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Alloying strategies for tuning product selectivity during electrochemical CO₂ reduction over Cu

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Excessive reliance on fossil fuels has led to the release and accumulation of large quantities of CO₂ into the atmosphere which has raised serious concerns related to environmental pollution and global warming. One way to mitigate this problem is to electrochemically recycle CO₂ to value-added chemicals or fuels using electricity from renewable energy sources. Cu is the only metallic electrocatalyst that has been shown to produce a wide range of industrially important chemicals at appreciable rates. However, low product selectivity is a fundamental issue limiting commercial applications of electrochemical CO₂ reduction over Cu catalysts. Combining copper with other metals that actively contribute to the electrochemical CO₂ reduction reaction process can selectively facilitate generation of desirable products. Alloying Cu can alter surface binding strength through electronic and geometric effects, enhancing the availability of surface confined carbon species, and stabilising key reduction intermediates. As a result, significant research has been undertaken to design and fabricate copper-based alloy catalysts with structures that can enhance the selectivity of targeted products. In this article, progress with use of alloying strategies for development of Cu-alloy catalysts are reviewed. Challenges in achieving high selectivity and possible future directions for development of new copper-based alloy catalysts are considered.

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1. Introduction

Carbon dioxide (CO₂) is a greenhouse gas that is a major contributor to global warming. In recent years, the concentration of CO₂ in the atmosphere has increased rapidly, resulting in the rise of global temperatures, which disturbs the ecological balance. In nature, the CO₂ released into the atmosphere by natural processes is utilised by terrestrial plants and aquatic species, thus achieving a carbon balance. However, excessive dependency on fossil fuels in the modern chemical industry, particularly the automobile and electricity generation sectors, has led to the accumulation of additional CO₂ in the atmosphere.¹ In a recent report, the National Oceanic and Atmospheric Administration (NOAA) global monitoring laboratory advised of a 2.3 ppm surge in CO₂ levels in 2020 to 2021 period leading to an average level of 419 ppm in 2021.² Fig. 1 shows the global monthly mean CO₂ emissions in the atmosphere from 1980 until May 2022 reported by NOAA which reveal the seriousness of the carbon accumulation problem. However, if an effective CO₂ capture, storage and conversion to chemical fuels technologies can be developed, the CO₂ can be recycled. Ideally, if CO₂ is sourced directly through air or bio-

energy carbon capture and storage and reacted with renewable H₂, the carbon cycle can be closed.³

A range of methods such as photochemical,^{4–7} biochemical,^{8–10} and photo-/thermochemical,^{11–15} have been explored for conversion of CO₂ to value added chemicals or fuels. However, their success is highly sensitive to the reaction environment.¹⁶ Another possibility is use of the electrochemical CO₂ reduction reaction (eCO₂RR) which can be per-

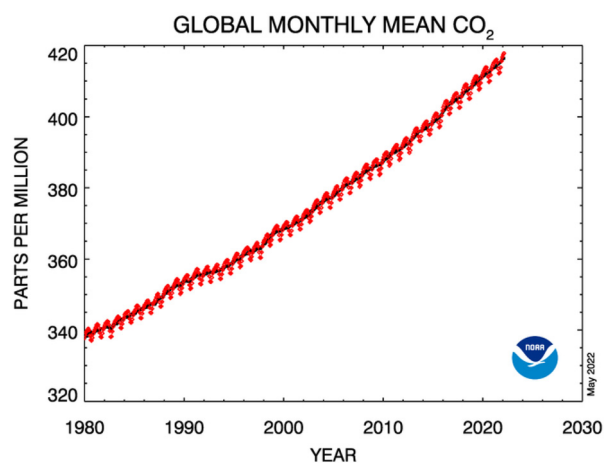


Fig. 1 Monthly mean CO₂ emissions from 1980 to May 2022 as reported by the NOAA global monitoring laboratory.²

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formed under ambient conditions using non-toxic materials and electrical energy from renewable sources, making it a highly desirable approach.¹⁷

The eCO₂RR can be performed in aqueous and non-aqueous media where the former are attractive in terms of non-toxicity, low cost, good conductivity, high proton availability and eco-friendliness. Aqueous electrolytes used are generally alkali salts comprising Na⁺, K⁺, Li⁺, Rb⁺, or Cs⁺ cations and OH⁻, HCO₃⁻, Cl⁻, H₂PO₄⁻ or SO₄²⁻ anions. Both cations and anions can have a strong impact on the efficiency of eCO₂RR as well as product selectivity.^{18–21} Additionally, other factors, such as temperature and pH (more critically local pH, since it can differ substantially from the bulk value, especially under high current density/low buffer capacity conditions) also influence the product selectivity during the eCO₂RR.^{22–24} Table 1 provides several examples of eCO₂RR products generated in aqueous media and their reversible potentials (E°) vs. the reversible hydrogen electrode (RHE). Significant drawbacks of eCO₂RR in aqueous solutions are the low solubility of CO₂ resulting in low current densities, the presence of a competitive hydrogen evolution reaction (HER) which suppresses the overall eCO₂RR efficiency, in addition to sluggish kinetics and hence a large overpotential required to produce commercially desirable products at a sufficient rate.^{17,21,25}

To date, three types of electrolyzers have been mainly used for the eCO₂RR. The H-cell is most commonly used since it is easy to construct and facilitates rapid screening of catalysts and electrolyte composition. Fig. 2(a) shows a conventional H-cell for eCO₂RR. In this, the aqueous electrolytes used in most cases in both anodic and cathodic compartments are pre-saturated with CO₂. The cell is then sealed prior to performing electrolysis allowing gaseous products to be collected and quantified. The solubility of molecular CO₂ in these electrolytes is around 30 mM at 1 bar and thus the rate of CO₂ reduction is limited with a current density of typically less than 100 mA cm⁻².²⁶ To overcome the drawback of the low mass transport rate associated with the H-cell, electrolyzers equipped with a gas diffusion electrode (GDE) have been developed. In these electrolyzers, the CO₂ gas is diffused through a thin electrolyte layer on the GDE, with the diffusion length being ~50 nm.²⁷ The much higher mass transport rate

Table 1 Electrochemical reactions relevant to eCO₂RR and their reversible potentials.¹⁷

Half-cell reaction	E° (V vs. RHE)
CO ₂ + 2H ⁺ + 2e ⁻ → HCOOH(aq)	-0.12
CO ₂ + 2H ⁺ + 2e ⁻ → CO(g) + H ₂ O	-0.10
CO ₂ + 6H ⁺ + 6e ⁻ → CH ₃ OH(aq) + H ₂ O	0.03
CO ₂ + 8H ⁺ + 8e ⁻ → CH ₄ (g) + 2H ₂ O	0.17
2CO ₂ + 8H ⁺ + 8e ⁻ → CH ₃ COOH(aq) + 2H ₂ O	0.11
2CO ₂ + 10H ⁺ + 10e ⁻ → CH ₃ CHO(aq) + 3H ₂ O	0.06
2CO ₂ + 12H ⁺ + 12e ⁻ → C ₂ H ₄ (g) + 4H ₂ O	0.08
2CO ₂ + 12H ⁺ + 12e ⁻ → C ₂ H ₅ OH(aq) + 3H ₂ O	0.09
2CO ₂ + 14H ⁺ + 14e ⁻ → C ₂ H ₆ (g) + 4H ₂ O	0.14
3CO ₂ + 16H ⁺ + 16e ⁻ → C ₂ H ₅ CHO(aq) + 5H ₂ O	0.09
3CO ₂ + 18H ⁺ + 18e ⁻ → C ₃ H ₇ OH(aq) + 5H ₂ O	0.10

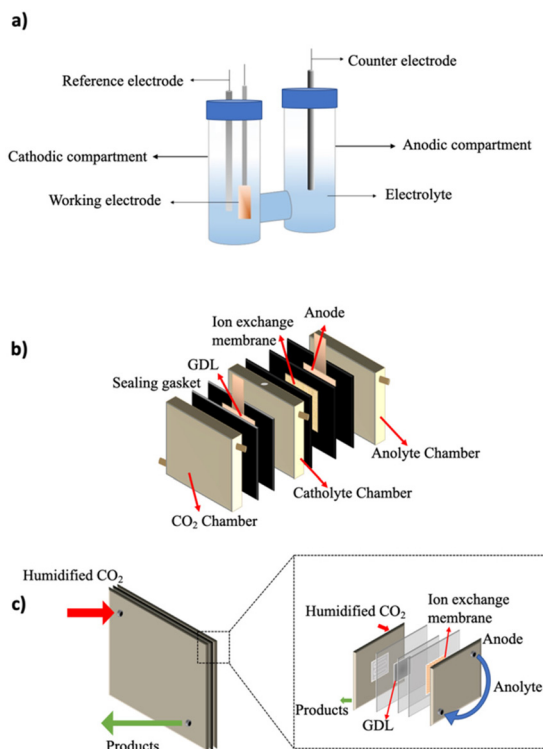


Fig. 2 Illustration of (a) H-cell, (b) flow cell and (c) zero gap MEA electrolyzers used for eCO₂RR.

achieved under these conditions allows an industrially appropriate current density of above 0.5 A cm⁻² to be achieved.^{27–29} Two types of such electrolyzers are commonly used; the flow-cell and membrane electrode assembly (also known as zero-gap) electrolyzers. To fabricate a GDE, a gas diffusion layer (GDL) is coated with a layer of catalyst. GDLs composed of carbon fibers possess high porosity.^{30,31} The structure of the GDE has an impact on the transport of both reactants and products, which affects chemical species adsorption and depletion at the catalyst surface, and hence the overall eCO₂RR performance.^{31,32}

Properties such as porosity, electrical conductivity and hydrophobicity of the GDLs are now being tailored to improve the cell performance.³⁰ A typical flow-cell electrolyzer is shown in Fig. 2(b) and consists of three channels, one for the CO₂ gas flow and the other two for electrolyte (catholyte and anolyte). The GDE is placed in between the CO₂ gas and catholyte channels. The catholyte and anolyte are separated by a polymer exchange membrane (PEM). The PEM used in the cell also greatly impacts the efficiency of the electrolyzer. Cation exchange membranes (CEM) and anion exchange membranes (AEM) are the two types of PEM used depending on the reaction environment. Nafion, a CEM, was explored for eCO₂RR by Delacourt *et al.*³³ and found to suffer from excessive H₂ production and lack of stability in long-term electrolysis. AEMs are better suited for eCO₂RR circumstances because neutral or alkaline media can be used to minimize the competing H₂ evolution reaction. Since anions are far less mobile than

protons, both in solution and in the membrane, an additional driving force (*i.e.* larger overpotential) is needed to overcome the resistance effect. Further, the transport of bicarbonate ions away from the cathode also diminishes the eCO₂RR performance.^{30,34,35} The ratio of CO₂ and the proton source in a flow cell electrolyser is significantly higher than that found in an H-cell electrolyser. Furthermore, during the eCO₂RR, CO₂ is reduced and protons are consumed. Hence, the local reaction environment under the flow-cell conditions may be drastically different from that found under H-cell conditions due to its inherently higher CO₂ mass transport rate and hence higher current density. Accordingly, the results obtained with an H-cell may not be applicable under flow-cell conditions. Consequently, eCO₂RR performance ideally should be assessed under commercially relevant conditions using a flow cell-electrolyser. H-cell and flow-cell electrolysers are mainly used for fundamental research. The presence of a reference electrode in these configurations allows the anode and cathode to be investigated separately. In commercial applications, a reference electrode is not needed. A membrane electrode assembly (MEA) electrolyser is commonly used, inspired by developments of fuel cells, a closely related research area. A MEA electrolyser is composed of two channels, one for CO₂ and the other for anolyte (Fig. 2c). The PEM is sandwiched between the anode and the cathode to minimize the cell resistance. Use of a MEA electrolyser allows the full cell performance, including energy efficiency, to be evaluated.

The catalyst employed in an eCO₂RR is critical in determining the eCO₂RR product selectivity.³⁶ During electrolysis, the CO₂ molecule commonly is adsorbed onto the catalyst present in the cathode. Accordingly, the CO₂ reduction pathway usually depends on the binding strength of CO₂ and its reduced intermediates over the catalyst surface, as depicted in Fig. 3 and 4. CO₂ adsorbed on the catalyst is initially reduced to either carbonyl (*COOH) (an asterisk is used to indicate an adsorbed species) or formyl (*OCHO) intermediates in aqueous electrolyte which can be further reduced to CO and formate, respectively. CO is electroactive. However, if CO binds

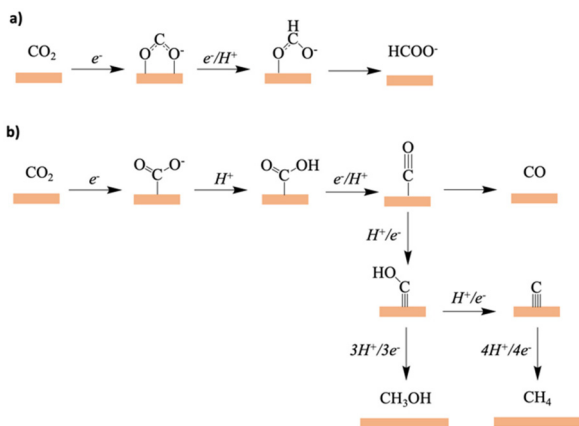


Fig. 3 Illustration of eCO₂RR pathways for the formation of (a) HCOO⁻, (b) CO, CH₃OH, CH₄.

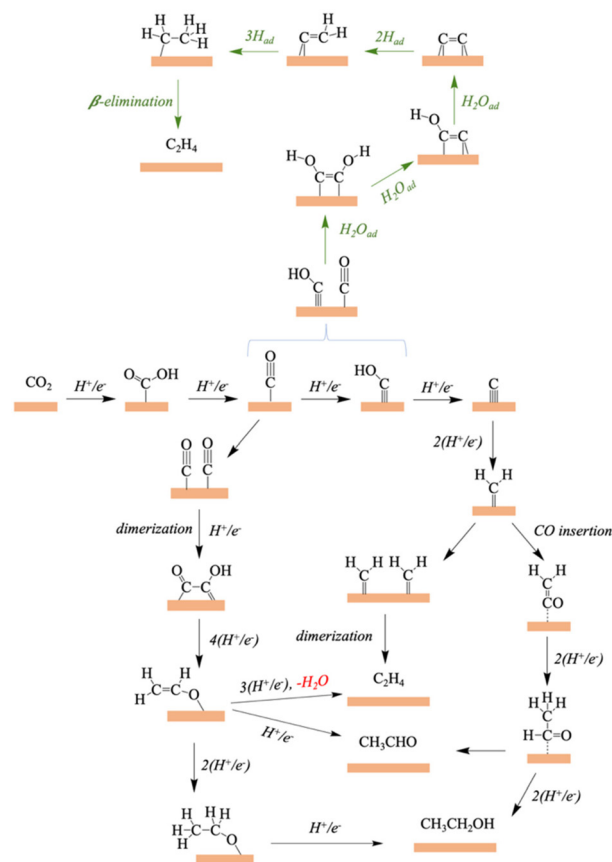


Fig. 4 Illustration of eCO₂RR pathways for the formation of C₂H₄, CH₃CHO and C₂H₅OH.

to the catalyst too strongly, then further reduction to desirable products and their subsequent desorption is inhibited. Poisoning of the catalyst also is likely to occur. Consequently, hydrogen evolution rather than CO₂ reduction may become the favoured reduction pathway. This scenario occurs at Pt, Ni, and Fe electrodes.³⁷ Alternatively, if the *CO binding ability to the catalyst is weak, as applies with metals such as Zn, Au, and Ag, then high selectivity for CO as the product is achieved.³⁷ If the catalyst has moderate binding strength for *CO, further reduction is facilitated, leading to more reduced carbon products. Cu is the only catalyst that has moderate *CO binding energy, leading to unique prospects for formation of a wide range of products. Metals which bind eCO₂RR intermediates too weakly are also likely to bind *H weakly. If these catalysts also possess strong *O affinity, an O-bound *OCHO intermediate will be formed during the initial reaction step which leads to formation of HCOOH.

