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Pillar[5]arene-based ion-pair recognition for constructing a [2]pseudorotaxane with supramolecular interaction induced LCST behavior†

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A LCST-type [2]pseudorotaxane based on perbromoethylated pillar[5]arene/imidazolium iodide ionic liquid ion-pair recognition is constructed in both solution and the solid state. The thermo-responsive behavior can be adjusted by altering the molar ratio of the host and guest as well as the concentrations of these two components in solution.

Introduction

Along with the advent and burgeoning of supramolecular chemistry, mechanically interlocked molecules (MIMs), 1-5 due to their sophisticated topological architectures and potential applications in artificial molecular machines, have attracted significant interest in the scientific community. 6-9 Various types of MIMs, such as rotaxanes, pseudorotaxanes and catenanes, have been widely used as smart building blocks for the creation of supramolecular systems and advanced functional materials.10-15 Pseudorotaxanes, which are supramolecular complexes but not compounds, can be considered as not only crucial fundamental precursors for the synthesis of rotaxanes and catenanes but also the prototypes of simple supramolecuchemical topologies and molecular devices. 16-19 Considering the particular molecular architectures with fascinating properties and extensive applications of pseudorotaxanes, the design and preparation of novel pseudorotaxanes with the ability to perform functions are valuable and highly desirable for chemists.20-24

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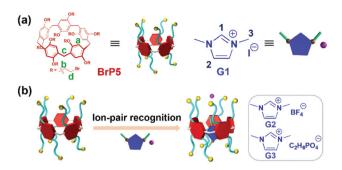
Endowing pseudorotaxanes with strong binding ability is also highly important in constructing pseudorotaxanes. Ionpair recognition, emerging from the coordination chemistry of cations and anions, has been proved to be an efficient strategy to construct robust pseudorotaxanes.²⁵ Typical ion-pair receptors contain disparate binding sites for complexing concurrently both cations and anions.²⁶ In particular, ion-pair recognition is believed to offer advantages in terms of strong affinity and selectivity during the recognition process, thus enhancing the binding strength and raising the association constants of the host-guest inclusion assemblies. 27-30 Due to the affirmative cooperative effects, ion-pair recognition-based pseudorotaxanes have been widely utilized in diverse areas such as ion extraction, self-assembly and logic gate design. 31-33 Therefore, exploring new systems for ion-pair recognition-based pseudorotaxanes with new properties and functionalities is still of great interest and importance, and will undoubtedly enrich the fundamental research and practical applications of ion-pair recognition chemistry.

Pillararenes, with unique symmetrical and rigid molecular architectures, have received significant attention and have experienced an extraordinary development in recent years. $^{34-40}$ Due to abundant host–guest recognition properties, pillar[n] arenes and their derivatives have been proven to serve as ideal hosts for the fabrication of various MIMs and more advanced supramolecular structures. $^{41-48}$ Herein, a new [2]pseudorotaxane based on perbromoethylated pillar[5]arene (BrP5) and 1,3-dimethylimidazolium iodide ionic liquid (G1) ion-pair recognition was successfully constructed (Scheme 1). Due to the introduction of ion-pair recognition, the binding ability of BrP5 \supset G1 was stronger than those of pseudorotaxanes formed between BrP5 and other imidazolium-based ionic liquids with larger counterions. The hydrogen bonds between the hydrogen atoms on the bromoethyl chains of BrP5 and the

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Scheme 1 (a) Chemical structures and cartoon representations of BrP5 and G1 used in this study. (b) Schematic representation of the ion-pair recognition process of BrP5 and G1 to form a [2]pseudorotaxane and chemical structures of G2 and G3 with different counterions.

iodide anions of G1 have greatly contributed to the formation of such a [2]pseudorotaxane. Interestingly, the [2]pseudorotaxane BrP5 \(G1 \) showed lower critical solution temperature (LCST) phase transition behaviors which can be modulated by altering the molar ratio of the host and guest as well as the concentrations of these two components in solution.

