

## PAPER

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# Degradation of CV dye by the as-synthesized Fe<sup>0</sup>–TiO<sub>2</sub> supported clinoptilolite under UV and solar irradiations

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Textile industries release toxic organic dyes into wastewater, harming aquatic ecosystems and affecting photosynthesis. This study aims to synthesize a photocatalyst for the efficient degradation of crystal violet (CV) dye in aqueous media. Titanium dioxide (TiO<sub>2</sub>) is the most preferable photo-catalyst, but its fast electron–hole recombination rate and low adsorption capacity have limited its applications on a large scale. To enhance the adsorption and degradation efficiency, a TiO<sub>2</sub>-supported clinoptilolite (CP) and a porous composite of zerovalent iron (Fe<sup>0</sup>) co-doped with titanium dioxide/c clinoptilolite (Fe<sup>0</sup>–TiO<sub>2</sub>/CP) were synthesized using sol–gel and borohydride reduction methods, respectively. The effects of various parameters like acidity, temperature and concentration on the photo-catalytic activity, morphological and micro-structural features and surface areas of different TiO<sub>2</sub>/CP and Fe<sup>0</sup>–TiO<sub>2</sub>/CP composites were characterized by various techniques such as X-ray diffraction (XRD), Scanning electron microscopy (SEM), energy dispersive X-ray analysis (EDXA), Fourier transform infrared (FTIR) spectroscopy, UV-Visible spectroscopy, thermogravimetric (TG) and differential thermogravimetric (DTG) analyses and Brunauer–Emmett–Teller (BET) isotherm. The as-synthesized composites (TiO<sub>2</sub>/CP and Fe<sup>0</sup>–TiO<sub>2</sub>/CP) were used as photocatalysts to remove CV dye from water. The parent CP and TiO<sub>2</sub> achieved 51% and 58% removal efficiency of CV dye under UV radiation, respectively. The 0.25 M TiO<sub>2</sub>/CP composite showed the highest degradation efficiency (92.3%) under UV radiation, while 0.1 M TiO<sub>2</sub>/CP performed best under solar radiation (88.5% removal) in 120 minutes. Theoretical analysis via kinetic models revealed that the adsorption and degradation processes of CV dye followed pseudo-second-order (PSO) and pseudo-first-order (PFO) kinetic models, respectively. The TiO<sub>2</sub>/CP composite mainly produces hydroxyl radicals (•OH) during dye degradation, while Fe<sup>0</sup>–TiO<sub>2</sub>/CP generates both •OH and superoxide radicals (•O<sub>2</sub><sup>–</sup>). This •O<sub>2</sub><sup>–</sup> radical enhances the degradation efficiency of Fe<sup>0</sup>–TiO<sub>2</sub>/CP due to Fe<sup>0</sup>'s favorable reduction potential. CV mineralization occurs through two pathways: N-demethylation and hydroxyl radical attack on the central carbon, leading to degradation and complete mineralization. Moreover, the structure, morphology and particle size of the composite play vital roles in the extent of their photocatalytic efficiencies. Therefore, a combination of compositional and structural engineering of TiO<sub>2</sub>-based photocatalysts is expected to give better device performance. However, further investigation is needed in the near future.

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## 1. Introduction

Water scarcity is the most serious challenge of today's growing world. The usage and disposal of different biological and chemical effluents in water bodies cause rapid quality degradation of water and chronically deplete the water resources.<sup>1</sup> Various non-pretreated dyes from the textile industries are continuously getting released into the water, endangering it to the point where it is unsafe for consumption and putting the population's health at risk.<sup>2</sup> CV dye is important in the artificial coloration of textile materials on a large scale. It is a water-

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soluble and non-biodegradable dye, and its improper discharge into water bodies has raised a major environmental concern.<sup>3</sup> By using various eco-friendly methods, wastewater containing dyes is efficiently treated. In this study, the combination of adsorption and heterogeneous photo-catalysis is employed for the degradation of CV dye under dark and light-assisted UV and solar irradiations.<sup>4</sup> The key photocatalyst responsible for the degradation is the semiconductor  $\text{TiO}_2$ .<sup>5–7</sup>  $\text{TiO}_2$  exhibits a number of drawbacks, including limited reaction efficiency under solar radiations, difficulty in the catalyst recovery steps, limited surface area that results in poor adsorption capacity, a large bandgap and fast rate of electron-hole recombination, which in turn limited its practical application as a photocatalyst on large scale.<sup>5–7</sup>

To enhance the photo-catalytic and adsorption efficiency of  $\text{TiO}_2$ , several strategies are employed, such as morphology controlling,<sup>8</sup> elemental doping,<sup>9</sup> surface loading<sup>7,10</sup> and energy band engineering.<sup>11,12</sup> Among these strategies, the most promising is the loading of  $\text{TiO}_2$  onto the surface of nanoporous adsorbents and elemental doping with  $\text{TiO}_2$  to overcome the drawbacks of  $\text{TiO}_2$  efficiency.<sup>7</sup> Clinoptilolite is an effective supporting adsorbent for  $\text{TiO}_2$  due to its characteristics, such as porosity (34%), excellent resistance to extreme temperatures, ion-exchange capacity ( $2.16 \text{ meq g}^{-1}$ ) and bulk density ( $1.15 \text{ g cm}^{-3}$ ).<sup>5,6,13</sup>

The energy difference between the valence band (VB) and conduction band (CB) of  $\text{TiO}_2$  is in the UV radiation region, so the radiation absorption capability in the visible region is still relatively low. Using solar light as an energy source is less effective in the  $\text{TiO}_2$  photocatalytic process because the composition of UV light in the sun is only 5–7%.<sup>14</sup> Thus, modification of the  $\text{TiO}_2$  photocatalyst is required in order to absorb light in the solar light region. To expand the photocatalyst response into the visible spectrum, an iron metal-doped  $\text{TiO}_2$  ( $\text{Fe}^0\text{-TiO}_2$ ) composite was found to be an effective photocatalyst for dye degradation when exposed to UV and visible light.<sup>15</sup>

In the present study, different molar  $\text{TiO}_2/\text{CP}$  composites were prepared using the sol-gel method. The second type of composites used for the degradation of CV dye was  $\text{Fe}^0\text{-TiO}_2$  and  $\text{Fe}^0\text{-TiO}_2/\text{CP}$ , which were synthesized using a simple borohydride reduction method. The adsorption rate and photocatalytic degradation efficiency of CV dye from aqueous solution by various molar composites of  $\text{TiO}_2/\text{CP}$ ,  $\text{Fe}^0\text{-TiO}_2$  and  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  and the role of active free radicals, such as  $\text{OH}^\bullet$ ,  $\text{O}_2^\bullet$  and  $\text{h}^+$ , in CV dye degradation were studied using a UV-visible spectrophotometer. The effects of various parameters, like temperature, concentration on photo-catalytic activity, micro-structural feature, surface area, pore size, pore volume, crystal phase and photo-catalytic capacity of  $\text{TiO}_2/\text{CP}$ ,  $\text{Fe}^0\text{-TiO}_2$  and  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  composites, were characterized using X-ray diffraction (XRD), scanning electron microscopy (SEM), energy dispersive X-ray analysis (EDXA), Brunauer-Emmett-Teller isotherm (BET), and Fourier transform infrared (FTIR) spectroscopy and thermogravimetric (TG) and differential thermogravimetric (DTG) analyses. The adsorption and degradation kinetic processes of CV dye were theoretically analyzed using pseudo-first-order and pseudo-second-order models, respectively. The degradation

pathway of CV dye was studied using heated electrospray ionization (H-ESI) coupled with a mass spectrometer.

## 2. Experimental section

### 2.1. Materials

Clinoptilolite was supplied by VSK ProZeo Zeolite, Slovakia. CV dye of high purity,  $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$  and sodium borohydride (98%) were purchased from Sigma Aldrich.  $\text{TiCl}_4$  was purchased from Fluka, and ammonia (32%) and HCl (37%) were purchased from Merck Schuchardt, Germany.

### 2.2. Synthesis and post-synthesis

$\text{TiO}_2$  and different molar ratio composites of  $\text{TiO}_2/\text{CP}$  were synthesized using the sol gel method, as reported by Ullah *et al.* in 2021.<sup>7</sup> In order to synthesize the  $\text{Fe}^0/\text{TiO}_2$  composite, a 30 mL solution was synthesized by mixing ethanol and water (5 : 25 v/v), and 0.800 g of already synthesized  $\text{TiO}_2$  was added to the mixture with continuous stirring. Then, 1.200 g of  $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$  was added to the  $\text{TiO}_2$  solution and stirred for 30 minutes under vacuum. After that, a 40 mL solution of sodium borohydride (1.200 g) was prepared in deionized water and added dropwise to the  $\text{TiO}_2$  solution under an inert atmosphere with continuous stirring to achieve maximum reduction. With the addition of the first drop of  $\text{NaBH}_4$ , a black precipitate of  $\text{Fe}^0/\text{TiO}_2$  was formed. The black precipitates were separated after complete reduction, washed with ethanol and then vacuum dried at  $60^\circ\text{C}$ . The sample was ground and stored at room temperature in an air-tight container to protect the composite powder from oxidation by air. The  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  was also prepared using the method mentioned above with already synthesized 0.1 M  $\text{TiO}_2/\text{CP}$  (0.800 g) powder.

### 2.3. Adsorption and photo-catalytic activity of the composites

The removal of CV dye through adsorption and photocatalytic degradation was conducted at room temperature in duplicate. A 100 mL solution of 10 ppm CV dye was prepared in a 250 mL beaker. Then, 3 mL of the solution was extracted for zero/standard reading. Afterwards, 50 mg of photocatalyst was added, and the mixture was stirred for 30 minutes at 140 rpm in the dark using a hot plate and magnetic stirrer to achieve adsorption/desorption equilibrium. Moreover, during the adsorption process, a 3 mL sample was withdrawn from the suspension at a specific time interval (10, 20 and 30 minutes). In order to study the dye degradation, the same solution was then placed under UV radiation, and 3 mL sample was withdrawn at a particular time interval (10, 20, 40, 60, 90 and 120 minutes). After irradiation, the withdrawn samples were centrifuged at 3000 rpm for 8 minutes to remove suspended particles of the composite from the solution.

The light source used was a UV lamp ( $0.02 \text{ mW cm}^{-2}$ ) enclosed in a rectangular wooden box. A UV-Vis spectrophotometer (PerkinElmer, Lambda 25) was used to analyze the absorbance of CV dye concentration at  $\lambda_{\text{max}}$  (585 nm). The same dye degradation procedure that was previously performed



under a UV source was employed using solar radiation in a photochemical reactor. The percentage removal efficiency and the adsorbed quantity of dye per unit mass ( $\text{mg g}^{-1}$ ) of adsorbent at equilibrium were calculated using the following equations:

$$\text{Removal\%} = \frac{(C_0 - C_t)}{C_0} \times 100 \quad (1)$$

$$\text{Adsorption capacity}(q_t) = C_0 - C_t \times \frac{\text{Vol of soln in Liter}}{\text{Mass of adsorbent in grams}} \quad (2)$$

where  $C_0$  and  $C_t$  are the concentrations ( $\text{mol L}^{-1}$ ) of CV dye at the initial stage ( $t = 0$ ) and at time  $t$  (minutes), respectively.

