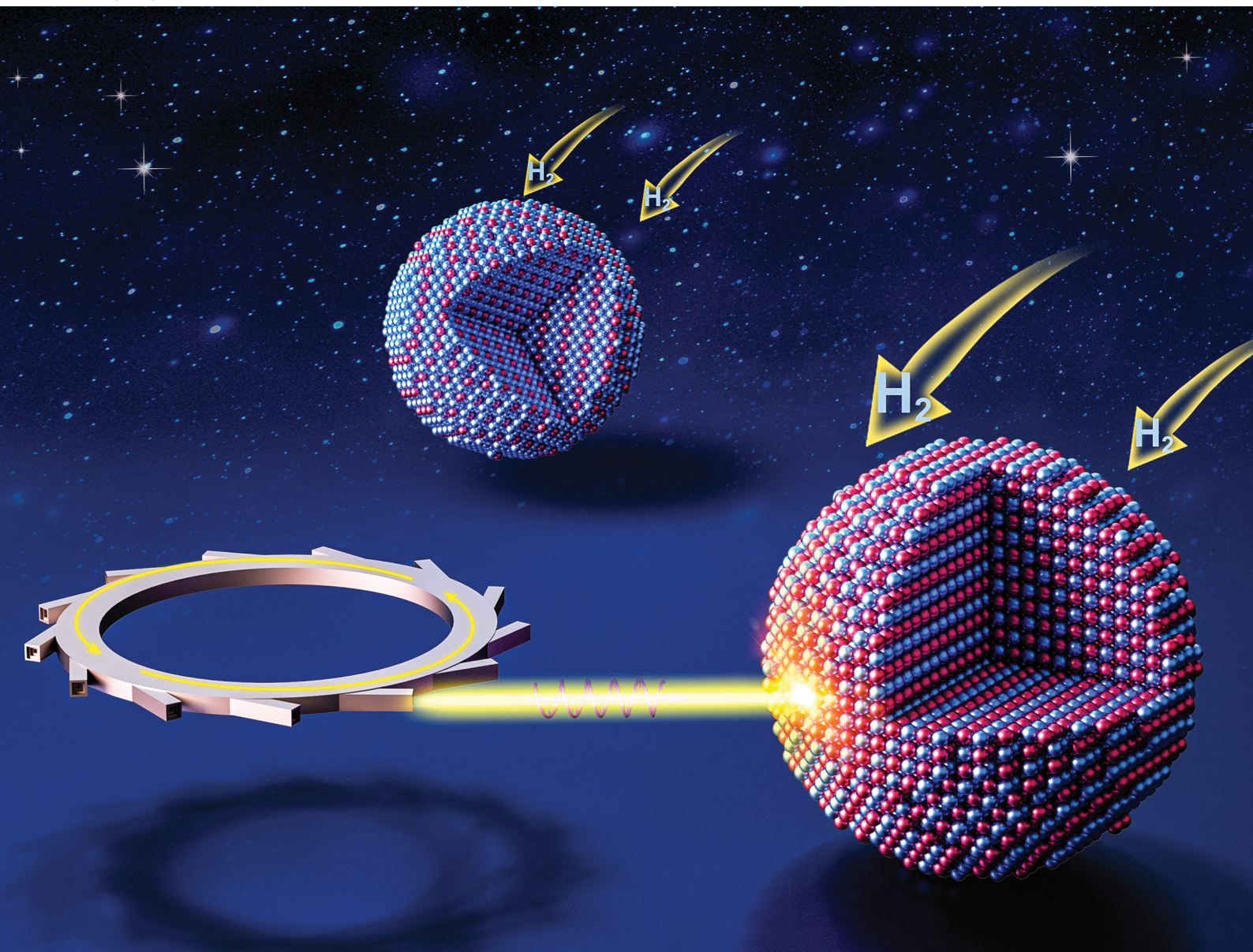


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Synthesis of Pt₃Zn₁ and Pt₁Zn₁ intermetallic nanocatalysts for dehydrogenation of ethane†

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Pt₃Zn₁ and Pt₁Zn₁ intermetallic nanoparticles supported on SiO₂ were synthesized by combining atomic layer deposition (ALD) of ZnO, incipient wetness impregnation (IWI) of Pt, and appropriate hydrogen reduction. The formation of Pt₁Zn₁ and Pt₃Zn₁ intermetallic nanoparticles was observed by both X-ray diffraction (XRD) and synchrotron X-ray absorption spectroscopy (XAS). STEM images showed that the 2–3 nm Pt-based intermetallic nanoparticles were uniformly dispersed on a SiO₂ support. The relationships between Pt–Zn intermetallic phases and synthesis conditions were established. *In situ* XAS measurements at Pt L₃ and Zn K edges during hydrogen reduction provided a detailed image of surface species evolution. Owing to a combined electronic and geometric effect, Pt₁Zn₁ exhibited much higher reactivity and stability than Pt₃Zn₁ and Pt in both the direct dehydrogenation and oxidative dehydrogenation of ethane to ethylene reactions.

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1. Introduction

Single-atom alloys (SAA) are a form of single-atom catalysts that consist of atomically dispersed active sites in a host metal.¹ When the metal active sites and the host metal form a random alloy (*i.e.*, solid solution), atomic dispersion can be achieved by maintaining a low density of active sites.² High-density, mono-dispersed active sites can be achieved if the binary system forms an intermetallic phase.³ Intermetallic alloys are a type of bimetallic compound that forms a finite stoichiometry and ordered atomic structure which warrants uniform catalytic active sites. Although intermetallic catalysts exhibit greater active site density than the solid solution, the thermodynamically favored intermetallic phases could theoretically resist segregation under heat and chemical adsorption.⁴ These features make intermetallic catalysts attractive for applications in well-defined catalysis systems. A Pt–Sn intermetallic alloy is the most studied alloy system in the dehydrogenation of alkanes and has been commercialized.⁵ Adding a second metal to Pt can reduce the particle size thus effectively suppressing side reactions, for example, hydrogenolysis

and coking.⁶ Various new Pt-based intermetallic nanocatalysts, such as Pt–Zn, Pt–In, and Pt–V, have demonstrated enhanced activity and selectivity in alkane dehydrogenation.⁷ However, the intrinsic contributions of these add-metals to Pt are still in question, due to the difficulties in the preparation of Pt-based intermetallic nanoparticles with a targeted crystal structure.⁸

