

Confined growth of pyridinic N–Mo<sub>2</sub>C sites on MXenes for hydrogen evolution†Cite this: *J. Mater. Chem. A*, 2020, 8, 7109Received 12th February 2020  
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Developing low-cost and high-performance hydrogen evolution reaction (HER) electrocatalysts is a key research area for scalable hydrogen production from water electrolysis. Here, a hybrid of nitrogen-doped carbon encapsulated Mo<sub>2</sub>C nanodots on Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene (Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC) is developed through *in situ* polymerization of dopamine and a Mo precursor on the Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene surface. During the annealing treatment, the polydopamine plays multiple roles in forming N-doped carbon, confining MoO<sub>4</sub><sup>2-</sup> ions into ultrasmall Mo<sub>2</sub>C nanodots, and stabilizing the MXene flakes against spontaneous oxidation. The as-synthesized hybrid exhibits excellent HER activity in acidic electrolyte with an overpotential of 53 mV at 10 mA cm<sup>-2</sup> and excellent stability over 30 hours. The combination of experiments and simulations demonstrates that pyridinic N-doped carbon coated Mo<sub>2</sub>C nanodots serve as the active sites and Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene facilitates the charge transfer, synergistically contributing to the superior HER performance.

Owing to the high energy density and zero carbon emission, hydrogen is considered as an optimal solution to address the energy crisis and alleviate the environmental pollution.<sup>1,2</sup> Although industrial hydrogen production still relies on fossil fuels, water electrolysis shows promise as a clean and sustainable route for hydrogen generation.<sup>3,4</sup> Thus, developing efficient hydrogen evolution reaction (HER) electrocatalysts is of great

importance to realize scalable electrocatalytic hydrogen production. Until now, platinum (Pt) and its derivatives remain the best HER electrocatalysts with low overpotentials and fast reaction kinetics, but their low abundance and high cost severely hinder their large-scale use.<sup>5</sup> Recently, intensive research efforts have been made to explore low-cost and earth-abundant alternatives with comparable HER performance.<sup>6-15</sup> Among them, molybdenum carbide (Mo<sub>2</sub>C) holds great promise due to its Pt-like d-band electronic structure.<sup>16</sup> In general, the HER activity of Mo<sub>2</sub>C mainly relies on the accessibility of active sites and conductivity. Mo<sub>2</sub>C nanoparticles, even ultrasmall nanodots, are preferable but tend to agglomerate when processed into electrodes. In this context, introducing a highly conductive substrate that could simultaneously optimize the dispersion of ultrasmall Mo<sub>2</sub>C would be of importance.<sup>17,18</sup>

Recently, an emerging family of two-dimensional (2D) transition metal carbides and nitrides, called MXenes, have attracted much attention due to their unique characteristics of metallic conductivity (up to 10 000 S cm<sup>-1</sup>), hydrophilicity and diversity stemming from various surface terminations.<sup>19,20</sup> Typically, the most studied Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene (T<sub>x</sub> denotes the surface –OH, –O and –F groups) shows great promise in various applications including supercapacitors,<sup>21</sup> sensors,<sup>22</sup> water purification,<sup>23</sup> electromagnetic interference shielding<sup>24</sup> and catalysis.<sup>25-29</sup> It would be natural to assume that coupling Mo<sub>2</sub>C and MXenes may deliver satisfactory HER performance, while it remains challenging due to the following factors. First, the synthesis of Mo<sub>2</sub>C always requires high annealing temperature. Proverbially, MXenes with high surface energy are prone to oxidation at evaluated temperatures, although protected by an inert atmosphere, leading to inferior conductivity.<sup>30</sup> In addition, the high synthesis temperature would accelerate the agglomeration of Mo<sub>2</sub>C nanoparticles and thus result in poor accessibility of active sites.<sup>31</sup> Accordingly, protecting MXenes from surface oxidation and Mo<sub>2</sub>C nanoparticles from aggregation during the calcination process is of great importance to realize MXene–Mo<sub>2</sub>C hybrid HER electrocatalysts.

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Here, we report a 2D coupled nanohybrid of Mo<sub>2</sub>C nanodots and Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene encapsulated by nitrogen-doped carbon layers (denoted as Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC) synthesized using molybdate ions chelated in polydopamine (Mo-PDA) as the precursor (Fig. 1). Notably, Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> flakes are fully capped with Mo-PDA by *in situ* polymerization, which could inhibit the surface oxidation of Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> during annealing treatment and thus retain the excellent conductivity.<sup>32</sup> Moreover, due to the confinement effect of the Mo-PDA structure, the ultrasmall Mo<sub>2</sub>C nanodots are uniformly dispersed in the carbon matrix without agglomeration. Demonstrated by theoretical and experimental analyses, the highly accessible pyridinic N-Mo<sub>2</sub>C sites coupled with the metallic Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> support of Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC can deliver remarkable HER activities with an overpotential of 53 mV at 10 mA cm<sup>-2</sup>, a Tafel slope of 40 mV dec<sup>-1</sup> and excellent stability over 30 hours in acidic electrolyte, which are superior to those of reported noble metal-free electrocatalysts.

As demonstrated previously, PDA may have a positive confinement effect on the dispersion of Mo<sub>2</sub>C nanodots, in which a hierarchical microflower carbon structure was formed (Fig. S1†). Even though Mo<sub>2</sub>C nanodots may be well dispersed, such a microflower structure would show inferior HER performance in terms of low conductivity. We assume that 2D Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene with various terminations may be a good candidate as a metallically conductive substrate to form a flat PDA layer, realizing both well dispersed Mo<sub>2</sub>C nanodots and fast electron transport. To verify this idea, we first check the morphology of polydopamine on the MXene by *in situ* polymerization. The ultrathin Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> with a lateral size of several microns exhibits single crystallinity (Fig. S2†). By *in situ* polymerization of dopamine on the surface of Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene, the polydopamine/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> (denoted as PDA/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>) shows a similar morphology to the pristine MXene (Fig. 2a). After annealing at 750 °C, the pristine MXene is severely oxidized with many TiO<sub>2</sub> nanoparticles/nanorods generated on the surface (Fig. S3a†), while the N-doped carbon coated Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene (denoted as Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC) derived from PDA/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> retains the flat surface without the presence of TiO<sub>2</sub> particles (Fig. 2b), which is confirmed from the enlarged TEM image (Fig. S3b†). The high-resolution TEM (HRTEM) image (Fig. 2c) clearly shows that the Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene layers are fully covered by carbon.

