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# Unifying CO<sub>2</sub>-to-fuel and biomass valorization over a metal-free 2D carbon nitride-fullerene heterostructure: a solar-driven chemical circular economy†

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Photocatalytic CO<sub>2</sub> photoreduction integrated with biomass oxidation is highly attractive to produce fuels and fine chemicals. Herein, for the first time, we manifested CO<sub>2</sub> photoreduction to CO (>95% sel.) in synergy with biomass-based alcohol oxidation *via* a metal-free fullerene/2D carbon-nitride (C<sub>60</sub>/TUCN) semiconductor under solar simulated light. We observed that the composite, 5-C<sub>60</sub>/TUCN showed the highest CO<sub>2</sub> to CO (selectivity >95%) conversion efficiency along with lignin biomass model substrate *p*-methoxybenzyl alcohol (*p*-MBA) oxidation to *p*-methoxybenzaldehyde (*p*-MBAL) under solar simulated light with an excellent CO production rate of 8.92 mmol h<sup>-1</sup> g<sup>-1</sup> and *p*-MBAL production rate of 0.65 mmol h<sup>-1</sup> g<sup>-1</sup>. The apparent quantum yield (AQY) of CO evolution was determined to be 3.38% at λ = 450 nm. DFT calculations confirmed that the C<sub>60</sub> loading improved the activation and reduction of CO<sub>2</sub> to CO while considerably lowering the formation barrier of COOH\* intermediates. Besides, the accelerated separation and charge transfer kinetics of TUCN after C<sub>60</sub> modification was confirmed from EPR, PL, and photocurrent studies. <sup>13</sup>C<sub>2</sub>O<sub>2</sub> labeling experiments and EPR studies established the mechanistic pathway of the CO<sub>2</sub> reduction reaction. Thus, the current study showed an excellent proof-of-concept for upscaling CO<sub>2</sub> and biomass synergistically into solar fuels and fine chemicals, featuring a sustainable approach to boost the overall (bio)-chemical economy.

Photocatalytic carbon cycling is the most appealing and offers a green methodology to convert CO<sub>2</sub> to fuels and chemical feedstock by effective utilization of the abundant renewable solar light.<sup>1</sup> A carbon neutral cycle of coupling CO<sub>2</sub> reduction with H<sub>2</sub>O oxidation, mimicking natural photosynthesis, *i.e.* artificial photosynthesis (AP) is an ideal green process.<sup>2</sup> Over the past few decades AP has gained huge scientific interest and insights to convert CO<sub>2</sub> into the corresponding gaseous products (for example, CO, CH<sub>4</sub>, and C<sub>2</sub>H<sub>4</sub>) and liquid products (for

example, HCOOH, CH<sub>3</sub>OH, and C<sub>2</sub>H<sub>5</sub>OH) with the assistance of a wide range of photocatalysts including metal oxides,<sup>3,4</sup> chalcogenides,<sup>5-7</sup> organic semiconductors,<sup>8-10</sup> organic-inorganic heterostructures,<sup>11,12</sup> metal complexes,<sup>13</sup> quantum dots,<sup>14</sup> and chalcogenide-metal(oxide) heterostructures.<sup>15</sup>

However, the thermodynamically and kinetically sluggish H<sub>2</sub>O oxidation half-reaction could be replaced by an organic oxidation reaction to make use of photo-excited holes for purposeful oxidation processes. Photocatalytic CO<sub>2</sub> reduction integrated with organic transformation offers the most fascinating process that features effective utilization of photo-generated electrons and holes to meet the criteria of sustainability and chemical economy overwhelmingly.<sup>16</sup>

To date few photocatalysts including CdS-TiO<sub>2</sub>,<sup>17</sup> Cu<sub>2</sub>O/Cu,<sup>18</sup> Cu/TiO<sub>2</sub>,<sup>19</sup> Ag/TiO<sub>2</sub>,<sup>20</sup> CdSe/CdS QDs,<sup>21</sup> CdSQD/NC,<sup>22</sup> Pd/TiO<sub>2</sub>,<sup>23</sup> and CsPbBr<sub>3</sub>/Cs<sub>4</sub>PbBr<sub>6</sub> (ref. 24) have shown the successful coupling strategy for CO<sub>2</sub> photoreduction along with the organic oxidation process. However, the above mentioned photocatalytic systems encompass the use of metals, ultra-violet light absorption, toxicity and poor stability. To address these challenges the choice of photocatalysts must be highly anticipated.

Layered two-dimensional (2D) graphitic carbon nitride (g-C<sub>3</sub>N<sub>4</sub>),<sup>25</sup> a metal-free photocatalyst turned out to be a wonderful choice; in particular with a moderate band gap (2.7 eV) it has become a promising photocatalyst for CO<sub>2</sub> reduction. The improvements in the photocatalytic activity of bulk g-C<sub>3</sub>N<sub>4</sub> by means of facile charge transfer are highly desirable and could efficiently be promoted by making 2D layered g-C<sub>3</sub>N<sub>4</sub> structure and (non)-metal doping to it. Moreover, coupling 2D g-C<sub>3</sub>N<sub>4</sub> with other organic functional materials including fullerene (C<sub>60</sub>), carbon nanotubes (CNTs), and graphene offers an excellent methodology towards facile charge transfer and advancing the photocatalytic process, and is much more benign compared to a metal-free system.

The dual photoredox process for coupling selective CO<sub>2</sub> photoreduction with an organic oxidation process over 2D g-C<sub>3</sub>N<sub>4</sub> based semiconductors is still in its infancy. Recently,

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Neumann's group synthesized a three-component hetero-structure consisting of  $g\text{-C}_3\text{N}_4$  as the semiconductor, polyoxometalate (POM,  $\text{H}_3\text{PW}_{12}\text{O}_{40}$ ) functioning as an electron shuttle to promote electron-hole pair separation, and an Re complex as the  $\text{CO}_2$  reduction co-catalyst.<sup>26</sup> A cascade study was initiated by the dehydrogenation of cyclohexene to 1,3-cyclohexadiene ( $\text{C}_6\text{H}_8$ , two electron dehydrogenation product) and benzene ( $\text{C}_6\text{H}_6$ , four-electron dehydrogenation product) integrated with the reduction of  $\text{CO}_2$  to CO. Furthermore, Xue and co-workers prepared a biphasic photocatalyst, *i.e.* pyrene (Py) functionalized polymeric  $g\text{-C}_3\text{N}_4$  (Py-PCN)<sup>27</sup> for photocatalytic reduction of  $\text{CO}_2$  in the aqueous solution with simultaneous oxidation of alkene ( $\text{C}=\text{C}$ ) in the organic phase.

