

## EDGE ARTICLE

[View Article Online](#)  
[View Journal](#) | [View Issue](#)
Cite this: *Chem. Sci.*, 2018, **9**, 3313

Received 15th December 2017  
 Accepted 20th February 2018

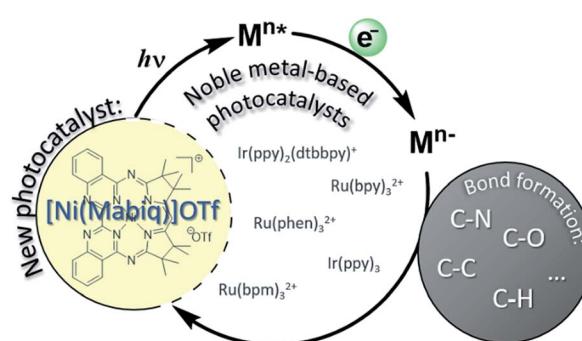
DOI: 10.1039/c7sc05320g  
[rsc.li/chemical-science](http://rsc.li/chemical-science)

## Introduction

Photoredox catalysis offers nascent opportunities to shift conventional chemical production methods to light-driven processes.<sup>1</sup> The design of new photoactive metal compounds is key to the development of new catalytic transformations. Chromophores can effect a wide array of functional group transformations and C–C bond formation reactions *via* the generation of active radical species, originating from excited state electron transfer processes (Scheme 1).<sup>2</sup> With very few exceptions, photoredox catalysis relies on the use of noble metal containing photosensitizers, mainly Ru- or Ir-polypyridyl complexes;<sup>3</sup> other heavy metal complexes (*e.g.* Os<sup>II</sup>, Re<sup>I</sup>, Mo<sup>0</sup>, W<sup>0</sup>) also have occasionally been employed.<sup>4,5</sup> [Ru(bpy)<sub>3</sub>]<sup>2+</sup> – with its long lived, charge separated excited state ([Ru<sup>III</sup>(bpy)<sup>–</sup>](bpy)<sub>2</sub>)<sup>2+\*</sup> – is the classic and universal photocatalyst for a plethora of applications.<sup>2</sup>

The use of less expensive, more abundant, late first row transition metal alternatives is generally precluded by their

inherently short excited-state lifetimes. In the first row, nickel complexes have been used in photocatalytic cross coupling reactions,<sup>6</sup> though the use of an added photosensitizer (commonly Ir) still is required in all of these tandem systems. A class of Ni<sup>II</sup> ligand-to-ligand charge transfer complexes recently were shown to possess advantageous properties as photosensitizers, but applications have not yet been demonstrated.<sup>7</sup> In fact, reports describing catalytic applications or reactivity of systems using only non-noble metal photosensitizers are exceedingly scarce.<sup>8–10</sup>



Department of Chemistry, Catalysis Research Center (CRC), Technische Universität München, 85747 Garching, Germany. E-mail: [thorsten.bach@ch.tum.de](mailto:thorsten.bach@ch.tum.de); [corinna.hess@ch.tum.de](mailto:corinna.hess@ch.tum.de)

† Dedicated to the memory of Inge Grübel.

‡ Electronic supplementary information (ESI) available: CIF files, synthetic procedures, spectroscopic, electrochemical and DFT computational data. CCDC 1538675, 1538676 and 1586948. For ESI and crystallographic data in CIF or other electronic format see DOI: [10.1039/c7sc05320g](https://doi.org/10.1039/c7sc05320g)

We now report a Ni<sup>II</sup> complex with a macrocyclic biquinazoline (Mabiq) ligand,<sup>11,12</sup> [Ni(Mabiq)]OTf (1), providing a rare example of a non-noble metal based photosensitizer. The divalent complex 1 can be photochemically reduced and the electronic structure of the product, [Ni(Mabiq)] (2), is described herein. Using a series of synthesized sacrificial donor molecules, the photo-excited state redox potential of 1 was assessed and compared to common noble-metal photocatalysts. We demonstrate the photoredox catalytic ability of 1 in a radical-based cyclization of a bromoalkyl-substituted indole. The reaction relies solely on the Ni-Mabiq photocatalyst, without the need for an additional noble metal photosensitizer.