The structures of these samples were further investigated by X-ray diffraction (XRD) and Raman spectroscopy. The pristine MXene exhibits the typical peak at ~7° corresponding to the (002) plane of Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene (Fig. S4a†).<sup>23</sup> The (002) peak of PDA/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> shifts to ~6° due to the expanded interlayer spacing of Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> layers. The annealed pristine MXene does not show characteristic Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene peaks but shows typical peaks of rutile TiO<sub>2</sub> (JCPDS no. 21-1276).<sup>34</sup> Notably, no TiO<sub>2</sub> peaks are observed in the Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC XRD pattern. In addition, the (002) peak of Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene disappeared, which should be attributed to the suppressed restacking of MXene flakes by carbon layers on the surface.<sup>35</sup> The Raman spectra (Fig. S4b†) show that all the pristine Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>, PDA/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> and Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC samples exhibit typical peaks corresponding to Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene while annealed Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> shows the characteristic peaks at 237.7, 446.9 and 609.8 cm<sup>-1</sup> of rutile TiO<sub>2</sub>.<sup>36</sup> In addition, X-ray photoelectron spectroscopy (XPS) analyses (Fig. S5†) confirm the oxidation of Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene into TiO<sub>2</sub> during the annealing process. Based on the above results, it can be concluded that dopamine self-polymerized on the surface of the MXene could protect it from oxidation during the annealing treatment by *in situ* forming N-doped carbon layers.

Previous studies have reported that molybdate ions can react with dopamine and then polymerize into the Mo-PDA complex, which can be used to fabricate nanostructured Mo compound/carbon composites for various applications.<sup>33,37,38</sup> Based on this inspiration, here we prepared Mo-PDA/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> through a facile *in situ* polymerization method, as depicted in Fig. 1. The SEM and TEM images (Fig. S6†) of Mo-PDA/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> show a similar ultrathin morphology to PDA/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>, confirming that Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> flakes are tightly covered by Mo-PDA layers. Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC was obtained by annealing Mo-PDA/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> under Ar at 750 °C. As shown in Fig. 2d and Fig. S7,† the morphology of Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC is similar to that of Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC (Fig. 2b). The enlarged TEM image (Fig. 2e) evidences that high-density ultrasmall Mo<sub>2</sub>C nanodots with an average diameter of 1.1 nm (Fig. 2f) are uniformly anchored on the surface of Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> flakes. The high-resolution TEM image (Fig. 2g) exhibits the lattice fringes of ~0.23 nm on the nanodots, corresponding to the *d*-spacing of the β-Mo<sub>2</sub>C (101) plane.<sup>17</sup> The background lattices belong to Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>, confirming its hexagonal structure. The high-angle annular dark field scanning TEM (HAADF-STEM) image (Fig. 2h) reveals numerous bright dots

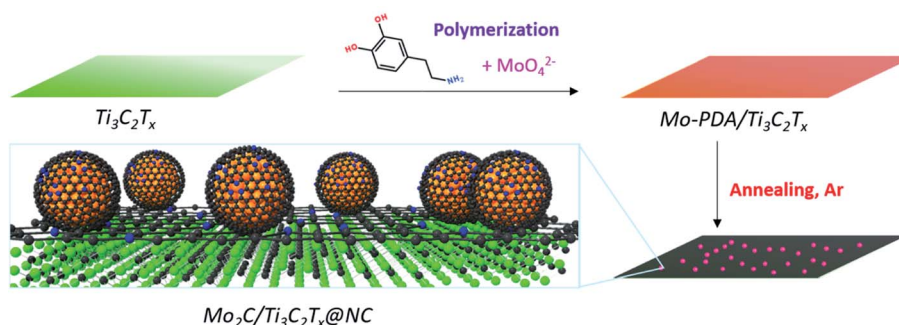


Fig. 1 Synthetic illustration of the preparation of Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC.

