

Cite this: *Chem. Sci.*, 2022, 13, 11577

All publication charges for this article have been paid for by the Royal Society of Chemistry

Stereomutation and chiroptical bias in the kinetically controlled supramolecular polymerization of cyano-luminogens†

Lucía López-Gandul,^a Cristina Naranjo,^a Cecilia Sánchez,^a Rafael Rodríguez,^b Rafael Gómez,^a Jeanne Crassous^{id}*^b and Luis Sánchez^{id}*^a

The synthesis of two pairs of enantiomeric cyano-luminogens **1** and **2**, in which the central chromophore is a *p*-phenylene or a 2,5-dithienylbenzene moiety, respectively, is described and their supramolecular polymerization under kinetic and thermodynamic control investigated. Compounds **1** and **2** form supramolecular polymers by quadruple H-bonding arrays between the amide groups and the π -stacking of the central aromatic moieties. In addition, the peripheral benzamide units are able to form intramolecularly H-bonded pseudocycles that behave as metastable monomer **M*** thus affording kinetically and thermodynamically controlled aggregated species AggI and AggII. The chiroptical and emissive features of compounds **1** and **2** strongly depend on the aggregation state and the nature of the central aromatic unit. Compounds **1** exhibit a bisignated dichroic response of different intensity but with similar sign for both AggI₁ and AggII₁ species, which suggests the formation of helical aggregates. In fact, these helical supramolecular polymers can be visualized by AFM imaging. Furthermore, both AggI and AggII species formed by the self-assembly of compounds **1** show CPL (circularly polarized light) activity of opposite sign depending on the aggregation state. Thienyl-derivatives **2** display dissimilar chiroptical, morphological and emissive characteristics for the corresponding kinetically and thermodynamically controlled aggregated species AggI and AggII in comparison to those registered for compounds **1**. Thus, a stereomutation phenomenon is observed in the AggI₂ $\rightarrow$ AggII₂ conversion. In addition, AggI₂ is arranged into nanoparticles that evolve to helical aggregates to afford AggII₂. The dissimilar chiroptical and morphological features of AggI₂ and AggII₂ are also appreciated in the emissive properties. Thus, whilst AggI₂ experiences a clear AIE (aggregation induced emission) process and CPL activity, the thermodynamically controlled AggII₂ undergoes an ACQ (aggregation caused quenching) process in which the CPL activity is cancelled.

Received 20th June 2022
Accepted 6th September 2022

DOI: 10.1039/d2sc03449b

rsc.li/chemical-science

Introduction

Self-assembly of suitable building blocks in an organized fashion is an attractive approach to achieve functional materials holding applicability in a number of research areas like biomedicine, photocatalysis, adhesives or energy conversion.¹ All these applications rely on the organized arrangement of the constitutive units to finally yield the functional aggregate. In this regard, supramolecular polymerization has emerged as a useful strategy to achieve the challenging goal of attaining

complex, functional materials.² Especially relevant in the field of supramolecular polymers is the incorporation of asymmetry elements in the self-assembling units. The supramolecular polymerization of these chiral monomeric species results in the efficient transfer of asymmetry from the monomers to the aggregated species that finally yield helical, and therefore chiral aggregates.³ These helical aggregates have been exploited to perform detailed studies on transfer and amplification of asymmetry,⁴ as well as in the preparation of optical and electronic devices.⁵ It is worth mentioning that a vast majority of these applications of functional supramolecular polymers are studied under thermodynamic control. However, kinetically controlled chemical processes are common in synthetic protocols, material science and in biological phenomena and have been also reported for supramolecular polymers.⁶ Thus, kinetically controlled supramolecular polymerization has derived in differentiation events,⁷ living and seeded processes,⁷ and also in stereomutation phenomena.^{7a,9} Importantly, a key issue on the characterization and study of the thermodynamic or kinetic

^aDepartamento de Química Orgánica, Facultad de Ciencias Químicas, Universidad Complutense de Madrid, 28040 Madrid, Spain. E-mail: lusamar@ucm.es

^bUniv Rennes, CNRS, ISCR (Institut des Sciences Chimiques de Rennes) – UMR 6226, F-35000, Rennes, France. E-mail: jeanne.crassous@univ-rennes1.fr

† Electronic supplementary information (ESI) available: These might include comments relevant not central to the matter under discussion, limited experimental and spectral data, and crystallographic data. See <https://doi.org/10.1039/d2sc03449b>

Fig. 1 Chemical structures of the described chiral cyano-luminogens (central part) and schematic representation of the CPL-emitting kinetic and thermodynamically controlled aggregated structures formed by **1** and **2**.

Results and discussion

Synthesis and self-assembly in solution

The synthesis of the cyano-luminogens **1** and **2** was accomplished by following a similar protocol to that carried out for the achiral congeners.¹⁴ Thus, the alkylation reaction of methyl 3,4,5-trihydroxybenzoate with the chiral (*S* or *R*)-3,7-dimethyloctyl 4-methylbenzenesulfonate, subsequent basic hydrolysis of the ester group and amidation reaction with ethane-1,2-diamine, yields the monoamides **3**. These monoamides are reacted with 4-(cyanomethyl)benzoic acid **4** to yield the corresponding cyanomethyl derivatives **5**. Finally, a double Knoevenagel reaction of these cyano compounds **5** with terephthalaldehyde or thiophene-2,5-dicarbaldehyde readily yields the chiral, cyano-luminogens **1** and **2** (Scheme S1†). All the new compounds have been characterized by utilizing the usual spectroscopic techniques (see, ESI†).

