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FEATURE ARTICLE

Recent progress on MN_4 macrocycle-derived oxygen reduction reaction electrocatalysts for polymer electrolyte fuel cell applicationZubair Ahmed,^a Karina Muñoz-Becerra,^{*b} Srinu Akula,^a José H. Zagal,^{*c} and Kaido Tammeveski^{*a}Received 00th January 20xx,
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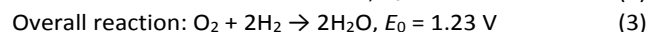
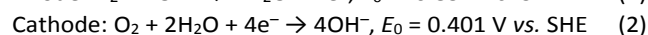
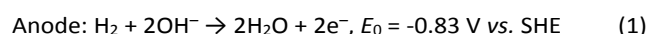
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Transportation and power generation sectors have been garnering strong support for fuel cell-based technologies in recent times for emission-free energy conversion. However, the performance of these devices largely relies on the costly and scarce Pt-group metal-based electrocatalysts, which hinder their widespread commercialization. For the past few decades, a vast research effort has been made to develop cost-effective and efficient non-precious metal electrocatalysts for low-temperature fuel cells. Single-atom catalysts (SACs) have attracted the most attention due to their excellent electrocatalytic performance. Moreover, SACs derived from transition metal MN_4 macrocycles have many advantages in terms of structural integrity of the active centers that are beneficial for superior electrocatalytic activity and durability. This feature article covers the recent reports focused on the development of the MN_4 macrocycle-derived oxygen reduction reaction (ORR) electrocatalysts for proton exchange membrane fuel cell (PEMFC) and anion exchange membrane fuel cell (AEMFC), and summarizes the viability in energy applications, including their stability issues. This review aims to discuss the various factors that influence the fuel cell performance of cathode catalysts based on transition-metal macrocyclic complexes (metal porphyrins and phthalocyanines).

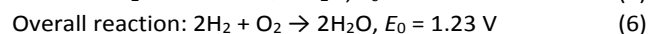
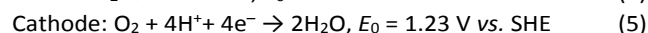
1. Introduction

Polymer electrolyte fuel cells are economically viable and sustainable electrochemical energy conversion devices that have received increasing attention during the last decade.^{1,2} The anion exchange membrane fuel cell (AEMFC) and proton exchange membrane fuel cell (PEMFC) consist of the anode and cathode coated with catalyst layers on the gas diffusion layer (GDL) electrode and with an ion exchange membrane. In AEMFCs, O_2 is reduced in the presence of water to OH^- ions at the cathode that migrate through the anion exchange membrane (AEM) to the anode side, reacting with H_2 to produce H_2O as the final product (reactions 1-3).³ During these processes, electrons that are generated at the anode are collected by current collectors and transferred to the cathode electrode through the external circuits. The ideal AEMs must have excellent OH^- conductivity and ion exchange capacity, chemical and thermal stability, mechanical strength, flexibility, and low gas permeability, while preventing fuel crossover and a

circuit break by blocking the electron transport through it.³ The theoretical thermodynamic cell voltage of a single AEMFC is 1.23 V at room temperature. However, the practical cell voltage is much less than 1.23 V due to the cathodic and anodic overpotentials, ohmic drop, mass-transport losses, and fuel crossover. Hydrogen-fueled AEMFCs:



PEMFCs share many technical similarities with AEMFCs. The main difference lies in the type of ion exchange membrane used in each system. However, as protons diffuse much faster than OH^- ions and as a result, PEMFC manages to obtain higher ionic conductivity and power density. Anodic and cathodic reactions in the PEMFC with H_2 gas as a fuel are given in reactions 4-6:



The oxygen reduction reaction (ORR) process may follow two main pathways in aqueous media. The desired $4e^-$ ORR pathway involving the splitting of the O-O bond shows a minimum yield of peroxide formation.⁴ In contrast, in a $2e^-$ ORR pathway, a high peroxide yield is involved, which is detrimental to the fuel cell catalyst and membrane stability.⁵⁻⁷ It is not only detrimental to the fuel cell components, but the peroxide route delivers much less energy as compared to the full 4-electron reduction

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pathway. A well-established methodology is available to obtain kinetic parameters from the rotating (ring)-disk electrode (RRDE/RDE) tests. However, due to several aspects such as a thin layer of catalyst in the RDE test applied to a glassy carbon electrode instead of a gas diffusion layer, differences in gas flow pattern, water flooding at micropores, etc., the experimental results in the RDE/RRDE system might mismatch the environment of polymer electrolyte fuel cells.^{8–17}

Platinum group metal-based electrocatalysts are mainly used at the polymer electrolyte fuel cell anode, the cathode, or both.^{18,19} Importantly, non-precious metal catalysts have been considered for years as low-cost promising alternatives, specially at the cathode electrode.^{20,21} There is a large amount of literature dealing with metal macrocyclic complexes (metalloporphyrins, metallophthalocyanines, metallocorroles) for ORR application,^{22–29} however, there are only a few reviews concentrating on their fuel cell performance.^{30–32} This review focuses on the rational design for the development of highly efficient transition metal MN₄ macrocycle-based ORR electrocatalysts to boost their performance in AEMFC and PEMFC applications.

2. ORR on transition metal macrocycle-derived M-N-C catalysts

Despite their excellent ORR activity, MN₄ macrocycles face important limitations, especially in acidic media, where metal center dissolution and degradation of the macrocyclic framework compromise their stability and long-term durability. Furthermore, their inherently low electrical conductivity restricts charge transport, thereby limiting overall electrocatalytic performance. To address these issues, various strategies have been developed to immobilize MN₄ macrocycles onto conductive carbon supports. These include simple physical adsorption through π - π stacking interactions and covalent or coordinative attachment via axial ligands such as pyridine.^{33–35} The physicochemical properties of the carbon support – particularly, specific surface area, porosity, and electrical conductivity – play a crucial role in determining the electrocatalytic performance of the resulting catalyst materials. These parameters have been extensively investigated for both pristine MN₄ macrocycles supported on nanostructured carbons and their corresponding pyrolyzed M-N-C derivatives. Particularly, high surface area and well-developed porosity enhance the accessibility and dispersion of active sites, leading to improved electrocatalytic ORR activity.^{33,34,36} A notable example involves single-walled carbon nanotubes functionalized with pyridine moieties, which were employed as axial ligands for pristine (16Cl)FePc complexes. This strategy led to improved ORR performance in alkaline media, attributed to electronic interactions between the macrocycle and the functionalized support, which modulate the electronic density and 3d-orbital configuration of the Fe-center in the FeN₄ core, thereby optimizing adsorption of oxygenated intermediates.³⁴ Such electronic perturbations resemble those observed in pyrolyzed M-N-C catalysts. In these systems, the N-doped

carbon matrix acts not only as support but also as an electronic modulator of the anchored MN₄ sites, which is fundamental to their enhanced ORR activity. Fe phthalocyanine (FePc) and Co porphyrin physically adsorbed on the reduced graphene oxide (rGO) nanosheets exhibited excellent electrocatalytic activity for ORR in alkaline media. This indicates efficient electron transfer between intact MN₄ macrocycles and rGO, and, therefore, these catalyst materials are very promising for ORR application.³⁷ A recent study by Zhao et al.³⁸ employing a model Fe-N-C catalyst based on a conjugated FePc=FePc dimeric system supported on carbon black XC-72 provided important insights into the nature of the ORR-active sites. The authors proposed that the primary ORR-active site corresponds to a five-coordinated OH-Fe³⁺N₄ moiety with a spin multiplicity of 5/2, in which all five Fe 3d-orbitals are singly occupied. Density functional theory (DFT) calculations demonstrated that a moderate *d*- π interaction between the Fe center and the conjugated carbon plane governs both the spin-state of the active centers and their adsorption energies with the oxygenated reaction intermediates, ultimately accounting for their superior ORR activity. The integration of MN₄ macrocyclic complexes with carbon substrates induces electronic modulation at the active site through intermolecular charge transfer. DFT calculations suggest that the highest occupied molecular orbital of the MN₄ complex in MN₄ macrocycle-carbon composites aligns favorably with the π^* orbital of O₂, thereby promoting oxygen activation and reduction.^{34,39} Although these supported macrocycles serve as valuable model systems for fundamental mechanistic studies in ORR electrocatalysis, their practical application remains constrained by insufficient activity and stability under operando conditions.^{9,38,40} A mechanistic study by Wan et al.⁴¹ provided important insights into the degradation pathway of FePc supported on KJ600 carbon black. Under acidic, O₂-saturated electrolyte at applied potentials, the reactive oxygen species, particularly ·OH and ·OOH radicals, were identified as the main contributors to degradation. These species initiate hydroxylation at the C=C bonds within the macrocycle, disrupting aromaticity, breaking C-N bonds, and ultimately leading to demetallation of the Fe center.

2.1. Thermal conversion of MN₄-macrocycles for obtaining M-N-C catalysts

Following Jasinski's pioneering report on Co phthalocyanine as an ORR catalyst,⁴² Jahnke and co-workers demonstrated that pyrolyzing MN₄ complexes supported on carbon under an inert atmosphere significantly improved both activity and stability in acidic media compared to their unpyrolyzed counterparts.^{43,44} This pivotal study laid the groundwork for developing pyrolyzed M-N-C catalysts, utilizing MN₄ macrocycles and other metal, nitrogen, and carbon precursors as starting materials.^{45–48} The thermal conversion of MN₄ macrocycles typically involves first supporting the pristine complex on a conductive, high-surface-area carbon matrix, such as low-cost carbon black^{47–51} or carbide-derived carbons (CDC),⁵² followed by pyrolysis at 400–1000 °C under inert or NH₃ atmospheres. The thermal



treatment methodology used to obtain pyrolyzed M-N-C catalysts is a critical factor in determining their final electrocatalytic activity. In many cases, a post-pyrolysis acid washing step is applied to remove unstable metal-based nanoparticles formed during thermal treatment (e.g., carbides, oxides, and pure-metallic phases), enhancing the exposure of catalytically active MN_x ($x = 2-5$) sites.^{45,47,50,53-56}

Pyrolysis under NH_3 can further increase microporosity and introduce basic N-functionalities, which can boost the ORR activity but often at the expense of long-term stability.^{53,57} Charretier et al.⁵⁷ showed that the pyrolysis atmosphere significantly impacts Fe-N-C catalyst performance. Using iron tetramethoxyphenylporphyrin (FeTMPP) supported on carbon black N-330 and pyrolyzed at 950 °C, they showed that NH_3 -treated catalyst delivered higher initial ORR activity at pH = 1 (17.1 mA mg^{-1} at 0.8 V_{RHE}), but poor stability. In contrast, the Ar-treated catalyst showed lower activity (1.9 mA mg^{-1} at 0.8 V_{RHE}), but superior durability. The enhanced electrocatalytic activity after NH_3 treatment was linked to increased microporosity, while improved stability in Ar was attributed to greater graphitization. Similar trends were observed by Meng et al.⁵⁸, reinforcing the trade-off between activity and stability. While NH_3 serves as an external nitrogen source, the macrocyclic ligands themselves also contribute to nitrogen doping, with varying effects on the electrocatalytic performance. Roncaroli et al.⁵⁹ systematically investigated how nitrogenated heterocyclic ligands (ranging from five to six-membered rings with one to four N atoms, including tetra-4-carboxyphenylporphyrin, TCPP) affect the ORR activity of pyrolyzed Fe- and Co-based catalysts supported on Vulcan carbon XC-72. The complexes were pyrolyzed at 500, 700, and 900 °C under N_2 , yielding the highest activity in acidic media at the optimal temperature of 700 °C. Polypyridinic ligands like phenanthroline (phen) outperformed TCPP, enhancing activity, kinetics, and selectivity of the final catalyst. Although Fe-phen and bimetallic FeCo-phen catalysts showed lower peroxide yield, they underperformed compared to Co-phen in onset potential and kinetics. The highest ORR activity was linked to a greater N-retention (notably pyridinic-N) and reduced CoO content. However, the stability of Fe- and FeCo-based catalysts was not assessed, despite their promising selectivity.

