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Metal-free transfer hydrochlorination of internal C–C triple bonds with a bicyclo[3.1.0]hexane-based surrogate releasing two molecules of hydrogen chloride†

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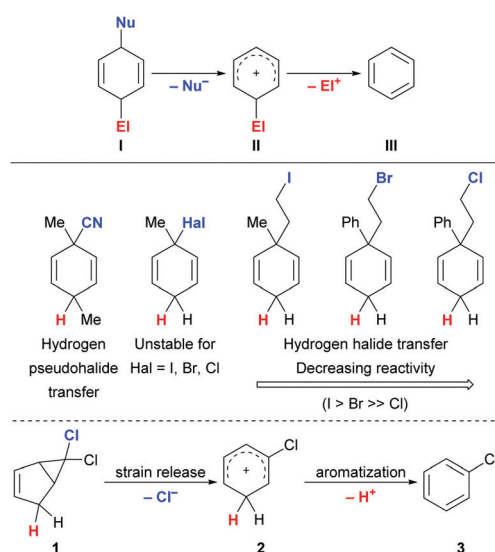
The development and application of a transfer hydrochlorination reagent based on a trichlorinated bicyclo[3.1.0]hexane core that transfers two molecules of HCl per molecule of surrogate to a π -basic substrate under $B(C_6F_5)_3$ catalysis is reported. Lewis acid-assisted chloride abstraction followed by thermal electrocyclic cyclopropyl-to-allyl cation ring opening releases ring strain as a previously unexploited driving force.

Alkenyl chlorides are important building blocks in organic chemistry both on industrial and laboratory scale. For their synthesis, the addition of HCl across C–C triple bonds is arguably the most atom-economic way of forming the C(sp²)–Cl bond. However, this direct approach requires the use of corrosive and highly toxic HCl gas. Moreover, in the industrial benchmark process of hydrochlorination of acetylene to generate vinyl chloride, less appealing mercury-based catalysts are the state of the art although newer protocols using expensive palladium-¹ or gold-based² catalysts have been developed. In more complex molecules, uncatalyzed, direct hydrochlorination with HCl is usually limited to electron-rich terminal alkynes,³ while ruthenium-⁴ palladium-⁵ platinum-⁶ and gold-based⁷ catalysts have been shown to also convert a wider range of substrates including electron-deficient alkynes.^{8,9} Utilizing readily activated surrogates of commonly employed reagents can both solve handling problems and allow for the use of cheaper and less toxic catalysts albeit at the expense of reduced atom efficiency. Surrogates that contain more than one equivalent of the compound to transfer could be even more attractive but are currently not established.

A broadly applicable surrogate with reduced acidity compared to HCl was introduced by Hammond and Xu.¹⁰ Although a gold catalyst was typically required, a handful of more activated substrates reacted under metal-free conditions. Based

on their research interest in transition-metal-catalyzed shuttle catalysis,¹¹ Morandi and coworkers established non-acidic hydrochlorination surrogates.¹² In a first protocol, 4-chlorobutan-2-one was used as surrogate, yielding highly reactive methyl vinyl ketone as stoichiometric byproduct. While iridium catalysis was required, high functional-group tolerance was observed. Under more forcing conditions, the second protocol enabled the use of cheap *tert*-butyl chloride and liberated less reactive isobutene yet at the cost of a limited substrate scope.

A general strategy for ionic transfer hydrofunctionalizations employed by our group relies on the use of strong Brønsted or Lewis acids such as $B(C_5F_6)_3$ to activate otherwise bench-stable surrogates.¹³ Typically, these processes employ pro-aromatic reagents **I** passing through a reactive Wheland intermediate **II** before affording an aromatic byproduct **III** (Scheme 1, top).^{13,14} In this vein, we have successfully developed protocols for



Scheme 1 Established hydrofunctionalization surrogates and development of a new surrogate for metal-free transfer hydrochlorination.