As an advancement towards the above discussed dual redox process, herein, for the first time we report a simple photocatalytic system in a metal-free fullerene/TUCN ( $\text{C}_{60}$ /TUCN; TUCN = thiourea derived carbon nitride) organic semiconductor to drive  $\text{CO}_2$  photoreduction process integrated with biomass alcohol oxidation under solar simulated light. Here, we utilized  $\text{C}_{60}$  as the co-catalyst<sup>28</sup> over 2D TUCN due to its excellent electron capture ability, which is beneficial for the  $\text{CO}_2$  reduction process and for achieving better charge separation ability. In our initial experiments, we found that amid all, 5 wt% of  $\text{C}_{60}$  on TUCN (5- $\text{C}_{60}$ /TUCN composite) showed efficient  $\text{CO}_2$  to CO conversion along with lignin biomass model substrate *para*-methoxybenzyl alcohol (*p*-MBA) oxidation to *para*-methoxybenzaldehyde (*p*-MBAL) under solar simulated light ( $100\text{ mW cm}^{-2}$ ) with an excellent rate of  $8.92\text{ mmol h}^{-1}\text{ g}^{-1}$  (CO selectivity of >95%) and  $0.65\text{ mmol h}^{-1}\text{ g}^{-1}$ , respectively.

Moreover, the state-of-the-art photocatalyst was compared with the existing literature having different photocatalytic systems; especially main emphasis was given to metal-free  $g\text{-C}_3\text{N}_4$  based photocatalysts as shown in Table S1.† The current strategy offers a new horizon for the dual photo-redox process to achieve an efficient  $\text{CO}_2$  reduction process to produce solar fuels and biomass oxidation to fine chemicals, thus featuring a sustainable approach to boost the overall (bio)-chemical economy.

Herein, Scheme S1† demonstrates the schematic illustration of the  $\text{C}_{60}$ /TUCN synthesis. Moreover, the detailed procedure for the  $\text{C}_{60}$ /TUCN heterostructure synthesis is provided in the ESI† under the Experimental section. It is pertinent to mention that based on superior photocatalytic performance, the 5- $\text{C}_{60}$ /TUCN composite was characterized and presented all throughout the manuscript.

The crystal structure of the layered TUCN and 5- $\text{C}_{60}$ /TUCN composite was analyzed by XRD (Fig. 1a). The XRD pattern of TUCN clearly showed a strong diffraction peak at  $27.3^\circ$  and weak diffraction peak at  $13.1^\circ$ . The diffraction peaks at  $13.1^\circ$  and  $27.3^\circ$  correspond to the (100) plane of the in-plane ordering in a tri-*s*-triazine sheet and the (002) plane of the inter-planar stacking of the conjugated tri-*s*-triazine sheets, respectively.<sup>29</sup> Compared with TUCN, the XRD pattern of 5- $\text{C}_{60}$ /TUCN exhibited diffraction peaks at  $10.7$ ,  $17.7$ ,  $20.7$ ,  $21.78$ ,  $30.8$ , and  $32.9$  corresponding to the (111), (220), (311), (222), (422), and (511) planes, respectively, which can be attributed to the cubic phase of  $\text{C}_{60}$ .<sup>30,31</sup> Moreover, the intensity and position of the peak at

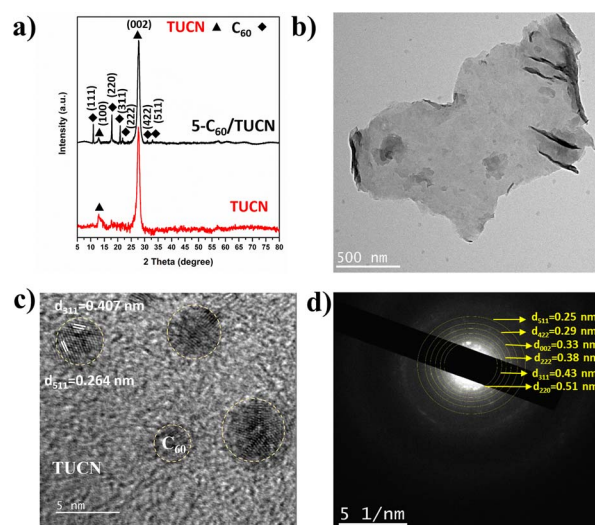


Fig. 1 (a) XRD patterns of TUCN and 5- $\text{C}_{60}$ /TUCN, and (b) TEM image, (c) HR-TEM image, and (d) TEM-SAED pattern of 5- $\text{C}_{60}$ /TUCN.

$27.3^\circ$  did not show any changes relative to that of bare TUCN, manifesting that immobilization of  $\text{C}_{60}$  on the layered structure of TUCN did not influence the crystal structure of TUCN.

The morphology of TUCN and the 5- $\text{C}_{60}$ /TUCN composite was investigated *via* Fe-SEM (Fig. S1†) and TEM analysis (Fig. 1b–d). As observed from Fig. 1b, 5- $\text{C}_{60}$ /TUCN showed a 2D sheet-like structure where  $\text{C}_{60}$  is immobilized on the TUCN sheets. Moreover, the high-resolution image of 5- $\text{C}_{60}$ /TUCN exhibited lattice fringes with a  $d$  spacing of  $0.407\text{ nm}$  and  $0.264\text{ nm}$  corresponding to the (311) and (511) planes of crystalline  $\text{C}_{60}$  (Fig. 1c), which was consistent with the XRD results. The crystallinity of 5- $\text{C}_{60}$ /TUCN was further proved by the TEM-SAED pattern (Fig. 1d). It was observed that small brighter spots of crystalline  $\text{C}_{60}$  were distributed over the surface of TUCN. Concentric rings of  $\text{C}_{60}$  planes (220), (311), (222), (422), and (511), were found to be closely surrounded by the TUCN (002) plane, revealing the formation of a composite with strong interfacial contact between  $\text{C}_{60}$  and TUCN. These results were found to be in good agreement with the XRD results.