Results and discussion

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Initially, to investigate the complexation behavior of BrP5 and G1, ¹H NMR spectroscopy was conducted as a useful method to study the host-guest interactions. Significant chemical shift changes of the proton signals on both BrP5 and G1 occurred upon mixing them in CDCl₃. Obviously, as shown in Fig. 1, all the peaks related to protons H2 and H3 on G1 displayed conspicuous upfield shifts and remarkable complexation-induced broadening effects compared with those of free G1. Moreover, the proton signal of H₁ on G1 could not be observed after the formation of the host-guest complex, indicating that the imidazolium part of G1 was fully located in the cavity of BrP5 and

3 (b) (c) 7.5 7.0 6.5 6.0 5.5 5.0 4.5

Fig. 1 Partial ¹H NMR spectra (600 MHz, CDCl₃, room temperature): (a) G1; (b) BrP5 (10.0 mM) and G1 (10.0 mM); and (c) BrP5. The peaks marked with orange dots are ascribed to chloroform-d.

shielded by the electron-rich cyclic structure (ESI, Fig. S2†). Besides, the peaks of protons H_a , H_b , H_c and H_d on **BrP5** also exhibited chemical shift changes due to the host-guest interactions. It could be speculated that the extensive changes of the chemical shifts and shapes of proton signals on BrP5 and G1 resulted from the formation of the inclusion complex, in which the guest protons were located within the cavity of BrP5. Then, 2D NOESY NMR spectroscopy was carried out to investigate the host-guest complex with an interpenetrated geometry. As shown in Fig. 2, nuclear Overhauser effect (NOE) correlations were observed between the signals related to proton H₂ on G1 and protons H_b-H_d on BrP5, indicating that G1 penetrated the cavity of BrP5.

Further proton NMR titration experiments (ESI, Fig. S3-S5†) were carried out to measure the complexation stoichiometry and association constant for the complexation between BrP5 and G1 in CDCl₃. A 1:1 stoichiometry for the complexation was obtained from a mole ratio plot based on the chemical shift changes of H_a on **BrP5**. The association constant (K_a) was estimated to be 967 \pm 72 M^{-1} in chloroform-d using the curve-fitting analysis method. The formation of the [2]pseudorotaxane was further confirmed by a low-resolution electrospray ionization mass spectrum (ESI-MS). The spectrum (ESI, Fig. S10†) showed a peak at m/z 1776.4, corresponding to $[BrP5 \supset G1 - I]^+$, which was in agreement with the result of the above-mentioned ¹H NMR experiments. All these phenomena provided strong evidence for the formation of a stable [2] pseudorotaxane in a 1:1 host-to-guest stoichiometry between BrP5 and G1 in solution.

To gain a deep insight into the [2]pseudorotaxane and its binding properties, single crystals suitable for X-ray diffraction analysis were obtained by slow diffusion of isopropyl ether into a solution of BrP5 and G1 in chloroform (ESI, Table S1†). The single crystal structure between the pillararene and the ionic liquid was obtained, providing direct evidence for their host-guest complexation behavior. 49-51 As shown in Fig. 3, the

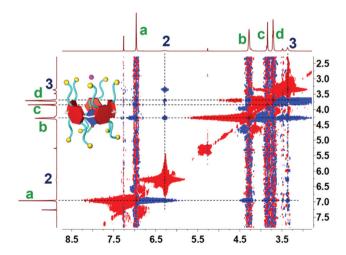


Fig. 2 Partial 2D ¹H-¹H NOESY spectrum (500 MHz, CDCl₃, room temperature) of the [2]pseudorotaxane BrP5 G1. [BrP5] = 10.0 mM. [G1]