As noted above, extensive eCO₂RR studies with many metals in the periodic table have revealed that Cu has the unique capability of producing an extensive range of reduced carbon products containing one carbon (C₁), two carbons (C₂) or multi-carbons (*i.e.* more than two carbons; C₂₊). The 16 products detected using Cu catalysts during eCO₂RR and their structures are shown in Fig. 5.³⁸⁻⁴¹ This on one hand demon-

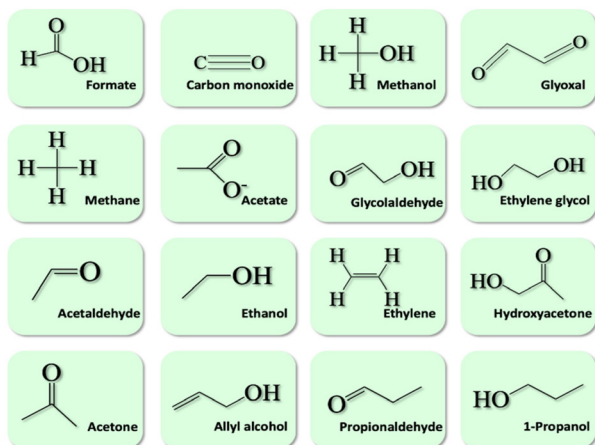


Fig. 5 Graphical representation of products detected during eCO₂RR at a Cu electrode. Adopted and redrawn.³⁸

strates the versatility of Cu catalysts, but on the other hand implies that poor product selectivity is a major challenge to address for commercial applications.

In attempts to address the selectivity issue, nanostructuring, alloying, defect engineering, atomizing, and the synthesis of oxide, sulphide, and nitride-derived Cu as well as other methods have been employed.^{17,21,42–49}

Since alloying is one of the most promising and fundamental strategies for manipulating the reaction pathways to obtain desired products, this review focuses on what has been achieved to date and what might be accomplished in the future by alloying Cu with other eCO₂RR active metals. Several review articles^{17,25,50–54} have been published on the use of alloyed-Cu catalysts. However, they focus on a much broader aspect of eCO₂RR with copper. In this review, the correlation between alloy type and selectivity is examined and future directions in the development of Cu based alloys as effective eCO₂RR catalysts are proposed. Finally, in order to provide an overview of the status of findings with alloyed-Cu catalysts, comparisons of what has been achieved with state-of-the-art Cu alloys and other Cu based catalysts for each eCO₂RR product also are provided in a table at the end of each section.

2. Catalyst performance descriptors

The effectiveness of eCO₂RR is determined by the following descriptors:

(a) Faradaic efficiency: faradaic efficiency (FE) is the indicator of the selectivity of the products obtained with a catalyst during eCO₂RR. FE (reported in %) is calculated by using the formula given in eqn (1)

$$\text{FE (\%)} = \frac{nNF}{Q} \times 100 \quad (1)$$

where N represents the number of moles of the product formed, n is the number of electrons transferred for the formation of one molecule of the product, F is the Faraday con-

stant ($96\,485\text{ C mol}^{-1}$) and Q is the amount of charge passed during electrolysis.

(b) Overpotential: overpotential (η) is the difference between the applied potential (E_{applied}) to obtain a product and the equilibrium potential (E°). It is calculated by using the formula given in eqn (2)

$$\eta = |E_{\text{applied}} - E^{\circ}|. \quad (2)$$

(c) Current density: current density (j) denotes the rate/s of electrochemical reaction/s. It is the catalytic current (i) generated per unit area. Current densities presented in many studies are calculated with respect to the geometric area of the electrode. However, ideally an electrochemical active surface area should be used. Partial current density (j_{product}) indicates the rate of formation of a given product and calculated by the formulae given in eqn (3) and (4)

$$j_{\text{total}} = \frac{i_{\text{total}}(\text{mA})}{\text{area}(\text{cm}^2)} \quad (3)$$

$$j_{\text{product}} = j_{\text{total}} \times \text{FE}_{\text{product}} \quad (4)$$

(d) Stability: the stability is the lifetime of a catalyst for performing eCO₂RR. In general, it is the time period for which a catalyst provides stable activity and selectivity.

3. Strategies for alloying copper

Watanabe *et al.*,⁵⁵ were the first to report eCO₂RR on Cu alloys formed by electroplating. They found that alloy catalysts often exhibited distinctly different catalytic properties from their constituent elemental metals. This feature was later attributed to electronic and geometric effects at the catalyst surface that arise due to alloying.^{17,36,50,56} The electronic effect is generated in an alloy due to the interaction between the constituent elemental metals with different electronic properties and/or the lattice mismatch between two adjacent metals, which alters the binding strength of the eCO₂RR intermediates. The geometric effect arises from the change in the arrangement of the alloyed metals in the catalyst which alters the environment near the active sites for the eCO₂RR intermediates and hence impacts their binding and subsequent reactions.

The structure of the alloy or in other words, its atomic arrangement play a crucial role in determining the selectivity of the products formed during eCO₂RR. Alloy systems can be ordered, intermetallic, disordered, core/shell structured, phase-separated and have high entropy structures (Fig. 6(a)). All these variations have been explored in the design of catalysts for eCO₂RR. Intermetallic alloy systems having an ordered atomic arrangement are the most stable form of alloys. With thermodynamically stable alloy forms, the electronic structure and atomic co-ordination environment can be controlled precisely as they possess a specific crystal structure.⁵⁷ By altering the alloy composition, bond lengths can be tuned, which effects the binding energies of molecules adsorbed on the surface of the catalysts and hence their catalytic behaviour.

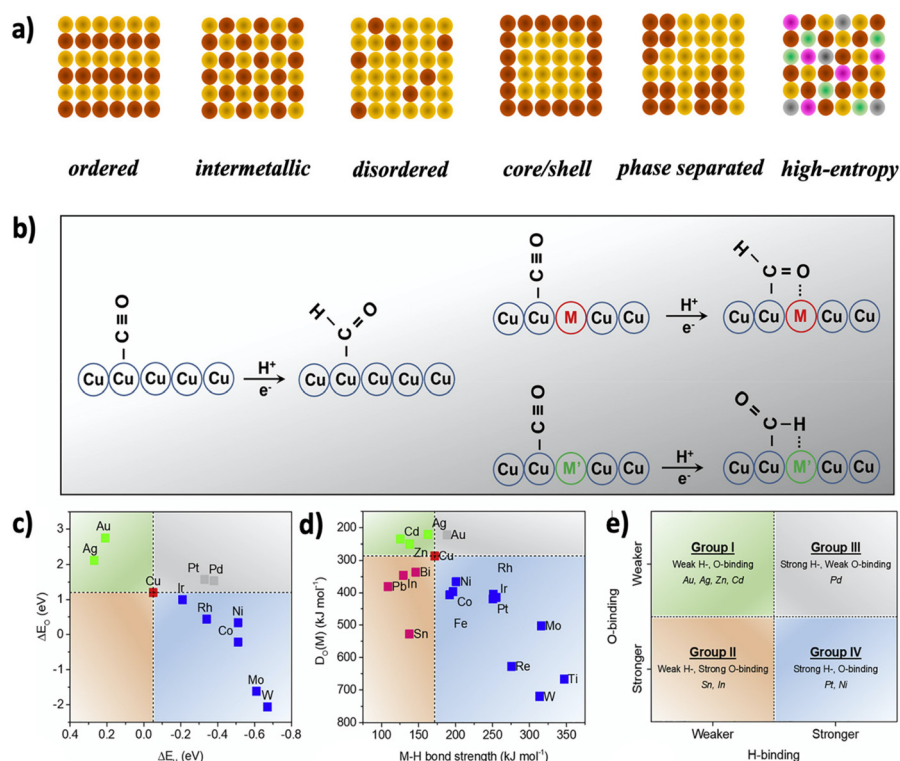


Fig. 6 (a) Alloy systems reported in the literature. (b) Illustration of binding of eCO₂RR intermediates to Cu, Cu–M and Cu–M' alloy catalysts where M is the guest metal with O-affinity and M' is the guest metal with H-affinity, (c) H and O binding energies for designated metals, (d) M–H bond strength shown as a function of bond dissociation enthalpy of metal oxides, (e) classification of metals into different groups with respect to Cu, based on their O and H affinities. Reprinted with permission.²⁵ Copyright 2018, Elsevier Inc.

Disordered alloys are solid-solution mixtures with irregular arrangements of binary atoms. They are generally thermodynamically unstable and their local electronic and the geometry of their structures often changes during the electrolysis process.^{57,58}

The catalytic properties in core/shell structured alloys are regulated by the strain and ligand effects that arise between the core and shell materials. Strain effects such as tensile stress on the surface of atoms arise due to the lattice mismatch between core/shell interface.^{59,60} Ligand effects are short-range and only exist within two or three atomic layers from the surface.⁶¹ Atomic vacancies and metal doping in the shell material also can occur. They are a function of the thickness of the shell and also impact on the product selectivity during electrolysis.⁶² In eCO₂RR, the major benefit of such alloys is the creation of phase interfaces where the products formed at one phase can migrate through the interface and be further reduced to C₂₊ products. These catalytic Cu alloy forms are being extensively studied as a means of increasing the CO environment at Cu and facilitating the production of highly reduced C₂₊ products such as ethylene and ethanol.^{63–65} Details of the efficiency of phase separated alloy catalysts and the CO spill-over effect are discussed below.

High entropy alloys (HEAs) are a novel class of alloys recently being investigated for eCO₂RR. These alloys are formed by about five metals, with each element constituting

between 5–35% and with variation in the elemental ratios. Accordingly, a large variety of elemental combinations is possible in the alloy and consequently a wide range of catalytic capabilities can be achieved.^{66,67}

Compared to bimetallic alloys, HEAs provide superior stability and durability particularly under harsh reaction conditions such as high temperature and high electrochemical potential due to their high mixing entropy and hence low Gibbs free energy as quantified by eqn (5).^{67,68}

$$\Delta G_{\text{mix}} = \Delta H_{\text{mix}} - T\Delta S_{\text{mix}} \quad (5)$$

where ΔG_{mix} is the change in the Gibbs free energy of the alloy system, ΔS_{mix} is the change in the entropy of the alloy system, ΔH_{mix} is the change in the enthalpy of the alloy system, and T is the temperature. In addition, the random atomic arrangement in the lattice leads to variation in lattice potential energy at different locations which provides a higher diffusion activation energy that contributes to the stability of the alloy.⁶⁹ Furthermore, due to lattice distortion effects resulting from the presence of multiple elements, HEAs have more flexibility to alter their chemical and physical properties.^{67,70,71} In terms of electrolysis, this effect helps the optimization of the coordination environment of atoms on the catalyst surface as well as the adsorption and desorption energies of key reaction intermediates.⁶⁸ Similar to the synergistic effects in bimetallic

alloys, HEAs possess a “cocktail effect” resulting from the interactions between multiple elements. As a result of a range of novel features, these alloys offer new properties with respect to corrosion and oxidation resistance, and other mechanical properties.^{68,70} The unique characteristics of HEAs change the scaling relationships and provide a wide window of opportunity for manipulating the reaction pathways and overcoming the limitations of conventional alloys.⁷²

Grouping of metals according to their binding affinities for *H and *O relative to copper was proposed to facilitate the selection of a guest metal to form a binary alloy with Cu.²⁵ In the alloy, the bimetallic catalyst material will have two binding sites available to achieve stabilisation of a desired intermediate. For example, metals which produce hydrogen have a more robust *H and *O affinity than Cu while those with weaker *O and *H binding strengths, produce CO *via* a *COOH intermediate. Metals having a lower H affinity but higher O-affinity than Cu favour formation of the *OCHO intermediate, which leads to formation of formate. Hence, metals with good O-affinity can be alloyed with Cu to stabilize O-bound intermediates, while those having H-affinity can stabilize H binding in the intermediates, as illustrated in Fig. 6(b). Fig. 6(c–e) show a comparison of O and H affinities of some metals with respect to copper and the grouping of metals. In the following sections, the factors that govern the selectivity of Cu alloy towards each product are summarized and discussed to guide the development of advanced alloy catalysts. Examples of the state-of-the-art Cu alloys and other Cu based catalysts for each product are also given.

4. Copper alloys for selective reduction of CO₂

4.1 Formate and CO

4.1.1 Formate. Catalysts obtained by alloying Cu with formate favouring metals, such as Bi,^{74–78} Sn,^{79–83} Pb,^{84–88} Pd,^{89–93} Mo⁹⁴ and Sb,^{95–97} particularly as intermetallics, composites and surface alloys, often produces formate with good efficiency.^{98,99}

Copper modified with Pd and Pd–H produces significant amounts of formate. For example, early studies by Fujishima and co-workers showed substantial formate production at the Pd modified Cu electrodes.^{100,101} Zhang *et al.*¹⁰² have synthesised a stannate derived Sn–Cu bimetallic catalyst *via* cation exchange with Na₂Sn(OH)₆ followed by annealing and electroreduction. This Cu–Sn catalyst produced 86% formate at –1.1 V *vs.* RHE in 0.5 M NaHCO₃ with a partial current density of ~11 mA cm^{–2}. These authors also reported that a slight increase in formate faradaic efficiency could be achieved by incorporating the Cu–Sn catalyst into reduced graphene oxide along with a major increase in partial current density. The Hod group¹⁰³ reported the favorable formation of formate with 3D structured Cu₂S catalysts prepared by an electrochemically driven cation exchange mechanism. To prepare such electrodes, pre-synthesised CoS_x nanosheets based 3D structures were

taken as templates and a cation exchange reaction with Cu²⁺ (where Co²⁺ ions in CoS₂ were replaced with Cu^{1+/2+} ions) was performed electrochemically in a 0.1 M LiClO₄-dimethylformamide electrolyte solution containing 3 mM Cu(NO₃)₂ by applying a potential of –0.47 V *vs.* normal hydrogen electrode for varied time periods. The amount of Co²⁺ replaced with Cu²⁺ ions was varied by controlling the charge passed during the electrochemical reaction. The formate production was found to be somewhat dependent on the percentage of Co²⁺ ions exchanged by Cu²⁺. With a Cu₂S sample obtained by passing 2.0 C of charge, 87.3% of formate was produced with a partial current density of 19 mA cm^{–2} at –0.9 V *vs.* RHE in 0.1 M NaHCO₃ solution saturated with CO₂. The authors reported that variation of the grain boundary obtained with the modulation of binary cations was the origin of the change in the product selectivity.