#### 2.4. Degradation pathway study of CV dye

Two 5 mL samples were withdrawn at two time intervals (20 and 120 minutes) after irradiation during the photocatalytic activity experiment. The withdrawn samples were centrifuged at 3000 rpm for 8 minutes to remove the suspended particles of the composite from the solution. The analysis of the CV dye degradation pathways was performed using a triple quadrupole mass spectrometer, model TSQ Quantis (Thermo Electron Scientific, USA), equipped with a heated electrospray ionization (H-ESI) source. The instrument was operated in both negative and positive ion modes, with a capillary voltage of 4.0 kV and a capillary temperature of 275 °C. Methanol was used as the solvent, and the sample was introduced using the direct insertion method at a flow rate of  $10 \mu\text{L min}^{-1}$ . The sheath gas flow rate was set to 15 units, and the auxiliary gas flow rate was set to 3 units. The mass spectrometer was set to scan a mass range of 30–1500  $m/z$ . Fragmentation (MS/MS) was performed on various peaks selected for fragmentation using collision energy ranging from 10 to 40 V. The instrument was controlled, and the data were acquired using Xcalibur software. This setup allowed for a detailed analysis of the CV dye degradation pathways, providing insights into the structures of the degradation products and the mechanisms of degradation.

#### 2.5. Physicochemical characterization techniques

The physical and chemical properties, composition, quality, stability, and purity of the synthesized composites were examined by physicochemical characterization.<sup>16</sup>

The crystalline structure of CP, as-synthesized  $\text{TiO}_2$ , composites of synthesized  $\text{TiO}_2/\text{CP}$ , and  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  were characterized using XRD (Bruker D8) with Cu Ka ( $\lambda = 0.1542 \text{ nm}$ ) radiation. The morphology of the samples mentioned above was studied using SEM (JEOL SEM IT200) operated at an accelerating voltage between 13.0 and 16.0 kV, and  $\text{TiO}_2$  nanoparticles were investigated using a JEOL JEM-2100 TEM. The TG/DTG analysis was performed using a simultaneous thermal analyzer (PerkinElmer STA 8000) at a heating rate of  $10 \text{ }^\circ\text{C min}^{-1}$  under  $\text{N}_2$  flow in the temperature range of 0–900 °C. A JW-BK-300C surface area and micropore size analyzer (Beijing JWGB Sci and Tech Co., Ltd) were used to assess the sample's BET surface area by  $\text{N}_2$  adsorption/desorption isotherms at 77 and 298 K.

The FTIR spectrometer (Bruker ALPHA-T) was used in the wave-number range of  $4000\text{--}400 \text{ cm}^{-1}$  to evaluate the chemical structures of the composites. The band gap of the composites was analyzed using the absorbance spectra of the samples recorded on a UV-VIS spectrometer (UV-2600, SHIMADZU).

### 3. Results and discussion

#### 3.1. Structural analysis

The XRD patterns of parent CP,  $\text{TiO}_2$  and their composites of various molarities are shown in Fig. 1. The characteristic peaks of CP, the anatase phase of  $\text{TiO}_2$  and  $\text{Fe}^0$  are confirmed by comparing their values with JCDPS values reported in the literature.<sup>5</sup> The specific peaks observed at  $25.4^\circ$ ,  $37.5^\circ$ ,  $47.7^\circ$ ,  $53.6^\circ$ ,  $54.7^\circ$  and  $62.3^\circ$  are assigned to the anatase phase of  $\text{TiO}_2$ , as shown in Fig. 1(a). The peaks present at  $9.8^\circ$ ,  $22.1^\circ$ ,  $26.4^\circ$ ,  $30.1^\circ$  and  $35.7^\circ$  depicted in Fig. 1(b) are attributed to the main features of CP. Moreover, in the composites of  $\text{TiO}_2/\text{CP}$  of different molarities (Fig. 1(A)–(c)–(f)), the characteristic peaks of both  $\text{TiO}_2$  and CP are observed, revealing that their individual crystalline structure is intact and unchanged during the composite formation. However, a decrease in the peak intensities and a slight shift of peaks of CP towards higher 2 theta angles are observed, which confirms the successful loading of  $\text{TiO}_2$  on the surface of CP.<sup>5</sup> This is attributed to the fact that during the synthesis, high calcination temperature may cause dealumination and desalination, which in turn leads to a decline in peak intensity of CP. Moreover, as shown in Fig. 1(A)–(c)–(f), the characteristic peak intensity of  $\text{TiO}_2$  at  $2\theta = 25.4^\circ$  increases linearly with an increase in molar ratios of  $\text{TiO}_2$  in 0.1, 0.25, 0.5 and 1.0 M  $\text{TiO}_2/\text{CP}$  composites.

In the  $\text{Fe}^0\text{-TiO}_2$  composite, as shown in Fig. 1(B)–(g), the apparent peaks of  $\text{Fe}^0$  at  $2\theta = 44^\circ$  with characteristic peaks of  $\text{TiO}_2$  are observed, which confirms the successful doping of  $\text{Fe}^0$  with  $\text{TiO}_2$ . Two very small peaks at  $2\theta = 31.9^\circ$  and  $33.7^\circ$ , as shown in Fig. 1(B)–(g), indicate the presence of iron oxide ( $\text{FeO}$ ) in the sample, which is formed during the doping of  $\text{Fe}^0$ , as reported in the literature. The crystalline structure of  $\text{TiO}_2$  was not affected by co-doping with  $\text{Fe}^0$ , which is in close agreement with a previous study conducted by Bibi *et al.*<sup>17</sup>

In the XRD pattern of the  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  composite (Fig. 1(B)–(h)), the characteristic peaks of  $\text{TiO}_2$  and  $\text{Fe}^0$  are relatively less intense compared to the peaks of CP. The loading of  $\text{TiO}_2$  and  $\text{Fe}^0$  onto the CP framework results in a decrease in the overall peak intensities of CP,  $\text{TiO}_2$  and  $\text{Fe}^0$ . Notably, the incorporation of  $\text{Fe}^0$  does not significantly alter the crystalline structure of the zeolite, except for a reduction in peak intensities. This suggests that the  $\text{Fe}^0$  particles are likely dispersed or incorporated within the CP framework without disrupting its underlying structure.<sup>18</sup>

The incorporation of  $\text{TiO}_2$  into the clinoptilolite zeolite framework led to a reduction in crystallinity, resulting in decreased sharpness and intensity of the XRD peaks of CP in all molar composites. This phenomenon occurs because the introduction of  $\text{TiO}_2$  disrupts the ordered arrangement of the zeolite's framework, causing distortions and a slight decrease



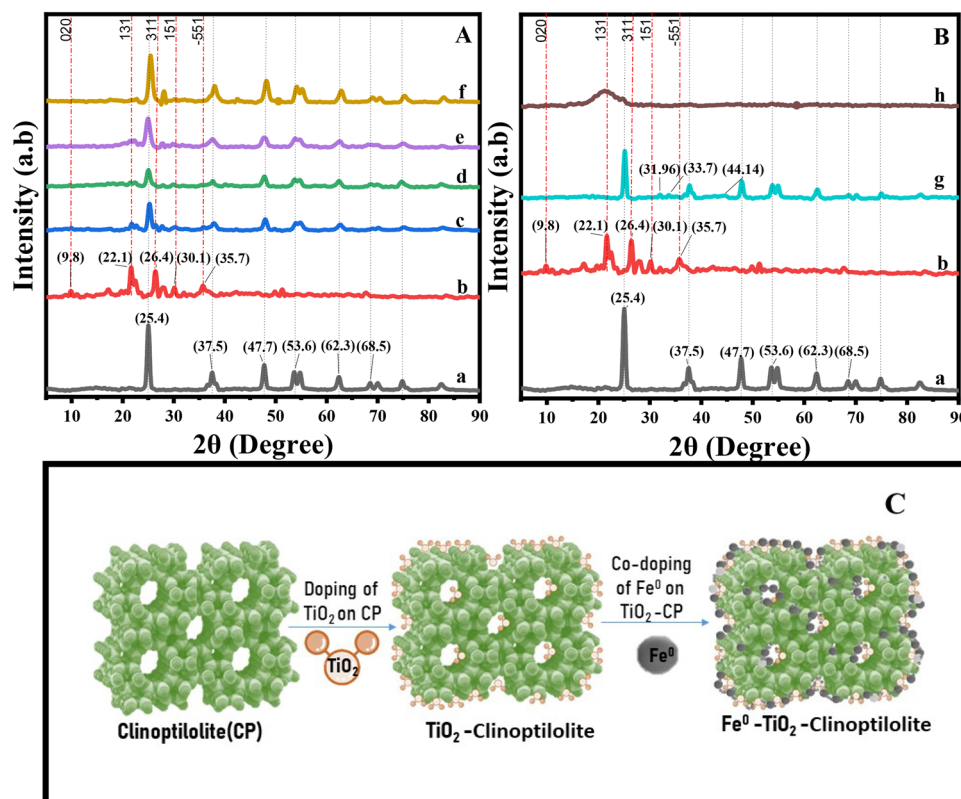


Fig. 1 XRD patterns of A: (a)  $\text{TiO}_2$ , (b) parent CP, (c) 0.1, (d) 0.25, (e) 0.5 and (f) 1.0 M  $\text{TiO}_2/\text{CP}$  and B: (g)  $\text{Fe}^0\text{-TiO}_2$  and (h)  $\text{Fe}^0\text{-TiO}_2/\text{CP}$ . C: doping and co-doping of  $\text{TiO}_2$  and  $\text{Fe}^0$ , respectively, on CP.

in crystallinity. As illustrated in Fig. 1(C), the surface and pores of CP become covered with  $\text{TiO}_2$  and  $\text{Fe}^0$  particles, leading to the formation of amorphous regions.

### 3.2. Morphological analysis

SEM images of CP,  $\text{TiO}_2$ , composites of  $\text{TiO}_2/\text{CP}$  and  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  are shown in Fig. 2. The CP exhibited a leaf-like structure with multiple layers or sheet morphology. A pure  $\text{TiO}_2$  particle has an irregular but somewhat spherical shape. The  $\text{TiO}_2$  particles are deposited on the surface of CP, as clearly demonstrated by SEM images of various  $\text{TiO}_2/\text{CP}$  composites.

Although  $\text{TiO}_2$  covers the surface of CP, it does not disrupt the structure; however,  $\text{TiO}_2$  covers the surface of CP and the sheet-like shape of CP that are clearly seen in the SEM images of all  $\text{TiO}_2/\text{CP}$  composites and confirms the intact structure of the composites after doping.<sup>5,7</sup>  $\text{Fe}^0$  particles exhibit a spherical shape and exist in chain form, which looks like spongy material.<sup>18,19</sup> The surface structure of CP in  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  composite looks highly rough and spongy due to the doping of  $\text{Fe}^0$ , as shown in Fig. 2. Due to its spongy nature, the  $\text{Fe}^0$  doping on the surface of  $\text{TiO}_2/\text{CP}$  composite covers the maximum area of CP zeolite. In terms of covering the maximum surface area of CP, the  $\text{Fe}^0$  doping decreases the adsorption capacity of the  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  composite due to the decreased average mesopore size of the composite, and this decreased efficiency is in good agreement with the decreased peak intensities obtained in the XRD pattern of the  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  composite. Meanwhile, the

EDXA spectrum of each sample shows that the main elements, such as Al, Si, and O, are spread almost throughout the CP structure, as illustrated in Fig. S1. In addition, the EDXA spectrum contains high concentrations of key elements, such as Al, Si, and O, along with metals like Na, K, Ca, and Mg, while the peaks for Ti and Fe metals confirm the successful doping of  $\text{TiO}_2$  and  $\text{Fe}^0$  in the CP composites.

### 3.3. Analysis of thermal properties

The TG and DTG curves of the composites heated in a temperature range of 0–900 °C are illustrated in Fig. 3. The TG/DTG analysis showed that CP,  $\text{TiO}_2$  and their composites are highly stable at higher temperatures up to 900 °C, as no changes occurred in the structure of any composites. However, a very slight weight loss occurred due to dehydration of water, but no other internal structural changes were observed.<sup>20,21</sup> The calculated Si/Al ratio of parent CP zeolite and all of its as-synthesized composites is around 4 that has re-confirmed their thermal stability at high temperatures and found in close agreement with the literature.<sup>21</sup> Si/Al ratio of CP is 4.76, which is confirmed by EDXA analysis, as shown in Fig. S1.