Pt–Zn intermetallic nanoparticles have exhibited excellent catalytic performance in electrooxidation,⁹ hydrogenation,¹⁰ and dehydrogenation.¹¹ Catalytic dehydrogenation of light alkanes, such as ethane and propane, is an on-purpose production method that exclusively yields the desired alkene of polymer-quality purity rather than a mixture of products.¹² Cybulskis and coworkers have shown that the intermetallic Pt₁Zn₁ nanoparticles are more active, selective, and stable than the un-promoted Pt nanoparticles in the dehydrogenation of ethane.^{7a} They ascribed the improved performance to both electronic and geometric effects. Nevertheless, how exactly these two factors affect the ethane dehydrogenation activity, selectivity, and stability were not well established with the limited examples of Pt₁Zn₁ and metallic Pt. Also, as their work and many recent studies on alkane dehydrogenation using PtZn nanocatalysts^{11,13} focused on more characterization on the Pt side, the information on the changes of Zn species was lacking, this might overlook some important aspects for the successful synthesis of active phases. The Pt₁Zn₁ catalyst they prepared utilized incipient wetness impregnation, a relatively rapid method widely used in catalyst synthesis. However, this method is known to produce nanoparticles with broad size distribution. A controlled synthesis method may offer a well-defined catalyst system to establish precise structure and performance relationships.

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† Electronic supplementary information (ESI) available: XRD patterns, EDX spectrum, EXAFS fitting, and N₂ absorption–desorption isotherms. See DOI: <https://doi.org/10.1039/d2cp04173a>

Chemical methods used to synthesize intermetallic alloy nanoparticles include impregnation, colloidal synthesis, and chemical vapor deposition. Atomic layer deposition (ALD) has emerged as an important technique for thin film deposition.¹⁴ It produces conformal coatings of materials *via* self-limiting chemical reactions between the precursors. By controlling ALD parameters such as dose time, the number of cycles, and deposition temperature, we can quantitatively design catalysts with the desired thickness, loading, and microstructures. ZnO ALD, as one of the well-developed ALD processes,¹⁵ can potentially aid in the design of mono-dispersed nanoscale Pt–Zn intermetallic alloys, such as Pt₃Zn₁, Pt₁Zn₁, Pt₃Zn₁₀, Pt₁Zn₈, *etc.*,¹⁶ which will provide more information regarding Pt site isolation and electronic states with the addition of Zn.

Herein, we synthesized silica-supported Pt₃Zn₁ and Pt₁Zn₁ intermetallic nanocatalysts by sequential deposition of ZnO *via* atomic layer deposition (ALD) and Pt *via* incipient wetness impregnation (IWI) on SiO₂ with H₂ reduction. The synthesis method was examined using standard characterization techniques as well as synchrotron X-ray absorption spectroscopy (XAS) under hydrogen reduction conditions. We revealed that ZnO reduction is a key step in the synthesis of the Pt₁Zn₁ phase by providing a detailed picture of Pt and Zn speciation kinetics under H₂ reduction *via in situ* XAS. Silica-supported Pt₃Zn₁, Pt₁Zn₁, and Pt monometallic nanocatalysts were tested in both dehydrogenation of ethane (DHE) and oxidative dehydrogenation of ethane (ODHE), and the electronic and geometric effects on the catalytic performance of PtZn intermetallic nanocatalysts are discussed thoroughly. The Pt₁Zn₁ intermetallic nanocatalysts exhibited improved selectivity and a production rate one order of magnitude greater than those of supported Pt and Pt₃Zn₁ with similar Pt loading and particle sizes for the dehydrogenation of ethane. Enhanced catalyst stability was also observed.

2. Experimental section

2.1 Synthesis

Synthesis of ZnO/SiO₂. SiO₂ (Silicycle S10040M, 100 m² g⁻¹) was used as the catalyst support. ZnO ALD was carried out in a viscous flow benchtop reactor (Gemstar-6, Arradience) at deposition temperatures of 150 °C and 200 °C. One gram of SiO₂ was spread out smoothly onto a metal tray and covered by a stainless steel wire cloth to prevent powder from spilling out while maintaining sufficient diffusion of the reactants. ZnO ALD was performed using alternating exposure to diethyl zinc (DEZ, Sigma Aldrich) and deionized water. Both DEZ and H₂O were kept in sealed stainless steel bottles at room temperature. One ALD cycle time sequence began with DEZ exposure time (*t*₁) followed by a N₂ purge time (*t*₂), then H₂O exposure time (*t*₃) followed by a second N₂ purge time (*t*₄).

Synthesis of Pt/SiO₂. Pt was deposited on SiO₂ (pore volume of 0.9 mL g⁻¹) by pH-modified incipient wetness impregnation (IWI). First, 0.4 g of (NH₃)₄Pt(NO₃)₂ (Alfa Aesar) was dissolved in 3.42 mL of deionized water. NH₄OH (Fisher Scientific) was

subsequently added to the solution to adjust the pH to 11. The solution was then added dropwise to 3.8 g of SiO₂ with constant stirring. The resulting powder was dried at 125 °C in an oven overnight, and then calcined at 300 °C in a tubular furnace for 3 h.

Synthesis of Pt/ZnO/SiO₂. Pt was deposited on ZnO/SiO₂ by impregnation of (NH₃)₄Pt(NO₃)₂ solution (pH adjusted to 11). The resulting powder was calcined at 300 °C in a tubular furnace for 3 h.

Synthesis of intermetallic PtZn/SiO₂. The pre-catalyst Pt/ZnO/SiO₂ was reduced in 5% H₂ with balance He (nexAir) at various temperatures between 400 °C and 800 °C to achieve different phases of PtZn intermetallic alloy nanoparticles. The H₂ flow was further treated using a Supelco gas purifier (Sigma Aldrich). In a typical experiment, the sample was heated from room temperature to the targeted temperature using a temperature ramping rate of 2.5 °C min⁻¹; the target temperature was held for a certain period of time before the sample was cooled to room temperature. The Pt₁Zn₁ and Pt₃Zn₁ phases were achieved by hydrogen reduction at 600 °C for 15 min and at 800 °C for 8 h, respectively. The actual loadings of Pt and Zn were determined by ICP-OES.