Among MN_4 macrocycles, the most studied correspond to Fe and Co phthalocyanines and porphyrins,^{45,47,48,60-62} which have well-known molecular structures enabling one to carry out systematic studies on pyrolysis processes to obtain efficient ORR electrocatalysts. Although M-N-C catalysts can initially match the ORR activity of Pt/C, their long-term stability, especially in PEMFCs, remains a major limitation. Fe-based pyrolyzed catalysts exhibit high initial activity and selectivity toward the 4-electron pathway, but their stability suffers under acidic conditions (pH < 4). This is due to Fe-mediated Fenton reactions, which decompose H_2O_2 into reactive oxygen species that degrade the carbon support.⁶³ To improve durability, bi- and tri-metallic M-N-C systems incorporating Co and/or Mn have been developed. While these metals enhanced stability by suppressing Fenton-type reactions, they often exhibit lower selectivity for the direct 4-electron pathway compared to Fe-N-

C catalysts, resulting in hydrogen peroxide yields under certain conditions.⁶³ Introducing both Fe and Co into a single pyrolyzed catalyst has shown synergistic effects, improved ORR activity, selectivity, and stability compared to monometallic counterparts.⁶⁴ Jiang and Chu⁶⁵ synthesized a CoFe-N-C catalyst by pyrolyzing a mixture of hemin and cobalt(II) tetramethoxyphenylporphyrin (CoTMPP) impregnated on Vulcan carbon XC-72R at 700 °C under N_2 . In half-cell tests, the catalyst exhibited onset potentials of 0.75 V_{RHE} (acidic) and 0.85 V_{RHE} (alkaline), following a 2+2 electron pathway (initial O_2 reduction at Co sites to H_2O_2 , followed by further reduction to H_2O at Fe sites). When integrated into a membrane electrode assembly (MEA), the catalyst showed a stability trend of acidic > neutral > alkaline, outperforming the Fe-N-C analogue. Similarly, Kumar et al.⁶⁶ reported the excellent ORR activity for pyrolyzed M-N-C catalysts (800 °C under N_2) derived from Fe-, Co-, Ni-, and Mn-phthalocyanine supported on multiwalled carbon nanotubes (MWCNTs). The FeCoN-MWCNT variant achieved the highest onset potential in alkaline media (0.93 V_{RHE}), linked to the formation of highly active MN_x sites. Sun et al.⁵⁵ developed a trimetallic Co_3Fe_7/N , Mn-PC catalyst by pyrolyzing a mixture of CoTMPP, hemin, manganese acetate, and dicyandiamide, followed by acid leaching. The resulting material, composed of Co_3Fe_7 alloy nanoparticles embedded in a Mn- and N-doped porous carbon matrix, exhibited superior ORR activity in alkaline media, with an onset potential of 0.98 V_{RHE} , outperforming mono- and bimetallic analogs. This improvement was attributed to intermetallic synergy, an optimal ratio of graphitic-N and pyridinic-N, and a high electrochemically active surface area, enabling a 4-electron pathway. The active sites were proposed to consist of Co_3Fe_7 nanoparticles and Fe/Co- N_x moieties, while Mn- and N-doping enhanced electrical conductivity by tuning the electronic structure. The catalyst also showed excellent methanol tolerance and stability, exceeding the performance of commercial Pt/C under alkaline conditions.

2.2. Structural evolution and site preservation during pyrolysis of MN_4 macrocycles for obtaining M-N-C catalysts

Although pyrolysis often introduces structural rearrangements and partial loss of molecular identity, several studies have shown that MN_4 macrocycle cores can be preserved or converted into well-dispersed MN_x ($x = 2-5$) sites embedded within nitrogen-doped graphitized carbon matrices.^{45,47,48,67} These matrices typically contain graphitic-N, pyridinic-N, and pyrrolic-N species, which contribute to electronic conductivity and activity. The resulting MN_x sites are regarded as the primary ORR-active centers and are commonly classified as single-atom catalysts (SACs), operating predominantly through a 4-electron pathway.² Among the different nitrogen species, pyridinic-N is particularly effective in promoting peroxide reduction to water, acting as a secondary active site.⁹ The MN_4 -site preservation during pyrolysis depends on the temperature and precursor design. Some studies suggest that pyrolyzing metallomacrocycles below 600-700 °C favors retention of MN_4 motifs with high atomic dispersion, whereas annealing at higher



temperatures ($>700\text{ }^{\circ}\text{C}$) often causes their decomposition into less-defined species and favors the formation of metal nanoparticles due to agglomeration.^{45,62,67–70} Such nanoparticle formation reduces the accessibility of the MN_x sites and may consequently impair ORR activity. Other studies have shown that MN_4 motifs can persist even after high-temperature pyrolysis, provided that the precursor design is carefully controlled.^{71–73} However, the final structural outcome strongly depends on multiple factors, including the synthesis route, precursor concentration, pyrolysis conditions, and axial ligand modification. Anchoring an axial ligand to an in-plane square-planar MN_4 site results in a distorted square-pyramidal MN_5 geometry. This axial coordination modifies the ligand field and electronic density at the metal center, thereby influencing the adsorption energies of oxygenated intermediates and tuning the M(III)/M(II) redox potential, ultimately impacting ORR activity.^{74,75}

More broadly, strategies aimed at incorporating axial ligands (e.g., O, S, N, or halogens) to expand the coordination sphere of MN_x sites have emerged as a promising approach to enhance the intrinsic ORR activity of the metal center. Such approaches include the rational design of precursors containing pre-coordinated axial ligands prior to pyrolysis, as well as post-synthetic adsorption or solution impregnation treatments applied to the pyrolyzed catalysts.^{74,75} However, achieving precise synthetic and thermally stable five-coordination environments under high-temperature pyrolysis conditions remains a significant synthetic challenge.

In parallel to coordination expansion strategies, careful engineering has also proven effective in preserving well-defined MN_4 motifs during pyrolysis. For instance, Zhai et al.⁵¹ demonstrated that ionic self-assembly of Fe(III) mesotetra(N -

methyl-4-pyridyl)porphine (FeNMePyP) and Co(III) mesotetra(4-sulfonatophenyl)porphine (CoTPPS_4),^{10394566G074783} impregnated onto carbon EC-600 (60% loading) and pyrolyzed at $650\text{ }^{\circ}\text{C}$ under inert atmosphere, produced a highly active ORR catalyst with preserved FeN_4 and CoN_4 sites (Fig. 1a). The catalysts exhibited an E_{onset} of $1.01\text{ V}_{\text{RHE}}$ in alkaline media (Fig. 1b), operating via a 4-electron pathway and showed excellent durability and methanol tolerance. The high performance was attributed to ionic interactions between the oppositely charged macrocycles and π - π stacking with the carbon support, promoting uniform self-assembly and controlled thermal transformation. This prevents the formation of metal nanoparticles, as confirmed by XRD, and preserves the MN_4 cores, as evidenced by XPS. In contrast, mechanically mixed controls showed lower onset potentials due to metal agglomeration.⁵⁷ ^{57}Fe Mössbauer spectroscopy further confirmed the exclusive formation of D1 and D2 Fe-sites (Fig. 1c), which are highly accessible and catalytically active. Similarly, Ye et al.⁷⁶ demonstrated that tailoring monomer connectivity (number of different sites to which the ligands connect) in macrocyclic precursors directly influences the coordination environment and electrocatalytic ORR performance in alkaline media, thereby validating the close correlation between catalyst structure and performance in these systems.

Recently, Muhyuddin et al.⁶² investigated the structural evolution of Fe-based sites during pyrolysis using in-situ X-ray absorption spectroscopy (XAS) combined with ex-situ analyses of FePc supported on EC-600JD carbon (10% loading), pyrolyzed at $600\text{ }^{\circ}\text{C}$ under inert atmosphere. The X-ray absorption near edge structure (XANES) linear combination fitting confirmed the preservation of atomically dispersed FeN_x species within defective carbon matrices up to $600\text{ }^{\circ}\text{C}$ (Fig. 2a). Beyond this

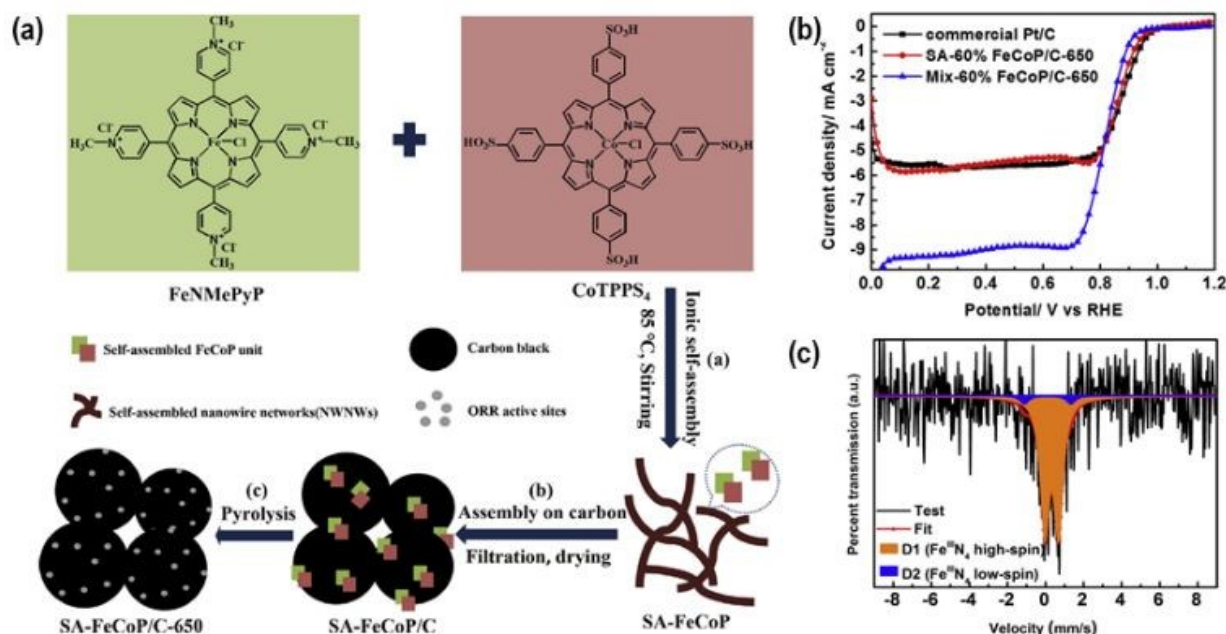


Fig. 1. (a) Schematic representation of the synthesis of ionic self-assembly and the pyrolysis process of Fe- and Co-macrocycles supported on carbon substrate. (b) ORR polarization curves in alkaline media of the pyrolyzed catalysts (SA = self-assembled synthesis methodology; Mix: mechanical mix synthesis methodology). (c) ^{57}Fe Mössbauer spectra of the pyrolyzed catalyst synthesized by the SA process. Reproduced from ref. 51 with permission from Elsevier, copyright 2021.



temperature, FeN₄ moieties degrade, favoring nanoparticle formation and nitrogen loss, both factors being detrimental to ORR performance in acid and alkaline media. The highest ORR activity (E_{onset} , $E_{1/2}$, and J_L) correlated with elevated FeN₄ content, accompanied by minor fractions of FeN₃ and FeN₂ species and abundant pyridinic-N functionalities (Fig. 2b-c). Similar retention of FeN₄ sites at 600 °C has also been reported in bi- and tri-metallic pyrolyzed M-N-C catalyst derived from Fe, Co, and Cu phthalocyanines.⁷⁰ These findings underscore the importance of carefully balancing thermal treatment conditions to maximize site preservation while minimizing coordination degradation.

