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# Conformationally-flexible and moderately electron-donating units-installed D–A–D triad enabling multicolor-changing mechanochromic luminescence, TADF and room-temperature phosphorescence†

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**A novel twisted donor–acceptor–donor (D–A–D)  $\pi$ -conjugated compound that contains flexible and moderately-electron-donating units has been designed and synthesized. It exhibited not only multi-color-changing mechanochromic luminescence and thermally activated delayed fluorescence, but also, unexpectedly, room-temperature phosphorescence in a host layer.**

Smart emissive organic materials that exhibit reversible changes in luminescence properties in the solid state in response to external stimuli, such as mechanical forces, temperature, and vapors (*i.e.*, mechanochromic luminescence: MCL), can find diverse applications in security inks, sensors, and optoelectronic devices.<sup>1</sup> Two-color-changing MCL systems have been intensively studied over the last decade.<sup>1</sup> In sharp contrast, the development of multi-color-changing MCL organic materials has been sluggish, due to the scarcity of design guidelines to meet demanding criteria to form multiple metastable states that must be also interconvertible and emissive in the solid state.<sup>2</sup> To address the issue and develop value-added MCL materials, taking into consideration that D–A–D  $\pi$ -conjugated triad scaffolds are widely used in materials science,<sup>3</sup> we have recently demonstrated that a twisted donor–acceptor–donor (D–A–D) triad (PTZ–DBPHZ, Fig. 1a) exhibits tricolor-changing

MCL properties<sup>4</sup> along with efficient thermally activated delayed fluorescence (TADF).<sup>5</sup> Most importantly, owing to the conformational flexibility of phenothiazine (PTZ) donors (quasi-equatorial and quasi-axial, Fig. 1a), the D–A–D molecule can adopt multiple metastable conformers, and they are switchable by external stimuli such as grinding, heating, and exposure to vapor. The interconversions among these conformers boost different emissive channels, depending on the charge transfer (CT) nature of the excited states associated with each conformer. To enhance the “conformation-dictated emission regulation” concept, herein we designed a novel twisted D–A–D triad (DPPZS–DBPHZ) **1** that comprises dibenzo[*a,j*]phenazine (DBPHZ)<sup>6</sup> as the A and dihydrophenolphosphazine<sup>7</sup> sulfide (DPPZS) as the Ds (Fig. 1a). Notably, the molecule is endowed not only with multi-color-changing MCL but also remarkable acid/base-responsive emission switching, single crystal transition-induced emission change, and room-temperature phosphorescence in a host matrix. The similarity in atomic radius between P (0.98 Å) and S (0.88 Å) would allow the donor (DPPZS) to mimic the flexible conformational change of PTZ–DBPHZ,<sup>4</sup> although this replacement would not completely mimic the behavior of PTZ–DBPHZ due to the

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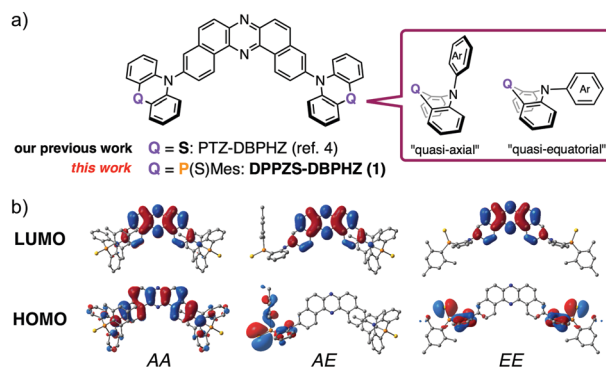


Fig. 1 (a) structures of D–A–D compounds cored with DBPHZ; (b) the frontier orbitals of the conformers of **1** (AA: quasi-axial and quasi-axial; AE: quasi-axial and quasi-equatorial; EE: quasi-equatorial and quasi-equatorial).



extra S=P and Mes-P moieties. Fig. 1b illustrates the optimized structures and the frontier orbitals of possible conformers of **1** calculated by the DFT method at the B3LYP/6-31+G(d,p) level.<sup>8</sup>