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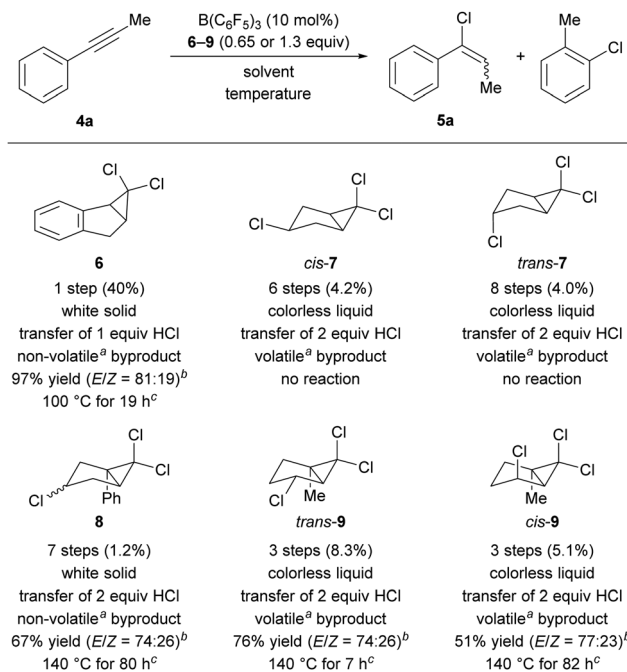
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transfer hydro(pseudo)halogenation such as transfer hydrocyanation,¹⁵ -iodination¹⁶ and -bromination¹⁷ using surrogates that show no Brønsted acidity and do not possess protic hydrogen atoms (Scheme 1, middle). While for transfer hydrocyanation a simple 3-substituted cyclohexa-1,4-diene scaffold could be utilized, this strategy had to be refined for the transfer of HI and HBr. In these cases, the corresponding 3-halocyclohexa-1,4-dienes were not suitable as surrogates due to their chemical instability at room temperature. The solution that enabled the transfer of HI was the installation of an ethylene tether to stabilize the molecule that would eventually be cleaved to generate ethylene gas and the aforementioned Wheland intermediate. When adapting this concept to transfer hydrobromination, we realized that activation of the surrogate was the easier the weaker the C(sp³)-X bond with X = I and Br as well as Cl. Overcoming this difficulty required higher temperatures and better stabilization of the intermediate carbenium ion with an additional phenyl group in the case of hydrobromination. These findings indicated that a transfer hydrohalogenation becomes increasingly complicated for the lighter halogen homologues and that a more sophisticated surrogate design might be necessary.

We commenced our quest for an HCl surrogate based on our experiences outlined above. Simple cyclohexa-1,4-diene surrogates were deemed as too unstable and, after preliminary experiments, we also learned that cleavage of the relatively strong C(sp³)-Cl bond in ethylene-tethered surrogates was impossible at synthetically useful temperatures.¹⁸ We then devised the platform **1** for releasing the cationic key intermediate based on the thermal, disrotatory electrocyclic ring opening of bicyclo[3.1.0]hexanes (Scheme 1, bottom).^{19–21} After chloride abstraction, we intended to harness release of ring strain as a driving force for the generation of the desired Wheland complex **2**. Subsequent aromatization to byproduct **3** would transfer the proton to the substrate. The parent system **1** is, however, described in the literature as thermally labile.¹⁹ For this reason, we considered and prepared the related surrogates **6–9** (Scheme 2, bottom) as potential alternatives.

Benzannulated surrogate **6** was obtained in one step from indene as a white solid in 40% yield by dichlorocarbene addition.²² When employed in a test reaction with 1-phenylprop-1-yne (**4a**) at 100 °C in C₆D₆ with 10 mol% of B(C₆F₅)₃ as catalyst, we found the formation of the desired vinyl chloride **5a** in 97% NMR yield after 19 h (Scheme 2, top). Encouraged by this initial result, we envisioned that the double bond present in **1** could also be generated during the transfer reaction by another elimination of HCl. A potential surrogate would thus carry two molecules of HCl, thereby allowing for the transfer of both. Surrogate *cis*-**7** bearing an additional chloro-substituent on the cyclopentane ring was prepared and tested in the model reaction.²³ Surprisingly, no product formation could be detected even after prolonged reaction time. The independently tested diastereomer *trans*-**7** was equally unreactive, indicating that it is not steric hindrance imposed by that chlorine atom that inhibits the reaction. We hypothesized that the activation of the surrogate is facilitated by better



