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Direct methylation of benzene with methane over Co/MFI catalysts generated by self-dispersion of $Co(OH)_2$†

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A methane–benzene reaction was performed over Co/MFI prepared with an aqueous solution of $Co(OAc)_2$ and MFI zeolite. $Co(OH)_2$ was deposited on the MFI zeolite when the $Co(OAc)_2$ solution was stirred at 343 K in the presence of MFI. A decline of the h-peak in NH_3 TPD occurred during the loading of Co on MFI. Further thermal treatment at 573–773 K caused the spontaneous formation of the atomically dispersed Co species, which was observed with Co K-edge XAFS. The toluene formation rate of the catalyst prepared with $Co(OAc)_2$ was proportional to the Co loading and was higher than that prepared with $Co(NO_3)_2$ due to the large amount of supported Co.

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1. Introduction

Methane, the main component of natural gas, is an abundant resource on earth, thus it is expected to be used not only as a fuel but also as a carbon resource. In order to use methane for non-fuel applications, it is usually converted to other molecules by C1 chemistry *via* synthesis gas ($CO + H_2$) obtained by steam reforming of methane with water vapor using Ni or Ru supported catalysts.[1,2] Meanwhile, if methane can be used directly as a carbon source, it is expected that a new synthetic route will be developed. The use of methane as a carbon source is beneficial over conventional ones using petroleum as the carbon source for production of plastics and so on from an economic viewpoint. Various catalysts that directly utilize methane as a carbon resource have been developed including Mo-loaded MFI for the formation of benzene.[3] Among them, Co/MFI has been found to be active in the direct methylation of benzene by methane to give toluene and hydrogen.[4] The reaction proceeded exclusively when MFI zeolite was employed as the support for Co among various structures of zeolites.[5]

The conventional Co/MFI catalysts have primarily been prepared with an ion-exchange method.[6] The Co located at the α sites was active in the methane–benzene reaction[7] among three kinds of ion-exchange sites (α, β, γ).[8] The adsorption energy generated at the adsorption of benzene on the Lewis acid sites of Co compensated the activation energy required for scission of the C–H bond in methane, so that the reaction between methane and benzene was promoted.[7] It could be supposed that one of the ways to improve the catalytic activity of Co/MFI zeolite is to increase the loading of Co with an atomically dispersed form using the ion-exchange sites of MFI. Another possible way to increase the loading of dispersed Co is to make use of the reaction between defect sites of MFI and Co. Such a method has been reported to prepare atomically dispersed species of various elements including Ti and Sb using volatile precursors.[9,10] In previous studies on Co/MFI catalysts, $Co(NO_3)_2$ has been primarily used as a precursor of Co for the preparation of Co/MFI by the above-mentioned ion-exchange method under non-hydration conditions.[11] The hydrolysis of Co^{2+} ions is inhibited under acidic or low temperature conditions.[12] Other possible methods for the preparation of Co/MFI is the impregnation of $Co(NO_3)_2$ and the sublimation of volatile Co precursors.[13] In the former case, aggregated and dispersed Co oxides formed simultaneously, resulting in a heterogeneous distribution of Co sizes.[4] In this research, we employed cobalt acetate tetrahydrate ($Co(OAc)_2{\cdot}4H_2O$) as the precursor of Co/MFI in place of $Co(NO_3)_2$ on purpose. The pH of the aqueous solution of $Co(OAc)_2$ was *ca.* 7.2 at 200 mmol L^{-1}. Under this condition, hydrolysis of Co^{2+} ions occurs to form $Co(OH)_2$ spontaneously at around 343 K.[14] We expected that the deposited $Co(OH)_2$ on the surface of MFI zeolites acted as the possible precursor of active Co species in the methane–benzene reaction, which was generated through the reaction with defect or ion exchange sites in MFI zeolites. This is because $Co(OAc)_2$ was reported to be a

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promising precursor to obtain well-dispersed Co species in MFI *via* solid-state ion-exchange reactions.[15] MFI treated with an aqueous NaOH solution was also used as the support for Co, taking into account that treatment with an aqueous NaOH solution has been reported to change the porosity of MFI.[16]

The dispersion process of aggregated to dispersed Co species was monitored with Co K-edge X-ray absorption fine structure (XAFS), and NH_3 temperature programmed desorption (NH_3-TPD). Particularly, Co K-edge XAFS is powerful technique to obtain information on the structural and electronic state of the Co center, which has been applied for the characterization of Co/zeolites so far.[17,18]

