

REVIEW

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Stimuli-responsive dynamic pseudorotaxane crystals

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This review summarizes recent advances in stimuli-responsive dynamic molecular crystals. For the construction of dynamic molecular crystals with response to external stimuli, such as light and/or heat, a number of strategies are available, which are based on (i) metal–organic frameworks, (ii) molecular rotors, (iii) photochromic molecules including azobenzenes and diarylethenes, (iv) cycloaddition reactions, and (v) rotaxanes and pseudorotaxanes. In particular, this review focuses on our recent results on dynamic pseudorotaxane crystals. Pseudorotaxane molecules comprising a crown ether as a ring and an ammonium cation as an axle respond to light and/or heat changing their molecular structures in the single-crystal state. Such structural changes at a molecular level influence the macroscopic properties of pseudorotaxanes, including mechanical motions of the crystals. Consequently, they are potential candidates for application as light-sensing and harvesting materials in microdevices for photomechanical conversion.

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1. Introduction

Stimulus-responsive molecular machines and switches have been recently attracting the interest of scientists in the fields of chemistry and interdisciplinary research.^{1–20} This has produced interesting developments, such as the first nanocar race, in which single-molecule nanocars are controlled on a two-dimensional surface by scanning tunneling microscopes.²¹ Meanwhile, rotaxanes and pseudorotaxanes comprising interlocked ring and axle

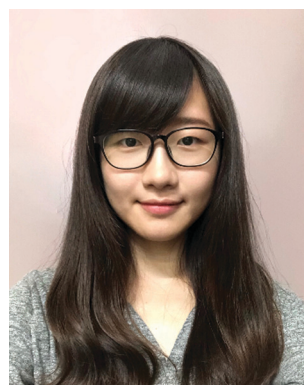
molecules exhibit translational and rotational motions of their components, which leads to changes in their relative positions.^{2–16} These inter- and intramolecular interactions and molecular motions can be activated and controlled by external stimuli, converting chemical, electrical, photochemical, and electrochemical energy into mechanical motion. The motion of rotaxanes and pseudorotaxanes is effectively controlled in various media, such as solution, aggregation, film, and crystal (Fig. 1). In solution and aggregation states, molecules have overall random alignment (isotropic), and the relative positions of the complexes can be changed. In contrast, molecules have well-defined two- and three-dimensional alignments in films and

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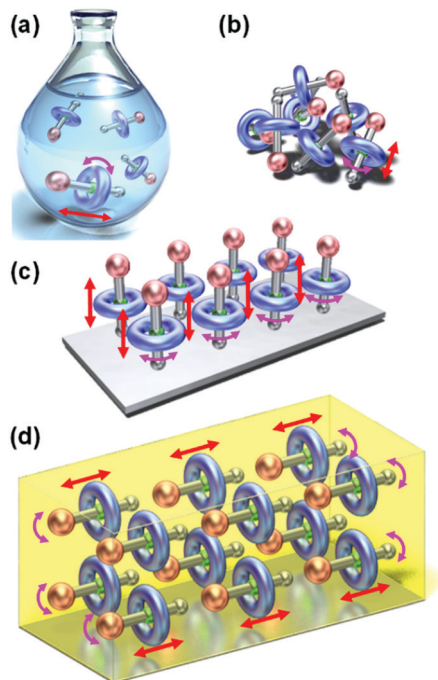


Fig. 1 Conceptual illustrations of the motions of rotaxanes in (a) solution, (b) aggregation, (c) film, and (d) crystal.

single crystals, respectively. In these cases, concerted molecular motions can be induced, leading to specific large crystal motions. In addition, the continuous minimization of these molecules has led to micro- and nanoscale applications *via* top-down approaches, whereas their self-assembled well-aligned architectures in the solid state provide large motions on a macroscopic level by bottom-up approaches.^{22–25}

The number of detailed studies on the molecular motions of rotaxanes and pseudorotaxanes in the crystal state is limited partly due to the lack of investigations on the molecular dynamics in such state. This review article aims to summarize the current literature on stimuli-responsive molecules with molecular dynamics in the crystal state. Furthermore, our recent results on pseudorotaxane **1** (Fig. 2) and related molecules are described. The dynamic motions of these molecules can be activated by heat and/or photoirradiation in the crystal state and are related to macroscopic motions of the crystals. Finally, new insights into the relationship between molecular structure and mechanical motions of pseudorotaxane crystals are discussed.

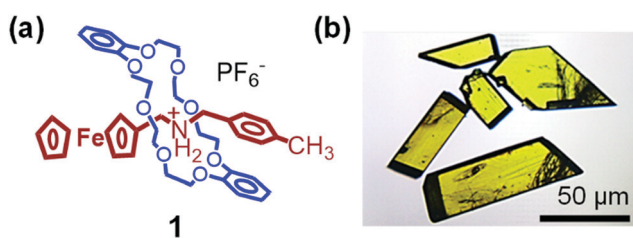


Fig. 2 (a) Chemical structure of pseudorotaxane **1** and (b) its single crystals.

2. Molecular motions induced in the solid state

This section describes various strategies that have been proposed to induce mechanical motions in the solid state, such as the construction of interlocked structures or the use of photochemical reactions.^{23,26–28}

2.1 MOFs incorporating rotaxanes and molecular rotors

The first approach involves constructing interlocked molecules, such as rotaxanes and pseudorotaxanes, which consist of ring and axle components. These interlocked molecules have been introduced into metal–organic frameworks (MOFs) with sufficient porosity and internal volume to accommodate the movement of the bulky axle component.^{29–34} Fig. 3 shows a conceptual illustration of these MOFs, in which rotaxanes are introduced by linking their ring component with the organic struts of the MOFs or by direct insertion of rotaxanes into the MOFs. For instance, Stoddart, Sauvage, Yaghi, *et al.* reported the incorporation of copper(i)-complexed [2]pseudorotaxane struts into a zinc oxide MOF.³⁰ They synthesized a Sauvage-type organic strut composed of a phenanthroline axle and an oligo-ether tethered phenanthroline ring using a Cu(i) template, which was then removed by demetalation. The resulting MOF was analyzed by powder X-ray diffraction techniques. Schurko, Loeb, *et al.* also investigated the molecular dynamics of a rotaxane in a zinc oxide MOF.^{31–33} They synthesized an organic strut composed of an imidazolium axle and a crown ether ring by ring-closing metathesis. In addition to powder X-ray diffraction techniques, the molecular dynamics were analyzed by ¹³C and ²H solid-state NMR spectroscopy focusing on the motion of the ring component. Unfortunately, despite these advances, direct observation of molecular structural changes by *in situ* single-crystal X-ray crystallography in this kind of MOF is still challenging due to their complicated molecular structures.

Molecular rotors and gyroscopes show dynamic rotational motions in the crystal state. Garcia-Garibay *et al.* reported various molecular rotators, in which typically a 1,4-diethynylbenzene axle is connected with two bulky moieties, such as triphenylmethyl (CPh₃) groups, to produce free volume (**2**, Fig. 4).^{35,36} When the diethynylbenzene axle is capped with bulkier fragments (*e.g.*, **3** in Fig. 4), the rotation of the axle is faster.^{35,36} In addition, two phenyl groups of **2** were connected with an azobenzene group to produce photochromic molecular gyroscope **4** (Fig. 4), in which

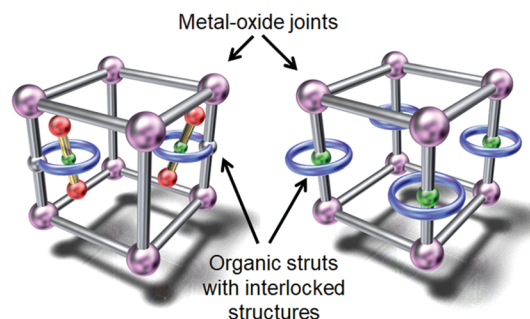


Fig. 3 Conceptual illustration of MOFs incorporating rotaxanes.

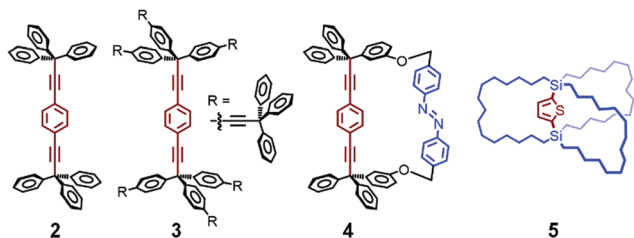


Fig. 4 Molecular rotors and gyroscopes.

the rotation rate of the axle part is controlled by photoisomerization of the azobenzene group in the solid state.³⁷ Their as-synthesized molecular structures were determined by single-crystal X-ray crystallography, and their rotation rates were estimated by ²H solid-state NMR spectroscopy at various temperatures.

