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# Photochemical CO<sub>2</sub> conversion on pristine and Mg-doped gallium nitride (GaN): a comprehensive DFT study based on a cluster model approach?

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The photochemical reduction of carbon dioxide (CO<sub>2</sub>) into methanol is very appealing since it requires sunlight as the only energy input. However, the development of highly selective and efficient photocatalysts is still very challenging. It has been reported that  $CO_2$  can be spontaneously activated on gallium nitride (GaN). Moreover, the photocatalytic activity for CO<sub>2</sub> conversion into methanol can be drastically enhanced by incorporating a small amount of Mg dopant. In this work, density functional theory (DFT) based on a cluster model approach has been applied to further explore the photocatalytic activity of bare GaN towards CO2 adsorption and conversion. We extended the investigation of Mg-doping replacing one Ga atom with Mg on three different sites and evaluated the consequent effects on the band gaps and CO2 adsorption energies. Finally, we explore different routes leading to the production of methanol and evaluate the catalytic activity of bare GaN by applying the energetic span model (ESM) in order to identify the rate-determining states which are fundamental for suggesting modifications that can improve the photocatalytic activity of this promising material.

### Introduction

The chemical transformation of carbon dioxide (CO<sub>2</sub>) into chemicals and fuel has been receiving growing attention since it can potentially mitigate the negative environmental impact of increasing greenhouse gas emissions in the atmosphere and furnish a new renewable carbon source. 1-7 In previous studies, we focused on homogeneous catalytic CO2 conversion to cyclic carbonates8,9 and methanol10 using sustainable alkali halideglycol complexes and Ru-complexes, respectively. Recently, our focus has been on photochemical CO2 reduction since it requires sunlight as the only energy input. Significant attention has been paid to this subject in recent years, and substantial advances in the hydrogenation of CO2 to hydrocarbons and oxygenates have been achieved. Among these products, methanol is one of the most desired since it can be used as a liquid fuel-like renewable energy source and a chemical feedstock for other useful chemicals. Due to the high thermodynamic stability of the CO2 molecule, the first step of its reduction involves the adsorption and its consequent activation on the photocatalyst surface.

Various semiconductor photocatalysts, including oxides, oxynitrides, sulfides, and phosphides, have been investigated over the last few decades. 11-18 It has been observed that the presence of oxygen vacancies in metal oxide surfaces dominates the CO2 adsorption. 19,20 However, this adsorption is hindered by the adsorption of O2, which occupies the vacancies leading to an inactive surface. Moreover, metal oxide semiconductors generally exhibit a large band gap, which limits the absorption of the visible and infrared solar spectrum. Therefore, it is of both fundamental and practical interest to explore alternative photocatalysts that can harvest a large part of the solar spectrum and can lead to spontaneous activation of the CO<sub>2</sub> molecule.

In this regard, the thermodynamically stable wurtzitestructure gallium nitride (GaN) has been proposed as a promising material for the reduction of highly stable CO<sub>2</sub> molecules due to its unique electronic and optical properties. GaN has a wide band gap of 3.4 eV which can be tuned to absorb nearly the entire solar spectrum by alloying it with In and Al. Mi and collaborators21 have demonstrated the photochemical reduction of CO<sub>2</sub> to CH<sub>3</sub>OH, CO, and CH<sub>4</sub> with sunlight by a joint experimental and computational study based on periodic boundary condition (PBC) calculations. The authors also discovered that, with the incorporation of Mg dopant, the rate of CO2 reduction to CH<sub>3</sub>OH, CH<sub>4</sub>, and CO is enhanced nearly 50-fold due to the reduced surface potential barrier and the enhanced adsorption of CO<sub>2</sub> molecules on the surface of the nanowire. In a very recent study, Liu and co-workers<sup>22</sup> demonstrated that bulk GaN

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is an efficient catalyst for the selective direct hydrogenation of CO2 to DME, and DME was rigorously revealed as the primary product. The authors have performed DFT calculations using PBC methods in order to investigate the details of the reaction conversion mechanism. According to the authors, the reaction proceeds through the formation of adsorbed H<sub>2</sub>COH\* species which can be hydrogenated to methanol or dissociate to CH2\* and OH\*. The energy barrier involved in the former process is calculated to be  $\sim 1$  eV, which is higher than the latter. Their DFT results well explain why DME rather than methanol is formed as the primary product on GaN surfaces, i.e. the difficult hydrogenation of CH<sub>2</sub>OH\* to CH<sub>3</sub>OH\* and the unfavourable formation of CH<sub>3</sub>O\*. The CO<sub>2</sub> photoreduction to hydrocarbons requires the presence of H<sup>+</sup> species which can be formed by water oxidation on GaN surfaces or via the dissociation of hydrogen molecules (H2). In this regard, the hydrogen evolution reaction<sup>23-26</sup> represents a competing reaction.

Calculations based on a cluster model approach have some advantages over PBC methods, such as the use of hybrid functionals at a lower computational cost, that allows for a highly accurate description of electronic properties; the ability to add/subtract a charge carrier without suffering from interactions with similar charge carriers due to periodic boundary conditions, 27,28 which makes cluster approaches suitable for the study of charged systems. Thanks to the reduced size, a cluster approach can be used to intercept all the stationary points, including the more complex transition states, involved in the catalytic cycle and to determine the rate determining states.<sup>29-31</sup> On the other hand, a PBC approach allows for the investigation of extended surfaces and provides an accurate description of the structural properties of the surface and molecular adsorbates. The complementarity of the two computational methods can therefore lead to an in-depth understanding of the structureproperty relationship.32