2.2 Characterization

The thickness of the ZnO film deposited onto the Si(100) wafer was determined by spectroscopic ellipsometry (SE) using a J. A. Woollam alpha-SE variable angle spectroscopic ellipsometer. X-ray Diffraction (XRD) was performed on a Rigaku MiniFlex 600 powder X-ray diffractometer operated at 40 kV and 15 mA. XRD patterns were scanned from 2θ = 10° to 90° with a scanning speed of 1° min⁻¹. Brunauer–Emmett–Teller (BET) surface area measurements were performed using a gas sorption analyzer (Autosorb iQ, Quantachrome Instruments) at 77 K. In a typical experiment, the sample was outgassed at 350 °C for 10 h. The specific surface area was calculated using the Brunauer–Emmett–Teller (BET) method and the pore volume and size were obtained using the Barrett–Joyner–Halenda (BJH) method. Pt dispersions based on H₂ chemisorption were also estimated using the Autosorb iQ gas sorption analyzer. All samples were heated *in situ* to 600 °C, held for 2 h under H₂ (Airgas, 99.999%), and evacuated for 2 h before cooling to 40 °C for chemisorption. The STEM images were taken at the Center for Nanoscale Materials within the Argonne National Laboratory using a FEI Talos scanning transmission electron microscope. A TEM grid was first mixed with the sample inside a vial by shaking rigorously. Images were then taken using a high-angle annular dark-field (HAADF) detector at 300 kV. The energy-dispersive X-ray (EDX) spectra of the catalysts were taken using a JEOL 2100 TEM equipped with a Gatan camera at an accelerating voltage of 200 kV. Inductively coupled plasma optical emission spectroscopy (ICP-OES) was carried out using an Agilent 5110 SVDV ICP-OES spectrometer.

X-ray absorption spectroscopy (XAS) measurements were conducted at 12-BM of the Advanced Photon Source (APS) at the Argonne National Laboratory. Transmission mode spectra were taken at the Pt L₃ edge (~11 564 eV) and the Zn K edge

(~ 9659 eV). For the *ex situ* measurements, Pt/SiO₂, Pt₁Zn₁/SiO₂, and Pt₃Zn₁/SiO₂ samples were first reduced in 3.5% H₂ with balance He at 600 °C for 30 min to remove any surface oxides which formed due to sample transfer. The samples were then cooled and measured in helium at RT. For the *in situ* XAS study, the Pt/ZnO/SiO₂ precatalyst was reduced by 3.5% H₂/He at the designated temperature for 15 min and then measured at RT under helium protection. The XANES linear combination fittings were processed from Athena in the normalized $\mu(E)$ space, with Pt metal foil and PtO₂ used as references. The fittings were done in the range of 10 eV below and 40 eV above the Pt L₃ edge. EXAFS regime data fittings were performed using Artemis in the IFEFFIT software package.

2.3 Catalytic performance

Direct dehydrogenation of ethane (DHE) and oxidative dehydrogenation of ethane (ODHE) were carried out in a quartz tube reactor (10 mm I.D.) at 550 °C under atmospheric pressure. The quartz tube was placed in an electric furnace and sealed with Ultra Torr fittings. A K-type thermocouple was used to measure the temperature of the catalyst bed. The catalyst particles sieved between 180 to 250 μm were mixed with 500 mg 60/80 mesh quartz chips. For DHE, the catalyst weight is fixed at 100 mg to achieve a weight hourly space velocity (WHSV) of 7.2 h⁻¹. The reactant feed used in DHE was composed of 18% C₂H₆ (Matheson, 99.99%) and 82% He (nexAir, 99.999%), for a total flow of 50 mL min⁻¹. For ODHE, the reactant feed contained 22.3% ethane and 1.7% oxygen with balance helium. Due to the rapid deactivation in ODHE, various amounts of catalysts were used to achieve $\sim 10\%$ initial conversion. The amount used for Pt, Pt₁Zn₁, and Pt₃Zn₁ were 250 mg, 10 mg, and 90 mg, respectively. The effluent line was heated to temperatures above 80 °C to prevent condensation of byproducts. Before each reaction, the catalyst was reduced by 5% H₂/He at 600 °C for 30 min. The effluents were analyzed on a gas chromatograph

(SRI, 310C) equipped with a ShinCarbon ST column and a thermal conductivity detector (TCD).

3. Results and discussion

Uniform ZnO thin films formed on SiO₂ support *via* atomic layer deposition acted as the source of Zn for the preparation of nano-sized Pt–Zn intermetallic alloys. As shown in Fig. 1(a), the linear growth behavior of ZnO thin films yielded a growth per cycle (GPC) of ~ 1.8 Å per cycle. This growth rate is consistent with the 1.5–1.9 Å per cycle range within the 50–300 °C temperature range found in the literature.¹⁷ The saturation exposure of DEZ for deposition of ZnO on high surface area SiO₂ powder was determined by varying the DEZ exposure time from 0.5 to 10 s, as shown in Fig. 1(b). The saturation weight gain was ~ 7 wt% when the dose time of DEZ reached 5 s (~ 6.7 wt% at 150 °C, ~ 7.0 wt% at 200 °C). The formation of metallic Zn within the thin films was observed at 200 °C (Fig. 1(c)). This incorporation of metallic Zn was likely due to the thermal decomposition of DEZ on the substrate,¹⁸ but the saturation weight gain suggests that the nature of the ZnO growth was within the ALD regime.¹⁹ Based on the surface area of SiO₂ and the density of ZnO, the theoretical saturated weight gain is calculated to be ~ 9.7 wt%. The weight gain is lower than expected, which indicates that ZnO ALD precursors cannot access the entire surface of SiO₂ due to steric hindrance.¹⁵ The time sequence of 5–200–37.5–200 s and deposition temperature of 200 °C were used to prepare ZnO/SiO₂. Pt nanoparticles were subsequently deposited on the ZnO-covered SiO₂ surface *via* incipient wetness impregnation (IWI).

The Pt–Zn intermetallic phases were generated by reducing Pt/ZnO/SiO₂ in hydrogen. The optimal reduction temperature was determined by varying the temperature between 600 °C and 800 °C, while the optimal reduction time was determined by varying the length of time at which Pt/ZnO/SiO₂ was reduced,