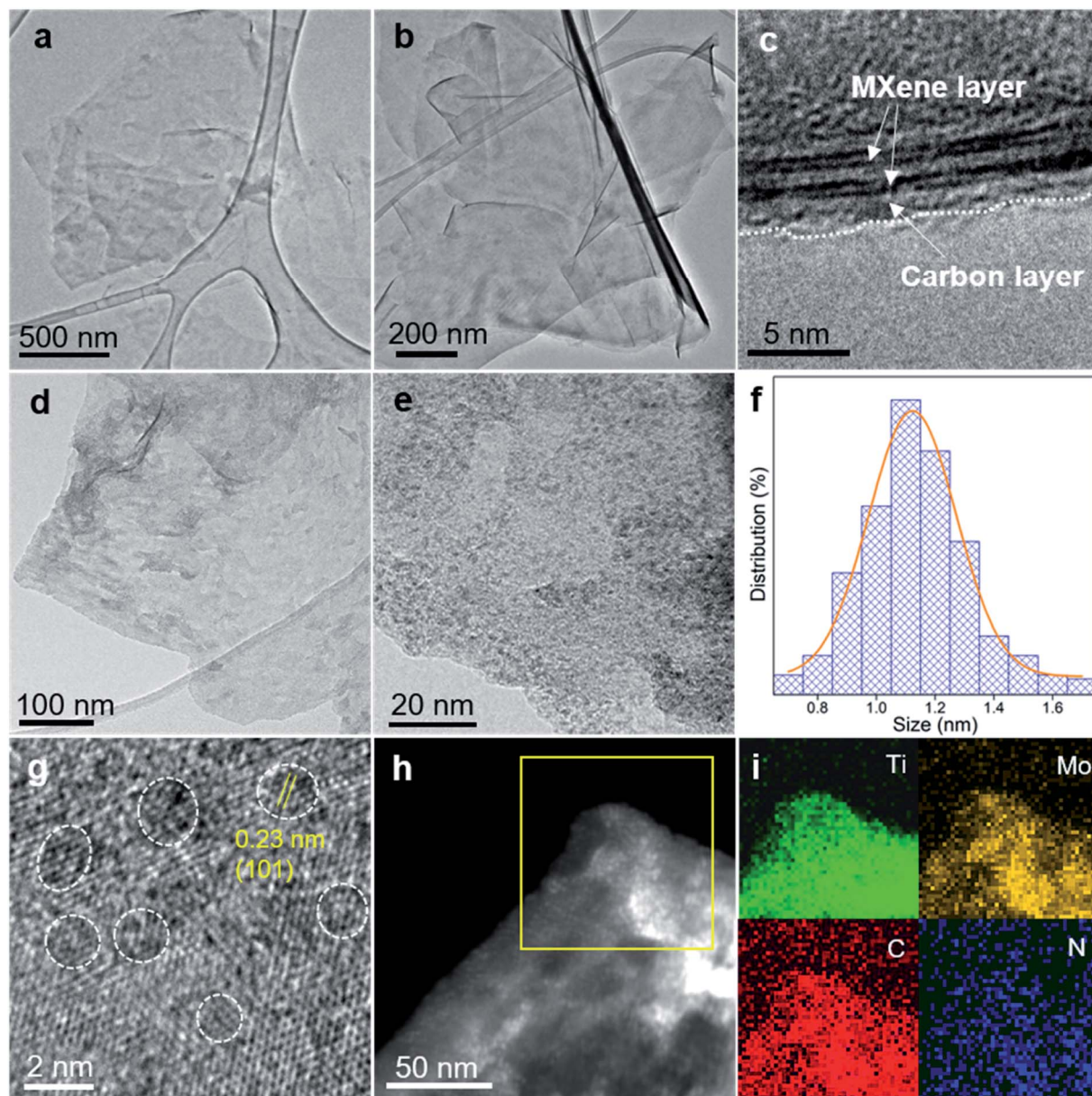


Fig. 2 TEM images of (a) PDA/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>, (b) Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC and (d) Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC. (c) HRTEM image of a Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC edge showing the carbon layer covered Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene structure. (e) Enlarged TEM image of Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC showing the uniform dispersion of Mo<sub>2</sub>C nanodots on the surface of Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene. (f) Size distribution of Mo<sub>2</sub>C nanodots in Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC counted in (e). HRTEM image (g) and HAADF-STEM image (h) of Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC. (i) EDX elemental mapping from the marked region in (h).

decorated on the MXene sheets, further demonstrating the uniform dispersion of Mo<sub>2</sub>C nanodots. The corresponding energy dispersive X-ray spectroscopy (EDX) elemental mapping images (Fig. 2i) confirm the distribution of Ti, Mo, C and N elements on the Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC surface.

Fig. 3a displays the XRD patterns of the samples. The characteristic peaks of both Mo-PDA and the MXene are observed in the XRD pattern of Mo-PDA/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>. Similar to PDA/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>, the (002) peaks of Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> negatively shift due to the expanded interlayer spacing. The absence of the (002) peak in Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC results from the separation of MXene layers by the Mo<sub>2</sub>C nanodots and carbon layers. Through investigation using

N<sub>2</sub> adsorption-desorption isotherms, as shown in Fig. 3b, the Brunauer-Emmett-Teller (BET) surface areas of Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC and Mo<sub>2</sub>C@NC were calculated to be 241.4 and 86.4 m<sup>2</sup> g<sup>-1</sup>, respectively. Both samples exhibit type IV isotherms of mesoporous structures, with pores in the range from 2 to 50 nm. The larger surface area and mesoporous structure of Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC could facilitate the accessibility of active sites and charge transfer during the HER process.

The valence states and compositions of Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC were elucidated by XPS analyses. The XPS full scan (Fig. S8a†) evidences the existence of Ti, Mo, C, N, and O in Mo<sub>2</sub>C/Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub>@NC. The Ti 2p core level can be fitted with four doublets for