In this work, we have used DFT calculations based on a cluster model approach in order to further explore the photocatalytic activity of bare GaN towards CO2 conversion. Based on the findings reported by Mi and collaborators, 21 we have extended the investigation of Mg-doping replacing one Ga atom with Mg on three different sites and evaluated the consequent effects on the band gaps and CO2 adsorption energies. This provides a more realistic picture of the stochastic distribution of dopant atoms within the crystal and on its surface. Based on the results published by Liu and co-workers, 22 we have explored different routes leading to the production of methanol that could potentially involve lower energy barriers. The catalytic activity of bare GaN has been evaluated by applying the energetic span model (ESM)<sup>33,34</sup> in order to identity the ratedetermining states and to suggest modifications that can improve the photocatalytic activity of this promising material. Eventually, the influence of Mg-doping at different sites on the catalytic CO<sub>2</sub> conversion into methanol has been investigated. As will be discussed in more detail in the next sections, the reliability of the selected cluster models has been confirmed by the good agreement of the geometric parameters, that include calculated bond distances and angles, compared to the ones

reported in ref. 21. Moreover, similar trends of the adsorption energies using the two different computational approaches have been computed, which further support the potential use of the selected cluster models to determine adsorption energies as a valid alternative to the more extended surfaces used in PBC calculations. The stability of the pure and Mg-doped GaN clusters has also been evaluated and the details are discussed in the ESI,† Therefore, these models can be used for further investigations allowing a better understanding of the photocatalytic activity of GaN-based materials at a notably reduced computational cost.

## Results and discussion

#### GaN cluster selection and CO2 adsorption

Two different cluster models have been selected in this study: A\_GaN built up of three Ga-N moiety rows and 2 layers for a total number of 58 atoms; B\_GaN, including four Ga-N moiety rows and three layers (112 atoms). Both the optimized structures are shown in Fig. 1. For the sake of simplicity, throughout the text, all the structures labeled with "A" and "B" will refer to the smaller and bigger size cluster models, respectively, while "GaN" and "Mg" will be used to indicate the pristine and Mg-doped models, respectively. To identify the three different studied doping sites, the superscripts "surf", "FL" and "bulk" will be adopted to refer to the surface, first layer and bulk, respectively.

As the first step of our investigation, we have focused on the adsorption of CO2 on the surface of both the cluster models. The relative adsorption energies,  $E(CO_2)_{ads}$ , are calculated as:

$$E(CO_2)_{ads} = E_{cluster+CO_2} - E_{cluster} - E_{CO_2}$$
 (1)

where  $E_{\text{cluster}}$  is the energy of the optimized cluster,  $E_{\text{CO}_2}$  is the energy of the  $CO_2$  molecule, and  $E_{cluster+CO_2}$  is the energy of the cluster when the CO2 molecule is coordinated to the surface. The linear CO2 has therefore been added to both A\_GaN and B\_GaN at an initial distance of about 3 Å from the surface. In both the systems, during the relaxation, the CO2 is spontaneously coordinated to the GaN surface via a single C-N bond and the consequent deformation of the O-C-O from the original angle of 180° to a bent conformation (see Fig. 1 and Table 1). The C coordination is also confirmed by NBO analysis which underlines the formation of a single bond between C and N2, with a higher contribution of the N hybrid orbitals (63%) with respect to the C hybrid orbitals (37%), as well as a higher polarization coefficient of N (0.80) than C (0.60). NBO analysis also shows the formation of a double bond between C and O1 and a single bond between C and O2. In 1B\_GaN, two O-Ga interactions are established between the negatively charged oxygen atoms O1 (-0.718) and O2(-0.790) and the positive Ga1 (+1.740) and Ga3 (+2.123) atoms, respectively. As for 1A\_GaN, the interaction between O1 and Ga1 becomes stronger as underlined by the O1-Ga1 bond whose distance is 0.342 Å shorter than in 1B\_GaN. However, the O2-Ga3 bond distance is 0.712 Å longer. Overall, increasing the size of the cluster has little influence on the CO<sub>2</sub> adsorption mode as also shown by the selected geometric parameters reported in Table 1.

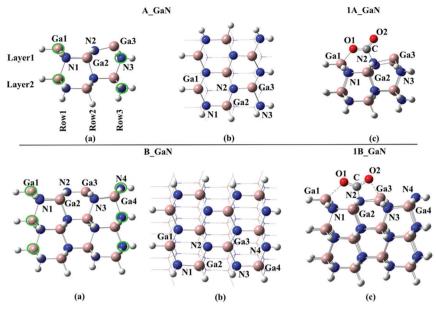


Fig. 1 Top: (a) side and (b) top view of the A\_GaN cluster model and (c) 1A\_GaN intermediate; bottom: (a) side and (b) top view of the B\_GaN cluster model and (c) 1B\_GaN intermediate.

Moreover, very similar adsorption energies of -2.11 and -2.05 eV have been calculated for the adsorption of CO2 on A\_GaN and B\_GaN, respectively, whose optimized structures 1A\_GaN and **1B\_GaN** are shown in Fig. 1c. Mi et al. 21 calculated an adsorption energy of 1.76 eV of one CO2 molecule on GaN(100) using PBC calculations. As shown in Table 1, the geometric parameters computed by the authors are in good agreement with those obtained using our selected cluster models.

The effect of Mg-doping on the CO2 adsorption energies has been evaluated on both the A GaN and B GaN models. Three different doping positions have been considered as shown in Fig. 2, while the corresponding adsorption energies are reported in Table 2. Mi et al. 21 computed an adsorption energy of -2.48 eV for  $CO_2$  on the Mg-doped GaN surface. The doping site considered by the authors can be compared to our A\_Mg<sup>surf</sup> and B\_Mgsurf systems, whose calculated adsorption energies are 3.23 and 3.12 eV for 1A\_Mgsurf and 1B\_Mgsurf, respectively. According to those results, our cluster models tend to overestimate the adsorption energies giving more negative values. However, both the methods show a much stronger CO<sub>2</sub> adsorption energy in the presence of the Mg dopant which corresponds to

Table 1 Adsorption energies (Ads energy) of CO<sub>2</sub> on A\_GaN and B\_GaN clusters, along with the main calculated geometric parameters. Values from ref. 21 are also shown for direct comparison