## Results and discussion

The yellow diamagnetic 1 (Scheme 2) was readily prepared by complexation of the ligand with Ni(OTf)<sub>2</sub> in ethanol solution (see ESI† for further details). The Ni ion adopts the expected square planar geometry in the solid state (Fig. S1†). The electronic spectrum of 1 in DCM exhibits intense absorption bands ( $\epsilon \approx 10^4 \text{ M}^{-1} \text{ cm}^{-1}$ ) in the visible light region ( $\lambda_{\text{max}} = 414, 435$ , and  $457 \text{ nm}$ ;  $\epsilon = 13.6, 16.4, 22.3 \times 10^3 \text{ M}^{-1} \text{ cm}^{-1}$ ; Fig. 2 inset). Similar features were observed in the spectra of other M-Mabiq complexes, notably in the spectrum of Zn(Mabiq)Cl.<sup>12a</sup> The Zn<sup>II</sup>-complex displayed two strong absorption bands at 471 and 502 nm that were assigned as Mabiq  $\pi \rightarrow \pi^*$  transitions. However, d-d or metal-to-ligand charge transfer (MLCT) processes may additionally contribute to the corresponding absorptions of 1. The cyclic voltammogram of 1 in MeCN (Fig. S9†) exhibits a reversible, formally Ni<sup>II/I</sup> redox couple at  $-1.05 \text{ V}$  vs. Fc<sup>+/-</sup> (Fc = ferrocene; Fc<sup>+/-</sup> = 0.4 V vs. SCE). Additional, seemingly reversible, reductive processes appear at potentials  $<-1.5 \text{ V}$ .

The one-electron reduced Ni(Mabiq) (2) was subsequently generated from 1 using CoCp<sub>2</sub> as the reductant (Scheme 2). The molecular structure of 2 (Fig. 1) reveals shorter Ni–N bond distances (Ni–N<sub>avg</sub> = 1.874 Å vs. Ni–N<sub>avg</sub> = 1.882 Å in 1), as well as the hallmark changes in the diketiminato C–N bonds that signify reduction of the Mabiq ligand (Table S8†). The  $S = 1/2$  ground state of the complex was verified by EPR spectroscopy. The spectrum is consistent with a ligand-centered radical, with  $g_{\text{iso}} = 1.995$  (Fig. S4†). Low energy features at  $\lambda_{\text{max}} = 641, 711$  and  $801 \text{ nm}$  ( $1.4, 3.5, 5.4 \times 10^3 \text{ M}^{-1} \text{ cm}^{-1}$ ) are apparent in the electronic spectrum of 2 in THF (Fig. 2 inset), accounting for the

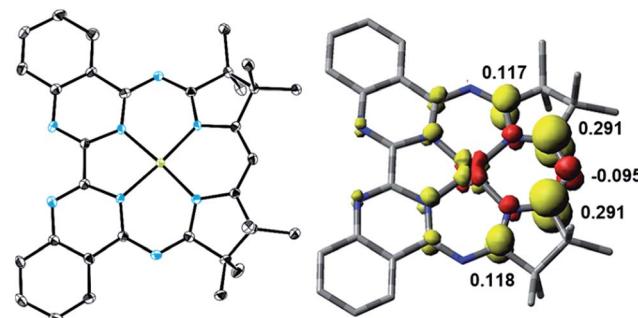


Fig. 1 Left: Molecular structure of 2 (50% probability ellipsoids; hydrogen atoms omitted for clarity). Right: DFT-derived (B3LYP) spin density plot for 2 based on Löwdin population analysis (isosurface value =  $\pm 0.005$ ).