The  $\text{N}_2$  adsorption–desorption isotherm analysis indicated that TUCN and 5- $\text{C}_{60}$ /TUCN displayed a type IV isotherm with narrow H3 hysteresis loops, indicating the development of a mesoporous structure in the form of parallel plates (Fig. S2a†).<sup>32,33</sup> Based on the BET analysis, the surface area of TUCN and 5- $\text{C}_{60}$ /TUCN was found to be identical *i.e.*  $28\text{ m}^2\text{ g}^{-1}$  and  $26\text{ m}^2\text{ g}^{-1}$ , respectively. This significantly demonstrated that the incorporation of  $\text{C}_{60}$  into the TUCN network did not affect the surface area, which indicated that it is not the main factor promoting the photocatalytic performance, while the enhanced light absorption and fast charge carrier transportation abilities helped to improve the photocatalytic activity.<sup>34</sup>

The surface compositions of the material were scrutinized *via* XPS analysis. The XPS survey spectrum of TUCN and 5- $\text{C}_{60}$ /TUCN showed that the surface of the material contains C, N, and O elements (Fig. S2b†). High resolution C 1s spectra of TUCN showed three peaks at  $288.4$ ,  $286.8$ , and  $284.8\text{ eV}$  which

corresponded to the existence of the N=C=N structure of tri-s-triazine, C-NH<sub>x</sub> groups, and adventitious carbon (C-C) used to calibrate the binding energies, respectively. The C 1s peaks of 5-C<sub>60</sub>/TUCN were found to be quite consistent with those of TUCN (Fig. S2c†), indicating that the existence of C<sub>60</sub> over the surface of TUCN and did not affect the C 1s electronic structure of TUCN. Fig. S2d† illustrates that the high resolution N 1s spectrum could be deconvoluted into four peaks. In TUCN, an intense peak at 398.5 eV showed the existence of an sp<sup>2</sup>-hybridized pyridinic nitrogen atom linked with a carbon atom (C=N-C), and the peaks at 399.8 eV and 400.8 eV are attributed to the tertiary N-(C)<sub>3</sub> and -NH<sub>x</sub> functional groups in the carbon nitride core, *i.e.* the tri-s-triazine motif, respectively. A weak N 1s peak at 404.1 eV was also observed which basically originated from π-excitation.<sup>35</sup> The N 1s peaks of 5-C<sub>60</sub>/TUCN were also found to be completely consistent with those of TUCN. It is evident that C<sub>60</sub> deposition over TUCN led to the enhancement in the carbon “C” content that was verified by using the XPS N/C ratio.<sup>30</sup> The decrease in the N/C ratio from TUCN (1.18) to 5-C<sub>60</sub>/TUCN (1.03) underpinned the high “C” content over the surface of 5-C<sub>60</sub>/TUCN (Table S2†). This explicitly showed the formation of a heterostructure between C<sub>60</sub> and TUCN. Moreover, the clear evidence of the enhanced “C” content over the surface of 5-C<sub>60</sub>/TUCN was further investigated by elemental CHNS analysis. The % of the elemental composition and N/C ratio are highlighted in Table S3.† The decreased N/C ratio from TUCN (1.96) to 5-C<sub>60</sub>/TUCN (1.65) significantly manifested the C<sub>60</sub> loading as “C” content in 5-C<sub>60</sub>/TUCN.

Of note, the optical and electronic properties of TUCN and 5-C<sub>60</sub>/TUCN were examined to overlay the difference in properties which is responsible for the photocatalytic CO<sub>2</sub> reduction and biomass alcohol oxidation. UV-vis diffuse reflectance absorption spectroscopy demonstrated broad and visible light absorption by TUCN and 5-C<sub>60</sub>/TUCN. As shown in Fig. 2a, bare

TUCN exhibited an absorption edge at 435 nm, which was assigned to the electronic transition from N 2p to C 2p orbitals of TUCN. On the other hand, the light absorption ability of 5-C<sub>60</sub>/TUCN in the visible region becomes more prominent, suggesting that C<sub>60</sub> loading enhanced the light absorption of TUCN.<sup>31</sup> Moreover, the Kubelka-Munk function equation was used to calculate the Tauc plot for the optical band gap ( $E_g$ ) energy estimation. The  $E_g$  of TUCN and C<sub>60</sub>/TUCN was found to be 2.81 eV and 2.73 eV, respectively (Fig. 2a: inset).

Mott-Schottky (MS) analysis was carried out to determine the conduction band (CB) position of TUCN and 5-C<sub>60</sub>/TUCN *via* flat-band potential ( $V_{fb}$ ) estimation (Fig. S2a and b†). The positive slope of the plots demonstrated the typical characteristics of n-type semiconductors. It is worth mentioning that  $V_{fb}$  generally lies 0.1 V below the CB of n-type semiconductors.<sup>35</sup> The  $V_{fb}$  of TUCN and 5-C<sub>60</sub>/TUCN was estimated to be -1.10 V and -1.20 V. On this account, the CB values of TUCN and 5-C<sub>60</sub>/TUCN were estimated to be -1.20 V and -1.30 V (*vs.* Ag/AgCl). On the NHE scale, the CB values were calculated to be -1.0 V and -1.10 V for TUCN and 5-C<sub>60</sub>/TUCN by using the  $E_{NHE} = E_{Ag/AgCl} + 0.197$  V equation. Moreover, the valence band (VB) positions of TUCN (+1.81 V) and 5-C<sub>60</sub>/TUCN (+1.63 V) were estimated by combining the obtained band gap from the Tauc plot with the relative CB positions from the  $E_{VB} = E_{CB} + E_g$  equation. The relative electronic band structures of TUCN and 5-C<sub>60</sub>/TUCN are shown in Fig. 2b.