The TGA curve of  $\text{TiO}_2$  reveals a total weight loss of 2.5%, which can be attributed to a three-step dehydration process as evidenced by the three distinct peaks of DTG curve (Fig. 3). The first step, occurring between 46–91 and 39–226 °C in DTG and TG profiles, is characterized by a broad endothermic peak corresponds to a weight loss of 0.9%. This step is associated

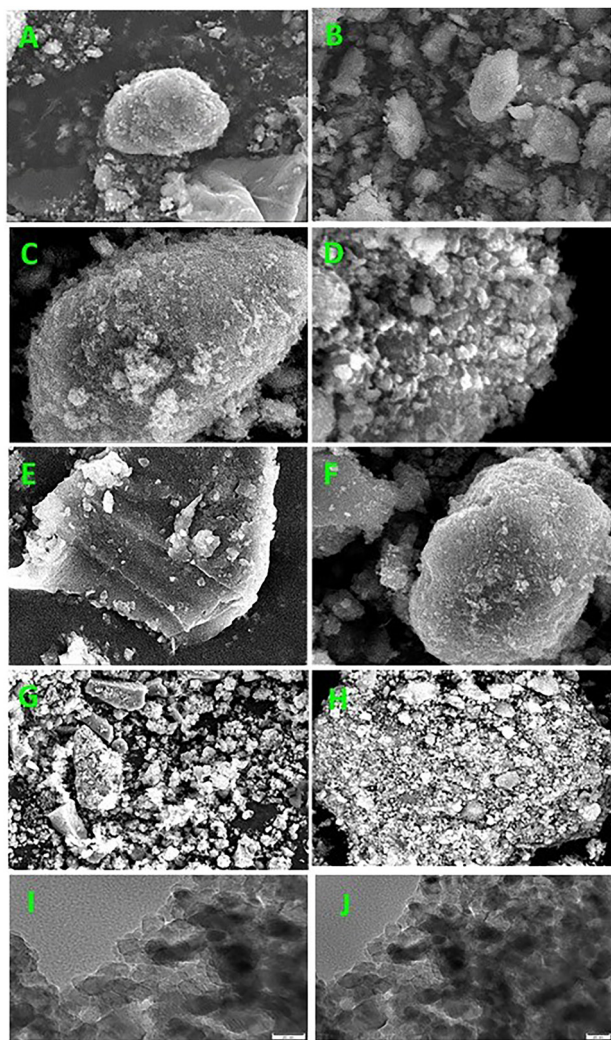


Fig. 2 SEM images of (A)  $\text{TiO}_2$  and (B) CP. Different molar composites: (C) 0.1, (D) 0.25, (E) 0.5, and (F) 1.0 M of  $\text{TiO}_2/\text{CP}$  (G)  $\text{Fe}^0\text{-TiO}_2$  and (H)  $\text{Fe}^0\text{-TiO}_2/\text{CP}$ , and (I) and (J) TEM of  $\text{TiO}_2$ .

with the desorption of loosely bound physisorbed water located on the external surface of  $\text{TiO}_2$ . The second step, observed between 196–273 and 226–381 °C of DTG and TG curves, exhibits a sharp endothermic peak and is linked to a weight loss of approximately 1.1%. This step is thought to be involved in the removal of tightly bound chemisorbed water from the framework structure of  $\text{TiO}_2$ . The third step, taking place over a higher temperature range of 381–900 °C of TG, is marked by a gradual weight loss of 0.5% and may be attributed to the condensation of  $\text{Ti}(\text{OH})_2$  to  $\text{TiO}_2$  in the  $\text{TiO}_2$  sample.<sup>22</sup>

In the TG curve of CP, a sharp decrease of 2.4% weight loss is observed due to the dehydration of physisorbed moisture from 40 to 225 °C, and a slow but continuous weight loss of 2.2% between 225 and 900 °C may be associated with the dehydration of water from internal cavities of the CP zeolite structure. The broad endothermic peak at 75–178 °C and a sharp endothermic peak at 224–272 °C in the DTG curve aid these two-step dehydration processes in CP zeolite.<sup>20</sup>

Total weight losses of about 6.0, 6.5, 6.3 and 5.6% in 0.1, 0.25, 0.5 and 1.0 M  $\text{TiO}_2/\text{CP}$  composites are calculated from the TG curves, respectively. The first steep TG curves show the loss of physisorbed water in the temperature range of 38–240 °C. The second slight TG curve shows the slow dehydration of water from the internal CP cavities in the temperature range of 240–900 °C. These two-step dehydration processes in different molar composites of  $\text{TiO}_2/\text{CP}$  are supported by the two endothermic peaks that appear in the DTG curves, as shown in Fig. 3.

The calculated weight losses of around 9.9 and 10.8% in  $\text{Fe}^0\text{-TiO}_2$  and  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  composites, respectively, are due to the evaporation of water and ethanol. A gradual and small increment in weight above 390 °C up to 760 °C is observed in both  $\text{Fe}^0\text{-TiO}_2$  and  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  composites possibly due to the formation of iron oxide at higher temperatures. This increment in the weight of  $\text{Fe}^0$ -doped composites is also confirmed by the small exothermic DTG curves, as shown in Fig. 3.

Alver *et al.* also reported that generally in all zeolites, the weight loss up to 500 °C is due to dehydration, while the slow weight loss at high temperature is due to dehydroxylation.<sup>20</sup> Generally, dehydroxylation is a slow process and occurs between 500 and 600 °C. At low temperatures, chemisorption took place between 28 and 41 °C on the surface of CP, as shown in Fig. 3. In the TG curves, a comparatively large peak is observed, leading to a weight gain of about 2% from 100 to 102% between 28 and 41 °C in almost all composites due to the absorption of water at low temperature. However, in  $\text{Fe}^0\text{-TiO}_2$  and  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  composites, a small peak is observed in the TG curves above 100%, indicating that the spongy structure of  $\text{Fe}^0$  decreases the adsorption capacity of the composites, which is why water is less adsorbed compared to other composites, reconfirming the above findings in SEM and XRD.

### 3.4. Surface area analysis

Table 1 and Fig. S2 present the BET surface area, micro, mesopore volume and average mesopore sizes of parent CP,  $\text{TiO}_2$ , composites of  $\text{TiO}_2/\text{CP}$  and  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  obtained from their  $\text{N}_2$  adsorption/desorption isotherms measured at 77 and 298 K. All the samples were degassed under helium gas at 150 °C/423 K for 6 h prior to  $\text{N}_2$  adsorption/desorption isotherms.

BET isotherms are used to investigate the porous structure of materials, and the form of the isotherm indicates the type of porosity in the material. The modifications made by doping of  $\text{TiO}_2$  and  $\text{Fe}^0$  have caused structural variations, which result in strong effects on the BET surface area, micro and mesopore volume and size in all types of  $\text{TiO}_2/\text{CP}$  composites due to the increasing molarity of  $\text{TiO}_2$ . Compared to the parent CP, the  $\text{TiO}_2$  doped CP samples exhibit significantly increased surface area, particularly in the 0.5 M  $\text{TiO}_2/\text{CP}$  composite ( $76.25 \text{ m}^2 \text{ g}^{-1}$ ). In all the  $\text{TiO}_2/\text{CP}$  and  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  composites, both the surface area and micropore volume increase compared to parent CP and pristine  $\text{TiO}_2$ , while the mesopore volume and average mesopore sizes decrease with increasing the molar ratio of  $\text{TiO}_2$  and with the co-doping of  $\text{Fe}^0$ . These changes are attributed to



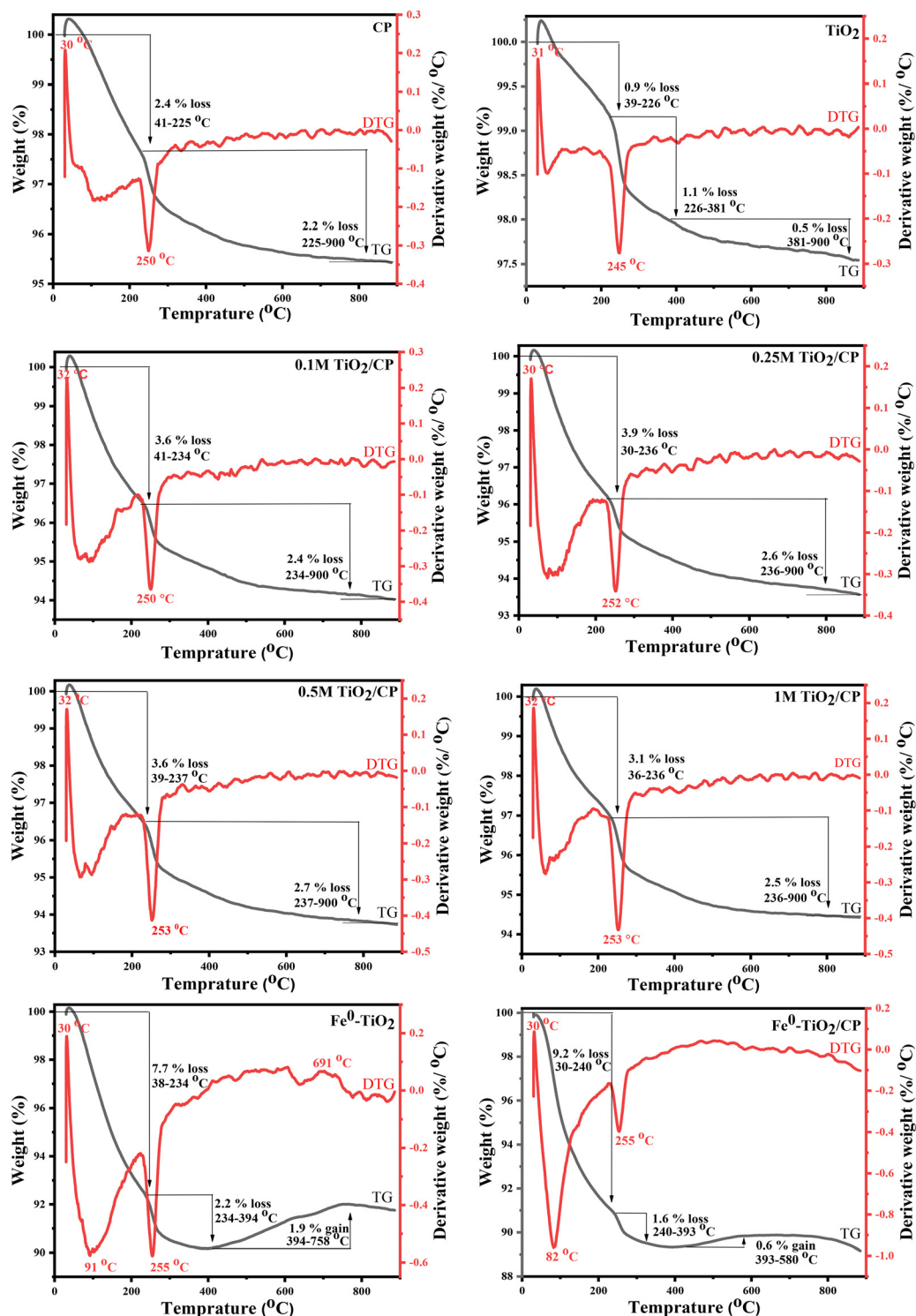


Fig. 3 TG/DTG profiles of pure CP,  $\text{TiO}_2$ , 0.1, 0.25, 0.5, 1.0 M  $\text{TiO}_2/\text{CP}$ , and  $\text{Fe}^0\text{-TiO}_2$  and  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  composites measured at a heating rate of  $10 \text{ K min}^{-1}$  under  $\text{N}_2$  flow.

the particle size of  $\text{TiO}_2$  ( $4.25 \mu\text{m}$ ) and  $\text{Fe}^0$  ( $0.90 \mu\text{m}$ ), as reported in the literature.<sup>18</sup>

The particle sizes of  $\text{TiO}_2$  and  $\text{Fe}^0$  are very small compared to the mesopore volume and size of the parent CP, where inside the mesopore both the  $\text{TiO}_2$  and  $\text{Fe}^0$  are adsorbed on the

surface of CP and cover the maximum surface area, which results in decreasing adsorption of CV dye and is found in close agreement with a previous study conducted by Panezai *et al.*<sup>23</sup>

As shown in Fig. S2, the profile for  $\text{TiO}_2$  demonstrates a clearly defined plateau at lower relative pressures, indicating



monolayer adsorption; as pressure increases, adsorption accelerates, indicating multilayer adsorption and capillary condensation at higher relative pressures. An obvious H3 hysteresis loop in the physisorption isotherm at higher relative pressures (0.4–0.1) is observed, which is a distinguishing feature of adsorption in mesoporous materials. The H3 type hysteresis loop according to the IUPAC classification observed in these composites is common for nonporous, macroporous, and mesoporous solids.<sup>23</sup> This is due to differences in the gas adsorption and desorption paths within the pores.<sup>24</sup> A steep desorption branch and a somewhat flat adsorption branch, with a hysteresis loop that closes at high relative pressure, are frequently connected with materials containing aggregation of platelet-like particles and slit-shaped pores.