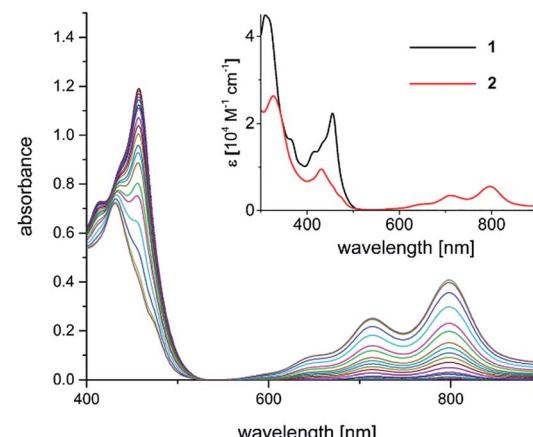
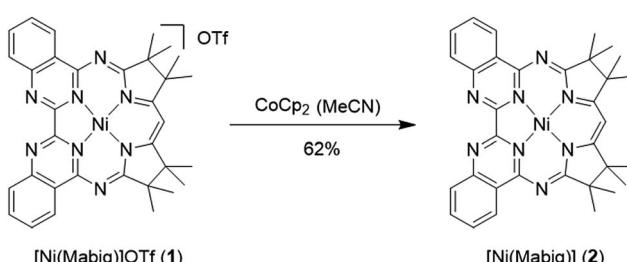


Fig. 2 Spectral evolution during photoconversion of 1 to 2 [ $c(1) = 0.05 \text{ mM}$ ;  $c(\text{NEt}_3) = 1.4 \text{ M}$ ;  $\lambda = 457 \text{ nm}$ , DMF]. Inset: electronic spectra of 1 (black trace;  $\text{CH}_2\text{Cl}_2$ ) and 2 (red trace; THF).

vibrant green color of the complex in solution. The spectrum again closely resembles that of the one-electron reduced Zn complex, Zn<sup>II</sup>(Mabiq<sup>·-</sup>).<sup>12a</sup> The spectroscopic data thus point to ligand-centered reduction of 1, such that the electronic structure of 2 corresponds to Ni<sup>II</sup>(Mabiq<sup>·-</sup>). Indeed, DFT calculations (B3LYP) on 2 further support this conclusion.

The DFT-derived (B3LYP) spin density plot (Fig. 1) describes a diamagnetic d<sup>8</sup> Ni<sup>II</sup> center with an unpaired electron localized primarily on the diketiminato unit of the Mabiq ligand. Four doubly occupied d-orbitals can be identified, while the SOMO possesses only *ca.* 4% d-orbital character and otherwise depicts a ligand  $\pi^*$  orbital (Fig. S5†). The latter molecular orbital is antibonding with respect to the diketiminato C–N p-orbital interactions, which explains the lengthening of these bonds in the structure of 2. It is noteworthy that ligand-centered reduction appears to prevail across the series of metal-Mabiq compounds we have examined thus far.<sup>12</sup>

The well-behaved redox chemistry of [Ni(Mabiq)]OTf (1), its high absorbance in the visible region and its relatively high reduction potential warranted a study of its photoredox properties. As mentioned above, the compound exhibits a strong



Scheme 2 Reaction of 1 with CoCp<sub>2</sub> (Cp = cyclopentadienyl) yields the one-electron reduced 2.



multi-structured absorption band with a maximum at  $\lambda = 457$  nm ( $\epsilon = 22\,300\text{ M}^{-1}\text{ cm}^{-1}$ ), which invites excitation with a visible light source and quenching studies with a suitable reductant. Gratifyingly, it was indeed found that irradiation of a DMF solution of **1** at  $\lambda = 457$  nm, in the presence of  $\text{NEt}_3$  (7.5 mM to 1.4 M), leads to a color change from yellow to green and to the formation of complex **2**.

The formation of the reduced compound was complete after 15 minutes (using 1.4 M  $\text{NEt}_3$ ;  $c$  (**1**) = 0.05 mM), as verified spectroscopically (Fig. 2). The photoconversion of **1** to **2** occurs on a much faster timescale in MeCN/THF or DMF/THF mixtures (the solvent combination solubilizes both forms), under identical conditions. The quantum yield for the photoconversion is  $\approx 10^{-4}$  (THF : DMF 4 : 1), and correlates with the  $\text{Et}_3\text{N}$  concentration (14–56 mM; Table S3 and Fig. S29†). From the data, the lifetime of the excited state form that reacts with the  $\text{Et}_3\text{N}$  can be estimated as  $\approx 1 \times 10^{-8}$  s (see ESI† for details). Steady-state emission spectra recorded at ambient temperature and at 77 K did not reveal any luminescence. Thus, we currently cannot comment in detail on the nature of the excited state processes involved in the photo-reduction of **1**. If one takes the longest wavelength absorption [ $\lambda \approx 510$  nm,  $E_0 \leq 235\text{ kJ mol}^{-1}$  (2.4 V)] of compound **1** to estimate the redox potential of photoexcited complex **1\***, a value of  $\leq +1.35$  V (vs.  $\text{Fc}^{+/0}$ ) is obtained.<sup>13</sup>

