



Cite this: *EES Catal.*, 2023,  
1, 54

## Direct carbonate electrolysis into pure syngas†

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Syngas, a mixture of carbon monoxide (CO) and hydrogen (H<sub>2</sub>), is a feedstock for a wide variety of chemical processes and is currently produced from fossil fuels. The need to reduce carbon dioxide (CO<sub>2</sub>) emissions motivates the production of syngas from atmospheric CO<sub>2</sub>, powered by renewable electricity. Current CO<sub>2</sub> electrolyzers require costly separation processes to purify the CO<sub>2</sub> reactant stream and to remove unreacted CO<sub>2</sub> from the product stream. We demonstrate direct carbonate electrolysis (DCE) in a reactive capture system that avoids the initial CO<sub>2</sub> purification process and produces pure syngas with sufficient CO content for direct industrial use (H<sub>2</sub>/CO ratios of 1–2). The DCE system incorporates a composite CO<sub>2</sub> diffusion layer (CDL) that attains high CO selectivity by achieving high alkalinity and available CO<sub>2</sub> concentration at the cathode. Applying this strategy, we produce pure syngas in the cathode outlet gas stream with a H<sub>2</sub>/CO ratio of 1.16 at 200 mA cm<sup>-2</sup>, corresponding to a CO faradaic efficiency (FE) of 46% and an energy intensity of 52 GJ tsyngas<sup>-1</sup>. By eliminating intensive upstream and downstream processes, DCE achieves syngas production with 13% less energy than CO<sub>2</sub> electrolysis combined with water electrolysis, 39% less energy than past carbonate reduction work, and 75% fewer emissions than the conventional fossil fuel based route.

Received 15th September 2022,  
Accepted 1st November 2022

DOI: 10.1039/d2ey00046f

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### Broader context

The electrochemical conversion of captured CO<sub>2</sub> into CO could reduce CO<sub>2</sub> emissions while producing the carbon content of a valuable feedstock, syngas, for upgrade into long-chain hydrocarbons. Most CO<sub>2</sub> electrolyzers require pure gaseous CO<sub>2</sub> streams, and, thus, incur substantial capital and operational costs for CO<sub>2</sub> capture liquid regeneration and CO<sub>2</sub> purification. Additionally, CO<sub>2</sub> reactant can be lost to carbonates in the electrolyzer which crossover to the anode, regenerate, and mix with the O<sub>2</sub> rich anode gas stream. Excess CO<sub>2</sub> in the cathode gas product stream also demands separation. These sequential purification steps are costly and limit the viability of electroproduced chemicals and fuels. We demonstrate a reactive capture system that produces pure syngas with an industrial H<sub>2</sub>/CO ratio through direct electrolysis of a CO<sub>2</sub> post-capture solution. We designed a CO<sub>2</sub> diffusion layer that achieves high local CO<sub>2</sub> reactant concentration and high alkalinity favourable for CO<sub>2</sub> conversion. This report illustrates the potential for the renewable electroproduction of syngas in a net-zero emissions future.

## Introduction

Syngas is a commodity feedstock used in the production of hydrocarbons and oxygenates *via* methanol routes and Fischer–Tropsch synthesis.<sup>1,2</sup> Syngas is currently produced from fossil-fuels *via* coal gasification and/or methane reforming.<sup>3,4</sup> These

pathways are energy intensive and have high CO<sub>2</sub> emission intensities (1.5 tCO<sub>2</sub>e tsyngas<sup>-1</sup>).<sup>5</sup>

Electrochemical syngas production methods use renewable electricity to produce syngas with a lower carbon footprint. These methods combine CO from CO<sub>2</sub> electrolysis and H<sub>2</sub> from water electrolysis.<sup>6–8</sup> However, current CO<sub>2</sub> electrolyzers require high-purity gaseous CO<sub>2</sub> feeds,<sup>9–12</sup> and sourcing this CO<sub>2</sub> from air – *via* direct air capture – or from industrial sources, is costly.<sup>13–16</sup> Within current electrolyzers, utilization of reactant CO<sub>2</sub> is low, and CO<sub>2</sub> is lost to carbonates and crossover to the anode.<sup>17–19</sup> As a result, both the anodic and the cathodic outlet streams require CO<sub>2</sub> separation.<sup>20,21</sup>

Reactive capture is an electrolysis approach that shortens this process (Fig. 1A). This pathway avoids the thermally-driven

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† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d2ey00046f>





**Fig. 1** Direct carbonate electrolysis to produce syngas enabled by a CDL. The chemical balance of carbon capture from air, anodic OER,  $\text{CO}_2$  regeneration, CO evolution reaction,  $\text{H}_2$  evolution reaction, and *in situ*  $\text{CO}_2$  recapture by  $\text{OH}^-$  are presented in eqn (1)–(6), respectively. (A) Schematic of DCE integrated with carbon capture in a reactive capture system. (B) Schematic of the conversion of carbonate into  $\text{CO}_2$  facilitated by the CDL and protons, and subsequent conversion of  $\text{CO}_2$  into CO via electroreduction on the Ag catalyst. Unreacted  $\text{CO}_2$  is recaptured by  $\text{OH}^-$  generated as a by-product of CO and  $\text{H}_2$  evolution. The carbonate electrolyte is recirculated.