As expected, depending on the conformation, the HOMO/LUMO distributions show significant variations, and this is also the case with the HOMO–LUMO gaps and excitation modes (Fig. S1, ESI†).<sup>8</sup> Furthermore, the distinct separation of the HOMOs and LUMOs of the AE and EE conformers implies a small single-triplet splitting ( $\Delta E_{ST}$ ) that would be promising for TADF emission.<sup>9</sup> Scheme S1 (ESI†) illustrates the synthetic route to **1**.<sup>8</sup> The double Br–Li permutation of *N*-Boc bis(*o*-bromophenyl)amine **2** generated a dilithiated species, which was subsequently trapped with P-Mes dihalophosphine<sup>10</sup> to give *N*-protected dihydrophenophosphanizine **3** in 67% yield. Since P(III) compound **3** was gradually oxidized under air into the corresponding phosphine oxide, it was subsequently converted into **4** by sulfurization with S<sub>8</sub> to protect the P center from oxidation. The detachment of the *N*-Boc group with trifluoroacetic acid (TFA) afforded **5** in 60% yield. The following Pd-catalyzed Hartwig-Buchwald double amination of dibromo DBPHZ **6**<sup>6</sup> with **5** successfully provided the D–A–D compound **1** in good yield.<sup>11</sup> The X-ray crystallographic analyses of the single crystals of **4** and **5** grown from *n*-hexane/AcOEt unambiguously determined their molecular conformations (Fig. S2 and S3, ESI†).<sup>12</sup> The central 6-membered heterocycle of **4** takes a boat-chair conformation, with both the *N*-Boc and P-Mes groups occupying the axial positions of the boat-chair (Fig. S2a, ESI†), which seems to be common for phenophosphanizine compounds.<sup>13</sup> The presence of the P=S functionality allows a dimeric structure through complementary C–H...S (2.97 Å) contacts (Fig. S2b, ESI†). The central heterocycle of **5** adopts a less-bent structure, and the P-Mes group occupies the equatorial site on the boat-chair (Fig. S3a, ESI†), which might due to the formation of a dimeric structure through complementary N–H...S (2.59 Å) contacts (Fig. S3b, ESI†).

The comparison of UV-vis absorption of diluted dichloromethane (DCM) solutions of D, A, a mixture of D and A (2:1 ratio), and D–A–D **1** clearly showed that an absorption observed at around 460 nm for **1** is assignable to an intramolecular charge-transfer (ICT) transition (Fig. 2a). The CT absorption showed a slight bathochromic shift as the solvent polarity increased (Fig. 2b), suggesting the mixing of CT and  $\pi$ – $\pi^*$  natures. The photoluminescence (PL) spectra of **1** exhibited a significant positive solvatochromism as a function of solvent polarity, with moderate photoluminescence quantum yields (PLQYs) ranging from 0.26 to 0.34 (Table S5, ESI†).<sup>8</sup> Since there would be significant differences in conformations in solution (all conformers are in equilibrium) and in solid state (specific conformers are dominant, *vide infra*), it would be difficult to correlate the photophysics of the solutions and solid state at this time.

As expected, D–A–D compound **1** showed multi-color-changing MCL properties in the solid state (Fig. 3). Crystals obtained from *n*-hex/CHCl<sub>3</sub> showed blue–green emission ( $\lambda_{em}$  497 nm) upon irradiation with UV light (**1-BG**, Fig. 3 and Fig. S7, ESI†).<sup>8</sup> Grinding of **1-BG** (*G*, Fig. 3) with a pestle caused a morphological transition from the crystalline state to

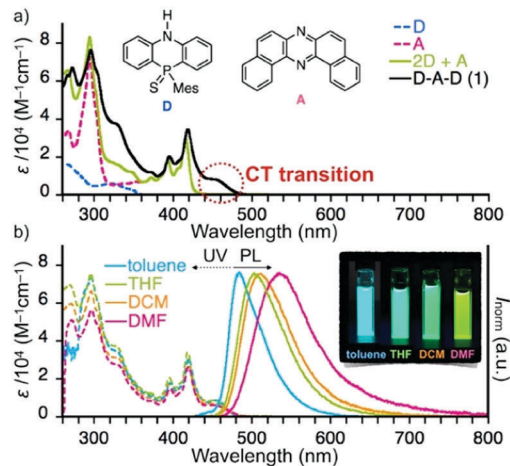


Fig. 2 (a) UV spectra of DCM solutions ( $c \times 10^{-5}$  M) of D, A, 2D + A, and D–A–D (**1**); (b) UV-vis and steady-state PL spectra of the solutions ( $c \times 10^{-5}$  M) prepared with different solvents. The inset photos were taken under irradiation with UV light ( $\lambda$  365 nm).