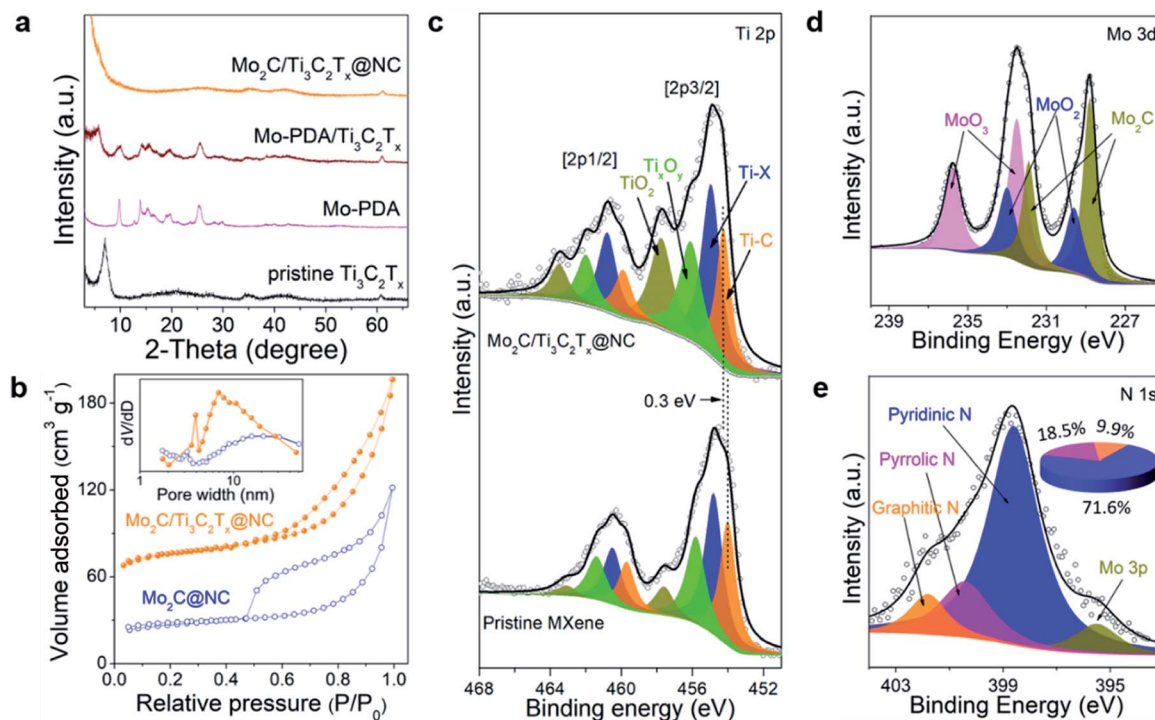


Fig. 3 (a) XRD patterns of pristine  $\text{Ti}_3\text{C}_2\text{T}_x$  MXene, Mo-PDA, Mo-PDA/ $\text{Ti}_3\text{C}_2\text{T}_x$  and  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$ . (b)  $\text{N}_2$  adsorption–desorption isotherms and the corresponding pore size distributions of  $\text{Mo}_2\text{C}@NC$  and  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$ . High-resolution XPS spectra of (c) Ti 2p, (d) Mo 3d and (e) N 1s of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$ . The Ti 2p survey of pristine  $\text{Ti}_3\text{C}_2\text{T}_x$  MXene is presented in (c) for comparison. The inset in (e) is the calculated contents of pyridinic N, pyrrolic N and graphitic N in  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$ .

Ti 2p<sub>1/2</sub> and Ti 2p<sub>3/2</sub>, as shown in Fig. 3c. Typically, the Ti 2p<sub>3/2</sub> components at 454.3, 455, 456.1, and 457.8 eV are assigned to Ti–C (Ti<sup>+</sup>), Ti–X (Ti<sup>2+</sup>), Ti<sub>x</sub>O<sub>y</sub> (Ti<sup>3+</sup>), and TiO<sub>2</sub> (Ti<sup>4+</sup>), respectively.<sup>22</sup> Compared to the pristine  $\text{Ti}_3\text{C}_2\text{T}_x$  MXene, the binding energy of Ti–C positively shifts by about 0.3 eV, indicating the electron transfer from  $\text{Ti}_3\text{C}_2\text{T}_x$  to  $\text{Mo}_2\text{C}$  nanodots in the hybrid.<sup>26</sup> The deconvoluted Mo 3d XPS spectrum (Fig. 3d) reveals that the peaks at 228.8 eV and 231.9 eV are attributed to  $\text{Mo}_2\text{C}$ . Meanwhile, due to inevitable surface oxidation, the peaks at 229.6, 232.5, 233, and 235.8 eV stem from  $\text{Mo}^{4+} 3d_{5/2}$ ,  $\text{Mo}^{6+} 3d_{5/2}$ ,  $\text{Mo}^{4+} 3d_{3/2}$  and  $\text{Mo}^{6+} 3d_{3/2}$ , respectively.<sup>33</sup> As shown in Fig. 3e, the deconvolution of the N1s XPS spectrum reveals the peaks at 398.6, 400.5 and 401.8 eV, which could be assigned to pyridinic, pyrrolic and graphitic N, respectively. Accordingly, the content is determined to be 71.64%, 18.43% and 9.93%, respectively. The C 1s spectrum (Fig. S8b†) was fitted with five components at 282, 284.6, 285.5, 287.1, and 289.1 eV, stemming from Ti–C, C–C, C–N, C–O, and –COO, respectively. Upon comparison, the XPS spectra of  $\text{Mo}_2\text{C}@NC$  are similar to those of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$  (Fig. S9†). The above XPS analyses imply that  $\text{Mo}_2\text{C}$  nanocrystals are intensely coupled on  $\text{Ti}_3\text{C}_2\text{T}_x$  MXene and encapsulated in the pyridinic N-rich carbon layer.

The HER performance of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$  was evaluated by investigating the polarization curves in 0.5 M  $\text{H}_2\text{SO}_4$  with all the potentials converted to the reversible hydrogen electrode (RHE). For comparison,  $\text{Mo}_2\text{C}@NC$ ,  $\text{Ti}_3\text{C}_2\text{T}_x@NC$  and commercial 20 wt% Pt/C catalyst were also measured. As shown in Fig. 4a, the commercial 20 wt% Pt/C exhibits the best activity

with an onset overpotential of zero.  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$  shows a low onset overpotential of 6 mV vs. RHE (Fig. S10a†), which is very comparable to that of Pt/C but much lower than that of  $\text{Mo}_2\text{C}@NC$  (53 mV) and  $\text{Ti}_3\text{C}_2\text{T}_x@NC$  (432 mV). Impressively,  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$  shows better performance than Pt/C at higher overpotentials (>106 mV). To elucidate the HER mechanism, Tafel plots of these catalysts were calculated based on the Tafel equation  $\eta = a + b \log|j|$ , where  $j$  represents the current density and  $b$  is the Tafel slope.<sup>39</sup> As shown in Fig. 4b, the Tafel slope of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$  is determined to be 40 mV dec<sup>-1</sup>, close to that of Pt/C (31 mV dec<sup>-1</sup>), and lower than those of  $\text{Mo}_2\text{C}@NC$  (67 mV dec<sup>-1</sup>) and  $\text{Ti}_3\text{C}_2\text{T}_x@NC$  (113 mV dec<sup>-1</sup>). The value implies that the HER operates through the Volmer–Heyrovsky mechanism on  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$ .<sup>40</sup> The exchange current density ( $j_0$ ) for these catalysts was determined by fitting the linear portion of the Tafel plots at low current density (Fig. S10b†).<sup>41</sup> As displayed in Fig. 4c, the  $j_0$  value of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$  is 0.52 mA cm<sup>-2</sup>, smaller than that of Pt/C (0.841 mA cm<sup>-2</sup>) but much higher than those of  $\text{Mo}_2\text{C}@NC$  (0.127 mA cm<sup>-2</sup>) and  $\text{Ti}_3\text{C}_2\text{T}_x@NC$  ( $2.6 \times 10^{-5}$  mA cm<sup>-2</sup>), indicating the superior intrinsic activity of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$ . To achieve a current density of 10 mA cm<sup>-2</sup>, the overpotentials of 27, 53, 121, and 631 mV are required for Pt/C,  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$ ,  $\text{Mo}_2\text{C}@NC$  and  $\text{Ti}_3\text{C}_2\text{T}_x@NC$ , respectively. As listed in Table S1,† such a low overpotential indicates that the  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$  exhibits one of the best HER performances among those of previously reported non-noble-metal electrocatalysts, such as MoC– $\text{Mo}_2\text{C}/\text{PNCDS}$ ,<sup>9</sup>  $\text{MoS}_2/\text{Ti}_3\text{C}_2\text{T}_x$ ,<sup>26</sup> Co-SAC,<sup>10</sup> and 1T′-