Beyond structural preservation, heat treatment also induces electronic modifications. Several studies have shown that pyrolysis shifts the M(III)/(II) redox potential of the MN₄ catalysts toward more positive values, making the metal center “more noble” and harder to oxidize.⁷⁷ The M(III)/(II) redox potential has thus been proposed as a reactivity descriptor for ORR activity, as shown in several publications.^{27,78} The fundamental understanding of molecular governing factors in nonprecious electrocatalysts for ORR was reported by Ramaswamy et al.⁷⁹

The mechanism of the ORR on the pyrolyzed nonprecious metal catalysts has been rationalized using an activity descriptor based on the principles of surface science, along with the coordination chemistry to establish a relationship with the redox potentials over the pyrolysis temperature (Fig. 3a). Upon heat-treatment of iron(III) meso-tetraphenylporphine chloride (FeTPPCI) at 800 °C, the Fe^{2+/3+} redox potential shifted significantly positive, and the Fe-N-C catalyst tended to be more ORR active. Interestingly, the electrocatalytic ORR activity of the pyrolyzed Fe-N-C catalyst depended also on the carbon support

material (Fig. 3c), which was explained by a change of the electron-withdrawing ability (Fig. 3b) and the degree of π -electron delocalization on disordered graphitic carbon. The intrinsic ORR activity, quantified in terms of turnover number at 0.8 V (Fig. 3d), further corroborates the structure-activity relationship. The proposed redox mechanism governing ORR on pyrolyzed Fe-N_x/C is illustrated in Fig. 3e.

Consequently, the ORR activity of MN₄ macrocycle-derived M-N-C catalysts is governed not only by site density but also by coordination geometry, local electronic structure, and support-induced electronic modulation. Collectively, these findings indicate that pyrolysis should be viewed as a coordination-engineering process capable of generating and stabilizing distinct MN_x environments, including FeN₅ configurations, with different electrocatalytic behavior. Therefore, elucidating the interplay between M-N coordination, electronic structure, and electrochemical descriptors remains essential for the rational design of high-performance M-N-C catalysts.

2.3. The ORR multiple reactivity descriptors in MN₄ macrocycle-derived M-N-C catalysts

The simplest reactivity descriptor for MN₄ macrocycles is the number of *d*-electrons when comparing MN₄ complexes sharing the same ligand.^{27,80} A volcano-shaped correlation has been observed for the ORR activity for both metallophthalocyanines and metalloporphyrins, but the optimum metal for phthalocyanines is Fe, while for porphyrins is Co. It is possible that the position of the maximum could change depending on the applied potential and the redox potential M(III)/(II) of the catalysts that will control the fraction of M(II) active states at a given potential. However, for Co seems to be the Co(II)/(I) redox potential that controls the activity.

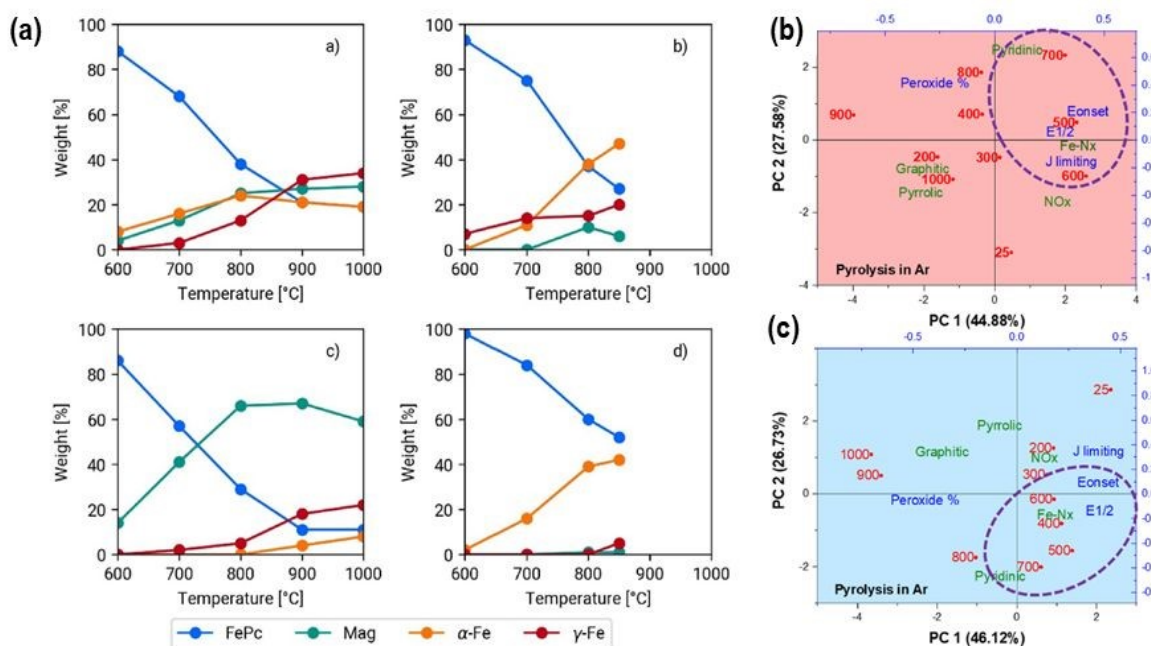


Fig. 2. (a) XANES linear combination fitting of the pyrolyzed (600 °C) FePc on EC-600JD carbon. Biplots acquired through principal component analysis (PCA) show the relationships among the surface-active moieties detected by XPS and performance parameters of FePc-derived catalysts at different temperatures during pyrolysis under Ar in (b) acidic and (c) alkaline electrolytes. Reproduced from ref. 62 with permission from Elsevier, copyright 2024.



A classical reactivity descriptor in electrocatalysis is the site-reactant intermediate binding energy, in this case using the $\text{M}-\text{O}_2$ binding energy, a plot of $\log(j)_E$ versus this descriptor gives volcano-type plots, where the highest activity is observed for intermediate binding energies, following the well-known Sabatier Principle, which states that optimum activity is observed for binding energies not too strong not too weak.⁸⁰ Too weak binding results in not sufficient activation of the reactant's bonds. For too strong binding, the reaction products are too stable and cannot desorb from the active sites. The optimum condition corresponds to a binding energy of bound reactant that is equal to half the bond energy, which is to be formed or broken, i.e.: $[\text{M(II)}]_{\text{ad}} + \text{O}_{2(\text{aq})} \rightleftharpoons [\text{M(III)}-\text{O}_2]_{\text{ad}} \Delta G_{\text{ad}} = 0$ (ligand is omitted for simplicity).

However, the binding energies are generally calculated in the absence of an electrolyte and electrode potential, and the latter can affect the surface concentration of M(II) active sites. A more practical reactivity descriptor is the $\text{M(III)}/\text{(II)}$ redox potential, which can be measured under the same experimental conditions at which the activities are determined,⁸⁰ i.e., an overpotential and the presence of the electrolyte. When $\log(j)_E$ is plotted versus the $\text{M(III)}/\text{(II)}$ redox potential, a volcano plot is observed. However, more recently it has been reported that, especially for Co intact complexes (not heat-treated), a better fit is observed if the $\text{Co(II)}/\text{(I)}$ redox potential is considered.⁸¹ The strong binding region or descending region in the volcano correlation is explained according to the Sabatier Principle does not seems to be realistic as discussed by Zagal and co-workers,^{80,82} the latter reports turnover frequency (TOF) values calculated on the basis of the surface concentration of M(II) active species at the electrode potential used for comparison

show an increase of this parameter in the descending region of the volcano and the volcano plot becomes a linear plot.

Recently, a new reactivity descriptor for ORR has been proposed, the "electrochemical hardness" (ΔE^0).^{83–85} This concept has been used to explain the reactivity of a series of FeN_4 macrocyclic complexes in a wide range of pH values. The *electrochemical hardness* is defined as the potential separation between the redox processes $E^0_{\text{M(III)}/\text{(II)}}$ and $E^0_{\text{M(II)}/\text{(I)}}$ of surface-confined MN_4 macrocycles, measured in the absence of O_2 and in aqueous media. These studies have been performed in both acidic and alkaline solutions. For the specific case of FePc , the second one-electron reduction redox process has been attributed to an electron-transfer occurring on the ligand, i.e., $E^0_{\text{Fe(II)Pc}/\text{Fe(II)(Pc-1)}}$.⁸⁴ Experimental results indicate that the ORR activity, expressed as $(\log j)_E$, diminishes when the "electrochemical hardness" increases.^{84,85} Intact MN_4 macrocyclic complexes, as discussed above, have served as useful model catalysts to identify several reactivity descriptors. However, for practical application, they lack the long-term stability required for polymer electrolyte fuel cell performance, especially in acid media.