2. Experimental

2.1. Catalyst preparation

The experiment was carried out primarily using the MFI zeolite supplied by Mizusawa Chemical Co. (52A). In addition to the 52A sample, MFI, FAU, LTL, MOR, and BEA zeolites purchased from Tosoh Co. or obtained from Catalysis Society of Japan and SiO_2 (Fuji-Silysia Co.) were employed as the support for Co as well. The composition, abbreviations of the employed zeolites are listed in Table 1. The letters "N" and "H" indicate the NH_4 and H forms of zeolites, respectively. The sample treated with an aqueous NaOH solution for zeolite is abbreviated as sample name + AT + treatment time. For example, 822N-AT6h indicated that Co was loaded on the NH_4-type 822 treated with an NaOH aqueous solution for 6 h. For the preparation of Co-loaded samples, the zeolite sample (3 g) was ion exchanged with an aqueous solution of NH_4NO_3 (200 mmol L^{-1}) at 343 K three times for 4 h to obtain the NH_4-form prior to the loading of Co. The amount of NH_4NO_3 used for each ion exchange treatment was ten times as much as the Al present in zeolites. In some sample preparations, Co was loaded on H-type zeolites. Some of the MFI samples were treated with an aqueous solution of NaOH (0.7 mol L^{-1}, 200 mL) at 313 K for 1–16 h. The treated samples were rinsed repeatedly with deionized water and subsequently dried in air. The NaOH-treated MFI samples were transformed to an NH_4-type after ion exchange with an NH_4NO_3 solution (0.5 mol L^{-1}) at 343 K three times, prior to the loading of Co. Cobalt(II) acetate tetrahydrate ($(CH_3COO)_2Co{\cdot}4H_2O$) was purchased from Fujifilm Wako Pure Chemical Industries. Zeolites (primarily NH_4-form) were added to the aqueous solution of $Co(OAc)_2$ (200 mL), and then they were continuously stirred at 343 K for 4 h to load Co on the zeolites. $Co(OAc)_2$ solutions with different Co concentrations (50–400 mmol L^{-1}) were used for the preparation of Co/MFI. The typical concentration of $Co(OAc)_2$ was 200 mmol L^{-1}. The Co loaded samples were separated by filtration, rinsed repeatedly with deionized water, and dried in an oven kept at 373 K. The loading of Co on silica was carried out in the same manner as for the preparation of Co/zeolites.

Table 1 Zeolites and silica samples employed as the support for Co

Sample name	Structure	Abbreviation	Supplier	Cation type	Si/Al_2
HSZ-820NHA	MFI	820	Tosoh Co.	NH_4^+	23
HSZ-822HOA	MFI	822	Tosoh Co.	H^+	24
HSZ-840HOA	MFI	840	Tosoh Co.	H^+	40
HSZ-860HOA	MFI	860	Tosoh Co.	H^+	72
HSZ-891HOA	MFI	891	Tosoh Co.	H^+	1500
52A	MFI	52	Mizusawa Chemical Co.	Na^+, H^+	52
JRC-Z5-90NA	MFI	90	Catalysis Society of Japan	Na^+	90
HSZ-360UHA	FAU	360	Tosoh Co.	H^+	15
HSZ-500KOA	LTL	500	Tosoh Co.	K^+	6.1
HSZ-640HOA	MOR	640	Tosoh Co.	H^+	380
HSZ-930NHA	BEA	930	Tosoh Co.	NH_4^+	27
Q3	SiO_2	SiO_2	Fuji-Silysia Co.	—	—

2.2. Sample characterization

Co K-edge XAFS data of the $Co(OAc)_2$-loaded supports were collected using synchrotron radiation. The data were recorded at the BL-9C beamline with the approval of the Photon Factory of the High Energy Accelerator Research Organization (KEK-PF). Proposal numbers were 2020G621 and 2022G581. The samples pressed into a wafer form (one pellet) were subjected to the XAFS measurement under ambient conditions. The data were obtained in a quick scan mode within 5 min using a Si (111) monochromator in the transmission mode. Ion chambers were used for the detection of primary (I_0, 100% N_2) and transmitted (15% Ar with the balance N_2) X-rays. The beam size at the sample position was 0.8 mm (horizontal) × 0.6 mm (vertical). The Co K-edge extended X-ray absorption fine structure (EXAFS) was analyzed by extracting oscillations using a spline smoothing method. The Fourier transform (FT) of the k^3-weighted EXAFS oscillations and $k^3\chi(k)$ from k-space to r-space was conducted in the range of 3–13 $Å^{-1}$ for curve-fitting analysis. The EXAFS data were analyzed using the REX software (Rigaku Co.). The parameters for the analysis of the Co–O and Co–Co bonds were extracted from Co K-edge EXAFS of CoO and Co foil, respectively. For the analysis of the Co–Si bond, parameters were obtained using the FEFF8.0 code.[19] In addition to the *ex situ* conditions, operando XAFS measurement was carried out at BL-9C of KEK-PF. For this purpose, a wafer sample (40 mg, 10 mm-ϕ) was mounted in a stainless-steel cross-shaped cell. XAFS and IR measurements were performed using crossed X-ray and infrared (IR) beams at the sample position (Fig. S1†). The sample pellet was placed on the optical path at an angle of 45° to the X-ray and IR. Data of the Co K-edge XAFS were obtained under flowing N_2 (30 mL min^{-1}) while increasing the temperature (heating rate: 10 K min^{-1}).