Fig. 4 (a) Polarization curves and (b) corresponding Tafel plots of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$ ,  $\text{Mo}_2\text{C}@\text{NC}$ ,  $\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$  and commercial 20% Pt/C in 0.5 M  $\text{H}_2\text{SO}_4$ . (c) Comparison of overpotentials at 10  $\text{mA cm}^{-2}$  ( $\eta_{10}$ ) and exchange current densities ( $j_0$ ) of the four catalysts. (d) Comparison of  $\Delta G_{\text{H}^*}$  of p-N-C, g-N-C,  $\text{Mo}_2\text{C}$ ,  $\text{Ti}_3\text{C}_2\text{T}_x$  as well as their heterostructures. (e) Time-dependent current density curve of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$  at a constant overpotential of 100 mV. The inset gives the polarization curves of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$  initially and after 1000 CV cycles.

ReSe QDs.<sup>42</sup> Moreover,  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$  also demonstrated excellent HER performance in basic (1 M KOH) and neutral (1 M PBS) electrolytes with an overpotential of 75 and 114 mV at 10  $\text{mA cm}^{-2}$  and a Tafel slope of 59.2 and 80.3  $\text{mV dec}^{-1}$ , respectively, comparable to that of state-of-the-art non-noble-metal electrocatalysts (Fig. S11<sup>†</sup>).<sup>29,43–45</sup>

The above results imply that  $\text{Ti}_3\text{C}_2\text{T}_x$  itself shows negligible HER activity but coupling of  $\text{Mo}_2\text{C}$  with conductive  $\text{Ti}_3\text{C}_2\text{T}_x$  can modulate the morphological structure to optimize the HER activity. The Nyquist plots of the electrochemical impedance spectroscopy (EIS) responses (Fig. S12<sup>†</sup>) show a much smaller charge transfer resistance for  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$  than that for  $\text{Mo}_2\text{C}@\text{NC}$  during the HER process, evidencing the faster HER kinetics due to the introduction of the conductive  $\text{Ti}_3\text{C}_2\text{T}_x$  support.<sup>46</sup> To further understand the intrinsic origin, the electrochemical surface area (ECSA) and per-site turnover frequency (TOF) of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$  and  $\text{Mo}_2\text{C}@\text{NC}$  were typically compared. The ECSA was estimated from the electrochemical double-layer capacitance ( $C_{\text{dl}}$ ) by performing cyclic voltammetry in the region from 0.2 to 0.3 V at rates varying from 10 to 100  $\text{mV s}^{-1}$  (Fig. S13<sup>†</sup>). The ECSA of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$  (603.75  $\text{cm}^2$ ) is larger than that of  $\text{Mo}_2\text{C}@\text{NC}$  (539  $\text{cm}^2$ ), indicating that the former has a higher number of active sites. As the  $\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$  has very poor HER activity, the active sites should arise from  $\text{Mo}_2\text{C}$ . Here, we assume that all the Mo atoms in the samples are active sites, which is determined by inductively coupled plasma optical emission spectroscopy (ICP-OES).  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$

has a TOF of 0.246  $\text{s}^{-1}$  at an overpotential of 100 mV (Fig. S14<sup>†</sup>), much higher than that of  $\text{Mo}_2\text{C}@\text{NC}$  (0.014  $\text{s}^{-1}$ ), which is comparable to those of reported non-noble-metal HER electrocatalysts, such as 3DHP- $\text{Mo}_2\text{C}$  (0.045  $\text{s}^{-1}$ )<sup>47</sup> and  $\text{CoN}_x/\text{C}$  (0.39  $\text{s}^{-1}$ ).<sup>48</sup>

The electrochemical results reveal that  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$  exhibits superior HER activity to that of its counterparts. To investigate the underlying mechanism, we performed density functional theory (DFT) calculations (Fig. S15–S17<sup>†</sup>). Here we considered a simple HER pathway containing three states: the initial state of  $\text{H}^+ + \text{e}^-$ , the intermediate state of adsorbed H ( $\text{H}^*$ ), and the final state of  $\text{H}_2$  molecule.<sup>39</sup> The Gibbs free energy of the  $\text{H}^*$  adsorption ( $\Delta G_{\text{H}^*}$ ) was calculated on the considered systems. Generally, a high-performance HER catalyst has a  $\Delta G_{\text{H}^*}$  value close to zero.<sup>41</sup> Fig. 4d shows the calculated free energy diagrams for the HER. Pristine  $\text{Mo}_2\text{C}$  and  $\text{Ti}_3\text{C}_2\text{T}_x$  MXene have a  $\Delta G_{\text{H}^*}$  of  $-0.78$  and  $-0.43$  eV, respectively, suggesting their strong interactions with H and consequently poor HER performance. A previous study demonstrated that pyridinic N and graphitic N in carbon contribute to the HER activity of composite catalysts.<sup>49</sup> The modeled pyridinic N doped graphene (p-N-C) and graphitic N doped graphene (g-N-C) possess a  $\Delta G_{\text{H}^*}$  of 0.49 and 0.69 eV, respectively, which implies their unfavorable interactions with H and poor HER activities. Coupling N-doped carbon with  $\text{Mo}_2\text{C}$  could lead to enhanced HER performance. Typically, the  $\Delta G_{\text{H}^*}$  values of  $\text{Mo}_2\text{C}@\text{p-N-C}$  and  $\text{Mo}_2\text{C}@\text{g-N-C}$  were 0.22 and 0.44 eV, respectively, which were