The charge-transfer ability of TUCN and 5-C<sub>60</sub>/TUCN was investigated by using photoluminescence (PL), time-resolved photoluminescence (TRPL), transient photocurrent responses (TPR), and the Nyquist plot. The recorded PL spectra of 5-C<sub>60</sub>/TUCN demonstrated a prominent decrease in the PL intensity compared with that of TUCN, favoring a low rate of photo-generated charge carrier recombination or feasible charge transfer in 5-C<sub>60</sub>/TUCN (Fig. 2c). The TRPL spectra were collected to estimate the prominent interfacial transition of the photogenerated charge carriers and average lifetime ( $\tau$ ) of photogenerated charge carriers (Fig. 2d). Comparatively highest ( $\tau$ ) for 5-C<sub>60</sub>/TUCN (4.9 ns) to that of TUCN (4.3 ns) is observed, revealing that 5-C<sub>60</sub>/TUCN could furnish a longer time for photogenerated charge carriers to take part in the surface photoreaction, thereby amplifying its photocatalytic performance.<sup>36</sup> Moreover, enhanced interfacial charge carrier separation and migration was also measured by TPR study under five periodic on/off (30 s) light cycles (Fig. S3c†), which exhibited significant and repeatable photocurrent responses. Interfacial catalytic charge transfer kinetics of the material was examined with the help of the electrochemical impedance spectroscopy (EIS)-Nyquist plot (Fig. S3d†). A decrease in the semicircle arc radius of 5-C<sub>60</sub>/TUCN compared to that of TUCN manifested the decrease in facial charge transfer resistance. This promoted the charge transfer efficiency on the catalyst interface of C<sub>60</sub> and TUCN. Invariably, the above results manifested that existence of the 2D layered structure of 5-C<sub>60</sub>/TUCN promoted facile charge separation and migration.

Adsorption behavior of CO<sub>2</sub> over the surface of pthe hotocatalyst was recognized by using CO<sub>2</sub>-temperature-programmed desorption (CO<sub>2</sub>-TPD) analysis (Fig. S4†). Basically, the 5-C<sub>60</sub>/



Fig. 2 (a) Diffuse reflectance UV-vis spectra of TUCN and 5-C<sub>60</sub>/TUCN; inset (Tauc plot), (b) band energy diagram of TUCN and 5-C<sub>60</sub>/TUCN, (c) PL spectra of TUCN and 5-C<sub>60</sub>/TUCN, and (d) TRPL spectra of TUCN and 5-C<sub>60</sub>/TUCN.

TUCN surface nitrogen active sites ( $-\text{NH}_2$  dangling groups) act as electron donors (Lewis base) for  $\text{CO}_2$  molecules (Lewis acid). Therefore, a  $\text{CO}_2$ -TPD analysis was employed to quantify the desorbed  $\text{CO}_2$  molecules over the surface of 5- $\text{C}_{60}$ /TUCN upon heating, which is regarded as explicit evidence of the structural basicity of the TUCN core. It is evident that the 5- $\text{C}_{60}$ /TUCN sample demonstrated two broad desorption peaks from 90–130 °C and 220–245 °C, manifesting physical and chemical adsorption of  $\text{CO}_2$  molecules on weak basic sites of TUCN ( $-\text{NH}_2$  groups).<sup>22</sup>

Obviously, the large negative CB potential of 5- $\text{C}_{60}$ /TUCN could feasibly promote the reduction of  $\text{CO}_2$  to CO *via* the two electron reduction process.<sup>37</sup> Moreover, the VB potential of 5- $\text{C}_{60}$ /TUCN was utilized for *p*-MBA oxidation (lignin model substrate) to the corresponding aldehyde (*p*-MBAL) in synergy with CO production. The illustrative diagram of the photocatalytic reaction setup is shown in Scheme S2.† The photocatalytic performance of different photocatalysts under a solar simulator (AM 1.5G, 100  $\text{mW cm}^{-2}$ ) was tested and is summarized in Table 1.

It was observed that among all  $\text{C}_{60}$ /TUCN composites, the 5- $\text{C}_{60}$ /TUCN composite showed the highest dual photo redox activity for simultaneous  $\text{CO}_2$  reduction and *p*-MBA oxidation to produce CO and *p*-MBAL in a synergistic manner (Table 1, entry 1–4). In the case of 5- $\text{C}_{60}$ /TUCN, production rates of CO and *p*-MBAL oxidation rates were found to be 8.92 ( $\text{mmol h}^{-1} \text{g}^{-1}$ ) and 0.65 ( $\text{mmol h}^{-1} \text{g}^{-1}$ ), respectively. The selectivity for CO and *p*-MBAL was found to be >95% and >99%, respectively. Furthermore, the requisite conditions such as light, photocatalyst, and  $\text{C}_{60}$  co-catalyst were found to be indispensable parameters to favor the photocatalytic performance (Table 1, entry 4–7). Moreover the liquid products were identified by GC-MS analysis (Fig. S5†). Meanwhile, along with the major gaseous CO product (>95% sel.), comparatively, a very low amount of  $\text{CH}_4$ , and  $\text{H}_2$  was also generated which was confirmed by GC analysis

(Fig. S6†). The photocatalytic activity of 5- $\text{C}_{60}$ /TUCN was further tested for  $\text{CO}_2$  reduction to CO in synergy with different biomass alcohol oxidations under similar reaction conditions as listed in Table S4.†

The photocatalytic efficiency for CO production was evaluated in terms of the apparent quantum yield (AQY). The AQY of CO evaluation was estimated to be 3.38% at  $\lambda = 450 \text{ nm}$ .

It was shown that all the biomass based alcohols showed good photocatalytic activity under a  $\text{CO}_2$  atmosphere, thus promoting the  $\text{CO}_2$  reduction rate for selective production of CO. *p*-MBA and veratryl alcohol displayed almost similar and highest % conversion of the alcohols to the corresponding aldehydes, while HMF showed quite a sluggish % conversion among all the biomass based alcohols with only a conversion of 7% into the corresponding aldehyde *i.e.* diformylfuran (DFF).