The self-assembly of compounds **1** and **2** has been investigated by utilizing a combination of spectroscopic techniques and taking into account that the supramolecular polymerization of the achiral congeners of the reported cyano-derivatives has been described as a cooperative process yielding highly stable, fibrillar aggregates.¹⁴ Thus, we have firstly registered the UV-vis spectra of (**S**)-**1** and (**R**)-**2** in a good solvent (CHCl₃) and in a bad solvent (methylcyclohexane, MCH). As expected, the absorption pattern of these chiral compounds shows maxima at $\lambda = 371$ and 411 nm for (**S**)-**1** and (**R**)-**2**, respectively, in CHCl₃. A clear hypsochromic effect is observed in MCH with maxima at $\lambda = 338$ and 386 nm for (**S**)-**1** and (**R**)-**2**, respectively, that suggests the formation of H-type aggregates and in good agreement with that reported for the achiral congeners (Fig. S1†).^{14,15} The formation of these H-type aggregates could be justified by considering the non-covalent interaction of the self-assembling units by the π -stacking of the aromatic units and the formation of H-bonding arrays between the four amide groups. In fact, concentration dependent ¹H-NMR spectra for (**S**)-**1** and (**R**)-**2** in CDCl₃ as solvent show the shielding of the corresponding

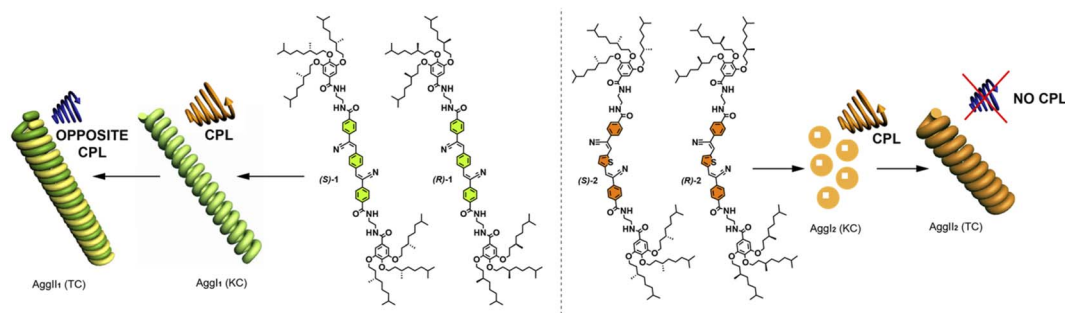


Fig. 1 Chemical structures of the described chiral cyano-luminogens (central part) and schematic representation of the CPL-emitting kinetic and thermodynamically controlled aggregated structures formed by **1** and **2**.

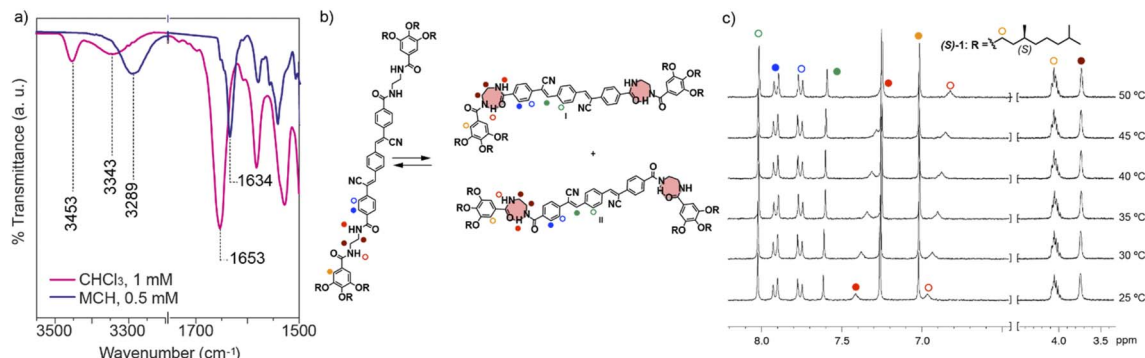


Fig. 2 (a) Partial FTIR spectra of (R)-1 in CHCl_3 and MCH; (b) schematic illustration of the equilibrium between the two monomeric species, the totally unbound and the intramolecularly H-bonded, of cyano-luminogens 1; (c) partial ^1H -NMR spectra of (S)-1 at different temperatures (300 MHz, CDCl_3 , $c_T = 1 \text{ mM}$).

resonances upon increasing the concentration and the downfield shift of the amide protons, that reinforces this hypothesis (Fig. S2†). To further demonstrate the formation of these H-bonding arrays, we have also registered FTIR spectra in these good and bad solvents (CHCl_3 and MCH, respectively). In the latter, the bands corresponding to the N–H stretching and amide I appear at ~ 3290 and $\sim 1634 \text{ cm}^{-1}$, corroborating the formation of the intermolecular H-bonds between the amide functional groups (Fig. 2a and S3†).¹⁶ In the former, however, two N–H stretching bands are observed at ~ 3453 and $\sim 3343 \text{ cm}^{-1}$ due to the presence of free monomeric species and also metastable monomeric species $\mathbf{M}^*$ formed by the intramolecular H-bonding of the inner or outer carbonyls of the amides and their outer and inner N–H (Fig. 2 and S3†).¹⁶

The formation of the intramolecularly H-bonded 7-membered pseudocycles has been corroborated by variable temperature (VT) ^1H -NMR spectroscopy in CDCl_3 . In these experiments, all the resonances observed in the spectra experience no shift upon heating the sample. This effect, however, is the contrary for those resonances corresponding to the inner and outer amides (filled and hollow red circles in Fig. 2c and S3c†). These resonances shift upfield upon heating the sample due to the rupture of the two possible intramolecular H-bonds between the amide functional groups of each peripheral side chain. The formation of the possible $\mathbf{M}^*$ can be justified by considering that both amides shield upon heating (Fig. 2c and S3c†).