capture liquid regeneration from the post-capture solution (carbonate electrolyte in the case of hydroxide-based direct air capture) and the subsequent  $\text{CO}_2$  dehydration, compression, and transportation steps.<sup>22–26</sup> In the electrolyzer, protons are generated by the anodic oxygen evolution reaction (OER) and transported to the cathode through a cation exchange membrane (CEM) (Fig. 1B). The protons then react with the carbonate ions to regenerate  $\text{CO}_2$  *in situ*. Syngas is produced through co-synthesis of CO from the regenerated  $\text{CO}_2$  and  $\text{H}_2$  from the aqueous solution. Unreacted  $\text{CO}_2$  is recaptured by hydroxide ions ( $\text{OH}^-$ ), a by-product of CO and  $\text{H}_2$  evolution, to form carbonate.

Previous studies of reactive capture using bicarbonate electrolyte have demonstrated high CO selectivity (CO faradaic efficiency (FE) > 50%). However, the limited  $\text{CO}_2$  recapture capacity of bicarbonate electrolyte results in a product gas stream diluted with  $\text{CO}_2$ .<sup>26–28</sup> Direct electrolysis of carbonate (rather than bicarbonate) electrolyte allows for the collection of high-purity gaseous products, evidenced by the lack of  $\text{CO}_2$  (<400 ppm) detected in the gas stream.<sup>29</sup> Carbonate electrocatalytic conversion into syngas has previously been achieved

with a  $\text{H}_2/\text{CO}$  ratio of 3 (CO FE of 25%) and energy intensity of  $86 \text{ GJ tsyngas}^{-1}$ . However, this mixture does not have sufficient CO content to meet industrial syngas standards.<sup>30–34</sup>

Here, we present an adlayer strategy that modulates the cathode pH and maximizes  $\text{CO}_2$  conversion to produce syngas with a  $\text{H}_2/\text{CO}$  ratio in the industrially relevant range (1–2, and corresponding to a CO FE between 33% to 50%). We develop a composite  $\text{CO}_2$  diffusion layer (CDL) that enables cathode alkalinity to favour  $\text{CO}_2$  electrolysis and increases CO selectivity by limiting the diffusion of protons to the cathode. We achieve a  $\text{H}_2/\text{CO}$  ratio of 1.16 (CO FE of 46%) at a current density of  $200 \text{ mA cm}^{-2}$ . An energy intensity of  $52 \text{ GJ tsyngas}^{-1}$  was achieved, resulting in a 39% energy saving compared to the previous carbonate electrolysis report.

## Results & discussion

### Increasing $\text{CO}_2$ conversion

Previous (bi)carbonate electrolyzers have employed a bipolar membrane (BPM) to dissociate water and provide protons and



$\text{OH}^-$  to the cathode and anode, respectively.<sup>26,29,35</sup> To reduce the membrane overvoltage, thereby reducing the required energy input, we used a cation exchange membrane (CEM) to transport protons directly from the anolyte or anodic OER, resulting in a voltage reduction of *ca.* 0.5 V (Fig. S1, ESI†). A zero-gap configuration was first investigated by assembling the electrolyzer with an Ag electrocatalyst directly in contact with the CEM. We found that the maximum FE towards CO was 17% ( $\text{H}_2/\text{CO}$  ratio of 4.88) at  $100 \text{ mA cm}^{-2}$  (Fig. 2A). The only other carbon-based product detected was methane with  $<0.2\%$  FE. Hydrogen evolution reaction (HER) accounted for the remaining FE.

We hypothesized that the low CO FE was caused by a CEM-induced acidic environment in which the kinetically more favourable HER outcompetes  $\text{CO}_2$  electrolysis,<sup>36–38</sup> and, thus, modulating the cathode pH would improve  $\text{CO}_2$  conversion.<sup>39–41</sup> We confirmed that  $\text{CO}_2$  was the electroreduction reactant, as opposed to the (bi)carbonate ions, by replacing the acidic anolyte with an alkaline electrolyte to suppress *in situ*  $\text{CO}_2$  regeneration. The only product detected was  $\text{H}_2$  (Fig. S2, ESI†). To increase  $\text{CO}_2$  conversion, we hypothesized that a  $\text{CO}_2$  diffusion adlayer between the CEM and the catalyst would limit proton diffusion to the cathode and separate the acidic  $\text{CO}_2$  regeneration region from the alkaline  $\text{CO}_2$  electrolysis region.