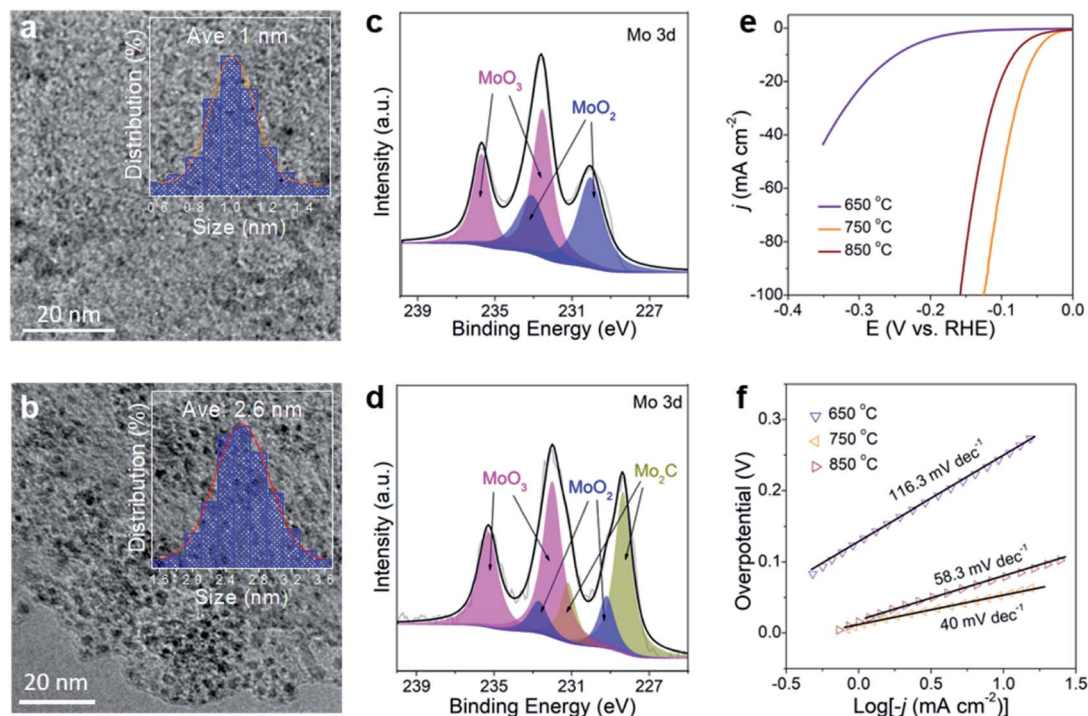


Fig. 5 TEM images of (a) M-650 and (b) M-850. The inset gives the size distributions of nanoparticles in the samples. High-resolution XPS Mo 3d spectra of (c) M-650 and (d) M-850. Polarization curves (e) and Tafel plots (f) of M-650, M-750, and M-850.

lower than those without N-doped carbon. Meanwhile, N-doped carbon coupled with  $\text{Ti}_3\text{C}_2\text{T}_x$  MXene shows high  $\Delta G_{\text{H}^*}$  with unfavorable HER activity. The DFT results are in good agreement with the above experimental results, revealing that  $\text{Mo}_2\text{C}@p\text{-N-C}$  serves as the active sites while the MXene support facilitates the charge transfer. Notably, Chen *et al.*<sup>50</sup> reported an efficient HER catalyst of Mo single atoms anchored on N-doped carbon synthesized by an appropriate annealing method, which could be formed in the as-prepared  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$ . Accordingly, the Mo–N–C active site, determined as  $\text{Mo}_1\text{N}_1\text{C}_2$ , is also considered. As shown in Fig. S18 and S19,† both  $\text{Mo}_1\text{N}_1\text{C}_2$  and  $\text{Mo}_1\text{N}_1\text{C}_2/\text{Ti}_3\text{C}_2\text{T}_x$  exhibit extremely low  $\Delta G_{\text{H}^*}$  values, suggesting that they are highly active toward the HER. The density of state near the Fermi level for  $\text{Mo}_1\text{N}_1\text{C}_2/\text{Ti}_3\text{C}_2\text{T}_x$  is higher than that of  $\text{Mo}_1\text{N}_1\text{C}_2$ , indicating that the  $\text{Ti}_3\text{C}_2\text{T}_x$  MXenes facilitate the charge transfer during the HER process.<sup>51</sup>

The stability of HER electrocatalysts is also a critical indicator for practical applications. As shown in Fig. 4e, the time-dependent current density curve of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$  at a constant overpotential of 100 mV shows a slight decrease over 30 h, which is much better than that of commercial 20 wt% Pt/C, demonstrating excellent durability. In addition, the stability of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$  was studied using CV cycles in the range of 0.1 to  $-0.2$  V vs. RHE. The polarization curve of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$  after 1000 cycles (inset of Fig. 4e) shows a negligible decrease in current density compared to the initial curve, implying its long-term stability. Detailed characterization of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$  after the stability test (Fig. S21†) further confirms the outstanding durability of its morphology and

structure, which could be attributed to the protection of the covered carbon layers.

In controlled experiments, the effect of carbonization temperature (650 °C and 850 °C) on HER performance was investigated. The samples are named M-650 and M-850. TEM images (Fig. 5a, b and S22†) reveal that  $\text{MoO}_x$  nanodots in M-650 (not completely carbonized) have an average size of 1 nm while the size of  $\text{Mo}_2\text{C}$  nanodots in M-850 is 2.6 nm. It is in good agreement with a previous report which showed that the particle size would increase with the carbonization temperature.<sup>52</sup> The XRD patterns of the two samples are similar to that of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$  (here defined as M-750) except for the presence of  $\text{TiO}_2$  signals in M-850 (Fig. S23†). The XPS results (Fig. 5c) suggest that M-650 contains  $\text{MoO}_x$  rather than  $\text{Mo}_2\text{C}$ , which could arise from the insufficient temperature. Meanwhile, as shown in Fig. 5d, M-850 has a similar Mo 3d XPS spectrum to M-750, suggesting the formation of  $\text{Mo}_2\text{C}$  nanodots. As shown in Fig. 5e, the overpotentials of M-650 and M-850 for achieving  $10 \text{ mA cm}^{-2}$  are 248 and 83 mV, respectively. And the corresponding Tafel slopes are 116.3 and 58.3  $\text{mV dec}^{-1}$ , respectively (Fig. 5f). Obviously, M-750 exhibits the best HER activity among the three samples, which could be attributed to the following two factors: the  $\text{MoO}_x$  generated at a lower temperature (650 °C) has poorer intrinsic activity than  $\text{Mo}_2\text{C}$ , and the higher temperature (850 °C) induces aggregation of  $\text{Mo}_2\text{C}$  nanodots and partial oxidation of the MXene which would reduce the density of active sites and electron conductivity.