The formation of these metastable monomeric species $\mathbf{M}^*$ has been reported to have a strong influence on the kinetics of the supramolecular polymerization of those self-assembling units endowed with the peripheral *N*-(2-benzamidoethyl)-3,4,5-trialkoxylbenzamide moieties.^{7c,8b,17} In fact, Fernández and coworkers have very recently demonstrated the ability of self-assembling units decorated with these peripheral moieties to generate supramolecular polymers with intermolecular H-bonding interactions in a consecutive manner, the first step being the formation of dimers in which intra- and intermolecular H-bonding interactions between the amide groups are operative.¹⁸ Consequently, we have not used the commonly utilized variable temperature spectroscopic (UV-vis or CD)

techniques to elucidate the supramolecular polymerization mechanism, but we have utilized denaturation experiments to derive, if possible, the thermodynamic parameters associated to the supramolecular polymerization of the reported cyano-luminogens (*vide infra*).

Chiroptical properties and morphology of the aggregated species of the cyano-luminogens 1 and 2

To elucidate the chiroptical properties of the cyano-luminogens 1 and 2, we have firstly carried out CD measurements in MCH as solvent. The high stability of the aggregated species of these compounds in MCH has been demonstrated by registering CD spectra at a total concentration, c_T , as low as $0.5 \mu\text{M}$. Only at these highly diluted conditions, the CD response is cancelled at high temperatures (Fig. S4†). Both compounds 1 and 2 present clear bisignated CD spectra, with maxima at $\lambda = 348$ and 325 nm and a zero-crossing point at $\lambda = 335 \text{ nm}$ for 1 and $\lambda = 396$ and 371 nm and a zero-crossing point at $\lambda = 379 \text{ nm}$ for 2, that suggest the formation of helical aggregates of opposite handedness for each enantiomer (Fig. S5†).

The high stability of the supramolecular polymers formed by cyano-derivatives 1 and 2 prompted us to utilize the denaturation (SD) model to achieve a detailed investigation of the supramolecular polymerization mechanism. In this SD model, diluted solutions at equal c_T of the investigated self-assembling units in a good (CHCl_3) and in a bad (MCH) solvent, are mixed together to achieve a complete binding isotherm.¹⁹ During the application of the denaturation model, we have noticed an uncommon pattern in the corresponding CD spectra of the mixtures that follow a dissimilar trend for compounds 1 and 2. In the case of *p*-phenylenevinylenes 1, the dichroic response of freshly prepared solutions in pristine MCH is less intense than that registered for MCH/ CHCl_3 mixtures with a low ratio of the good solvent. The dichroic response gradually increases until the MCH/ CHCl_3 mixtures reaches around a 25% of the good solvent, the dichroic sign being the same in all the experimental conditions (Fig. 3a and S6†). Increasing the percentage of CHCl_3 gradually decreases the intensity of the dichroic signal, that is completely cancelled when the ratio of the good solvent is higher than 50% (Fig. 3a and S6†). The corresponding UV-vis





Fig. 3 CD (a and c) and UV-vis (b and d) spectra of luminogen (S)-1 in MCH/CHCl₃ mixtures freshly prepared (a and b) and upon aging for 24 h (c and d) (20 °C; $c_T = 10 \mu\text{M}$). Arrows indicate the changes upon adding the solutions of 1 in CHCl₃ (e and f) AFM images of the kinetically controlled AggI (e) and thermodynamically controlled AggII (f) of compound (S)-1 (MCH/CHCl₃ 9/1 as solvent; HOPG as surface; $c_T = 10 \mu\text{M}$).

spectra at these conditions show the formation of three different species: the monomeric species (red line in Fig. 3b) and two different aggregated species (green and blue lines in Fig. 3b). Interestingly, aging these solutions for 24 h affords a gradual decreasing of the dichroic response upon adding increasing amounts of CHCl₃ (Fig. 3c and S6†). In addition, the corresponding UV-vis spectra show only the absorption pattern ascribable to the monomeric (red line in Fig. 3d) and aggregated (green and blue lines in Fig. 3d) species. This behaviour is diagnostic of the operation of a kinetically controlled supramolecular polymerization that, in certain conditions of solvent composition and time, yields an aggregated species AggI that evolves with time to a different aggregated species AggII (Fig. 1).

Upon reaching the thermodynamic equilibrium, it is possible to apply the SD model to derive the thermodynamic parameters. Thus, plotting the degree of aggregation α , calculated from the variation of the intensity of the dichroic response, observed from the corresponding CD spectra versus the molar fraction of CHCl₃ affords a non-sigmoidal curve that implies a cooperative supramolecular polymerization with a relatively large degree of cooperativity σ and a large value of m , that expresses the dependence of ΔG on the solvent (Fig. 3a and S7a† and Table 1). Similar findings have been extracted from the SD model applied to the UV-vis spectra (Fig. 3b and S7b†).