Yang et al.⁸⁶ reported that pyrolysis of axial imidazole-coordinated iron 5,10,15,20-tetra(2-thienyl)porphyrin (FeTThP) supported on carbon black at 800 °C under an inert atmosphere yielded dispersed FeN_5 sites. The resulting catalyst showed enhanced ORR activity in alkaline media ($E_{1/2} = 0.89 \text{ V}_{\text{RHE}}$) compared to its sulfur-free analogue (Fig. 4a). The extended X-ray absorption fine structure (EXAFS) analysis confirmed the formation of axially coordinated FeN_5 sites (Fig. 4b), suggesting that the combined effect of axial coordination and S-doping within the graphitized carbon matrix synergistically modulates the electronic structure of the active Fe centers. This work

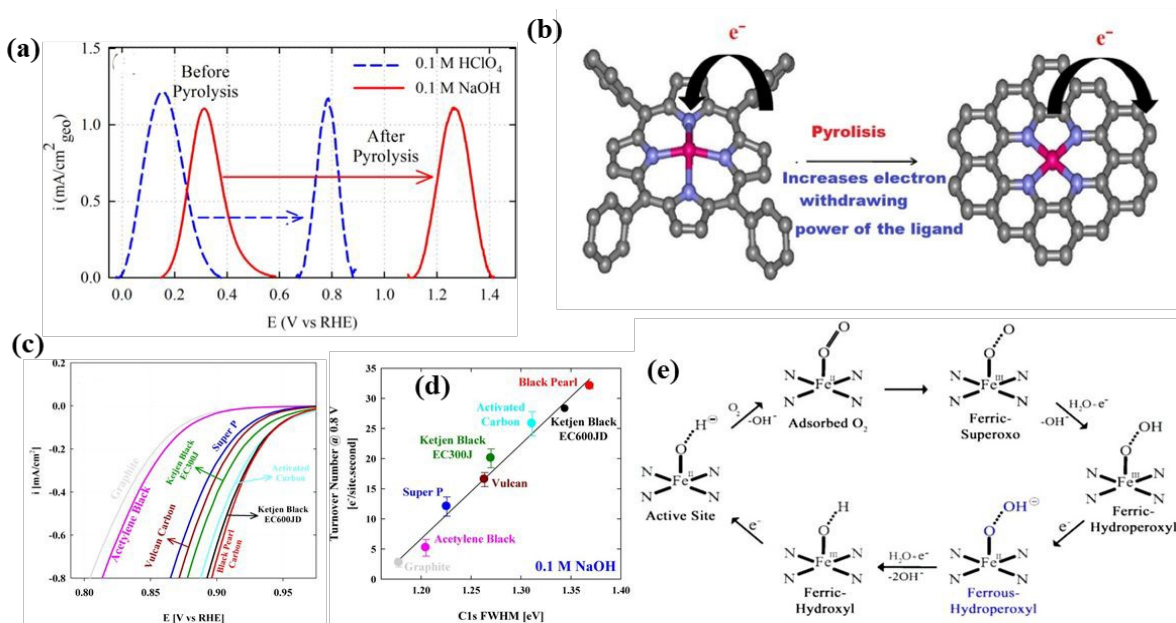


Fig. 3. (a) Square-wave voltammetry profiles of $\text{Fe}-\text{N}_4$ catalyst supported on black pearl carbon before and after pyrolysis at 800 °C in 0.1 M HClO_4 and 0.1 M NaOH electrolytes. (b) Structural models of active site structures before and after pyrolysis at 800 °C. Color codes in structural models: red, iron; blue, nitrogen; gray, carbon; white, oxygen. (c) ORR activity of $\text{Fe}-\text{N}_4/\text{C}$ catalyst (pyrolyzed at 800 °C) on various carbon supports in O_2 -saturated 0.1 M NaOH electrolyte at 900 rpm rotation rate and 20 mV s^{-1} scan rate. (d) Linear relationship between ORR turnover numbers in 0.1 M NaOH electrolyte versus full-width at half-maximum of C 1s photoemission spectra. (e) Catalyst cycle showing the redox mechanism involved in ORR on pyrolyzed $\text{Fe}-\text{N}_4/\text{C}$ active sites in dilute alkaline medium. Reproduced from ref. 79 with permission from the American Chemical Society, copyright 2013.



highlights the importance of precursor architecture in inducing coordination expansion within Fe-N-C catalysts.⁷⁴ Pyrolysis at 800 °C in an inert atmosphere of iron(III) 2,3,7,8,12,13,17,18-octaethylporphyrin chloride (8EtFeP) and iron(III) 5,10,15,20-tetraphenylporphyrin chloride (4PhFeP) supported on Vulcan XC-72 significantly enhances the ORR activity compared to unpyrolyzed precursors (Fig. 4c).⁸⁷ The ORR onset potentials of

these materials align closely with the Fe(III)/Fe(II) redox couple as discussed above.^{77,80} This improvement is attributed to the stabilization of FeN₄ units within N-doped graphitized carbon matrices, which exert electron-withdrawing effects on the metal centers (Figs. 3b and 4d). The electron-withdrawing effect shifts the Fe(III)/(II) redox potential to more positive values, which increases the electrocatalytic activity for ORR.^{77,79,88}

Pyrolysis shifts the onset potentials by ~100 mV to more positive values relative to the pristine complexes (Fig. 4e,f).⁸⁷ A similar trend is observed for Co-based analogs, where pyrolysis raises the $E^{\circ}_{\text{Co(III)/Co(II)}}$ and, accordingly, the ORR onset potential in both acidic and alkaline media (Fig. 4g).⁸⁹ In the Fe-based systems, the ORR follows either a direct 4-electron or 2×2 pathway, while Co-based catalysts tend to show mixed selectivity, producing both H₂O₂ and H₂O. In all cases, the ORR activity increases as the redox potential of the MN₄ (or MN_x) sites becomes more positive, following a linear $\log(j) - E^{\circ}_{\text{redox}}$ correlation (Fig. 4h), consistent with non-pyrolyzed macrocycle trends.⁸⁰ These results confirm that M(II)N₄ or M(II)N_x sites remain the dominant active centers in both acidic and alkaline media, with the graphitized N-doping lowering the M-O₂ binding energy and enhancing catalytic efficiency.^{80,90}

Domínguez et al.⁹¹ impregnated Fe, Co, and Mn phthalocyanines onto N-doped activated carbon and pyrolyzed the mixture at 800 °C under N₂. After acid-washing, the ORR activity in acidic media followed the trend Fe > Co > Mn. This was attributed to higher N-content and increased surface area, which improved the accessibility and dispersion of MN_x active sites. Notably, the trend also correlates with the number of d-electrons, a well-established reactivity descriptor for MN₄ macrocyclic complexes such as phthalocyanines and porphyrins.^{80,90}

2.4. Sustainable approaches for obtaining MN₄ macrocycle-derived M-N-C catalysts

In an effort to lower the production costs and align with circular economy principles, Giulini et al.⁹² explored the use of waste polypyrromethenes, by-products from the Adler-Longo synthesis of porphyrins (80% yield), as a precursor for Fe-based pyrolyzed ORR catalysts. Both purified porphyrins and the corresponding waste were impregnated onto EC-600 carbon and pyrolyzed at 600 °C under an inert atmosphere. The resulting Fe-N-C catalysts showed comparable ORR activity, with onset potentials of 0.937 and 0.972 V_{RHE} in alkaline electrolyte, proceeding via 2×2 and 4-electron pathway with minimal hydrogen peroxide formation. This study demonstrated the viability of macrocycle synthesis by-products for fabricating efficient ORR electrocatalysts. Similarly, biomass-derived carbon supports have gained attention for enhancing surface area and improving FeN_x site accessibility.^{93–95} For instance, FeTPP was combined with pre-carbonized, Si-rich rice husk and pyrolyzed at 800 °C in an inert atmosphere to yield Fe-N-C catalysts co-doped with Si, S, and P. Among them, FeN/SSIRC catalyst exhibited the highest performance, with a half-wave potential of 0.885, 0.773, and 0.682 V_{RHE} in alkaline, acidic, and neutral media, respectively. This was attributed to a

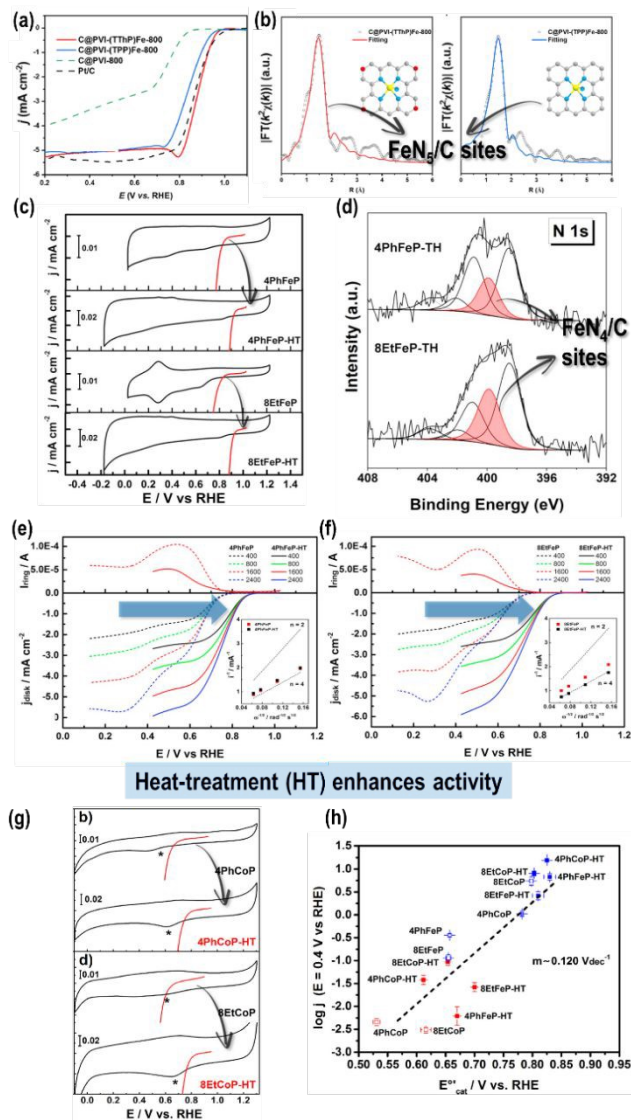


Fig. 4. (a) ORR polarization curves for the Fe-N-C catalysts in O₂-saturated 0.1 M KOH solution at 1600 rpm. (b) FT-EXAFS R-space fitting curves of the obtained Fe-N-C catalysts, including a structural model of Fe-N₄. Reproduced from ref. 86 with the permission from the American Chemical Society, copyright 2024. (c) Cyclic voltammograms for unpyrolyzed and pyrolyzed porphyrinic-based 4PhFeP and 8EtFeP catalysts in N₂-saturated alkaline media (black line), including kinetic currents (1600 rpm, 0.005 V s⁻¹) for ORR (red lines) in O₂-saturated alkaline media. (d) N1s XPS spectra of 4PhFeP and 8EtFeP pyrolyzed at 800 °C. (e) ORR activity comparison toward ORR for unpyrolyzed and pyrolyzed 4PhFeP and (f) 8EtFeP catalysts in O₂-saturated 0.1 M NaOH. Reproduced from ref. 87 with the permission from Elsevier, copyright 2019. (g) Cyclic voltammograms for unpyrolyzed and pyrolyzed phthalocyanine-based 4PhCoP and 8EtCoP catalysts in N₂-saturated acidic media (black line), including kinetic currents for ORR (red lines) in O₂-saturated 0.1 M H₂SO₄ (1600 rpm, 0.005 V s⁻¹). (h) Linear correlation between $\log(j)$ at 0.4 V (vs RHE) and E°_{redox} of unpyrolyzed Fe- and Co-based porphyrinic catalyst (open symbols) and pyrolyzed (closed symbols). Red and blue symbols indicate the values measured under acidic and alkaline media, respectively. Reproduced from ref. 89 with the permission from Elsevier, copyright 2020.



high density of exposed FeN_x sites and synergistic effects from pyridinic-N, Si, and S dopants, which were shown via DFT calculations to lower the ORR energy barrier.