Scheme 2 Developed surrogates and their properties. ^a Volatile was arbitrarily defined as having a boiling point below 160 °C. ^b Yield and d.r. obtained in the model reaction were determined by ¹H NMR spectroscopy with dimethyl terephthalate as an internal standard. ^c Reactions at 100 °C and 140 °C were run in C₆D₆ and C₆D₅Cl, respectively.

stabilization of the carbenium ion intermediate. We therefore prepared the phenyl-substituted surrogate **8**, available in seven steps from furfuryl alcohol in 1.2% overall yield. Indeed, this surrogate successfully transferred two molecules of HCl. The reaction time was significantly longer than that seen with surrogate **6**. We attribute this to differences in orbital overlap of the aromatic system with the carbenium ion. On the one hand, the rigid geometry of **6** enforces a planar benzyl cation. Orbitals are then perfectly aligned for carbenium ion stabilization by the phenyl ring and the adjacent double bond. On the other hand, the phenyl ring in the cation derived from **8** can twist out of plane to avoid steric hindrance. This renders a system that rather resembles a mere allyl cation. In contrast, previous transfer-hydrobromination surrogates benefitted from installation of a phenyl group.¹⁷ As a general disadvantage of the productive surrogates **6** and **8**, the byproducts 2-chloronaphthalene and -biphenyl are difficult to remove, and the latter is assumed to be highly toxic.

To provide carbenium ion stabilization without introducing a “heavy” phenyl group, we considered two effects: (1) a methyl group at the bridgehead position should lend further stabilization and (2) an adequately positioned chlorine atom could allow for the formation of a chloronium ion.²⁴ We found that we could test the influence of both effects with the trichloroalkane *trans*-**9**. The requisite *cis*-configured alcohol for its synthesis was described in the literature as a rare example where a hydroxy function acts as a directing group for dichlorocarbene addition.²⁵ Whereas any of the previous syntheses did



not tolerate unprotected alcohols, a straightforward three-step synthesis was possible in this case (see the ESI† for details). Having established the synthetic access, we investigated its reactivity in our model transfer hydrochlorination. Importantly, even 0.65 equivalents of the surrogate allowed for a yield of 76% of **5a** after 7 h at 140 °C. This confirmed the planned transfer of two molecules of HCl from one molecule of the surrogate. The lower yield compared to that obtained with surrogate **6** is counterbalanced by the unprecedented two-fold transfer and the generation of a harmless, volatile byproduct. As before, we were curious about the influence of the surrogate's relative stereochemistry and prepared *cis*-**9**. In the model reaction, we observed substantially lower reaction rate, and the product was obtained in lower yield (51% vs. 76%). This could be due to increased steric hindrance around the chlorine atoms in *cis*-**9** compared to the situation in *trans*-**9**. For optimal transfer efficiency, diastereopure *trans*-**9** was therefore used in all following experiments.

We next turned to the optimization of the reaction parameters. The need for an unreactive, high-boiling, and polar solvent left little room for variation. With the surrogate and the reaction temperature largely set, we thus focused on testing different catalysts in chlorobenzene at 130 °C (Table 1). The weak boron Lewis acid BPh₃ failed to promote the reaction (entry 2). B(C₆Cl₅)₃, considered a stronger Lewis acid than B(C₆F₅)₃, was also ineffective, probably due to the increased steric hindrance around the boron center (entry 3). The Brønsted acid Tf₂NH, successfully applied to hydrobromination and -iodination,^{16,17} led to decomposition (entry 4). This finding points to the role of B(C₆F₅)₃ as a catalyst instead of being a simple initiator.²⁶ As expected, no conversion of the substrate was detected in the absence of catalyst (entry 5). An increased catalyst loading of B(C₆F₅)₃ accelerated the reaction but was detrimental to the yield (entry 6 vs. entry 1). Increasing the reaction temperature from 130 °C to 140 °C drastically reduced the reaction time and further improved the yield to 76% (entry 7 vs. entry 1).

