Excellence Initiative. A. D. S. thanks the Royal Society for a Wolfson Research Merit Award. C. J. thanks the DAAD for a scholarship, and C. J. and J. R. thank the TUM Graduate School for financial support. We would also like to thank Dr Colin Logan Mackay of the University of Edinburgh for mass spectrometry and the University of Edinburgh SIRCAMS facility for use of the FTICRMS.

Notes and references

- D. H. Gibson, *Chem. Rev.*, 1996, **96**, 2063–2096.
- J. Artz, T. E. Müller, K. Thenert, J. Kleinekorte, R. Meys, A. Sternberg, A. Bardow and W. Leitner, *Chem. Rev.*, 2018, **118**, 434–504.
- Q. Liu, L. Wu, R. Jackstell and M. Beller, *Nat. Commun.*, 2015, **6**, 5933–5948.
- S. Ozaki, *Chem. Rev.*, 1972, **72**, 457–496.
- B. Bantu, G. M. Pawar, U. Decker, K. Wurst, M. Schmidt Axel and R. Buchmeiser Michael, *Chem.–Eur. J.*, 2009, **15**, 3103–3109.
- B. C. Norris and C. W. Bielawski, *Macromolecules*, 2010, **43**, 3591–3593.
- H. A. Duong, T. N. Tekavec, A. M. Arif and J. Louie, *Chem. Commun.*, 2004, 112–113.
- B. R. Van Ausdall, J. L. Glass, K. M. Wiggins, A. M. Aarif and J. Louie, *J. Org. Chem.*, 2009, **74**, 7935–7942.
- H. A. Duong, M. J. Cross and J. Louie, *Org. Lett.*, 2004, **6**, 4679–4681.
- B. C. Norris, D. G. Sheppard, G. Henkelman and C. W. Bielawski, *J. Org. Chem.*, 2011, **76**, 301–304.
- H. Zhou, W.-Z. Zhang, C.-H. Liu, J.-P. Qu and X.-B. Lu, *J. Org. Chem.*, 2008, **73**, 8039–8044.
- A. J. Arduengo, M. Tamm, S. J. McLain, J. C. Calabrese, F. Davidson and W. J. Marshall, *J. Am. Chem. Soc.*, 1994, **116**, 7927–7928.
- R. D. Fischer, *Angew. Chem., Int. Ed. Engl.*, 1994, **33**, 2165–2168.
- (a) P. L. Arnold and I. J. Casely, *Chem. Rev.*, 2009, **109**, 3599–3611; (b) H. Schumann, M. Glanz, J. Winterfeld, H. Hemling, N. Kuhn and T. Kratz, *Angew. Chem., Int. Ed.*, 1994, **33**, 1733–1734.
- T. Mehdoui, J.-C. Berthet, P. Thuéry and M. Ephritikhine, *Chem. Commun.*, 2005, 2860–2862.
- Y. Haisheng, Z. Yong, S. Hongmei and S. Qi, *Eur. J. Inorg. Chem.*, 2009, **2009**, 1920–1925.
- H. Yao, J. Zhang, Y. Zhang, H. Sun and Q. Shen, *Organometallics*, 2010, **29**, 5841–5846.
- J. Zhang, H. Yao, Y. Zhang, H. Sun and Q. Shen, *Organometallics*, 2008, **27**, 2672–2675.
- Z.-G. Wang, H.-M. Sun, H.-S. Yao, Q. Shen and Y. Zhang, *Organometallics*, 2006, **25**, 4436–4438.
- P. L. Arnold, I. A. Marr, S. Zlatogorsky, R. Bellabarba and R. P. Tooze, *Dalton Trans.*, 2014, **43**, 34–37.
- ThermoFisher Scientific, Safety Data Sheet: FeCl₃, <https://www.fishersci.com/msds?productName=AC217091000&product>, accessed 9 July, 2018.
- ThermoFisher Scientific, Safety Data Sheet: CeCl₃, <https://www.fishersci.com/shop/msdsproxy?productName=AC369750050&productDescription=CERIUM>, accessed 9 Jul, 2018.
- T. Montini, M. Melchionna, M. Monai and P. Fornasiero, *Chem. Rev.*, 2016, **116**, 5987–6041.
- P. L. Arnold, Z. R. Turner, A. I. Germeroth, I. J. Casely, G. S. Nichol, R. Bellabarba and R. P. Tooze, *Dalton Trans.*, 2013, **42**, 1333–1337.
- W. Ding and D. Wang, *Organometallics*, 2014, **33**, 7007–7010.
- G. Prabusankar and P. Suresh, *ChemistrySelect*, 2017, **2**, 9920–9923.
- H. Ren, P. Yao, S. Xu, H. Song and B. Wang, *J. Organomet. Chem.*, 2007, **692**, 2092–2098.
- M. Nirmala, G. Prakash, P. Viswanathamurthi and J. G. Malecki, *J. Mol. Catal. A: Chem.*, 2015, **403**, 15–26.
- L. Arnold Polly, R. Turner Zoë, N. Kaltsoyannis, P. Pelekanaki, M. Bellabarba Ronan and P. Tooze Robert, *Chem.–Eur. J.*, 2010, **16**, 9623–9629.
- I. J. Casely, S. T. Liddle, A. J. Blake, C. Wilson and P. L. Arnold, *Chem. Commun.*, 2007, 5037–5039.
- C. R. Martinez and B. L. Iverson, *Chem. Sci.*, 2012, **3**, 2191–2201.
- M. O. Sinnokrot, E. F. Valeev and C. D. Sherrill, *J. Am. Chem. Soc.*, 2002, **124**, 10887–10893.
- A. Tudose, A. Demonceau and L. Delaude, *J. Organomet. Chem.*, 2006, **691**, 5356–5365.
- E. Brule, V. Guérineau, P. Vermaut, F. Prima, J. Balogh, L. Maron, A. M. Z. Slawin, S. P. Nolan and C. M. Thomas, *Polym. Chem.*, 2013, **4**, 2414–2423.
- H. Büttner, L. Longwitz, J. Steinbauer, C. Wulf and T. Werner, *Top. Curr. Chem.*, 2017, **375**, 50.
- J. Eppinger, E. Herdtweck and R. Anwender, *Polyhedron*, 1998, **17**, 1195–1201.