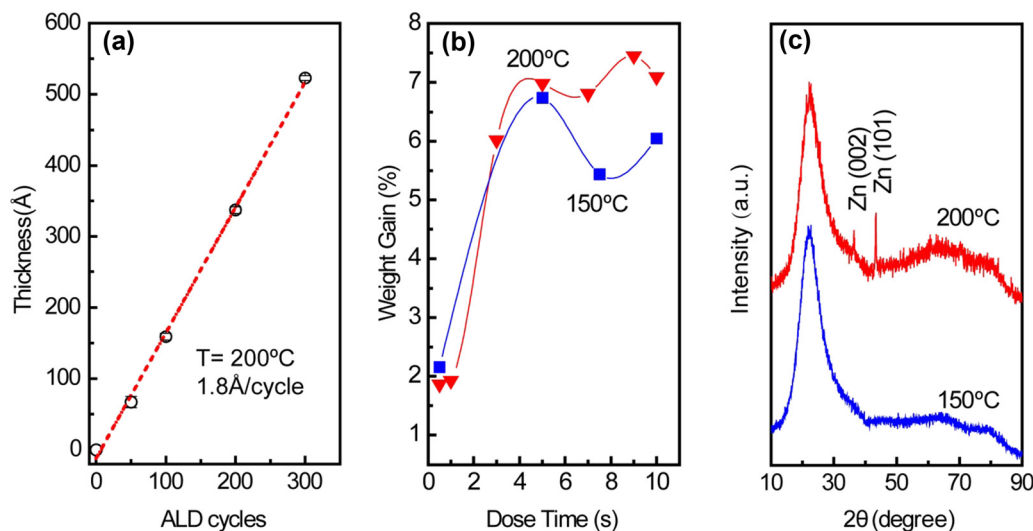


Fig. 1 (a) Growth rate of ZnO ALD at 200 °C with time sequence 0.03–10–0.1–10 s, (b) weight gain as a function of DEZ dose time with time sequence t_1 –200–40–200 s, (c) XRD of the ALD ZnO/SiO₂ sample prepared at 150 °C and 200 °C.

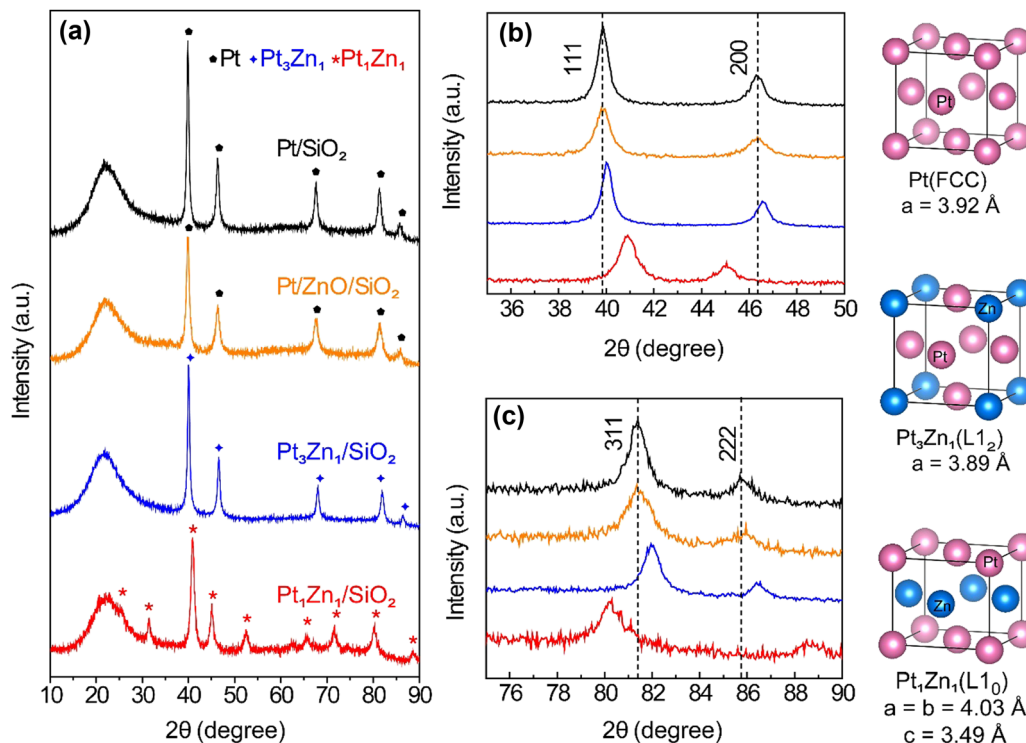


Fig. 2 (a) XRD patterns of Pt/SiO₂, pre-catalyst Pt/ZnO/SiO₂, Pt₃Zn₁/SiO₂ and Pt₁Zn₁/SiO₂. Detailed scans of XRD patterns of the four samples between (b) 35–50° and (c) 75–90°.

from 15 min to 10 h, as shown in Fig. S1 (ESI[†]). The best preparation conditions to generate Pt₁Zn₁ and Pt₃Zn₁ phases were determined to be 6 h reduction at 600 °C and 8 h reduction at 800 °C, respectively. The stoichiometry of the as-prepared Pt₁Zn₁ and Pt₃Zn₁ intermetallics were confirmed using EDX (Fig. S2, ESI[†]) and ICP-OES. Fig. 2(a) shows that the shape of the XRD pattern of Pt₃Zn₁ (JCPDS no. 65-3257) resembles that of the monometallic Pt phase (JCPDS no. 65-2868) but the peaks are positioned at higher diffraction angles, as shown in the detailed scans in Fig. 2(b) and (c). The shifts in peak location in the Pt₃Zn₁ XRD pattern are evidence that Zn has been incorporated into Pt's face-centered cubic (FCC) structure since Zn has a smaller atomic radius.²⁰ Based on the XRD patterns, the lattice parameters are calculated to be 3.92 Å for FCC Pt and 3.89 Å for the L₁₂ Pt₃Zn₁, respectively, and 4.03 Å and 3.49 Å for the *a* and *c* unit vectors, respectively, of the AuCu L₁₀ tetragonal structured Pt₁Zn₁ (JCPDS no. 06-0604). The Pt–Pt bond distance for Pt/SiO₂, Pt₃Zn₁/SiO₂, and Pt₁Zn₁/SiO₂ nanoparticles are calculated to be 2.77 Å, 2.75 Å, and 2.85 Å (Table S1, ESI[†]), respectively, which are similar to literature values.^{7a,9c,21}

The formation of Pt–Zn intermetallic alloys was further investigated using XAS measurements at Pt L₃ (11.564 keV) and Zn K (9.659 keV) edges. Prior to carrying out XAS measurements, the Pt-containing samples were reduced by 3.5% hydrogen at 600 °C for 30 min to remove any surface oxides that formed due to sample transfer. As compared with the Pt L₃ XANES spectrum of Pt/SiO₂, the whiteline intensities (2p → 5d transition) of Pt₃Zn₁/SiO₂ and Pt₁Zn₁/SiO₂ remain nearly the

same but their edge positions shift to higher photon energies by 1.5 and 2.4 eV, respectively. These positive shifts indicate changes in hybridization and lower 5d electronic density (Fig. 3(a)).^{7a,21,22} At the Zn K-edge, whitelines of Pt₃Zn₁/SiO₂ and Pt₁Zn₁/SiO₂ do not show obvious shifts as compared with metallic Zn, but their intensities are higher (Fig. 3(c)). This feature, corresponding to the 1s → 4p transition, is sensitive to the change of coordination of Zn and bond distance due to the formation of the Pt–Zn intermetallic phase.²³