To visualize the morphologic differences of AggI and AggII, we have registered the corresponding AFM images. Thus, isolated fibrillar aggregates of several micrometers of length and typical height of $\sim 3 \text{ nm}$ are visualized by spin-coating a freshly prepared solution of (S)-1 in a 9/1 MCH/CHCl₃ mixture at $c_T =$

$10 \mu\text{M}$ (Fig. 3e and S8†). In the case of AggII, the morphology visualized by AFM imaging demonstrates the bundling effect experienced by the initially formed fibers, corresponding to AggI. Thus, AggII appears as thicker fibers of several micrometers of length and typical height of $\sim 15 \text{ nm}$ that would imply a helix to super-helix transition to afford chiral structures (Fig. 3f and S9†).²⁰ In both cases, the rope-like morphology of the fibrillar aggregates corroborates the helical character of both AggI and AggII formed by *p*-phenylenevinyls 1.

Upon investigating the dynamic character of the supramolecular polymerization of cyano-luminogens 1, we have also investigated the kinetically-controlled self-assembly of the thienyl-based luminogens 2. In good analogy to that observed for compounds 1, the denaturation experiments performed with 2 evidence drastic spectroscopic changes, especially in the dichroic response. Noteworthy, a first dichroic characteristic of compounds 2 in comparison to 1 is that the CD response is much less intense in pristine MCH (Fig. S5†). Initially, the addition of increasing amounts of 2 in the good solvent CHCl₃ to a solution in the bad solvent MCH provokes a slight increase

Table 1 Thermodynamic parameters for the supramolecular polymerization of (S)-1 and (R)-2

Comp.	ΔG° ^a	m	σ
(S)-1 ^b	-39.9 ± 1	35.6	1.1×10^{-4}
(R)-2 ^c	-37.9 ± 3	31.7	4.5×10^{-3}

^a In kJ mol^{-1} . ^b From CD measurements. ^c From UV-vis measurements.

in intensity of the CD response. However, upon adding around 30% of solutions of **2** in CHCl_3 , a complete inversion of the dichroic response is clearly observed (Fig. 4 and S10†). This stereomutation, observed for both enantiomers (*S*)-**2** and (*R*)-**2**, could be ascribable to the kinetically controlled self-assembly of these cyano-luminogens (Fig. S11†).⁹ In fact, the dichroic response of compounds **2** experiences a signal inversion concomitant with a strong increase in intensity only after 30 min upon mixing both solutions (Fig. S11a and S11b†) or after heating the solutions up to 60 °C, to favour the complete disassembly, and cooling them down to 10 °C to induce the reassembly (Fig. S10b and S11c†). In all the samples with relatively low percentage of CHCl_3 (10–30%), the intensity of the dichroic response drastically increases in comparison to that registered for pristine MCH. The evolution of the dichroic response *versus* time is fairly visualized by registering the CD spectra of (*S*)-**2** and (*R*)-**2** in a 7/3 MCH/ CHCl_3 mixture that reaches the maximum intensity upon aging the sample for ~30 min (Fig. S13†). As expected, the addition of higher amounts of chloroform solutions of **2** provokes the gradual disassembly of the aggregated species (Fig. S12†). Unfortunately, the way variation of the dichroic response *versus* the molar fraction of CHCl_3 impedes to utilize these data for deriving the corresponding thermodynamic parameters by applying the SD model. However, it is possible to apply the SD model to the denaturation experiments performed upon 24 h by using the corresponding UV-vis spectra. In good accord to compounds **1**, the UV-vis spectra of the freshly prepared solutions show the presence of three different species: the monomers (red line in Fig. 4b and d), AggI, that is not detected upon

24 h (green line in Fig. 4b) and AggII (blue line in Fig. 4b and d). The hypsochromic effect observed in the absorption spectra upon increasing the fraction of the bad solvent MCH suggests the formation, as in the case of compounds **1**, of H-type aggregates. Plotting the variation of the degree of aggregation α *versus* the molar fraction of CHCl_3 displays a non-sigmoidal curve that can be fitted to the SD model (Fig. S14†). The corresponding thermodynamic parameters are collected in Table 1. In this case, the $\Delta G'$ value is higher to that derived for compound (*S*)-**1**, but the degree of cooperativity, σ , is higher.

We have also visualized the morphology of the aggregated species generated from compounds **2** by AFM imaging onto HOPG as surface. Unlike the fibrillar nature of AggI formed by compounds **1**, the AFM images of AggI of compound (*R*)-**2** exhibits the formation of a large number of nanoparticles (Fig. 4e and S15†) that evolve with time to form long fibrillar aggregates ascribable to AggII species (Fig. 4f and S16†). These fibrillar aggregates present several micrometers length and heights of ~2.5 and ~5 nm, the latter implying the intertwining effect also observed in the AggII species from phenylenevinyls **1**.

The spectroscopic and morphologic differences found for AggI and AggII species of luminogens **1** and **2** could be justified by considering the above-mentioned ability of the metastable monomeric species M^* , in which the intramolecularly H-bonded pseudocycle is able to form intermolecular H-bonding interactions, to generate the corresponding AggI species.¹⁸ These AggI species would be characterized by the intermediate spectra observed in the freshly prepared mixtures (green lines in Fig. 3b and 4b). Furthermore, the presence of two pseudocycles



Fig. 4 CD (a and c) and UV-vis (b and d) spectra of luminogen (*R*)-**2** in MCH/ CHCl_3 mixtures freshly prepared (a and b) and upon aging for 24 h (c and d) (20 °C; $c_T = 10 \mu\text{M}$). Arrows indicate the changes upon adding the solutions of **2** in CHCl_3 . (e and f) AFM images of the kinetically controlled AggI (e) and thermodynamically controlled AggII (f) of compound (*R*)-**2** (MCH/ CHCl_3 9/1 as solvent; HOPG as surface; $c_T = 10 \mu\text{M}$).

per monomeric unit, and the subsequent intermolecular interaction between them, would yield the kinetically controlled AggI species (Fig. S17†). The further evolution of the AggI species relies on the rupture of the intramolecular H-bonding interactions to form highly stable tetrafold intermolecular H-bonding arrays that afford to AggII species (Fig. S17†).