Co loadings were measured with the atomic absorption (AA) instrument AA-6200 (Shimadzu Co.). The sample was prepared after dissolving Co/MFI with an aqueous solution of HF and subsequently with conc. H_2SO_4, followed by

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evaporation. The residue dissolved with hydrochloric acid was subjected to the AA measurements for quantitative analysis of Co. Data of the temperature programmed desorption (TPD) of ammonia were obtained with a BELCAT II equipment (Microtrac Bel Co.). The sample was treated at 773 K in an N_2 flow prior to the measurement. Then ammonia (5%) diluted with He was equilibrated with the pretreated sample (0.1 g) at 373 K. The TPD data were collected with a temperature ramping rate of 10 K min^{-1}. The carrier gas (He) flow rate was 30 mL min^{-1}. A mass spectrometer (BELMASS, Microtrac Bel Co.) was used to measure the desorbed NH_3. In the measurement, m/z = 16 was monitored to analyze the desorbed NH_3. X-ray diffraction (XRD) patterns of the powder samples were obtained using a MiniFlex X-ray diffractometer (Rigaku Co.) with Cu Kα radiation in the 2θ range from 5 to 40° with 10° min^{-1} scanning speed. N_2 adsorption isotherms were recorded on the BELSORP-mini-X (Microtrac Bel Co.) instrument. The samples were dehydrated in a vacuum at 413 K in advance. Transmission electron microscopy (TEM) images were obtained using a JEOL-JEM-2100 microscope. For the preparation of the sample, an ethanol suspension of Co/MFI was dropped onto Cu grids coated with a C-coated porous thin membrane (NEM, Japan) and dried. TEM observations were performed at an operating voltage of 200 kV. Field emission scanning electron microscopy (FE-SEM) images were taken with a JEOL JSM-6701F microscope with an accelerating voltage of 5 kV.

2.3. Catalytic reaction

The catalytic reaction between methane and benzene was performed using a fixed-bed flow reactor. The prepared catalyst was mounted in a reaction tube filled with quartz wool, which was placed at the bottom of the powder catalyst. The catalyst (0.30 g) was pretreated at 823 K in an N_2 flow for 1 h prior to the reaction. Neat methane gas (99.995%) was passed through a bubbler containing liquid benzene (Fuji Film Wako Pure Chemical Industries) kept at 273 K with a flow rate of 30 mL min^{-1}. The gas mixture was reacted with the pretreated Co/MFI catalyst at 823 K. The ratio of methane to benzene was 29 : 1. The effluent gas was analyzed with an on-line FID type gas chromatograph (Shimadzu, GC-2014) equipped with a capillary column (GL Sciences Inc., InertCap 1). The toluene formation rate was calculated based on the benzene feed flow rate (2.7 mmol h^{-1}).

3. Results and discussion

3.1. Operando Co K-edge XAFS and IR studies

First, the local structure of Co/MFI(52N) was analyzed with Co K-edge XAFS for the as-prepared samples and the samples pretreated in an N_2 flow. Fig. 1(a) shows the Co K-edge X-ray absorption near edge structure (XANES) of the sample in which Co was loaded on MFI using $Co(OAc)_2$ as a precursor (as-prepared) and reference compounds. The XANES of the as-prepared Co/MFI sample was close to that of $Co(OH)_2$,

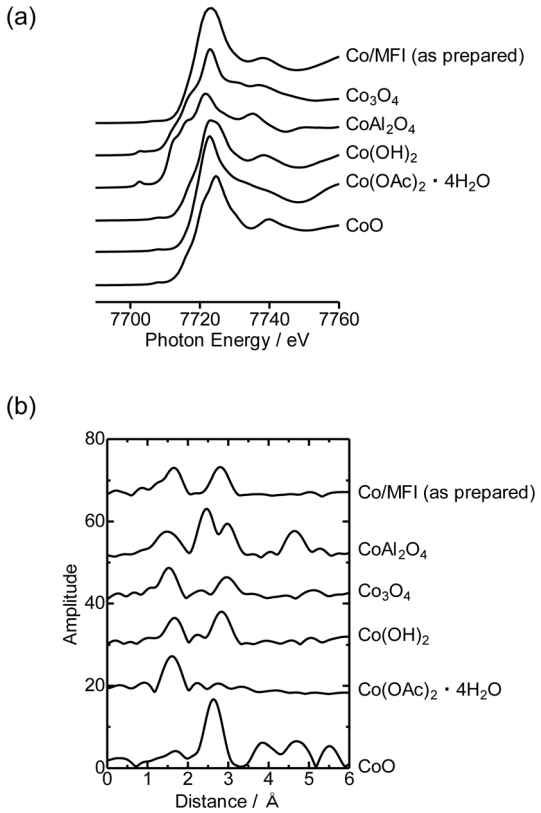

Fig. 1 Co K-edge (a) XANES and (b) EXAFS-FT of Co/MFI(52N) (Co: 0.5 mmol g^{-1}) and reference compounds measured at 298 K. Fourier transform range: 3–13 $Å^{-1}$.

indicating that $Co(OAc)_2$ was hydrated to give $Co(OH)_2$ during heat treatment of a $Co(OAc)_2$ solution in the presence of MFI at 343 K. As shown in Fig. 1(b), the Co K-edge EXAFS of the Co/MFI samples was also close to that of $Co(OH)_2$. The corresponding $k^3\chi(k)$ data are displayed in Fig. S2.† That is to say, two peaks assignable to the Co–O and Co–(O)–Co bonds appeared at 1.7 and 2.8 Å (phase shift uncorrected) in the EXAFS Fourier transform (EXAFS-FT), respectively. It has been reported that the dehydration of $Co(OAc)_2$ occurred at high pH or during heat treatment under neutral conditions.[14] Therefore, it was likely that hydration of $Co(OAc)_2$ proceeded during the preparation of Co/MFI to form $Co(OH)_2$ and acetic acid. The loading of Co in Co/MFI(52N) increased gradually with increasing concentration of $Co(OAc)_2$ measured by AA (Fig. 2(a)). At the same time, the Al content of MFI decreased and the Co loading increased (Fig. 2(b)). However, no shift of the diffraction in the XRD pattern was observed in the two 2θ regions before and after the loading of Co on MFI (Fig. S3†) when the 2θ angle was corrected with MgO as an internal standard. This fact indicated that the dealumination of framework Al hardly occurred. Instead, considering that MFI(52A) contains a large amount of extra framework Al (0.37 mmol g^{-1}), which was larger than that of framework Al (0.25 mmol g^{-1}), the extra framework Al species were probably removed during Co loading. The Co/Al ratio of