Fig. 3 Illustrative summary of the MCL property of **1**.

an amorphous meta-stable state emitting yellow light (**1-Y**, Fig. 3), which was clearly suggested by the PXRD and DSC measurements (Fig. S8 and S9, ESI†).<sup>8</sup> This morphological transition caused a significant red shift in the emission ( $1394 \text{ cm}^{-1}$ ) with an increase in PLQY (Fig. 3). By fuming (*F1*, Fig. 3) or thermal annealing (*H*, Fig. 3), the meta-stable solid (**1-Y**) underwent a phase transition to the more stable crystalline state **1-G1** (Fig. 3 and Fig. S8, S9, ESI†), showing slightly blue-shifted emission ( $\lambda_{em}$  520 nm). Notably, **1-BG** exhibited a drastic response toward acid (*F2*, Fig. 3). Upon treatment with TFA vapor, pale yellow **1-BG** immediately turned to pale purple **1-DR**, exhibiting near-infrared (NIR) emission in the solid state ( $\lambda_{em}$  740 nm, Fig. S7, ESI†), while treatment of **1-DR** with triethylamine (TEA) vapor (*F3*, Fig. 3) allowed a blue-shift of the emission to bright green (**1-G2**,  $\lambda_{em}$  518 nm), which is similar to that of **1-G1** (Fig. 3). Notably, these acid–base responsive emission color changes were reversible (*F2* and *F3*, Fig. 3), and were accompanied by a very large  $\Delta\lambda_{em}$ , as large as 222 nm ( $5791 \text{ cm}^{-1}$ ). It should be noted that there are few reports on emissive solids that exhibit remarkable stimuli-responsive emission switch between the visible and NIR regions.<sup>14</sup> All the solids reverted back to **1-BG** through recrystallization (*R*, Fig. 3).



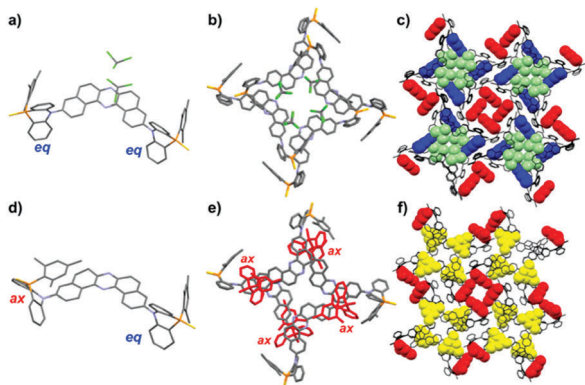


Fig. 4 Molecular structures of (a) **1-CHCl<sub>3</sub>** and (d) **1-apo**; packing structures of (b and c) **1-CHCl<sub>3</sub>** and (e and f) **1-apo**; space-filling models shown in red, blue, and green in (c) indicate eq-P-Mes, eq-P-Mes, and CHCl<sub>3</sub>, respectively; space-filling models shown in red and yellow in (f) indicate eq-P-Mes and ax-P-Mes, respectively.