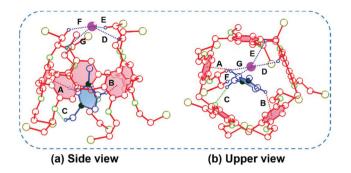


Fig. 3 Ball-stick views of the crystal structure of BrP5⊃G1: (a) side view and (b) upper view. Host BrP5 is red, guest G1 is blue, hydrogen atoms are blue, oxygen atoms are green, nitrogen atoms are solid and dark green, bromine atoms are dark yellow, and iodine atoms are solid and magenta. The red dashed lines indicate $C-H\cdots\pi$ interactions (A and B), the green dashed line indicates C-H...O interaction (C) and the blue dashed lines indicate $C-H\cdots I^-$ interactions (D, E, F and G). $C-H\cdots \pi$ interaction parameters are as follows: $H \cdots \pi$ -plane distances (angstroms) and $C-H\cdots\pi$ angles (degrees): A, 2.48, 152.28; B, 2.72, 141.30. $C-H\cdots O$ interaction parameters: C-H...O distance (angstrom) and C-H...O angle (degree): C, 2.60, 139.21. C-H···I⁻ interaction parameters: C-H···I⁻ distances (angstroms): D, 2.98; E, 2.98; F, 2.71; and G, 3.14. C-H····I⁻ angles (degrees): D, 126.45; E, 112.06; F, 134.56; and G, 116.83. Hydrogen atoms that do not have interactions with other atoms are removed for

[2]pseudorotaxane was indeed formed with the imidazolium part of G1 located in the central cavity of the host BrP5, which was in accordance with the NMR experimental results in solution. In this crystal structure, the imidazolium part inside the cavity established a couple of C-H $\cdots\pi$ interactions (A and B) with the benzene rings of BrP5 with $H \cdots \pi$ plane distances of 2.48 Å and 2.72 Å and C-H... π angles of 152.28° and 141.30°, respectively. Furthermore, a C-H···O interaction (C) was also observed between a hydrogen atom on G1 and an oxygen atom on BrP5 with a C-H···O distance of 2.60 Å and an angle of 139.21°. These interactions were believed to contribute to the stability of this [2]pseudorotaxane. Another essential contribution to the formation of these structures was that the iodide anion was captured in the cavity of BrP5, mainly stabilized by four C-H···I⁻ interactions (D, E, F and G). The distances between the iodide anion and bromoethoxy hydrogen atoms of BrP5 were calculated to be 2.98, 2.98, 2.71 and 3.14 Å, respectively, which were smaller than the sum of the van der Waals radii of the iodide anion and the hydrogen atom (3.40 Å), suggesting the presence of C–H \cdots I $^-$ interactions. 52,53

Besides, it should be noted that the binding affinity of the [2]pseudorotaxane BrP5 \(\)G1 was promoted and enhanced by the ion-pair recognition, and different counterions of the imidazolium-based ionic liquid have a huge influence on the association behavior. To prove this, imidazolium guests G2 with a large tetrafluoroborate anion and G3 with a dimethylphosphate anion were used as guests to investigate the binding properties of BrP5. The association constant K_a between G2 and BrP5 was determined to be $89 \pm 20 \text{ M}^{-1}$ in chloroform-d, which was smaller than that of **BrP5**⊃**G1** (ESI,

Fig. S6-S8†). Besides, in the titration process of a solution of G3 with a larger dimethylphosphate anion into BrP5, no significant shifts for the host protons Ha were observed even after the addition of 6.0 equiv. of G3 (ESI, Fig. S9†). This observation implies very weak complexation between BrP5 and G3. Accordingly, the association constant could not be calculated for this complex. These results indicated that the counterions on the guests played significant roles in the host-guest association process, and demonstrated that the ion-pair recognition showed a positive effect during molecular recognition.

Intriguingly, after the establishment of the ion-pair recognition motif between BrP5 and G1, the [2]pseudorotaxane showed supramolecular interaction-induced LCST phase behavior. Next, we aimed to investigate its functionality in LCST behavior. Firstly, due to the relatively poor solubility of the ionic liquid G1 in chloroform at room temperature, an immiscible solution was obtained by mixing a relatively high concentration of G1 with chloroform. In Fig. 4a, we can see that the mixture transformed from two separated phases into one transparent solution after adding BrP5 to establish the hostguest complexation with G1. The possible reason is that the host-guest complex has better solubility in chloroform. Notably, the solution exhibited a significant thermoresponsive behavior. The clear solution became opaque on heating and then turned back to transparent again upon cooling (Fig. 4a). The course of change was fully reversible. The explanation of