2.5. Porous and structured supports used for the thermal treatment of MN_4 macrocycle-derived M-N-C catalysts

In some cases, hard templates such as silica or magnesium oxide have been employed to produce highly active catalysts with enhanced mesoporosity, which improves mass transport and accessibility to active sites.^{93,96,97} More recently, metal-organic frameworks (MOFs) have been explored as precursors to further increase surface area, improve site dispersion, and boost the density of atomically dispersed MN_x active sites. For instance, MOF-5 and ZIF-8 have been employed as supports for metallomacrocycles such as FeTPP^{98} and FePc^{99} to produce efficient ORR catalysts in alkaline media. Their high electrocatalytic activity is attributed to the large surface area, which enhances FeN_x site accessibility, and the synergistic effect of secondary N-doping and residual metal nanoparticles. Additionally, porphyrinic-MOFs have gained interest as precursors for obtaining pyrolyzed M-N-C catalysts. Jiao et al.¹⁰⁰ synthesized a Fe-N-C catalyst by pyrolyzing MOF-545 at 800 °C in inert atmosphere (Fig. 5a), yielding a high density of isolated FeN_4 sites. The absence of Fe nanoparticles and the presence of oriented mesoporous channels boosted active site exposure,

leading to E_{onset} and $E_{1/2}$ values of 1.00 V_{RHE} and 0.89 V_{RHE} , respectively, in alkaline media, and 0.93 V_{RHE} and 0.78 V_{RHE} in acidic media (Fig. 5b-c). In a related study, a Zn-CoTCPP catalyst was prepared from $[\text{Zn}_2(\text{COO})_4]$ paddle-wheel SBUs and cobalt-tetrakis(4-carboxyphenyl)porphyrin (CoTCPP), followed by pyrolysis at 400 °C in inert atmosphere. The resulting material exhibited high CoN_4 site density (Fig. 5d), uniformly dispersed in a covalent-bonded quasi-carbon matrix, with the $E_{1/2}$ value of 0.82 V_{RHE} in alkaline media (Fig. 5e), making it a promising candidate for ORR applications.¹⁰¹

3. MN_4 macrocycle-derived cathodes for AEMFC and PEMFC applications

Since a decade, numerous researches have been conducted using different MN_4 macrocycle-derived catalyst materials for the improvement of polymer electrolyte fuel cell performance in search of alternative cathodes for Pt-based catalysts although the pioneering work by Jasinski on cobalt phthalocyanine-based ORR electrocatalyst dates back to 1964.⁴² On the account of the structural assets of MN_4 moieties, the efficient ORR activity and stability drags attraction in recent times to address the durability hurdles, specifically in acidic conditions.^{102–104} The ORR activity data determined from RDE experiments and the AEMFC^{32,105,106} and PEMFC¹⁰⁷ performances of the cathode

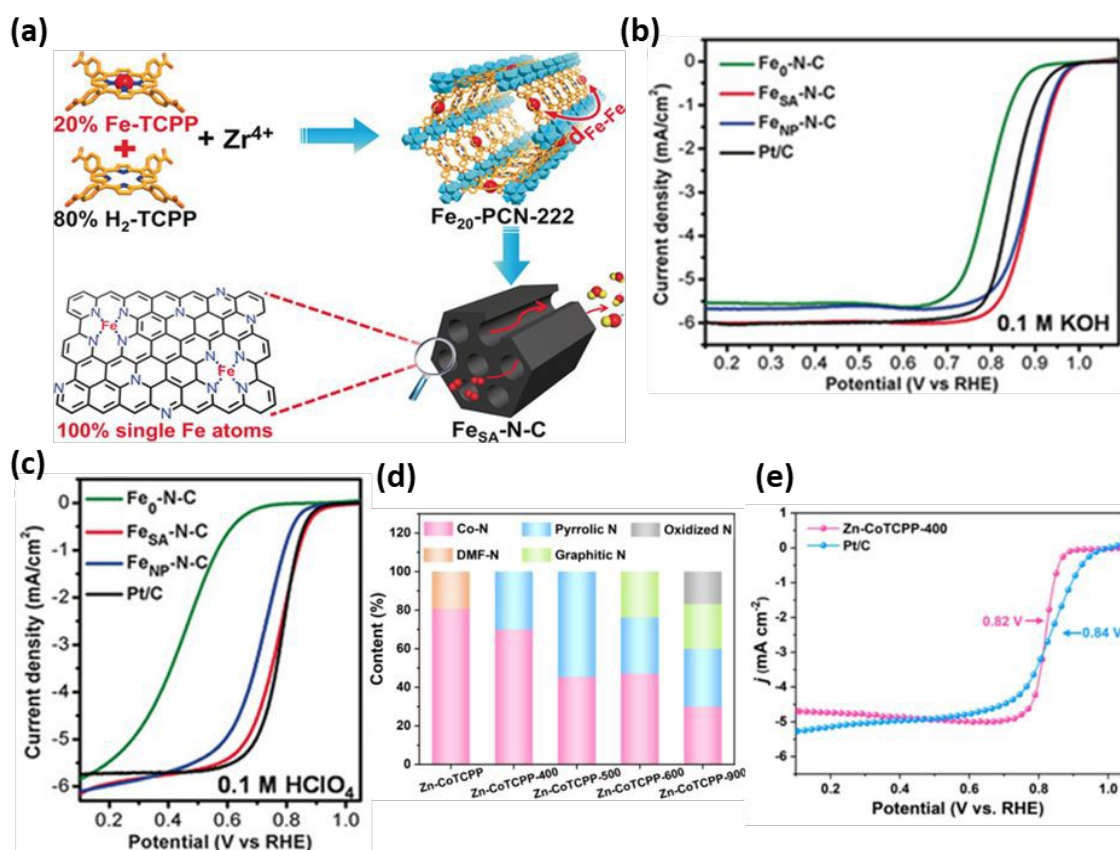


Fig. 5. (a) Schematic representation of the synthetic route for obtaining Fe-N-C catalyst using MOF-545, including FeTCPP and H₂TCPP moieties. Linear sweep voltammeteries obtained for the MOF-545-derived catalysts in (b) alkaline and (c) acidic media. Reproduced from ref. 100 with the permission from John Wiley and Sons, copyright 2018. (d) Normalized content of deconvoluted N-species quantified by high-resolution N 1s XPS. (e) LSV curves for Zn-CoTCPP-400 electrocatalyst in O₂-saturated 0.1 M KOH. Reproduced from ref. 101 with the permission from Elsevier, copyright 2025.



catalysts are summarized in Tables 1 and 2, respectively. The following subsections (3.1 and 3.2) specifically discuss recent advances in MN₄ macrocycle-derived catalysts investigated as cathodes for AEMFCs and PEMFCs, with emphasis on their structural design, performance optimization, and durability under relevant operating conditions.

Table 1. Summary of the ORR parameters from half-cell test and H₂-O₂ AEMFC performance for the MN₄ macrocycle-derived non-precious metal cathode catalysts.

Half-cell RDE data			H ₂ /O ₂ AEMFC data		
Catalyst	<i>E</i> _{onset} (V) vs. RHE	<i>E</i> _{1/2} (V) vs. RHE	OCV (V)	<i>P</i> _{max} (mW cm ⁻²)	Ref.
D-MN ₄ -CNF-IL-A	0.97	0.86	1.05	231	108
FePc/KB	0.96	-	0.98	524	109
MnPc/KB	0.86	-	0.93	248	109
CoPc/KB	0.84	-	0.95	605	109
NiPc/KB	0.82	-	0.81	350	109
FeNiN-MWCNT	0.94	0.87	1.0	406	110
CNT/PC (Fe)	0.97	0.88	0.96	380	111
FePc/C (600°C)	0.95	0.88	0.88	123	112
CoFeNx/C	0.95	0.84	0.87	38	65
MnPc/C	0.72	0.55	0.90	97	113
FePc/C	0.87	0.70	0.83	87	113
CoPc/C	0.79	0.70	0.87	105	113
NiPc/C	0.69	0.59	0.90	88	113
Fe-Phen/CNT	0.99	0.90	1.00	410	114
Fe-P-Phen/CNT	0.99	0.91	1.00	570	114
Fe-S-Phen/CNT	0.99	0.91	1.00	635	114
FeCoN-MWCNT	0.93	0.86	0.98	692	66
FeMnN-MWCNT	0.93	0.85	0.95	582	66
Fe@Fe-N-C	0.99	0.82	-	242	115
SiCDC/CNT(1:3)/CoPc	0.91	0.81	0.94	473	116
SHUb-Fe/N-A	0.94	0.80	1.01	234	54
SHUb-Co/N-A	0.93	0.80	0.98	232	54
SiCDC/CNT(1:3)/FePc	0.87	0.74	1.00	164	117
TiCDC/CNT(1:3)/FePc	0.93	0.77	1.00	182	117
pyrolyzed KB/FePc	0.94	0.76	1.00	186	117
FePc/MWCNT	0.86	-	0.98	60	118
CoPc/MWCNT	0.91	-	0.99	100	118
HT800-FeP aerogel	0.96	0.86	0.93	580	119

Table 2. Summary of the ORR activity from the half-cell test and PEMFC performance for the MN₄ macrocycle-derived non-precious metal cathode catalysts.

Catalyst	<i>E</i> _{1/2} (V) vs RHE	Cell temp. (°C)	Back pressure (MPa)	Anode gas - cathode gas	Catalyst loading (mg cm ⁻²)	<i>P</i> _{max} (mW cm ⁻²)	Ref.
1000-III-NH ₃ -g	-	80	0.2	H ₂ -O ₂	4	380	122
Co/N/C	-	50	0.2	H ₂ -O ₂	10-12	150	123
NC Por_0.8-1050	-	80	0.2	H ₂ -O ₂	4	600	124
Fe-N _x /C-NC1100	0.811	80	0.1	H ₂ -O ₂	4	656	125
Fe-N _x /C-NC1100	0.811	80	0.1	H ₂ -air	4	312	125
COP _{BTC} -F	0.74	80	0.16	H ₂ -O ₂	2	201	126
COP _{BTC} -F	0.74	80	0.16	H ₂ -air	2	147.46	126
L_FeMn	0.77	80	-	H ₂ -air	4.4	72	127
Fe-14MR/FePI	-	80	0.2	H ₂ -air	4,5	486	128

3.1. Template-assisted and confinement strategies for MN₄ macrocycle-derived M-N-C cathodes

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Sa et al.¹¹¹ reported the fabrication of Fe-N-C single-atom catalysts by proposing a silica-protective-layer-mediated pyrolysis method for preferential formation of catalytically active Fe-N₄ moieties in Fe-N-C material by effectively preventing the formation and agglomeration of inorganic Fe-based nanoparticles during the pyrolysis process (Fig. 6a-b). The resulting CNT/PC catalyst showed good ORR activity in both alkaline and acidic media (Fig. 6c), and demonstrated better fuel cell performances in both AEMFC and PEMFC (Fig. 6d,e). Moreover, structural uniformity of Fe-N_x sites made them a model system to precisely control their coordination environments and configurations, exposing more active sites, and promoting the efficient mass transport for a deeper understanding of the electrocatalytic properties of the catalysts. Similarly, Woo et al.¹¹⁴ synthesized Fe,N-doped carbon catalysts via silica coating-mediated pyrolysis-based strategy using 2,1,3-benzothiadiazole and triphenylphosphine as S and P precursors, respectively, with iron acetate/1,10-phenanthroline, and carbon nanotubes (CNTs). Both S and P-containing catalysts showed high ORR electrocatalytic activity, but interestingly S-containing catalyst outperformed the P-doped catalyst in a single-cell AEMFC test. The XANES and EXAFS analyses suggested that the silica-coating strategy proved to be useful for inhibiting the Fe aggregation and improves the density of Fe-N_x sites, resulting in improved ORR activity than the uncoated one. The Fe-S-Phen-CNT catalyst achieves an impressive maximum power density of 635 mW cm⁻². By integrating a nitrogen-rich precursor such as spent coffee grounds (SCGs) with the synergistic interaction of atomically dispersed Fe and Mn, bimetallic SCG-derived electrocatalysts were developed with enhanced ORR activity.¹²⁰ The graphene aerogel-supported materials act as efficient electrocatalysts by effectively suppressing the restacking and agglomeration of graphene sheets. This work not only shows the promise of a pyrolysis-free synthesis strategy for developing cost-effective catalysts, but also highlights the critical role of catalyst layer architecture in fully realizing the performance potential of non-precious metal catalysts for AEMFC applications.¹²¹