A single crystal (SC) of **1** was obtained by rapidly diffusing a CHCl<sub>3</sub> solution of **1** into *n*-hexane. The X-ray crystallographic analysis revealed that **1** adopts a double quasi-equatorial conformation (EE), holding two CHCl<sub>3</sub> molecules per unit (**1-CHCl<sub>3</sub>**, Fig. 4a).<sup>12</sup> The molecules form ordered micro pores, where CHCl<sub>3</sub> are included inside the channel (Fig. 4b and c). Notably, desorption of CHCl<sub>3</sub> from **1-CHCl<sub>3</sub>** under vacuum caused an SC-to-SC transition, allowing for the conformational change in a DPPZS unit from equatorial to axial (**1-apo**, Fig. 4d). Importantly, the SC-to-SC transition did not cause a significant disturbance of the supramolecular channel structures (Fig. 4e and f). Furthermore, the spaces filled with CHCl<sub>3</sub> in **1-CHCl<sub>3</sub>** were replaced by the P-Mes group on the axial donor in **1-apo** crystals (yellow in Fig. 4f). Most importantly, the transition caused a significant change in emission profiles (Fig. S10, ESI<sup>†</sup>),<sup>8</sup> and the PL spectra of **1-CHCl<sub>3</sub>** and **1-apo** are very similar to those of **1-BG** and **1-Y**, respectively. Since no significant change, other than the molecular conformation of a donor unit (from EE to AE), was observed through the transition, this would be clear evidence for how molecular conformations in the D-A-D scaffold significantly influence its emission properties. In conjunction with these experimental data and the TD-DFT calculations (Fig. S1, ESI<sup>†</sup>),<sup>8</sup> the relationship between the conformations and MCL properties of **1** would be explained as follows: the blue-green emission of **1-BG** comes from the EE conformer, while the yellow emission of **1-Y** derives from the AE conformer, as clarified with the above-mentioned results; the green emissions of **1-G1** and **1-G2** would be ascribed to the AA conformer or a mixture of AE and AA conformers (*vide infra*). Since the D-A geometry of the EE conformer is perpendicular, the HOMO-LUMO transition in the EE conformer is basically forbidden ( $\lambda$  516 nm,  $f$  = 0.000, Fig. S1, ESI<sup>†</sup>).<sup>8</sup> Since the electron-donating ability of a DPPZS unit is much weaker than that of PTZ, photo-induced electron transfer to generate CT states is less likely to occur. Instead, the HOMO-6 to LUMO transition is feasible, due to its  $\pi$ - $\pi^*$  nature ( $\lambda$  390 nm,  $f$  = 0.474). On the other hand, in the AE form, a smaller excitation energy is required, which is ascribed to the

allowed HOMO-1 to LUMO transition ( $\lambda$  450 nm,  $f$  = 0.364, CT and  $\pi$ - $\pi^*$  hybrid transition), qualitatively rationalizing why the EE conformer showed a higher-energy (bluer) emission than that of the AE conformer. Furthermore, the TD-DFT calculation indicated that the HOMO-LUMO transition is allowed in the AA conformer ( $\lambda$  444 nm,  $f$  = 0.652,  $\pi$ - $\pi^*$  transition), and the excitation energies of AA and AE are very close (Fig. S1, ESI<sup>†</sup>), implying that the emission spectra of AE and AA would overlap. It should be noted that the relationship between conformers (AA, AE, and EE) and the order of  $\lambda_{em}$  for D-A-D compound **1** is different from that for PTZ-DBPHZ,<sup>4</sup> where the CT nature of each conformer exclusively governs the  $\lambda_{em}$  order (*i.e.*,  $\lambda_{em}$ : EE > AE > AA). The discrepancy between PTZ- and DPPZS-installed D-A-D compounds could be ascribed to the difference in electron-donating abilities of the Ds, which determine whether or not the populated locally excited singlet state of the A unit (<sup>1</sup>LE<sub>A</sub>) can be converted to more energetically-relaxed charge-transfer states (<sup>1</sup>CT).<sup>9</sup> The influences of the P=S and P-Mes moieties, which might form aggregates in solution and solid state, on their emission properties are not totally excluded.

To obtain deeper insights into the photophysical properties of **1**, time-resolved luminescence spectroscopic measurements were performed in an inert and non-polar Zeonex<sup>®</sup> host (Fig. 5). Compound **1** showed emissions mainly in two distinct time regions (Fig. 5a). The first component that decays within the order of nanoseconds is attributed to prompt fluorescence (PF) from the singlet excited state (S<sub>1</sub>). The PL spectra of Zeonex<sup>®</sup>-based film with a time delay of 3 ns showed two distinct species: a structured peak in the range 400–450 nm that decays within the first few nanoseconds, and a broad Gaussian peak in the range 450–650 nm that decays over a longer time. These two species can be attributed to the emissions from the <sup>1</sup>LE<sub>A</sub> and <sup>1</sup>CT singlet states, respectively. With longer delay times (micro-to-milliseconds), delayed emissions were observed (Fig. 5a). Depending on the experimental temperature, both delayed emissions from S<sub>1</sub> and the emissions from T<sub>1</sub> were observed on similar millisecond time-scales. To differentiate these, spectral inspections were conducted at different temperatures. In all the hosts, an emission from T<sub>1</sub> (phosphorescence: PH) was observed at a low temperature (80 K) (blue lines in Fig. 5b), with the T<sub>1</sub> energy being 2.41 eV in Zeonex<sup>®</sup>. At an ambient temperature (320 K), a delayed emission observed between 440 and 530 nm in the Zeonex<sup>®</sup> was identified as TADF,