The distinct increase in adsorbate volume in the low  $P/P_0$  region and a barely noticeable plateau, as well as an irregular shape, are observed in Type II and IV isotherms, revealing the presence of micropores associated with mesopores in the CP and its composites. An important feature is the closed hysteresis loop, with less steepness corresponding to a more uniform pore system containing capillaries with wider profile bodies and narrow, short necks. These phenomena are obviously illustrated in Table 1 and Fig. S2, where both the BET surface area and the relevant pore volume ( $0 < P/P_0 < 0.1$ ) of all the composites, with the exception of 0.1 M  $\text{TiO}_2/\text{CP}$  are higher than parent CP (19.16  $\text{m}^2 \text{g}^{-1}$  and 0.10  $\text{cm}^3 \text{g}^{-1}$ ). This increase in both the surface area and pore volume could be attributed to increased nitrogen adsorption into  $\text{TiO}_2$ ,  $\text{Fe}^0$  and particularly on their CP support.

### 3.5. Optical characterization

Fig. 4 presents the FTIR spectra of CP,  $\text{TiO}_2$  and their combined composites in the wave number range of 4000–400  $\text{cm}^{-1}$ . The intense peak at around 1020  $\text{cm}^{-1}$  is associated with the asymmetric stretching vibration of O–Si(Al)–O in the CP structure. As shown in Fig. 4(A)–(b), the characteristic peaks appeared at 786 and 459  $\text{cm}^{-1}$  in CP, and their composites are attributed to stretching vibrations of  $\text{AlO}_4$  and  $\text{SiO}_4$  tetrahedral atoms present in zeolite structures.<sup>7</sup> As shown from c to f in Fig. 4(A), the peaks present in CP zeolite at 1020, 786 and 459  $\text{cm}^{-1}$  decrease slightly due to the increasing molar ratio of  $\text{TiO}_2$ , and these peaks appear very weak in Fig. 4(A)–(g), (h) due to doping of two different materials (titania and  $\text{Fe}^0$  metal), which is found in good agreement with the XRD patterns.

The adsorption bands at 1600 and 1560  $\text{cm}^{-1}$  in pure  $\text{TiO}_2$  represent the bending vibrations of Ti–OH, as shown in Fig. 4(B), and the very small absorption bands generated at 419 and 408  $\text{cm}^{-1}$  in  $\text{TiO}_2$  and all molar composites of  $\text{TiO}_2/\text{CP}$  are attributed to the bending vibrations of Ti–O–Ti bonds, as shown in Fig. 4(C).<sup>18</sup>

### 3.6. Bandgap analysis

In order to investigate the band gap of CP,  $\text{TiO}_2$ ,  $\text{TiO}_2/\text{CP}$  and  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  composites, solid UV Vis spectroscopy absorption peaks were used with the Tauc relation, as mentioned in the following equation:

$$(Ah\nu)^n = B(h\nu - E_g), \quad (3)$$

where  $A$  is absorbance,  $h$  is Planck's constant,  $\nu$  is the speed of light,  $E_g$  is the band gap energy of material,  $B$  is the proportionality constant, and superscript  $n$  has a value of 2 for allowed direct transition and 1/2 for allowed indirect transition. The energy equation of quantum mechanics is as follows:

$$E = \frac{h\nu}{\lambda}, \quad (4)$$

where energy ( $E$ ) represents the band gap, Planck's constant ( $h$ ) is  $6.626 \times 10^{-34}$  Joules per second, velocity of light ( $\nu$ ) is  $2.99 \times 10^8$  meters per second, and wavelength ( $\lambda$ ) is the absorption peak value. The band gaps of the CP,  $\text{TiO}_2$  and  $\text{Fe}^0$ -doped  $\text{TiO}_2$  and  $\text{TiO}_2/\text{CP}$  composites are determined by plotting  $(Ah\nu)^{1/2}$  against energy eV ( $h\nu/\lambda$ ) for the indirect band gap and  $(Ah\nu)^2$  against energy eV ( $h\nu/\lambda$ ) for the direct band gap.<sup>25</sup> The linear part of the curve  $(Ah\nu)^n$  against energy eV is extrapolated (Fig. S3), and the band gap energy is listed in Table 1.

In general, it is observed that rutile phase  $\text{TiO}_2$  has less photocatalytic property compared to the anatase phase. However, anatase phase  $\text{TiO}_2$  has a bandgap of 3.11 eV, which is best suited under UV radiation. Many studies have been conducted in order to tune and decrease the bandgap of  $\text{TiO}_2$  and shift its efficiency towards solar radiation by various metals and non-metal co-doping.<sup>12,18</sup> The bandgap in  $\text{TiO}_2$ -doped CP zeolite composites decreases with the increasing surface area. It can be observed in Table 1 and Fig. S3 that the bandgap of  $\text{TiO}_2$  decreased from 3.11 eV to 2.98 eV in the 0.1 M  $\text{TiO}_2/\text{CP}$  composite. There is a very slight decrease observed in the bandgap of 0.25, 0.5 and 1 M composites; this indicates that if the molar ratio of  $\text{TiO}_2$  increases, then the amount (1 g) of the CP zeolite is not enough to decrease the bandgap.<sup>26</sup>

Table 1 BET surface area, pore volume and size and band gap of different photocatalysts

| S. no | Sample                               | Surface area ( $\text{m}^2 \text{g}^{-1}$ ) | Mesopore volume ( $\text{cm}^3 \text{g}^{-1}$ ) | Average mesopore size (nm) | Micropore volume ( $\text{cm}^3 \text{g}^{-1}$ ) | Median micropore size (nm) | Direct bandgap (eV) | Indirect bandgap (eV) |
|-------|--------------------------------------|---------------------------------------------|-------------------------------------------------|----------------------------|--------------------------------------------------|----------------------------|---------------------|-----------------------|
| 1     | CP                                   | 19.16                                       | 0.10                                            | 19.87                      | 0.007                                            | 0.72                       | 3.90                | 3.78                  |
| 2     | $\text{TiO}_2$                       | 40.22                                       | 0.41                                            | 38.81                      | 0.015                                            | 0.84                       | 3.11                | 2.96                  |
| 3     | 0.1 M $\text{TiO}_2/\text{CP}$       | 38.58                                       | 0.23                                            | 23.15                      | 0.015                                            | 0.88                       | 2.98                | 2.62                  |
| 4     | 0.25 $\text{TiO}_2/\text{CP}$        | 53.53                                       | 0.28                                            | 19.79                      | 0.020                                            | 0.89                       | 3.02                | 2.66                  |
| 5     | 0.5 $\text{TiO}_2/\text{CP}$         | 76.25                                       | 0.36                                            | 18.11                      | 0.028                                            | 0.86                       | 3.10                | 2.95                  |
| 6     | 1.0 $\text{TiO}_2/\text{CP}$         | 71.61                                       | 0.39                                            | 20.58                      | 0.028                                            | 0.88                       | 3.01                | 2.91                  |
| 7     | $\text{Fe}^0\text{-TiO}_2$           | 55.79                                       | 0.21                                            | 14.60                      | 0.021                                            | 0.85                       | 1.91                | 1.39                  |
| 8     | $\text{Fe}^0\text{-TiO}_2/\text{CP}$ | 69.73                                       | 0.26                                            | 13.95                      | 0.026                                            | 0.86                       | 2.24                | 1.10                  |



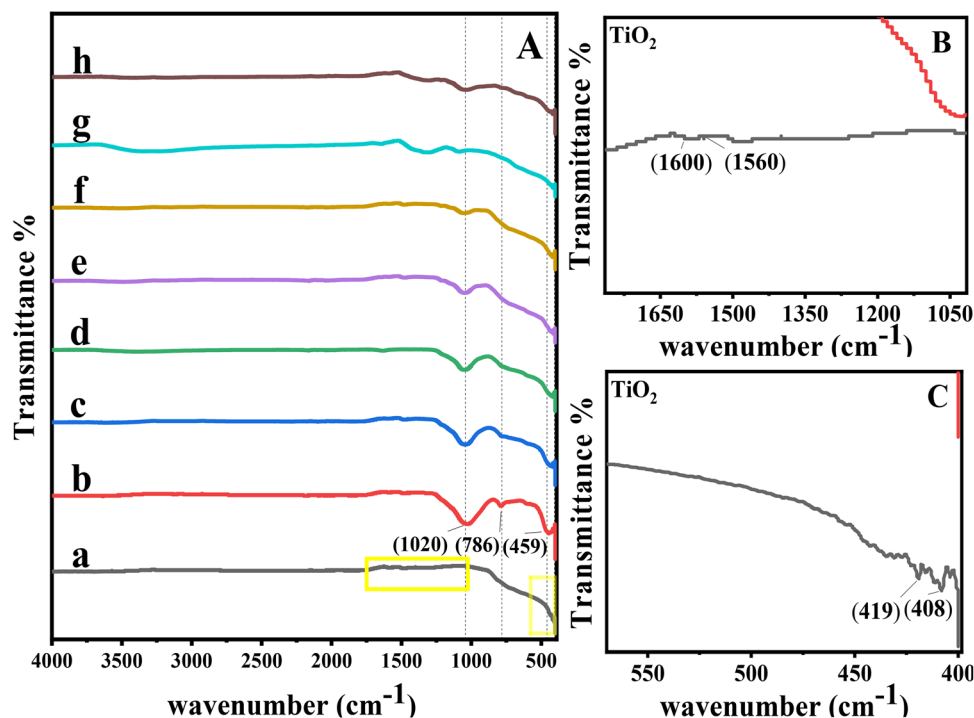


Fig. 4 FTIR spectra of A: (a)  $\text{TiO}_2$ , (b) parent CP, (c) 0.1, (d) 0.25, (e) 0.5, (f) 1.0 M  $\text{TiO}_2/\text{CP}$ , (g)  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  and (h)  $\text{Fe}^0\text{-TiO}_2$ . B and C are the insets of A.

$\text{TiO}_2$ -supported clinoptilolite ( $\text{TiO}_2/\text{CP}$ ) composites can reduce  $\text{TiO}_2$ 's electron-hole pair recombination rate *via* numerous mechanisms:

- **Electron trap:** the clinoptilolite zeolite structure can serve as an electron trap, trapping electrons and preventing their recombination with holes. Zeolites can function as electron sinks, delaying recombination and increasing the lifespan of reactive charge carriers, which improves photocatalytic activity.