Based on the aforementioned considerations, the extraordinary HER performance of  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@\text{NC}$  is postulated to

arise from the following synergistic aspects: (1) *in situ* polymerization of dopamine retains the structural stability of  $\text{Ti}_3\text{C}_2\text{T}_x$  MXene at high carbonization temperature, facilitating the charge transfer; (2) polydopamine chelated  $\text{MoO}_4^{2-}$  delivers highly disperse ultrasmall  $\text{Mo}_2\text{C}$  nanodots on the surface of the MXene, favoring the high accessibility of active sites for the HER; (3) the pyridinic N dopants in carbon coupled with  $\text{Mo}_2\text{C}$  induce activated active sites for the HER; (4) the intensely coupled  $\text{Mo}_2\text{C}$  on  $\text{Ti}_3\text{C}_2\text{T}_x$  provides a resistance-less path favorable for fast electron transfer; (5) the geometric confinement of  $\text{Mo}_2\text{C}$  and the MXene within the carbon layer guarantees the excellent stability of the HER performance during long-term operation.

## Conclusions

In summary, we developed an *in situ* polymerization strategy to stabilize MXenes against oxidation during high-temperature treatment, which was extended to the fabrication of a hierarchical  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$  composite. Serving as a HER catalyst,  $\text{Mo}_2\text{C}/\text{Ti}_3\text{C}_2\text{T}_x@NC$  exhibits excellent activity in the entire pH range and robust stability. Theoretical simulations demonstrated that the coupled  $\text{Mo}_2\text{C}$  and pyridinic N doped carbon serve as the most active sites for H adsorption, while the strongly incorporated metallic  $\text{Ti}_3\text{C}_2\text{T}_x$  MXene provides a fast charge transfer pathway, which synergistically lead to much superior HER performance. This strategy not only provides a highly active HER catalyst, but also opens a new avenue for designing MXene-based nanocomposites that require high synthesis temperature, which could be applied in various energy-related fields.

## Conflicts of interest

There are no conflicts to declare.

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## References

- S. Chu, Y. Cui and N. Liu, *Nat. Mater.*, 2017, **16**, 16.
- J. A. Turner, *Science*, 2004, **305**, 972.
- Z. W. Seh, J. Kibsgaard, C. F. Dickens, I. Chorkendorff, J. K. Nørskov and T. F. Jaramillo, *Science*, 2017, **355**, eaad4998.
- C. Niether, S. Faure, A. Bordet, J. Deseure, M. Chatenet, J. Carrey, B. Chaudret and A. Rouet, *Nat. Energy*, 2018, **3**, 476.
- M. S. Faber and S. Jin, *Energy Environ. Sci.*, 2014, **7**, 3519.
- Y. Zang, S. Niu, Y. Wu, X. Zheng, J. Cai, J. Ye, Y. Xie, Y. Liu, J. Zhou and J. Zhu, *Nat. Commun.*, 2019, **10**, 1217.
- H. Jin, X. Liu, S. Chen, A. Vasileff, L. Li, Y. Jiao, L. Song, Y. Zheng and S.-Z. Qiao, *ACS Energy Lett.*, 2019, **4**, 805.
- H. Wang, L. Ouyang, G. Zou, C. Sun, J. Hu, X. Xiao and L. Gao, *ACS Catal.*, 2018, **8**, 9529.
- X. F. Lu, L. Yu, J. Zhang and X. W. Lou, *Adv. Mater.*, 2019, **31**, 1900699.
- M. D. Hossain, Z. Liu, M. Zhuang, X. Yan, G. L. Xu, C. A. Gadre, A. Tyagi, I. H. Abidi, C. J. Sun and H. Wong, *Adv. Energy Mater.*, 2019, **9**, 1803689.
- Z. Fang, L. Peng, Y. Qian, X. Zhang, Y. Xie, J. J. Cha and G. Yu, *J. Am. Chem. Soc.*, 2018, **140**, 5241.
- V. Jose, E. Edison, W. W. Manalastas Jr, S. Sreejith, J. M. Vianney Nsanzimana, M. Srinivasan and J.-M. Lee, *ACS Appl. Mater. Interfaces*, 2019, **11**, 39798.
- H. Wang, Q. Yi, L. Gao, Y. Gao, T. Liu, Y.-B. Jiang, Y. Sun and G. Zou, *Nanoscale*, 2017, **9**, 16342.
- H. Wang, X. Xiao, S. Liu, C.-L. Chiang, X. Kuai, C.-K. Peng, Y.-C. Lin, J. Zhao, J.-H. Choi, Y.-G. Lin, J.-M. Lee and L. Gao, *J. Am. Chem. Soc.*, 2019, **141**, 18578.
- J. Yang, A. R. Mohamad, Y. Wang, R. Fullon, X. Song, F. Zhao, I. Bozkurt, M. Augustin, E. J. Santos and H. S. Shin, *Nat. Mater.*, 2019, **18**, 1309.
- J. R. Kitchin, J. K. Nørskov, M. A. Barteau and J. G. Chen, *Catal. Today*, 2005, **105**, 66.
- H. Wang, C. Sun, Y. Cao, J. Zhu, Y. Chen, J. Guo, J. Zhao, Y. Sun and G. Zou, *Carbon*, 2017, **114**, 628.
- Z. Zhou, Z. Yuan, S. Li, H. Li, J. Chen, Y. Wang, Q. Huang, C. Wang, H. E. Karahan and G. Henkelman, *Small*, 2019, **15**, 1900358.
- B. Anasori, M. R. Lukatskaya and Y. Gogotsi, *Nat. Rev. Mater.*, 2017, **2**, 16098.
- K. Maleski, C. E. Ren, M.-Q. Zhao, B. Anasori and Y. Gogotsi, *ACS Appl. Mater. Interfaces*, 2018, **10**, 24491.
- M. Ghidui, M. R. Lukatskaya, M.-Q. Zhao, Y. Gogotsi and M. W. Barsoum, *Nature*, 2014, **516**, 78.
- S. J. Kim, H.-J. Koh, C. E. Ren, O. Kwon, K. Maleski, S.-Y. Cho, B. Anasori, C.-K. Kim, Y.-K. Choi and J. Kim, *ACS Nano*, 2018, **12**, 986.
- W. Bao, X. Tang, X. Guo, S. Choi, C. Wang, Y. Gogotsi and G. Wang, *Joule*, 2018, **2**, 778.
- F. Shahzad, M. Alhabeab, C. B. Hatter, B. Anasori, S. M. Hong, C. M. Koo and Y. Gogotsi, *Science*, 2016, **353**, 1137.
- J. Zhang, Y. Zhao, X. Guo, C. Chen, C.-L. Dong, R.-S. Liu, C.-P. Han, Y. Li, Y. Gogotsi and G. Wang, *Nat. Catal.*, 2018, **1**, 985.
- J. Liu, Y. Liu, D. Xu, Y. Zhu, W. Peng, Y. Li, F. Zhang and X. Fan, *Appl. Catal., B*, 2019, **241**, 89.
- X.-D. Zhu, Y. Xie and Y.-T. Liu, *J. Mater. Chem. A*, 2018, **6**, 21255.
- D. Zhao, Z. Chen, W. Yang, S. Liu, X. Zhang, Y. Yu, W.-C. Cheong, L. Zheng, F. Ren, G. Ying, X. Cao, D. Wang, Q. Peng, G. Wang and C. Chen, *J. Am. Chem. Soc.*, 2019, **141**, 4086.
- X. Wu, S. Zhou, Z. Wang, J. Liu, W. Pei, P. Yang, J. Zhao and J. Qiu, *Adv. Energy Mater.*, 2019, **9**, 1901333.