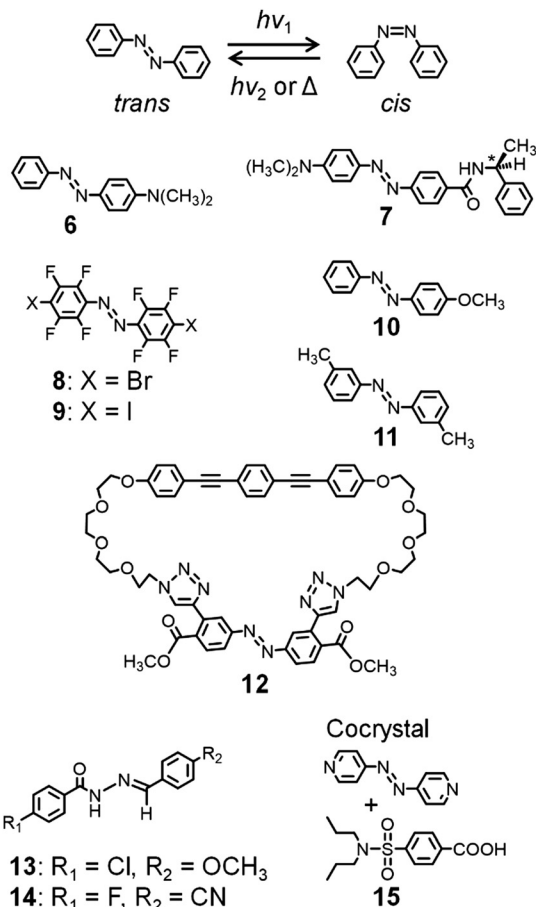
Setaka *et al.* reported a crystalline molecular gyroscope 5 (Fig. 4), which was synthesized by ring-closing metathesis of 1,4-bis(tri-7-octenyl)silylthiophene.³⁸ This molecule shows thermally induced order-disorder phase transition in a single-crystal state, as demonstrated by single-crystal X-ray crystallography, polarized optical microscopy, and ²H solid-state NMR spectroscopy. In addition, they replaced the 2,5-thienylene unit of the axle core with biphenylene³⁹ or thiophene dioxide-diyl⁴⁰ and modified the outer chain length to control the axle rotation in the crystal state.

2.2 Photochromic molecules

Mechanical motions in the solid state can also be induced by subjecting photochromic molecules, such as azobenzenes and diarylethenes, to photochemical reactions. Azobenzene-containing polymer films exhibit reversible and controllable bending based on *cis-trans* photoisomerization, which has been exploited for applications involving the bending of films and the rotation of a light-driven plastic motor.^{22,41–43}

Photoresponsive dynamic azobenzene crystals have also been described (Fig. 5). Koshima *et al.* reported that plate-shaped crystals of 4-dimethylaminoazobenzene 6 show reversible bending upon UV irradiation.⁴⁴ Recently, the same group reported the photoinduced bending, twisting, and rolling of chiral azobenzene 7.⁴⁵ This crystal shows crystal-to-crystal thermal phase transition upon heating at 145 °C and cooling at 143 °C, exhibiting flipping of the phenyl group, whose structural changes were analyzed by single-crystal X-ray crystallography.⁴⁶

Barrett and Frišević *et al.* reported the *cis* and *trans* forms of octafluorinated 4,4'-dihalogenated azobenzenes 8 (X = Br) and 9 (X = I) (Fig. 5), which were synthesized by oxidative coupling using MnO₂.⁴⁷ The *cis* forms show very long half-lives (*ca.* 2 months), allowing their isolation as crystals. In addition, their *cis* forms and *cis-trans* isomerization in the crystals were analyzed by single-crystal X-ray crystallography. These crystals show irreversible bending and crystal-to-crystal isomerization upon 457 nm irradiation. Furthermore, they cocrystallized these molecules with *cis*- and *trans*-1,2-bis(4-pyridyl)ethylene and with *trans*-4,4'-azopyridine, affording supramolecular structures based on halogen bond-driven cocrystallization.⁴⁸

Fig. 5 *cis-trans* isomerization of azobenzene and chemical structures of azo-molecules exhibiting dynamic crystal motions.

Norikane *et al.* demonstrated the photoinduced translation (so-called swimming) of a single crystal of 4-methoxyazobenzene 10 (Fig. 5) on the water surface, which is caused by its *trans-cis* photoisomerization.⁴⁹ In addition, a crystal of 10 was used to carry a boat on the water surface under UV light. Furthermore, they observed photoinduced translational motions, also called crawling, of crystals of 3,3'-dimethylazobenzene 11 (Fig. 5), caused by retraction induced by UV light irradiation and crystal growth prompted by visible light.⁵⁰

In contrast with the relatively simple chemical structures of the abovementioned molecules, Muraoka and Kinbara *et al.* reported a cyclic azobenzene tethered with two tetra(ethylene glycol) chains and a 1,4-bis(phenylethynyl)benzene group (12, Fig. 5).⁵¹ Interestingly, crystals of 12 show crystal-to-crystal thermal phase transition at 60 °C and 56 °C on heating and cooling, respectively, caused by structural changes in the tetra(ethylene glycol) chains due to their flexible features.

Naumov and Nath *et al.* described the occurrence of multiple mechanical responses of crystals of acylhydrazone derivatives 13 and 14 (Fig. 5).⁵² Thus, a needle-shaped crystal of 13 undergoes plastic bending by mechanical stress, with storing of the deformed shape after releasing the stress, whereas 14 undergoes elastic bending by mechanical stress. Furthermore, these crystals

exhibit photoinduced bending, caused by *E*-to-*Z* photoisomerization upon exposure to 365 nm UV light. In addition, they reported that cocrystals of 4,4'-azopyridine and probenecid **15** (Fig. 5) in a 1:2 molar ratio respond to multiple external stimuli (heat, UV light, and mechanical pressure), exhibiting twisting, bending, and elastic deformation.⁵³ The photo- and thermomechanical effects and self-healing properties of **15** result from the *cis*-*trans* photoisomerization of the azopyridine unit and crystal-to-crystal thermal phase transition. They proposed that such crystalline materials could find application as crystal actuators.

Meanwhile, diarylethene derivatives have been shown to exhibit significant color changes accompanied by ring-closing and ring-opening photochemical reactions, being both processes photochemically reversible but thermally irreversible (Fig. 6).⁵⁴ Irie and Kobatake *et al.* reported various diarylethene derivatives whose crystals show reversible dynamic motions based on photoisomerization induced by UV or visible light irradiation.⁵⁴⁻⁶⁰ Kobatake and Bardeen *et al.* reported the control of the bending direction of a ribbon-shaped crystal of **16** (Fig. 6) by irradiation direction.⁵⁸ Furthermore, they prepared organic-inorganic hybrid materials as photon-powered actuators using nanocrystals of **16** and porous alumina, which was used for the alignment control of **16**.⁵⁹ A crystal of *n*-octyloxy-containing **17** (Fig. 6) was found to exhibit bending and crystal-to-crystal thermal phase transitions on heating and cooling, which is exceptional for the diarylethene series.⁵⁷ Such thermally induced behavior of **17** was combined with photochromism to control crystal bending.

Fluorescence switching could also be controlled by irradiating a crystal of **18** tethered with a 4,7-diphenylbenzothiadiazole unit with UV and visible light.⁶⁰ Due to intermolecular energy transfer in the dense molecular packing structure of **18**, its open-ring state is fluorescent, whereas its closed-ring state is nonfluorescent.

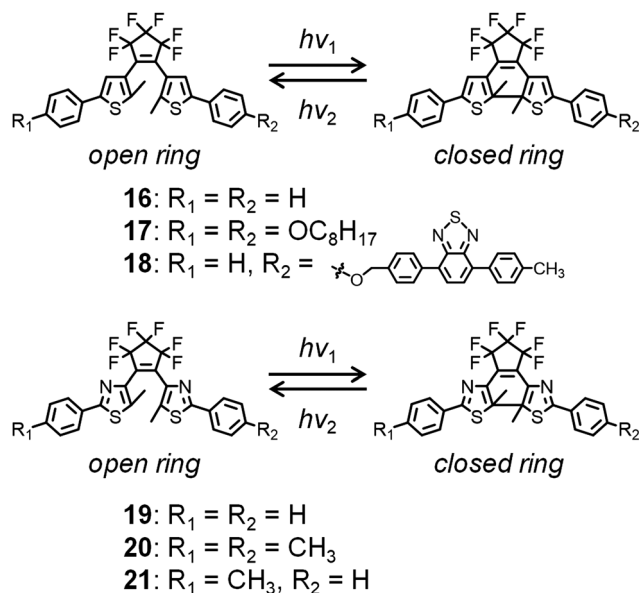


Fig. 6 Photochromism of diarylethene derivatives exhibiting dynamic crystal motions.