Fig. 5 (a) Temperature-variable plots of emission intensity of **1** in Zeonex<sup>®</sup> (1% w/w) against delay time at different temperatures; (b) normalized PL spectra of **1** in Zeonex<sup>®</sup> at 320 K and 80 K.



which was confirmed by the linear dependence of the emission intensity in the millisecond delayed region on laser fluence (Fig. S11, ESI†). Notably, in Zeonex<sup>®</sup>, a significant PH was observed even at room temperature (*i.e.*, room temperature phosphorescence: RTP) without quenching.<sup>15</sup> This makes a nice contrast with our previously reported D–A–D compounds,<sup>4,9</sup> which do not exhibit RTP. The temperature-dependent emission behaviour of **1** in a Zeonex<sup>®</sup> host indicates the presence of an interplay between TADF and RTP processes. From the viewpoint of molecular design, the lower electron-donating ability of DPPZS than those of POZ and PTZ destabilizes the <sup>1</sup>CT energy. As the result, the  $\Delta E_{ST}$  (0.30 eV) of **1** is moderately larger than those of our previously reported D–A–D triads.<sup>4,9</sup> In addition to the  $\Delta E_{ST}$  value, the host environments would constrain the vibrational and rotational movements of the D–A–D compound, which would inhibit non-radiative channels from T<sub>1</sub> and allow the boosting of both radiative (TADF and RTP) pathways.<sup>16</sup>

In conclusion, we have developed a novel D–A–D  $\pi$ -conjugated compound that shows unique emission properties beyond multi-color-changing MCL behavior induced by conformational interconversion. The D–A–D compound exhibited a remarkable reversible emission change between the visible and NIR region in response to acid or base. Notably, an analysis of the SC-to-SC transition phenomena provided us with clear evidence for conformation-dependent emission changes in the D–A–D triad. Furthermore, the phosphorus compound unexpectedly exhibited temperature-dependent TADF and RTP emissions in a host matrix. We believe that the results presented here would further guide us to more diverse designs of value-added MCL-materials. Moreover, the photophysical studies shown here would provide us with opportunities for designing metal-free purely organic RTP materials that can be utilized in optoelectronics, such as phosphorescence organic light-emitting diodes (PHOLEDs),<sup>17</sup> in the future.

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## Conflicts of interest

There are no conflicts to declare.