We developed multi-physics models of the carbonate electrolyzer with varying CDL thicknesses of 0 (zero-gap configuration), 10, 25, and  $50 \mu\text{m}$  (Supplementary note 2, ESI†). We found that increasing the CDL thickness increased the pH at the cathode which favoured  $\text{CO}_2$  reduction over HER (Fig. 2B).<sup>42,43</sup> However, increasing the CDL thickness also reduced the  $\text{CO}_2$  concentration at the cathode due to the recapturing of *in situ*  $\text{CO}_2$  within the extended alkaline region (Fig. 2C). The CDL must achieve high local cathode alkalinity and  $\text{CO}_2$  concentration to produce syngas with sufficient CO content for direct industrial application.

### CDL design strategy

We first inserted commercially available hydrophilic microporous filters of different thicknesses between the catalyst and the CEM

to separate the acidic  $\text{CO}_2$  regeneration region from the alkaline electrolysis region. We found that the filters increased the selectivity towards CO (Fig. S6, ESI†). However, the performance of the inserted filters was inconsistent due to trapped  $\text{CO}_2$  bubbles and material incompatibilities with the high and low pH extremes in this system ( $11 < \text{pH} < 2$ ). We were motivated to design a robust and tuneable composite CDL. The engineered CDL needed to facilitate (bi)carbonate diffusion, hinder proton transport to the catalyst, sustain both high and low pH conditions, and have porous networks that allow mass transfer of *in situ* generated  $\text{CO}_2$ . We selected  $\text{TiO}_2$  particles as the main substrate in view of their chemical stability and hydrophilic nature.<sup>44,45</sup> For a substrate binder, we chose a hydrophilic ionomer that is anion-permeable to transport (bi)carbonate ions to the CEM.<sup>46,47</sup> The wettability of the substrate and ionomer is important because hydrophobic elements will hinder ion transport, limiting the availability of carbonate ions in the  $\text{CO}_2$  regeneration region.<sup>48</sup> When a hydrophobic substrate or ionomer was incorporated into the CDL, the CO FE was lower than the fully hydrophilic system (Fig. S7A and B, ESI†). The  $\text{TiO}_2$  and ionomer mixture was air-brushed evenly onto an Ag-catalyst layer (Fig. 3A and B) with corresponding energy dispersive X-ray (EDX) spectroscopy images confirming a distinct and uniform CDL (Fig. 3C).

To optimize the CDL for high CO FE, we screened  $\text{TiO}_2$  particle sizes (5, 25, 200, and  $1500 \text{ nm}$ ) and  $\text{TiO}_2/\text{ionomer}$  weight ratios between 5–25 (Fig. S7C and D, ESI†). We found that a combination of  $25 \text{ nm}$   $\text{TiO}_2$  and a  $\text{TiO}_2/\text{ionomer}$  ratio of 15 balanced the diffusion of (bi)carbonate ions and protons and enabled the local generation of  $\text{CO}_2$  to result in peak CO FE. The size of  $\text{TiO}_2$  nanoparticles and the ionomer volume fraction contribute to the permeability of the CDL. A high permeability failed to sufficiently hinder proton transport and resulted in hydrogen generation. A low permeability resulted in insufficient *in situ* regeneration of reactant  $\text{CO}_2$ .<sup>49</sup>