Emissive properties of the cyano-luminogens 1 and 2

The achiral congeners of cyano-luminogens **1** and **2** have been reported to behave as highly emissive soft-materials with color-tunable features and aggregation-induced enhanced emission (AIE).^{14,21} Noteworthy, the emission studies developed for such achiral cyano-luminogens were carried out in mixtures 7/3 MCH/1,2-dichloroethane (DCE) without detecting the formation of the two aggregated species AggI and AggII.¹⁴ Taking into account these emissive characteristics, we have also registered the emission spectra of kinetically and thermodynamically aggregated species formed by compounds **1** and **2**. Similarly to that observed for the dichroic features described in the previous part, the emissive features of the *p*-phenylenevinyls **1** differ to those observed for the thienyl-derivatives **2**. Fig. 5a displays the photoluminescence (PL) spectra of freshly prepared and aged diluted solutions ($c_T = 10 \mu\text{M}$) of compound (**S**)-**1** in a 9/1 MCH/ CHCl_3 mixture, together with the PL spectrum registered in CHCl_3 , that corresponds to the monomeric species. The latter spectrum shows two peaks at $\lambda = 486$ and 436 nm , in good agreement with that previously described for the achiral congener (Fig. 5a).¹⁴ The supramolecular polymerization of (**S**)-**1** provokes a strong increase of the PL intensity accompanied by a bathochromic shift, diagnostic of the operation of an AIE phenomenon.²¹ Importantly, the PL spectrum of both AggI and AggII differs in shape but not in intensity. Thus, the PL spectrum of AggI species shows a broad emission band centered at

$\lambda = 543 \text{ nm}$. Aging this sample to generate AggII changes the PL spectrum which, for this aggregated species, displays a fine structure with an intense band centered at $\lambda = 490 \text{ nm}$ and two shoulders at $\lambda = 524$ and 466 nm (Fig. 5a). The different emissive properties are also well observed in the corresponding Commission International de l'Eclairage (CIE) that present CIE values for the monomeric, AggI and AggII species of (0.15, 0.15), (0.39, 0.52) and (0.21, 0.40), respectively (Fig. S18†).

The emissive characteristics of the chiral cyano-derivatives (**S**)-**1** and (**R**)-**1** prompted us to register the CPL activity displayed by both the monomeric and aggregated species. A first remarkable finding is that, whilst the monomeric species shows no CPL activity in solution, the supramolecular polymerization of (**S**)-**1** and (**R**)-**1** results in intense CPL spectra with opposite sign for AggI and AggII species (Fig. 5b). Thus, the CPL spectra of the kinetically controlled AggI species shows a broad CPL band centered at $\lambda \sim 560 \text{ nm}$ with a positive sign for (**R**)-**1** and a negative sign for (**S**)-**1**, in good agreement to that observed in the CD spectra (Fig. S5†). These kinetically controlled AggI species exhibit a remarkable dissymmetry factor $g_{\text{lum}} \sim 0.007$ ($\lambda_{\text{exc}} = 365 \text{ nm}$) (Fig. 5b). Interestingly, aging the 9/1 MCH/ CHCl_3 solution of AggI for 24 h, to ensure the complete conversion to AggII, results in CPL active supramolecular aggregates in which a sign inversion is detected. Thus, the AggII species of compound (**R**)-**1** shows negative CPL signal, the opposite being detected for (**S**)-**1**. Furthermore, the g_{lum} slightly decreases to values ~ 0.004 (Fig. 5b). The CPL spectra demonstrate that the CPL activity of the kinetically controlled AggI species follows the helicity dictated by the stereogenic centres located at the peripheral side chains. The intertwining effect responsible for the formation of the AggII species changes the CPL sign. This CPL sign inversion, well documented for molecular species,²² is rarely observed for supramolecular polymers.²³

Finally, we have also monitored the emissive features of thienyl-derivatives **2** that, similarly to that observed in the CD spectra, present noticeable changes regarding compounds **1**. The PL spectra of compounds **2** in the molecularly dissolved state present two maxima of low intensity at $\lambda = 520$ and 488 nm that becomes a very broad and intense band centered at $\lambda = 607 \text{ nm}$ for the kinetically controlled AggI₂ species (Fig. 5c). This strong increase in the PL intensity is diagnostic of an AIE process as it has been reported for the achiral congener.¹⁴ To our surprise, the evolution of AggI₂ species to the thermodynamically controlled AggII₂ species results in a significant quenching of the intensity that implies an aggregation quenching effect (ACQ) related to the kinetically aggregated species (Fig. 5c). In this case, the colour changes are less pronounced with CIE coordinates closer for the monomeric, AggI₂ and AggII₂ species [(0.51, 0.32), (0.42, 0.52) and (0.41, 0.67), respectively] (Fig. S18†).