Moreover, it should be noted that the moderate oxidation ability of the photogenerated holes ( $\text{h}^+$ ) over the TUCN surface can prevent the formation of non-selective  $\cdot\text{OH}$  radicals due to the less positive VB potential of 5- $\text{C}_{60}$ /TUCN ( $E_{\text{VB}} + 1.63 \text{ V}$ ). In addition, non-aqueous acetonitrile as a solvent also prevents the generation of  $\cdot\text{OH}$  *via*  $\text{h}^+$ . Therefore, high selectivity >99% of the oxidized products was observed as corroborated through liquid mixture analysis with the help of the GC-MS technique (Table S4†).

The recyclability/stability of 5- $\text{C}_{60}$ /TUCN was checked subsequently by performing five cyclic runs. The dual photo redox performance over 5- $\text{C}_{60}$ /TUCN indicated that the production rates of CO and *p*-MBAL were slightly decreased to 7.8  $\text{mmol h}^{-1} \text{g}^{-1}$  and 0.49  $\text{mmol h}^{-1} \text{g}^{-1}$ , respectively, after five runs (Fig. S7a†). Moreover, the selectivity of *p*-MBAL (>99%) remained constant throughout all cycling runs. On comparing the XRD and TEM analysis of 5- $\text{C}_{60}$ /TUCN before and after the cycle, no significant changes were observed in the crystal structure and surface morphologies of 5- $\text{C}_{60}$ /TUCN, respectively (Fig. S7b and c†).

Table 1 Photocatalytic coupling strategy of *p*-MBA oxidation in synergy with  $\text{CO}_2$  photoreduction to CO<sup>a</sup>

Entry	Catalyst	Light	Conv. <sup>b</sup> %	<i>p</i> -MBAL sel. <sup>b</sup> %	Production rate ( $\text{mmol h}^{-1} \text{g}^{-1}$ ) CO <i>p</i> -MBAL
1	2.5- $\text{C}_{60}$ /TUCN	✓	20	>99	3.24 0.33
2	5- $\text{C}_{60}$ /TUCN	✓	32	>99	9.80 0.65
3	7.5- $\text{C}_{60}$ /TUCN	✓	25	>99	7.61 0.42
4	10- $\text{C}_{60}$ /TUCN	✓	21	>99	4.80 0.35
5 <sup>c</sup>	5- $\text{C}_{60}$ /TUCN	×	N.A.	N.A.	N.A. N.A.
6 <sup>d</sup>	×	✓	N.A.	N.A.	N.A. N.A.
7 <sup>e</sup>	TUCN	✓	<5	>99	0.10 0.04

<sup>a</sup> Reaction conditions: catalyst (10 mg),  $\text{CO}_2$  purging for 30 min. *p*-MBA (0.1 mmol), 3 mL ACN, reaction time 6 h, room temperature (25 °C) and light source: solar simulator (AM 1.5G, 100  $\text{mW cm}^{-2}$ ). <sup>b</sup> The conversion and selectivity were obtained by GC-MS. <sup>c</sup> In the absence of light. <sup>d</sup> Without a photocatalyst. <sup>e</sup> In the absence of co-catalyst,  $\text{C}_{60}$ . N.A. – not applicable.

A plausible mechanism for dual photo redox performance of 5-C<sub>60</sub>/TUCN illustrating the coupling strategy of CO<sub>2</sub> reduction to CO with *p*-MBA oxidation by effective utilization of electrons and holes could be underlined by performing some control experiments (Table S5†). Under an argon (without CO<sub>2</sub>) atmosphere, a trace amount of H<sub>2</sub> and only 5% conversion of *p*-MBA to *p*-MBAL were observed (Table S5,† entry 1), implying that CO could be realized from CO<sub>2</sub> photoreduction. To corroborate this, an isotopic labeling experiment was performed by using a <sup>13</sup>CO<sub>2</sub> gas source instead of <sup>12</sup>CO<sub>2</sub> to authenticate the origin of the produced CO. Gas chromatography-mass spectrometry (GC-MS) analysis distinctly showed the observed signal of <sup>13</sup>CO with an *m/z* value of 29 (Fig. S8†) in the reaction system, manifesting that the detected CO originated from CO<sub>2</sub>.

Moreover, when an electron scavenger (AgNO<sub>3</sub>) was added to the reaction medium saturated with a CO<sub>2</sub> atmosphere, no gaseous CO species were observed, implying that the photo-generated electrons were utilized by AgNO<sub>3</sub> and were barely available for the CO<sub>2</sub> photoreduction (Table S5,† entry 2). Although in the case of IPA, the CO production rate was enhanced comparatively (good electron donor) but *p*-MBA observed mere oxidation/conversion to *p*-MBAL which could be ascribed to the excellent hole scavenging ability of IPA (Table S5,† entry 3).

On comparison, the more negative *V*<sub>fb</sub> value of 5-C<sub>60</sub>/TUCN (−1.10 V) than that of TUCN (−1.0 V), enables significant band bending that eventually facilitates the transport of photo-generated electrons efficiently to drive the CO<sub>2</sub> reduction reaction. In addition, the electron transfer mechanism involved in CO<sub>2</sub> photoreduction to CO over the surface of the 5-C<sub>60</sub>/TUCN photocatalyst was verified by electron paramagnetic resonance (EPR), demonstrating the excellent structure–activity relationship between the 5-C<sub>60</sub>/TUCN photocatalyst and CO<sub>2</sub> molecules. Under dark and inert conditions, no peak of the products was observed. Moreover, when the 5-C<sub>60</sub>/TUCN photocatalyst was exposed to visible light and an inert atmosphere, a new EPR signal appeared at *g* = 2.004, which was attributed to the surface −NH<sub>2</sub> amino groups. Interestingly, when the EPR spectra of the 5-C<sub>60</sub>/TUCN photocatalyst were recorded under a CO<sub>2</sub> atmosphere, a comparatively reduced EPR signal appeared at *g* = 2.004, manifesting a strong interaction between the adsorbed CO<sub>2</sub> species (Lewis acid) and the surface active sites (−NH<sub>2</sub> groups, Lewis base) of the 5-C<sub>60</sub>/TUCN photocatalyst (Fig. S9†).