Ag is an excellent CO formation catalyst.¹⁰⁴ However, AgCu alloy can be designed to favour formate formation. Recent work on CuAg bimetallic nanoarchitectures demonstrated the importance of electronic and geometric modulation of the catalyst with respect to product selectivity.⁷³ The sponge-like Ag₉₁Cu₉ and coralline Ag₆₅Cu₃₅ nanoalloys prepared by anodising Ag₅₂Cu₃₉Sn₉ alloy foil induced a dramatic change in the selectivity of CO *versus* HCOOH. Changing the composition of the AgCu alloy, modifies the electronic arrangement of Ag and Cu and hence the binding strengths of intermediates as demonstrated by *in situ* Raman spectroscopic measurements shown in Fig. 7. Thus, on increasing the Cu content of the AgCu alloy, the selectivity shifted from favouring CO to forming formate with a faradaic efficiency of 96%.

Table 2 provides details of some state-of-the-art Cu alloys and other Cu based catalysts for selective formate production.

4.1.2 CO. Metals such as Zn,^{105–109} Ag,^{110–114} Au,^{115–119} Pd,^{89,120–123} which are highly selective for CO, have been alloyed with copper to enhance CO production in many studies. Post-transition metal such as In^{124–128} and Sn^{129–133} that are selective for formate generation also have been com-

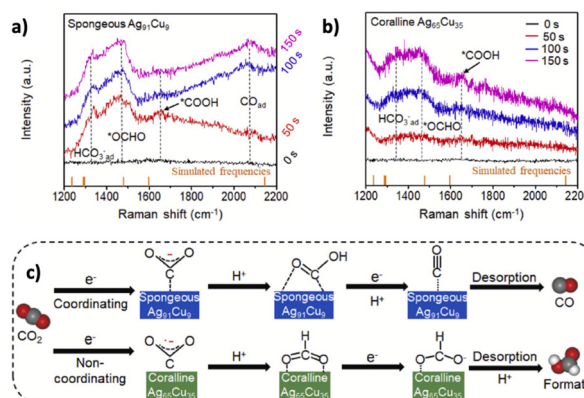


Fig. 7 (a and b) *In situ* Raman spectra obtained during electrolysis and (c) proposed reaction pathways on sponge-like Ag₉₁Cu₉ and coralline Ag₆₅Cu₃₅ alloy nanoarchitectures. Reprinted with permission.⁷³ Copyright 2020, Elsevier Ltd.

Table 2 Examples of formate production with state-of-the-art Cu alloys and other Cu based catalysts

Catalyst	Electrolysis cell	Electrolyte	Potential (V vs RHE) ^a	FE (%)	j_{HCOO^-} (mA cm ⁻²)	Ref.
Cu–Sn/rGO	H-cell	0.5 M NaHCO ₃	–1.0	87.4	20.7	102
Cu ₂ S-2.0C	H-cell	0.1 M NaHCO ₃	–0.9	87.3	19.1	103
Coralline Ag ₆₅ Cu ₃₅	H-cell	0.1 M KHCO ₃	–1.0	91.8	16.8	73
CuCd@Cu-20,45	H-cell	0.5 M NaHCO ₃	–1.1	70.1	26.8	99
Porous Cu _{6,26} Sn ₅	H-cell	0.1 M KHCO ₃	–1.1	97.8	~30	136
CuBi-100	H-cell	0.5 M KHCO ₃	–1.0	94.7	13	137
Cu/Au	H-cell	0.5 M KHCO ₃	–0.6	81	10.4	138
Sn–Cu	H-cell	0.1 M KHCO ₃	–0.95	92	~10	139
Cu@Sn nanocones	H-cell	0.1 M KHCO ₃ + 0.3 M KCl	–1.1	90.4	52	140
CuSn alloy@Cu doped SnO	H-cell	0.5 M KHCO ₃	–1.2	95.4	30.3	141
Cu _{11.5} In _{88.5} -OH	H-cell	0.1 M KHCO ₃	–1.1	85	~10	142
MOF derived CuBi	H-cell	0.5 M KHCO ₃	–0.77	100	67.9	143
Cu ₂ O/CuO/CuS	H-cell	0.1 M KHCO ₃	–0.7	84	20	144
CuBi NPs	H-cell	0.1 M KHCO ₃	–0.99	96	12.5	145
CuS	H-cell	0.1 M KHCO ₃	–0.8	80	18	146
HCS/Cu-0.12	H-cell	0.5 M KHCO ₃	–0.81	82.4	26	147
S-doped OD-Cu	H-cell	0.1 M KHCO ₃	–0.8	73.6	13.9	148
Cu ₆ Sn ₅ /Sn	Flow cell	1 M KOH	–1.0	86.7	103	149
CuSn/Sn	Flow cell	1 M KHCO ₃	–0.7	84.2	~26	150
CuBi	MEA	0.5 M KHCO ₃	–1.07	98.3	55.6	151

^a Potential at which maximum FE was obtained.

combined with copper to tune the selectivity towards CO. This was achieved by alloying copper with metals where in most of the cases, at least one of the following situations exists: the alloy possesses (i) copper in an oxidized form, (ii) a lattice mismatch between the copper phase and the guest metal phase, (iii) a higher shared boundary between the copper and guest metal phases. Consequently, synergistic geometric and electronic effects emerge which influence product formation. Various factors influencing the CO selectivity over the copper-based bimetallic alloys are discussed below.

In early studies in the 1990s, Watanabe *et al.*^{55,134} prepared CuZn alloys electrodes by electroplating onto gold electrodes and studied the eCO₂RR efficiency in 0.05 M KHCO₃ electrolyte. They reported that higher concentrations of Zn in the CuZn alloy led to an increase in CO generation compared to the use of pure Cu. The only other eCO₂RR product found was HCOO[–]. The Berlinguette group¹³⁵ studied eCO₂RR using brass and bronze catalysts. They prepared 21 catalysts with varying Cu : Zn : Sn ratios by drop casting methanolic solutions of Cu, Zn, and/or Sn in the required proportions onto a titanium substrate followed by multiple exposures to near-infrared radiation and electroreduction to produce alloy films. Four-hour eCO₂RR experiments with these film electrodes generated H₂, CO and HCOOH as products. Cu–Zn–Sn alloys with Sn < 20% generated mainly syngas with the H₂ : CO ratio depending on the ratio between Cu and Zn. Hu *et al.*¹⁵² prepared Cu–Zn catalysts by annealing a brass substrate at 500 °C in an Ar atmosphere prior to electroreduction of ZnO formed in the Zn metal annealing step, as shown in Fig. 8(a). On annealing, uniform 200 nm particles formed over the brass substrate. Zn and Cu have different melting points of 419.5 °C and 1085 °C, respectively. Accordingly, Zn melts at 500 °C and migrates all over the surface and is then converted to ZnO on exposure to air while Cu remains unchanged. The resultant catalyst used

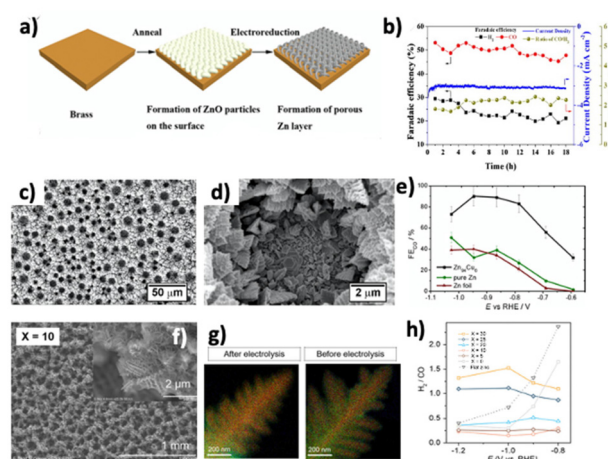


Fig. 8 (a) Schematic illustration of the preparation of Zn–Cu catalysts from brass foil, (b) faradaic efficiencies, current density and CO/H₂ ratio obtained during eCO₂RR as a function of time. Reprinted with permission.¹⁵² Copyright 2018, Elsevier B.V. (c and d) SEM images of the Zn₉₄Cu₆ catalyst, (e) CO faradaic efficiencies obtained with designated CuZn and Zn catalysts. Reprinted with permission.¹⁵³ Copyright 2018, the American Chemical Society. (f) SEM, (g) scanning transmission electron microscope – energy-dispersive X-ray spectroscopy (STEM-EDXS) mapping images of Zn₉₀Cu₁₀ catalyst (h) ratio of H₂/CO faradaic efficiencies obtained with designated Zn_{100–x}Cu_x catalysts. Reprinted with permission.¹⁵⁴ Copyright 2019, Wiley-VCH Verlag GmbH & Co. KGaA, Weinheim.

for eCO₂RR over a range of applied potentials produced syngas with a 2 : 1 ratio of CO : H₂ (Fig. 8(b)) which is suitable for the Fischer–Tropsch process. The Broekmann group¹⁵³ synthesised Zn₉₄Cu₆ alloy foams, as displayed in Fig. 8(c and d), by dynamic hydrogen bubble template assisted electrodeposition onto a Cu plate. eCO₂RR with this catalyst produced 90% CO at –0.95 V vs. RHE in aqueous 0.5 M KHCO₃ electrolyte

(Fig. 8(e)). Lamaison *et al.*¹⁵⁴ also designed Cu–Zn alloy foam catalysts by electrodeposition of Cu onto Zn plates in their eCO₂RR application. These foams were composed of a dendritic CuZn alloy structure (Fig. 8(f and g)). As shown in Fig. 8(h), bulk electrolysis experiments performed under eCO₂RR conditions with 0.1 M CsHCO₃ as the electrolyte revealed an increase in the H₂ : CO ratio with increasing Cu content.

CuAu alloys, like CuZn ones, have been extensively studied for eCO₂RR generation of CO. For example, enhancement of CO at Au_xCu_{100-x} alloy catalysts has been reported by Christophe *et al.*¹⁵⁵ In particular, a significant increase in the production of CO was observed with use of an Au₅₀Cu₅₀ catalyst. CO desorption at Au sites is promoted by the presence of adsorbed CO on Cu sites due to dipole repulsion between CO molecules at adjacent sites. Consequently, the presence of both Cu and Au in an optimal ratio enhances CO product formation. Bimetallic AuCu catalysts also induce an electronic effect which results in the shift in the d-band centre and hence changes the binding strength of the reaction intermediates. As a result of the shift in the d-band centre, the way the metals in the AuCu alloy interact with adsorbed intermediate also changes. The implication of this feature was examined by Kim *et al.*¹⁵⁶ who found that the evolution of CO increases with increasing Au content in the AuCu alloy. In the case of Au₃Cu, they reported a mass activity (*i.e.* current density normalised to the amount of catalyst loaded) exceeding 200 A g⁻¹ at -0.73 V *versus* RHE, which represents a ten-fold increase in comparison with pure Cu. Surface valence band photoemission spectral analysis suggested that the improved selectivity towards CO with Au₃Cu could also be due to the synergistic electronic and geometric effects associated with AuCu nanoparticle formation that occurred during the eCO₂RR. The atomic arrangement of Au and Cu in AuCu bimetallic alloy nanoparticles (either ordered or disordered) also can have an impact. Among AuCu alloy nanoparticles, the nanoparticles with an ordered arrangement of Au and Cu (o-AuCu) exhibited a factor of two improvement in selectivity for CO (Fig. 9(a)).¹⁵⁷ STEM (Fig. 9(b)) and X-ray absorption spectroscopy (XAS) (Fig. 9(c)) analyses revealed that a thin layer of segregated Au atoms was formed on the surface of o-AuCu nanoparticles during electrolysis. DFT calculations suggested that the lattice strain of 6% associated with the Au layer due to the underlying AuCu lattice affected the catalytic activity of o-AuCu nanoparticles.

Intermetallic CuPd alloy catalysts provide another example of a Cu alloy that enhances CO generation during the eCO₂RR process.^{159–161} Excellent selectivity for CO with a faradaic efficiency of 86% and a partial current density of 6.9 mA cm⁻² was obtained on Pd₈₅Cu₁₅ alloy nanoparticles (Fig. 10(a and b)).¹⁵⁸ On increasing the Cu content in the alloy, the efficiency decreased. X-ray absorption near edge structure (XANES) analysis shown in Fig. 10(c and d) for the PdCu alloy nanoparticles revealed that variation in Pd–Pd and Cu–Cu bond lengths alters the binding strengths of the CO₂ reduction intermediate (*CO). Further, the formation of Pd and Cu oxides at higher Cu to Pd ratios leads to suppression of CO evolution during

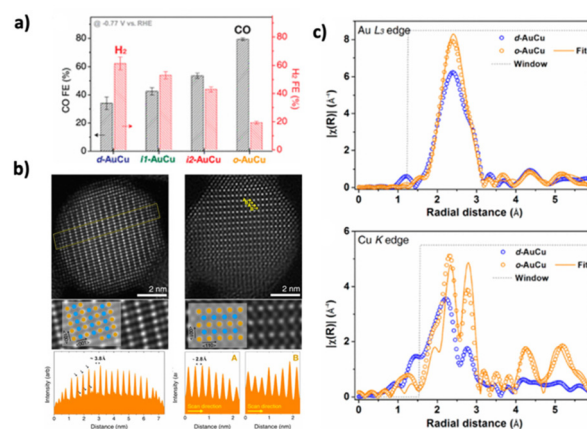


Fig. 9 (a) Faradaic efficiencies for generation of H₂ and CO using AuCu alloy nanoparticles with designated atomic ordering, (b) high-angle annular dark-field imaging – scanning transmission electron microscope (HAADF – STEM) analysis of ordered AuCu nanoparticles, and (c) XAS analysis of ordered and disordered AuCu nanoparticles. Reprinted with permission.¹⁵⁷ Copyright 2017, the American Chemical Society.