- **Charge carrier separation:** the interaction between  $\text{TiO}_2$  and CP allows for charge carrier separation. The CP structure can take electrons from  $\text{TiO}_2$ , while the nanoparticles can deliver holes, minimizing recombination and enhancing photocatalytic performance. Zeolite-supported  $\text{TiO}_2$  composites promote the generation of reactive species like hydroxyl radicals ( $\cdot\text{OH}$ ), which are necessary for the breakdown of organic contaminants.

- **Better light absorption:** particularly in the UV-Vis region, the unique structure of the CP can improve light absorption because of its high surface area and pore volume, which can scatter light in a way that lengthens the optical path within the composite. This scattering effect can increase the absorption of light by  $\text{TiO}_2$  nanoparticles, which can increase the generation of electron-hole pairs and improve photocatalytic activity.<sup>27</sup>

- $\text{Fe}^0$  acts as a new conduction band, lies below the conduction band of  $\text{TiO}_2$  and reduces the bandgap difference of  $\text{TiO}_2$ , which previously worked efficiently only under UV radiation. The electron present in the valence band of  $\text{TiO}_2$  can now easily jump to the new conduction band, *i.e.*  $\text{Fe}^0$ , under solar radiation, as shown Fig. 5.<sup>28</sup>

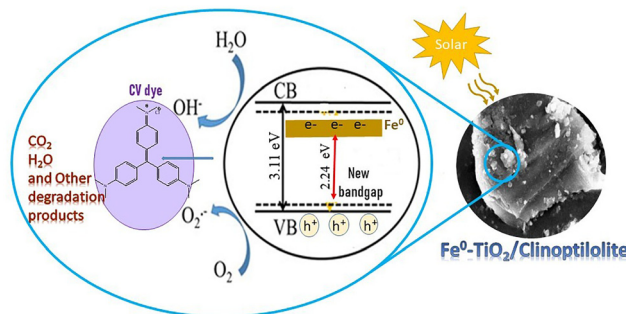


Fig. 5 Reduction in the bandgap of  $\text{TiO}_2/\text{CP}$  by  $\text{Fe}^0$  doping on  $\text{TiO}_2/\text{CP}$ .

### 3.7. Application of $\text{TiO}_2/\text{CP}$ photocatalyst

**a. Removal of CV dye from aqueous medium *via* adsorption and photocatalytic degradation.** In a heterogeneous photocatalysis, the first process is adsorption, which takes place in the dark, and then, the degradation process starts upon exposure of the solution containing dyes and photocatalyst to radiation. The percentage of dye removed from an aqueous solution is not significantly affected by adsorption alone. A substantial portion of the dye in an aqueous solution is removed using the synthesized  $\text{TiO}_2/\text{CP}$  photocatalyst through a combined effect of adsorption and photocatalytic degradation.

A calibration curve for CV dye was constructed by plotting absorbance against concentration (in ppm), revealing a linear relationship, as shown in Fig. S4 Chart A. The resulting analytical equation and correlation coefficient ( $R^2$ ) were derived

and utilized for data analysis, enabling the calculation of the removal percentage and adsorption capacity:

$$y = 0.0776x - 0.0096, \quad (5)$$

$$R^2 = 0.9956. \quad (6)$$

**b. Effect of the calcination temperature on the removal efficiency of the photocatalyst.** Calcination significantly influences the phase, crystallinity, and properties of TiO<sub>2</sub>, with anatase being the dominant phase at both 500 and 600 °C. At 500 °C, TiO<sub>2</sub> nanoparticles exhibit improved crystallinity compared to lower temperatures, but some amorphous content may remain, potentially reducing degradation efficiency. In contrast, calcination at 600 °C enhances crystallization and particle growth, resulting in higher crystallinity and larger particle sizes, as confirmed by TEM analysis. The TEM images, as shown in Fig. 2(I) and (J), reveal nearly spherical nanoparticles consistent with the anatase phase.<sup>29–31</sup> Anatase TiO<sub>2</sub>, with its tetragonal structure and 3.11 eV band gap,<sup>32,33</sup> is particularly suitable for photocatalysis due to its stability up to 712 °C.<sup>34</sup> As the primary catalyst, TiO<sub>2</sub> plays a crucial role in the photocatalytic degradation of CV dye, leveraging its crystalline structure and properties to facilitate efficient degradation.

TiO<sub>2</sub>-500 °C has a band gap energy of 3.30 eV.<sup>35</sup> This high band energy is not suitable for the degradation of CV dye under solar radiation. The TiO<sub>2</sub> synthesized *via* the sol-gel method was calcined at two different temperatures (500 and 600 °C). The photocatalytic performance of TiO<sub>2</sub> calcined at 500 and 600 °C (TiO<sub>2</sub>-500 °C and TiO<sub>2</sub>-600 °C) temperatures was investigated, and the results summarized in Fig. S4 Chart B revealed that TiO<sub>2</sub>-600 °C has a high removal efficiency of CV dye compared to TiO<sub>2</sub>-500 °C.

**c. Effect of the concentration of Titania in TiO<sub>2</sub>/CP on the removal of dye.** The adsorptive and degradation photocatalytic efficiency of different as-synthesized photocatalysts were comparatively investigated. The bare anatase phase TiO<sub>2</sub> has low adsorption activity and needs to be combined with materials with a high surface area and adsorption capacity. The CP zeolite played a positive role, as an adsorbent, for CV dye degradation, but bare CP exhibited no degradation properties. The combined composite of TiO<sub>2</sub> and CP enhances the adsorption and degradation properties of each other and gives a high removal% of CV dye. Table 2 shows that among the four different molarities (0.1, 0.25, 0.5 and 1.0 M) of TiCl<sub>4</sub> in TiO<sub>2</sub>/CP composites, 0.25 M TiO<sub>2</sub>/CP performed the highest degradation of CV dye in aqueous solution. However, when the concentration of titania increased, it occupied the maximum surface of CP zeolite and led to a reduction in the adsorption capacity of the composite for CV dye.

**d. Effect of pH on reaction media.** The TiO<sub>2</sub>/CP composites showed the highest degradation efficiency of CV dye at pH 6.<sup>7</sup> At first instance, the photo-catalysis process was carried out without adjusting the pH of the solution, which led to less removal of CV dye from aqueous solution. Then, the same process was repeated in duplicate, with the pH adjusted to 6

**Table 2** Removal% and adsorption capacity of CV dye using different photocatalysts

| S. no | Photocatalyst/sample                  | Radiation source | Removal% | Adsorption capacity |
|-------|---------------------------------------|------------------|----------|---------------------|
| 1     | CP                                    | UV               | 51.19    | 13.82               |
| 2     | TiO <sub>2</sub> -500 °C              | UV               | 39.29    | 20.22               |
| 3     | TiO <sub>2</sub> -600 °C              | UV               | 58.67    | 24.13               |
| 4     | TiO <sub>2</sub> -600 °C              | Solar            | 48.45    | 10.47               |
| 5     | 0.1 M TiO <sub>2</sub> /CP            | UV               | 90.00    | 21.51               |
| 6     | 0.25 TiO <sub>2</sub> /CP             | UV               | 92.36    | 22.08               |
| 7     | 0.5 TiO <sub>2</sub> /CP              | UV               | 88.06    | 21.05               |
| 8     | 1.0 TiO <sub>2</sub> /CP              | UV               | 81.87    | 19.57               |
| 9     | 0.1 TiO <sub>2</sub> /CP              | Solar            | 88.52    | 21.16               |
| 10    | 0.25 TiO <sub>2</sub> /CP             | Solar            | 84.24    | 20.14               |
| 11    | 0.5 TiO <sub>2</sub> /CP              | Solar            | 86.99    | 20.79               |
| 12    | 1.0 TiO <sub>2</sub> /CP              | Solar            | 71.40    | 17.07               |
| 13    | Fe <sup>0</sup> -TiO <sub>2</sub> /CP | Solar            | 77.89    | 17.91               |
| 14    | Fe <sup>0</sup> -TiO <sub>2</sub> /CP | UV               | 80.19    | 12.38               |
| 15    | Fe <sup>0</sup> -TiO <sub>2</sub>     | Solar            | 76.01    | 16.37               |
| 16    | Fe <sup>0</sup> -TiO <sub>2</sub>     | UV               | 84.77    | 16.89               |

using a 1% HCl solution. The results summarized in Table S1 illustrate that all the TiO<sub>2</sub>/CP composites give maximum removal efficiency of CV dye in the pH range of 5.5–6.5.

**e. Effect of time interval of reaction.** As shown in Fig. 6(A), the pure parent CP gives total adsorption of 13.8% and total removal of 51% in 120 minutes. The bar of TiO<sub>2</sub> (Fig. 6(A) and (B)) gives the maximum 58 and 48% removal of CV dye in 120 minutes under UV and solar radiations, respectively. The 0.25 M TiO<sub>2</sub>/CP composite, as depicted in Fig. 6(A), performed the highest removal efficiency of about 92.3% in 120 minutes under UV radiation. Additionally, 0.1 M TiO<sub>2</sub>/CP, as shown in Fig. 6(B), revealed the highest CV dye removal of 88.5% under solar radiation in 120 minutes.

The UV-vis spectra of CV dye degradation using various photocatalysts show a decrease in absorbance over time, indicating a reduction in dye concentration (Fig. 6(C) and (D) and Fig. S5).

As irradiation time increases, the characteristic peak's absorbance decreases, and the peak shifts downward, signifying enhanced removal efficiency. After 120 minutes, the peaks nearly flatten, suggesting that optimal removal is achieved within this timeframe.

**f. Effect of the type of radiation on degradation.** The percentage degradation rates of TiO<sub>2</sub>, CP, and various molar composites of TiO<sub>2</sub>-CP were investigated under two sources, such as UV and solar irradiation. Table 2 shows the removal efficiencies and adsorption capacities. The results demonstrate that the degradation of CV dye under a UV source is much faster than solar, as shown in Fig. S4 Chart C, and TiO<sub>2</sub> exhibits minimal activity in the presence of visible light compared to UV rays, which is found in close agreement with the literature.<sup>36</sup> This occurs due to the TiO<sub>2</sub>'s broad band gap, which limits its potential but can be increased through surface modification techniques.

In order to improve TiO<sub>2</sub> efficiency under solar radiation, the strategy used was to synthesize the Fe<sup>0</sup>-TiO<sub>2</sub>/CP composite. The co-doping of Fe<sup>0</sup> with TiO<sub>2</sub> reduces the band gap of TiO<sub>2</sub>, allowing it to operate efficiently under solar radiation. The degradation efficiency of Fe<sup>0</sup>-TiO<sub>2</sub>/CP composite under solar