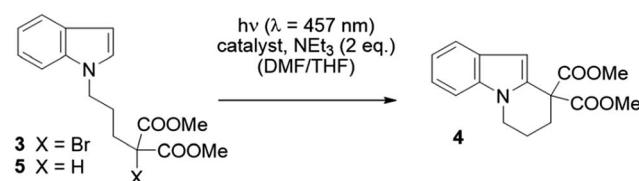
TDDFT (B3LYP) computational studies provide some insight into the nature of the absorptions in the visible region. The calculated transitions correlate well with the experimentally obtained absorbance spectrum of **1** (Fig. S6†). The absorptions at 400–500 nm include a prominent LL/CT transition that corresponds to the HOMO to LUMO transition. The HOMO is localized on the bipyrimidine moiety of the Mabiq ligand, while the LUMO is a diketiminato based  $\pi^*$  orbital (Fig. S7†). Other, less intense, transitions in the vicinity possess d–d ( $\text{Ni d}_{z^2} \rightarrow \text{Ni}$

$\text{d}_{x^2-y^2}$ ) and MLCT ( $\text{Ni d}_{z^2} \rightarrow \text{L}_{\pi^*}$ ) character. These states may contribute to the unique photochemical properties of **1**. However, a detailed investigation regarding the photochemistry and excited state kinetics of this compound is warranted, and will be the subject of future investigations.

We examined whether the photochemical properties of **1** might render it a suitable photoredox catalyst for C–C bond formation reactions. The radical cyclization of the *N*-( $\omega$ -bromoalkyl)-substituted indole **3** was chosen as a test reaction.<sup>14</sup> The reaction had been previously studied by the Stephenson group and was found to produce mainly product **4** by C–C bond formation under optimized conditions.<sup>15</sup> Under non-optimized conditions, hydro-de-bromination was a competing side reaction and varying product ratios of **4** and **5** were observed. Optimal conditions were reported to include the use of  $[\text{Ru}(\text{bpy})_3]\text{Cl}_2$  as the catalyst (1 mol%) and  $\text{NEt}_3$  (2 equiv.) in DMF solution and gave product **4** in 60% yield.<sup>15</sup>

Given the limited solubility of **2** in DMF, the reaction was initially attempted in a DMF/THF mixture (v/v = 1/2) with 2 mol% of catalyst **1** and 2 equiv.  $\text{NEt}_3$  as the quencher (Table 1, entry 1).<sup>16</sup> We were pleased to find that the desired cyclization proceeded smoothly and delivered with high chemoselectivity the desired product **4**. The inseparable hydro-de-brominated byproduct **5** was detectable in minor quantities but the ratio of products was 95/5 in favor of cyclization product **4**. When increasing the relative volume of THF in the solvent mixture both conversion and yield improved slightly (entry 2). The selectivity towards the desired reaction was high with a yield of 86% at 94% conversion, *i.e.* 91% yield based on conversion. For comparison, the  $\text{Ru}^{II}$  complex  $[\text{Ru}(\text{bpy})_3](\text{PF}_6)_2$  was employed under identical conditions (entry 3). Although the ratio **4**/**5** was identical with this catalyst, the reaction suffered from a lower conversion and a lower chemoselectivity (63% yield based on conversion). Similar observations were made when the catalyst

**Table 1** Photoredox-catalyzed cyclization of bromide **3** to tricyclic product **4** and reduction to hydro-de-brominated product **5**; influence of the catalyst and the reaction parameters on the yield and chemoselectivity



Entry <sup>a</sup>	Catalyst <sup>a</sup>	mol%	DMF/THF [v/v]	Conv. <sup>b</sup> [%]	Yield <sup>c</sup> [%]	<b>4/5</b> <sup>d</sup>
1	<b>1</b>	2	1/2	93	84	95/5
2	<b>1</b>	2	1/4	94	86	95/5
3	$[\text{Ru}(\text{bpy})_3](\text{PF}_6)_2$	2	1/4	73	46	95/5
4	<b>1</b>	1	1/4	95	84	95/5
5	$[\text{Ru}(\text{bpy})_3](\text{PF}_6)_2$	1	1/4	59	23	95/5
6	<b>1<sup>e</sup></b>	2	1/4	n.d.	—	—
7	— <sup>f</sup>	—	1/4	12	12	55/45
8	<b>1<sup>g</sup></b>	2	1/4	<5	<5	—