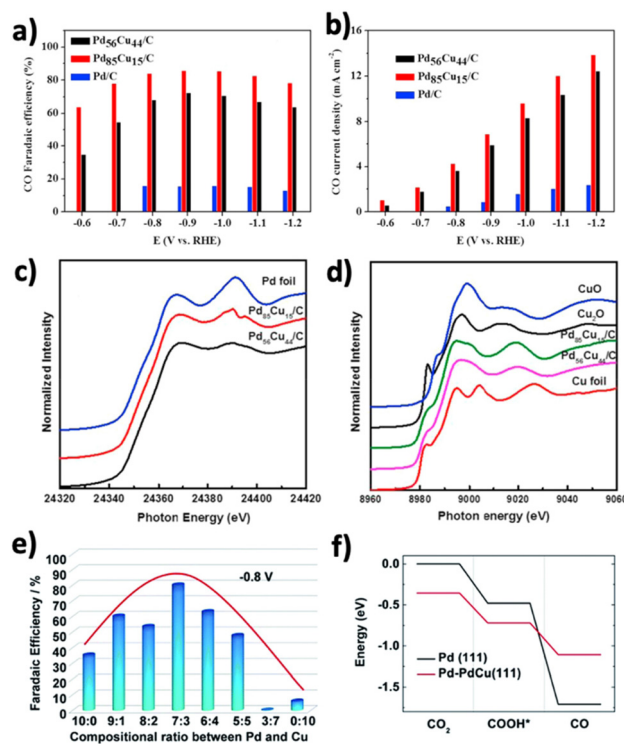


Fig. 10 (a) CO faradaic efficiency Pd_xCu_{1-x} nanoparticles, (b) CO partial current densities obtained using designated Pd_xCu_{1-x} nanoparticles. (c) Pd K edge, and (d) Cu K edge XANES spectrum of a PdCu/C catalyst. Reprinted with permission.¹⁵⁸ Copyright 2016, Elsevier Ltd. (e) A comparison of CO selectivity at PdCu NPs with specified compositions, (f) free energy diagram for CO₂, *COOH and CO on a Pd terminated PdCu (111) surface. Reprinted with permission.¹⁵⁹ Copyright 2016, the Royal Society of Chemistry.

electrolysis. CuPd alloys with a 1.5 : 8.5 atomic ratio bind CO less strongly on the surface than with pure Pd, resulting in an enhancement of CO production.

The study by Li *et al.*¹⁵⁹ with mesoporous CuPd alloys also shows composition-dependent activity for CO₂ conversion (Fig. 10(e)). An alloy with a Cu to Pd ratio of 3 : 7 gives high CO yields, which diminishes with an increase in the Cu content. DFT calculations predict that Pd terminated Pd₇Cu₃ alloy surfaces faceted with (111) planes will favour adsorption of a *COOH intermediate and weaken adsorption ability of *CO resulting in excellent CO selectivity (Fig. 10(f)). Due to the differences in electronegativity of Cu and Pd, the presence of Cu atoms adjacent to Pd in a Cu–Pd alloy increases the adsorption of *COOH and CO on Pd surfaces, making Pd sites the active sites for the eCO₂RR.

Oxide derived Cu–M (M = In/Sn) catalysts have been surveyed for their ability to convert CO₂ to CO. In and Sn in their metallic,¹⁶² oxide^{163–165} or chalcogenide^{166–168} forms produce HCOO[−] from the eCO₂RR. However, when oxides of these post-transition metals are combined with copper, scaling relationships can be altered due to synergistic effects associated with the lattice mismatch between Cu/CuO and MO₂ (M = In/Sn). The result is that product selectivity is shifted from formate to

CO with remarkably high faradaic efficiencies. For example, the study by Li *et al.*,¹⁶⁹ shows that the change in product distribution as a function of the thickness of the SnO₂ coated onto Cu nanoparticles (Fig. 11(b and c)). With a thin SnO₂ shell (0.8 nm), the product selectivity for CO had a FE of >90% at −0.6 V vs. RHE with little formate produced. In contrast, with a thick SnO₂ shell (1.8 nm), the eCO₂RR behaviour is similar to that at Sn and formate is the major product. According to the results of DFT calculations, the synergistic effect emerged because of the compression strain and self-doping of Cu into the thinner SnO₂ shell layer during electrolysis which shifted selectivity from HCOO[−] to CO. A related study using Cu–In₂O₃ core–shell nanoparticles (Fig. 11(d))¹⁷⁰ also showed that the thinnest shell provided a shift in selectivity to CO (Fig. 11(e and f)). With the thinner In₂O₃ shell, the lattice mismatch between Cu and In₂O₃ facilitates self-doping of Cu into the In₂O₃ shell resulting in alteration of binding energies of *COOH and *OCHO intermediates, which leads to the generation of CO and HCOOH, respectively. Cu oxide surfaces decorated with Sn also enhance conversion of CO₂ to CO.

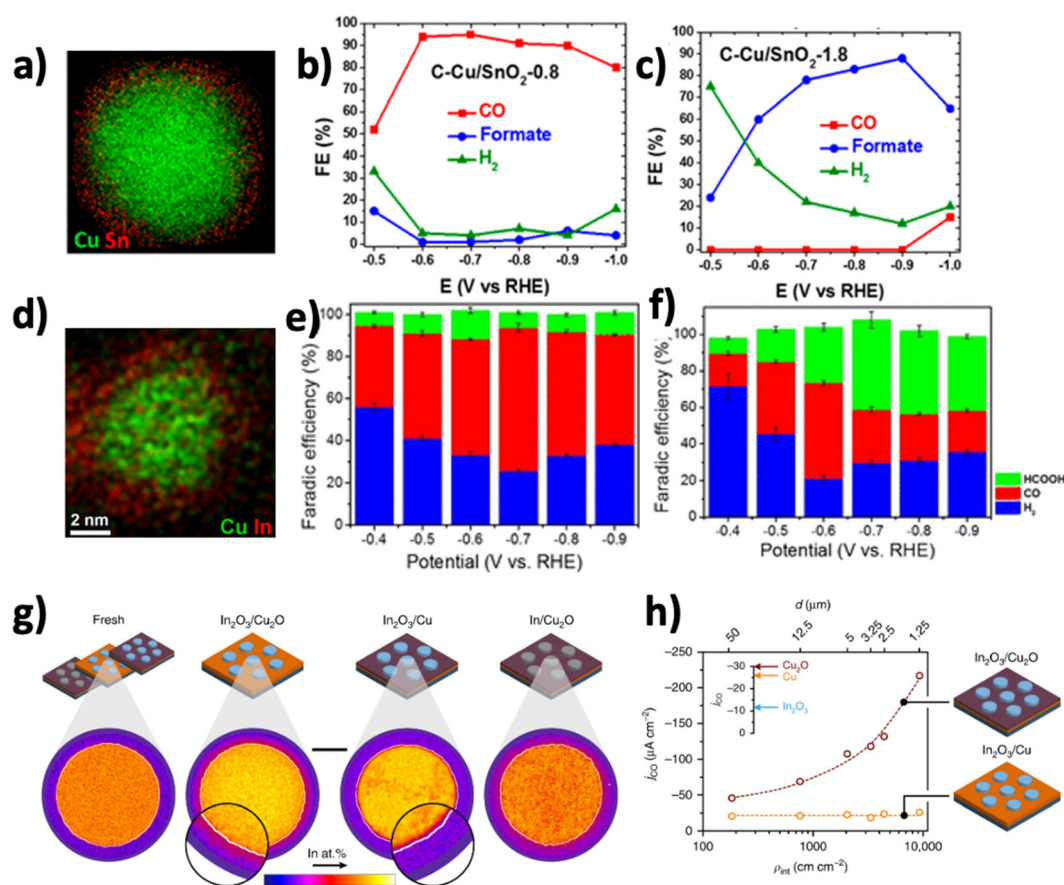


Fig. 11 (a) Electron energy loss spectroscopy (EELS) elemental mapping of Cu/SnO₂ nanoparticles. (b) and (c) faradaic efficiencies of products obtained during eCO₂RR with Cu/SnO₂ NPs having a shell thickness of (b) 0.8 nm and (c) 1.8 nm. Reprinted with permission.¹⁶⁹ Copyright 2017, the American Chemical Society. (d) EDS elemental mapping of C–Cu/In₂O₃ nanoparticles. (e) and (f) faradaic efficiencies of products obtained during eCO₂RR with Cu/In₂O₃ nanoparticles having a shell thickness of (e) 0.4 nm and (f) 1.5 nm. Reprinted with permission.¹⁷⁰ Copyright 2018, the American Chemical Society. (g) elemental mapping showing the formation of indium islands when undertaking eCO₂RR with CuIn electrodes, (h) CO partial current density obtained during eCO₂RR at −0.6 V with In₂O₃/Cu₂O and In₂O₃/Cu electrodes as a function of island diameter and interfacial density. Reprinted with permission.¹⁷¹ Copyright 2018, Nature Publishing Group.

Zhao *et al.*¹⁷² decorated electrochemically generated CuO nanowires with Sn. An optimal loading of Sn produced 90% of CO at -0.8 V vs. RHE in 0.1 M KHCO₃ electrolyte. Sarfraz *et al.*¹⁷³ electrochemically deposited Sn on a Cu sheet. With a Sn loading of 3.9 mol cm⁻², 90% of CO was obtained at -0.6 V vs. RHE in 0.1 M KHCO₃. However, with higher Sn loadings, selectivity reverted to HCOO⁻. Zeng *et al.*¹⁷⁴ used Cu–Sn foam based dendritic structures, where the Cu dendrites were decorated with small amounts of SnO_x to form a Cu/CuO_x–SnO_x core/shell structure. This allowed 93–94% CO to be formed between -0.75 V and -0.9 V vs. RHE in 0.1 M KHCO₃.

Studies by the Takanabe group^{175,176} employing In coated on oxide derived (OD)-Cu¹⁷⁵ and CuInO₂ derived Cu–In alloy¹⁷⁶ catalysts again revealed a drastic increase in CO selectivity. DFT calculations suggest that the replacement of a Cu atom with an In atom would suppress *H adsorption with *CO adsorption energy unchanged, resulting an increase in the FE of CO while decreasing HER. A study by Larrazábal *et al.*¹⁷⁷ proposed that an *in situ* generated metastable In(OH)₃ phase in the oxide plays an important role in the use of Cu–In catalysts. The core–shell structured Cu–In catalysts, with In(OH)₃ as the shell evolved upon repeated voltammetric cycles of potential on the initial Cu–In in a 0.1 M KHCO₃ solution, exhibited a significant increase in the CO selectivity. The same group also reported the dependence of CO selectivity on the nature of the Cu–In interfaces.¹⁷¹ In their study, the authors microfabricated In₂O₃ dots on Cu and Cu₂O surfaces. In₂O₃ dots on Cu left the selectivity unaltered, whereas those on Cu₂O led to a drastic improvement in selectivity for CO. When the interfacial density is low, both catalysts showed similar activity. However, when the density of metal–oxide interface was increased, the catalytic activity towards CO was enhanced with In₂O₃/Cu₂O, as shown in Fig. 11(h). The irregular layer of Cu formed by reduction of Cu₂O enhances the diffusion of In (Fig. 11(g)) giving rise to a synergistic effect at Cu–In interfaces. This study reveals that both the interfaces and Cu₂O are essential for achieving synergistic effects that tune the selectivity pathway.

State of the art Cu alloys and other Cu based catalysts that favoured CO production are tabulated in Table 3.

4.2 Highly reduced C₁ products

4.2.1 Methane. The formation of methane on Cu based bi-metallic catalysts with the eCO₂RR was reported in early literature using Cu modified Pd and PdH catalysts.^{100,101} The amount of hydrogen absorbed in PdH has a significant effect on the selectivity for methane. For instance, the methane production was enhanced initially at Cu–PdH catalysts with lower coverages of adsorbed H on Pd. However, there was no further increase in CH₄ at higher concentrations of adsorbed H.

Electrochemical conversion of CO₂ to methane has been explored on Cu–Zn catalysts. Cuenya and co-workers¹⁸⁴ studied the eCO₂RR product selectivity dependence on the composition and structure of CuZn nanoparticles derived from adoption of the inverse micelle encapsulation method. Their bulk electrolysis eCO₂RR experiments with Cu_{100-x}Zn_x revealed the presence of Zn from 10% to 50% increased CH₄ formation while higher 50% Zn levels suppressed CH₄ formation and increased that of CO (Fig. 12(a)). The XANES studies suggested that the Cu–ZnO interface is needed to generate CH₄ or other hydrocarbons while CuZn (brass) alloy favours formation of CO/H₂ mixtures. Due to the faster reduction rate with CuO than with ZnO, in samples with a lower Zn content (Cu_{100-x}Zn_x, $x < 50$), the Cu–ZnO interface prevailed in the initial period of time during eCO₂RR, when the hydrogenation of CO species was favoured at Cu sites. When both oxides were fully reduced, brass nanoparticles were generated so that Cu loses its ability to stabilise *CH_y ($y = 1-3$) intermediates due to ligand effects and hence CO and H₂ were released as the electrolysis products.

Ceria (CeO₂) is another material that facilitates CH₄ production in the presence of Cu sites during eCO₂RR.¹⁸⁵ Doping CeO₂ nanorods with Cu establishes a strong interaction between atomically dispersed Cu sites and CeO₂. Accordingly, multiple oxygen vacancies are induced around Ce under eCO₂RR conditions, which provides an effective site for electro-

Table 3 Examples of the state-of-the-art Cu alloys and other Cu based catalysts that favour CO production

Catalyst	Electrolysis cell	Electrolyte	Potential (V vs. RHE) ^a	FE (%)	j_{CO} (mA cm ⁻²)	Ref.
Zn ₉₄ Cu ₆ nanofoam	H-cell	0.5 M KHCO ₃	-0.9	90	~4.5	153
o-AuCu NPs	H-cell	0.1 M KHCO ₃	-0.8	~80	1.4	157
Mesoporous Pd ₇ Cu ₃	H-cell	0.1 M KHCO ₃	-0.8	80	~1.5	159
Pd ₈₅ Cu ₁₅ /C	H-cell	0.1 M KHCO ₃	-0.9	86	6.9	158
C–Cu/SnO ₂ -0.8	H-cell	0.5 M KHCO ₃	-0.7	93	~13	169
Cu–Sn10 nanowires	H-cell	0.1 M KHCO ₃	-0.8	90	4.5	172
Cu–Sn	H-cell	0.1 M KHCO ₃	-0.6	90	1.0	173
Cu–Sn foam	H-cell	0.1 M KHCO ₃	-0.8	94	3	174
PTFE-Cu	H-cell	0.1 M KHCO ₃	-0.4	71	1.5	178
Cu NPs-700	H-cell	0.1 M KHCO ₃	-0.6	75.6	3.8	179
Cu–N ₂ /GN	H-cell	0.1 M KHCO ₃	-0.5	81	~2	180
Cu/CNT	PEM Cell ^b	0.1 M KHCO ₃	-3.5^c	75.7	12.2	181
CuPd nanosheets	Flow cell	1 M KOH	-0.6	71	58	161
2.7 nm PdCu NPs	Flow cell	1 M KOH	-1.0	82	80	182
Sb–Cu ₂ O	Flow cell	0.1 M KOH	-0.8	96	60	183

^a Potential at which maximum FE was obtained. ^b Polymer electrolyte membrane electrolysis cell. ^c Cell voltage.