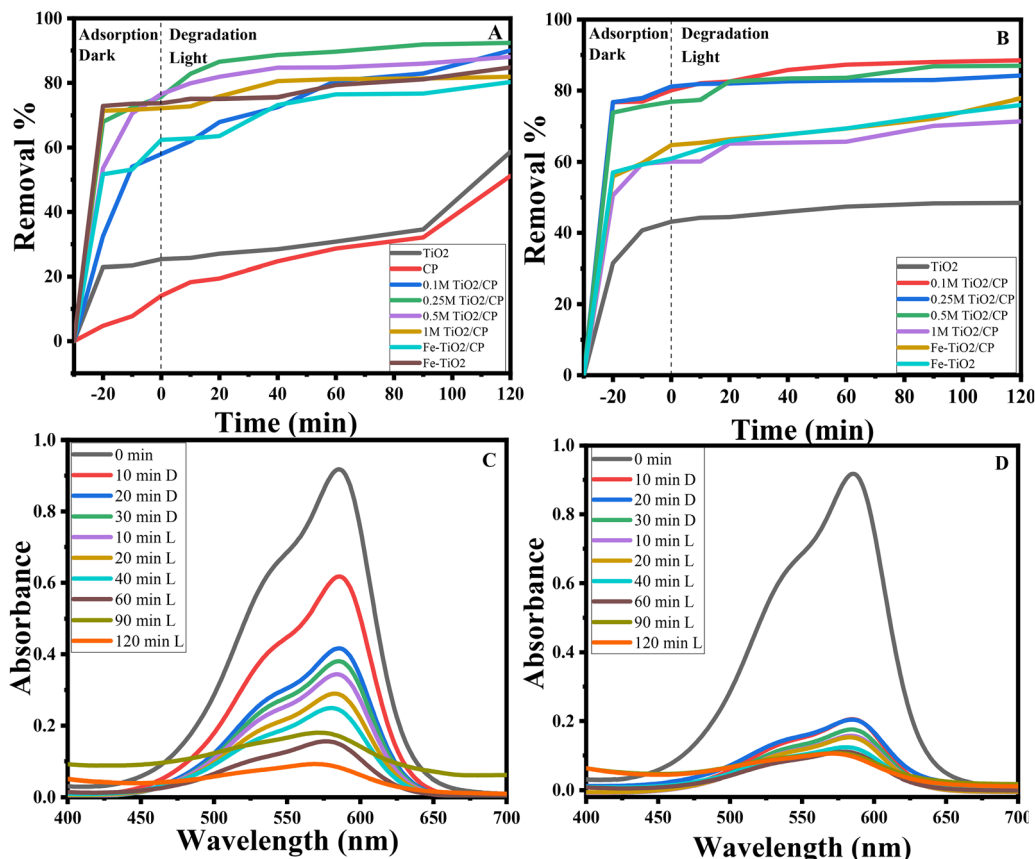


Fig. 6 Removal percent vs. time in UV (A) and solar radiations (B). UV-Visible spectra of CV dye in UV (C) and solar radiations (D) D = dark and L = light.

radiation was not enhanced after doping with  $\text{Fe}^0$ . As shown in Table 2, the maximum degradation of  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  composite is 77.8%, which is less than 0.1 M  $\text{TiO}_2/\text{CP}$  with 88.5% removal of CV dye under solar radiation in 120 minutes. However, the degradation efficiency of bare  $\text{TiO}_2$  from  $\text{TiO}_2\text{-600 } ^\circ\text{C}$  (48%) to  $\text{Fe}^0\text{-TiO}_2$  (76%) under solar radiation is increased, as shown in Table 2. The  $\text{Fe}^0$  doping did not considerably increase the degradation efficiency of  $\text{TiO}_2/\text{CP}$ , which is confirmed by the XRD patterns, SEM images, TG/DTG analysis and BET analysis. As the zerovalent iron particles do exist in chain like structure, which have a sponge-like morphology and cover the external surface area of the composite, it results in a decreased average pore size of the  $\text{Fe}^0\text{-TiO}_2/\text{CP}$ , as shown in Table 1. The BET analysis showed that the average pore size of 0.1 M  $\text{TiO}_2/\text{CP}$  (23.1 nm) decreased in the  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  (13.9 nm) composite, leading to less adsorption of CV dye. This was also confirmed by the results reported in Table 2, which showed that 0.1 M  $\text{TiO}_2/\text{CP}$  composite showed 21 and 21.5  $\text{mg g}^{-1}$  of maximum adsorption capacity under UV and solar radiations, respectively, while the same composite with  $\text{Fe}^0$  doped ( $\text{Fe}^0\text{-TiO}_2/\text{CP}$ ) offered maximum adsorption capacities of 17 and 12  $\text{mg g}^{-1}$  under UV and solar radiations, respectively.

### 3.8. Kinetic study of adsorption and photocatalytic degradation of CV dye

The adsorption process of CV dye on the surface of the photocatalyst is dependent on both the amount of CV dye in aqueous

solution and the available active sites on titania-supported clinoptilolite composites. Adsorption kinetics studies were carried out using pseudo-first order (PFO) and pseudo-second order (PSO) models. The adsorption process of CV dye was found to follow the PSO model, which explains why adsorption is a two-site occupancy. The following two equations are applied to the adsorption process:

$$\ln(q_e - q_t) = \ln q_e - k_1 t, \quad (7)$$

$$\frac{t}{q_t} = \frac{1}{k_2 \cdot q_e^2} + \frac{t}{q_e}, \quad (8)$$

where  $q_e$  and  $q_t$  are the adsorbed amount of dye at equilibrium and time  $t$ , respectively, and  $k_1$  and  $k_2$  are the PFO and PSO constants, respectively.<sup>7</sup>

For PSO, the  $R^2$  values are higher than those for PFO, as shown in Table 3. The adsorption of CV dye on all  $\text{TiO}_2/\text{CP}$  composites showed the best fit to PSO. As depicted in Fig. 7(A) and (B), the experimental curve is a straight line plotted with PSO and PFO, respectively, for the adsorption of CV dye on a 0.1 M  $\text{TiO}_2/\text{CP}$  composite.  $R^2$  values for PSO is 0.9685 and for PFO is 0.6973, confirming that the adsorption process best fits the PSO kinetic model.

The degradation process follows the PFO model, in which the reaction depends on the initial concentration of the dye in the solution. In the process of CV dye degradation,  $\cdot\text{OH}$  and



Table 3 Kinetics study of the adsorption and degradation of CV dye

| S. no | Photocatalyst/sample                  | Radiation source used | $q_{\text{e,exp}}$ | Adsorption          |       |       |                     |       |       | Degradation      |       |
|-------|---------------------------------------|-----------------------|--------------------|---------------------|-------|-------|---------------------|-------|-------|------------------|-------|
|       |                                       |                       |                    | Pseudo 1st order    |       |       | Pseudo 2nd order    |       |       | Pseudo 1st order |       |
|       |                                       |                       |                    | $q_{\text{e,calc}}$ | $k_1$ | $R^2$ | $q_{\text{e,calc}}$ | $k_2$ | $R^2$ | $K$              | $R^2$ |
| 1     | CP                                    | UV                    | 3.81               | 5.37                | 0.06  | 0.88  | 6.04                | 0.02  | 0.69  | —                | —     |
| 3     | TiO <sub>2</sub> -600 °C              | UV                    | 10.43              | 5.03                | 0.07  | 0.59  | 10.36               | 0.16  | 0.99  | 0.004            | 0.77  |
| 4     | TiO <sub>2</sub> -600 °C              | Solar                 | 9.32               | 6.48                | 0.08  | 0.73  | 9.57                | 0.06  | 0.99  | 0.0009           | 0.92  |
| 5     | 0.1 M TiO <sub>2</sub> /CP            | UV                    | 19.86              | 9.13                | 0.10  | 0.69  | 16.98               | 0.03  | 0.99  | 0.011            | 0.98  |
| 6     | 0.25 TiO <sub>2</sub> /CP             | UV                    | 20.44              | 8.10                | 0.09  | 0.62  | 19.19               | 0.13  | 0.99  | 0.008            | 0.86  |
| 7     | 0.5 TiO <sub>2</sub> /CP              | UV                    | 19.99              | 11.30               | 0.09  | 0.80  | 19.38               | 0.04  | 0.99  | 0.005            | 0.89  |
| 8     | 1.0 TiO <sub>2</sub> /CP              | UV                    | 14.08              | 6.06                | 0.09  | 0.53  | 15.95               | 0.04  | 0.98  | 0.003            | 0.80  |
| 9     | 0.1 TiO <sub>2</sub> /CP              | Solar                 | 19.13              | 5.45                | 0.08  | 0.42  | 19.19               | 0.30  | 0.99  | 0.005            | 0.89  |
| 10    | 0.25 TiO <sub>2</sub> /CP             | Solar                 | 19.39              | 11.05               | 0.09  | 0.81  | 19.05               | 0.04  | 0.99  | 0.001            | 0.92  |
| 11    | 0.5 TiO <sub>2</sub> /CP              | Solar                 | 18.39              | 9.09                | 0.09  | 0.69  | 18.45               | 0.05  | 0.99  | 0.005            | 0.87  |
| 12    | 1.0 TiO <sub>2</sub> /CP              | Solar                 | 14.54              | 7.45                | 0.09  | 0.62  | 14.59               | 0.09  | 0.99  | 0.001            | 0.94  |
| 13    | Fe <sup>0</sup> -TiO <sub>2</sub> /CP | Solar                 | 13.17              | 5.71                | 0.08  | 0.55  | 13.97               | 0.15  | 0.99  | 0.004            | 0.93  |
| 14    | Fe <sup>0</sup> -TiO <sub>2</sub> /CP | UV                    | 9.63               | 6.17                | 0.07  | 0.78  | 9.49                | 0.09  | 0.98  | 0.006            | 0.91  |

D = dark and L = light.

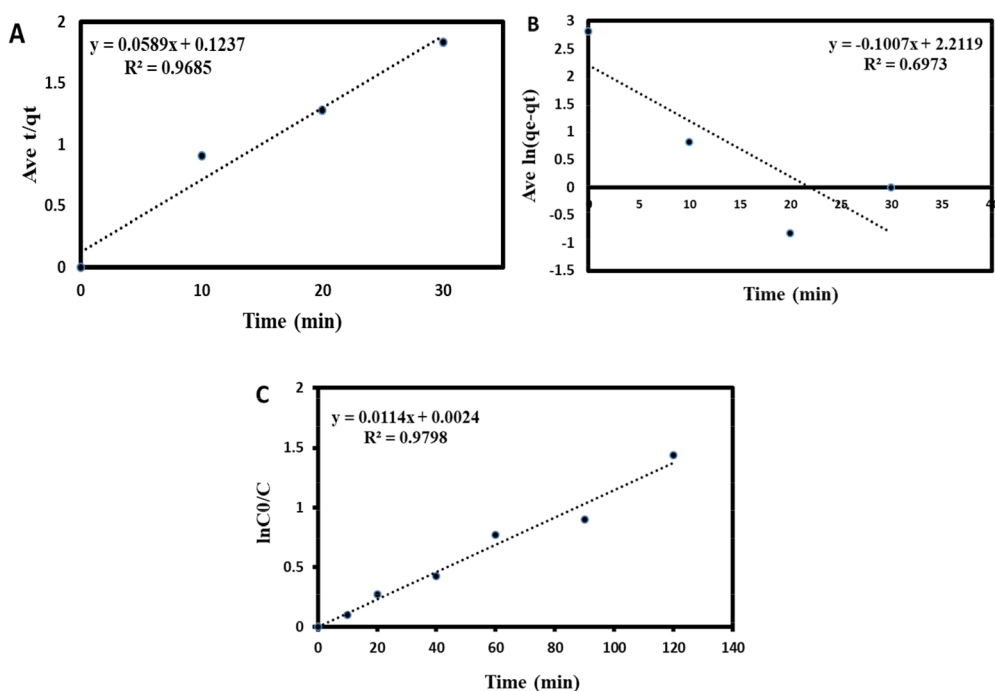


Fig. 7 (A) Adsorption study of CV dye using pseudo-second-order model and (B) pseudo-first-order model, and degradation kinetics of CV dye using (C) pseudo-first-order model.