The Pt L₃ edge EXAFS spectrum of Pt/ZnO/SiO₂ in Fig. 3(b) shows a Pt–O peak at ~1.7 Å and three peaks between 2 Å and 3 Å, indicating the Pt was partially oxidized. The EXAFS spectrum of Pt₁Zn₁ shows a Pt–Zn single peak at ~2.4 Å (uncorrected) while the Pt₃Zn₁ spectrum contains a cluster of three peaks between 2 Å and 3 Å resembling that of Pt owing to their similar crystalline structure. Data fitting in Table S2 and Fig. S3 (ESI[†]) shows that each Pt atom on Pt/SiO₂ has an average of 10.5 nearest neighbors at a bond distance of 2.76 Å. With the incorporation of Zn, the number of nearest Pt neighbors surrounding one Pt atom in Pt₃Zn₁/SiO₂ decreases to 6.1 and the Pt–Pt bond distance is shortened to 2.73 Å. In Pt₁Zn₁ nanoparticles, the Pt–Pt bond distance is elongated to 2.85 Å and each Pt atom has an average of 3.1 nearest Pt neighbors, indicating that these Pt atoms were structurally more isolated. The bond distances determined by EXAFS data fittings are comparable to those obtained from XRD patterns. In the Zn K edge EXAFS spectra (Fig. 3(d)) and data fitting (Table S3 and Fig. S4, ESI[†]), ZnO and Zn have single peaks at ~1.5 Å (Zn–O) and 2.3 Å (Zn–Zn), respectively. In the bulk structure of metallic

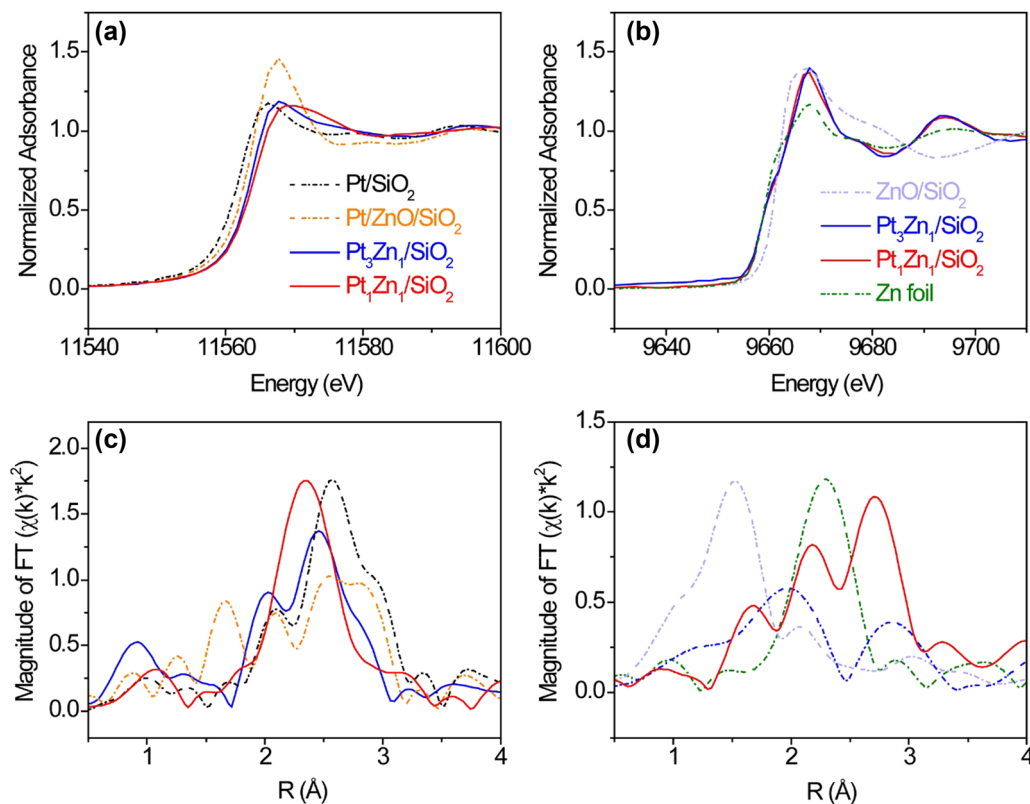


Fig. 3 Pt L₃ edge (a) XANES and (b) EXAFS Fourier transform of Pt/SiO₂, Pt/ZnO/SiO₂, Pt₁Zn₁/SiO₂, Pt₃Zn₁/SiO₂. Zn K edge (c) XANES and (d) EXAFS Fourier transform of Zn, ZnO/SiO₂, Pt₁Zn₁/SiO₂, and Pt₃Zn₁/SiO₂.

Zn, each Zn atom has 6 Zn neighbors, with a bond distance of 2.64 Å between atoms. The Zn–Pt bond distances in Pt₃Zn₁ and Pt₁Zn₁ are 2.73 Å and 2.65 Å, respectively, consistent with the bond distances derived from Pt L₃ EXAFS data fitting (Table S2, ESI[†]) and XRD patterns (Table S1, ESI[†]).

Fig. 4 shows the STEM images and particle size distribution histograms of Pt/SiO₂, Pt₃Zn₁/SiO₂, and Pt₁Zn₁/SiO₂. The nanoparticles were well dispersed on the SiO₂ support. The particle size distribution, calculated for more than 500 nanoparticles, demonstrates that the average diameter of Pt/SiO₂, Pt₁Zn₁/SiO₂, and Pt₃Zn₁/SiO₂ nanoparticles are all about 2.2 nm. In the EDX mapping, all Pt and Zn co-locate at the same spots, suggesting a homogeneous formation of Pt–Zn alloys.