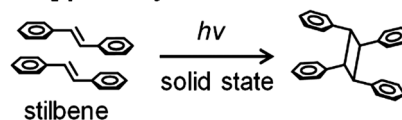
Finally, a crystal of **19** and mixed crystals of **20** and **21** (Fig. 6) were exploited for applications involving an electric switch⁵⁶ and rotation of gears, respectively.⁵⁵

2.3 Cycloaddition in crystals

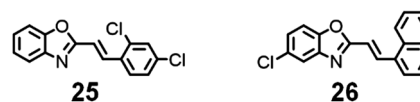
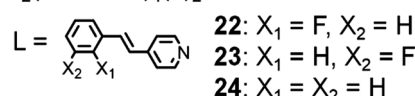
The [2+2] photocycloaddition of two aligned olefins (*e.g.*, stilbene in Fig. 7) induced by UV or visible light provides carbocyclic products.⁶¹⁻⁶³ In the crystal state, the regular alignment of the target molecules allows this [2+2] photocycloaddition to proceed with high regio- and stereocontrol. The photocycloaddition requires that the two reactive sites are located within a short distance, preferably shorter than 4.2 Å, from each other.⁶¹⁻⁶³ To achieve this, various strategies using weak intermolecular interactions, hydrogen bonds, coordination complexes, external templates, and MOFs have been reported.⁶¹⁻⁶⁸

In addition, photocycloaddition can induce dynamic motions, such as bending, curling, rolling, and salient behaviors in crystals. For instance, Vittal and Naumov *et al.* reported photosalient effects of Zn-benzoate (PhCOO⁻) complexes, [Zn₂(benzoate)₄(L)₂] (**22-24**), where the ligand (L) is 2'-fluoro-4-styrylpyridine for **22**,

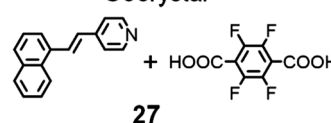
[2 + 2] photocycloaddition



Zn₂(benzoate)₄(L)₂



Cocrystal



[4 + 4] photocycloaddition

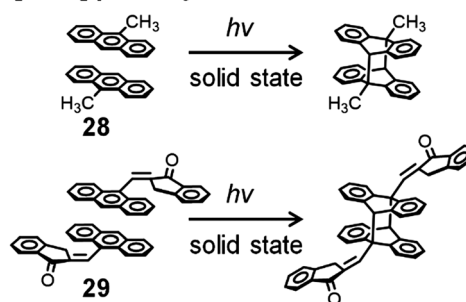


Fig. 7 Photocycloaddition of stilbene and anthracene derivatives exhibiting dynamic crystal motions.

3'-fluoro-4-styrylpyridine for **23**, or 4-styrylpyridine for **24** (Fig. 7).⁶⁹ In addition, using the same ligands, they synthesized Zn-cinnamate ($\text{PhCH}=\text{CHCOO}^-$) complexes,⁷⁰ Zn-isothiocyanate ($\text{N}=\text{C}=\text{S}^-$) complexes,⁷¹ and silver complexes (AgL_2X_2) with various counter-anions ($\text{X} = \text{BF}_4^-, \text{ClO}_4^-, \text{or } \text{NO}_3^-$).⁷² Upon exposure to weak UV light, with the concomitant [2+2] photocycloaddition, these silver crystals burst violently and are propelled several millimeters. The crystals of silver complexes exhibit photosalient, popped, hopped, and/or leaped motions, depending on their head-to-tail or head-to-head arrangement.

Lu *et al.* investigated the motion of (*E*)-2-(2,4-dichlorostyryl)-benzo[d]oxazole **25** and (*E*)-5-chloro-2-(2-(naphthalen-1-yl)vinyl)-benzo[d]oxazole **26** (Fig. 7). Their crystals show bending, curling, and rolling caused by [2+2] photocycloaddition.^{73,74} In particular, crystals of **26** exhibit green photoluminescence with a maximum photoluminescence wavelength (λ_{PL}) of 492 nm, which changes to blue with $\lambda_{\text{PL}} = 454$ nm after UV irradiation at 365 nm. They proposed that such photoluminescence color change could be useful for the remote detection of photomechanical reactions. Yan *et al.* reported that cocrystals of naphthylvinylpyridine and tetrafluoroterephthalic acid **27** with the molar ratio of 2:1 show photosalient phenomena resulting from [2+2] photocycloaddition, whereas the 1:1 derivative shows bending motions.⁷⁵ The photochemical reaction also affects the photoluminescence intensity of these crystals.

Bardeen *et al.* reported photoinduced dynamics of crystal-line 9-methylantracene **28**, whose motions are induced by the [4+4] photodimerization of two molecules in the crystal state (Fig. 7).^{76,77} A ribbon-like crystal of **28** shows photoinduced rolling and unrolling during UV exposure and during extended UV exposure, respectively, whereas a sheet-like crystal of **28** shows twisting and untwisting during UV exposure and during extended UV exposure, respectively. They demonstrated translational motion of the sheet-like crystal induced by a magnetic field in the presence of superparamagnetic colloidal nanocrystal clusters.⁷⁷ Koshima *et al.* also reported photoinduced crystal bending motions induced by the [4+4] photodimerization of (*E*)-2-(9-anthrylmethylene)-1-indanone **29** (Fig. 7).⁷⁸

Sureshan *et al.* reported the thermal dimerization of **30** (Fig. 8) in the crystal state by the intramolecular reaction of azide and alkyne groups, which releases strain that provides mechanical output.⁷⁹ They observed instant explosion of the crystals upon fast heating. In contrast, a silent crystal-to-crystal pathway occurs during the absorption of water molecules at a slow rate. Such phenomena were detected by a unique method using an electric circuit. Furthermore, they reported the thermal oligomerization of dipeptide to triazole-linked oligopeptides in the crystal state by the intermolecular reaction of azide and alkyne groups, which causes twisting of crystals.⁸⁰

3. Pseudorotaxane crystals

3.1 Thermal phase transition

With the idea of controlling the formation of pseudorotaxane by an electrochemical reaction, we prepared [2]pseudorotaxane

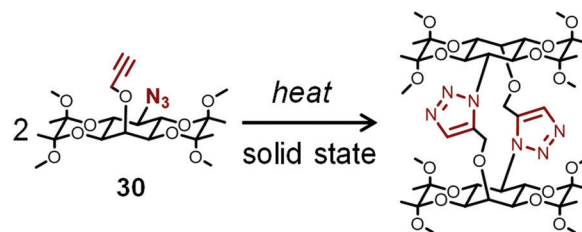


Fig. 8 Thermal dimerization of molecules containing azide and alkyne in the crystal state.

1 (Fig. 2), which is composed of ferrocenylmethyl(4-methylphenyl)ammonium as an axle and dibenzo[24]crown-8 ether (DB24C8) as a ring molecule.^{81,82} In the electrochemical reaction, the ferrocenyl group in the neutral axle precursor is oxidized to form an ammonium cation that interacts ionically with DB24C8.^{81,82}

The crystal-to-crystal phase transition of **1** was serendipitously discovered when the crystals were analyzed by differential scanning calorimetry (DSC) and polarized optical microscopy (Fig. 9).⁸³ The DSC scans of the crystals are shown in Fig. 9, in which endothermic and exothermic peaks at 128 °C and 115 °C, respectively, can be observed. Upon heating up to 190 °C, irreversible endothermic peaks appear around 160 °C and 170 °C, being the higher temperature peak assigned to the melting.⁸⁴ The crystal exhibits interference color change under cross-polarized light at 128 °C. The color switches from green to orange above 128 °C, to return to green at 115 °C on cooling. Although an irreversible change before the melting seemed to occur, the instability of this state prevented its analysis by X-ray crystallography at 160 °C. On the other hand, we were interested in unveiling the nature of the first peak at 128 °C, which is not due to melting.

Therefore, we conducted single-crystal X-ray crystallography to directly observe the molecular structure of **1**.^{83,84} As shown in Fig. 10 (left), **1** is stabilized by intramolecular ionic interactions between oxygen atoms of the crown ether and the ammonium cation of the axle molecule. In addition, intra- and intermolecular π - π interactions between the catechol group of the crown ether and the tolyl group (PhCH_3) in the axle molecule are also

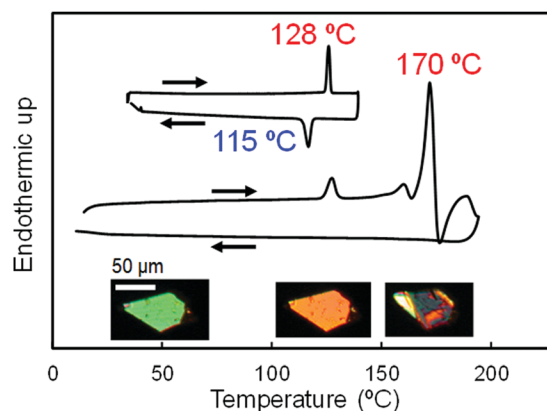


Fig. 9 DSC heating and cooling scans of **1** up to 130 °C (top) and 190 °C (bottom). Optical micrographs of a single crystal at 30 °C, 128 °C, and 160 °C under cross-polarized light are inserted.