This significantly entailed that the strong CO<sub>2</sub> adsorption over the surface of the 5-C<sub>60</sub>/TUCN photocatalyst was also confirmed *via* CO<sub>2</sub>-TPD analysis. The reduced EPR signal under light irradiation facilitated easy interfacial charge transfer from the photocatalyst, promoting the facile photoreduction process of the CO<sub>2</sub> molecule to CO.<sup>38,39</sup>

To further elucidate the interfacial charge transfer in the optimized 5-C<sub>60</sub>/TUCN heterostructure, the density of states (DOS) was computed, as shown in Fig. S10,† and the band gap of 5-C<sub>60</sub>/TUCN was calculated to be 2.02 eV, which was found to be lower than the band gap of TUCN (2.82 eV), as calculated by the GGA method. Notably, the band gap of TUCN (2.82 eV) matched well with the experimental result (2.81 eV); however, the band gap of 5-C<sub>60</sub>/TUCN (2.02 eV) was found to be lower

than the experimental result (2.73 eV). This might be due to the delocalization of the electron density with an inaccuracy of the HSE06 hybrid functional, causing underestimation of the band gap using this method.<sup>40</sup> The band gap reduction after the heterostructure formation between C<sub>60</sub> and TUCN resulted in facile charge carrier generation, and facilitates efficient charge separation and migration. Fig. S11† illustrates the optimized geometric structures of the TUCN monolayer, the C<sub>60</sub> molecule, and their heterostructure with a top view and side view, respectively.

Theoretical studies demonstrated that in the 5-C<sub>60</sub>/TUCN heterostructure, due to the weak van der Waals interaction of C<sub>60</sub> with 2D TUCN,<sup>41</sup> a strong bend occurred in the 2D layer of TUCN, implying a strong strain in the TUCN monolayer (Fig. S11c†). This may alter the atomic structure, leading to change in its electronic properties, which seemed to be favorable for facile charge transfer<sup>42</sup> which in turn, resulted in enhancement of the photocatalytic activity of the 5-C<sub>60</sub>/TUCN photocatalyst.

Moreover, the elementary steps of CO<sub>2</sub> photoreduction selectively to CO on TUCN and the 5-C<sub>60</sub>/TUCN surface were further revealed by DFT calculations. The energies directly obtained from DFT and calculated Gibbs free energies (Δ*G*) are listed in Tables S6 and S7.† As shown in Fig. 3a, the calculated Δ*G* of CO<sub>2</sub>-to-CO conversion involves the following steps: (i) first, the CO<sub>2</sub> molecules are adsorbed on the photocatalyst surface; (ii) generation of a COOH\* intermediate through a proton–electron pair transfer process on the CO<sub>2</sub> adsorbed surface. Since the Δ*G* for this step is uphill by 1.43 eV for TUCN and 1.19 eV for 5-C<sub>60</sub>/TUCN, this result significantly corroborated that the CO<sub>2</sub> reduction reaction catalyzed by 5-C<sub>60</sub>/TUCN is more favorable than that catalyzed by TUCN. Afterward, (iii) the reaction intermediate (COOH\*) further couples with a proton–electron pair to generate H<sub>2</sub>O and CO\* detached from the surface to generate the CO product which is characterized by the release of the energy.

Additionally, the photoreduction process of CO<sub>2</sub> for selective CO production over CH<sub>4</sub> and CH<sub>3</sub>OH could further be verified *via* the thermodynamic free energy (Δ*G*) barrier. As shown in Scheme 1, four reaction paths are proposed for CO<sub>2</sub> photoreduction on the surface of 5-C<sub>60</sub>/TUCN, where \* denotes the catalytically active sites. The transformation of COOH\* species into CO\* is easily realized compared to that of COOH\* into HCOOH\* (path-1) over 5-C<sub>60</sub>/TUCN due to the downhill free



Fig. 3 (a) Calculated free energy diagram corresponding to the reaction path followed by the CO<sub>2</sub> conversion on TUCN and 5-C<sub>60</sub>/TUCN and (b) key steps of CO<sub>2</sub> photoreduction to CO for the 5-C<sub>60</sub>/TUCN system.



Scheme 1 Proposed reaction paths for CO<sub>2</sub> reduction on the surface of 5-C<sub>60</sub>/TUCN.

energy profiles of COOH\* to CO\* as shown in Fig. 3b. The free energy change of CO\* to CO (path-2) is downhill by 0.13 eV while in contrast, the formation of CHO\* demands an input energy of 1.89 eV. This is the key selectivity step between path-2 and path-3 & 4. This result indicates that CO is a more favorable product of the CO<sub>2</sub> reduction reaction than CH<sub>4</sub> and CH<sub>3</sub>OH over the 5-C<sub>60</sub>/TUCN surface, which was elucidated in the experimental analysis described above.

In the synergistic photo-redox process, it is challenging to maintain the stoichiometry of the photoredox products. Herein, it was observed that the CO production rate was found to be multifold higher than the oxidation rate of the alcohols to the corresponding aldehydes. The multifold CO production rate over the *p*-MBAL production rate could be substantially induced by C<sub>60</sub>, as it acts as an excellent electron acceptor in the ground state and can accept, reversibly, up to six electrons.<sup>43</sup> Moreover, the moderate HOMO–LUMO energy gap (~1.5 eV) of C<sub>60</sub> separates a maximally degenerate and fully occupied HOMO (5-fold) from a degenerate LUMO (3-fold). Particularly, in the presence of a suitable electron-donor or reducing agent, the electron transfer and reduction reaction can be accelerated efficiently.<sup>44</sup> Therefore, in the present photo-redox process, excited state electron generation over C<sub>60</sub> coupled with surplus electron transfer from the CB of TUCN to C<sub>60</sub> resulted in a high CO production rate, in the presence of *p*-MBA as the electron-donor or reducing agent. Notably, the other possible reason is that the

large difference between e<sup>-</sup> and h<sup>+</sup> migration rates leads to the accumulation of photogenerated h<sup>+</sup> in the bulk and makes the h<sup>+</sup> short-lived.<sup>45,46</sup> Therefore, a comparatively low production rate of *p*-MBAL was observed during the photoredox process.