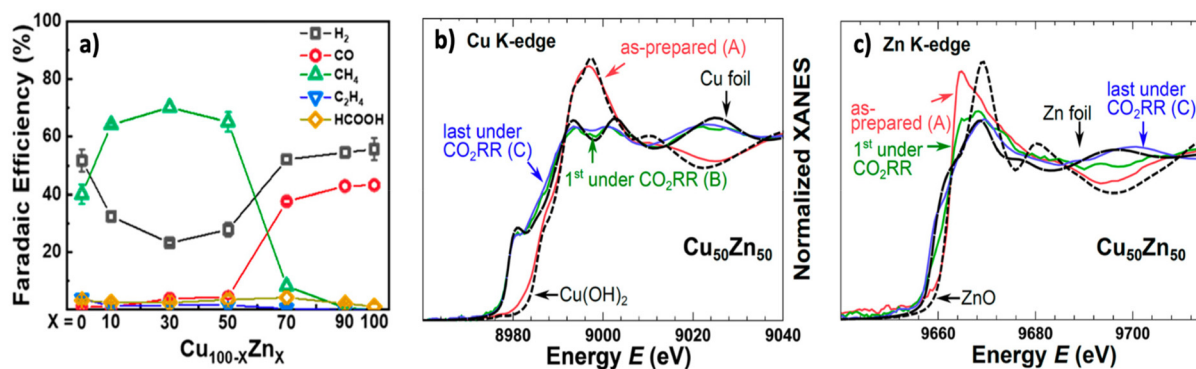


Fig. 12 (a) A comparison of faradaic efficiencies for the products obtained by eCO₂RR with Cu_{100-x}Zn_x catalysts, (b) Cu K-edge and (c) Zn K-edge XANES spectra of Cu₅₀Zn₅₀ nanoparticles: spectrum A – as-prepared sample, spectrum B – immediately after the onset of eCO₂RR and spectrum C – after 7 h of electrolysis. Reprinted with permission.¹⁸⁴ Copyright 2019, the American Chemical Society. Source: <https://pubs.acs.org/doi/10.1021/jacs.9b10709>. Further permissions related to the material excerpted should be directed to the ACS.

catalytic reduction of CO₂ to CH₄. CeO₂ nanorods doped with 4% Cu produce ~58% of CH₄ at -1.8 V vs. RHE with a current density of 56 mA cm⁻² in 0.1 M KHCO₃, as shown in Fig. 13(b). The Buonsanti group¹⁸⁶ has explained the importance of the interaction between CeO₂ and Cu in facilitating CH₄ formation by examining outcomes of the eCO₂RR with Cu/CeO₂ heterodimer (HD) nanoparticles as well as physical mixtures of Cu and CeO₂ nanoparticles. In addition to the presence of CeO₂ interactions with Cu nanoparticles described above, the size of the Cu nanoparticles was shown to play a significant role in the improvement in CH₄ selectivity. The 36 nm sized Cu/CeO₂ HD nanoparticles enhanced eCO₂RR per-

formance over HER and the selectivity achieved for CH₄ reached 54% at -1.2 V vs. RHE in 0.1 M KHCO₃. In contrast, the Cu and CeO₂ physical mixture resulted in only <10% of methane. XAS data (Fig. 13(d)) confirmed that during eCO₂RR, Ce⁴⁺ in Cu/CeO₂ HD is reduced to Ce³⁺. On the basis of DFT calculations, the eCO₂RR intermediates (*COOH, *CHO, H₂CO*, H₃CO*) were postulated to be adsorbed onto both Cu and Ce sites with an O-vacancy in order to form CH₄. Thus, oxygen vacancies played a crucial role in enhancing methane selectivity.

Chang *et al.*¹⁸⁷ have studied the impact of structural reconstruction of Cu_{100-x}Ag_x catalysts on product selectivity using the techniques of *in situ* grazing-angle X-ray scattering/diffraction, X-ray absorption spectroscopy, and Raman spectroscopy. The authors showed that Cu₆₈Ag₃₂ nanowires underwent dynamic oxidation–reduction cycles on the nanowire surface resulting in an inter diffusion of Ag and Cu atoms prior to stabilization of the metallic states of Cu and Ag. The high methane FE of 60% is three times to that achieved with pristine Cu nanowire.

4.2.2 Methanol. Early studies by Watanabe *et al.*^{55,193} on the electrochemical generation of CH₃OH by reduction of CO₂ used Cu based alloys, particularly CuNi ones. Enhancement of methanol selectivity was reported by Jia *et al.*¹⁹⁴ using a AuCu alloy catalyst deposited electrochemically onto a nanoporous Cu film (NCF). The FE for methanol production was dependent on the alloy composition. With a Cu_{63.9}Au_{36.1}/NCF catalyst, a FE of 15.9% was achieved which was 19 times higher than with pristine Cu.

Albo *et al.*¹⁹⁵ studied the use of Cu₂O/ZnO catalysts for eCO₂RR in a flow cell using electrodes prepared by airbrushing commercially available Cu₂O and ZnO nanoparticles over carbon paper. They found that the ratio of Cu₂O/ZnO has a significant effect on methanol generation during eCO₂RR. Cu₂O/ZnO with 2 : 1 ratio generated a high selectivity for methanol with a FE of 25.2% at -1.3 V vs. Ag/AgCl (sat. KCl) in 0.5 M KHCO₃. Recently, Bagchi *et al.*¹⁹² reported excellent selectivity towards methanol could be achieved with an intermetallic

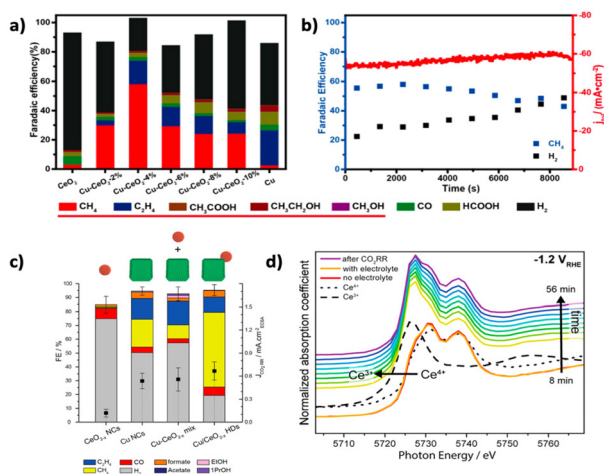


Fig. 13 (a) Comparison of product selectivity obtained with Cu doped CeO₂ catalysts, (b) long term stability of a Cu–CeO₂-4% (4% of Cu doped in CeO₂) catalyst at -1.2 V vs. RHE. Reprinted with permission.¹⁸⁵ Copyright 2018, the American Chemical Society. (c) Product selectivity comparison and CO₂RR partial current densities obtained with Cu/CeO_{2-x} HDs, Cu–CeO_{2-x} mixture, Cu nanocrystals (NCs), and CeO_{2-x} NCs, (d) Ce L_{III}-edge XANES spectra collected every eight minutes when using Cu/CeO_{2-x} HDs under eCO₂RR conditions at -1.2 V vs. RHE in 0.1 M KHCO₃. Reprinted with permission.¹⁸⁶ Copyright 2019, the American Chemical Society.

CuGa₂ electrocatalyst. An exceptional FE of 77.6% for methanol was achieved at -0.3 V vs. RHE. X-ray photoelectron spectroscopy (XPS) and *in situ* X-ray absorption fine structure (XAFS) analysis revealed the importance of surface and subsurface oxides of Ga at low potentials in enhancing methanol selectivity. However, when more negative potentials were applied, lattice expansion occurs, indicating reduction of Ga₂O₃, which resulted in the diminution of methanol selectivity.

Table 4 summarises state-of-the-art Cu alloys and other Cu based catalysts, which are selective for methane and methanol in H type and flow cells.

4.3 C₂/C₂₊ products

C–C coupling is required to produce C₂/C₂₊ products. Electrochemically reducing CO₂ to C₂ products with significant selectivity is a great challenge. Typically, copper is combined with transition metals that generate CO as their primary product during eCO₂RR to improve the selectivity for C₂ products. With Cu–M (M = guest metal) bimetallic alloy catalysts, the formation of C₂ hydrocarbons is impacted by two main effects: the spillover of CO at the boundaries and electronic effects. The geometric arrangement of the Cu and guest metal

Table 4 Examples of the state-of-the-art Cu alloys and other Cu catalysts that favour methane or methanol production

Catalyst	Product	Electrolysis cell	Electrolyte	Potential (V vs. RHE) ^a	FE (%)	j_{product} (mA cm ⁻²)	Ref.
Cu–ZnO NPs	CH ₄	H-cell	0.1 M KHCO ₃	-1.35	70	~40	184
4% Cu doped CeO ₂	CH ₄	H-cell	0.1 M KHCO ₃	-1.8	58	33.6	185
Cu ₆₈ Ag ₃₂ nanowires	CH ₄	H-cell	0.5 M KHCO ₃	-1.2	60	~27	187
Cu foil	CH ₄	H-cell	0.5 M NaHCO ₃ + methyl carbamate	-2.1	85	31	188
Single atom Cu/GYD	CH ₄	H-cell	0.1 M KHCO ₃	-1.3	66	20	189
Single atom Cu/TCNFs	CH ₃ OH	H-cell	0.1 M KHCO ₃	-0.9	44	41	190
Cu (111) nanospheres	CH ₄	Flow cell	1 M KOH	-0.91	53	53	191
CuGa ₂	CH ₃ OH	Flow cell	1 M KOH	-0.3	78	~16	192

^a Potential at which maximum FE was obtained.

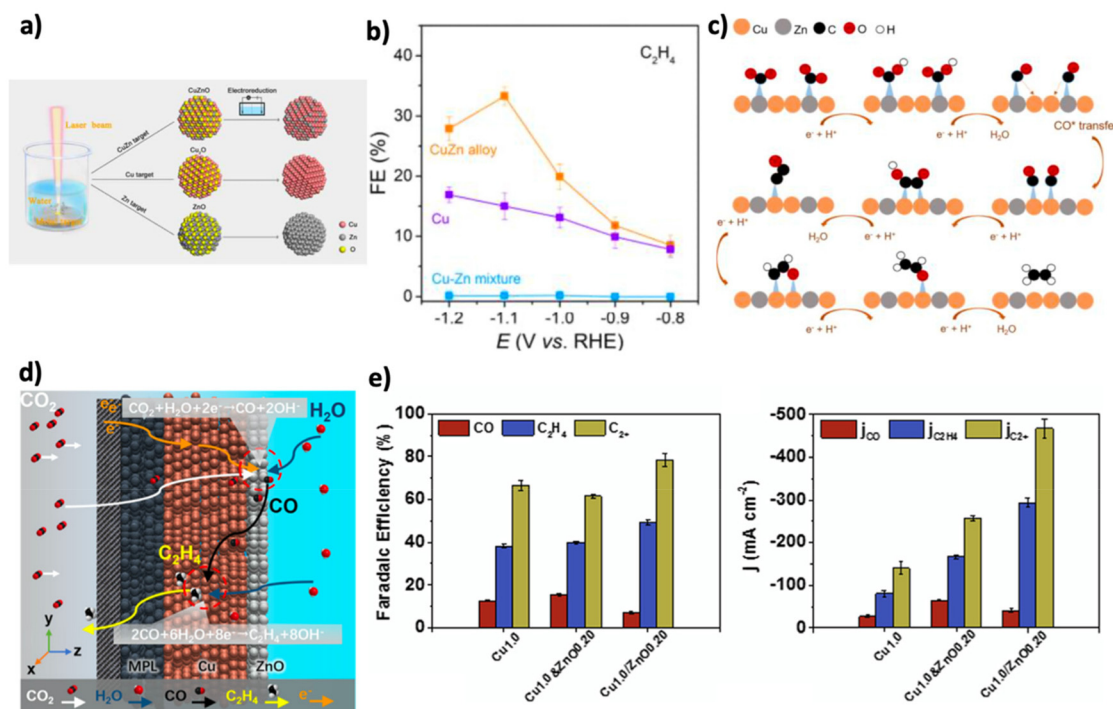


Fig. 14 (a) Schematic representation of the preparation of CuZn alloy nanoparticles, (b) faradaic efficiency of C₂H₄ obtained by eCO₂RR using CuZn alloy, Cu–Zn mixtures and Cu catalysts, (c) mechanism proposed for C₂H₄ formation with CuZn catalysts. Reprinted with permission.²⁰¹ Copyright 2018, the American Chemical Society; (d) schematic representation of use of a Cu–ZnO tandem catalyst to generate CO₂ → C₂₊, (e) faradaic efficiency and current density obtained with Cu, Cu and ZnO mixtures and Cu–ZnO tandem catalysts. Reprinted with permission.²⁰³ Copyright 2020, Elsevier Inc.

in the catalyst leads to the emergence of phase boundaries which promote C–C coupling by increasing CO coverage and stabilising reaction intermediates. The identity and availability of the guest metal near Cu affects the concentration of CO at copper. The oxidation state of Cu in the catalyst also is significant in providing a selective pathway for ethylene or ethanol production during electrolysis. In particular, stabilisation of the Cu(I) state of the catalyst during eCO₂RR has been suggested to be crucial for the generation of highly reduced carbons,^{36,196–199} but not without debate.²⁰⁰

4.3.1 Ethylene. CuZn alloy nanoparticles, prepared by Feng *et al.*,²⁰¹ have been reported to produce ethylene with considerable selectivity, as depicted in Fig. 14(a and b). In their study, the authors initially prepared CuZn nanoparticles having different Cu : Zn ratios by the pulsed laser ablation method and electrochemically reduced them to CuZn before performing bulk electrolysis experiments. They found that CuZn nanoparticles with a 4 : 1 Cu : Zn ratio generated 33.3% ethylene at -1.15 V vs. RHE in 0.1 M KHCO₃. The reason for enhanced ethylene formation is the generation of abundant *CO associated with the presence of Zn and also the homogeneous distribution of Cu and Zn in CuZn nanoparticles that aids the transfer of *CO from Zn to Cu sites followed by CO–CO coupling over the Cu sites and then finally the release of C₂H₄ (see Fig. 14(c)).