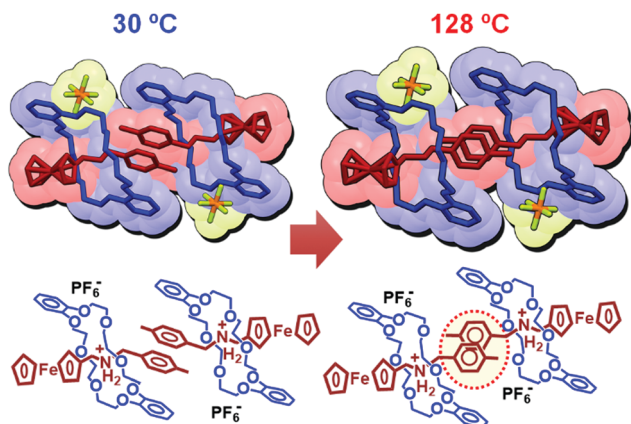


Fig. 10 Molecular structures of **1** obtained by single-crystal X-ray crystallography at 30 °C and 128 °C.

observed. At 128 °C, **1** shows a different structure from that at 30 °C and undergoes crystal-to-crystal thermal phase transition. This change in the molecular structure results from the twisting of a pair of tolyl groups of two axle molecules (Fig. 10, right), which changes the crystal anisotropy, leading to interference color changes (Fig. 9).

Such phase transition and reversible interference color changes were also observed in thin films of **1** deposited on substrates (Fig. 11).⁸⁵ Highly ordered thin films of **1** were prepared on glass or SiO₂ substrates *via* casting of chloroform solution followed by annealing. Grazing-incidence wide-angle X-ray scattering (GIWAXS) experiments indicated that **1** forms a highly ordered crystalline film during solvent evaporation. In addition, **1** has an edge-on orientation of the DB24C8 rings, which are perpendicular to the substrate, forming an interlocked pseudorotaxane structure with the axle molecules. Upon heating at 130 °C, the pseudorotaxane molecules reorganize to

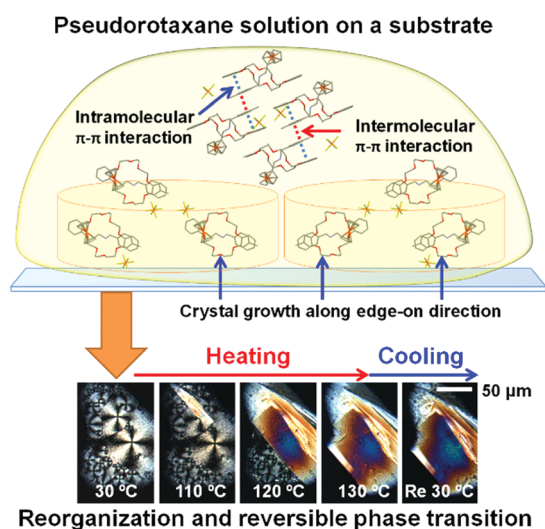


Fig. 11 Schematic illustration of the film formation and optical micrographs of a film of **1** under polarized light. The figure was modified using ref. 85 and reprinted with permission from the Royal Society of Chemistry, Copyright (2014).

form large domains with sizes in the range of 50–100 µm. These crystal domains exhibit reversible crystal-to-crystal thermal phase transition upon heating and cooling, which were observed *via in situ* GIWAXS and polarized optical microscopy.

We then explored the thermal properties of other ferrocene-containing pseudorotaxanes (Fig. 12), which are composed of AsF₆⁻ counter anions (**31**),⁸⁴ phenyl (**32**),⁸⁴ ethyl (**33**),⁸⁴ tetrabromo (**34**),⁸⁶ biferrocene (**35**),⁸⁶ biferrocene and tetrabromo (**36**),⁸⁶ [3]pseudorotaxane structure (**37**),⁸² or ruthenocene (**38**).⁸⁷ Among them, only **38** shows crystal-to-crystal phase transition.⁸⁷ In the crystal state, **38** has a similar alignment to that of **1** owing to π - π interactions between catechol and tolyl groups. The crystal-to-crystal thermal phase transition of **38** was observed at 86 °C and 55 °C on heating and cooling, respectively, which is significantly lower than for **1** by 42 °C and 60 °C, respectively. This is due to the strain provided by the bulky ruthenocene group in the unit cell. As a result, the formation of the high-temperature structure of the Ru complex occurs at lower temperature than for the Fe complex. Interestingly, these phase transition temperatures could be tailored by mixing **1** and **38** in different ratios during crystallization.⁸⁷

Mixed crystals of **1** and **31** also show crystal-to-crystal phase transition induced by the presence of π - π interactions between catechol and tolyl groups, whereas the twisted conformation of the tolyl group in pure **31** hinders such phase transition.⁸⁴ Crystals of **33**–**37** melt upon heating without the occurrence of crystal-to-crystal phase transition. In the case of **32** and **37**, this is most likely due to the twisted conformation of phenyl and

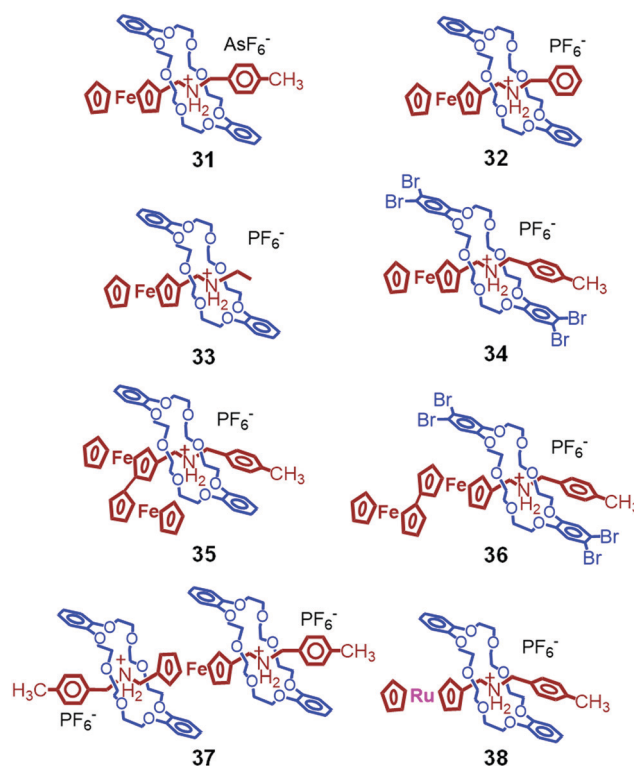


Fig. 12 Chemical structures of **31**–**38**. Only **38** shows crystal-to-crystal thermal phase transition.

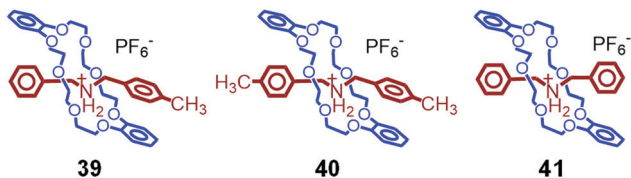


Fig. 13 Chemical structures of 39–41.

tolyl groups, respectively. On the other hand, 34 and 36 have tight intermolecular π - π interactions between tetrabromocatechol groups, possibly restricting the molecular motions for the thermal phase transition.

In addition, we investigated the thermal properties of crystals of various Stoddart-type pseudorotaxanes. Some examples are shown in Fig. 13.⁸⁸ However, these crystals only melt, and no crystal-to-crystal thermal phase transitions are observed. From these results, we concluded that the presence of (i) a ferrocenyl group, (ii) π - π interactions between the catechol and tolyl groups, (iii) PF_6^- anions, and (iv) DB24C8 (not tetrabromo-DB24C8) is required for the crystal-to-crystal thermal phase transition to occur.