In situ XAS spectra at Pt L₃ and Zn K edge were measured during hydrogen reduction to determine the surface species evolution. The measurements were obtained by quickly switching between Pt L₃ and Zn K edges (<1 min), and thus the data on both edges were taken almost simultaneously. Fig. 5 shows the XANES and EXAFS spectra at Pt L₃ and Zn K edges during reduction of Pt/ZnO/SiO₂ from 25 °C to 800 °C. The *in situ* XAS spectra clearly show the transition from the initial oxidation state to the final intermetallic phase. In Fig. 5(a) and (b), the Pt L₃ edge whiteline gradually shifts to a higher energy, accompanied by the shift of Zn K edge to a lower energy. In the EXAFS spectra in Fig. 5(c) and (d), Pt–Zn bonds arise with the elimination of Pt–O and Zn–O bonds. The quantitative changes to an atom's nearest neighbors and the associated bond distance

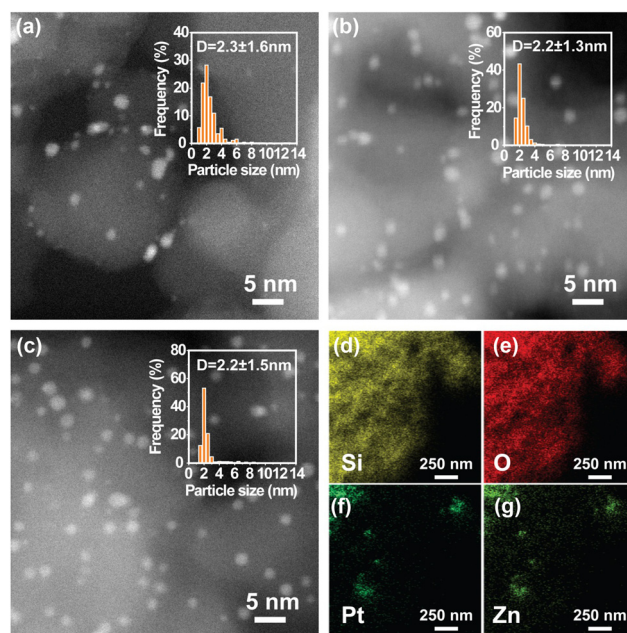


Fig. 4 (a) and (b) HAADF-STEM images and particle size distribution histograms of (a) Pt/SiO₂, (b) Pt₃Zn₁/SiO₂, (c) Pt₁Zn₁/SiO₂, and (d)–(g) EDX mapping of Pt₁Zn₁/SiO₂.

with respect to Pt and Zn are summarized in Tables S4 and S5 (ESI[†]), and the fitting quality is shown in Fig. S5–S8 (ESI[†]).

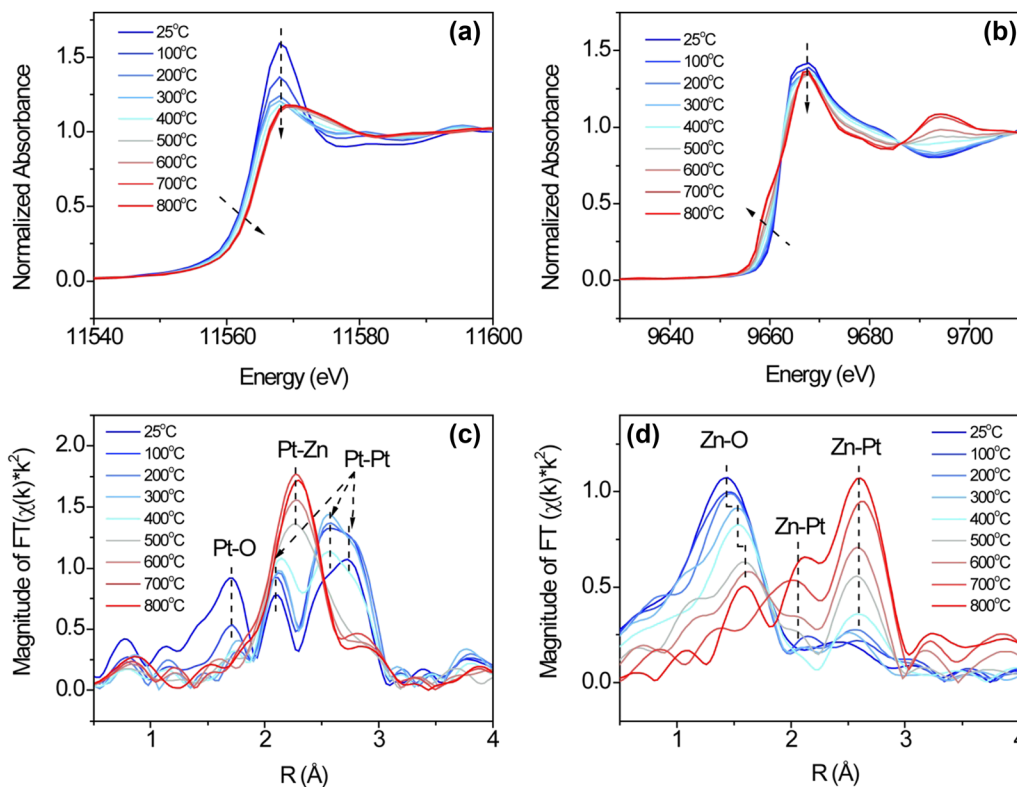


Fig. 5 Pt L₃ edge (a) XANES, (b) EXAFS Fourier transform of Pt/ZnO/SiO₂ during *in situ* hydrogen reduction at elevated temperatures and their Zn K edge XANES (e) and (d) EXAFS counterparts.

XANES linear combination fittings (LCF, Fig. 6) are consistent with the EXAFS fittings summarized in Tables S4 and S5 (ESI[†]). The as-prepared Pt/ZnO/SiO₂ sample was initially partially oxidized, containing 69% metallic Pt and 31% PtO₂, while

all the zinc was in the form of ZnO. PtO₂ was completely reduced to the metallic state at 200 °C, coinciding with the formation of the Pt-Zn phase and conforming to previous studies on hydrogen reduction of Pt/ZnO/Al₂O₃ and Pd/ZnO.²⁴