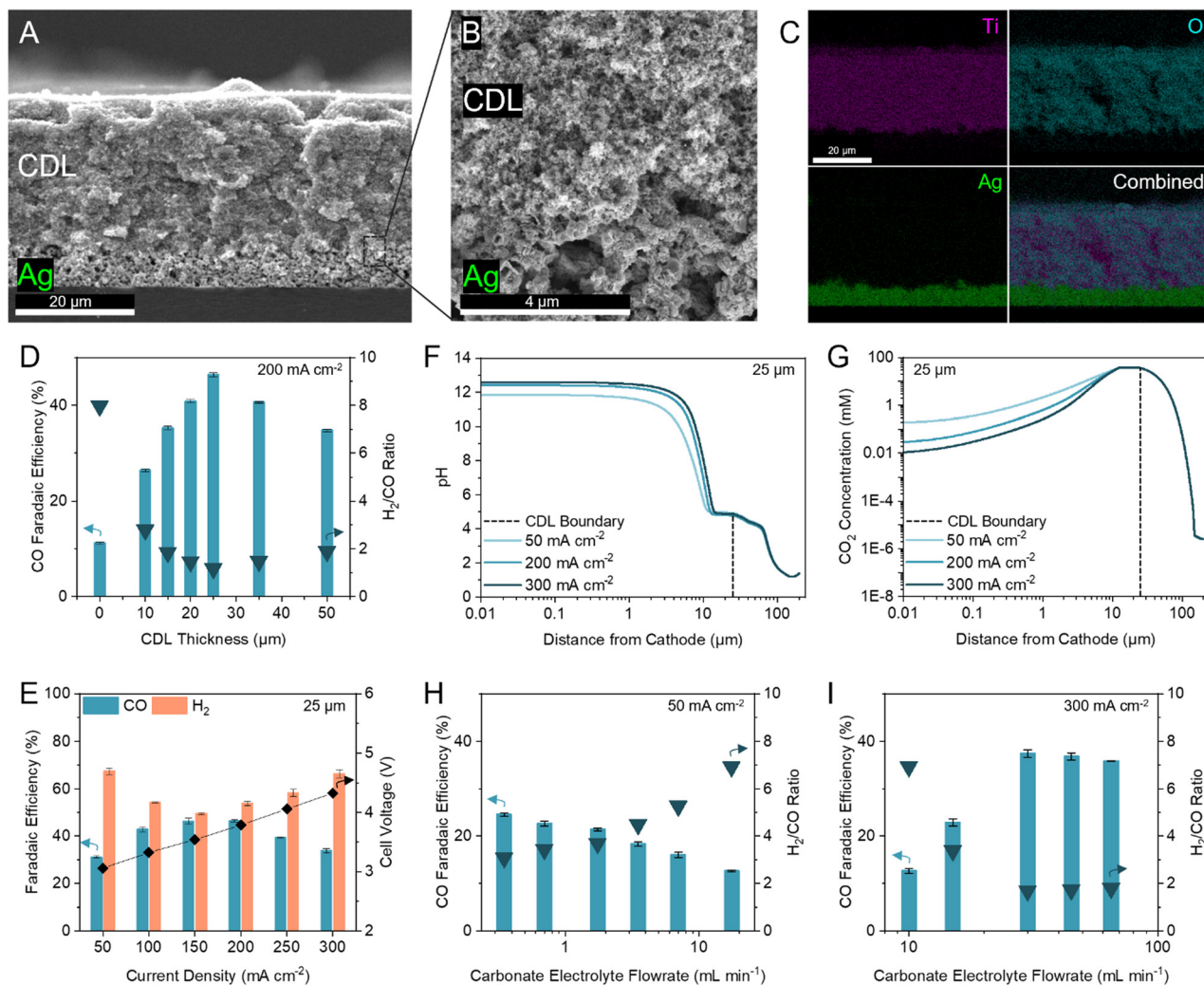
We varied the CDL thickness between 10 to  $50 \mu\text{m}$  and achieved a maximum CO FE of 46% ( $\text{H}_2/\text{CO}$  ratio of 1.16) at  $200 \text{ mA cm}^{-2}$  with a  $25 \mu\text{m}$  CDL (Fig. 3D). Achieving a high CO FE requires both a sufficiently alkaline local pH and adequate



**Fig. 2** CDL thickness modulates cathode pH and  $\text{CO}_2$  concentration. (A) FE towards CO and  $\text{H}_2$  in a zero-gap configuration at current densities between  $50$  to  $200 \text{ mA cm}^{-2}$ . Corresponding full cell voltages are noted on the secondary y-axis. Error bars represent the standard deviation of at least three samples measured under identical conditions. (B) One-dimensional multi-physics modelling of pH at distances from the cathode and current density of  $200 \text{ mA cm}^{-2}$  for CDL with thickness of 0 (zero-gap), 10, 25, and  $50 \mu\text{m}$ . (C) One-dimensional multi-physics modelling of  $\text{CO}_2$  concentration at distances from the cathode and current density of  $200 \text{ mA cm}^{-2}$  for CDL with thickness of 0 (zero-gap), 10, 25, and  $50 \mu\text{m}$ . Accompanying models of  $\text{HCO}_3^-$  and  $\text{CO}_3^{2-}$  concentrations are provided in the ESI† (Fig. S4).







**Fig. 3** Optimization of the CDL for industrial  $\text{H}_2/\text{CO}$  ratio. (A–C) Cross-sectional scanning electron microscopy (SEM) image of the CDL evenly air-brushed onto the Ag catalyst atop a silicon wafer (A), with seamless interfacial contact between the CDL and Ag catalyst (B), and corresponding energy dispersive X-ray (EDX) spectroscopy elemental mapping of Ti, O, and Ag (C). (D) FE towards CO and  $\text{H}_2/\text{CO}$  ratio at  $200 \text{ mA cm}^{-2}$  with thicknesses of the CDL between 0 and  $50 \mu\text{m}$ . (E) FE towards CO and  $\text{H}_2$  at current densities between  $50$  to  $300 \text{ mA cm}^{-2}$  with  $25 \mu\text{m}$  CDL. Corresponding full cell voltages are noted on the secondary y-axis. (F) One-dimensional multi-physics modelling of pH at distances from the cathode and current densities of  $50$ ,  $200$ , and  $300 \text{ mA cm}^{-2}$  for  $25 \mu\text{m}$  CDL. (G) One-dimensional multi-physics modelling of  $\text{CO}_2$  concentration at distances from the cathode and current densities of  $50$ ,  $200$ , and  $300 \text{ mA cm}^{-2}$  for  $25 \mu\text{m}$  CDL. Accompanying models of  $\text{HCO}_3^-$  and  $\text{CO}_3^{2-}$  concentrations are provided in the ESI† (Fig. S5). (H) FE towards CO and  $\text{H}_2/\text{CO}$  ratio at  $50 \text{ mA cm}^{-2}$  with carbonate electrolyte flowrates between  $0.35$  to  $17.5 \text{ mL min}^{-1}$  showing an increase in CO FE with decreasing flowrate. (I) FE towards CO and  $\text{H}_2/\text{CO}$  ratio at  $300 \text{ mA cm}^{-2}$  with carbonate electrolyte flowrates between  $10$  to  $65 \text{ mL min}^{-1}$  showing an increase in CO FE with increasing flowrate. Typical carbonate electrolyte flowrates are provided in the ESI† (Table S5). Error bars represent the standard deviation of at least three samples measured under identical conditions.