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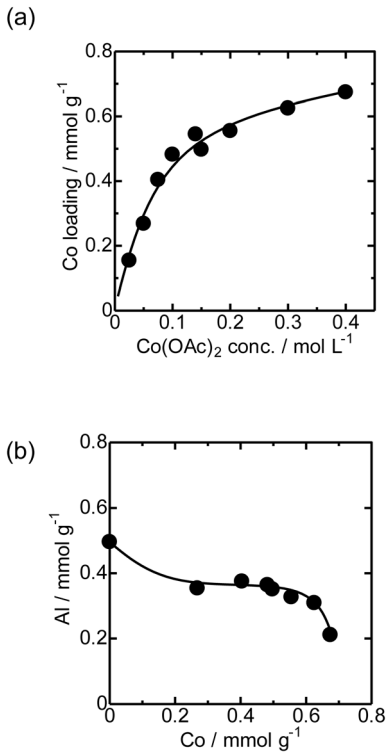

Fig. 2 (a) Relationship between the Co loading of Co/MFI(52N) and concentration of $Co(OAc)_2$. (b) Relationship between Co loading and the Al concentration of Co/MFI(52N).

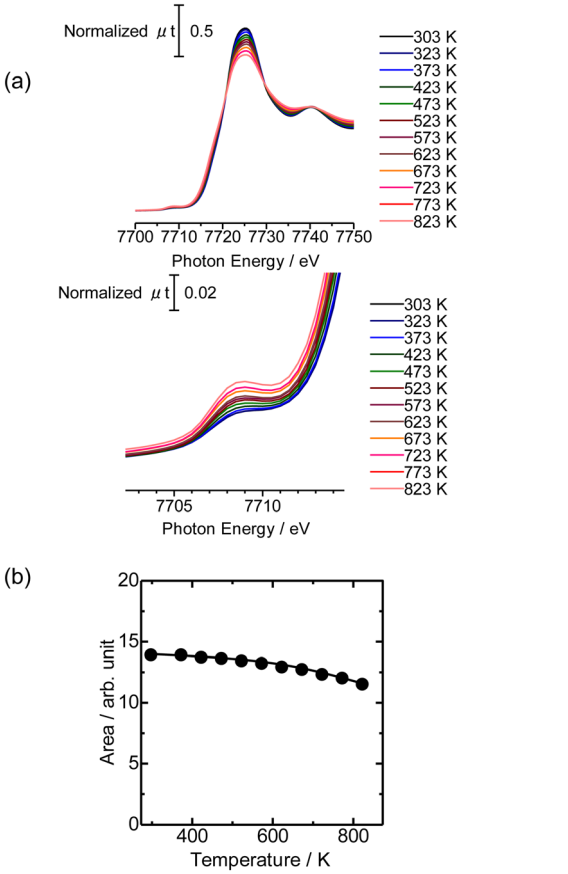

Fig. 3 (a) Co K-edge XANES of Co/MFI(52N) (Co: 0.5 mmol g^{-1}) measured under a N_2 flow from 303 to 823 K measured under *in situ* conditions. (b) White line area in Co K-edge XANES of Co/MFI(52N) plotted as a function of the treatment temperature.

Co/MFI exceeded 0.5 or 1.0 in many samples, suggesting that $Co(OH)_2$ deposition occurred independently of the Al sites in the MFI (Fig. S4†).

The loading of Co on MFI(52N) was 0.18 mmol g^{-1} when a 100 mmol L^{-1} $Co(NO_3)_2$ aqueous solution was used as the precursor for Co under acidic (non-hydration) conditions, which has been generally employed as the Co precursor. The loaded amount of Co (0.18 mmol g^{-1}) was comparable to that of ion-exchangeable sites of 52N (0.25 mmol g^{-1}) assuming that the ion-exchange occurred between NH_4^+ and $Co(OAc)_2$ or $Co(OH)_2$ to form $Co(OH)^+$ species. On the other hand, the Co loading increased to 0.48 mmol g^{-1} when a 100 mmol L^{-1} $Co(OAc)_2$ aqueous solution was used for the Co/MFI preparation, which exceeded the ion exchange capacity of 52N (0.25 mmol g^{-1}). The higher Co loading with the use of the $Co(OAc)_2$ precursor suggested that the formation of dispersed Co occurred through not only the ion exchange but also the reaction with defect sites with $Co(OH)_2$.