	A_GaN	B_GaN	Ref. 21	
	B3LYP/6-31+G**	B3LYP/6-31+G**	PBC	
Ads energy (eV)	-2.11	-2.05	-1.76	
C-N2 (Å)	1.475	1.418	1.41	
O1-Ga1 (Å)	1.918	2.323	2.24	
O2-Ga3 (Å)	2.793	2.081	2.10	
Ga1-N2-Ga3 (°)	170.4	167.0		
O1-C-O2 (°)	123.8	129.4	128.3	

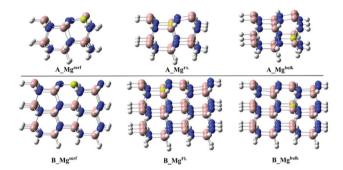


Fig. 2 Top: side view of the A\_Mg cluster models doped by replacing one Ga atom of the surface (A\_Mg<sup>surf</sup>), first layer (A\_Mg<sup>FL</sup>) and bulk (A\_Mg<sup>bulk</sup>) with Mg. Bottom: side view of the **B\_Mg** cluster models doped by replacing one Ga atom of the surface  $(B\_Mg^{surf})$ , first layer  $(B\_Mg^{FL})$  and bulk  $(B\_Mg^{bulk})$  with Mg.

0.7 eV for the PBC calculations in ref. 21, and 1.1 eV for our cluster model calculations.

As shown in Table 2, the stronger adsorption of the CO<sub>2</sub> molecule in the presence of the Mg dopant is particularly enhanced when Mg replaces a gallium atom of the surface as in A\_Mgsurf and B\_Mgsurf in Fig. 2.

On the other hand, when Mg occupies a "bulk-like" position (A\_Mgbulk and B\_Mgbulk in Fig. 2), the calculated adsorption energy is much more similar to the undoped case. For a better understanding of the influence of Mg doping on the adsorption energy, the total adsorption energy,  $E_{ads}$ , of  $CO_2$  on the B\_Mg

Table 2 Adsorption energies (eV) of CO<sub>2</sub> on A\_Mg and B\_Mg clusters, for the three studied doping sites

1A_Mg <sup>surf</sup>	1B_Mg.surf	1A_Mg <sup>FL</sup>	1B_Mg <sup>FL</sup>	1A_Mg <sup>bulk</sup>	1B_Mg <sup>bulk</sup>
3.23	3.12	2.95	2.54	2.65	2.29

clusters has been decomposed into two main contributions, i.e., interaction and distortion energies:

$$E_{\rm ads} = E_{\rm int} + E_{\rm dist}$$
 (2)

The interaction energy,  $E_{\rm int}$ , is given by the difference between the energy of the whole system and that of the single components taken at the final geometry of the adsorbed state, thus accounting for the bare electronic effects occurring at the adsorbate-surface interface:

$$E_{\rm int} = E_{\rm system} - \left(E_{\rm adsorbate}^* + E_{\rm surface}^*\right) \tag{3}$$

The distortion energy,  $E_{\rm dist}$ , is given by eqn (4):

$$E_{\text{dist}} = E_{\text{adsorbate}}^* + E_{\text{surface}}^* - E_{\text{adsorbate}} - E_{\text{surface}}$$
 (4)

where the terms with and without the \* are taken, respectively, at the geometry of the final adsorbed state and the unperturbed equilibrium structures of each single component. This contribution accounts for the structural arrangements of both molecule and surface occurring upon adsorption.

As seen in Table 3, a positive contribution of about 4 eV, due to the distortion of the bare and Mg-doped GaN as a consequence of the CO<sub>2</sub> adsorption, has been calculated. Moreover, the positive deformation energy confirms also that the CO2 molecule is activated upon coordination, changing from a linear structure to a bent one. On the other hand, the very negative terms of the interaction energy indicate a strong interaction between the CO<sub>2</sub> molecule and the surface favoring the overall adsorption process. Particularly, following the same trend of the adsorption energy, the interaction energy becomes stronger when the cluster models are doped with Mg, and a more negative interaction energy is calculated when Mg replaces a surface Ga atom as in 1B\_Mgsurf, while a more positive value is obtained for the "bulk-like" replacement. Therefore, the replacement of Mg on the surface will lead to a stronger interaction energy and an overall stronger adsorption energy.

Table 3 Calculated  $E_{\rm int}$  and  $E_{\rm dist}$  contributions to the CO<sub>2</sub> adsorption energy on the pristine **B\_GaN** and B\_Mg cluster models. Energies are given in eV

1B_GaN		1B_Mgsurf		1B_MgFI		1B_Mgbulk	
$E_{ m int}$	$E_{ m dist}$	$E_{ m int}$ $E_{ m dist}$		$E_{ m int}$ $E_{ m dist}$		$E_{ m int}$	$E_{ m dist}$
-5.94	3.89	-6.95	3.83	-6.63	4.10	-6.22	3.93

Table 4 Methanol adsorption energies, E, (eV) on the pristine A\_GaN and A\_Mg cluster models along with the calculated Ga-O distance (Å)

	A <sub>MeOH</sub> _GaN	$A_{MeOH\_}Mg^{surf}$	$A_{MeOH\_}Mg^{FL}$	A <sub>MeOH</sub> _Mg <sup>bulk</sup>
E Ga-O	-1.69 $2.009$	-1.52 $2.006$	-1.53 $2.012$	-1.54 $2.013$

Methanol production: The adsorption energies of one methanol molecule on the selected cluster models have been calculated and the corresponding values are reported in Table 4. Thanks to the agreement between the small and big cluster models on the computed CO2 adsorption energies, only the small clusters have been considered.