- 30 X. Wu, Z. Wang, M. Yu, L. Xiu and J. Qiu, *Adv. Mater.*, 2017, **29**, 1607017.
- 31 J.-S. Li, Y. Wang, C.-H. Liu, S.-L. Li, Y.-G. Wang, L.-Z. Dong, Z.-H. Dai, Y.-F. Li and Y.-Q. Lan, *Nat. Commun.*, 2016, **7**, 11204.
- 32 H. Huang, J. Cui, G. Liu, R. Bi and L. Zhang, *ACS Nano*, 2019, **13**, 3448.
- 33 Y. Huang, Q. Gong, X. Song, K. Feng, K. Nie, F. Zhao, Y. Wang, M. Zeng, J. Zhong and Y. Li, *ACS Nano*, 2016, **10**, 11337.
- 34 M. Han, X. Yin, X. Li, B. Anasori, L. Zhang, L. Cheng and Y. Gogotsi, *ACS Appl. Mater. Interfaces*, 2017, **9**, 20038.
- 35 M. Yu, S. Zhou, Z. Wang, J. Zhao and J. Qiu, *Nano Energy*, 2018, **44**, 181.
- 36 X. Wang, Z. Wang, M. Zhang, X. Jiang, Y. Wang, J. Lv, G. He and Z. Sun, *J. Alloys Compd.*, 2017, **725**, 1166.
- 37 Y. Yang, M. Luo, Y. Xing, S. Wang, W. Zhang, F. Lv, Y. Li, Y. Zhang, W. Wang and S. Guo, *Adv. Mater.*, 2018, **30**, 1706085.
- 38 J. Zhang, L. Zhou, Q. Sun, H. Ming, L. Sun, C. Wang, Y. Wu, K. Guan, L. Wang and J. Ming, *Chem.–Eur. J.*, 2019, **25**, 8813.
- 39 Y. Li, H. Wang, L. Xie, Y. Liang, G. Hong and H. Dai, *J. Am. Chem. Soc.*, 2011, **133**, 7296.
- 40 T. F. Jaramillo, K. P. Jørgensen, J. Bonde, J. H. Nielsen, S. Horch and I. Chorkendorff, *Science*, 2007, **317**, 100.
- 41 J. K. Nørskov, T. Bligaard, A. Logadottir, J. Kitchin, J. G. Chen, S. Pandelov and U. Stimming, *J. Electrochem. Soc.*, 2005, **152**, 23.
- 42 Z. Lai, A. Chaturvedi, Y. Wang, T. H. Tran, X. Liu, C. Tan, Z. Luo, B. Chen, Y. Huang and G.-H. Nam, *J. Am. Chem. Soc.*, 2018, **140**, 8563–8568.
- 43 X. Zhang, X. Yu, L. Zhang, F. Zhou, Y. Liang and R. Wang, *Adv. Funct. Mater.*, 2018, **28**, 1706523.
- 44 I. K. Mishra, H. Zhou, J. Sun, F. Qin, K. Dahal, J. Bao, S. Chen and Z. Ren, *Energy Environ. Sci.*, 2018, **11**, 2246.
- 45 K. Liang, S. Pakhira, Z. Yang, A. Nijamudheen, L. Ju, M. Wang, C. I. Aguirre-Velez, G. E. Sterbinsky, Y. Du and Z. Feng, *ACS Catal.*, 2018, **9**, 651.
- 46 J. Liang, C. Ding, J. Liu, T. Chen, W. Peng, Y. Li, F. Zhang and X. Fan, *Nanoscale*, 2019, **11**, 10992.
- 47 H. Ang, H. Wang, B. Li, Y. Zong, X. Wang and Q. Yan, *Small*, 2016, **12**, 2859.
- 48 H.-W. Liang, S. Brüller, R. Dong, J. Zhang, X. Feng and K. Müllen, *Nat. Commun.*, 2015, **6**, 7992.
- 49 X. Wang, A. Vasileff, Y. Jiao, Y. Zheng and S. Z. Qiao, *Adv. Mater.*, 2019, **31**, 1803625.
- 50 W. Chen, J. Pei, C.-T. He, J. Wan, H. Ren, Y. Zhu, Y. Wang, J. Dong, S. Tian, W.-C. Cheong, S. Lu, L. Zheng, X. Zheng, W. Yan, Z. Zhuang, C. Chen, Q. Peng, D. Wang and Y. Li, *Angew. Chem., Int. Ed.*, 2017, **56**, 16086.
- 51 M. Yu, S. Zhou, Z. Wang, J. Zhao and J. Qiu, *Nano Energy*, 2018, **44**, 181.
- 52 Y. Liu, G. Yu, G. D. Li, Y. Sun, T. Asefa, W. Chen and X. Zou, *Angew. Chem., Int. Ed.*, 2015, **54**, 10752.