$\text{CO}_2$  availability. Thinner CDLs have a smaller gap between the catalyst and the CEM which shortens the proton diffusion distance and results in a lower cathode pH. Despite having the highest  $\text{CO}_2$  concentrations in our simulations, the selectivity of the system with thinner CDLs was not optimal and approached the performance of the zero-gap configuration due to insufficient cathode alkalinity (Fig. 2B). As the CDL thickness is increased, the pH at the cathode increased but the local  $\text{CO}_2$  concentration decreased (Fig. 2C). With the cathode pH plateauing at thicknesses greater than  $25 \mu\text{m}$ , this CDL thickness provided sufficient  $\text{CO}_2$  availability while maintaining an alkaline cathode environment to suppress HER.

We compared the difference in  $iR$ -compensated voltage of a zero-gap configuration to an otherwise identical system with a  $25 \mu\text{m}$  CDL, with both experiments using an identical hydrogen-evolving catalyst (Supplementary note 3, ESI†). At  $200 \text{ mA cm}^{-2}$ , a voltage increase of  $117 \text{ mV}$  was observed for the CDL system which corresponded to an increase in pH of  $2$ , consistent with the multi-physics model which predicted a pH increase of  $1.7$  (Fig. 2B and Fig. S8, ESI†). These findings suggest that the CDL increases the pH of the cathode environment to favour  $\text{CO}_2$  conversion and thereby yielded a higher CO FE.

We screened the selectivity and full cell voltage of the optimized  $25 \mu\text{m}$  CDL at current densities between  $50$  to



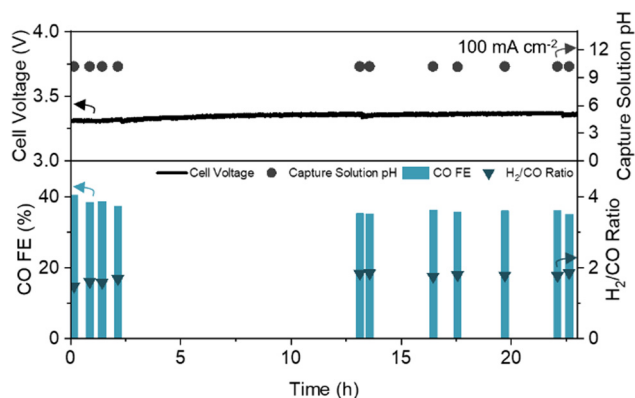


Fig. 4 Long-term operation of DCE over 23 hours with CO FE,  $H_2/CO$  ratio, full cell voltage, and capture solution pH noted. Experiment conducted at a constant current density of  $100 \text{ mA cm}^{-2}$  with a  $1 \text{ cm}^2$  active area. Schematic and picture of the experimental set-up are provided in the ESI† (Fig. S10A and B).

$300 \text{ mA cm}^{-2}$  and found that the CDL resulted in minimal voltage penalties while significantly increasing CO FE compared to the zero-gap configuration (+0.23 V and +35.3% CO

FE at  $200 \text{ mA cm}^{-2}$ ) (Fig. 3E). Above  $200 \text{ mA cm}^{-2}$ , the CO selectivity decreased, and the CO partial current plateaued, which indicated a  $CO_2$  mass transfer limit at the cathode (Fig. S9, ESI†). We hypothesized that the decrease in CO FE at the lower and higher current densities were due to an imbalance of cathode alkalinity and  $CO_2$  concentration at these extremes. The multi-physics model showed that compared to operating at  $200 \text{ mA cm}^{-2}$ , the pH is lower at  $50 \text{ mA cm}^{-2}$ , while the  $CO_2$  concentration is lower at  $300 \text{ mA cm}^{-2}$  (Fig. 3F and G). To investigate further, we decreased the carbonate flow rate while operating at  $50 \text{ mA cm}^{-2}$  and found that the CO FE increased (Fig. 3H). A slower flowrate increases the local pH due to the accumulation of  $OH^-$ . At  $300 \text{ mA cm}^{-2}$ , the CO FE increased as the carbonate electrolyte flowrate increased up to  $30 \text{ mL min}^{-1}$ ; however, further increases in flowrate resulted in similar, or slightly decreased, CO FE (Fig. 3I). This result suggests that at high current densities, the local environment is excessively alkaline, and  $CO_2$  availability is low. Increasing the carbonate electrolyte flowrate lowered the local pH and thereby increased the availability of  $CO_2$  for reaction.