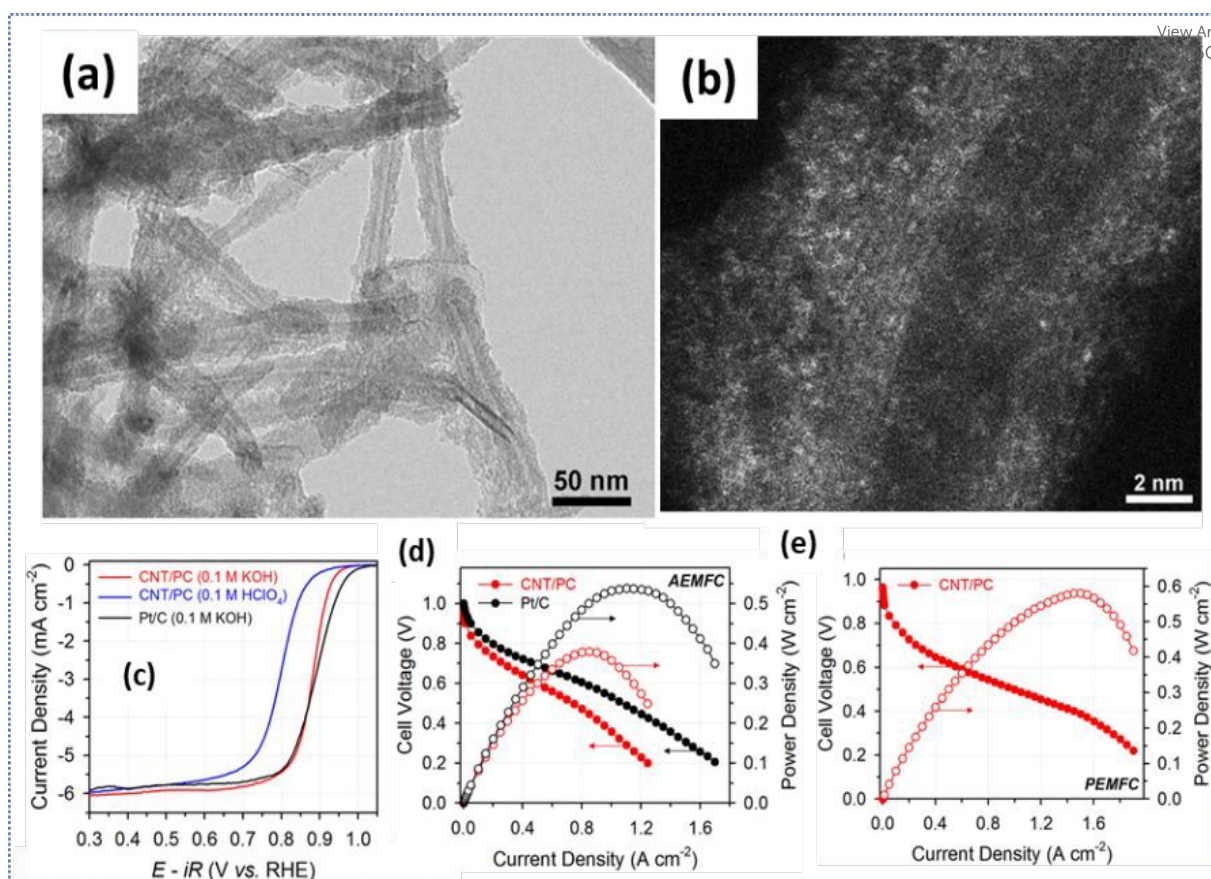


Fig. 6 (a) TEM and (b) HAADF-STEM image of CNT/PC. (c) ORR activity of CNT/PC. (d) AEMFC performances using CNT/PC and Pt/C catalysts, and (e) PEMFC performance with CNT/PC cathode catalyst. Reproduced from ref. 111 with the permission from the American Chemical Society, copyright 2016.

The strategy of tailoring the catalyst layer architecture through polymer composite platforms offers a new and versatile pathway for implementing a wide range of molecular catalysts in fuel cell applications.¹²⁹ By employing a complete suite of analytical techniques that establishes the pivotal role of tailored material design, structural characteristics, and synthesis strategies in defining electrocatalytic performance toward the ORR in AEMFCs. Especially, the controlled integration of silica and metal species, the removal of structure-directing agents, and the use of nitrogen-rich precursors are shown to regulate the catalyst morphology, pore architecture, and chemical composition.¹³⁰ Recently, Dekel and co-workers,¹⁰⁹ reported the pyrolysis-free preparation of Ketjen Black (KB) supported MPC-based catalysts (M = Fe, Mn, Co, Cu, Ni, Li, Sn) through single-step, room-temperature ultrasonication process that fetches valuable insight for efficient catalyst design and the broader adoption of transition metal phthalocyanine-based catalysts and AEMFC technology.

Praats et al.¹¹⁶ synthesized a cobalt-containing catalyst using Co phthalocyanine (CoPc) as a precursor on carbon nanotube and carbide-derived carbon (CDC) support prepared from two sources (i.e., titanium carbide and silicon carbide), where the Si-CDC-supported catalyst was found to possess enhanced ORR activity. Detailed analysis suggests a higher amount of nitrogen species on Si-CDC-supported catalyst as compared to the Ti-CDC counterpart. The CoPc-derived cathode catalyst revealed

excellent AEMFC performance. In another study, Praats et al.¹¹⁷ used FePc as a precursor for the preparation of Fe-N-C catalyst. The catalyst material consisting of TiC-derived carbon showed higher ORR activity than the catalyst based on the SiC-derived support. This leaves the scope to probe the electronic properties of such materials to unlock their potential for better electrocatalytic ORR behavior. However, it is envisioned that this can be achieved by connecting material characterization and theoretical calculations in such a way that it can substantially impact and predict the performance of M-N-C catalyst materials.

3.2. Influence of pyrolysis on the structure and activity of MN_4 macrocycle-derived M-N-C cathodes

Our early work on heat-treated MWCNT-supported MPC-based catalysts has shown a decent ORR electrocatalytic activity in 0.1 M KOH and promising AEMFC performance ($P_{\max}$ of 100 and 60 $mW\ cm^{-2}$ for CoPc/MWCNT and FePc/MWCNT, respectively).¹¹⁸ The MPC-MWCNT mixture was pyrolyzed in inert atmosphere at 800 °C to prepare ORR active and stable catalysts. Chung et al.¹⁰⁶ reported a low-temperature and gram-scale approach to synthesize FePc-embedded two-dimensional carbon sheets by annealing at 450 °C for an efficient cathode at AEMFC. Generally, the employment of Fe-N-C electrocatalysts has been found to deliver high electrocatalytic ORR performance owing to a suitable electronic structure and interaction with O_2 when



the iron is coordinated to nitrogen in FeN_x active sites.^{106,131} Furthermore, it is well established that the electrocatalytic activity of the FePc is strongly associated with the specific interactions of the cations/anions at the electrolyte/electrode interface, which influence the energy barriers and reorganization energies of the intermediates formed during the reaction.⁸⁴ Consequently, it becomes excellent models for studying and assessing the electronic factors on the Fe-N_x site that affect the ORR performance, serving as an excellent functional molecular block to study reactivity descriptors for Fe-N-C catalysts. Zhu et al.¹³² fabricated metal phthalocyanine (MPc) on a carbon support and tested for AEMFC application. The MPc (M = Co, Cu, Zn, and Ni) on Vulcan carbon were pyrolyzed at 600 and 800 °C and compared with non-pyrolyzed MPcs. The untreated MPcs revealed the following order of the ORR activities: CoPc/C > CuPc/C > NiPc/C > ZnPc/C. For 600 °C pyrolyzed catalysts, CoPc/C was found to be the most active one among the other catalysts studied. For 800 °C treated materials, the activity order changed, and CuPc/C showed the highest ORR activity. The study reflects the large influence of heat-treatment temperature on the MPc/C catalyst activity. The MPc/C catalysts showed much lower fuel cell performances than that of Pt/C due to the complexity of the MEA preparation and the other influencing factors (such as the catalyst ink, spraying technology, hot-press temperature as well as the amount of binder).

Interestingly, several authors have tuned the redox properties of FePc by chemical substituents in their molecular structure with electronegative fluorine atoms.¹³³ It was observed that the powerful electron-withdrawing nature of F-atoms promotes a Fe-O₂ interaction during the ORR process as a result of a shift of the Fe(III)/(II) redox potential to more positive values compared to unsubstituted FePc. Elbaz and co-workers¹³⁴ found the enhanced electrocatalytic ORR activity in the electrochemical half-cell test, which was also translated into AEMFC performance as shown in Fig. 7a-d. The enhanced ORR activity and fuel cell performance were attributed to the synergistic interactions in the macrocyclic complex and efficient active centers created by the uniformly distributed atomic iron moieties. However, the study lacks a detailed elucidation of the physicochemical nature of the complex governing the ORR selectivity, activity, and durability in F-substituted FePc-based catalyst. However, increasing the site density of M-N₄ centers always remains challenging; this limitation arises from the random distribution of M-N₄ sites during high-temperature pyrolysis, making precise control over their site density complicated. Yang et al.¹³⁵ introduced a high content of F-doping into the carbon substrate (6.29 at%), and assembled poly(iron phthalocyanine) on the prepared substrate to afford a highly efficient electrocatalyst for ORR. Owing to the induced orbital coupling between electron-withdrawing fluorine atoms in the F-doped carbon substrates and metal centers, thereby promoting the delocalization of Fe d-orbital electrons. The optimized PFePc/FC catalyst demonstrated enhanced ORR kinetics.