From the above results, we proposed the following mechanistic pathway as shown in Scheme 2. First, upon light excitation charge carriers (e<sup>-</sup>/h<sup>+</sup>) were generated over the photocatalyst surface. The e<sup>-</sup> got excited from the VB of TUCN to its CB, leaving h<sup>+</sup> at the VB of TUCN. The C<sub>60</sub> co-catalyst acted as an electron trapper, and e<sup>-</sup> from the CB of TUCN transferred immediately to C<sub>60</sub>. In contrast, the photogenerated h<sup>+</sup> over the surface of the photocatalyst was exploited feasibly to promote the oxidation of *p*-MBA to *p*-MBAL.

The discussion made above reinforced the excellent potency of the C<sub>60</sub> molecule towards the photoredox process and such an ability significantly provided a proof-of-concept for the effective utilization of e<sup>-</sup>/h<sup>+</sup> pairs synergistically for solar fuel generation and organic transformations over a metal-free surface.

## Conclusions

In summary, we successfully demonstrated the first ever coupling strategy of CO<sub>2</sub> reduction to CO with biomass based aromatic alcohol oxidation under solar simulated light irradiation (AM 1.5G, 100 mW cm<sup>-2</sup>) over the metal-free 5-C<sub>60</sub>/TUCN photocatalyst by effective utilization of e<sup>-</sup> and h<sup>+</sup>, respectively. Deposition of C<sub>60</sub> over 2D TUCN not only promoted an efficient charge transfer ability as observed in PL, TRPL and TPR studies, but also offered an excellent CO production rate (8.92 mmol h<sup>-1</sup> g<sup>-1</sup>) and *p*-MBA oxidation rate (0.65 mmol h<sup>-1</sup> g<sup>-1</sup>), synergistically. Maximum 3.38% AQY was achieved for CO evolution at λ = 450 nm. <sup>13</sup>CO<sub>2</sub> labelling experiments, EPR studies and DFT calculations established the mechanistic pathway of the CO<sub>2</sub> reduction reaction. This work thus provides an insight for the fabrication of metal-free organic semiconductors with carbon nitrides for solar energy utilization which in turn, achieves the sustainable dual photoredox process simultaneously for up scaling CO<sub>2</sub> to solar fuels coupled with biomass conversion.

## Conflicts of interest

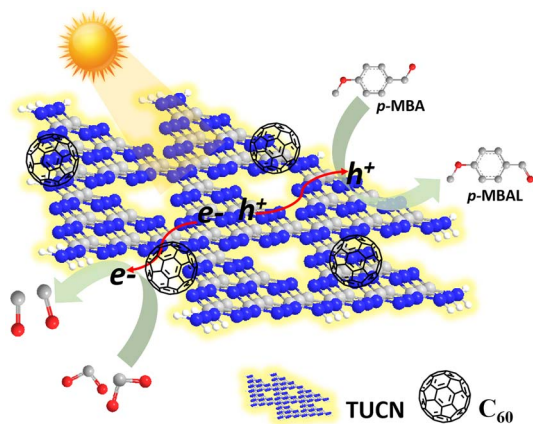
There are no conflicts to declare.

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## References

- 1 S. Yao, J. He, F. Gao, H. Wang, J. Lin, Y. Bai, J. Fang, F. Zhu, F. Huang and M. Wang, *J. Mater. Chem. A*, 2023, **11**, 12539–12558.
- 2 S. Yoshino, T. Takayama, Y. Yamaguchi, A. Iwase and A. Kudo, *Acc. Chem. Res.*, 2022, **55**, 966–977.



Scheme 2 Plausible mechanistic illustration of dual photoredox CO<sub>2</sub> reduction to CO coupled with *p*-MBA oxidation over the 5-C<sub>60</sub>/TUCN photocatalyst.