Then the Co K-edge XAFS data were collected in an N_2 flow while the temperature was increased from 303 to 823 K in order to monitor the local structure of Co species accompanied by heat treatments. The XANES data obtained under *in situ* conditions every 50 K are displayed in Fig. 3(a). A gradual decrease in the white line intensity of the Co K-edge XANES was observed up to 823 K as plotted in Fig. 3(b). The changes up to 573 K observed in XANES may be due to the dehydration of $Co(OH)_2$ to form CoO, consistent with the literature.[20] The change occurred progressively in the temperature range between 303 and 823 K, as shown in Fig. S5;† fitting the intermediate-stage XANES with a linear combination of the spectra obtained at room temperature and 823 K showed that the former component continuously decreased and the latter increased instead with increasing temperature. At the same time, the intensity of the pre-edge peak (7708 eV) increased with increasing temperature up to 823 K. This pre-edge peak is attributed to the 1s to 3d transition, which is normally dipole forbidden but allowed when the metal has tetrahedral coordination due to the mixing with p-character.[21,22] Therefore, the change may be due to the transformation of the local structure from 6- to 4-coordinated Co centers. This could be caused by the removal of water molecules coordinated to the Co centers during the heat treatment and subsequent reaction of Co with OH groups or acid sites in MFI.

Fig. 4(a) shows the Co K-edge EXAFS-FT of the sample in which Co was loaded on MFI(52N) using $Co(OAc)_2$ as a precursor measured during heat treatment. Raw XAFS and $k^3\chi(k)$ data are displayed in Fig. S6(a) and (b),† respectively. The Co loading of the measured sample was 0.5 mmol g^{-1}. When these samples were heated from 573 to 823 K, the

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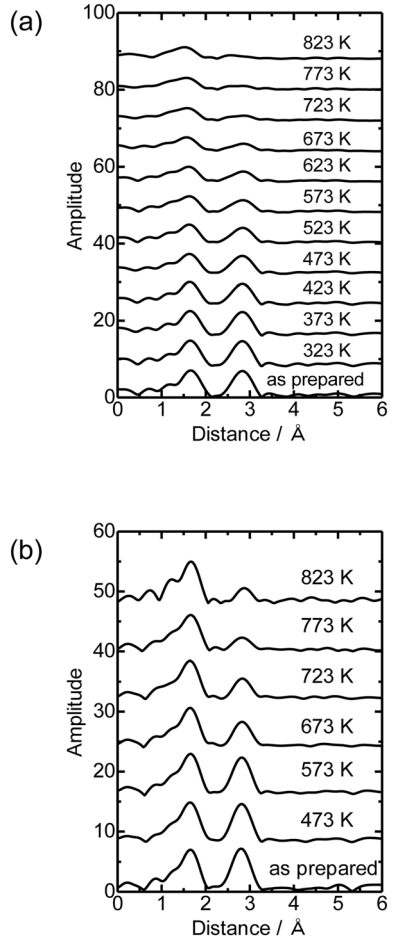

Fig. 4 Co K-edge EXAFS-FT of Co/MFI(52N) (Co: 0.5 mmol g^{-1}) (a) measured under a N_2 flow from 323 to 823 K under *in situ* conditions, the as-prepared sample and (b) measured under ambient conditions.

intensity of the Co–O–Co bond (2.8 Å) decreased while the Co–O peak remained almost intact. In the IR spectra measured at the same time, the intensity of isolated hydroxyl groups appearing at 3730 cm^{-1} remained unchanged, while the NH_4^+ stretching vibration peak at 3370 cm^{-1} decreased,[23] indicating that the decomposition of NH_4^+ progressed during heat treatment. The intensity of the acidic hydroxyl group at 3590 cm^{-1} tended to decrease with increasing temperature, suggesting that ion exchange between H^+ or NH_4^+ occurred with Co^{2+}, but at the same time, the formation of acidic hydroxyl groups progressed as a result of the decomposition of NH_4^+ (Fig. S7†). The EXAFS oscillation is greatly reduced as the temperature rises because the Debye–Waller factor increases at high temperature. To prevent this temperature effect, XAFS data of the heat-treated samples were collected at 298 K (Fig. 4(b)). The corresponding $k^3\chi(k)$ data are displayed in Fig. S8(a).† A significant decrease in the intensity of the second peak appearing at 2.8 Å (phase shift uncorrected) was observed in the temperature range of 573–773 K. A similar change was observed in the Co/MFI containing lower Co loading (0.3 mmol g^{-1}, Fig. S8(b) and (c)†). The EXAFS of the as-prepared Co/MFI sample was successfully fitted after assuming Co–O and Co–(O)–Co bonds, while that of the 773 K-treated one was fitted after assuming Co–O and Co–(O)–Si bonds as listed in Table 2 and shown in Fig. S9.† The change observed after treatment at 773 K meant that the aggregated $Co(OH)_2$ underwent self-dispersion through the interaction with MFI zeolite. Such dispersion was also found in Pd on MFI or MOR zeolites.[24] Other examples are the formation and segregation of Ir–MgO, and Pt–MgO solid solutions observed in the temperature range between 773 and 1273 K.[25,26] A similar change (decline) of the Co–(O)–Co bond appearing at 2.8 Å was observed in Co/SiO_2 (Co loading: 0.5 mmol g^{-1}) as displayed in the EXAFS-FT (Fig. S10(b)†); the peak was significantly reduced after the treatment at 773 K. This structural change suggested that the self-dispersion of Co occurred on Co/SiO_2, similar to the case of Co/MFI.