We then set out to explore the substrate scope (Scheme 3, top). Although initially optimized for the hydrochlorination of unfunctionalized 1-phenylprop-1-yne (**4a**), we quickly found



Scheme 3 Substrate scope of aryl/heteroaryl-substituted internal alkynes suitable for the transfer hydrochlorination. ^a 74% (E/Z = 71:29) were obtained on a 2.0 mmol scale.

Table 1 Selected examples of the optimization of the transfer hydrochlorination of an internal C–C triple bond (**4a** → **5a**; cf. Scheme 2, top)^a

Entry	Catalyst (mol%)	Time (h)	d.r. ^b (E:Z)	Yield ^b (%)
1	B(C ₆ F ₅) ₃ (10 mol%)	83	74:26	72
2	BPh ₃ (10 mol%)	—	—	<5 (Conv)
3	B(C ₆ Cl ₅) ₃ (10 mol%)	—	—	<5 (Conv)
4	Tf ₂ NH (10 mol%)	—	—	Decomp
5	None	—	—	<5 (Conv)
6	B(C ₆ F ₅) ₃ (20 mol%)	2	74:26	65
7 ^c	B(C ₆ F ₅) ₃ (10 mol%)	7	74:26	76

^a Unless otherwise noted, all reactions were performed in C₆D₅Cl at 130 °C using surrogate *trans*-**9** (0.65 equiv.). ^b Diastereomeric ratio, yield, and conversion were determined by ¹H NMR spectroscopy with dimethyl terephthalate as an internal standard. ^c Reaction performed at 140 °C.

that polar functional groups are tolerated. Their presence also made the chromatographic separation of the alkenyl chlorides from generated side products easier. As expected, aryl-substituted internal alkynes not having electron-withdrawing groups at the aryl ring participated in this ionic process. Various phenol derivatives **4b–g** either as an ether, ester or sulfonate reacted in good yields; the dimethoxyaryl group in **4c** is likely too electron-rich for the acidic reaction medium. Substrate **4e** containing two carboxyl groups was also amenable. A primary alkyl chloride as in **4f** resulted in significant decomposition. A phthalimide group as in **4h** was compatible, furnishing the desired vinyl chloride in high yield. The naphth-2-yl-substituted alkyne **4i** was converted into the product in good yield, while a similar substrate having a naphth-1-yl group was found to decompose under the reaction conditions (see the ESI† for details).



We next investigated a variety of heteroaryl groups, potentially not compatible with $\text{B}(\text{C}_6\text{F}_5)_3$ at high temperature (Scheme 3, bottom). Notably, benzofurans turned out to be suitable substrates as exemplified by **10a** and the 2-substituted derivative **10b**. The reaction of a furyl-substituted derivative was successful but gave an unstable product that could not be isolated (see the ESI† for details). Thiophene derivatives such as **10c** and **10d** underwent the hydrochlorination without any problems.

The stereoselectivity of the reaction was generally modest, with the *E/Z* ratio ranging from 74 : 26 to 39 : 61 in most cases. The obtained major isomer was normally *E*-configured which is in accordance with a vinyl cation as an intermediate accepting a chloride anion from the less hindered side either from the *in situ*-formed borate $[\text{ClB}(\text{C}_6\text{F}_5)_3]^-$ or the surrogate itself.²⁷

The present work is conceptual rather than being an advance in synthetic methodology. We have described a design for storable surrogates of HCl as an alternative to unstable or unreactive cyclohexa-1,4-diene-based platforms. Aromatization of the multiply chlorinated bicyclo[3.1.0]hexanes can be initiated by $\text{B}(\text{C}_6\text{F}_5)_3$ -mediated chloride abstraction from a geminally dichlorinated cyclopropane unit. An electrocyclic ring opening driven by the release of strain followed by loss of a proton leads to a chlorobenzene byproduct with low molecular weight. The surrogate can be designed in such a way that another elimination of HCl is required to achieve aromatization. By this, two molecules of HCl can be generated from one molecule of the surrogate. With an internal C–C triple bond as a proton acceptor,²⁸ a $\text{B}(\text{C}_6\text{F}_5)_3$ -catalyzed transfer hydrochlorination has become possible in moderate yields.