Garcia *et al.*²⁰² have synthesized CuO/ZnO particles by a water/oil microemulsion method. Bulk electrolysis experiments using this catalyst in 0.1 M KHCO₃ using a flow cell, produced ethylene with a faradaic efficiency of 91.1% at highly negative potentials (-2.5 V vs. Ag/AgCl). Recently, Zhang *et al.*²⁰³ prepared a Cu/ZnO tandem catalyst by airbrushing commercial Cu and ZnO nanoparticles over carbon paper and performed bulk electrolysis with a flow cell in 1 M KOH solution (Fig. 14(d)). With this Cu/ZnO tandem catalyst, the combination of ZnO acting as a CO generator and Cu as a C₂ intermediate stabilizer enhances C₂ and C₂₊ product formation. With an optimal loading of Cu and ZnO, Cu/ZnO tandem catalyst produces 48% ethylene at -0.7 V vs. RHE in 1 M KOH (Fig. 14(e)).

Use of uniformly distributed CuAg alloy nanowires prepared by electroplating in 3,5-diamino-1,2,4-triazole baths, significantly enhances C₂ product generation (ethylene in particular) during eCO₂RR in a flow reactor.²⁰⁴ The faradaic efficiency of ~60% for ethylene at -0.7 V vs. RHE and a total current density of 300 mA cm⁻² achieved with CuAg alloy nanowires containing 6% Ag, demonstrates superior CO₂ reduction activity as shown in Fig. 15(a). Examination of *in situ* Raman spectra shown in Fig. 15(b and c) led to the conclusion that high stability of Cu₂O layers under electrolysis conditions and the optimal availability of CO were the factors that resulted in high C₂H₄ selectivity. The

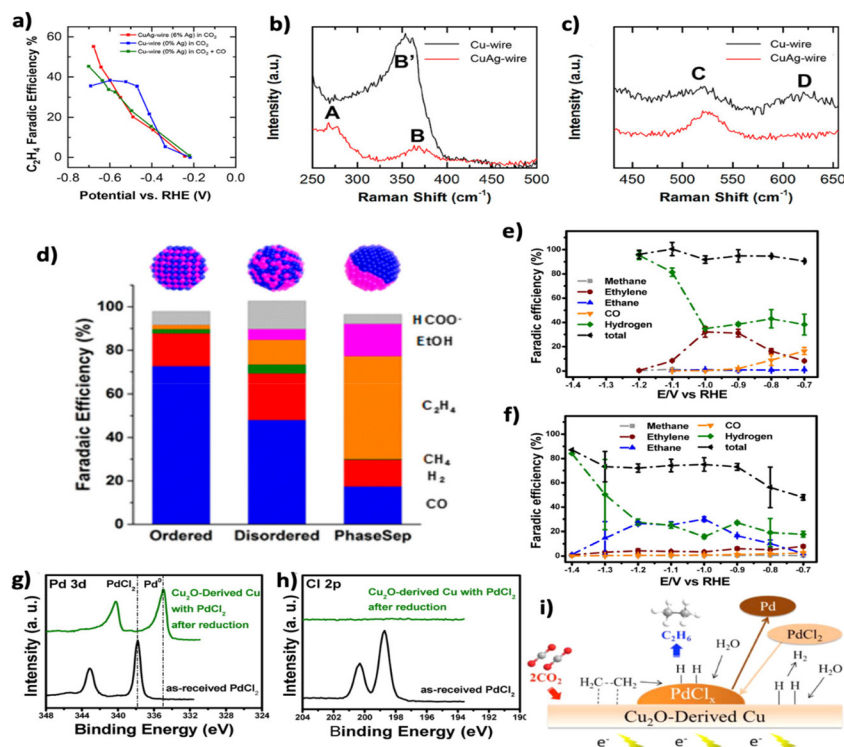


Fig. 15 (a) Faradaic efficiencies for ethylene formation obtained during electrolysis at AgCu and Cu catalysts under different experimental conditions, (b and c) *in situ* Raman spectra showing (b) the Cu–CO stretch, and (c) the Cu–O stretch regions obtained from the eCO₂RR using electrolysis AgCu and Cu catalysts at -0.7 V in 0.1 M KOH solution. Reprinted with permission.²⁰⁴ Copyright 2018, American Chemical Society. (d) Faradaic efficiencies of products obtained with three CuPd catalysts. Reprinted with permission.⁶⁴ Copyright 2017, the American Chemical Society. (e and f) eCO₂RR product distribution obtained using Cu catalysts (e) with and (f) without PdCl₂ in the electrolyte, (g and h) XPS of (g) Pd 3d and (h) Cu 2p regions of PdCl₂ and a Cu catalyst with PdCl₂ after reduction, (i) schematic illustration of ethane formation from surface adsorbed ethylene and hydrogen using a Cu catalyst in the presence of PdCl₂. Reprinted with permission.²⁰⁵ Copyright 2015, the American Chemical Society.

importance of the geometric arrangement in tuning the product selectivity at copper-based bimetallic catalysts is emphasized by eCO₂RR studies using alloys with different alloying configurations.⁶⁴ CuPd alloys with ordered, disordered, and phase-separated geometric structures have been examined for CO₂ reduction in 1 M KOH solution using a flow reactor. The phase-separated sample exhibited superior activity for C₂ product selectivity. A total FE for C₂ products of 63% was achieved at -0.8 V vs. RHE, in which the contribution from ethylene was $\sim 50\%$, as shown in Fig. 15(d). On the other hand, the selectivity for ethylene at disordered alloy sample is ~ 4 times lower than the phase separated sample. In contrast, the ordered CuPd alloy produced only C₁ products, mainly CO.

The significant role of PdCl₂ electrolyte in modifying the product selectivity achieved with Cu₂O derived Cu catalysts was identified by Yeo and co-workers.²⁰⁵ With a pure Cu₂O derived Cu electrode, a 32% yield of ethylene was achieved -1.0 V vs. RHE in an aqueous bicarbonate solution. With addition of PdCl₂ into the electrolyte, a complete shift in the major product was observed with selectivity changing from C₂H₄ to C₂H₆. Now the highest faradaic efficiency was 30% for C₂H₆ at the same potential (Fig. 15(e and f)). Analysis of XPS data after electrolysis (Fig. 15(g and h)) revealed that PdCl₂ provides the source of a sacrificial dopant by forming Pd⁰ on Cu₂O during the electrolysis reaction. According to the authors, the Pd⁰ sites adsorb hydrogen effectively and provide a hydrogen source that facilitates reduction of C₂H₄ to C₂H₆, as shown in Fig. 15(i).

Tables 5 and 6 summarises the state-of-the-art Cu alloys and Cu based catalysts, which are selective for ethylene in H-type and flow cells, respectively.

4.3.2 Ethanol. The Yeo group²⁴³ introduced oxide derived Cu_xZn catalysts for improving the selectivity for ethanol. In their study, Cu_xZn catalysts were prepared by electrodeposition of Cu₂O and ZnO layers onto a polished Cu disk from a solution containing CuSO₄ and ZnCl₂ with Cu : Zn ratios of 10, 4, 2 at a current density of -0.92 mA cm⁻² for 600 s. Detailed characterization with selected area electron diffraction (SAED) and X-ray diffraction (XRD) techniques showed that the electrodeposited Cu_xZn consisted of phase segregated Cu and Zn crystallites rather than a CuZn alloy. eCO₂RR experiments with phase segregated Cu_xZn catalysts revealed that Cu₄Zn produced 29% ethanol at -1.05 V vs. RHE in 0.1 M KHCO₃ with a partial current density of 8.2 mA cm⁻². The ethanol/ethylene ratio was found to be positively correlated with the Zn content with the tested samples. On increasing the Zn content from 0% to 30% in Cu_xZn, the ethanol/ethylene ratio increased from 0.48 to 6. As shown in Fig. 16(a), it was proposed that CO generated at the Zn site spilt over to the Cu site and inserted into the *CH₂ intermediate on the Cu surface, leading to the enhancement of ethanol generation.

The Grätzel group²⁴⁴ reported that a CuO/ZnO core/shell structure derived CuZn bimetallic alloy catalyst for eCO₂RR achieved 41% and 48% of C₂₊ product selectivity in H-cell and flow cell configurations respectively, with ethanol being the primary C₂ product. To achieve this outcome, CuO nanowires

Table 5 Examples of state-of-the-art Cu alloys and other Cu based catalysts selective for ethylene generation in an H-cell by eCO₂RR

Catalyst	Electrolysis cell	Electrolyte	Potential (V vs. RHE) ^a	FE (%)	$j_{C_2H_4}$ (mA cm ⁻²)	Ref.
OD-Cu + PdCl ₂	H-cell	0.1 M KHCO ₃	-1.0	32	~ 10	205
ZrO ₂ /Cu-Cu ₂ O	H-cell	0.1 M KCl	-1.3	62.5	15	206
4H crystal phased Au/Cu nanoribbon	H-cell	0.1 M KHCO ₃	-1.1	44.9	14.4	207
Ag/Nafion-Cu ₂ O	Single compartment cell	0.1 M NaHCO ₃	-1.9	80	16	208
Ag-Cu nanodimers	Custom made two compartment cell	0.1 M KHCO ₃	-1.1	~ 40	10	209
Anodized Cu	PEEK ^b	0.1 M KHCO ₃	-1.08	38.1	7.3	210
Plasma activated Cu	Custom made two compartment cell	0.1 M KHCO ₃	-0.9	60	>8	211
Cu ₂ O derived Cu	Teflon cell ^b	0.1 M KHCO ₃	-1.0	42.6	13.3	212
Nano-defective Cu nanosheets	H-Cell	0.1 M K ₂ SO ₄	-1.18	83.2	~ 50	213
Cu ₂ O film	Teflon cell ^b	0.1 M KHCO ₃	-0.99	37.5	12.9	214
Cu mesocrystal	PTFE cell ^b	0.1 M KHCO ₃	-0.99	27.2	7	215
B-doped Cu	H-cell	0.1 M KCl	-1.1	52	36.4	216
Plasma treated Cu	H-cell	0.1 M KHCO ₃	-1.0	45	15.3	217
<i>t</i> -Cu ₂ O NPs/C	H-cell	0.1 M KHCO ₃	-1.1	59	24	218
Cu ₂ O NPs/C	H-cell	0.1 M KHCO ₃	-1.1	57.3	12	219
Cu ₃ N nanocubes	H-cell	0.1 M KHCO ₃	-1.6	60	18	220
Pulsed-Cu	H-cell	0.1 M KHCO ₃	-1.0	48.6	20	221
CuO _x	H-cell	0.1 M KHCO ₃	-1.3	53	14	222
UiO-66-derived amorphous ZrO _x /Cu	H-cell	0.1 M KHCO ₃	-1.05	43.3	15.7	223
CuBr derived Cu nanodendrites	H-cell	0.1 M KHCO ₃	-0.85	~ 40	~ 9	224
Cu on Cu ₃ N	H-cell	0.1 M KHCO ₃	-1.05	43	~ 16.5	225
Reconstructed Cu-I	H-cell	0.1 M KHCO ₃	-1.09	59.9	~ 15.7	226
Cu _{1.8} Se nanowires	H-cell	0.1 M KHCO ₃	-1.1	55	8.3	227
A-Cu NWs	H-cell	0.1 M KHCO ₃	-1.0	69.79	~ 18	228
Cu/TiNT	H-cell	0.5 M KCl	-1.6^c	55	132	229

^a Potential at which maximum FE was obtained. ^b Custom made two compartment cell. ^c V vs. Ag/AgCl.

Table 6 Examples of state-of-the-art Cu alloys and other Cu based catalysts that favour ethylene formation in flow/MEA cells

Catalyst	Electrolysis cell configuration	Electrolyte	Potential (V vs. RHE) ^a	FE (%)	$j_{C_2H_4}$ (mA cm ⁻²)	Ref.
Cu/ZnO	Plug flow reactor	1 M KOH	-0.9	~50	~300	203
CuAg nanowires	Flow cell	1 M KOH	-0.7	60	~180	204
Phase separated CuPd NPs	Flow cell	1 M KOH	-0.8	~50	~180	64
Ce doped Cu NPs	Flow cell	1 M KOH	-0.7	53	~80	230
Atomic Ni decorated Cu	Flow cell	1 M KOH	-0.88	31.8	85	231
Cu (100) nanocubes	Flow cell	1 M KOH	-0.7	60	120	191
Cu-DAT nanowires	Flow cell	1 M KOH	-0.6	38.2	90	232
Cu NPs	Flow cell	1 M KOH	-0.58	35	150	233
Graphite/carbon NPs/Cu/PTFE	Flow cell	7 M KOH	-0.55	70	~70	234
DVL-Cu	Flow cell	1 M KCl	-0.81	84.4	92.5	235
Fluorinated-Cu	Flow cell	0.75 M KOH	-0.89	65	1040	236
CuS/Cu-V	Flow cell	1 M KOH	-0.92	21.1	84	237
Dendritic Cu	Flow cell	0.1 M KHCO ₃	-1.2	36	162	238
Nanoporous Cu	Flow cell	1 M KOH	-0.67	38.6	252	239
PTFE-Cu NPS	MEA	0.1 M KHCO ₃	3.8 ^b	56.7	85	240
Molecular tuned Cu	MEA	1 M KHCO ₃	3.65 ^b	72	230	241
Graphite/carbonNP/Cu/PTFE	MEA	1 M KHCO ₃	4.2 ^b	46	92	242

^a Potential at which maximum FE was obtained. ^b Cell voltage.

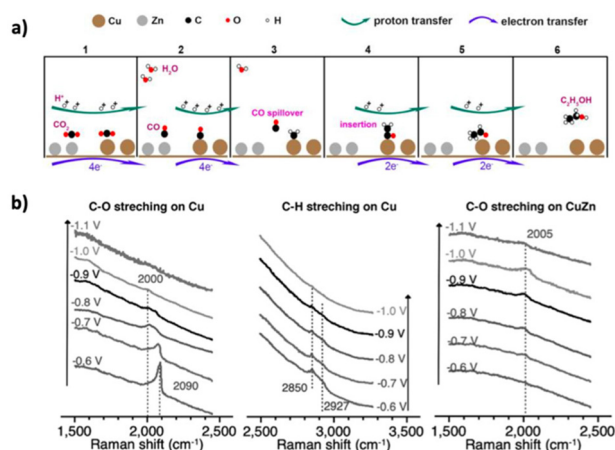


Fig. 16 (a) Mechanism proposed for ethanol formation using CuZn catalysts. Reprinted with permission.²⁴³ Copyright 2016, the American Chemical Society. (b) Operando Raman spectra obtained during electrolysis of saturated CO₂ in 0.1 M KHCO₃ solution using Cu and CuZn catalysts. Reprinted with permission.²⁴⁴ Copyright 2019, Wiley-VCH Verlag GmbH & Co. KGaA, Weinheim.

were coated with ZnO (~30 nm thick) by atomic layer deposition. Subsequently, electrochemical reduction was undertaken to form a CuZn bimetallic layer (~90 nm) over the Cu nanorods. XRD analysis after electroreduction showed the presence of Cu and Zn/ZnO but not CuZn alloy. This CuZn catalyst was used to catalyze eCO₂RR in 0.1 M KHCO₃ using an H-cell. Ethanol was produced at -1.15 V vs. RHE with a FE of 32% and a partial current density of 10.5 mA cm⁻². The mechanism proposed in this study on the basis of analysis of Raman spectra (Fig. 16(b)) is similar to that suggested by Ren *et al.*²⁴³ for the formation of ethanol using a phase segregated Cu_xZn catalyst.