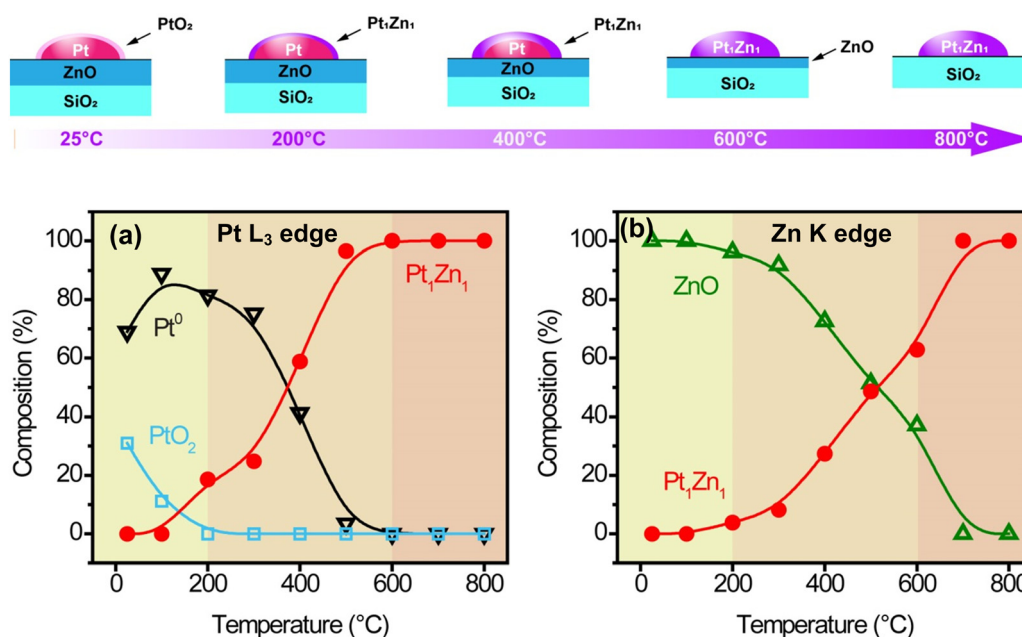


Fig. 6 (a) Scheme of surface species evolution, and XANES linear combination fitting (LCF) of (b) Pt L₃ and (c) Zn K edges.

The hydrogen reduction was likely catalyzed by platinum, which was able to reduce Zn at such a low temperature. Upon increasing temperature from 200 °C to 600 °C, more Pt–Zn phase formed as more ZnO was reduced. The formation of Pt₁Zn₁ seems to be limited by the catalytic reduction rate of ZnO instead of Zn atom diffusion into Pt. At 600 °C, all platinum existed in the form of Pt₁Zn₁ phase while excess ZnO remained. The excess ZnO could be removed by long reduction or vaporization at higher temperature, *i.e.*, 800 °C. Note that it was still in the Pt₁Zn₁ phase as it took 8 hours at 800 °C to form the Pt₃Zn₁ phase (Fig. S1b, ESI†).

There are three main concerns regarding the reduction of the Pt/ZnO/SiO₂ sample to Pt₃Zn₁/SiO₂ at 800 °C in hydrogen. The first concern is the possible collapse of the porous SiO₂ support, which will cause a close encapsulation of platinum and prevent reactants from accessing the catalyst's active sites. The second concern is the occurrence of sintering of platinum nanoparticles due to diffusion-coalescence at high temperatures and a hydrogen environment, again leading to the loss of active sites for catalytic reactions. The last concern is the potential formation of the Pt–Si phase induced by platinum-catalyzed SiO₂ reduction. The latter two concerns were observed in an *in situ* TEM study of Pt/SiO₂ reduction at 800 °C in hydrogen.²⁵ The diameter of Pt nanoparticles increased from 2 nm to 4.8 nm, which is associated with the formation of Pt₃Si after being reduced in hydrogen at 800 °C. In our study, these three concerns were ruled out as evidenced by combining BET surface area analysis, XRD patterns, and EXAFS spectra. As shown in Fig. S9 and Table S6 (ESI†), the surface area, pore diameter, and pore volume are consistent between Pt/ZnO/SiO₂, Pt₁Zn₁/SiO₂, and Pt₃Zn₁/SiO₂, indicating that the structural integrity of SiO₂ support was maintained. Neither any Pt–Si phase was observed in XRD patterns, nor was Pt–Si present in the EXAFS spectra. As the SiO₂ support was uniformly covered by ZnO deposited by ALD, it may have prevented the contact between Pt and SiO₂ and the possible subsequent formation of Pt₃Si.

The dehydrogenation of ethane reaction at 550 °C with a WHSV of 7.2 h⁻¹ was used to evaluate the catalytic performance of the Pt–Zn intermetallic nanoparticles. Catalytic performance results are shown in Fig. 7. Based on the Pt dispersion

(Table S6, ESI†), the initial ethylene turnover frequency (TOF), normalized per surface Pt was 0.56 s⁻¹ for Pt₁Zn₁/SiO₂, much higher than the Pt-rich Pt₃Zn₁/SiO₂ (0.48 s⁻¹) and the monometallic Pt/SiO₂ (0.04 s⁻¹). Although all three catalysts had close to 100% C₂H₄ selectivity, the Pt₁Zn₁ catalyst exhibited significantly improved stability. The Pt₁Zn₁ catalyst maintained at a higher conversion (~24.5%) during the reaction period as compared to the Pt₃Zn₁ whose conversion rapidly decreased from 7% to 1.7%.

As shown in Fig. 8, the Pt₁Zn₁ catalysts also exhibited the best performance in terms of C₂H₄ TOFs, conversion, and selectivity in the ODHE. The time-on-stream catalytic tests demonstrate that Pt/SiO₂ has an initial C₂H₄ TOF of 0.04 s⁻¹ and a C₂H₄ selectivity of 70.2%, and the selectivity decays rather quickly during the first hour, with 39.3% left after 5 h. The Pt-rich Pt₃Zn₁/SiO₂ exhibited better initial C₂H₄ TOF (0.66 s⁻¹) and selectivity (77%) as compared to Pt/SiO₂, but after 2 h the selectivity drops steeply and almost follows the decay path of Pt/SiO₂. Pt₁Zn₁/SiO₂ shows not only improved initial C₂H₄ TOF (2.79 s⁻¹, 70 times higher than Pt/SiO₂) and C₂H₄ selectivity (89.4%), but also more stable. Its selectivity did not show a drastic decrease and remained at 67.2% after 5 h. It is shown in Fig. 8d that the main by-products for ODHE are CO, CO₂, and CH₄, the CO plus CO₂ initial selectivity for Pt/SiO₂ is the highest, revealing that Pt monometallic is favorable for the combustion reaction path.