Kumar et al. prepared Fe and Ni phthalocyanine-derived bimetal electrocatalysts of various nanocarbon supports.¹¹⁰ The

mixture of MPCs and carbon nanomaterials was heat-treated at 800 °C in N₂ atmosphere to obtain highly active electrocatalysts for ORR. FeNiN-MWCNT and FeNiN-MC were reported to be the most active catalysts in 0.1 M KOH among the nanocarbon-supported MPC-based materials studied. In AEMFC conditions, the FeNiN-MWCNT cathode exhibited the highest $P_{\max}$ of 406 mW cm⁻², followed by FeNiN-MC ($P_{\max}$ = 386 mW cm⁻²). Also, the current density at the cell voltage of 0.6 V was higher for FeNiN-MWCNT (638 mA cm⁻²) as compared to the FeNiN-MC catalyst (528 mA cm⁻²). Muuli et al.¹⁰⁸ employed polyacrylonitrile-based electrospun carbon nanofiber (CNF) as a support material for FePc and CoPc to achieve high electrocatalytic activity and good electrical conductivity for AEMFC performance. The influence of catalyst preparation conditions (i.e., MPc content, presence of porogen, pyrolysis temperature, acid leaching) was investigated. The best electrocatalyst revealed a high $E_{1/2}$ value (0.857 V_{RHE}), good durability ($\Delta E_{1/2}$ = 6 mV) in 0.1 M KOH and substantial $P_{\max}$ value of 231 mW cm⁻² in AEMFC conditions. Muhyuddin et al.¹²⁷ used lignin-based biomass to produce activated porous char, which was functionalized with FePc and MnPc via a simplistic pyrolysis process to obtain monometallic and bimetallic electrocatalysts. The developed electrocatalysts performed well in both alkaline and acidic media, and affirm the utility in both PEMFC and AEMFC.

However, the commonly used pyrolysis procedure typically destroys MN₄ centers present in metal phthalocyanines and porphyrins, but results in improved stability if the pyrolysis procedure is designed to prevent the loss of nitrogen species.¹²² Miller et al.¹¹² studied the ORR activity of a FePc-based catalyst prepared at different pyrolysis temperatures. They observed that the ORR activity of FePc/C catalyst decreased as the pyrolysis temperature increased above 600 °C. To get into insights, they utilized the EXAFS analysis and reported that FeN₄ as the active sites for ORR remained thermally stable up to 700

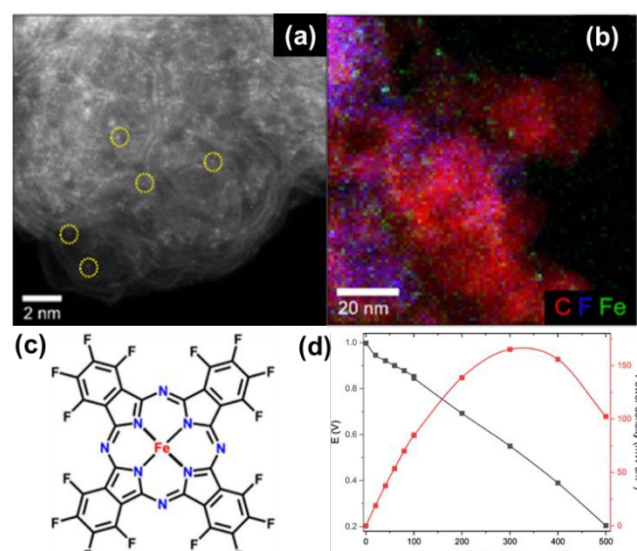


Fig. 7. (a) HAADF image and (b) corresponding EDS mapping. (c) Proposed structure of iron phthalocyanine substituted with fluorine. (d) Polarization curve (black) and power density curve (red) of FePc@BP2000 in AEMFC. Operating conditions: 80 °C, 0.5 L min⁻¹, H₂/O₂, 2 mg cm⁻² catalyst loading at the cathode, and 0.6 mg_{Pt} cm⁻² at the anode. Reproduced from ref. 134 with the permission from the American Chemical Society, copyright 2021.



°C, but above 800 °C, the coordination number decreased subsequently forming metallic iron nanoparticles causing decomposition of FeN_x species at higher temperatures and is responsible for a decrease in the electrocatalytic ORR activity (Fig. 8a-d).

Chougule et al.¹³⁶ did a systematic optimization of pyrolysis conditions to establish the critical relationship between catalyst structure, active site configuration, and electrocatalytic performance. The optimized Fe-N-C/C catalyst exhibited high ORR activity and durability over prolonged electrochemical testing and displayed competitive single-cell AEMFC performance. Kumar et al.⁶⁶ presented a one-step pyrolysis method for the synthesis of mixed M-N-C-type catalysts, where iron phthalocyanine was used in combination with different metal phthalocyanines (e.g., FeNi, FeMn, and FeCo) and was studied for the ORR in alkaline conditions. The results suggest that FeCoN-MWCNT is the most active catalyst for the ORR. Also, the FeCoN-MWCNT catalyst exhibited a high peak power density value (692 mW cm⁻²) in an AEMFC as compared to the other prepared catalysts (Fig. 9a-c). Persky et al.¹³⁷ reported another innovative bioinspired catalyst using a directed synthetic pathway to create adjacent Cu and Fe sites composed of a covalent 3D framework in an aerogel form. The aerogel-

based catalyst exhibits high performance in a half-cell in 0.1 M KOH, with an onset potential of 0.94 V vs RHE and a half-wave potential of 0.80 V vs RHE, high selectivity toward four-electron ORR, and an *i*R-corrected peak power density of 0.51 W cm⁻² was delivered in an AEMFC with an open circuit potential of 0.97 V (Fig. 9d). This study allowed the development of even more active, durable, and suitable electrocatalytic systems for high-performing fuel cells.

Nevertheless, if the afore-mentioned active moieties are buried within carbon layers, they will not actively participate in the ORR process.¹³⁸ Nabae et al.¹²⁸ utilized an organosynthetic approach to establish Fe-N-C catalyst as a support material for the Fe-14MR complex. This approach was designed to increase Fe-N₄ centers both on the outer surface and within the meso and micropores. As a result, their most optimized catalyst led to superior ORR kinetics in half-cell testing and showed promising results in PEMFC testing in terms of the power density and stability for longer run (Fig. 10a,b). Carbon aerogels are promising candidates for carbonaceous substrates for the large-scale preparation of M-N-C catalysts, owing to their porous nature and high electrical conductivity. Elbaz and co-workers¹³⁹ combined the advantages of the FeN₄ sites in iron porphyrin and a 3D network of covalent organic framework and carbon

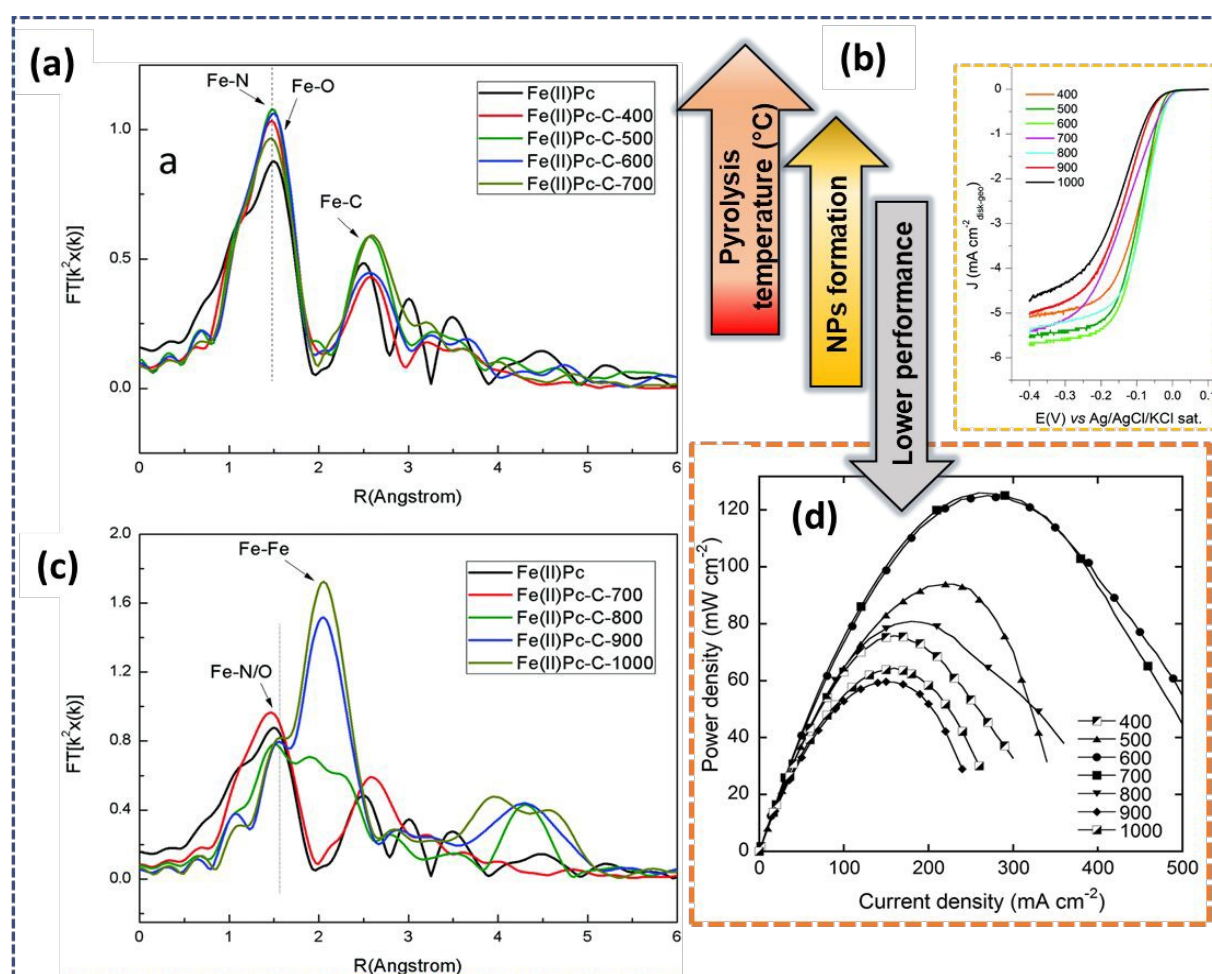


Fig. 8. Fourier transform of EXAFS of unpyrolyzed FePc and FePc samples pyrolyzed (a) up to 700 °C, (c) samples pyrolyzed from 700–1000 °C. (b) ORR polarization curve in an O₂-saturated 0.1 M KOH solution, (d) power density curves for H₂-O₂ AMFCs. T = 55 °C. H₂-O₂ (100% relative humidity, RH) with the flow rate of 100 and 200 mL min⁻¹, respectively. Reproduced from ref. 112 with the permission from RSC Publishing, copyright 2016.



aerogel, which manage to achieve high PEMFC performance with $P_{\max}$ of 292 mW cm^{-2} (Fig. 10c) owing to increased surface area and high density of FeN_4 sites. The dispersion of the Fe atoms before and after the heat-treatment was demonstrated through the ADF-STEM images of the pristine FeTEPC aerogel (Fig. 10d) and the heat-treated Fe800TEPC aerogel (Fig. 10e). In order to design high-performing catalysts and understand the nature of active sites in M-N-C type catalysts, it is important to expand understanding of how the density and location of heteroatoms and MN_x sites and pyrolysis temperature directly affect the electrocatalytic ORR behavior of the M-N-C materials. This could involve high-temperature pyrolysis of different precursors in a controlled manner to interpret the formation of active sites, as well as prevent the agglomeration of FeN_x centers during the annealing process. Qu et al.¹²⁵ utilized a stepwise pyrolysis strategy to construct atomically dispersed FeN_x sites on highly graphitized carbon support, resulting in significantly enhanced electrochemical stability in acid media and a high-power density and stability in PEMFC testing conditions.