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Conflicts of interest

There are no conflicts to declare.

Notes and references

- 1 J. Zhao, Y. Yue, G. Sheng, B. Wang, H. Lai, S. Di, Y. Zhai, L. Guo and X. Li, *Chem. Eng. J.*, 2019, **360**, 38–46.
- 2 J. Zhao, B. Wang, X. Xu, Y. Yu, S. Di, H. Xu, Y. Zhai, H. He, L. Guo, Z. Pan and X. Li, *J. Catal.*, 2017, **350**, 149–158.
- 3 (a) P. J. Kropp and S. D. Crawford, *J. Org. Chem.*, 1994, **59**, 3102–3112; (b) C.-X. Xu, C.-H. Ma, F.-R. Xiao, H.-W. Chen and B. Dai, *Chin. Chem. Lett.*, 2016, **27**, 1683–1685.
- 4 (a) H. Klein, T. Roisnel, C. Bruneau and S. Dérien, *Chem. Commun.*, 2012, **48**, 11032–11034; (b) S. Dérien, H. Klein and C. Bruneau, *Angew. Chem., Int. Ed.*, 2015, **54**, 12112–12115.
- 5 (a) J. Dupont, N. R. Basso, M. R. Meneghetti, R. A. Konrath, R. Burrow and M. Horner, *Organometallics*, 1997, **16**, 2386–2391; (b) G. Zhu, D. Chen, Y. Wang and R. Zheng, *Chem. Commun.*, 2012, **48**, 5796–5798; (c) J. Derosa, A. L. Cantu, M. N. Boulous, M. L. O'Duill, J. L. Turnbull, Z. Liu, D. M. de La Torre and K. M. Engle, *J. Am. Chem. Soc.*, 2017, **139**, 5183–5193;
- (d) D. A. Petrone, I. Franzoni, J. Ye, J. F. Rodríguez, A. I. Poblador-Bahamonde and M. Lautens, *J. Am. Chem. Soc.*, 2017, **139**, 3546–3557.
- 6 C.-Y. Lo, M. P. Kumar, H.-K. Chang, S.-F. Lush and R.-S. Liu, *J. Org. Chem.*, 2005, **70**, 10482–10487.
- 7 J. Oliver-Meseguer, A. Doménech-Carbó, M. Boronat, A. Leyva-Pérez and A. Corma, *Angew. Chem., Int. Ed.*, 2017, **56**, 6435–6439.
- 8 For a review on transition-metal-catalyzed carbon–halogen bond formation, see: D. A. Petrone, J. Ye and M. Lautens, *Chem. Rev.*, 2016, **116**, 8003–8104.
- 9 The uncatalyzed, Michael-type addition of HCl to electron-deficient alkynes is long known. See: (a) R. Friedrich, *Justus Liebigs Ann. Chem.*, 1883, **219**, 368–374; (b) A. Michael and H. Pendleton, *J. Prakt. Chem.*, 1889, **40**, 63–68.
- 10 (a) R. Ebule, S. Liang, G. B. Hammond and B. Xu, *ACS Catal.*, 2017, **7**, 6798–6801; (b) X. Zeng, Z. Lu, S. Liu, G. B. Hammond and B. Xu, *J. Org. Chem.*, 2017, **82**, 13179–13187; (c) S. Liang, G. B. Hammond and B. Xu, *Green Chem.*, 2018, **20**, 680–684; (d) X. Zeng, S. Liu, G. B. Hammond and B. Xu, *ACS Catal.*, 2018, **8**, 904–909.
- 11 For general aspects of shuttle catalysis, see: (a) B. N. Bhawal and B. Morandi, *Angew. Chem., Int. Ed.*, 2019, **58**, 10074–10103; for an aromatization-driven transfer dichlorination, see: (b) X. Dong, J. L. Roeckl, S. R. Waldvogel and B. Morandi, *Science*, 2021, **371**, 507–514.
- 12 P. Yu, A. Bismuto and B. Morandi, *Angew. Chem., Int. Ed.*, 2020, **59**, 2904–2910.
- 13 (a) J. C. L. Walker and M. Oestreich, *Synlett*, 2019, **30**, 2216–2232; (b) S. Keess and M. Oestreich, *Chem. Sci.*, 2017, **8**, 4688–4695.