<sup>a</sup> All reactions were performed on a scale of 0.08 mmol ( $c = 25$  mM) with a LED lamp (3 W power output) as light source. Irradiation time: 13 h. <sup>b</sup> The conversion was calculated from recovered starting material. <sup>c</sup> Total yield of isolated products **4** and **5**. <sup>d</sup> Ratio of cyclized to hydro-de-brominated product as determined by  $^1\text{H-NMR}$ . <sup>e</sup> Attempted reaction without irradiation. <sup>f</sup> No catalyst was added. <sup>g</sup> No  $\text{NEt}_3$  was added. n.d. = not detected.



loading was further decreased to 1 mol%: while the performance of  $\text{Ni}^{II}$  catalyst **1** remained unchanged (entry 4) the reaction with the  $\text{Ru}^{II}$  catalyst was sluggish and a decrease in yield was observed (entry 5).

The above reaction is induced by visible light as no conversion occurs without irradiation (entry 6). In the absence of the  $\text{Ni}^{II}$  catalyst,<sup>17</sup> only 12% of a product mixture was obtained, which was composed of the cyclized product **4** and the reduced product **5** in a 55/45 ratio (entry 7). In the absence of the reductant, no reaction was observed (entry 8). The free HMabiq ligand is not photocatalytically active. The quantum yield for the  $[\text{Ni}(\text{Mabiq})]\text{OTf}$  catalysed cyclization reaction was determined to be  $\Phi = 0.006$ .

Mechanistically, it is suggested that the photoreduction of complex **1** by  $\text{NEt}_3$  (reductive quenching cycle)<sup>1</sup> generates complex **2**, which may engage in SET to the bromide **3** (Scheme 3). Complex **2** was shown to be competent to reduce **3**. We note that modification of the Ni-Mabiq complex during the cyclization reaction was not observed, as verified by ESI-MS and  $^1\text{H}$ -NMR (Fig. S8 and S41<sup>‡</sup>).

To experimentally verify the estimated photoexcited state redox potential of  $\text{1}^*$ , we synthesized a series of amines with oxidation potentials in the range of 0.78–1.59 V (vs. SCE; Tables 2 and S1<sup>‡</sup>), as determined by CV. The amines were employed as sacrificial donors in the cyclization reaction. Excellent yields of **4** were obtained using the donor molecules with oxidation potentials up to 1.25 V (Table 2, amines **6a**–**6c**), whereas a drastic decrease in yield was observed using those with higher redox potentials. Only 40% and 20% product yields were obtained with **6d** ( $E_{\text{ox}} = 1.41$  V vs. SCE) and **6e** ( $E_{\text{ox}} = 1.59$  V vs. SCE), respectively. The product yield in the control reaction using **6e** in the absence of photocatalyst **1** was 7%, a result that is comparable to that obtained using  $\text{NEt}_3$  (Table 1, entry 7). The results confirm that the excited state redox potential of  $\text{1}^*$  is at least 1.25 V (vs. SCE). Thus, our new Ni-Mabiq complex is a more powerful oxidant than  $[\text{Ru}(\text{bpy})_3]^{2+}$  ( $E_{1/2} [\text{Ru}^{2+}/\text{Ru}^+] = 0.78$  V vs. SCE),<sup>18</sup> and comparable to  $[\text{Ir}(\text{dF}(\text{CF}_3)\text{ppy})_2(\text{dtbbpy})]^+$  ( $E_{1/2} [\text{Ir}^{3+}/\text{Ir}^{2+}] = 1.21$  V vs. SCE).<sup>19</sup>

We additionally generated a sterically hindered amine,  $\text{N}(\text{CH}_2\text{Mes})\text{Cy}_2$  (**6f**), to assess whether coordination of the sacrificial donor molecules influences the reactivity of **1**. In contrast to the bpy- and phen-based photosensitizers, the Ni-Mabiq complexes are coordinatively unsaturated and intramolecular electron transfer from a coordinated amine to the