3.2 Photoinduced molecular motions

Next, in our investigation of pseudorotaxane crystals, we applied photoirradiation to induce molecular motions in the crystal state. These experiments were conducted using a polarized optical microscope equipped with a $\times 20$ objective lens, a continuous-wave 445 nm diode laser, a hot stage, top- and side-

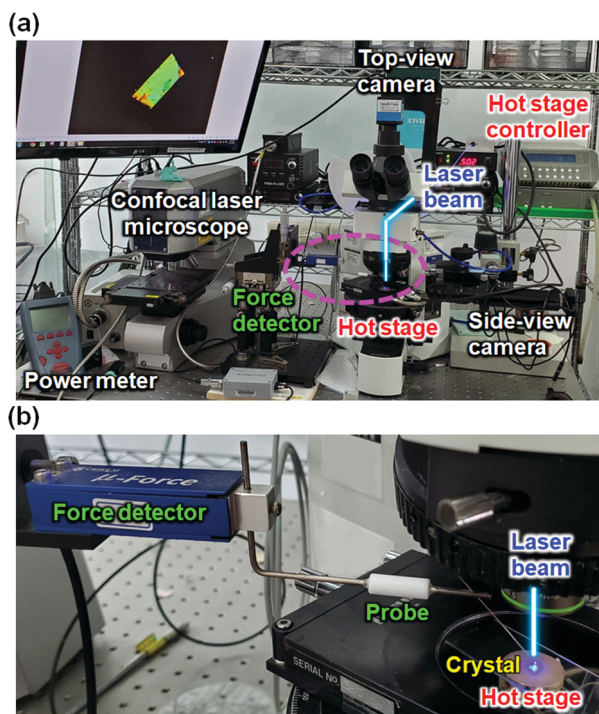


Fig. 14 Photographs of the setup for the photoirradiation experiments. (a) Overall view and (b) expanded view around the crystal sample. A laser beam line is illustrated.

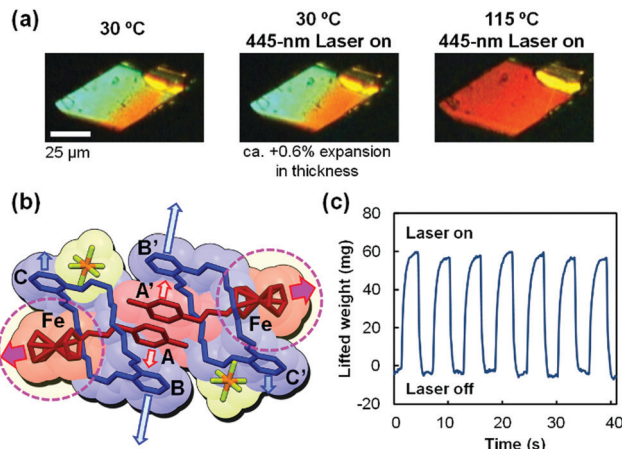


Fig. 15 (a) Optical micrographs of a single crystal of **1** under polarized light. (b) Molecular structure of **1** and its photoinduced motion by 445 nm laser irradiation. (c) Time dependence of lifting weight of the crystal induced by 445 nm focused laser irradiation at 30 °C.

view CCD cameras, and a force detector (Fig. 14).^{87,89} The crystal dimensions were measured by a confocal laser microscope with/without laser irradiation.

Upon subjecting a single crystal of **1** to irradiation by a focused 445 nm laser at 30 °C, the crystal exhibits slightly subtractive interference color compared with the nonirradiated crystal (Fig. 15a, left and middle).⁸⁹ This change is accompanied by a quick anisotropic deformation of the crystal; its length, width, and thickness increase by 0.2%, 1.0%, and 0.6%, respectively. After turning off the laser, the crystal immediately returns to the original interference color and size. When the crystal is heated at 115 °C, a much lower temperature than that of the thermal phase transition temperature (128 °C), a photo-triggered crystal-to-crystal phase transition is observed under focused laser irradiation at 445 nm (Fig. 15a, right). The significant interference color change from green to red is indicative of the formation of a high-temperature phase. Since the laser beam spot (*ca.* 1 μm diameter) was much smaller than the top face of the crystal, the local excitation is rapidly propagated by concerted molecular structural changes in the entire crystal.

The photoinduced molecular structural changes in **1** were analyzed by single-crystal X-ray crystallography with 445 nm laser irradiation at 30 °C (Fig. 15b).⁸⁹ We found that the photoinduced three-dimensional expansion of the crystal is related to the elongation of the distances A-A', B-B', and A-Fe. These molecular structural changes are different from the heat-induced motions of **1**. Upon heating at 105 °C in the absence of laser irradiation, the distances A-A' and B-B' increase, whereas A-Fe slightly decreases. Thus, we concluded that the elongation of the A-Fe distance is due to the photo-activation of the ferrocenyl group leading to the expansion of the whole crystal.

The force produced by the photoinduced expansion of the crystal of **1** was monitored using a microforce analyzer containing two strain gauges. As shown in Fig. 15c, the crystal can lift

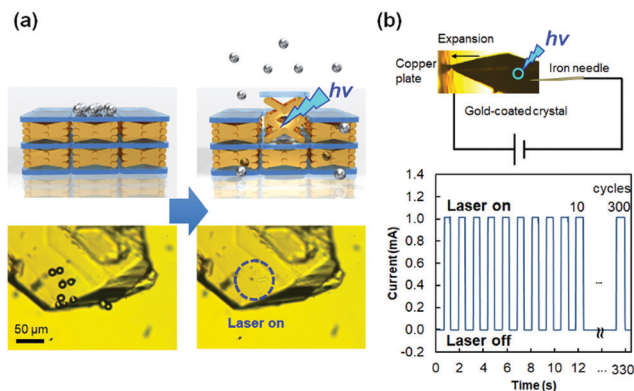


Fig. 16 (a) Transport of silica microparticles induced by expansion of a crystal of **1** triggered by focused 445 nm laser irradiation. (b) Electric current switching controlled by deformation of a gold-coated crystal of **1** with/without 445 nm laser irradiation.

ca. 68 mg under laser irradiation. The crystal weight was estimated to be 41 μg from its volume by confocal laser microscopy and from its density by single-crystal X-ray crystallography. Therefore, the ratio (lifted weight)/(crystal weight) was estimated to be 1650.⁸⁹ Such mechanical force produced by a crystal was used for the demonstration of transport of silica microparticles (Fig. 16a). According to this experiment, silica microparticles placed on a crystal of **1** jump from the crystal upon irradiation by a 445 nm laser, due to the quick expansion of the crystal. In addition, a crystal of **1** can be used as a mechanical microswitch to control an electric circuit (Fig. 16b).⁸⁹ Thus, a gold-coated crystal of **1** and a nearby copper plate are introduced in an electric circuit, and the current is remotely controlled by turning the laser on and off. This is accompanied by connection and disconnection of the crystal with the copper plate due to the expansion/contraction motion of the crystal. The current switching is robust and repeated 300 times in 330 s (0.91 times s^{-1} , Fig. 16b) or 560 times in 60 s (9.3 times s^{-1}).⁸⁹ This switching is much faster and more facile than the current switching using diarylethene crystals,⁵⁶ because the expansion/contraction motion of the crystal of **1** only requires one laser wavelength to change the state; moreover, the backward process is spontaneous.

Next, the magnitude of the photoinduced crystal expansion of **1** was compared with that of other molecules. The laser

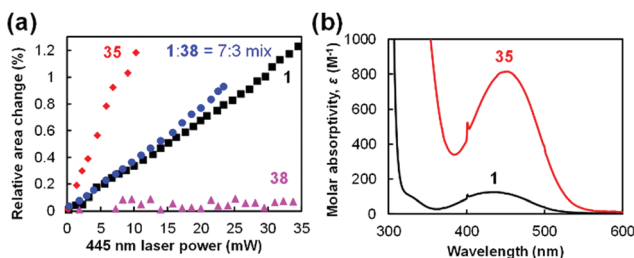


Fig. 17 (a) Laser power dependence of the change in the relative area of the top face of the single crystals induced by focused 445 nm laser irradiation. (b) UV-visible absorption spectra of mixed crystals in chloroform solutions.

power dependence of the relative area change of the top face of crystals of **1**, **35**, and **38** and a mixed crystal **1** : **38** in a 7 : 3 molar ratio is plotted in Fig. 17a. As can be seen, the crystal of biferrocenyl pseudorotaxane **35** shows the highest sensitivity to the 445 nm laser light (+1.2% expansion at 10 mW), whereas **1** shows +0.35% expansion at the same laser power.⁸⁶ The mixed crystal shows a slightly higher magnitude of expansion (+0.8%) than **1** at a laser power of 10 mW.⁸⁷ In contrast, the change of **38** is negligible (<+0.1%).⁸⁷

The magnitude of the crystal expansion can be accounted for by the photoabsorption of the molecules. Fig. 17b shows the UV-vis absorption spectra of **1** and **35**. The maximum absorption wavelength (λ_{max}) of **1** is observed at 433 nm with a molar absorptivity (ϵ) of 125 $\text{M}^{-1} \text{cm}^{-1}$, whereas **35** shows $\lambda_{\text{max}} = 453 \text{ nm}$ with $\epsilon = 810 \text{ M}^{-1} \text{cm}^{-1}$ (Fig. 17b).⁸⁶ These peaks are attributed to the absorption of the ferrocenyl or biferrocenyl group. Therefore, **35** shows a much higher sensitivity than **1** in the relative area changes. On the other hand, no photoabsorption around 400–500 nm is observed for ruthenocene complex **38**, leading to significantly low photosensitivity in the relative area changes.⁸⁷

Interestingly, the mixed crystal **1** : **38** shows larger expansion of the crystal than **1**, even though the absorbance of the former crystal is lower than that of the latter.⁸⁷ This might be explained by the crystal packing structures. The structures of the mixed crystal are less packed than those of the crystals of pure **1** and **38**, which results in the pseudorotaxane molecules in the mixed crystal having larger free volume, molecular flexibility, and higher dynamics.