To assess the long-term stability of the engineered CDL, we operated the DCE system with continuous  $CO_2$  capture and

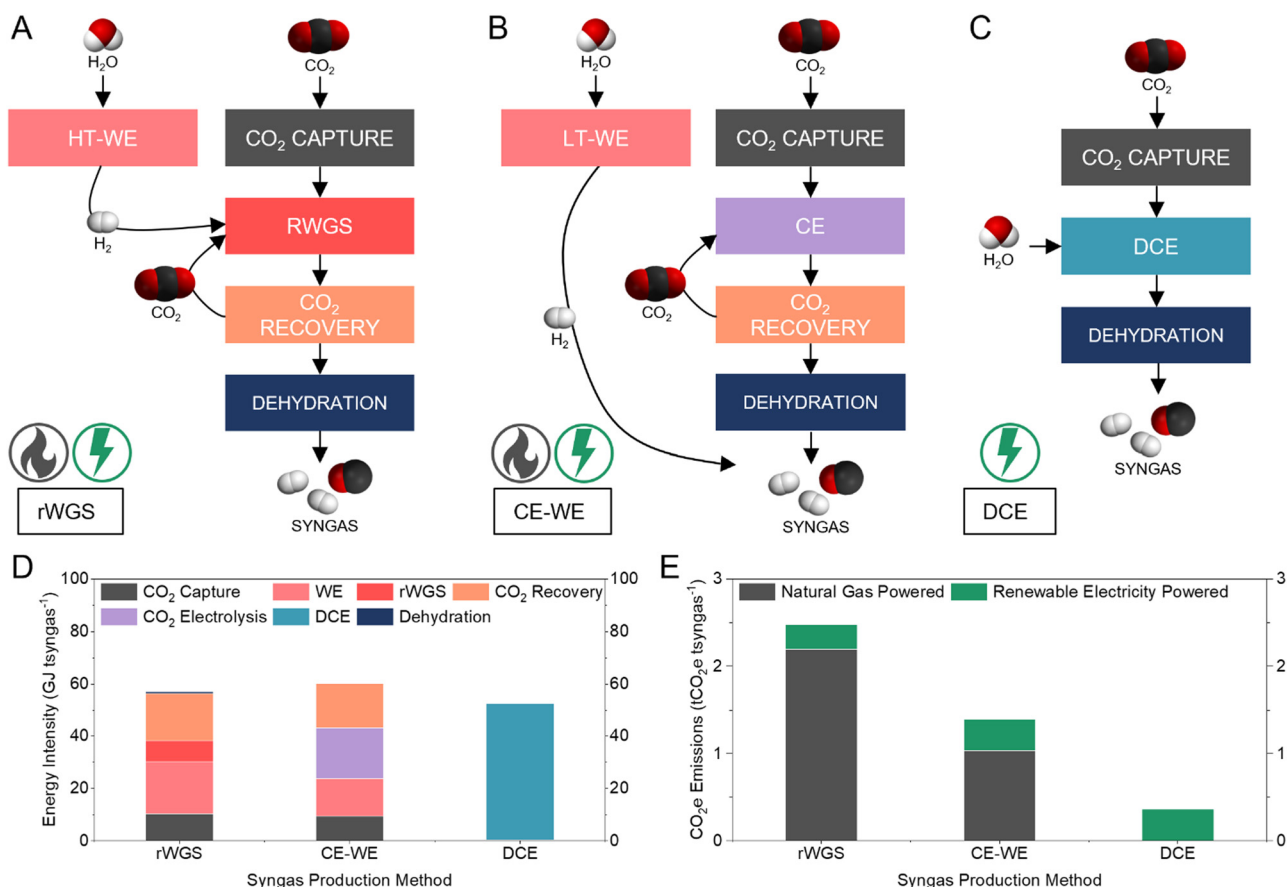


Fig. 5 Comparison of three syngas production methods. (A–C) Schematic showing process pathways, major chemical inputs and outputs, and energy source of rWGS (A), CE-WE (B), and DCE (C). (D) Energy intensity comparison to produce one tonne syngas using rWGS, CE-WE, and DCE. Syngas dehydration energies for CE-WE and DCE are too small to be seen in this figure. Detailed breakdown available in the ESI† (Table S3). (E)  $CO_2e$  emissions to produce one tonne syngas using rWGS, CE-WE, and DCE. The  $CO_2e$  associated with the energy input in each process is considered. Detailed breakdown available in the ESI† (Table S4).



recycling of regenerated alkaline capture fluid (Fig. S10A, ESI<sup>†</sup>). At a constant current density of 100 mA cm<sup>-2</sup>, the full cell voltage, CO FE, H<sub>2</sub>/CO ratio, and capture solution pH were stable for over 23 hours of operation (Fig. 4). The pH of the anode electrolyte remained constant, (Fig. S10C, ESI<sup>†</sup>) and negligible CO<sub>2</sub> was detected in the cathode and anode gas streams (<400 ppm), yielding a pure syngas (99.91 vol%, dry basis). Scanning electron microscopy (SEM) before and after prolonged operation showed minimal change in the CDL structure (Fig. S11, ESI<sup>†</sup>).