Table 2 Comparison of the oxidation potentials for different sacrificial donors **6** with the yields obtained for the catalytic reaction of **3** to **4** and **5**

<b>6b</b>		<b>6f</b>		<b>6c</b> R = Me
				<b>6d</b> R = $\text{CH}_2\text{CO}_2\text{Et}$
				<b>6e</b> R = $\text{CH}_2\text{C}_6\text{F}_5$

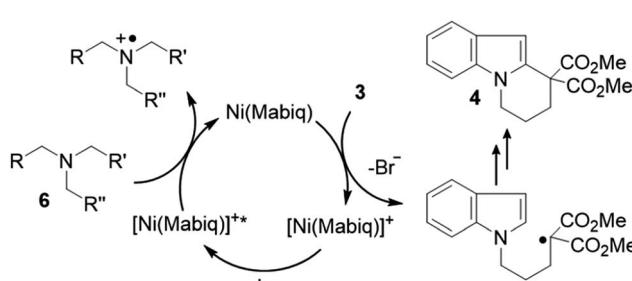
Amine <sup>a</sup>	$E_{\text{ox}}$ (vs. SCE)	Yield <sup>b</sup> [%]
$\text{Et}_3\text{N}$ ( <b>6a</b> )	0.83	84
<b>6b</b>	1.05	97
<b>6c</b>	1.25	95
<b>6d</b>	1.41	40
<b>6e</b>	1.59	20 <sup>c</sup>
<b>6f</b>	0.78	84

<sup>a</sup> All reactions were performed on a scale of 0.08 mmol ( $c = 25$  mM) with a 457 nm LED lamp (3 W power output) as the light source. Irradiation time: 13 h. <sup>b</sup> Total yield of isolated products **4** and **5**, with a 95 : 5 ratio of cyclized to hydro-de-brominated product as determined by  $^1\text{H}$ -NMR. <sup>c</sup> Average of two runs.  $\text{Fc}^{+/-} = 0.4$  V vs. SCE. Cy = cyclohexyl.

divalent metal center could potentially occur. The molecular structure of **6f** (Fig. S3<sup>‡</sup>) suggests that coordination of this amine group to the Ni center is unlikely; the cyclohexyl and mesityl groups encapsulate the nitrogen atom rendering the lone pair inaccessible. Whereas noticeable changes in the absorption spectrum of **1** were observed upon addition of  $\text{Et}_3\text{N}$  (400 equiv.) to a solution of the complex in THF/DMF, the addition of **6f** has no effect (Fig. S30 and S31<sup>‡</sup>). However, the yield of the catalytic reaction using **6f** as sacrificial donor was found to be 84%, a value that is comparable to the one obtained using  $\text{Et}_3\text{N}$  (**6a**) and **6b**–**6e** as sacrificial donors.

## Conclusions

In summary, we have discovered a new photoredox catalyst,  $[\text{Ni}(\text{Mabiq})]\text{OTf}$  (**1**) that is based on the earth-abundant metal nickel. The diamagnetic, bench-stable compound was readily prepared, its redox properties were studied and the one-electron reduced form  $\text{Ni}(\text{Mabiq})$  (**2**) likewise was isolated. Further studies to elucidate the detailed photophysical properties of  $\text{1}^*$  are warranted. However, we have already demonstrated that the photoexcited complex is a strong oxidant, with the capacity to induce C–C bond formation in an initial test reaction. The Ni-Mabiq compound may offer an alternative to noble metal photosensitizers for other synthetic transformations in organic photoredox chemistry, as well as for energy conversion processes. The Mabiq ligand also features a second metal binding site that could be exploited for tandem catalysis. Thus, the macrocycle represents a new type of platform for the development of photoactive systems. With evidence of the ability of **1** to act as a photosensitizer and photoredox catalyst, the broader photocatalytic applications of our system subsequently will be investigated.



Scheme 3 Proposed catalytic cycle for the cyclization of **3** to **4**.



## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

We thank Dr Alex Pöthig for contributions to the crystallography, and Dr Andreas Bauer for assistance with fluorescence and quantum yield measurements. We are grateful to Prof. Klaus Köhler for the use of his EPR instrument, and to Dr Carmen Haeßner and Manuel Kaspar for technical assistance. I. B. acknowledges the Ramón Areces Foundation (Becas para Estudios Postdoctorales) and the Alexander von Humboldt Foundation for a postdoctoral fellowship. M. G acknowledges the TUM Graduate School for financial support.

## Notes and references

- 1 For seminal studies, see: (a) M. A. Ischay, M. E. Anzovino, J. Du and T. P. Yoon, *J. Am. Chem. Soc.*, 2008, **130**, 12886–12887; (b) D. A. Nicewicz and D. W. C. MacMillan, *Science*, 2008, **322**, 77–80; (c) J. M. R. Narayanan, J. W. Tucker and C. R. J. Stephenson, *J. Am. Chem. Soc.*, 2009, **131**, 8756–8757.
- 2 Reviews: (a) D. M. Schultz and T. P. Yoon, *Science*, 2014, **343**, 985; (b) Y. Xi, H. Yi and A. Lei, *Org. Biomol. Chem.*, 2013, **11**, 2387–2403; (c) M. Reckenthaler and A. G. Griesbeck, *Adv. Synth. Catal.*, 2013, **355**, 2727–2744; (d) C. K. Prier, D. A. Rankic and D. W. C. MacMillan, *Chem. Rev.*, 2013, **113**, 5322–5363; (e) J. M. R. Narayanan and C. R. J. Stephenson, *Chem. Soc. Rev.*, 2011, **40**, 102–113; (f) K. Zeitler, *Angew. Chem., Int. Ed.*, 2009, **48**, 9785–9789.
- 3 D. M. Arias-Rotondo and J. K. McCusker, *Chem. Soc. Rev.*, 2016, **45**, 5803–5820.
- 4 (a) L. A. Büldt and O. S. Wenger, *Angew. Chem., Int. Ed.*, 2017, **56**, 5676–5682; (b) L. A. Büldt, X. Guo, A. Prescimone and O. S. Wenger, *Angew. Chem., Int. Ed.*, 2016, **55**, 11247–11250; (c) J.-P. Sauvage, J.-P. Collin, J.-C. Chambron, S. Guillerez and C. Coudret, *Chem. Rev.*, 1994, **94**, 993–1019; (d) D. J. Stufkensa and A. Vlcek Jr, *Coord. Chem. Rev.*, 1998, **177**, 127–179; (e) K. S. Schanze, D. B. MacQueen, T. A. Perkins and L. A. Cabana, *Coord. Chem. Rev.*, 1993, **122**, 63–89; (f) S. Meister, R. O. Reithmeier, M. Tschurl, U. Heiz and B. Rieger, *ChemCatChem*, 2015, **7**, 690–697.
- 5 For the use of organic compounds as photoredox catalysts, see: N. A. Romero and D. A. Nicewicz, *Chem. Rev.*, 2016, **116**, 10075–11166.
- 6 (a) Z. Zuo, D. T. Ahneman, L. Chu, J. A. Terrett, A. G. Doyle and D. W. C. MacMillan, *Science*, 2014, **345**, 437–440; (b) D. T. Ahneman and A. G. Doyle, *Chem. Sci.*, 2016, **7**, 7002–7006; (c) M. Jouffroy, D. N. Primer and G. A. Molander, *J. Am. Chem. Soc.*, 2016, **138**, 475–478; (d) B. P. Woods, M. Orlandi, C.-Y. Huang, M. S. Sigman and A. G. Doyle, *J. Am. Chem. Soc.*, 2017, **139**, 5688–5691; (e) M. K. Nielsen, B. J. Shields, J. Liu, M. J. Williams, M. J. Zacuto and A. G. Doyle, *Angew. Chem., Int. Ed.*, 2017, **56**, 7191–7194; (f) J. Twilton, C. Le, P. Zhang, M. H. Shaw, R. W. Evans and D. W. C. MacMillan, *Nat. Rev. Chem.*, 2017, **1**, 1–18; (g) E. R. Welin, C. Le, D. M. Arias-Rotondo, J. K. McCusker and D. W. C. MacMillan, *Science*, 2017, **355**, 380–385.