The photoinduced weight lifting by crystals of biferrocene and tetrabromo complexes **34–36** and ruthenocene complex **38** was also measured by using the microforce analyzer. When a crystal of **35** is irradiated by a 445 nm laser, the expansion of the crystal produces a ratio (lifted weight)/(crystal weight) of 2900, which is nearly two times higher than that of **1**.⁸⁶ However, crystals of tetrabromo complexes **34** and **36** afford a much lower weight lifted, possibly due to the tight molecular packing derived from strong π - π interactions between the tetrabromocatechol rings.⁸⁶

Similar to **1**, the mixed crystal **1** : **38** produces a ratio (lifted weight)/(crystal weight) of 1600.⁸⁷ In contrast, a negative force is observed for this crystal when it is irradiated by a 405 nm laser at 85 °C, due to a decrease in the thickness of the crystal that accompanied the laser-assisted thermal phase transition. When the laser is switched off, the crystal returns to the low-temperature phase producing a strong force corresponding to a ratio (lifted weight)/(crystal weight) of 6400.⁸⁷ Under the same conditions, **1** does not undergo such phase transition due to its much higher thermal phase transition temperature than that of the mixed crystal. The observed (lifted weight)/(crystal weight) ratio resulting from the laser-assisted thermal phase transition is 1–2 orders of magnitude higher than those of previously reported materials, such as diarylethene crystals,^{55,90} azobenzene polymer films,^{91,92} polymer composite actuators,⁹³ and electro-triggered hydrogels.⁹⁴

On the other hand, crystals of **39–41** show insignificant changes by photoirradiation at 445 nm due to the lack of photoabsorbing groups in their molecules.

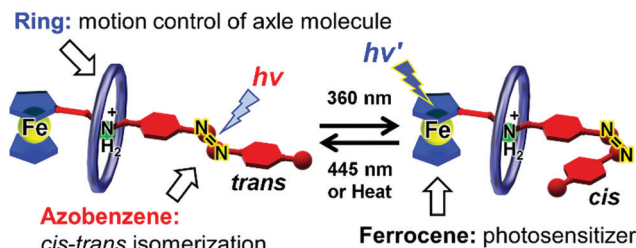


Fig. 18 Illustration of *cis-trans* isomerization of azobenzene- and ferrocene-containing pseudorotaxanes.

3.3 Azobenzene-containing pseudorotaxanes

Recently, we introduced azobenzene groups into ferrocene-containing [2]pseudorotaxanes.⁹⁵ Fig. 18 shows a conceptual illustration. The azobenzene group undergoes reversible *cis-trans* photoisomerization, transducing photoenergy into mechanical bending motion. Using pseudorotaxanes, such motions are expected to be flexible and diverse. In addition, pseudorotaxanes introduce structural diversity because they consist of various components.

Fig. 19a shows the chemical structures of azobenzene- and ferrocene-containing [2]pseudorotaxanes **42–44**.⁹⁵ Their molecular structures in a crystal state were determined by single-crystal X-ray crystallography. Despite having very similar chemical structures, their intra- and intermolecular interactions

in the crystal state differ largely from one another. Thus, **42**, which lacks a substituent on the 4-position of the azobenzene group, adopts strong intermolecular π - π interactions between a pair of aromatic rings of the azobenzene groups. These azobenzene groups have a planar *trans*-azobenzene structure. In contrast, methyl-substituted **43** has intramolecular π - π interactions between azobenzene and catechol groups. Therefore, one of the aromatic rings of the azobenzene group is relatively free, and thus, the azobenzene group has a twisted structure. On the other hand, **44**, which contains tetrabromo-substituted DB24C8, shows only one intermolecular π - π interaction between one of the aromatic rings of the azobenzene group. As a result, the azobenzene group in **44** is also expected to be flexible.

The structures of these crystals have a significant influence on their molecular motions and, consequently, on their crystal motions. Thus, a crystal of **42** remains immobile when it is irradiated at 445 nm, whereas it vibrates slightly upon irradiation by a 360 nm laser. In contrast, crystals of **43** and **44** show significant downward bending under 445 nm irradiation (Fig. 19b).⁹⁵ These bended crystals undergo slow upward bending under 360 nm laser irradiation.

These crystal bending motions are related to the *cis-trans* photoisomerization of the azobenzene groups. In CH_2Cl_2 solutions, all [2]pseudorotaxanes **42–44** show a strong absorption peak at 324 nm and a weak peak around 450 nm attributable to *trans*-azobenzene and ferrocene, respectively. The peak at 324 nm decreases significantly after irradiation at 360 nm due to the formation of *cis*-azobenzene. Thus, the photoisomerization of the azobenzene group is very efficient for all [2]pseudorotaxanes **42–44** in solution. To estimate their photoisomerization, single crystals were irradiated at 360 nm and dissolved in CH_2Cl_2 to measure the corresponding UV-vis spectra. On the basis of the absorption intensities, the photoisomerization conversion was estimated to be 4% for **42**, 28% for **43**, and 40% for **44** in their crystal state.⁹⁵

The low photoisomerization conversion for **42** is due to the stable planar *trans*-azobenzene group with strong intermolecular π - π interactions, which restricts the molecular motions and immobilizes the crystals. In contrast, the decent degrees of photoisomerization conversion for **43** and **44** can be explained in terms of the twisted azobenzene structures and weak π - π interactions providing significant flexibility of the azobenzene group, which results in macroscopic crystal bending.

It is worth mentioning that the crystals undergo a fast *cis-trans* backward process within 0.3 s, which contrasts with the fast forward and slow backward isomerizations previously reported for azobenzene derivatives.^{45,96} Such fast backward process is probably due to the presence of the ferrocenyl group, which absorbs at 445 nm converting photoenergy into heat, thereby accelerating the *cis-trans* backward process.

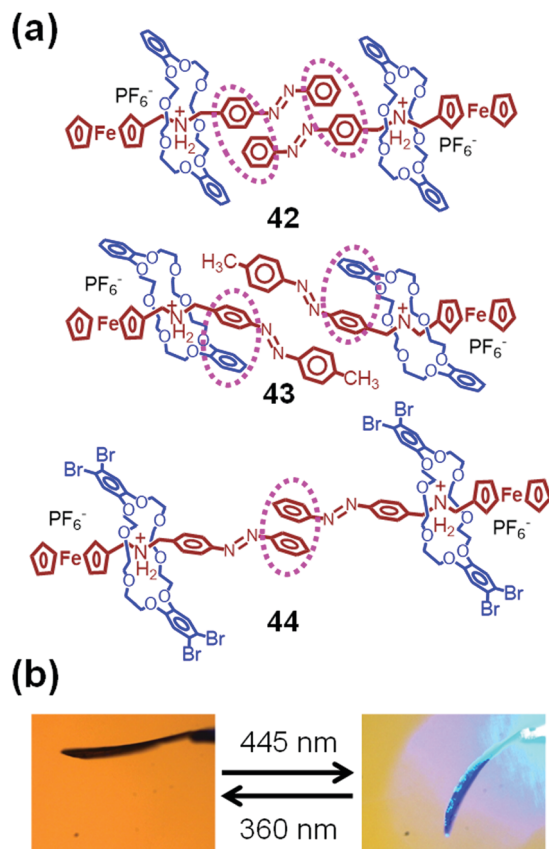


Fig. 19 (a) Chemical structures of **42–44**. (b) Optical micrographs of **44** under 360 nm (left) and 445 nm (right) laser irradiation.

4. Conclusions and perspectives

This review covers recent literature on stimuli-responsive dynamic molecular crystals comprising MOFs, molecular rotors,

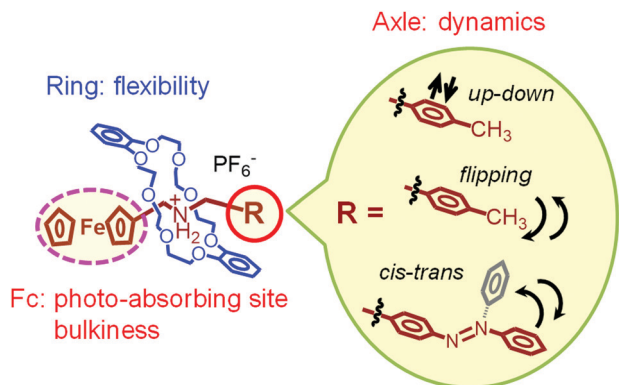


Fig. 20 Schematic illustration of the molecular motions and functions of ferrocene-containing pseudorotaxanes.

photochromic molecules such as azobenzenes and diarylethenes, cycloaddition reactive units, and pseudorotaxanes. To achieve light- and/or heat-responsive dynamic crystal motions, efficient reactive sites, free volume, and mobile parts are needed. Rotaxanes and pseudorotaxanes have complicated interlocked structures composed of ring and axle molecules. Consequently, reports on the direct observation of molecular structural changes in rotaxanes and pseudorotaxanes by single-crystal X-ray crystallography are limited. The pseudorotaxanes investigated in our group overcome such limitation.