These values show that the adsorption of the methanol molecule is less strong (more positive) than that of CO<sub>2</sub> for the bare and Mg-doped clusters. However, the methanol adsorption energy on A\_GaN is slightly higher than the Mg-doped cases. The adsorption mode of the methanol molecule on the GaN is not particularly influenced by the Mg doping. Moreover, very similar adsorption energies have been calculated for the Mg-doping at the three different sites (see Table 4). Our calculations show that the adsorption of CO2 on pristine A\_GaN is 0.42 eV stronger than that of methanol (see Tables 1 and 4). This difference increases when the Mg-doped clusters are taken into account and it becomes 1.71, 1.42 and 1.11 eV (see Tables 2 and 4) for A Mg<sup>surf</sup>, A Mg<sup>FL</sup> and A\_Mg<sup>bulk</sup>, respectively. In other words, replacing one Ga atom with Mg favors the CO<sub>2</sub> adsorption (particularly when a surface Ga atom is replaced) but has no influence on the methanol adsorption facilitating its release.

#### Band gaps and HOMO/LUMO orbitals

For all six Mg-doped and two bare GaN clusters, the band gaps have been calculated as the energy difference between the Lowest Unoccupied Molecular Orbital (LUMO) and Highest Occupied Molecular Orbital (HOMO). The computed band gap of the pristine **B\_GaN** is 3.5 eV which is very similar to the experimental one (3.4 eV). According to our results, reported in Table 5, the presence of Mg lowers the band gap. This effect is much enhanced when Mg replaces a surface Ga as in B\_Mgsurf. Even though the smaller A\_GaN and A\_Mg models dramatically underestimate the band gaps, the obtained trend confirms that calculated with the bigger cluster model.

The band gaps can also be derived from the gap between the photoemission (PES) and inverse-photoemission (IPES) energy levels. 27,35 It is indeed possible to compute the electron affinity/ ionization potential of pristine and Mg-doped GaN by the addition/removal of an electron from the neutral system. This would be analogous to the measurement of the fundamental gap arising from the PES/IPES process, as explained in more detail in the ESI.† By applying this method to the smaller cluster models, A\_GaN and A\_Mgsurf, we obtain band gap values of 2.86 and 2.27 eV, respectively. Even though these values still underestimate the band gaps, they give more realistic values showing that the smaller clusters can be used for qualitative analysis at a very reduced computational cost.

Table 5 Band gaps of all the selected clusters calculated as the difference between the HOMO and LUMO orbitals. Energies are given in eV

A_GaN	B_GaN	A_Mg <sup>surf</sup>	B_Mg <sup>surf</sup>	$A\_Mg^{FL}$	$B\_Mg^{FL}$	A_Mg <sup>bulk</sup>	B_Mg.bulk
1.24	3.52	0.59	2.49	1.13	3.29	1.15	3.27

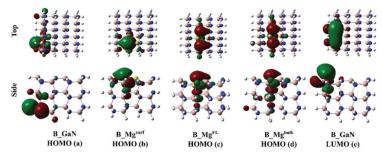


Fig. 3 Top and side view of the HOMO orbitals of **B\_GaN** (a), **B\_Mg<sup>surf</sup>** (b), **B\_Mg<sup>FL</sup>** (c) and **B\_Mg<sup>bulk</sup>** (d); top and side view of the LUMO orbital of the pristine **B\_GaN** (e).

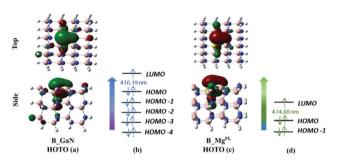


Fig. 4 Top and side view of the HOTO of  $B\_GaN$  (a) and  $B\_Mg^{FL}$  (c) cluster models; canonical orbitals involved in the transition to the first excited state of  $B\_GaN$  (b) and  $B\_Mg^{FL}$  (d).

Fig. 3 shows the HOMO orbitals of the **B\_GaN** and **B\_Mg** clusters. The HOMO orbital of the pristine GaN is localized on the bottom-left part of the cluster. On the other hand, when the surface Ga atom is replaced with Mg, as in **B\_Mg**<sup>surf</sup>, the HOMO orbital involves mostly the atoms of the surface layer. A more extended character of the HOMO orbital is calculated in the case of Mg-doping as in **B\_Mg**<sup>FL</sup> and **B\_Mg**<sup>bulk</sup>. The analysis of the LUMO orbitals, instead, shows a very similar character of the LUMO for all the studied clusters. For this reason, in Fig. 4 only the undoped case is shown.

Therefore, both the "A" and "B" cluster models have underlined the spontaneous coordination of the  $\mathrm{CO}_2$  molecule on the pristine and Mg-doped surfaces and its consequent activation is thanks to the formation of a strong C–N single bond. Moreover, both the "A" and "B" clusters confirm that replacing a surface Ga atom with Mg leads to stronger (more negative) adsorption and interaction energies, along with notably reducing the band gap.

#### Absorption spectra

TD-DFT, which is a popular tool for computing the electronic excitation energies and the properties related to the optical spectra, <sup>36–38</sup> has been used to investigate the light absorption of pure and Mg-doped GaN materials determining their vertical electronic absorption energies of optical transitions. The corresponding simulated spectra are shown in the ESI,† while the main excitation energies and the related oscillator strengths are listed in Table 6.

We report all excitations falling in the energy interval in terms of wavelengths between  $\sim\!700\text{--}350$  nm (corresponding to  $\sim\!1.8\text{--}3.5$  eV). This energy interval includes the visible spectrum (400–700 nm). The main difference between pristine and Mg-doped clusters is that the inclusion of Mg leads to absorption features at lower energies and this effect is particularly evident for the  $B\_Mg^{surf}$  cluster which includes a Mg surface atom: pure GaN shows peaks that are located at about 3 eV, which means  $\sim\!400$  nm in the visible region of violet light, and some further peaks are present out of the visible window (less than 400 nm); on the other hand, the presence of Mg shifts the absorption up to 666.32 nm, extending the range of the absorption of visible light and, consequently, improving the photocatalytic performance of gallium nitride.