- 7 L. A. Cameron, J. W. Ziller and A. F. Heyduk, *Chem. Sci.*, 2016, **7**, 1807–1814.
- 8 For a recent review of earth-abundant photosensitizers see: C. B. Larsen and O. S. Wenger, *Chem.-Eur. J.*, 2018, **24**, 2039–2058.
- 9 (a) F. J. Sarabia and E. M. Ferreira, *Org. Lett.*, 2017, **19**, 2865–2868; (b) S. M. Stevenson, M. P. Shores and E. M. Ferreira, *Angew. Chem., Int. Ed.*, 2015, **54**, 6506–6510; (c) A. Gualandi, M. Marchini, L. Mengozzi, M. Natali, M. Lucarini, P. Ceroni and P. G. Cozzi, *ACS Catal.*, 2015, **5**, 5927–5931; (d) P. Zimmer, P. Müller, L. Burkhardt, R. Schepper, A. Neuba, J. Steube, F. Dietrich, U. Flörke, S. Mangold, M. Gerhards and M. Bauer, *Eur. J. Inorg. Chem.*, 2017, 1504–1509; (e) S. Otto, A. M. Nauth, E. Ermilov, N. Scholz, A. Friedrich, U. Resch-Genger, S. Lochbrunner, T. Opatz and K. Heinze, *ChemPhotoChem*, 2017, **1**, 344–349.
- 10 S. J. Hwang, D. C. Powers, A. G. Maher, B. L. Anderson, R. G. Hadt, S.-L. Zheng, Y.-S. Chen and D. G. Nocera, *J. Am. Chem. Soc.*, 2015, **137**, 6472–6475.
- 11 E. Müller, G. Bernardinelli and A. von Zelewsky, *Inorg. Chem.*, 1988, **27**, 4645–4651.
- 12 (a) P. Banerjee, A. Company, T. Weyermann, E. Bill and C. R. Hess, *Inorg. Chem.*, 2009, **48**, 2944–2955; (b) E. V. Puttock, P. Banerjee, M. Kaspar, L. Drennan, D. S. Yufit, E. Bill, S. Sproules and C. R. Hess, *Inorg. Chem.*, 2015, **54**, 5864–5873; (c) M. Kaspar, P. J. Altmann, A. Pöthig, S. Sproules and C. R. Hess, *Chem. Commun.*, 2017, **53**, 7282–7285.
- 13 The redox potential of  $1^*/2$  was calculated from the redox potential of  $1/2$  ( $-1.05$  V vs.  $\text{Fc}^{+/-}$ ) by adding the energy of the excited state  $E_0$  (see ref. 3).
- 14 For recent work on photoredox-induced alkylation reactions at the indole core, see: D. Alpers, M. Gallhof, J. Witt, F. Hoffmann and M. Brasholz, *Angew. Chem., Int. Ed.*, 2017, **56**, 1402–1406 and refs cited therein.
- 15 J. W. Tucker, J. M. R. Narayanan, S. W. Krabbe and C. R. J. Stephenson, *Org. Lett.*, 2010, **12**, 368–371.
- 16 For the reaction set-up, see: (a) D. Rackl, V. Kais, P. Kreitmeier and O. Reiser, *Beilstein J. Org. Chem.*, 2014, **10**, 2157–2165; (b) D. Lenhart, A. Pöthig and T. Bach, *Chem.-Eur. J.*, 2016, **22**, 6519–6523.
- 17 Radical reactions of amines can be initiated in the absence of a photoredox catalyst if an alkyl halide is present: (a) J. F. Franz, W. B. Kraus and K. Zeitler, *Chem. Commun.*, 2015, **51**, 8280–8283; (b) A. Böhm and T. Bach, *Chem.-Eur. J.*, 2016, **22**, 15821–15928; (c) A. M. Nauth, J. C. O. Pacheco, S. Pusch and T. Opatz, *Eur. J. Org. Chem.*, 2017, 6966–6974.
- 18 A. Juris, V. Balzani, P. Belser and A. von Zelewsky, *Helv. Chim. Acta*, 1981, **64**, 2175–2182.
- 19 M. S. Lowry, J. I. Goldsmith, J. D. Slinker, R. Rohl, R. A. Pascal Jr, G. G. Mailliaras and S. Bernhard, *Chem. Mater.*, 2005, **17**, 5712–5719.