The development of dynamic rotaxane and pseudorotaxane crystals is still challenging and would be a pioneering step in the study of solid-state molecular motions, compared with other sophisticated dynamic molecular systems. To develop sophisticated crystal molecular machines using rotaxanes and pseudorotaxanes, fundamental studies of the systematic tailoring of the molecular structures and packing properties of these molecules in the crystal state are required.

Fig. 20 summarizes the molecular motions and functions of our pseudorotaxanes in the crystal state. The ring molecule in the pseudorotaxane structure provides molecular flexibility by adjusting its structure with the structural changes of the axle molecule. The latter, in turn, undergoes dynamics, such as up-down, flipping, and *cis-trans* isomerizations induced by heat and/or light. Meanwhile, photoabsorption by the organometallic ferrocenyl group converts photoenergy into thermal energy, leading to photomechanical conversion in the crystals. In addition, the bulkiness of the ferrocenyl group is an important factor to provide molecular dynamics in the crystal state, since it provides flexible molecular packing structures.

We believe that this review paves the way for new findings in molecular machines and rotaxanes with interesting novel motions in the crystal state. In addition, we expect that such stimuli-responsive rotaxane crystals will be used as light-sensing and harvesting materials in photomechanical microdevices and switches, such as miniaturized optical circuit elements, miniaturized chemical and biosensing tools, and future crystal actuators.

Conflicts of interest

There are no conflicts to declare.

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References

- 1 S. Shinkai, *Pure Appl. Chem.*, 1987, **59**, 425–430.
- 2 J. F. Stoddart, *Angew. Chem., Int. Ed.*, 2017, **56**, 11094–11125.
- 3 B. L. Feringa, *Angew. Chem., Int. Ed.*, 2017, **56**, 11060–11078.
- 4 J. P. Sauvage, *Angew. Chem., Int. Ed.*, 2017, **56**, 11080–11093.
- 5 A. Coskun, M. Banaszak, R. D. Astumian, J. F. Stoddart and B. A. Grzybowski, *Chem. Soc. Rev.*, 2012, **41**, 19–30.
- 6 C. J. Bruns and J. F. Stoddart, *Acc. Chem. Res.*, 2014, **47**, 2186–2199.
- 7 R. Barat, T. Legigan, I. Tranoy-Opalinski, B. Renoux, E. Péraudeau, J. Clarhaut, P. Poinot, A. E. Fernandes, V. Aucagne, D. A. Leigh and S. Papot, *Chem. Sci.*, 2015, **6**, 2608–2613.
- 8 C. Y. Cheng, P. R. McGonigal, S. T. Schneebeli, H. Li, N. A. Vermeulen, C. F. Ke and J. F. Stoddart, *Nat. Nanotechnol.*, 2015, **10**, 547–553.
- 9 S. Erbas-Cakmak, D. A. Leigh, C. T. McTernan and A. L. Nussbaumer, *Chem. Rev.*, 2015, **115**, 10081–10206.
- 10 H. K. Bisoyi and Q. Li, *Chem. Rev.*, 2016, **116**, 15089–15166.
- 11 S. Erbas-Cakmak, S. D. P. Fielden, U. Karaca, D. A. Leigh, C. T. McTernan, D. J. Tetlow and M. R. Wilson, *Science*, 2017, **358**, 340–343.
- 12 A. Martinez-Cuezva, A. Saura-Sanmartin, T. Nicolas-Garcia, C. Navarro, R.-A. Orenes, M. Alajarin and J. Berna, *Chem. Sci.*, 2017, **8**, 3775–3780.
- 13 S. Kassem, T. van Leeuwen, A. S. Lubbe, M. R. Wilson, B. L. Feringa and D. A. Leigh, *Chem. Soc. Rev.*, 2017, **46**, 2592–2621.
- 14 K. Zhu, G. Baggi and S. J. Loeb, *Nat. Chem.*, 2018, **10**, 625–630.
- 15 D. Sluysmans, F. Devaux, C. J. Bruns, J. F. Stoddart and A.-S. Duwez, *Proc. Natl. Acad. Sci. U. S. A.*, 2018, **115**, 9362–9366.
- 16 J. J. Yu, L. Y. Zhao, Z. T. Shi, Q. Zhang, G. London, W. J. Liang, C. Gao, M. M. Li, X. M. Cao, H. Tian, B. L. Feringa and D. H. Qu, *J. Org. Chem.*, 2019, **84**, 5790–5802.
- 17 K.-Y. Chen, O. Ivashenko, G. T. Carroll, J. Robertus, J. C. M. Kistemaker, G. London, W. R. Browne, P. Rudolf and B. L. Feringa, *J. Am. Chem. Soc.*, 2014, **136**, 3219–3224.
- 18 L. Greb and J.-M. Lehn, *J. Am. Chem. Soc.*, 2014, **136**, 13114–13117.
- 19 S. Kassem, A. T. L. Lee, D. A. Leigh, A. Markevicius and J. Sola, *Nat. Chem.*, 2016, **8**, 138–143.
- 20 S. K. M. Nalluri, J. W. Zhou, T. Cheng, Z. C. Liu, M. T. Nguyen, T. Y. Chen, H. A. Patel, M. D. Krzyaniak, W. A. Goddard, M. R. Wasielewski and J. F. Stoddart, *J. Am. Chem. Soc.*, 2019, **141**, 1290–1303.
- 21 G. Rapenne and C. Joachim, *Nat. Rev. Mater.*, 2017, **2**, 17040.
- 22 T. Ikeda and T. Ube, *Mater. Today*, 2011, **14**, 480–487.
- 23 M. D. Manrique-Juárez, S. Rat, L. Salmon, G. Molnár, C. M. Quintero, L. Nicu, H. J. Shepherd and A. Bousseksou, *Coord. Chem. Rev.*, 2016, **308**, 395–408.