The analysis of the TD-DFT calculations underlines that in the first excited state for both  $B\_Mg^{surf}$  and  $B\_Mg^{bulk}$  the HOMO–LUMO transition dominates with a transition coefficient of about 0.70. In the case of  $B\_Mg^{FL}$ , two transitions are involved: the HOMO–1  $\rightarrow$  LUMO and HOMO  $\rightarrow$  LUMO, with transition coefficients of 0.6 and 0.4, respectively (see Fig. 4d). Eventually, in the  $B\_GaN$  cluster, several transitions contribute to the first excited state with very similar coefficients (see Fig. 4b). In these two latter cases, examining the relevant canonical orbitals is not

Table 6 Main excitation energies (eV and nm) and related oscillator strengths (f) for the first five excited states of the **B\_GaN** and **B\_Mg** cluster models calculated with the TD-DFT method

	B_GaN			B_Mg <sup>st</sup>	B_Mg <sup>surf</sup> B_		B_Mg <sup>F1</sup>	B_Mg <sup>FL</sup>			B_Mg <sup>bulk</sup>		
	eV	nm	f	eV	nm	f	eV	nm	f	eV	nm	f	
ES1	2.98	416.2	0.0143	1.86	666.3	0.0033	2.61	474.9	0.0135	2.63	471.6	0.0237	
ES2	3.06	405.7	0.0114	2.36	525.6	0.0446	2.67	463.5	0.0167	2.71	456.8	0.0018	
ES3	3.08	402.1	0.0020	2.65	467.2	0.0102	2.85	434.7	0.0118	2.88	429.7	0.0036	
ES4	3.25	381.8	0.0038	2.68	461.8	0.0050	2.92	425.1	0.0061	2.95	420.1	0.0115	
ES5	3.28	377.6	0.0006	2.70	458.9	0.0217	2.95	419.9	0.0040	3.00	412.7	0.0188	

particularly helpful and therefore the analysis of the Natural Frontier Orbitals (NTOs) has been performed and the corresponding Highest Occupied Transition Orbitals (HOTOs) are shown in Fig. 4a and c for B\_GaN and B\_MgFL, respectively.

#### CO2 conversion on A\_GaN

As the first step of our investigation of the CO2 conversion into methanol, we have considered the coordination of: (a) a CO<sub>2</sub> molecule followed by the coordination of a hydrogen molecule (mechanism 1, Fig. 5) and (b) a hydrogen molecule followed by CO<sub>2</sub> (mechanism 2, Fig. 6). For a direct comparison with the adsorption energy given in the previous section, the electronic energies calculated at B3LYP/6-31+G\*\* will be discussed while comparing the reaction mechanisms. However, free energies are also reported in the Potential Energy Surfaces (PESs) and they will be discussed in more detail in the next section where the energetic span model will be applied in order to identify both the rate-determining intermediate, TDI, and the rate-determining transition state, TDTS.

**Mechanism 1.** As mentioned in the previous section, the addition of a CO2 molecule to A\_GaN leads to its spontaneous

coordination to the surface through the C-N1 bond as in 1A GaN, which lies at 2.11 eV below the reaction coordinate. When a hydrogen molecule is added to 1A GaN to form 2A GaN, no spontaneous adsorption is observed and the calculated H2-Ga bond distance is 2.573 Å. The reaction can proceed via TS2A GaN in which the calculated imaginary frequency of -1032.54i cm<sup>-1</sup> is associated with the breaking of the H-H bond and the consequent formation of the O-H and Ga-H bonds, as shown in 3A GaN. The energy barrier is 0.4 eV with respect to the previous intermediate. The second transition state, TS2A\_GaN, involves a hydride nucleophilic attack at the carbon atom and breaking of the C-OH bond. The calculated energy barrier with respect to 3A GaN is 2.9 eV. The IRC calculations show that in the subsequent intermediate, 4A GaN, the HCO group is rotated in order to form an O-Ga bond, while the OH group is transferred to the adjacent Ga surface atom. This intermediate is calculated to be  $\sim 1$  eV more stable than the previous one. A second uncoordinated molecule of H<sub>2</sub> is added to the system as in 5A GaN and it is heterolytically broken as shown by the imaginary frequency of -1094.62i cm<sup>-1</sup>

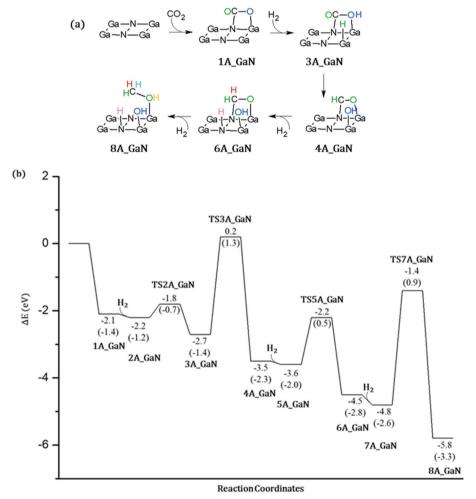


Fig. 5 (a) Schematic representations of all the intercepted stationary points along with the (b) calculated B3LYP energy profile for the CO<sub>2</sub> conversion to methanol on the A\_GaN cluster model according to mechanism 1. Energies are in eV and relative to the asymptote of the reactant. Free energies (eV) are shown in parentheses