An eCO₂RR study by Lee *et al.*⁶³ also reveals the importance of Cu-Ag biphasic boundaries in achieving high ethanol

selectivity. The structure of this catalyst allows the CO concentration near the Cu-Ag phase blended catalyst surface to be increased, facilitating the insertion of CO into the Cu bound intermediates (*CH₂) to form *COCH₂. Upon further transfer of protons and electrons, *COCH₂ is hydrogenated, resulting in the formation of ethanol *via* the acetaldehyde route. A recent eCO₂RR study using CuO_x coated Ag nanowires revealed an interesting relationship between the thickness of the CuO_x layer and ethanol selectivity.²⁴⁵ While the Ag/CuO_x-y (core/shell-y, where y represents the average thickness of the shell in nm) samples with lower CuO_x thickness (Ag/CuO_x-10) showed substantial ethanol formation (Fig. 17(c)), ones with thicker CuO_x favoured ethylene. *In situ* XAS studies (Fig. 17(d)) revealed that Ag/CuO_x-10 initially contained Cu₂O which was stable in the potential range of -0.5 to -0.7 V vs. RHE (Cu-O(Cu₂O) and Cu-Cu(Cu₂O) bands preserved) but reduced to Cu⁰ at potentials more negative than -0.7 V resulting in formation of a Cu-Cu(Cu) bond (Cu-Cu band of metallic Cu). However, CuAg alloy was not formed during this reduction process. With Ag/CuO_x-32, a less negative potential is sufficient to trigger this structural change. At a potential of -0.6 V vs. RHE, reduction of Cu₂O to metallic Cu⁰ was observed. Further, in both cases this led to a stable atomic arrangement in the potential range of -0.7 to -1.3 V vs. RHE resulting in enhanced C₂ product selectivity. The authors found that the existence of a majority of Cu in the +1 state facilitated the CO insertion mechanism and favoured ethanol production at Ag/CuO_x-10 while the presence of more Cu⁰ in Ag/CuO_x-32 favoured C-C coupling mechanism which led to the formation of ethylene. This result signifies that the transformation of Cu (I) to Cu(0) is controlled by the CuO_x thickness, which in turn influences the eCO₂RR product selectivity.

Tables 7 and 8 summarise the state-of-the-art Cu alloys and other Cu based catalysts, which are selective for ethanol in H type and flow/membrane electrode assembly (MEA) cells, respectively.

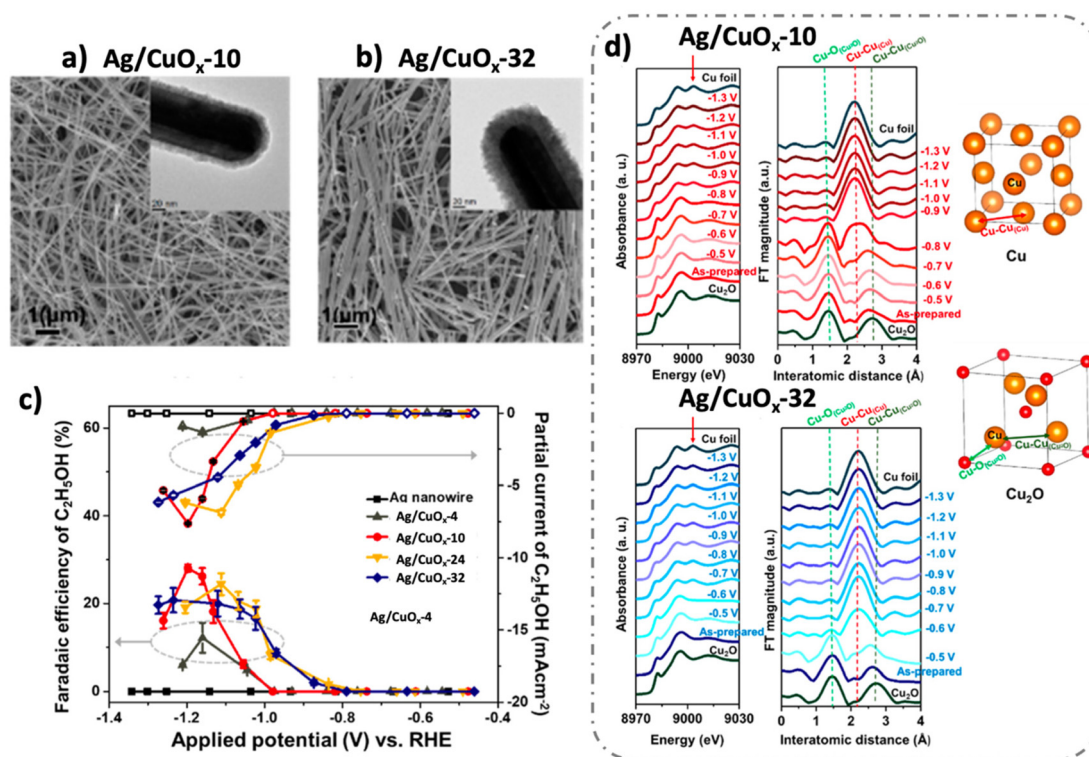


Fig. 17 (a and b) SEM images of Ag/CuO_{x-y} (core/shell-y, where y represents the average thickness of the shell in nm) catalysts, (c) ethanol faradaic efficiencies and partial current densities obtained with Ag/CuO_{x-y} catalysts, (d) *in situ* XANES and EXAFS data obtained with Ag/CuO_{x-10} and Ag/CuO_{x-32} catalysts. Reprinted with permission.²⁴⁵ Copyright 2019, the American Chemical Society. Source: <https://pubs.acs.org/doi/10.1021/acscentsci.9b01142>. Further permissions related to the material excerpted should be directed to the ACS.

Table 7 Examples of state-of the art Cu alloy catalysts that favour ethanol production by eCO₂RR in an H-cell

Catalyst	Electrolysis cell	Electrolyte	Potential (V vs. RHE) ^a	FE (%)	<i>j</i> _{C₂H₅OH} (mA cm ⁻²)	Ref.
Cu ₄ Zn	H-cell	0.1 M KHCO ₃	-1.05	29	8.2	243
CuO/ZnO	PEEK ^b	0.1 M KHCO ₃	-1.15	32	10	244
Ag/CuO _{x-10}	H-cell	0.1 M KHCO ₃	-1.2	29	7.6	245
Cu ₃ Sn	H-cell	0.1 M KHCO ₃	-1.0	64	5.7	246
Ag ₂₀ Cu ₁	H-cell	0.1 M KHCO ₃	-1.1	16.5	4.1	247
Au ₁ Cu ₃	H-cell	0.5 M KHCO ₃	-1.0	29	5.6	248
Cu ₅ Zn ₈	H-cell	0.1 M KHCO ₃	-0.8	46.6	2.3	249
Dendritic Cu-Cu ₂ O	H-cell	0.1 M KCl	-0.4	26	2.99	250
CuO NPs	H-cell	0.2 M KI	-1.7 ^c	36.1	N.R	251
Cu/carbon nanospire	H-cell	0.1 M KOH	-1.2	63	~3	252
Cu-I	H-cell	0.1 M KHCO ₃	-0.9	25	11	253
Cu-GNC-VL	H-cell	0.5 M KHCO ₃	-0.9	70.5	9	254
Cu ₂ /N _{0.14} C	H-cell	0.1 M KHCO ₃	-1.1	51	14.4	255
UiO-66-derived amorphous ZrO _x /Cu	H-cell	0.1 M KHCO ₃	-1.1	22.4	12.7	223
Cu on Cu ₃ N	H-cell	0.1 M KHCO ₃	-0.95	18.4	~8	225
Cu _{1.8} Se nanowires	H-cell	0.1 M KHCO ₃	-1.1	24	3.5	227

N.R – not reported. ^a Potential at which maximum FE was obtained. ^b Custom made two compartment cell. ^c Potential (V) vs. SCE.

4.3.3 Other C₂/C₂₊ products. Apart from ethanol, acetate, propanol, and acetaldehyde are the other major oxygenates derived from CO₂ using the eCO₂RR with copper containing electrocatalytic materials. However, despite significant efforts to enhance the yields of these other highly reduced products, faradaic efficiencies reported to date are still very low.

Acetate formation with reasonable selectivity (faradaic efficiency of 21% at -1.3 V vs. RHE), at 0 °C was achieved by the Meyer group by employing ultra-small (Cu)_m(Ag)_n (*m*, *n* denote the atomic ratios of Cu and Ag) bimetallic nanoparticles immobilised on a polymer. In 0.5 M KHCO₃ electrolyte containing 8 ppm benzotriazole at 0 °C, significant

Table 8 Examples of state-of the art Cu alloy catalysts that favour ethanol production by eCO₂RR in flow/MEA cells

Catalyst	Electrolysis cell	Electrolyte	Potential (V vs. RHE) ^a	FE (%)	$j_{C_2H_5OH}$ (mA cm ⁻²)	Ref.
CuO/ZnO	Flow cell	1 M KOH	N.R	41	82	244
Phase separated CuPd NPs	Flow cell	1 M KOH	-0.8	~50	~180	64
Cu ₂ Zn ₁	Flow cell	1 M KOH	-1.0	26.2	~60	256
CuAg nanowires	Flow cell	1 M KOH	-0.7	25	75	204
Binding-site diverse Ag/Cu	Flow cell	1 M KOH	-0.67	41	102	231
Ag decorated Cu/Cu ₂ O	Flow cell	1 M KOH	N.R	19.2	304.5	257
Atomic Ni decorated Cu	Flow cell	1 M KOH	-0.88	22.9	61.4	258
N-C/Cu	Flow cell	1 M KOH	-0.68	52	156	259
Cu ₂ O spheres	Flow cell	2 M KOH	-0.61	26.9	71.8	250
Cu-DAT nanowires	Flow cell	1 M KOH	-0.69	27.3	~35	232
Fluorinated-Cu	Flow cell	2.5 M KOH	-0.54	16	128	236
CuS/Cu-V	Flow cell	1 M KOH	-0.92	24.7	99	237
Nanoporous Cu	Flow cell	1 M KOH	-0.67	16.6	~108	239
Cu- _{DS}	Flow cell	1 M KOH	-0.95	52	52	260
Cu ₂ O NPs	Flow cell	2 M KOH	-0.6	~27	93	198
FeTPP[Cl]/Cu	Flow cell	1 M KHCO ₃	-0.82	41	124	261
FeTPP[Cl]/Cu	MEA	0.1 M KHCO ₃	3.7 ^b	45	100	261
Cu ₃ Sn	MEA	1 M KOH	3 ^b	40	361	246
PTFE-Cu NPS	MEA	0.1 M KHCO ₃	4.2 ^b	17.1	~40	240
Graphite/carbonNP/Cu/PTFE	MEA	0.1 M KHCO ₃	4.2 ^b	~15	~30	242
Cu- _{DS}	MEA	0.1 M KHCO ₃	3.5 ^b	50	95	260

N.R – not reported. ^a Potential at which maximum FE was obtained. ^b Cell voltage.

enhancement of acetate with a FE of 21.2% at -1.3 V vs. RHE was achieved on (Cu)_m, (Ag)_n compared to pure Cu where less than 1% of acetate was produced. The CO spillover mechanism from Ag to Cu was again proposed. As for ethanol production described above, CO inserted onto the Cu surface and then coupled with *CH₂ to give the *COCH₂ intermediate, but which in this case was further reduced to acetate *via* the acetaldehyde route. Since the availability of CO at copper is the limiting factor for C-C coupling, Lum *et al.*²⁶² have introduced sequential Cu-Au and Cu-Ag catalysts. These were microfabricated electrodes where Cu and Au or Ag were organized in a sequential order with predefined thickness of each metal and the distance between the two metals as shown in Fig. 18(a-c). In these catalysts, Cu is in close proximity with Au or Ag and favours formation of CO during eCO₂RR. The generated CO spills over to the Cu surface and facilitates production of more highly reduced carbon fuels. By this strategy, these authors have improved the formation of oxygenate products to 41.3% which surpasses the selectivity of 26.1% achieved for hydrocarbons. The microfabricated Cu-Au or Cu-Ag electrodes used in these studies consisted of integrated but independent Cu and Au/Ag electrodes arranged adjacent to each other (Fig. 18(a-c)). As a result, the CO generated during eCO₂RR was accumulated near adjacent Cu electrodes. By varying the distance between the Cu and Ag probes, the oxygenate to ethylene ratio can be varied from 0.5 to 2.4, with the highest faradaic efficiency of oxygenates reaching 41.4%, with total C₂/C₂₊ products being 65% (Fig. 18(d)). This study paves the way for new catalyst design strategies with controlled bimetallic arrangements for selective generation of C₂/C₂₊ products other than ethylene. Morales-Guio *et al.*²⁶³ developed an Au/Cu tandem catalyst by depositing Au nanoparticles on Cu foil. This catalyst design provided more than a 100-fold increase in the rate of

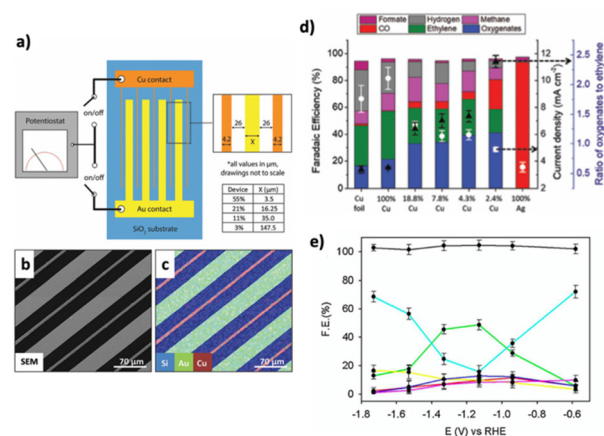


Fig. 18 (a) Schematic diagram of a custom-made electrochemical device with microfabricated Au and Cu electrodes, (b) SEM and (c) EDX mapping of AuCu electrodes with 11% of geometric area occupied by Cu, (d) faradaic efficiencies and current densities obtained with AgCu electrodes. Reprinted with permission.²⁶² Copyright 2018, the Royal Society of Chemistry. (e) Faradaic efficiencies of products obtained with a Mo₉/Cu heterostructure catalyst (cyan: hydrogen, red: ethanol, green: acetate, yellow: methane, blue: ethylene, purple: ethane). Reprinted with permission.²⁶⁴ Copyright 2021, Elsevier B.V.