The improved performance may be caused by the combined electronic and geometry effects due to the presence of Zn in Pt.^{7a,24a} Cybulskis observed that Pt₁Zn₁/SiO₂ was more active and stable than Pt/SiO₂ in dehydrogenation of ethane. By using density functional theory (DFT) calculations and synchrotron resonant inelastic X-ray spectroscopy (RIXS), they attributed the improvement to the electronic effect where the structure of Pt 5d valence orbitals changed due to the presence of Zn. Our studies by including the performance of an additional Pt₃Zn₁/SiO₂ intermetallic catalysts support their conclusion. The improved performance of Pt₁Zn₁/SiO₂ is more likely due to a combined electronic and geometric effect. In other words, the electronic states of Pt 5d are changed by adding the Zn element. As shown by XANES in Fig. 3(a), the Pt L₃ edge

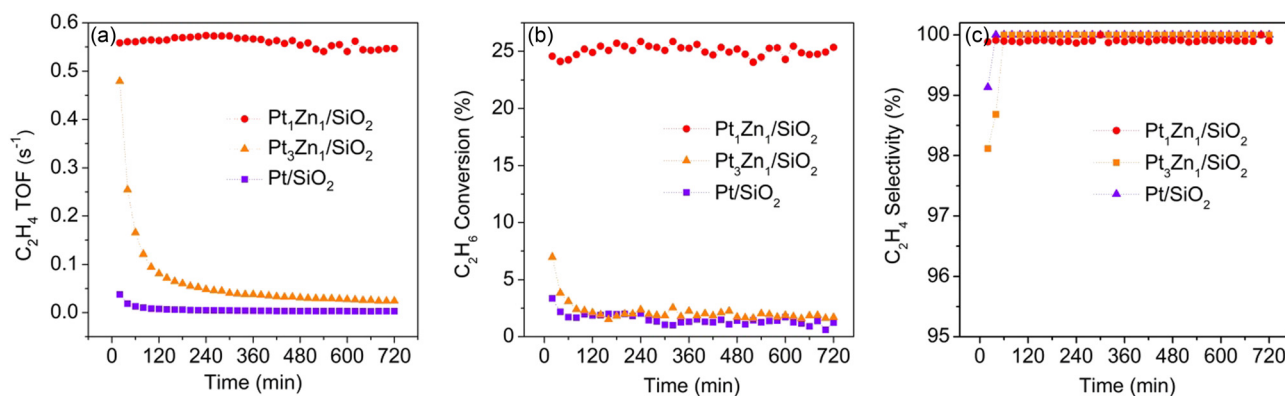


Fig. 7 (a) C₂H₄ TOF s⁻¹, (b) C₂H₆ conversion and (c) C₂H₄ selectivity as a function of reaction time (reaction conditions: *T* = 550 °C, C₂H₆ = 18% with balance He. 100 mg catalyst was used for each test. All catalysts were pre-treated with 5% H₂/He for 30 min at 600 °C before DHE, WHSV = 7.2 h⁻¹).



Fig. 8 (a) C₂H₄ TOF, (b) C₂H₆ conversion and (c) C₂H₄ selectivity as a function of reaction time. (d) Initial selectivity of the product gases on the catalysts. (Reaction conditions: $T = 550$ °C, C₂H₆ = 22.3%, O₂ = 1.7%, He = 76%, total flowrate = 50 ml min⁻¹. All catalysts were pre-treated with 5% H₂/He for 30 min at 600 °C before ODHE, initial conversion was at ~10% for all catalysts).

position of Pt₁Zn₁ and Pt₃Zn₁ shifted to higher energy by 2.4 eV and 1.5 eV, respectively, compared with Pt monometallic. Those shifts in edge energy reflect an upward shift of unfilled Pt 5d bands, which also cause a downward shift of filled 5d bands. These energy changes are related to the bonding of Pt and the adsorbate, which in turn affects the catalytic turnover on the Pt active site. These results are consistent with the initial performance of the catalyst in the order of Pt₁Zn₁ > Pt₃Zn₁ > Pt/SiO₂ in both DHE and ODHE reactions. Regarding the geometric effect, the Pt atoms are atomically more isolated in Pt₁Zn₁ (CN_{Pt-Pt} = 3.1, $R_{\text{Pt-Pt}}$ = 2.85 Å) than in either Pt₃Zn₁ (CN_{Pt-Pt} = 6.1, $R_{\text{Pt-Pt}}$ = 2.73 Å) or Pt (CN_{Pt-Pt} = 10.5, $R_{\text{Pt-Pt}}$ = 2.76 Å). The C₂H₄ TOF for both DHE and ODHE are compared with those state-of-the-art Pt-based catalysts from the literature (Tables S7 and S8, ESI[†]), Pt₁Zn₁/SiO₂ shows comparable or enhanced C₂H₄ TOF with other Pt–Zn or Pt–Sn nanocatalysts, considering our reaction temperature and space velocity. The more isolated Pt in Pt₁Zn₁ may be responsible for the improved catalyst stability as it slows the coke deposition. The use of intermetallic alloys could be a promising approach to synthesize well-defined, atomically dispersed catalysts.

4. Conclusions

Pt₁Zn₁/SiO₂ and Pt₃Zn₁/SiO₂ intermetallic nanocatalysts were synthesized by sequential deposition of ZnO (by ALD) and Pt

(by IWI) on a SiO₂ support, followed by hydrogen reduction with suitable temperature and length of time. In particular, the hydrogen reduction of Pt/ZnO/SiO₂ was investigated using *in situ* synchrotron X-ray absorption spectroscopy at both Pt L₃ and Zn K edges. The Pt₁Zn₁ intermetallic phase was formed at temperatures as low as 200 °C by the platinum catalyzed reduction. The formation rate of Pt₁Zn₁ was controlled by the ZnO reduction, not by diffusion of the metallic Zn into Pt. All Pt was completely in the Pt₁Zn₁ phase at 600 °C while excess ZnO still existed on the support. A higher temperature (at 800 °C for 15 min) or longer reduction time (at 600 °C for 6 h) is necessary to remove the excess ZnO and reach the Pt:Zn = 1:1 stoichiometry on the surface. Further reduction at 800 °C for 8 h will result in the Pt₃Zn₁ phase.

Pt₁Zn₁/SiO₂ exhibited an ethylene TOF much higher than those of Pt₃Zn₁/SiO₂ and Pt/SiO₂ in direct dehydrogenation and oxidative hydrogenation of ethane to ethylene. The Pt₁Zn₁/SiO₂ catalysts also exhibited considerably enhanced stability. The presence of Zn in Pt–Zn intermetallic alloys leads to both electronic structure and geometric structure changes of Pt. The electronic structure of Pt 5d orbitals, which concerns and influences the nanoparticle's catalytic properties, is similar in both Pt₁Zn₁ and Pt₃Zn₁, as determined by the Pt L₃ XAS spectra. This may explain the initial high TOF of both catalysts as compared to the Pt catalyst. The improved stability of Pt₁Zn₁ is likely associated with the geometric effect, as the Pt atoms in Pt₁Zn₁ are significantly more atomically isolated than the Pt atoms in Pt₃Zn₁ and Pt.

Conflicts of interest

There are no conflicts of interest to declare.

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