$\text{O}_2^{\bullet-}$  are continuously generated by UV and solar radiations falling on the photo-catalyst containing semiconductor TiO<sub>2</sub>. These radicals are in excess, and the reaction rate depends merely on the dye concentration. The following equation is applied for degradation:

$$\ln\left(\frac{C_0}{C}\right) = k \cdot t, \quad (9)$$

where  $C_0$  and  $C$  are the concentrations of dye at time 0 and time  $t$ , respectively,  $k$  is the rate constant and  $t$  is the interval time.<sup>7</sup>

### 3.9. Scavenger study using EDTA, IPA and *p*-BQ

The photocatalytic mechanisms of TiO<sub>2</sub>/CP and Fe<sup>0</sup>-TiO<sub>2</sub>/CP composites were investigated through free radical scavenger experiments to identify the key active species involved in the degradation of CV dye. In each experiment, 10 mL of 0.1 M solutions of specific scavengers, including isopropyl alcohol (IPA) to capture hydroxyl radicals ( $\bullet\text{OH}$ ), ethylenediamine tetraacetic acid (EDTA) to capture holes ( $\text{h}^+$ ), and benzoquinone (BQ) to capture superoxide radicals ( $\bullet\text{O}_2^-$ ), was added to the CV dye solution before photodegradation.<sup>37</sup> The results showed that the addition of BQ moderately reduced the degradation of



## SCAVENGER STUDY

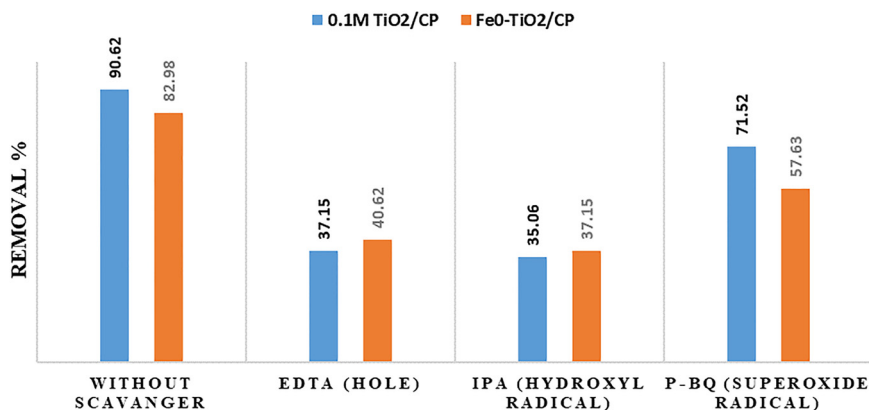


Fig. 8 Scavenger study using EDTA, IPA and *p*-BQ.

CV dye, indicating that  $\bullet\text{O}_2^-$  plays an important role in both reaction systems. In contrast, the addition of EDTA and IPA significantly reduced the degradation rate of CV dye, by 37 and 40% for EDTA and by 35 and 37% for IPA in the  $\text{TiO}_2/\text{CP}$  and  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  systems, respectively. This suggests that holes and  $\bullet\text{OH}$  are crucial for the photocatalytic process. Overall, Fig. 8 illustrates that  $\text{h}^+$ ,  $\bullet\text{OH}$ , and  $\bullet\text{O}_2^-$  are the main active radicals responsible for the degradation of CV dye in both photocatalytic systems.

The migration direction of photogenerated charge carriers in a composite is closely related to the band edge positions of the semiconductors. To determine these positions, the following formulas are used:

$$E_{\text{CB}} = X - E_{\text{C}} - 0.5 E_{\text{g}} \quad (10)$$

for the conduction band ( $E_{\text{CB}}$ ) edge potential,

$$E_{\text{VB}} = E_{\text{CB}} + E_{\text{g}} \quad (11)$$

for the valence band ( $E_{\text{VB}}$ ) edge potential.

Here,  $X$  represents the absolute electronegativity of the semiconductor (5.81 eV), and  $E_{\text{C}}$  is the energy of free electrons on the hydrogen scale, which is 4.50 eV. Using these formulas, the calculated edge potentials for  $\text{TiO}_2/\text{CP}$  are 2.8 eV for the valence band ( $E_{\text{VB}}$ ) and  $-0.18$  eV for the conduction band ( $E_{\text{CB}}$ ).<sup>37,38</sup>

**a. Photocatalytic mechanism of  $\text{TiO}_2/\text{CP}$ .** According to the traditional photocatalytic degradation mechanism, when  $\text{TiO}_2$  in the  $\text{TiO}_2/\text{CP}$  composite is exposed to light irradiation, it generates photoelectrons that jump from the valence band (VB) to the conduction band (CB). The holes ( $\text{h}^+$ ) in the VB of  $\text{TiO}_2$ , with an edge potential of 2.8 eV, can react with  $\text{H}_2\text{O}$  through an oxidation process to produce hydroxyl radicals ( $\bullet\text{OH}$ ), given that the VB edge potential is more positive than the  $\text{H}_2\text{O}/\text{OH}$  potential (2.40 eV vs. NHE).<sup>38</sup> However, the electrons in the CB of  $\text{TiO}_2$ , with an edge potential of  $-0.18$  eV, are not sufficiently negative to reduce  $\text{O}_2$  and form superoxide radicals ( $\bullet\text{O}_2^-$ ) since the CB edge potential is more positive than the  $\text{O}_2/\bullet\text{O}_2^-$  potential ( $-0.046$  eV vs. NHE).<sup>2</sup> As a result, the  $\text{TiO}_2/\text{CP}$

photocatalyst does not produce a significant amount of  $\bullet\text{O}_2^-$  to oxidize CV dye, which is consistent with the findings from the free radical scavenger experiments. This analysis supports the proposed photocatalytic mechanism for the  $\text{TiO}_2/\text{CP}$  composite.

**b. Photocatalytic mechanism of  $\text{Fe}^0\text{-TiO}_2/\text{CP}$ .** The XRD pattern of the used  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  sample, as shown in Fig. 1(B)-(g), clearly exhibits a peak corresponding to  $\text{Fe}^0$  at a  $2\theta$  degree of  $44^\circ$ . This confirms the presence of  $\text{Fe}^0$  in the composite. According to the proposed mechanism illustrated in Fig. 5, when  $\text{TiO}_2$  is irradiated, it generates photoelectrons ( $\text{e}^-$ ) and holes ( $\text{h}^+$ ). The electrons in the conduction band of  $\text{TiO}_2$  are then transferred to  $\text{Fe}^0$ , which acts as an electronic medium. This electron transfer is facilitated by the electron trap constructed by  $\text{Fe}^0$ , allowing for the spatial isolation of photo-generated electron-hole pairs. As a result, the undesirable recombination of electrons and holes is greatly limited, enhancing the photocatalytic efficiency of the  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  composite.

In the  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  composite, the electrons associated with  $\text{Fe}^0$ , which have a reduction potential of approximately  $-0.41$  to  $-0.44$  V,<sup>39,40</sup> can combine with oxygen in water to form superoxide radicals ( $\bullet\text{O}_2^-$ ). The potential for forming superoxide radicals is  $-0.046$  eV vs. NHE, and these radicals can mineralize CV dye. Simultaneously, the holes in the valence band of  $\text{TiO}_2$ , with an edge potential of 2.8 eV, can directly react with CV dye to mineralize it. Furthermore, the holes in the  $\text{TiO}_2$  valence band can react with water to form hydroxyl radicals ( $\bullet\text{OH}$ ) due to the valence band potential being higher than the  $\text{H}_2\text{O}/\bullet\text{OH}$  potential (2.40 eV vs. NHE). These hydroxyl radicals can then oxidize CV dye into degradation products. This proposed mechanism is consistent with the results of free radical scavenger experiments, which identify the reactive species involved in the photocatalytic degradation process using the  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  composite.

The formation of superoxide radical anions ( $\bullet\text{O}_2^-$ ) occurs when electrons in the CB of  $\text{TiO}_2$  or  $\text{Fe}^0$  combine with oxygen through a reduction process. A key reason why the  $\text{Fe}^0\text{-TiO}_2/\text{CP}$  system produces more superoxide radicals than the  $\text{TiO}_2/\text{CP}$



system lies in the difference in their reduction potentials. Specifically, the conduction band potential of  $\text{TiO}_2$  ( $-0.18$  eV) is not sufficiently negative to reduce oxygen molecules to form superoxide radicals, given that the  $\text{O}_2/\bullet\text{O}_2^-$  potential is  $-0.046$  eV vs. NHE. In contrast, the reduction potential of  $\text{Fe}^0$  ( $-0.41$  to  $-0.44$  V)<sup>39,40</sup> is more negative than the  $\text{O}_2/\bullet\text{O}_2^-$  potential, allowing it to easily reduce oxygen molecules and produce superoxide radicals. The redox potentials of the conduction band and valence band of  $\text{TiO}_2/\text{CP}$  are  $-0.18$  and  $2.8$  eV, respectively, indicating that  $\text{TiO}_2/\text{CP}$  acts as a mild reducing and strong oxidizing photocatalyst.<sup>41</sup>

### 3.10. Degradation pathway and identification of intermediates using H-ESI-MS analysis

The various intermediates formed during the mineralization of CV dye by  $\text{Fe}^0$ - $\text{TiO}_2/\text{CP}$  and  $0.1$  M  $\text{TiO}_2/\text{CP}$  photocatalysis are shown in Fig. S6. The reaction intermediates were examined using the H-ESI-MS technique. The results of the obtained H-ESI mass spectra are summarized in Table 4, and the mechanism is shown in Fig. 9, Scheme 1. Nineteen intermediates were identified.

Route 1: the attack of radical species on the CV dye molecule is a crucial step in the *N*-de-methylation pathway. This process begins with a radical species, such as  $\bullet\text{OH}$ , abstracting a hydrogen atom from a methyl substituent of the amino group, which forms a carbon-centered radical. Depending on the specific reaction conditions, other radical species such as superoxide ( $\text{O}_2\bullet^-$ ) radicals may also be involved in the attack on the dimethylamine group. This carbon-centered radical is highly reactive and can undergo further reactions. Another radical species then attacks the dimethylamine group, leading to the formation of a mono-de-methylated intermediate.

After the second radical attack, the intermediate molecule remains positively charged due to the delocalized positive charge over the conjugated triphenylmethane system of the original CV molecule. Although a neutral methyl group is removed during demethylation, the positive charge persists

albeit with a slightly altered distribution. This positively charged intermediate can still be adsorbed on the photocatalyst surface, where further stepwise demethylation occurs.<sup>42</sup>

The cleavage of the C–N bond during the second radical attack is a heterolytic process. The first attack involves homolytic cleavage, in which a radical abstracts a hydrogen atom from a methyl group, leaving a carbon-centered radical intermediate. The second radical species then interacts with this intermediate, triggering a rearrangement in which the C–N bond is broken unevenly.

The nitrogen atom retains both electrons from the bond, while the carbon-centered radical remains a smaller neutral molecule, such as formaldehyde. This heterolytic cleavage results in a final, less-methylated cationic intermediate and a neutral byproduct, with the positive charge remaining on the larger triphenylmethane structure.<sup>43</sup>

The *N*-de-methylated intermediates of CV were identified using H-ESI mass spectrometry. The molecular ion peaks, as shown in Fig. S6 and listed in Table 4, confirmed the presence of various intermediates. These species correspond to three pairs of isomeric molecules with two to four fewer methyl groups than CV. For instance, **B** is formed by the removal of a methyl group from the CV molecule. In the first pair of isomers, **C** is formed by the removal of two methyl groups from two sides of the CV molecule, while **D** is produced by the removal of two methyl groups from the same side of the CV structure.