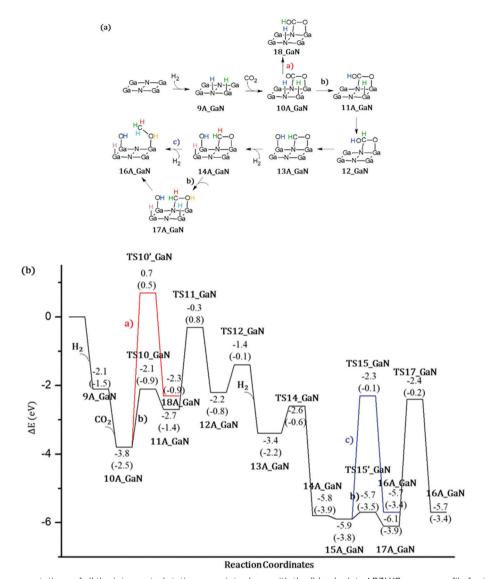


Fig. 6 (a) Schematic representations of all the intercepted stationary points along with the (b) calculated B3LYP energy profile for the CO2 conversion to methanol on the A\_GaN cluster model according to mechanism 2. Energies are in eV and relative to the asymptote of the reactant. Free energies (eV) are shown in parentheses

calculated in TS5\_GaN. The overcoming of an energy barrier of 1.3 eV leads to the formation of the 6A GaN intermediate in which a N-H bond and a C-H bond are formed. In this intermediate, the carbon atom of the H<sub>2</sub>C-O\* species coordinates to a surface nitrogen atom via a sp<sup>3</sup> carbon with a calculated bond distance of 1.572 Å, while the oxygen atom binds to a surface Ga atom via a Ga-O bond of 1.868 Å. The C-O bond is 1.409 Å which is longer than a C=O double bond. Therefore, these results show that this mechanism does not lead to the formation of acetaldehyde. Eventually the third molecule of hydrogen is added to the system which remains uncoordinated as in 7A GaN. The last transition state, TS7\_GaN, of the proposed mechanism involves the hydrogenation of the coordinated H<sub>2</sub>CO thanks to the breaking of the H-H bond and the formation of both CH and OH. The calculated imaginary frequency is -1307.49i cm<sup>-1</sup> and the involved activation barrier with respect to 7A\_GaN is 3.44 eV. Eventually, the 8A\_GaN intermediate is formed in which the

methanol molecule is coordinated via a Ga-OH bond whose calculated distance is 1.951 Å, while a hydrogen bond is also formed between the hydrogen of the methanol and the oxygen of the adsorbed OH, whose distance is 1.423 Å. This intermediate lies at 4.4 eV below the reaction asymptote.

Mechanism 2. As seen before, when a H<sub>2</sub> molecule is added to 1A\_GaN, neither spontaneous molecular nor dissociative adsorption is observed. Moreover, no transition state associates to the coordination of H2 or its dissociation and further coordination has been intercepted. On the other hand, our calculations show that when the H2 molecule is added to the **A\_GaN** in the absence of the coordinated CO<sub>2</sub>, it spontaneously gives dissociative adsorption leading to the formation of N-H and Ga-H bonds as in 9A\_GaN, which lies at 2.08 eV below the reactants' asymptote. Liu and co-workers<sup>22</sup> have performed PBC calculations showing that H2 molecules dissociate at Ga-N pairs in a heterolytic pathway with low activation barriers of

0.09 eV on GaN(100). The different result can be associated to the higher reactivity of cluster models in comparison with PBC calculations. However, both the approaches underline the ease of the hydrogen adsorption process. Our calculations show that 9A\_GaN is almost isoenergetic with 1A\_GaN and therefore there is no preference in the adsorption of one molecule over the other as the first step of the reaction mechanism. This result was also confirmed by Liu and collaborators, 22 who underlined that the comparable adsorption of H2 and CO2 paves the way for the facile hydrogenation of CO2 on the catalyst surface. The presence of the adsorbed hydrogen atoms on the GaN cluster does not prevent the further coordination, and consequent activation, of the approaching CO2 molecule. Indeed, our calculation shows that when the linear inactivated molecule of CO2 is added to 9A\_GaN at an initial C-N1 distance of 2.710 Å, during the optimization it coordinates as described in 1A\_GaN, leading to the formation of the intermediate 10A\_GaN which is found to be 1.68 eV more stable than the previous intermediate. From 10A\_GaN, the reaction can proceed via two different routes: (a) proton transfer from the surface nitrogen to the oxygen of the activated CO<sub>2</sub> (TS10 GaN); (b) nucleophilic attack at the activated carbon of CO2 by the hydride bound to the surface Ga atom (TS10'\_GaN). The former transition state is found to be 1.47 eV more stable than the latter, with a calculated activation energy of 1.62 eV with respect to 10A\_GaN. This result is also in agreement with those reported by Liu and co-workers<sup>22</sup> who report that the hydrogenation of CO2 to COOH\* is more favorable than that to HCOO. The formed intermediate 11A\_GaN is 2.66 eV more stable than the separated reactants. The reaction proceeds via TS11\_GaN in which the hydride attacks the electrophilic carbon coordinated to the surface nitrogen as underlined by the imaginary frequency of -1235.38i cm<sup>-1</sup>. In the formed intermediate, 12A\_GaN, the carbon atom shows sp<sup>3</sup> hybridization which involves a single bond with the previously formed OH group, a second single bond with the transferred hydrogen, a third single bond with the second oxygen of the original CO2 molecule, and eventually a fourth single bond with the surface nitrogen atom. Once the intermediate 12A\_GaN is formed, two different routes have been explored: (a) the former consists of the addition of a further molecule of hydrogen whose H-H bond breaks forming Ga-H and O-H bonds. As will be discussed in more detail in the ESI,† following this route which involves TS12'\_GaN, we did not succeed in intercepting the final transition state leading to the formation of methanol; (b) the latter involves the breaking of the C-OH bond which is broken via TS12\_GaN by overcoming an energy barrier of 0.72 eV with respect to 12A\_GaN, and leads to the formation of 13A\_GaN in which the OH group is transferred to the next Ga surface atom, while the carbon atom gets back to sp<sup>2</sup> hybridization. As seen for 1A\_GaN, when a second molecule of hydrogen is added to 13A\_GaN, it does not spontaneously coordinate to the surface. Three different mechanisms have been considered. We will describe the details of the selected one in this section while the details of the other two routes can be found in the ESI.† The H-H bond of the uncoordinated H<sub>2</sub> molecule is broken via