Non-traditional heat-treatment methods such as microwave-based methods, have also been explored to synthesize M-N-C catalysts, which have shown some success. For example, Ma et al.¹²³ prepared carbon-supported CoTMPP electrocatalyst by

applying the microwave method and reported decent power density and stability under fuel cell operation at a current density of 200 mA cm^{-2} for 15 h operation. Low-temperature synthesis has attracted wide attention to enhance the stability and ORR activity as this method can retain the well-defined chemical structures of the MN_4 macrocyclic compounds. This method offers a uniform distribution of MN_4 sites on carbon nanosheets. However, these composite catalysts could have durability problems because of their dissolution/dispersion into the acidic or alkaline electrolytes. Alternatively, MN_x macrocyclic covalent organic polymers are mostly fabricated by networking the MN_x macrocyclic monomers with inorganic and/or organic linking. In addition, this would also reduce the density of the MN_x active sites in the catalyst materials. Zhang et al.¹⁴⁰ reported a facile linker-free formation of the FePc conjugate structure by one-pot microwave heating. They employed in situ polymerization of FePc on carbon scaffolds to enhance π - π stacking, which strengthens the interaction between FePc and carbon support by increasing the delocalization of π electrons. Moreover, the use of advanced methods has allowed a better understanding of M-N-C catalysts, and significant progress can be seen in the field of fuel cell research and development.

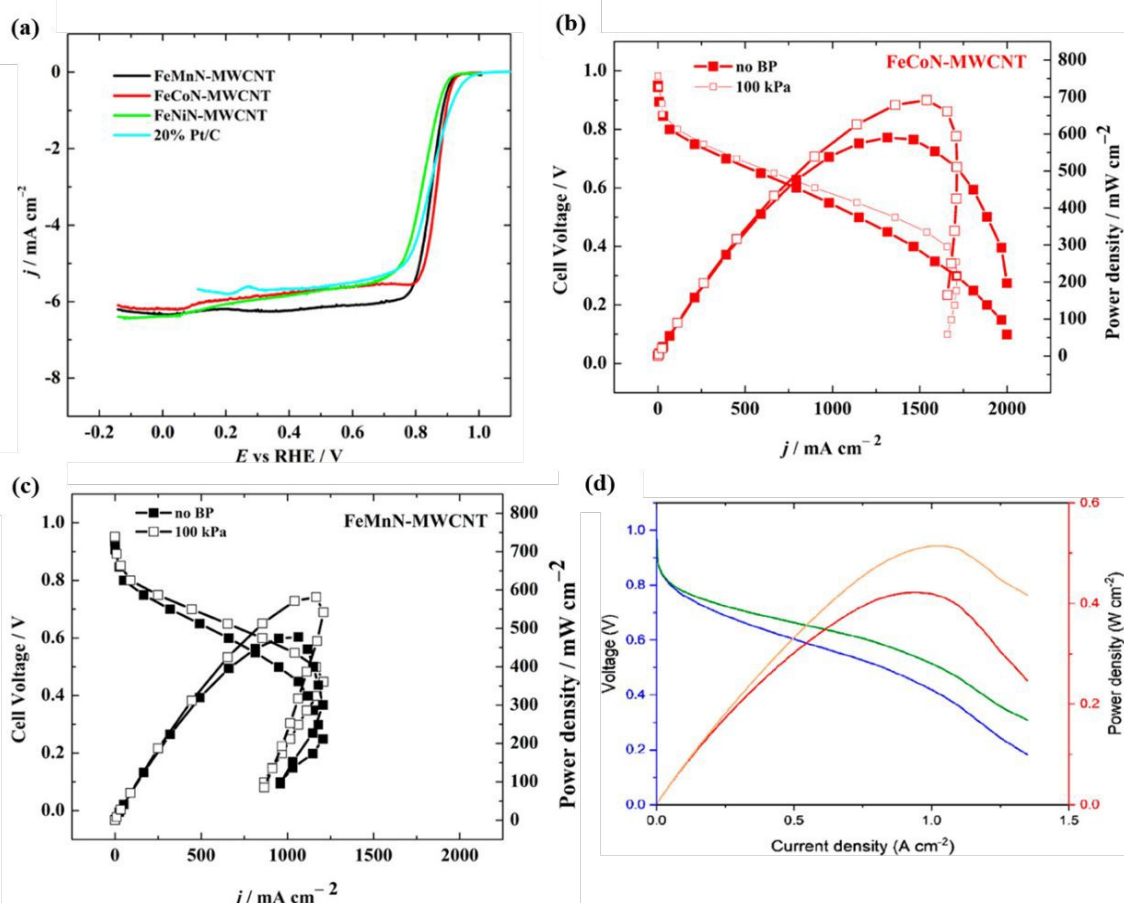


Fig. 9. (a) ORR polarization curves recorded in O_2 -saturated 0.1 M KOH. AEMFC polarization and power density curves using (b) FeCoN-MWCNT and (c) FeMnN-MWCNT cathode catalysts. Reproduced from ref. 66 under a Creative Commons Attribution Licence (CC-BY), copyright 2021 the authors. (d) AEMFC performance using HT-FeCu porphyrrole aerogel-based cathode catalyst. Testing conditions: temperature $75/77/80^\circ\text{C}$, H_2/O_2 , 150 sccm. Performances with (green and orange) and without (blue and red) iR correction. Reproduced from ref. 137 under a Creative Commons Attribution Licence (CC-BY), copyright 2023 the authors.



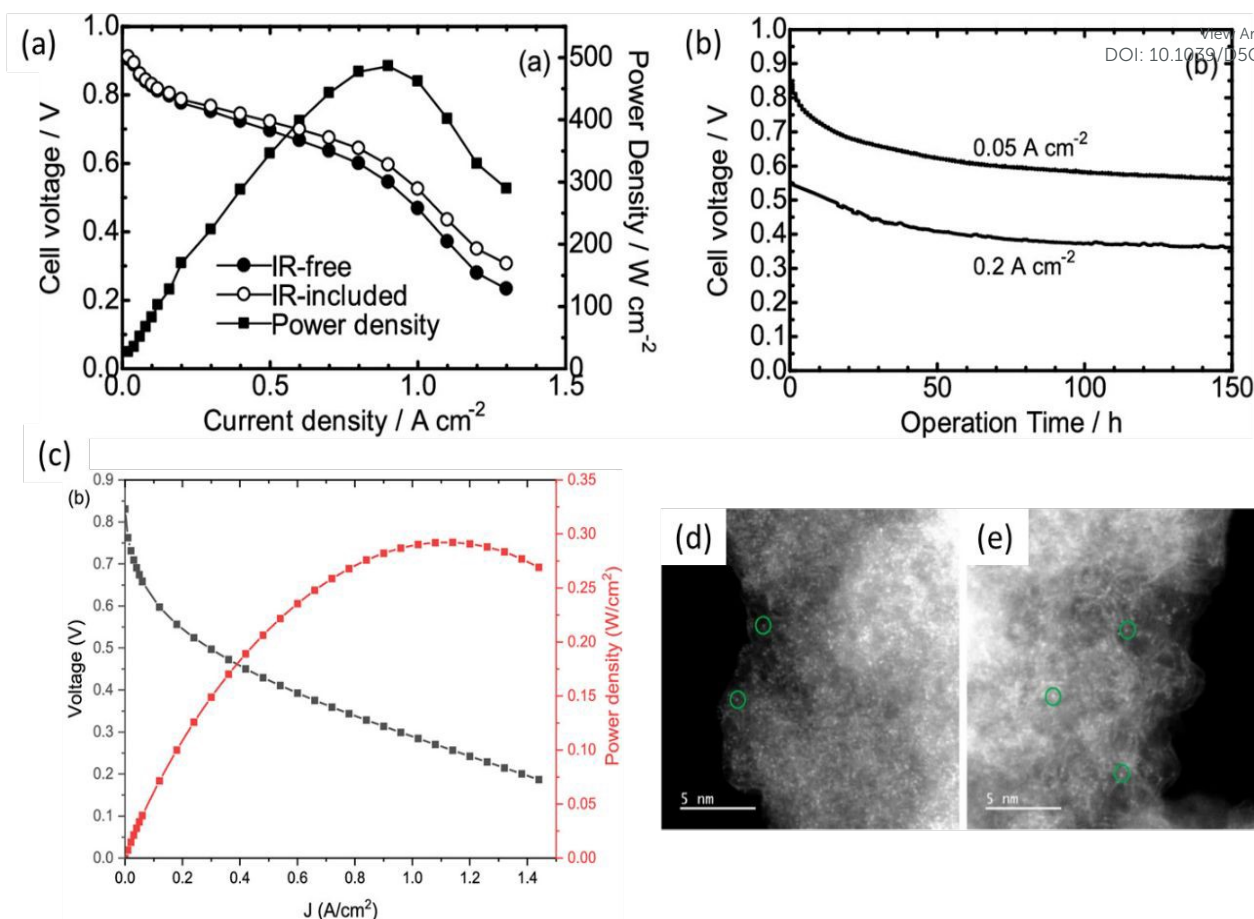


Fig. 10. (a) Polarization and power density curves, and (b) changes in the cell voltage during continuous operation with air at 0.2 and 0.05 A cm⁻² in PEMFC testing of Fe14-MR/FePi. Reproduced from ref. 130 under a Creative Commons Attribution Licence (CC-BY), copyright 2025 the authors. (c) H₂-O₂ PEMFC performance of the Fe800TEPc aerogel, (d,e) ADF-STEM images of pristine FeTEPc aerogel and heat-treated (at 800 °C) Fe800TEPc aerogel (the green circles show examples of atomically dispersed iron within the aerogel samples). Reproduced from ref. 139 with the permission from the American Chemical Society, copyright 2022.

4. Summary and prospects

Despite all these advances in the field of MN₄ macrocycle-derived non-precious metal catalysts for AEMFCs and PEMFCs, more efforts are needed to develop a strategy that maximizes their active site utilization. Additionally, the degradation of carbon supports brings the loss of FeN_x sites. Therefore, studying and optimizing the degree of graphitization can effectively restrain the corrosion of the carbon material, and adding this with computational studies can provide stability descriptors for M-N-C catalysts that can enhance their electrocatalytic performance. Apart from this, less emphasis has been given to *ex-situ* and *in-situ* electrochemical methods for active site quantification in the majority of Fe-N-C catalyst reports. This would enable comparison of turnover frequencies (TOFs) and formation of activity descriptors among reported catalysts.^{28,82} Additionally, the involvement of operando studies in Fe-N-C catalyzed electrocatalytic reactions could provide key insights into active site degradation mechanisms. Research effort should also be focused on exploring more sustainable precursors, such as biomass-based materials, as well as environmentally friendly and cost-effective synthesis routes, such as non-pyrolyzed methods. Lastly, the synthesis methods

should also focus on scalability parameters while also attaining state-of-the-art performance with minimal steps and negative environmental impact. Therefore, leaves big avenues for the development of cost-effective electrocatalysts for sluggish ORR in polymer electrolyte fuel cells.

Data Availability

No new primary research data were generated or included as part of this review article.

Conflicts of interest

There are no conflicts to declare.

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No new primary research data were generated or included as part of this review article.