## Notes and references

- (a) Y. Sagara and T. Kato, *Nat. Chem.*, 2009, **1**, 605–610; (b) A. Pucci, R. Bizzarri and R. Giacomo, *Soft Matter*, 2011, **7**, 3689–3700; (c) Z. Chi, X. Zhang, B. Xu, X. Zhou, C. Ma, Y. Zhang, S. Liu and J. Xu, *Chem. Soc. Rev.*, 2012, **41**, 3878–3896; (d) Y. Sagara, S. Yamane, M. Mitani, C. Weder and T. Kato, *Adv. Mater.*, 2016, **28**, 1073–1095; (e) C. Wang and Z. Li, *Mater. Chem. Front.*, 2017, **1**, 2174–2194.
- (a) Y. Sagara and T. Kato, *Angew. Chem., Int. Ed.*, 2011, **50**, 9128–9132; (b) Z. Ma, Z. Wang, M. Teng, Z. Xu and X. Jia, *Chem-PhysChem*, 2015, **16**, 1811–1828 and references therein.
- (a) B. J. Eckstein, F. S. Melkonyan, E. F. Manley, S. Fabiano, A. R. Mouat, L. X. Chen, A. Facchetti and T. J. Marks, *J. Am. Chem. Soc.*, 2017, **139**, 14356–14359; (b) T. J. Aldrich, S. M. Swick, F. S. Melkonyan and T. J. Marks, *Chem. Mater.*, 2017, **29**, 10294–10298.
- M. Okazaki, Y. Takeda, P. Data, P. Pander, H. Higginbotham, A. P. Monkman and S. Minakata, *Chem. Sci.*, 2017, **8**, 2677–2686.
- (a) H. Uoyama, K. Goushi, K. Shizu, H. Nomura and C. Adachi, *Nature*, 2012, **492**, 234–238; (b) Y. Tao, K. Yuan, T. Chen, P. Xu, H. Li, R. Chen, C. Zheng, L. Zhang and W. Huang, *Adv. Mater.*, 2014, **26**, 7931–7958; (c) Z. Yang, Z. Mao, Z. Xie, Y. Zhang, S. Liu, J. Zhao, J. Xu, Z. Chi and M. P. Aldred, *Chem. Soc. Rev.*, 2017, **46**, 915–1016; (d) F. B. Dias, T. J. Penfold and A. P. Monkman, *Methods Appl. Fluoresc.*, 2017, **5**, 012001.
- Y. Takeda, M. Okazaki and S. Minakata, *Chem. Commun.*, 2014, **50**, 10291–10294.
- (a) A. I. Bokanov and B. I. Stepanov, *Russ. Chem. Rev.*, 1977, **46**, 855–860; (b) L. D. Freedman and H. S. Freeman, *Chem. Rev.*, 1987, **87**, 289–306.
- For the details, see the ESI†.
- (a) P. Data, P. Pander, M. Okazaki, Y. Takeda, S. Minakata and A. P. Monkman, *Angew. Chem., Int. Ed.*, 2016, **55**, 5739–5744; (b) D. de Sa Pereira, P. L. dos Santos, J. S. Ward, P. Data, M. Okazaki, Y. Takeda, S. Minakata, M. R. Bryce and A. P. Monkman, *Sci. Rep.*, 2017, **7**, 6234.
- Y. Takeda, T. Nishida and S. Minakata, *Chem. – Eur. J.*, 2014, **20**, 10266–10270.
- Although reduction of the P(v) compound **1** to the corresponding P(III) compound as reference compound using a variety of reductants such as DIBAL, LiAlH<sub>4</sub>, and P(NMe<sub>2</sub>)<sub>3</sub> was attempted, any reduced P(III) compound was not obtained.
- For a summary of the crystal data of **4**, **5**, **1-CHCl<sub>3</sub>**, and **1-apo**, see the ESI†.
- (a) S. N. Gurkova, A. I. Gusev, V. A. Sharapov, L. A. Yagodina, A. I. Bokanov and B. I. Stepanov, *Zh. Strukt. Khim.*, 1977, **18**, 62–66; (b) A. I. Gusev, S. N. Gurkova, V. K. Bel’skii, V. E. Zavadnik and L. A. Yagodina, *J. Struct. Chem.*, 1980, **20**, 537–541.
- (a) Q. Xiao, J. Zheng, M. Li, S.-Z. Zhan, J.-H. Wang and D. Li, *Inorg. Chem.*, 2014, **53**, 11604–11615; (b) M. Tanioka, S. Kamino, A. Muranaka, Y. Ooyama, H. Ota, Y. Shirasaki, J. Horigome, M. Ueda, M. Uchiyama, D. Sawada and S. Enomoto, *J. Am. Chem. Soc.*, 2015, **137**, 6436–6439; (c) T. Seki, N. Tokodai, S. Omagari, T. Nakanishi, Y. Hasegawa, T. Iwasa, T. Taketsugu and H. Ito, *J. Am. Chem. Soc.*, 2017, **139**, 6514–6517.
- S. Mukherjee and P. Thilagar, *Chem. Commun.*, 2015, **51**, 10988–11003.
- J. S. Ward, R. S. Nobuyasu, A. S. Batsanov, P. Data, A. P. Monkman, F. B. Dias and M. R. Bryce, *Chem. Commun.*, 2016, **52**, 2612–2615.
- D. Chaudhuri, E. Sigmund, A. Meyer, L. Rock, P. Klemm, S. Lautenschlager, A. Schmid, S. R. Yost, T. Vanvoorhis, S. Bange, S. Hoger and J. M. Lupton, *Angew. Chem., Int. Ed.*, 2013, **52**, 13449–13452.