### Comparison to alternative syngas production methods

We compared the energy intensity of syngas production *via* three CO<sub>2</sub> electrolysis pathways: thermocatalytic reverse water gas shift (rWGS) (Fig. 5A); low-temperature CO<sub>2</sub> electrolysis combined with water electrolysis (CE-WE) (Fig. 5B); and DCE (Fig. 5C) (Supplementary note 4, ESI<sup>†</sup>). In all cases, the feedstock CO<sub>2</sub> is captured from the atmosphere using an alkaline capture liquid.<sup>13</sup> The rWGS pathway produces syngas with H<sub>2</sub>/CO ratio of 1 and requires 57 GJ tsyngas<sup>-1</sup> (Fig. 5D). The CE-WE and DCE pathways produce syngas with H<sub>2</sub>/CO of 1.16, requiring 60 and 52 GJ tsyngas<sup>-1</sup>, respectively. Syngas production *via* DCE circumvents intensive upstream and downstream processes and thereby results in a 13% energy saving compared to the CE-WE pathway and 8% energy savings compared to the rWGS method. The capture and dehydration steps in the DCE required only 0.85% of the energy demand for the DCE pathway, because the most energy intensive step of the capture process (regeneration) is avoided, and the output syngas stream is pure.

Of the three syngas production methods, DCE is the only pathway that is fully electrically driven, whereas CE-WE requires thermal energy input during CO<sub>2</sub> capture and rWGS requires thermal energy in two major processes. Comparing the operational CO<sub>2</sub>e emissions of the three syngas production pathways, DCE was the only method that offered a low CO<sub>2</sub> intensity (0.36 tCO<sub>2</sub>e tsyngas<sup>-1</sup>), while both CE-WE and rWGS exhibited high net CO<sub>2</sub> emissions (1.39 and 2.48 tCO<sub>2</sub>e tsyngas<sup>-1</sup>, respectively) even when using renewable electricity (Fig. 5E). Compared to the fossil-based method (1.5 tCO<sub>2</sub>e tsyngas<sup>-1</sup>), DCE offers a 75% reduction in CO<sub>2</sub>e emissions.

## Conclusions

We developed a strategy for the efficient electroproduction of syngas with sufficient CO-content and purity for direct industrial application. Through a one-dimensional multi-physics model, we found that the selectivity towards CO can be improved by separating the acidic CO<sub>2</sub> regeneration region from the alkaline CO<sub>2</sub> electrolysis region. We engineered a composite CDL positioned between the cathode and CEM of a DCE system to modulate the local pH and improve local CO<sub>2</sub> conversion. The CDL was comprised of TiO<sub>2</sub> nanoparticles bound by hydrophilic ionomer and was conformally coated onto the Ag catalyst. By determining the optimal CDL thickness that balanced the cathode alkalinity with CO<sub>2</sub> concentration, a H<sub>2</sub>/CO ratio of 1.16

(CO FE 46%) was achieved with an energy intensity of 52 GJ tsyngas<sup>-1</sup> and a CO<sub>2</sub> intensity of 0.36 tCO<sub>2</sub>e tsyngas<sup>-1</sup>. Syngas production *via* DCE can provide a 13% energy saving compared to conventional CO<sub>2</sub> electrolysis methods, and a 75% CO<sub>2</sub>e emissions reduction compared to fossil-fuel methods. These savings suggest that DCE is a promising pathway toward energy efficient syngas – a foundational feedstock for renewable chemicals and fuels in a net-zero emissions future.

## Experimental

### Reagents

Potassium hydroxide (KOH) (>85%) and sulfuric acid (H<sub>2</sub>SO<sub>4</sub>) were purchased from Bioshop. All reagents were of analytical grade and used without further purification. All solutions were prepared using Milli-Q grade water (18.2 MΩ).

### Electrode preparation

The cathode was fabricated by air-brushing an Ag nanoparticle ink onto commercially available hydrophilic carbon paper (AvCarb MGL190, Fuel Cell Store) on a hot plate at 75 °C to achieve a loading of approximately 3 mg cm<sup>-2</sup>. The loading was measured by weighing the carbon paper before and after air-brushing. The catalyst ink was prepared with 150 mg of Ag nanoparticle (99.99%, 20 nm, metal basis, US Research Nanomaterials), 150 mg of Nafion dispersion (5 wt%, Fuel Cell Store), and 6 mL of methanol for a 25 cm<sup>2</sup> substrate and sonicated for 1 hour prior to air-brushing. A commercially available titanium-based anode was used (Magneto Special Anodes, Evoqua Water Technologies).