3.2. NH_3 TPD and XRD patterns

Fig. 5 shows the NH_3 TPD profile of Co/MFI(52N) with different Co loadings. The intensity of the h-peak appearing around 730 K was much lower than that of unloaded MFI(52N), while the change in the l-peak at *ca.* 500 K was small. The NH_3 TPD profile did not change depending on the Co loading. The decline of the h-peak (730 K) observed after the loading of Co might be caused by the ion exchange of H^+ with Co^{2+}.

NH_3 TPD profiles of the employed MFI zeolites (NH_4^+ form) and those of untreated 822N and those treated with an aqueous solution of NaOH, followed by treatment with an

Table 2 Curve fitting analysis of the Co K-edge EXAFS data measured at 298 K for Co/MFI(52N) (Co loading: 0.5 mmol g^{-1})[g]

Sample	Scatterer	CN[a]	R/Å[b]	ΔE_0[c]/eV	DW[d]/Å	R_f[e]/%
As-prepared Co/MFI(52N)	O	5.2 ± 1.1	2.07 ± 0.02	1 ± 3	0.084 ± 0.027	0.6
	Co	5.1 ± 1.3	3.14 ± 0.02	4 ± 2	0.086 ± 0.023	
Co/MFI(52N) treated at 773 K	O	4.9 ± 0.8	2.08 ± 0.01	4 ± 2	0.087 ± 0.020	2.3
	Si	2.0 ± 1.0	3.27 ± 0.03	−4 ± 5	0.062 ± 0.067	
$Co(OH)_2$[f]	O	(6)	(2.13)			
	Co	(6)	(3.20)			

[a] Coordination number. [b] Bond distance. [c] Difference in the origin of photoelectron energy between the reference and the sample. [d] Debye–Waller factor. [e] Residual factor. [f] Data of X-ray crystallography. [g] Fourier transform range: 3–13 $Å^{-1}$. Fourier filtering range: 1.0–3.2 Å.

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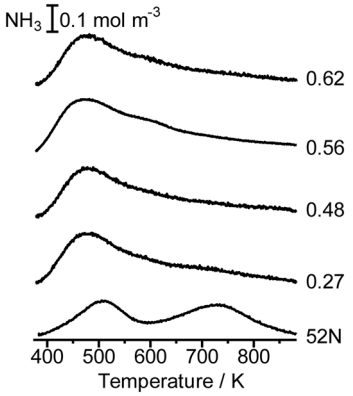

Fig. 5 NH_3 TPD profile of unloaded MFI(52N) and Co/MFI(52N). The samples were treated in an N_2 flow at 773 K prior to the TPD measurement. The numbers indicate the loading of Co (mmol g^{-1}).

NH_4NO_3 solution at 343 K are displayed in Fig. S11(a) and (b),† respectively. In the latter case, the intensity of the h-peak decreased in the MFI treated with an NaOH solution for 16 h. This means that the number of Al atoms in the MFI framework decreased after the NaOH treatment. In agreement with this, the XRD patterns revealed that partial collapse of the MFI crystals occurred after treatment with NaOH for 16 h as will be shown later. In general, the treatment of zeolite with NaOH solution resulted in the preferential dissolution of Si.[27] Consistent with this, the total Al concentration in the MFI increased with treatment time, as shown in Fig. S12.† As for Al distribution, the amount of framework Al decreased, while that of extra-framework Al increased with increasing treatment time with an aqueous solution of NaOH (Table S1†).

XRD diffraction patterns of the as-prepared Co/MFI(52N) and those treated at 773 K as well as unloaded MFI(52N) are displayed in Fig. 6(a) and (b), respectively. For Co-loaded MFI(52N), no diffraction other than those ascribed to the MFI zeolite was observed after the loading of $Co(OAc)_2$ (Fig. 6(a)), while the Co–(O)–Co bond appeared at 2.8 Å in the Co K-edge EXAFS as already shown in Fig. 4(a and b), indicating that Co existed on the MFI in an aggregated form. The contradictory data suggested that the crystallite size of $Co(OH)_2$ was less than several nanometers. In agreement with the observation, no clear aggregate of $Co(OH)_2$ was observed in the TEM image as will be noted later. Meanwhile, the overall intensity of the diffraction peaks decreased after the loading of Co. The extent was enhanced by increasing the Co loadings. This feature did not change after the thermal treatment in N_2 at 773 K (Fig. 6(b)).

XRD patterns of the as-received MFI(822N) and those treated with NaOH are displayed in Fig. S13.† The overall intensity of the diffraction declined accompanied by an increase in the concentration of the NaOH solutions, suggesting that the partial collapse of the MFI framework took place during the treatment with the NaOH solution and the extent of which was enhanced as the concentration of the NaOH solution increased.

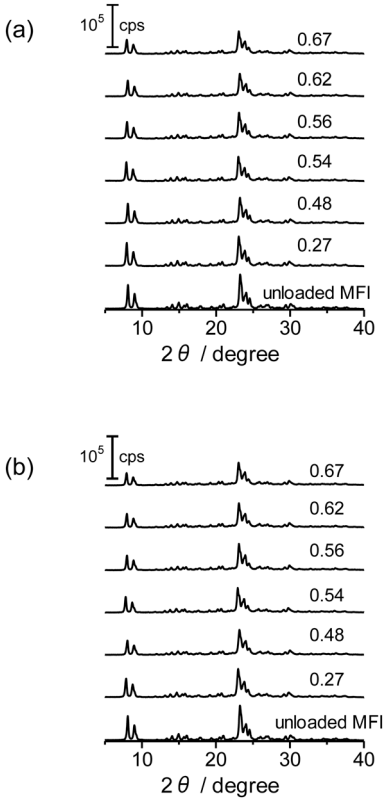

Fig. 6 XRD patterns of Co-unloaded MFI(52N) and Co/MFI(52N). (a) The as-prepared and (b) 773 K-treated samples. The numbers indicate the Co loading (mmol g^{-1}).