TS13\_GaN: one hydrogen atom adsorbs on the GaN surface

while the second binds the carbon atom forming the intermediate 14A\_GaN, in which the H2C-O species coordinates to a surface nitrogen atom in a similar way as seen for the 6A GaN intermediate and no acetaldehyde is formed. These results support the experimental findings based on which no acetaldehyde was observed. The calculated imaginary frequency is -520.83i cm<sup>-1</sup>, while the activation barrier is 0.8 eV with respect to 14A GaN. Also in this case, the addition of a further hydrogen molecule to the reaction system does not lead to its spontaneous adsorption and the molecule is kept at a distance of  $\sim 2.600 \text{ Å}$ from the surface as in intermediate 15A\_GaN. From this intermediate, two different mechanisms can occur which can be defined as (a) concerted and (b) step-wise mechanisms. The concerted mechanism consists of the direct hydrogenation of the adsorbed H<sub>2</sub>C-O species via the final TS15\_GaN whose imaginary frequency is -1284.50i cm<sup>-1</sup> and which involves the highest activation energy of 3.63 eV. The last intermediate 16A\_GaN lies at 5.70 eV below the reactants' asymptote and shows that the methanol molecule coordinates to a surface Ga atom whose calculated bond distance is 2.000 Å, along with a hydrogen atom adsorbed on a second surface Ga with a computed distance of 1.593 Å and eventually a Ga-OH bond of 1.793 Å. In the step-wise mechanism the added uncoordinated hydrogen molecule first breaks forming O-H and Ga-H bonds via TS15'\_GaN with a very low activation energy of 0.17 eV and forming the intermediate 17A GaN that lies at 2.44 eV below the separated reactants. In the last step, a C-H bond is formed through TS16\_GaN which involves an energy barrier of 3.70 eV.

It is important to note that the stabilization energies of the last intermediates (8A\_GaN and 16A\_GaN) of the proposed catalytic cycles are much more negative than the ones discussed in the previous section (A<sub>MeOH</sub> GaN and A<sub>MeOH</sub> Mg). Since the adsorption mode of the methanol molecule is the same, the more negative stabilization energies are due to the presence of more adsorbed species in the former case. Therefore, in order to get a more realistic view, the coverage effect should be investigated on a more extended surface, which will be the focus of our further investigation.

Energetic span model and rate determining states: According to the energetic span model (ESM), the turnover frequency (TOF) can be derived from electronic-structure calculations by identifying the energetic span of the cycle,  $\delta E$ , which is defined as the energy difference between the points of highest and lowest energy along the reaction profile corresponding to the complete catalytic cycle. Kozuch and Shaik<sup>33</sup> extended this model including the  $\Delta G_r$  of the reaction and defining  $\delta E_{ii}$  as:

$$\begin{cases} \delta E_{ij} = G(\text{TS})j - G(\text{I})i & j > i \\ \delta Eij = G(\text{TS})j - G(\text{I})i - |\Delta G_{\text{r}}| & j \leq i \end{cases}$$

Therefore, if the j-th transition state lies after the i-th intermediate, then  $\delta E_{ij}$  is just the energy difference between them; if the j-th transition state precedes the i-th intermediate, then  $\delta E_{ij}$  is their energy difference minus the  $\Delta G_r$ . The highest  $\delta E_{ii}$  corresponds to the energy span of the cycle connecting the

TOF-determining intermediate (TDI) and the TOF-determining transition state (TDTS). By applying this method to the potential energy surface of Fig. 5 and considering the free energies given in parentheses, we can identify 6A GaN as the TDI and TS5A\_GaN as the TDTS with a corresponding  $\delta E$  of 3.7 eV. Therefore, the rate determining state involves the last hydrogenation for the production of methanol. The same  $\delta E$ , involving very similar structures of TDI and TDTS, is calculated when the (c) route of the potential energy surface in Fig. 6 is considered. Along the (b) route, the application of the ESM leads to the identification of the rate-determining states, which involve the formation of the third C-H bond and the consequent production of methanol. Therefore, our calculations show that the last step leading to the methanol formation is the most challenging. Interestingly, no difference in terms of energy barrier has been found between the hydrogenation of the adsorbed H<sub>2</sub>CO species with an approaching inactivated hydrogen molecule, and the transfer of an adsorbed hydrogen atom. This is due to the very strong bonds that the adsorbed species establish with the GaN surface atoms that prevent easy bond breaking/formation. The application of the ESM to the mechanism proposed by Liu and collaborators<sup>22</sup> leads to a  $\delta E$  of 3.3 eV. In their investigation, the TDI is the intermediate in which the CO2 molecule is activated on GaN as the bent COO\* species while H2 is adsorbed dissociatively on Ga-N Lewis pairs (an intermediate that can be compared to our 10A GaN). The TDTS of the catalytic cycle proposed by Liu and collaborators<sup>22</sup> is the transition state associated to the C-H bond formation via an adsorbed hydrogen transfer to the H2COH\* (comparable to TS17\_GaN). Therefore, their study also identifies the last step of the methanol production as the rate-determining state. The authors indeed suggest that the dissociation of CH2OH\* into CH2\* (which is further hydrogenated to CH3\*) and OH\* is significantly more favorable than the formation of methanol, explaining why their experimental findings show that methanol is not the primary product of the CO<sub>2</sub> hydrogenation.

Even though we were not able to identify an alternative mechanism for the methanol production involving lower energy barriers, our investigation furnishes a wider picture of the possible conversion route which contributes to a deeper understanding of the complex reaction mechanism. Moreover, the application of the ESM allows us to identify the rate determining states which is key information for the development of catalysts with improved performances.