In the second pair of isomers, **E** is formed by the removal of three methyl groups from each side of the CV molecule, while **F** is produced by the removal of two methyl groups from one side and one methyl group from the other side. In the third pair of isomers, **H** is formed by the removal of two methyl groups from two sides of the CV molecule, while **G** is produced by the removal of two methyl groups from the same side and one methyl group from the remaining two sides. Further, *N*-de-methylation leads to the formation of **I** and **J**. The *N*-de-methylated intermediates (**A–J**) are shown in Fig. 9, Scheme 1

Table 4 Degradation intermediates of CV dye

| H-ESI-MS peaks | Intermediates                                                                | ESI-MS spectrum ions ( <i>m/z</i> ) |
|----------------|------------------------------------------------------------------------------|-------------------------------------|
| <b>A</b>       | <i>N,N,N',N',N'',N''</i> -Hexamethylpararosanine CV                          | 372.42                              |
| <b>B</b>       | <i>N,N</i> -Dimethyl- <i>N',N'</i> -dimethyl- <i>N''</i> -methylpararosanine | 358.50                              |
| <b>C</b>       | <i>N,N</i> -Dimethyl- <i>N'</i> -methyl- <i>N''</i> -methylpararosanine      | 344.67                              |
| <b>D</b>       | <i>N,N</i> -Dimethyl- <i>N',N'</i> -methylpararosanine                       | 344.33                              |
| <b>E</b>       | <i>N</i> -Methyl- <i>N'</i> -methyl- <i>N''</i> -methylpararosanine          | 332.83                              |
| <b>F</b>       | <i>N,N</i> -Dimethyl- <i>N'</i> -methylpararosanine                          | 330.67                              |
| <b>G</b>       | <i>N</i> -Methyl- <i>N'</i> -methylpararosanine                              | 316.75                              |
| <b>H</b>       | <i>N,N</i> -Dimethylpararosanine                                             | 316.50                              |
| <b>I</b>       | <i>N</i> -Methylpararosanine                                                 | 302.25                              |
| <b>J</b>       | Pararosanine                                                                 | 288.83                              |
| <b>K</b>       | 4-( <i>N,N</i> -dimethylamino)phenol                                         | 138.08                              |
| <b>L</b>       | 4-( <i>N,N</i> -Dimethylamino)-4-( <i>N,N</i> -dimethylamino)benzophenone    | 269.92                              |
| <b>M</b>       | 4-( <i>N</i> -Methylamino)phenol                                             | 124.08                              |
| <b>N</b>       | 4-aminophenol                                                                | 110.83                              |
| <b>O</b>       | 4-( <i>N,N</i> -Dimethylamino)-4-( <i>N</i> -methylamino)benzophenone        | 255.00                              |
| <b>P</b>       | 4-( <i>N,N</i> -Dimethylamino)-4-aminobenzophenone                           | 240.75                              |
| <b>Q</b>       | 4-( <i>N</i> -Methylamino)-4-( <i>N</i> -methylamino)benzophenone            | 240.67                              |
| <b>R</b>       | 4-( <i>N</i> -Methylamino)-4-aminobenzophenone                               | 226.00                              |
| <b>S</b>       | 4,4-Bis-aminobenzophenone                                                    | 213.83                              |



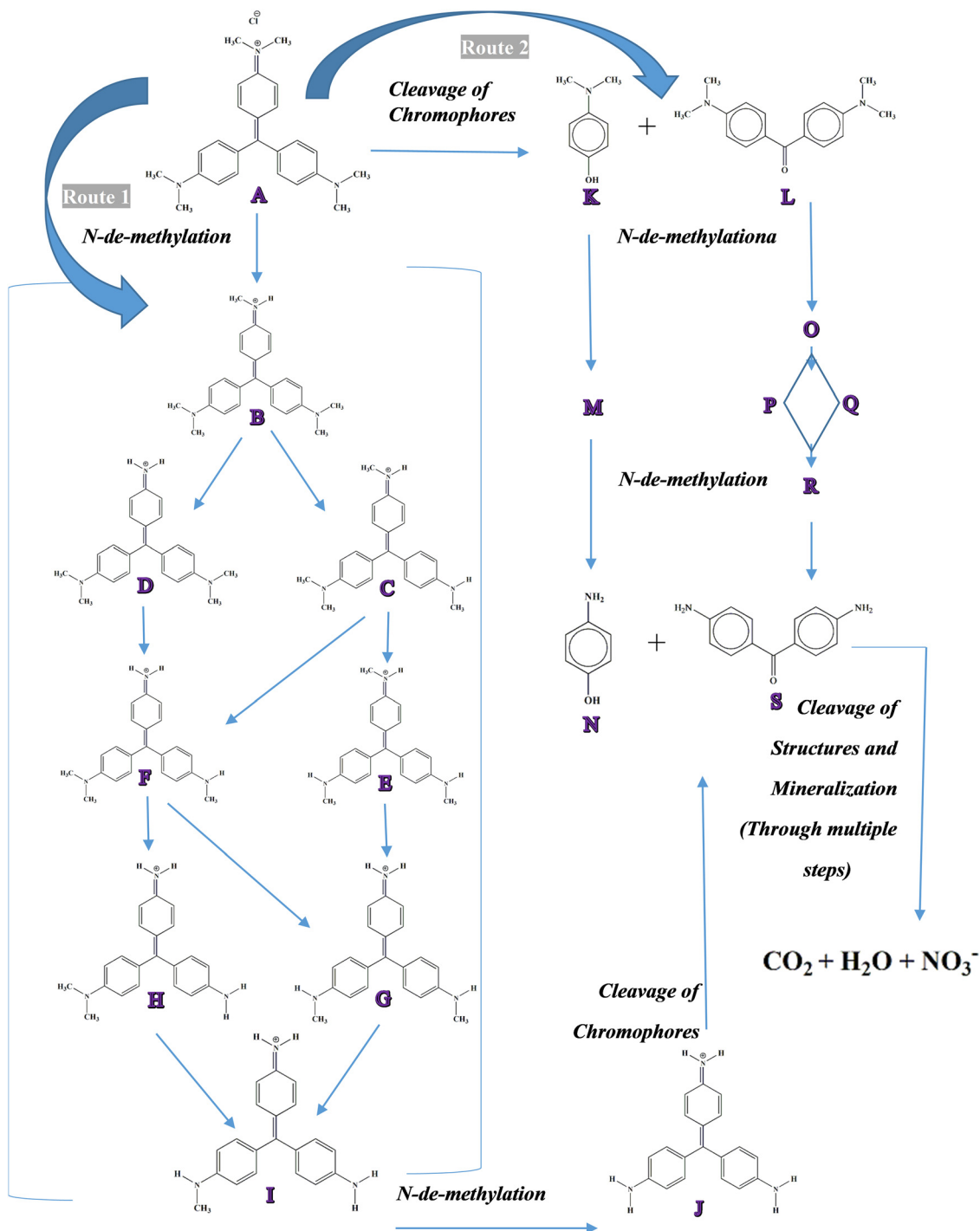


Fig. 9 Scheme 1: Degradation pathways of CV dye.

(Route 1). After that, the direct attack of the ( $\bullet\text{OH}$ ) radical on the central carbon of J produces N and S through chromophore cleavage.<sup>44</sup>

Route 2: however, the degradation of CV is initiated by the highly reactive hydroxyl radical ( $\bullet\text{OH}$ ), which attacks the central carbon atom of the molecule. This central carbon is part of the extensively conjugated triphenylmethane chromophore responsible for CV's deep purple color. The attack likely occurs

due to the electrophilic nature of the central carbon, making it susceptible to nucleophilic species like the hydroxyl radical or radical attack, leading to bond breaking. The hydroxyl radical's attack initiates the cleavage of a carbon-carbon bond connecting one of the *N,N*-dimethylaminophenyl rings to the central carbon atom, directly breaking the extended conjugation system of the CV molecule. This bond breaking results in the formation of intermediates, including 4-(dimethylamino)



phenol (**K**), which is formed from the cleavage of one of the *N,N*-dimethylaminophenyl groups. Another intermediate, 4-(*N,N*-dimethylamino)-4-(*N,N*-dimethylamino)benzo-phenone (**L**), also known as Michler's ketone, is formed from the remaining two *N,N*-dimethylaminophenyl groups still attached to the central carbon, which is now part of a ketone structure after bond rearrangement and oxygen incorporation. Afterwards, compounds **N** and **S** were possibly formed by a series of *N*-de-methylated intermediates in a stepwise manner from **K** and **L**, respectively, as shown in Fig. 9, Scheme 1 (Route 2). The degradation process does not necessarily stop at these intermediates. They can undergo further reactions, such as *N*-demethylation, ring-opening, and eventual mineralization to simpler, less toxic compounds, like CO<sub>2</sub> and H<sub>2</sub>O, or other intermediates, like carboxylic acids, depending on the specific reaction conditions and the presence of other oxidizing agents.<sup>45,46</sup>

## 4. Conclusion

In order to enhance the removal efficiency of CP as support and synthetic TiO<sub>2</sub>, the TiO<sub>2</sub>/CP, Fe<sup>0</sup>-TiO<sub>2</sub> and Fe<sup>0</sup>-TiO<sub>2</sub>/CP composites were successfully synthesized *via* sol-gel and borohydride reduction methods. The characteristic peaks of CP at  $2\theta = 9.8^\circ$ ,  $22.1^\circ$ ,  $26.4^\circ$ ,  $30.1^\circ$ , and  $35.7^\circ$  revealed that the CP crystalline structure remained intact in TiO<sub>2</sub>/CP after TiO<sub>2</sub> doping. The sheet-like/multi-layers/plate-like morphology of CP was unaffected under a high calcination temperature of 600 °C, confirming that it is thermally stable in all TiO<sub>2</sub>/CP composites. The SEM and BET results showed that Fe<sup>0</sup> particles with a spherical shape and spongy nature covered the maximum surface area of CP zeolite after doping on TiO<sub>2</sub>/CP composites. Furthermore, the Fe<sup>0</sup> doping reduces the adsorption capacity and removal efficiency in Fe<sup>0</sup>-TiO<sub>2</sub>/CP composite by reducing the composite's average mesopore size from TiO<sub>2</sub>/CP (23.1 nm) to Fe<sup>0</sup>-TiO<sub>2</sub>/CP (13.9 nm), and this decreased efficiency is also reconfirmed by the decreased peak intensities of Fe<sup>0</sup>-TiO<sub>2</sub>/CP composite's in XRD patterns.

According to TG/DTG profiles, the CP, TiO<sub>2</sub> and their composites showed thermal stability at higher temperatures up to 900 °C. Dehydration of water caused a very minor loss of weight, but no additional structural or internal changes were noticed. The TiO<sub>2</sub> removed the maximum amount of CV dye (58 and 48%) under UV and solar irradiations. In addition, a maximum 92% removal of CV dye was achieved by 0.25 MTiO<sub>2</sub>/CP composite under UV radiation compared to solar radiation, whereas a maximum CV dye removal of 88% was achieved by 0.1 M TiO<sub>2</sub>/CP under solar radiation. Furthermore, the band gap of 0.1 M TiO<sub>2</sub>/CP was significantly reduced, which in turn showed higher removal efficiency in solar radiations.

Scavenger study showed that the TiO<sub>2</sub>/CP composite primarily generates •OH radicals in CV dye degradation, while the Fe<sup>0</sup>-TiO<sub>2</sub>/CP composite produces both •OH and •O<sub>2</sub><sup>-</sup> radicals, which have played a vital role in enhancing photocatalytic efficiency due to the favorable reduction potential of Fe<sup>0</sup>.

The degradation ability of Fe<sup>0</sup>-TiO<sub>2</sub>/CP composites to produce •O<sub>2</sub><sup>-</sup> radicals significantly contributes to enhancing their efficiency.

The degradation of CV occurs through two routes: route 1 involves stepwise *N*-de-methylation *via* radical attacks, breaking C–N bonds and forming less-methylated intermediates, while Route 2 involves hydroxyl radical attack on the central carbon, cleaving C–C bonds and disrupting the conjugated chromophore structure. Both routes lead to the breakdown of the CV dye molecule. In conclusion, using CP as a support material for TiO<sub>2</sub> is a significant option because CP is neutral, abundant in nature, and inexpensive, making TiO<sub>2</sub>/CP composites an eco-friendly photocatalyst that offers a promising solution for the removal of CV dye.

## Author contributions

Nazia Aziz: writing – original draft, investigation, methodology, validation, visualization, data curation, formal analysis. Hamida Panezai: writing – review & editing, visualization, supervision, project administration, conceptualization, resources. Jihong Sun: validation, resources. Noor Samad Shah: Validation, resources. Raza Ullah: data analysis, validation. Ruohan Xu: Resources. Zakira Jomezai: formal analysis.

## Conflicts of interest

There are no conflicts to declare.

## Data availability

The data supporting this article are summarized in this manuscript and are included in the Supplementary information (SI). Supplementary information is available. See DOI: <https://doi.org/10.1039/d5ma00658a>.

Raw data are available upon request from the corresponding author.

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