3.3. N_2 adsorption isotherms, SEM and TEM images

The N_2 adsorption isotherms of the unloaded MFI(52N) and Co-loaded ones are displayed in Fig. S14(a).† The isotherms of the Co/MFI (0.4 mmol g^{-1}) was identical to that of the Co-unloaded one. The specific surface areas derived from the isotherm are plotted as a function of Co loadings in Fig. 7. A gradual decrease in surface area was observed with increasing Co loading between 0.4 and 0.7 mmol g^{-1}. The tendency was consistent with the XRD pattern; a decrease in diffraction intensity was observed with increasing Co loading (Fig. 6).

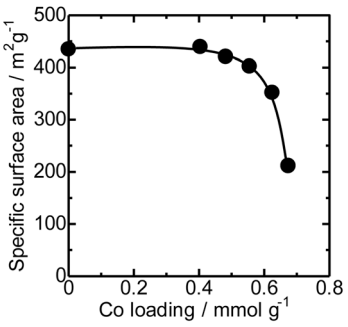

Fig. 7 Specific surface area plotted as a function of Co loadings of Co/MFI(52N).

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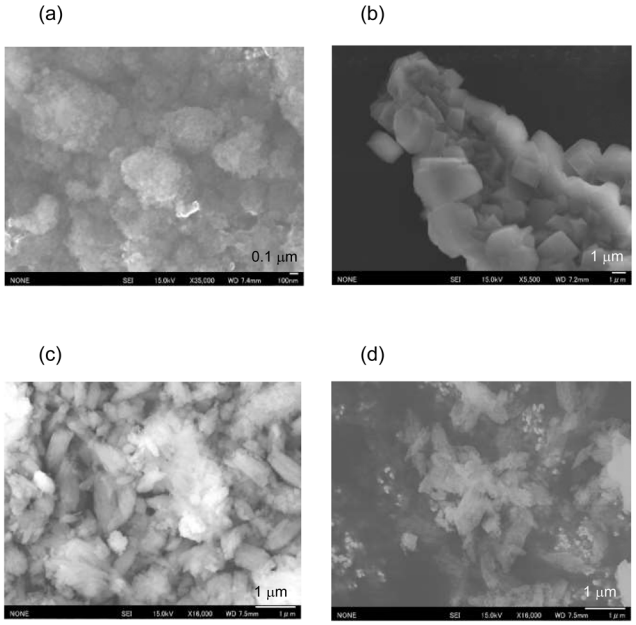

Fig. 8 FE-SEM images of (a) 52N, (b) 840N, (c) 822N, and (d) 822N-AT16h.

FE-SEM images of the representative MFI(52N, 840N, 822N, and 822N-16h) employed as the support for Co are displayed in Fig. 8. The average crystal size of 840N was approximately several micrometers (Fig. 8(b)). The large crystal size was associated with the appearance of the N_2 adsorption isotherm in that the isotherm of the samples was assignable to type I as shown in Fig. S14(b).† The size of the primary particle of 52N and 822N was much smaller than that of 840N; in the former case, the size was *ca.* 50 nm. Consistent with this SEM figure, a progressive increase in the nitrogen adsorption isotherm curve was observed at pressures above $p/p_0 = 0.01$ (Fig. S14(b)†).

No difference was found between TEM images of Co/MFI(52N) (as-prepared sample and those thermally treated at 823 K) and Co-unloaded MFI(52N) as shown in Fig. S15,† suggesting that the degree of the aggregation of the deposited $Co(OH)_2$ was limited in the preparation step. This fact was consistent with the XRD patterns in that no diffraction assignable to the crystalline $Co(OH)_2$ was observed (Fig. 6(a)).

3.4. Catalytic reaction

Fig. 9 shows the time-course change in the rate of toluene formation over Co loaded on different kinds of zeolites. A slight decrease in formation rate was observed in the highly active samples of Co/MFI, but the activity of the catalyst was relatively stable. It can be seen from Fig. 9(a) and (b) that the formation rate was dependent on the kinds of employed MFI zeolites. Co loaded on zeolites other than MFI was almost inactive in the reaction (Fig. 9(c)). The specific high activity of Co supported on MFI was also found with the catalyst using $Co(NO_3)_2$ as the precursor.[4] The H^+- and Na^+-MFI have