The emissive properties of compounds **2** are also reflected in the CPL activity of the aggregated species of these cyano-luminogens. Thus, neither the monomeric species nor the thermodynamically controlled supramolecular polymers AggII₂ present CPL activity. In fact, the CPL spectra of AggII₂ for both (**R**)-**2** and (**S**)-**2** are very small and not mirror images and only



Fig. 5 PL (a and c) and CPL/PL (b and d) of (**S**)-**1** (a and b) and (**S**)-**2** (c and d) spectra in monomeric and aggregated states (CHCl_3 and 9/1 MCH/ CHCl_3 mixture, respectively) (experimental conditions: $c_T = 10 \mu\text{M}$; $\lambda_{\text{exc}} = 365 \text{ nm}$ for (**S**)-**1** and 390 nm for (**S**)-**2**). Dashed arrows in panel (b) indicate the inversion of the CPL sign.



a positive activity is observed for these aggregates ascribable to potential artefacts, linear anisotropic effects or photoselection, derived from the aggregation process (Fig. S19†).²⁴ Nevertheless, the large increase of the PL observed for the kinetically controlled supramolecular polymers formed by compounds **2** is also observed in the CPL activity shown by these aggregates whose CPL spectra are almost mirror images for (*R*)-**2** and (*S*)-**2** with $g_{\text{lum}} \sim 0.002$ ($\lambda_{\text{exc}} = 390 \text{ nm}$) (Fig. 5d).

Conclusions

In conclusion, we report herein the synthesis of two pairs of enantiomeric cyano-luminogens **1** and **2** in which the central chromophore is a *p*-phenylene or a 2,5-dithienylbenzene moiety, respectively. The decoration of these chromophoric units with four benzamide units per monomeric unit favours their supramolecular polymerization in paraffinic solvents like MCH and also in mixtures of MCH/CHCl₃ by the formation of quadruple H-bonding arrays between the amide groups and the π -stacking of the central aromatic moieties. These two benzamide units, separated by two methylenes, are able to form intramolecularly H-bonded pseudocycles that behave as metastable monomeric species **M***. These **M*** species generate kinetically and thermodynamically controlled aggregated species AggI and AggII. The presence of point chirality in the peripheral side chains gives rise to helical aggregates exhibiting rich chiroptical features depending on the aggregation state and the central aromatic unit. The *p*-phenylene-based derivatives **1** exhibit a clear bisignated dichroic response of different intensity but similar sign for both AggI and AggII species and remarkable emissive features. The formation of helical (AggI₁) and superhelical (AggII₁) aggregates is confirmed by AFM imaging. Furthermore, both AggI₁ and AggII₁ species show CPL activity of opposite sign depending on the aggregation state. Despite the structural similarity between compounds **1** and **2**, the latter displays a dissimilar chiroptical, morphological and emissive characteristics for the corresponding kinetically and thermodynamically controlled aggregated species AggI and AggII in comparison to those registered for compounds **1**. Thus, a stereomutation phenomenon is observed in the AggI₂ → AggII₂ with opposite dichroic response of each enantiomer depending on the aggregation state. In addition, AggI₂ is arranged into nanoparticles that evolve to helical aggregates to afford AggII₂. The dissimilar chiroptical and morphological features of AggI₂ and AggII₂ species are also appreciated in the emissive properties. Thus, whilst AggI₂ experiences a clear AIE process and CPL activity, the thermodynamically controlled AggII₂ undergoes an aggregation caused quenching (ACQ) process regarding AggI₂ in which the CPL activity vanishes. Therefore, the results shed light on the modulation of relevant optical and chiroptical properties of luminescent self-assembling units able to form highly organized supramolecular structures depending on the aggregation state and contribute to expand the knowledge and potential applicability of kinetically and thermodynamically-controlled supramolecular polymers.

Data availability

All data associated with this article have been included in the main text and ESI.†

Author contributions

Lucía López-Gandul, Cristina Naranjo and Cecilia Sánchez: synthesis, spectroscopic characterization, supramolecular polymerization mechanisms, AFM imaging and writing. Rafael Rodríguez: chiroptical and emissive properties and writing; Rafael Gómez, Jeanne Crassous and Luis Sánchez: conceptualization, writing, supervision, and funding acquisition.

Conflicts of interest

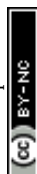
There are no conflicts to declare.

Acknowledgements

Financial support by the MCIU of Spain (PID2020-113512GB-I00 and RED2018-102331-T) and the Comunidad de Madrid (S2018/NMT-4389) is acknowledged. R. R. thanks Xunta de Galicia for postdoctoral fellowship. J. C. acknowledges the Ministère de l'Education Nationale, de la Recherche et de la Technologie, the Centre National de la Recherche Scientifique (CNRS). J. C. warmly thanks Dr Kais Dhbaibi for his technical assistance.