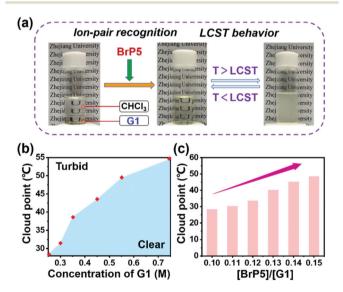
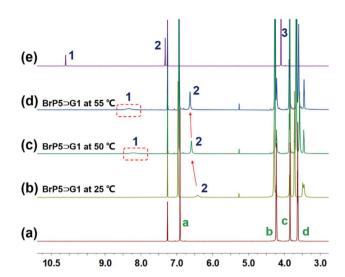


Fig. 4 (a) Illustration of host-guest ion-pair recognition and supramolecular interaction induced LCST behavior of the [2]pseudorotaxane BrP5 G1: the transparent solution was formed after the addition of BrP5 into the two immiscible phases of G1 and chloroform and its LCST-type phase changes. (b) Concentration dependence of the cloud point of G1 in the presence of BrP5 in chloroform. The initial concentrations of G1 were 0.25 M, 0.30 M, 0.35 M, 0.45 M, 0.55 M and 0.75 M, from left to right, respectively, and the molar ratio of BrP5 to G1 was kept constant at 0.10. (c) Change of the cloud point upon altering the molar ratio of BrP5 to G1 (0.10, 0.11, 0.12, 0.13, 0.14 and 0.15, from left to right, respectively, while the concentration of G1 was maintained at 0.25 M).



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Fig. 5 Partial temperature-dependent ¹H NMR spectra (600 MHz, CDCl₃): (a) BrP5 at 25 °C; (b) BrP5⊃G1 at 25 °C; (c) BrP5⊃G1 at 50 °C; (d) BrP5⊃G1 at 55 °C; and (e) G1 at 25 °C.

such a tunable LCST behavior was ascribed to the fact that the complexation affinity between BrP5 and G1 in the [2]pseudorotaxane was temperature dependent. When the temperature increased, the interactions between BrP5 and G1 were weakened or even destroyed, and the disassembly of the [2]pseudorotaxane resulted in the release of G1 from the cyclic cavity of **BrP5** to the chloroform solution. Temperature-variable ¹H NMR experiments (Fig. 5) showed that as the temperature increased, changes occurred in the chemical shifts related to the protons on both BrP5 and G1. In particular, the signals of protons H₁ on G1 of the [2]pseudorotaxane gradually appeared, which were completely shielded by the cyclic structure before.

The concentration-dependent characteristic of the complexation-induced LCST behavior was investigated next. As the concentration of G1 increased from 0.25 M to 0.75 M (the molar ratio of BrP5 to G1 was kept constant at 0.10), the cloud point ($T_{\rm cloud}$) gradually increased from 28.4 to 54.6 °C (Fig. 4b and ESI, Fig. S11a†). Furthermore, the system also showed molar ratio-dependent behavior (Fig. 4c and ESI, Fig. S11b†). While we fixed the concentration of G1 at 0.25 M, upon altering the molar ratio of BrP5 to G1 from 0.10 to 0.15, the $T_{\rm cloud}$ changed accordingly from 28.4 to 48.4 °C. Besides, DLS results revealed that the size of the nanostructure of this LCST system changed immensely at different temperatures. For example, a system with the T_{cloud} found to be around 28.4 °C by monitoring the transmittance change showed average sizes of 7.87 \pm 0.89 nm and 703 \pm 85.1 nm below and above the LCST, respectively (ESI, Fig. S12†).

Conclusions

In summary, we developed a novel [2] pseudorotaxane based on perbromoethylated pillar[5]arene/imidazolium iodide ionic liquid ion-pair recognition in both solution and the solid state. The introduction of ion-pair recognition contributed greatly to the stability of the [2]pseudorotaxane. In addition, experimental results demonstrated that the thermosensitive phase behavior of the [2]pseudorotaxane can be affected by the molar ratio of the host to guest at a certain concentration or the concentrations of these two components at a certain molar ratio. This work not only enriched pillararene-based threaded structures but also constructed a smart functional supramolecular material. The exploration of such interesting pseudorotaxanes with extra functions will provide some helpful hints for the design of more advanced functional materials such as molecular sensors^{54,55} and stimuli-responsive functional supramolecular systems.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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