### CDL preparation

Four sizes of TiO<sub>2</sub> nanoparticle were used: 5 nm (anatase, 99.5%, US Research Nanomaterials), 25 nm (Aeroxide TiO<sub>2</sub> P25, Evonik), 100 nm (anatase, 99.9%, US Research Nanomaterials), and 1500 nm (anatase, 99.9%, US Research Nanomaterials). Two ionomers were used: Nafion (Hydrophobic) and Aemion (Hydrophilic). Nafion ionomer can create strongly hydrophobic nanoporous networks due to its polytetrafluoroethylene (PTFE) (-CF<sub>2</sub>-CF<sub>2</sub>-) back-bone.<sup>50,51</sup> Aemion ionomer contains poly-(benzimidazole) units which are more hydrophilic than the PTFE in Nafion.<sup>52</sup> The Aemion dispersion was prepared by adding 388 mg of ionomer powder (AP1-CNN5-00-X, Ionomer) to 40 mL of ethanol and 10 mL of acetone and sonicated until fully dissolved.

The CDL was fabricated by air-brushing a TiO<sub>2</sub> nanoparticle ink onto the fabricated cathode to achieve the desired thickness. For the optimized CDL, the ink was prepared with 50 mg of 25 nm TiO<sub>2</sub>, 333 mg of the prepared Aemion dispersion, and 4 mL of ethanol for a 6.25 cm<sup>2</sup> cathode and was sonicated for 1 hour prior to air-brushing. The CDL coated cathode was cut to a 1 cm<sup>2</sup> size prior to electrolyzer assembly. CDLs were characterized using SEM at the Centre for Nanostructure Imaging at the University of Toronto using an FEI Quanta FEG 250 environmental SEM.





Two hydrophilic microporous membrane filters were used: 125  $\mu\text{m}$  PVDF (Filter 1, 0.45  $\mu\text{m}$  pore size) was purchased from Sigma Aldrich and 100  $\mu\text{m}$  nylon (Filter 2, 5  $\mu\text{m}$  pore size) was purchased from Sterlitech. Both filters were used as received.

### Operation of the electrochemical cell

The carbonate electrolysis experiments were performed in a 1  $\text{cm}^2$  electrolyzer with serpentine flow channels ingrained in both the stainless-steel cathode and the titanium anode current collectors. The electrolyzer was assembled by placing a CEM (Nafion 117) over the cathode, then placing the anode on the membrane. In all experiments, unless otherwise specified, the cathode feedstock was a carbonate electrolyte prepared by purging  $\text{CO}_2$  at 80 sccm into 85 mL of 2 M KOH for 40 minutes, similar to a previous report.<sup>29</sup> The anode was fed with 0.05 M  $\text{H}_2\text{SO}_4$ . After starting the experiment, the carbonate electrolyte was continuously purged with Ar gas flowing at 20  $\text{mL min}^{-1}$ . The first gas sample is typically collected 20 minutes after the start of the experiment to ensure complete purging of excess  $\text{CO}_2$  from the electrolyte preparation process and even mixing of gaseous products. The electrochemical measurements were performed with a potentiostat (Autolab PGSTAT204) and the full cell voltages reported are not iR compensated unless otherwise specified.

### Product analysis

The cathode gas outlet stream was analyzed in 1 mL sample volumes by a gas chromatograph (PerkinElmer Clarus 590) coupled with a thermos conductivity detector (TCD) and flame ionization detector (FID). The gas chromatograph used Ar gas as the carrier (99.999%, Linde) and was equipped with a Molecular Sieve 5A Capillary Column and a packed Carboxen-1000 Column. Liquid product detection was performed using proton nuclear magnetic resonance spectroscopy ( $^1\text{H}$  NMR) on an Agilent DD2 600 spectrometer in  $\text{D}_2\text{O}$  using water suppression mode, with dimethyl sulfoxide (DMSO) as the internal standard. Calculation of FE and energy efficiency are included in Supplementary note 1 in the ESI.† pH measurements were conducted using Apera Instruments AI311 Premium Series PH60.

## Author contributions

D. S. and E. H. S. supervised the project. Y. C. X. and C. M. G. conceived the study. Y. C. X. designed and performed all the electrochemical experiments and generated the figures. G. L., C. P. O., R. K. M., and M. F. assisted with electrochemical experiments. C. M. G., G. L., C. P. O., and J. Z. assisted with CDL design. T. A. assisted with pH investigation experiment design. Y. C. X. and Y. X. performed gaseous product analysis. J. L., P. P., and J. E. H. performed liquid product analysis. J. P. E. and R. K. M. assisted with energy analysis. Y. Z. performed SEM imaging. S. L. performed system modelling. A. S. R. assisted with system modelling. Y. C. X. and D. S. wrote the manuscript. All authors discussed the results and assisted during manuscript preparation.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

The authors would like to thank Dr Paul J. Corbett, Dr Ben Rowley, and Dr Kai Han for helpful discussions. The authors acknowledge support from the Natural Sciences and Engineering Research Council (NSERC) of Canada and Natural Resources Canada – Clean Growth Program. Support from Canada Research Chairs Program is gratefully acknowledged, as is support from an NSERC E. W. R. Steacie Fellowship to D. S. Y. C. X. thanks Hatch for their support through graduate scholarships.

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