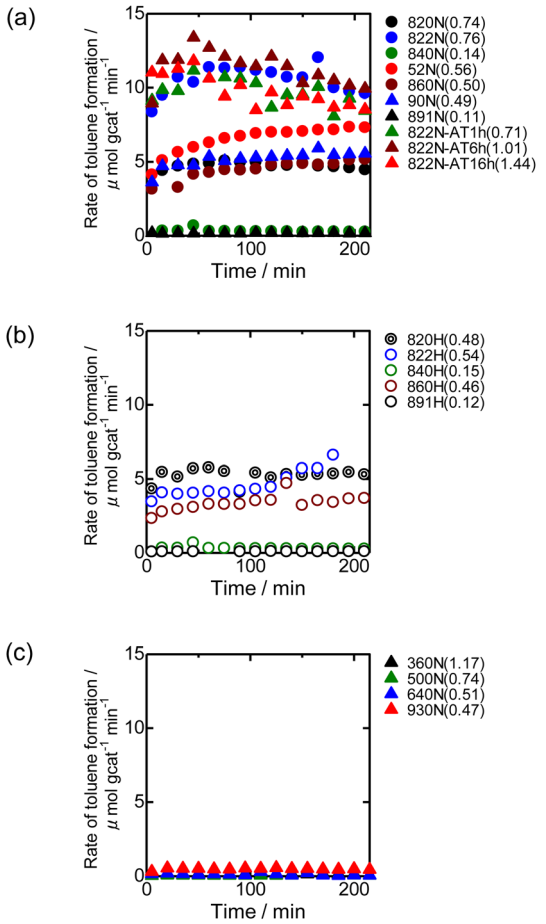

Fig. 9 Time course change of the toluene formation rate over Co loaded on different kinds of zeolite supports: (a) NH_4-form MFI, (b) H-form MFI, and (c) zeolites other than MFI. The Co/MFI catalysts were prepared using a 200 mmol L^{-1} $Co(OAc)_2$ solution at 343 K for 4 h. The numbers in the parentheses indicate the loading of Co (mmol g^{-1}).

strong Brønsted and Lewis acid character, respectively.[28,29] Therefore, we speculate that the strong adsorption of benzene on the Lewis sites of Co^{2+} promoted the reaction with methane, the presence of which was confirmed with IR spectra of adsorbed NH_3.[7] The rate of toluene formation at 3 h of time on stream from the beginning of the reaction in Co/MFI is plotted as a function of Co loading in Fig. 10(a). The toluene formation rate was almost proportional to the loading of Co up to 1.0 mmol g^{-1} except for the Co/MFI(822N-AT16h), in which excessive Co was loaded on the MFI support (Co loading: 1.5 mmol g^{-1}). Since the formation rate of H-type MFI and Co-loaded samples of NH_4-type MFI follow the same line, cationic species of MFI during catalyst preparation had no influence on the catalytic activity of Co. The linear relationship indicated that the catalytic activity of Co/MFI depended only on Co loading. The highest toluene formation rate (11.3 μmol $gcat^{-1}$ min^{-1}) was higher than that (6.2 μmol $gcat^{-1}$ min^{-1}) of Co/MFI(52N) (Co loading: 0.34 mmol g^{-1}) prepared using $Co(NO_3)_2$ as the Co precursor with an impregnation method. The turnover frequency (TOF) value of Co/MFI prepared with $Co(OAc)_2$ was distributed in the

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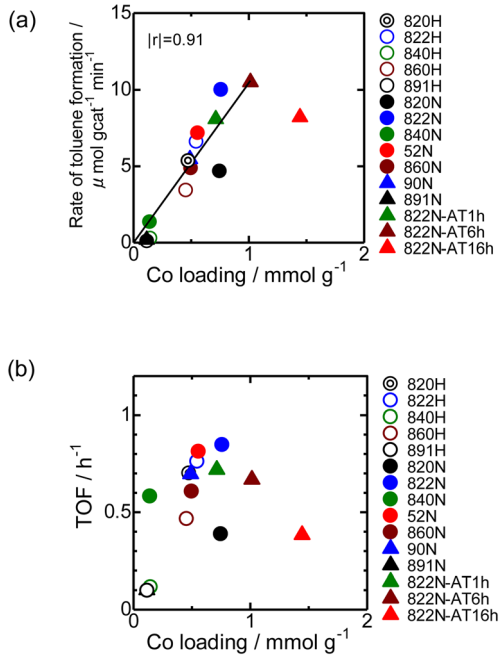

Fig. 10 (a) Toluene formation rate at 3 h of time on stream plotted as a function of Co loading for Co/MFI catalysts. The Co/MFI catalysts were prepared using a $Co(OAc)_2$ solution (200 mmol L^{-1}) at 343 K for 4 h. The linear plot was made without the addition of the data of 822N-AT16h. (b) TOF (h^{-1}) plotted as a function of Co loading. The TOF values were calculated based on the toluene formation rate at 3 h of time on stream from the beginning of the reaction and the Co loading.

range of 0.58–0.85 h^{-1} with the exception of several samples (Fig. 10(b)).

4. Conclusions

In this study, we found that $Co(OH)_2$ precipitated on MFI when $Co(OAc)_2$ was used as a precursor for Co/MFI. Further heat treatment led to self-dispersion of Co^{2+} species between 573 and 773 K. The toluene formation rate in the Co/MFI-catalyzed methane–benzene reaction was proportional to Co loading up to 1.0 mmol g^{-1}, suggesting that the catalytic activity was simply determined by the Co loading. Traditionally, atomically dispersed active species using zeolite defect sites have been prepared using volatile precursors. However, as found here, more readily available metal salt ($Co(OAc)_2{\cdot}4H_2O$) can be another precursor by making use of the self-dispersion of active species.

Conflicts of interest

There are no conflicts to declare.

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