- 3 B. Pan, J. Qin and C. Wang, ed. X. Wang, M. Anpo, X. Fu, *Current Developments in Photocatalysis and Photocatalytic Materials New Horizons in Photocatalysis*, Elsevier, 2020, pp. 77–87.
- 4 Y. Cui, X. Lian, L. Xu, M. Chen, B. Yang, C. Wu, W. Li, B. Huang and X. Hu, *Materials*, 2019, **12**(2), 276.
- 5 S. Adabala and D. P. Dutta, *J. Environ. Chem. Eng.*, 2022, **10**, 107763.
- 6 M. T. Leng Lai, C. W. Lai and J. C. Juan, in *Micro and Nano Technologies*, ed. M. M. B. T.-C.-B. N. as P. Khan, Elsevier, 2021, pp. 295–306.
- 7 Y. Wang, B. Ren, J. Zhen Ou, K. Xu, C. Yang, Y. Li and H. Zhang, *Sci. Bull.*, 2021, **66**, 1228–1252.
- 8 H. L. Nguyen and A. Alzamy, *ACS Catal.*, 2021, **11**, 9809–9824.
- 9 Y. Zhang, H. Liu, F. Gao, X. Tan, Y. Cai, B. Hu, Q. Huang, M. Fang and X. Wang, *EnergyChem*, 2022, **4**, 100078.
- 10 W. Zhan, H. Gao, Y. Yang, X. Li and Q.-L. Zhu, *Adv. Energy Sustainability Res.*, 2022, **3**, 2200004.
- 11 P. K. Prajapati, N. Saini, D. K. Chauhan and K. Kailasam, *J. Mater. Chem. A*, 2023, **11**, 385–400.
- 12 N. Ahmad, C.-F. Jeffrey Kuo, M. Mustaqeem, M. K. Hussien and K.-H. Chen, *Mater. Today Phys.*, 2023, **31**, 100965.
- 13 H. Kumagai, Y. Tamaki and O. Ishitani, *Acc. Chem. Res.*, 2022, **55**, 978–990.
- 14 H.-L. Wu, X.-B. Li, C.-H. Tung and L.-Z. Wu, *Adv. Mater.*, 2019, **31**, 1900709.
- 15 X. Li, Z. Wang, J. Zhang, K. Dai, K. Fan and G. Dawson, *Mater. Today Phys.*, 2022, **26**, 100729.
- 16 L. Yuan, M.-Y. Qi, Z.-R. Tang and Y.-J. Xu, *Angew. Chem., Int. Ed.*, 2021, **60**, 21150–21172.
- 17 G. Song, F. Xin, J. Chen and X. Yin, *Appl. Catal., A*, 2014, **473**, 90–95.
- 18 Y. Chen, M. Wang, Y. Ma, Y. Li, J. Cai and Z. Li, *Catal. Sci. Technol.*, 2018, **8**, 2218–2223.
- 19 T. Yang, Q. Yu and H. Wang, *Catal. Lett.*, 2018, **148**, 2382–2390.
- 20 L. Wang, X. Zhang, L. Yang, C. Wang and H. Wang, *Catal. Sci. Technol.*, 2015, **5**, 4800–4805.
- 21 Q. Guo, F. Liang, X.-B. Li, Y.-J. Gao, M.-Y. Huang, Y. Wang, S.-G. Xia, X.-Y. Gao, Q.-C. Gan, Z.-S. Lin, C.-H. Tung and L.-Z. Wu, *Chem*, 2019, **5**, 2605–2616.
- 22 F. Wang, T. Hou, X. Zhao, W. Yao, R. Fang, K. Shen and Y. Li, *Adv. Mater.*, 2021, **33**, 2102690.
- 23 R. Zhang, H. Wang, S. Tang, C. Liu, F. Dong, H. Yue and B. Liang, *ACS Catal.*, 2018, **8**, 9280–9286.
- 24 G.-X. Dong, W. Zhang, Y.-F. Mu, K. Su, M. Zhang and T.-B. Lu, *Chem. Commun.*, 2020, **56**, 4664–4667.
- 25 X. Wang, K. Maeda, A. Thomas, K. Takane, G. Xin, J. M. Carlsson, K. Domen and M. Antonietti, *Nat. Mater.*, 2009, **8**, 76–80.
- 26 H. Yu, E. Haviv and R. Neumann, *Angew. Chem., Int. Ed.*, 2020, **59**, 6219–6223.
- 27 X. Gong, S. Yu, M. Guan, X. Zhu and C. Xue, *J. Mater. Chem. A*, 2019, **7**, 7373–7379.
- 28 M. Chen, R. Guan and S. Yang, *Adv. Sci.*, 2019, **6**, 1800941.
- 29 S. Cao, J. Low, J. Yu and M. Jaroniec, *Adv. Mater.*, 2015, **27**, 2150–2176.
- 30 L. Song, T. Li and S. Zhang, *J. Phys. Chem. C*, 2017, **121**, 293–299.
- 31 X. Chen, K. Deng, P. Zhou and Z. Zhang, *ChemSusChem*, 2018, **11**, 2444–2452.
- 32 M. O. Fuentez-Torres, F. Ortiz-Chi, C. G. Espinosa-González, M. Aleman, A. Cervantes-Urbe, J. G. Torres-Torres, M. K. Kesarla, V. Collins-Martínez, S. Godavarthi and L. Martínez-Gómez, *Top. Catal.*, 2021, **64**, 65–72.
- 33 X. Song, Q. Yang, M. Yin, D. Tang and L. Zhou, *RSC Adv.*, 2018, **8**, 7260–7268.
- 34 J. Cheng, C. Li, Z. Yu and H. Liu, *J. Colloid Interface Sci.*, 2023, **629**, 739–749.
- 35 D. K. Chauhan, V. R. Battula, A. Giri, A. Patra and K. Kailasam, *Catal. Sci. Technol.*, 2022, **12**, 144–153.
- 36 P. Sharma, S. Kumar, O. Tomanec, M. Petr, J. Zhu Chen, J. T. Miller, R. S. Varma, M. B. Gawande and R. Zbořil, *Small*, 2021, **17**, 2006478.
- 37 J. Ran, M. Jaroniec and S.-Z. Qiao, *Adv. Mater.*, 2018, **30**, 1704649.
- 38 D. Sun, Y. Fu, W. Liu, L. Ye, D. Wang, L. Yang, X. Fu and Z. Li, *Chem. – Eur. J.*, 2013, **19**, 14279–14285.
- 39 W. Tu, Y. Yang, C. Chen, T. Zhou, T. Li, H. Wang, S. Wu, Y. Zhou, D. O'Hare, Z. Zou and R. Xu, *Small Struct.*, 2023, **4**, 2200233.
- 40 J. Xue, M. Fujitsuka and T. Majima, *ACS Appl. Mater. Interfaces*, 2019, **11**, 40860–40867.
- 41 Q. Li, L. Xu, K.-W. Luo, W.-Q. Huang, L.-L. Wang, X.-F. Li, G.-F. Huang and Y.-B. Yu, *Phys. Chem. Chem. Phys.*, 2016, **18**, 33094–33102.
- 42 G. Gao, Y. Jiao, F. Ma, Y. Jiao, E. Waclawik and A. Du, *J. Catal.*, 2015, **332**, 149–155.
- 43 D. M. Guldi and M. Prato, *Acc. Chem. Res.*, 2000, **33**, 695–703.
- 44 L. Brunet, D. Y. Lyon, E. M. Hotze, P. J. J. Alvarez and M. R. Wiesner, *Environ. Sci. Technol.*, 2009, **43**, 4355–4360.
- 45 X. Zhang, P. Wang, X. Lv, X. Niu, X. Lin, S. Zhong, D. Wang, H. Lin, J. Chen and S. Bai, *ACS Catal.*, 2022, **12**, 2569–2580.
- 46 Q. Liu, S. Wang, W. Mo, Y. Zheng, Y. Xu, G. Yang, S. Zhong, J. Ma, D. Liu and S. Bai, *Small*, 2022, **18**, 2104681.