Notes and references

- (a) T. Aida, E. W. Meijer and S. I. Stupp, *Science*, 2012, **335**, 813; (b) Z. Álvarez, A. N. Kolberg-Edelbrock, I. R. Sasselli, J. A. Ortega, R. Qiu, Z. Syrgiannis, P. A. Mirau, F. Chen, S. M. Chin, S. Weigand, E. Kiskinis and S. I. Stupp, *Science*, 2021, **374**, 848; (c) O. Dumele, L. Đorđević, H. Sai, T. J. Cotey, M. H. Sangji, K. Sato, A. J. Dannenhoffer and S. I. Stupp, *J. Am. Chem. Soc.*, 2022, **144**, 3127; (d) W. Zhao, J. Tropp, B. Qiao, M. Pink, J. D. Azoulay and A. H. Flood, *J. Am. Chem. Soc.*, 2020, **142**, 2579; (e) A. T. Haedler, K. Kreger, A. Issac, B. Wittmann, M. Kivala, N. Hammer, J. Köhler, H. W. Schmidt and R. Hildner, *Nature*, 2015, **523**, 196.
- (a) T. F. A. de Greef, M. M. J. Smulders, M. Wolffs, A. P. H. J. Schenning, R. P. Sijbesma and E. W. Meijer, *Chem. Rev.*, 2009, **109**, 5687; (b) M. Wehner and F. Würthner, *Nat. Rev. Chem.*, 2020, **4**, 38.
- (a) A. R. A. Palmans and E. W. Meijer, *Angew. Chem., Int. Ed.*, 2007, **46**, 8948; (b) Y. Dorca, E. E. Greciano, J. S. Valera, R. Gómez and L. Sánchez, *Chem.-Eur. J.*, 2019, **25**, 5848.
- (a) E. E. Greciano, J. Calbo, J. Buendía, J. Cerdá, J. Aragón, E. Ortí and L. Sánchez, *J. Am. Chem. Soc.*, 2019, **141**, 7463; (b) F. García and L. Sánchez, *J. Am. Chem. Soc.*, 2012, **134**, 734; (c) M. M. J. Smulders, I. A. W. Filot, J. M. A. Leenders, P. van der Schoot, A. R. A. Palmans, A. P. H. J. Schenning and E. W. Meijer, *J. Am. Chem. Soc.*, 2010, **132**, 611.
- (a) R. Sethy, J. Kumar, R. Métivier, M. Louis, K. Nakatani, N. M. T. Mecheri, A. Subhakumari, K. G. Thomas, T. Kawai