CO₂ reduction to the higher reduced carbon products. The improvement in C₂/C₂₊ alcohol production at low overpotentials was attributed to the enhanced local CO concentration provided by Au present in close proximity to copper.

Recently, the Wu group²⁰³ have reported that enhancement in oxygenate production also can be achieved with a Cu/ZnO tandem catalyst. Tandem catalysts with variable ZnO loading were prepared by air brushing ZnO onto a Cu electrode pre-synthesised on a GDL electrode and dried under vacuum. By

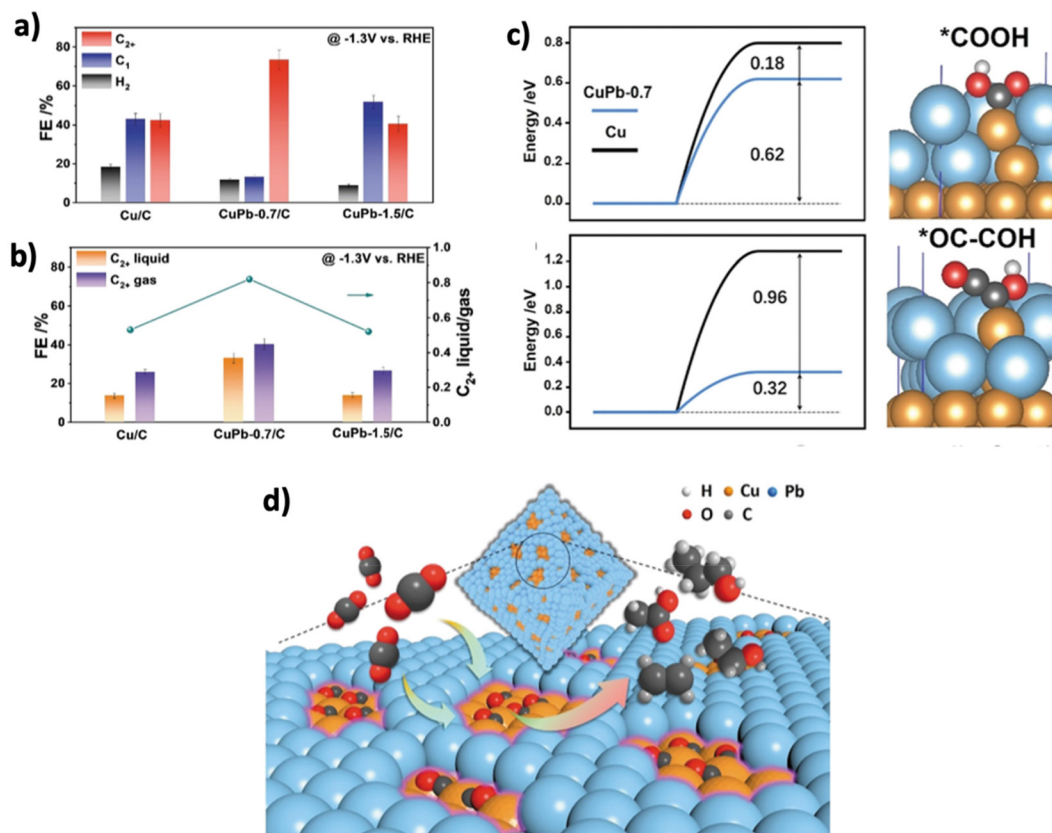


Fig. 19 (a and b) Faradaic efficiencies of C_1 and C_2 products obtained using a CuPb-0.7/C (where 0.7 represents the thickness of Pb shell) catalyst, (c) simulated data and (d) schematic illustration of the production of C_{2+} products via the eCO₂RR route. Reprinted with permission.²⁶⁵ Copyright 2021, the American Chemical Society.

controlling the ZnO loading, effective spatial management of CO transport was achieved to optimise CO utilisation. High selectivity for C_{2+} products with a faradaic efficiency reaching 80% is achieved as shown in Fig. 14(e) with a partial current density of $\sim 470 \text{ mA cm}^{-2}$ obtained with Cu_{1.0}Zn_{0.2}, values which respectively are 1.2 and 3.4 times higher than obtained with pure Cu.

An excellent conversion yield for CO₂ to acetate was reported recently at polyoxometalate modified Cu cubes. The Mo₈/Cu

heterostructure catalyst was rich in Cu–O–Mo interfaces, which led to the generation of acetate with a FE of 48.8% at -1.13 V (Fig. 18(e)) and a partial current density of 68.9 mA cm^{-2} in saturated NaHCO₃.²⁶⁴ The selectivity for C_{2+} products was found to be dependent on the thickness of the shell in Cu/Pb core/shell nanocubes. Cu/Pb nanocubes with an optimal shell thickness of 0.7 nm gave 73.5% C_{2+} products with a partial current density of 294.4 mA cm^{-2} at -1.3 V in 1 M KOH solution under flow cell conditions (Fig. 19(a and b)).²⁶⁵ DFT calculations

Table 9 Examples of the state-of-the-art Cu alloys and other Cu based catalysts for other C_2/C_{2+} product formation with eCO₂RR

Catalyst	Product	Electrolysis cell	Electrolyte	Potential (V vs. RHE) ^a	FE (%)	j_{product} (mA cm ⁻²)	Ref.
(Cu) _m , (Ag) _n NPs	CH ₃ COO ⁻	H-cell	0.5 M KHCO ₃ + 8 ppm benzotriazole	-1.33	21	N.R.	267
Dendritic Cu–Cu ₂ O	CH ₃ COO ⁻	H-cell	0.1 M KCl	-0.4	40	4.6	250
4% Cu dots/Ag	Oxygenates	PEEK ^b	0.1 M CsHCO ₃	-1.0	41	4.4	262
Mo ₈ /Cu	CH ₃ COO ⁻	Electrochemical batch cell	Saturated NaHCO ₃	-1.13	49	57	264
CuBi	C ₃ H ₈	Filter press cell	0.45 M KHCO ₃ + 0.5 M KCl.	N.R.	85	38	266
Single atom Cu/NPC	CH ₃ COCH ₃	H-cell	0.1 M KHCO ₃	-0.36	36.7	N.R.	268
CuI derived Cu nanofibres	C ₂ H ₆	H-cell	0.1 M KHCO ₃	-0.73	~ 30	~ 9	224
CuS _x -DSV	<i>n</i> -C ₃ H ₇ OH	H-cell	0.1 M KHCO ₃	-1.05	15.4	~ 10	269
Graphene/ZnO/Cu ₂ O	<i>n</i> -C ₃ H ₇ OH	H-cell	0.5 M NaHCO ₃	-0.9 ^c	30	N.R.	270
MOF derived Cu	<i>n</i> -C ₃ H ₇ OH	Five port electrochemical cell	0.1 M KHCO ₃	-2.4 ^c	~ 16	~ 4	271

N.R. – not reported. ^a Potential at which maximum FE was obtained. ^b Custom made two compartment cell. ^c V vs. Ag/AgCl.

revealed that the synergistic effects due to the core/shell structure decreased the formation energies of *COOH and *OCCOH intermediates, thereby facilitating production of highly reduced carbon products as shown in Fig. 19(c and d). Azenha *et al.*²⁶⁶ reported the production of propane with a remarkable FE of 84.6% by employing Bi coated CuO nanowires on a filter press cell in 1 M KOH. Bidentate carbonate formation on the catalyst surface was an important requirement for the formation of propane. Detailed investigations revealed that the presence of Cu(I) sites and oxygen defects influenced the binding affinity of CO₂ to the Bi/Cu NW catalyst in a manner which led to high selectivity for propane.

Table 9 summarises the state-of-the-art Cu alloys and other Cu based catalysts, which are selective for other C₂/C₂₊ products.

5. Conclusions and future directions

The development of rationally designed Cu based alloy catalysts is helping to address the problem of low product selectivity achieved by eCO₂RR with pristine Cu. The reaction pathways can be manipulated by geometric and electronic effects that arise upon alloying a material. The choice of the guest elements alloyed with Cu and the spatial distribution of the constituent elements are crucial in determining the selectivity of eCO₂RR products. In this review, the different strategies reported for alloying and their influence on the properties of the catalyst are surveyed. Achievements based on state-of-the-art Cu based alloy catalysts that improve the selectivity of the targeted product are highlighted.

Commonly, the selectivity of the product is related to the binding energies of key reaction intermediates which can be tuned by varying the structure of the alloy. For example, in alloys with ordered atomic structures, the elemental composition influences the bond length between the elements. This, in turn, influences the binding energies of eCO₂RR intermediates adsorbed onto the catalyst surface, providing a systematic route to tuning product selectivity. On the other hand, a disordered alloy system can lead to changes in electronic and geometric structures during electrolysis which also influences the products that are formed. In the case of a core/shell type alloy structure, it is the strain and ligand effects, atomic vacancies and doping that affect the binding strength of reaction intermediates, again providing systematic routes to selectively tuning the product formation pathway. Phase separated alloys enhance the product selectivity in a unique way. In this class of alloy there is a clear separation in the phases of the different elements in the alloy, allowing each phase to individually generate the product that is characteristic of the element in the particular phase. Migration of products or the coupling of intermediates at the boundary of different phases is possible. This helps in producing highly reduced carbon products such as ethylene or ethanol with enhanced selectivity generally facilitated by the CO spillover mechanism.

Strategic alloying can introduce specific synergistic effects that are useful to tune the selectivity towards each product. However, other inevitably coupled effects could diminish the product selectivity. Hence, despite the ability to manipulate the reaction pathways by altering the binding energies, complete selectivity for one product remains exceedingly difficult to achieve when an alloying strategy is used alone. Therefore, integration with other catalyst design strategies is highly recommended.

Alloy structures also tend to be unstable under electrolysis conditions. The lifetime for most of the catalysts used to date is typically within a day, which is far less than that required for commercial applications. To address the problem of stability, high entropy alloys are now receiving considerable attention. These are a new class of alloy formed with multiple elements. HEAs are highly stable under extreme conditions, such as high temperatures and prolonged electrolysis conditions due to an exceptionally high kinetic solid diffusion barrier. They are active for eCO₂RR, and in principle the electronic properties and hence selectivity can be precisely tuned as they possess exceptional flexibility for adjusting the binding energy compared to primitive alloys. HEAs consist of a minimum of five different elements, resulting in a wide range of compositions and atomic arrangements. Hence, powerful machine learning tools^{272,273} are needed to identify the optimal composition and structure required to achieve selective generation of a desired eCO₂RR product. Even though the suitable composition and structure of the HEAs are predicted theoretically, the currently used synthesis strategies for HEAs require high temperatures, pressures and an inert environment. Precisely controlling the microstructures or the local atomic arrangement also is highly complicated.^{274,275}

The synergistic, so called “cocktail effect” displayed by HEAs with their multiple elemental combinations exhibit exceptional binding energies for reaction intermediates. The availability of different binding sites in these catalysts also can provide the possibility of wide range of binding energies for eCO₂RR intermediates and thus provide the possibility to tune the selectivity of the products.²⁷⁴ This property can be advantageous for cascade reactions, where the product formed at one binding site is desorbed from the surface and re-bound at the other for further reaction. Hence the choice of elements and atomic arrangement in the catalyst should be given priority when designing these catalysts. Copper based high entropy alloys are a focus as they provide the possibility of producing highly reduced hydrocarbon fuels with improved selectivity by means of cascade reactions. Support from theoretical studies is critical to extricate the full potential of such alloys.^{66,72,276} Despite these highly attractive properties, application of HEAs in the eCO₂RR is still limited.^{274,275} Nellaippan *et al.*²⁷⁷ achieved the first experimental realization of the application of nanocrystalline HEAs for eCO₂RR. Nanocrystalline AuAgPtPdCu produces 100% of gaseous products including CO, CH₄ and C₂H₄ at a low potential of -0.3 V vs. RHE with high hydrocarbon (CH₄ + C₂H₄) selectivity (~70%) in aqueous 0.5 M K₂SO₄. First principle based DFT calculations attribute this outcome to the destabilisation of

*OCH₃ and highly enhanced stabilisation of *O intermediates. Pedersen *et al.*⁶⁶ investigated the possibility of the formation of multicarbon products at HEAs by combining DFT with machine learning with (111) facets of CoCuGaNiZn and AgAuCuPdPt as model systems. The authors established a correlation between composition and the variation in H and CO binding energies.

In addition to achieving a more rational design of catalyst material, understanding the mechanism of eCO₂RR at the atomic level should make it possible to establish more precise structure–property relationships. From this perspective, advanced *in situ* characterisation techniques are receiving increased attention.^{278–280} Several studies have shown that mixed oxidation states of Cu stabilise reaction intermediates that produce highly reduced C₂ fuels.^{43,65,198,210,281} However, recent *in situ* studies through a combination of synchrotron based grazing incidence X-ray absorption and X-ray diffraction spectroscopy have suggested that Cu is present in the metallic 0 oxidation state during electrolysis.²⁰⁰ Such contradictory results may be due to structural differences in the catalytic materials. Clearly, morphological and structural changes in the catalyst during electrolysis need to be monitored. Development of more sensitive *in situ* spectroscopic and microscopic structural characterisation tools is needed.

Finally, the electrolysis system design must be optimised to enhance mass transport rates for both reactants and products. Recent research has demonstrated that use of gas diffusion electrodes and membrane electrode assemblies allow current densities to be increased drastically compared to those achievable with conventional H-shaped electrolysis cells.^{234,282,283} Improved cell designs and reaction conditions that allow much higher current densities, lower resistance and higher stability are essential to reach commercially viable efficiencies. So far, high purity CO₂ gas has been commonly used in eCO₂RR studies. To fully realize the potential of the eCO₂RR for commercial application, the effect of impurities, such as SO_x, NO_x, which are commonly present in the industrial flue gases, should be investigated. However, research towards this direction is still limited.²⁸⁴

In summary, eCO₂RR is a viable method for utilising CO₂. However, even though eCO₂RR can produce a variety of commercially valuable fuels, manufacturing of targeted products with high selectivity remains difficult. Using catalysts derived from alloying with Cu addresses some of the limitations since it allows considerable control over product selectivity by changing the guest element and the alloy structure. Despite many significant achievements emerging from the development of efficient Cu based alloy catalysts, fabrication of catalysts that have high stability and can generate significantly reduced carbon products in commercially viable quantities at competitive prices remains highly challenging.

Conflicts of interest

There are no conflicts to declare.

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