## Experimental

All the calculations have been performed using a cluster model approach. Two cluster models of 58, **A\_GaN**, and 112, **B\_GaN**, atoms have been carved out from the crystallographic structure of wurzite GaN, cutting along the (100) plane. We retained the cluster's overall zero spin and neutral charge by passivating with capping H atoms only those dangling bonds of our cluster that would have formed covalent bonds to Ga or N atoms in a bulk surface structure. The capping H atoms along with the outermost N and Ga atoms, underlined in green in Fig. 1, have

been kept fixed while the rest of the atoms were allowed to relax. These optimizations were performed at the B3LYP39 level of theory as implemented in the Gaussian 16 quantum chemistry package<sup>13,40</sup> and using the LANL2DZ<sup>14,41</sup> pseudo potential for Ga atoms and the fully electron basis sets 6-31G for N and H. To check the accuracy of the selected computational protocol, preliminary tests were performed using the more accurate double hybrid PBE0DH<sup>15,42</sup> functional that combines exact HF exchange with an MP2-like correlation to a DFT calculation. Moreover, the Counterpoise error as implemented in Gaussian 16 software has been evaluated. These results, whose details can be found in the ESI,† have confirmed the reliability of the B3LYP functional, which was therefore selected in our study. Frequency calculations have been done to check the nature of the stationary points (minimum and transition state). In order to obtain more accurate adsorption energy values, single-point calculations using the more extended 6-31+G\*\* basis sets, except for Ga, on the fully optimized geometry, have also been performed and the obtained energies will be discussed throughout the text (unless differently specified). To evaluate the influence of Mg on CO2 adsorption, we replaced the three-coordinated Ga(+3) atom on the surface of the pristine A\_GaN and B\_GaN cluster models by Mg(+2), leading to the formation of A\_Mg and B\_Mg clusters, whose net charge is -1.

### Conclusions

In this work, DFT calculations based on a cluster model approach have been used to explore the photocatalytic activity of bare GaN towards CO2 conversion. Two different cluster models have been selected: A\_GaN built up of three Ga-N moiety rows and 2 layers for a total number of 58 atoms and B\_GaN, including four Ga-N moiety rows and three layers (112 atoms). We extended the investigation of the Mg-doping effect considering three different doping sites. Our calculations have shown that the CO2 molecule spontaneously coordinates on the pristine and Mg-doped surfaces for both the selected cluster models. This coordination implies the consequent activation of CO2 thanks to the formation of a strong C-N single bond and the change from the CO2 linear structure to a bent one. The activation of the CO<sub>2</sub> molecule upon coordination is also confirmed by a similar positive contribution of about 4 eV for all the investigated cases related to its deformation along with the distortion of the bare and Mg-doped GaN clusters as a consequence of the CO2 adsorption. On the other hand, both the clusters confirm that replacing a surface Ga atom with Mg leads to stronger (more negative) adsorption and interaction energies. The most negative interaction energy is calculated when Mg replaces a surface Ga atom as in 1B\_Mgsurf, while the more positive values are obtained for the "bulk-like" replacement. Therefore, the replacement of Mg on the surface leads to a stronger interaction energy and an overall stronger adsorption energy.

Small pristine **A\_GaN** and Mg-doped **A\_Mg** clusters have been used to calculate the adsorption energies of one molecule of methanol which is the final product of the catalytic cycle.

The calculated values have shown that the adsorption of a methanol molecule is less strong (more positive) than that of CO2 with the bare A\_GaN showing a stronger MeOH adsorption energy. The adsorption mode of the methanol molecule on GaN is not particularly influenced by the Mg doping and doping sites, as underlined by the very similar calculated adsorption energies. The highest difference between CO2 and the MeOH adsorption energy of 1.71 eV was calculated when the two molecules were adsorbed on the A Mgsurf cluster. Therefore, our calculations show the important contribution of replacing one surface Ga atom with Mg, which favors the CO<sub>2</sub> adsorption and its consequent activation, but has no influence on the methanol adsorption facilitating its release.

Using the bigger cluster models, the band gaps have been calculated as the difference between the HOMO/LUMO orbital energies. The calculated band gap of 3.52 eV obtained for the pure B\_GaN cluster is in good agreement with the experimental band gap (3.4 eV). Our calculations show that Mg-doping contributes to decrease the band gaps and this effect is particularly important in the case of B\_Mgsurf for which a band gap value of 2.49 eV has been calculated. On the other hand, the small cluster models tend to drastically underestimate the band gaps even though they confirm the trend and, therefore, they could be used for a qualitative analysis.

TD-DFT has been used to calculate the vertical electronic absorption energies at optical wavelengths for the pure and Mg-doped GaN model clusters. The main difference between pristine and Mg-doped clusters is that the inclusion of Mg leads to absorption features at lower energies and this effect is particularly evident for the B\_Mgsurf cluster which includes a Mg surface atom: pure GaN shows peaks that are located at about 3 eV, which means ~400 nm in the visible region of violet light, and some further peaks are present out of the visible window (less than 400 nm); on the other hand, the presence of Mg shifts the absorption up to 666.32 nm extending the range of the absorption of visible light and, consequently, improving the photocatalytic performance of gallium nitride.

Thanks to the agreement of the two size cluster models, the smaller and less computationally demanding A\_GaN was selected for the investigation of CO<sub>2</sub> conversion into methanol. The catalytic activity of bare GaN was evaluated by applying the energetic span model in order to identify the ratedetermining states. Following the proposed mechanism 1, the rate determining state involves the last hydrogenation for the production of methanol with a computed  $\delta E$  of 3.7 eV. The same  $\delta E$ , involving very similar structures of TDI and TDTS, is calculated when the (c) route is considered. Along the (b) route the application of the ESM leads to the identification of the rate-determining states which involve the formation of the third C-H bond and the consequent production of methanol. Therefore, our calculations show that the last step leading to methanol formation is the most challenging.

Therefore, the selected cluster models can be used for further investigations allowing a better understanding of the photocatalytic activity of GaN-based materials at a notably reduced computational cost.

## **Author contributions**

V. B. performed the DFT simulations and analysis. Both the authors contributed to the manuscript.

## Conflicts of interest

There are no conflicts to declare.

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