- and T. Nakashima, *Angew. Chem., Int. Ed.*, 2017, **56**, 15053; (b) A. K. Mondal, M. D. Preuss, M. L. Ślęczkowski, T. K. Das, G. Vantomme, E. W. Meijer and R. Naaman, *J. Am. Chem. Soc.*, 2021, **143**, 7189.
- 6 (a) J. Matern, Y. Dorca, L. Sánchez and G. Fernández, *Angew. Chem., Int. Ed.*, 2019, **58**, 16730; (b) A. Sorrenti, J. Leira-Iglesias, A. J. Markvoort, T. F. A. de Greef and T. M. Hermans, *Chem. Soc. Rev.*, 2017, **46**, 5476.
- 7 (a) P. A. Korevaar, S. J. George, A. J. Markvoort, M. M. J. Smulders, P. A. J. Hilbers, A. P. H. J. Schenning, T. F. A. de Greef and E. W. Meijer, *Nature*, 2012, **481**, 492; (b) T. Fukui, S. Kawai, S. Fujinuma, Y. Matsushita, T. Yasuda, T. Sakurai, S. Seki, M. Takeuchi and K. Sugiyasu, *Nat. Chem.*, 2017, **9**, 493; (c) E. E. Greciano, J. Calbo, E. Ortí and L. Sánchez, *Angew. Chem., Int. Ed.*, 2020, **59**, 17517.
- 8 (a) J. Kang, D. Miyajima, T. Mori, Y. Inoue, Y. Itoh and T. Aida, *Science*, 2015, **347**, 646; (b) S. Ogi, V. Stepanenko, K. Sugiyasu, M. Takeuchi and F. Würthner, *J. Am. Chem. Soc.*, 2015, **137**, 3300; (c) J. S. Valera, R. Gómez and L. Sánchez, *Small*, 2018, **14**, 1702437; (d) S. Sarkar, A. Sarkar, A. Som, S. S. Agasti and S. J. George, *J. Am. Chem. Soc.*, 2021, **143**, 11777.
- 9 J. S. Valera, R. Gómez and L. Sánchez, *Angew. Chem., Int. Ed.*, 2019, **58**, 510.
- 10 (a) J. Han, S. Guo, H. Lu, S. Liu, Q. Zhao and W. Huang, *Adv. Opt. Mater.*, 2018, **6**, 1800538; (b) M. Schulz, F. Balzer, D. Scheunemann, O. Arteaga, A. Lützen, S. C. J. Meskers and M. Schiek, *Adv. Funct. Mater.*, 2019, **29**, 1900684; (c) L. Wan, J. Wade, F. Salerno, O. Arteaga, B. Laidlaw, X. Wang, T. Penfold, M. J. Fuchter and A. J. Campbell, *ACS Nano*, 2019, **13**, 8099; (d) Y. Sang, J. Han, T. Zhao, P. Duan and M. Liu, *Adv. Mater.*, 2020, **32**, 1900110.
- 11 (a) E. E. Greciano, R. Rodríguez, K. Maeda and L. Sánchez, *Chem. Commun.*, 2020, **56**, 2244; (b) J. S. Kang, S. Kang, J.-M. Suh, S. M. Park, D. K. Yoon, M. H. Lim, W. Y. Kim and M. Seo, *J. Am. Chem. Soc.*, 2022, **144**, 2657.
- 12 (a) E. S. Gauthier, L. Abella, N. Hellou, B. Darquie, E. Caytan, T. Roisnel, N. Vanthuyne, L. Favereau, M. Srebro-Hooper, J. A. G. Williams, J. Autschbach and J. Crassous, *Angew. Chem., Int. Ed.*, 2020, **59**, 8394; (b) E. M. Sánchez-Carnerero, F. Moreno, B. L. Maroto, A. R. Agarrabeitia, M. J. Ortiz, B. G. Vo, G. Muller and S. de la Moya, *J. Am. Chem. Soc.*, 2014, **136**, 3346; (c) M. A. Medel, R. Tapia, V. Blanco, D. Miguel, S. P. Morcillo and A. González Campaña, *Angew. Chem., Int. Ed.*, 2021, **60**, 6094.
- 13 (a) H. Maeda, K. Bando, K. Shimomura, I. Yamada, M. Naito, H. Nobusawa, H. Tsumatori and T. Kawai, *J. Am. Chem. Soc.*, 2011, **133**, 9266; (b) H. Fan, H. Jiang, X. Zhu, Z. Guo, L. Zhang and M. Liu, *Nanoscale*, 2019, **11**, 10504; (c) M. Louis, R. Sathy, J. Kumar, S. Katao, R. Guillot, T. Nakashima, C. Allain, T. Kawai and R. Metivier, *Chem. Sci.*, 2019, **10**, 843; (d) J. Han, P. Duan, X. Li and M. Liu, *J. Am. Chem. Soc.*, 2017, **139**, 9783; (e) D. Niu, Y. Jiang, L. Ji, G. Ouyang and M. Liu, *Angew. Chem., Int. Ed.*, 2019, **58**, 5946.
- 14 F. Aparicio, S. Cherumukkil, A. Ajayaghosh and L. Sánchez, *Langmuir*, 2016, **32**, 284.
- 15 (a) M. A. Martínez, A. Doncel-Giménez, J. Cerdá, J. Calbo, R. Rodríguez, J. Aragón, J. Crassous, E. Ortí and L. Sánchez, *J. Am. Chem. Soc.*, 2021, **143**, 13281; (b) N. Bäumer, K. K. Kartha, S. Buss, I. Maisuls, J. P. Palakkal, C. A. Strassert and G. Fernández, *Chem. Sci.*, 2021, **12**, 5236; (c) H. Wang, Y. Zhang, Y. Chen, H. Pan, X. Ren and Z. Chen, *Angew. Chem., Int. Ed.*, 2020, **59**, 5185.
- 16 (a) M. Wehner, M. I. S. Röhr, M. Bühler, V. Stepanenko, W. Wagner and F. Würthner, *J. Am. Chem. Soc.*, 2019, **141**, 6092; (b) E. E. Greciano, S. Alsina, G. Ghosh, G. Fernández and L. Sánchez, *Small Methods*, 2020, **4**, 1900715; (c) E. E. Greciano, M. A. Martínez, S. Alsina, A. Laguna and L. Sánchez, *Org. Chem. Front.*, 2021, **8**, 5328.
- 17 E. E. Greciano and L. Sánchez, *Chem.-Eur. J.*, 2016, **22**, 13724.
- 18 J. Matern, Z. Fernández, N. Bäumer and G. Fernández, *Angew. Chem., Int. Ed.*, 2022, **61**, e202203783.
- 19 P. A. Korevaar, C. Schaefer, T. F. A. de Greef and E. W. Meijer, *J. Am. Chem. Soc.*, 2012, **134**, 13482.
- 20 M. Hifsudheen, R. K. Mishra, B. Vedhanarayanan, V. K. Praveen and A. Ajayaghosh, *Angew. Chem., Int. Ed.*, 2017, **56**, 12634.
- 21 J. Mei, N. L. C. Leung, R. T. K. Kwok, J. W. Y. Lam and B. Z. Tang, *Chem. Rev.*, 2015, **115**, 11718.
- 22 (a) K. Takaishi, K. Iwachido and T. Ema, *J. Am. Chem. Soc.*, 2020, **142**, 1774; (b) M.-Y. Zhang, X. Liang, D.-N. Ni, D.-H. Liu, Q. Peng and C.-H. Zhao, *Org. Lett.*, 2021, **23**, 2; (c) K. Takaishi, M. Yasui and T. Ema, *J. Am. Chem. Soc.*, 2018, **140**, 5334; (d) K. Takaishi, K. Iwachido, R. Takehana, M. Uchiyama and T. Ema, *J. Am. Chem. Soc.*, 2019, **141**, 6185; (e) E. Yen-Pon, F. Buttard, L. Frédéric, P. Thuéry, F. Taran, G. Pieters, P. A. Champagne and D. Audisio, *JACS Au*, 2021, **1**, 807.
- 23 (a) R. Rodríguez, C. Naranjo, A. Kumar, P. Matozzo, T. K. Das, Q. Zhu, N. Vanthuyne, R. Gómez, R. Naaman, L. Sánchez and J. Crassous, *J. Am. Chem. Soc.*, 2022, **144**, 7709; (b) S. Fa, T. Tomita, K. Wada, K. Yasuhara, S. Ohtani, K. Kato, M. Gon, K. Tanaka, T. Kakuta, T. Yamagishi and T. Ogoshi, *Chem. Sci.*, 2022, **13**, 5846; (c) F. Gan, C. Shen and H. Qiu, Circularly Polarized Luminescence of Aggregation-induced Emission Materials, *Handb. Aggregation-Induced Emiss.*, 2022, vol. 3, pp. 27–56.
- 24 (a) I. Tinoco Jr, B. Ehrenberg and I. Z. Steinberg, *J. Chem. Phys.*, 1977, **66**, 916; (b) H. P. J. M. Dekkers, P. F. Moraal, J. M. Timper and J. P. Riehl, *Appl. Spectrosc.*, 1985, **39**, 818; (c) E. Castiglioni, S. Abbate, F. Lebon and G. Longhi, *Chirality*, 2012, **24**, 725; (d) G. Albano, G. Pescitelli and L. Di Bari, *Chem. Rev.*, 2020, **120**, 10145.