- 14 For radical-based approaches, see: A. Bhunia and A. Studer, *Chem*, 2021, **7**, 2060–2100.
- 15 P. Orecchia, W. Yuan and M. Oestreich, *Angew. Chem., Int. Ed.*, 2019, **58**, 3579–3583.
- 16 W. Chen, J. C. L. Walker and M. Oestreich, *J. Am. Chem. Soc.*, 2019, **141**, 1135–1140.
- 17 W. Chen and M. Oestreich, *Org. Lett.*, 2019, **21**, 4531–4534.
- 18 H. Fang and M. Oestreich, *Unpublished results*, Technische Universität, Berlin, 2019.
- 19 M. S. Baird, D. G. Lindsay and C. B. Reese, *J. Chem. Soc. C*, 1969, 1173–1178.
- 20 For the electrocyclic ring opening of cyclopropyl cations, see: (a) R. B. Woodward and R. Hoffmann, *J. Am. Chem. Soc.*, 1965, **87**, 395–397; (b) P. v. R. Schleyer, T. M. Su, M. Saunders and J. C. Rosenfeld, *J. Am. Chem. Soc.*, 1969, **91**, 5174–5176; Orbital symmetry dictates that in the present system, the chloride ion is released from the axial (= “endo”) position. This has been experimentally verified: (c) P. S. Skell and S. R. Sandler, *J. Am. Chem. Soc.*, 1958, **80**, 2024–2025; (d) M. S. Baird and C. B. Reese, *Tetrahedron Lett.*, 1967, **8**, 1379–1382.
- 21 For a similar approach exploiting the ring strain of pro-aromatic bicyclo[2.2.0]hexenes, see: B. Wu, J. Wang, X. Liu and R. Zhu, *Nat. Commun.*, 2021, **12**, 3680.
- 22 (a) W. E. Parham and H. E. Reiff, *J. Am. Chem. Soc.*, 1955, **77**, 1177–1178; (b) W. E. Parham, H. E. Reiff and P. Swartzentruber, *J. Am. Chem. Soc.*, 1956, **78**, 1437–1440; for a seminal report, see: (c) G. L. Ciamician and M. Dennstedt, *Ber. Dtsch. Chem. Ges.*, 1881, **14**, 1153–1163.
- 23 I. Fleming and E. J. Thomas, *Tetrahedron*, 1972, **28**, 4989–5001.
- 24 (a) G. A. Olah and J. M. Bollinger, *J. Am. Chem. Soc.*, 1967, **89**, 4744–4752; (b) G. A. Olah, J. M. Bollinger and J. Brinich, *J. Am. Chem. Soc.*, 1968, **90**, 2587–2594.
- 25 (a) D. Seyferth and V. A. Mai, *J. Am. Chem. Soc.*, 1970, **92**, 7412–7424; (b) R. H. Ellison, *J. Org. Chem.*, 1980, **45**, 2509–2511.
- 26 S. Banerjee and K. Vanka, *ACS Catal.*, 2018, **8**, 6163–6176.
- 27 As exceptions, the electron-rich substrates **4b** and **10d** gave almost pure *Z*-isomers as products. Presumably, thermodynamic control favors isomerization to the *Z* isomer. As a general trend, the stereoselectivity of the reaction was found to considerably vary with the alkyne substrate but to be almost independent of the used surrogate. This indicates that chloride is released from the surrogate before the $\text{C}(\text{sp}^2)\text{--Cl}$ bond is formed. As an alternative explanation for the observed diastereomeric ratios, thermodynamic equilibration of the alkenyl chlorides cannot be ruled out.
- 28 Terminal alkynes decomposed and dialkyl-substituted internal alkynes did not react.