- 24 Y. Hu, Z. Li, T. Lan and W. Chen, *Adv. Mater.*, 2016, **28**, 10548–10556.
- 25 M. A. Watson and S. L. Cockroft, *Chem. Soc. Rev.*, 2016, **45**, 6118–6129.
- 26 N. K. Nath, M. K. Panda, S. C. Sahoo and P. Naumov, *CrystEngComm*, 2014, **16**, 1850–1858.
- 27 O. Sato, *Nat. Chem.*, 2016, **8**, 644–656.
- 28 O. S. Bushuyev, T. Friščić and C. J. Barrett, *CrystEngComm*, 2016, **18**, 7204–7211.
- 29 H. X. Deng, M. A. Olson, J. F. Stoddart and O. M. Yaghi, *Nat. Chem.*, 2010, **2**, 439–443.
- 30 A. Coskun, M. Hmadeh, G. Barin, F. Gándara, Q. Li, E. Choi, N. L. Strutt, D. B. Cordes, A. M. Z. Slawin, J. F. Stoddart, J.-P. Sauvage and O. M. Yaghi, *Angew. Chem., Int. Ed.*, 2012, **51**, 2160–2163.
- 31 K. Zhu, C. A. O’Keefe, V. N. Vukotic, R. W. Schurko and S. J. Loeb, *Nat. Chem.*, 2015, **7**, 514.
- 32 N. Farahani, K. L. Zhu, C. A. O’Keefe, R. W. Schurko and S. J. Loeb, *ChemPlusChem*, 2016, **81**, 836–841.
- 33 V. N. Vukotic, K. Zhu, G. Baggi and S. J. Loeb, *Angew. Chem.*, 2017, **129**, 6232–6237.
- 34 D. K. Shen, G. Wang, Z. C. Liu, P. Li, K. Cai, C. Y. Cheng, Y. Shi, J. M. Han, C. W. Kung, X. R. Gong, Q. H. Guo, H. L. Chen, A. C. H. Sue, Y. Y. Botros, A. Facchetti, O. K. Farha, T. J. Marks and J. F. Stoddart, *J. Am. Chem. Soc.*, 2018, **140**, 11402–11407.
- 35 X. Jiang, Z. J. O’Brien, S. Yang, L. H. Lai, J. Buenaflor, C. Tan, S. Khan, K. N. Houk and M. A. Garcia-Garibay, *J. Am. Chem. Soc.*, 2016, **138**, 4650–4656.
- 36 M. E. Howe and M. A. Garcia-Garibay, *J. Org. Chem.*, 2019, **84**, 9835–9849.
- 37 P. Commins and M. A. Garcia-Garibay, *J. Org. Chem.*, 2014, **79**, 1611–1619.
- 38 W. Setaka and K. Yamaguchi, *J. Am. Chem. Soc.*, 2013, **135**, 14560–14563.
- 39 A. Fujiwara, Y. Inagaki, H. Momma, E. Kwon, K. Yamaguchi, M. Kanno, H. Kono and W. Setaka, *CrystEngComm*, 2017, **19**, 6049–6056.
- 40 M. Tsurunaga, Y. Inagaki, H. Momma, E. Kwon, K. Yamaguchi, K. Yoza and W. Setaka, *Org. Lett.*, 2018, **20**, 6934–6937.
- 41 Y. L. Yu, M. Nakano and T. Ikeda, *Nature*, 2003, **425**, 145.
- 42 N. Hosono, T. Kajitani, T. Fukushima, K. Ito, S. Sasaki, M. Takata and T. Aida, *Science*, 2010, **330**, 808–811.
- 43 T. Ube and T. Ikeda, *Angew. Chem., Int. Ed.*, 2014, **53**, 10290–10299.
- 44 H. Koshima, N. Ojima and H. Uchimoto, *J. Am. Chem. Soc.*, 2009, **131**, 6890–6891.
- 45 T. Taniguchi, J. Fujisawa, M. Shiro, H. Koshima and T. Asahi, *Chem. – Eur. J.*, 2016, **22**, 7950–7958.
- 46 T. Taniguchi, H. Sugiyama, H. Uekusa, M. Shiro, T. Asahi and H. Koshima, *Nat. Commun.*, 2018, **9**, 538.
- 47 O. S. Bushuyev, T. C. Corkery, C. J. Barrett and T. Friščić, *Chem. Sci.*, 2014, **5**, 3158–3164.
- 48 O. S. Bushuyev, A. Tomberg, T. Friščić and C. J. Barrett, *J. Am. Chem. Soc.*, 2013, **135**, 12556–12559.
- 49 Y. Norikane, S. Tanaka and E. Uchida, *CrystEngComm*, 2016, **18**, 7225–7228.
- 50 E. Uchida, R. Azumi and Y. Norikane, *Nat. Commun.*, 2015, **6**, 7310.
- 51 T. Shima, T. Muraoka, N. Hoshino, T. Akutagawa, Y. Kobayashi and K. Kinbara, *Angew. Chem., Int. Ed.*, 2014, **53**, 7173–7178.
- 52 P. Gupta, T. Panda, S. Allu, S. Borah, A. Baishya, A. Gunnam, A. Nangia, P. Naumov and N. K. Nath, *Cryst. Growth Des.*, 2019, **19**, 3039–3044.
- 53 P. Gupta, D. P. Karothu, E. Ahmed, P. Naumov and N. K. Nath, *Angew. Chem.*, 2018, **130**, 8634–8638.
- 54 M. Irie, T. Fulcaminato, K. Matsuda and S. Kobatake, *Chem. Rev.*, 2014, **114**, 12174–12277.
- 55 F. Terao, M. Morimoto and M. Irie, *Angew. Chem., Int. Ed.*, 2012, **51**, 901–904.
- 56 D. Kitagawa and S. Kobatake, *Chem. Commun.*, 2015, **51**, 4421–4424.
- 57 D. Kitagawa, K. Kawasaki, R. Tanaka and S. Kobatake, *Chem. Mater.*, 2017, **29**, 7524–7532.
- 58 D. Kitagawa, H. Tsujioka, F. Tong, X. N. Dong, C. J. Bardeen and S. Kobatake, *J. Am. Chem. Soc.*, 2018, **140**, 4208–4212.
- 59 X. M. Dong, F. Tong, K. M. Hanson, R. O. Al-Kaysi, D. Kitagawa, S. Kobatake and C. J. Bardeen, *Chem. Mater.*, 2019, **31**, 1016–1022.
- 60 S. Ishida, D. Kitagawa, S. Kobatake, S. Kim, S. Kurihara and T. Fukaminato, *Chem. Commun.*, 2019, **55**, 5681–5684.
- 61 K. Biradha and R. Santra, *Chem. Soc. Rev.*, 2013, **42**, 950–967.
- 62 V. Ramamurthy and J. Sivaguru, *Chem. Rev.*, 2016, **116**, 9914–9993.
- 63 J. J. Vittal and H. S. Quah, *Dalton Trans.*, 2017, **46**, 7120–7140.
- 64 H. Ying-Feng, J. Guo-Xin, C. G. Daniliuc and H. F. Ekkehardt, *Angew. Chem., Int. Ed.*, 2015, **54**, 4958–4962.
- 65 R. Medishetty, Z. Bai, H. Yang, M. W. Wong and J. J. Vittal, *Cryst. Growth Des.*, 2015, **15**, 4055–4061.
- 66 P. In-Hyeok, M. Raghavender, L. Hyeong-Hwan, M. C. Evania, Q. H. Sheng, L. S. Sung and J. J. Vittal, *Angew. Chem., Int. Ed.*, 2015, **54**, 7313–7317.
- 67 J.-M. Chen, Y.-X. Hou, Q.-K. Zhou, H. Zhang and D. Liu, *CrystEngComm*, 2017, **19**, 2603–2607.
- 68 Y. X. Shi, W. H. Zhang, B. F. Abrahams, P. Braunstein and J. P. Lang, *Angew. Chem., Int. Ed.*, 2019, **58**, 9453–9458.
- 69 R. Medishetty, A. Husain, Z. Bai, T. Runčevski, R. E. Dinnebier, P. Naumov and J. J. Vittal, *Angew. Chem., Int. Ed.*, 2014, **53**, 5907–5911.
- 70 K. Yadava and J. J. Vittal, *Cryst. Growth Des.*, 2019, **19**, 2542–2547.
- 71 C. E. Mulijanto, H. S. Quah, G. K. Tan, B. Donnadiu and J. J. Vittal, *IUCrJ*, 2017, **4**, 65–71.
- 72 R. Medishetty, S. C. Sahoo, C. E. Mulijanto, P. Naumov and J. J. Vittal, *Chem. Mater.*, 2015, **27**, 1821–1829.
- 73 H. Wang, P. Chen, Z. Wu, J. Zhao, J. Sun and R. Lu, *Angew. Chem., Int. Ed.*, 2017, **56**, 9463–9467.
- 74 J. Peng, K. Q. Ye, C. Liu, J. B. Sun and R. Lu, *J. Mater. Chem. C*, 2019, **7**, 5433–5441.

- 75 S. Z. Li and D. P. Yan, *ACS Appl. Mater. Interfaces*, 2018, **10**, 22703–22710.
- 76 T. Kim, L. Zhu, L. J. Mueller and C. J. Bardeen, *J. Am. Chem. Soc.*, 2014, **136**, 6617–6625.
- 77 F. Tong, W. J. Xu, M. Al-Haidar, D. Kitagawa, R. O. Al-Kaysi and C. J. Bardeen, *Angew. Chem., Int. Ed.*, 2018, **57**, 7080–7084.
- 78 H. Koshima, H. Uchimoto, T. Taniguchi, J. Nakamura, T. Asahi and T. Asahi, *CrystEngComm*, 2016, **18**, 7305–7310.
- 79 A. Ravi and K. M. Sureshan, *Angew. Chem., Int. Ed.*, 2018, **57**, 9362–9366.
- 80 R. Rai, B. P. Krishnan and K. M. Sureshan, *Proc. Natl. Acad. Sci. U. S. A.*, 2018, **115**, 2896–2901.
- 81 M. Horie, Y. Suzaki and K. Osakada, *J. Am. Chem. Soc.*, 2004, **126**, 3684–3685.
- 82 M. Horie, Y. Suzaki and K. Osakada, *Inorg. Chem.*, 2005, **44**, 5844–5853.
- 83 M. Horie, T. Sassa, D. Hashizume, Y. Suzaki, K. Osakada and T. Wada, *Angew. Chem., Int. Ed.*, 2007, **46**, 4983–4986.
- 84 M. Horie, Y. Suzaki, D. Hashizume, T. Abe, T. Wu, T. Sassa, T. Hosokai and K. Osakada, *J. Am. Chem. Soc.*, 2012, **134**, 17932–17944.
- 85 Y.-C. Tsai, K.-J. Chen, C.-J. Su, W.-R. Wu, U. S. Jeng and M. Horie, *J. Mater. Chem. C*, 2014, **2**, 2061–2068.
- 86 K. J. Chen, A. C. Tan, C. H. Wang, T. S. Kuo, P. L. Chen and M. Horie, *Cryst. Growth Des.*, 2019, **19**, 17–22.
- 87 K.-J. Chen, P.-L. Chen and M. Horie, *Sci. Rep.*, 2017, **7**, 14195.
- 88 P. R. Ashton, M. C. T. Fyfe, S. K. Hickingbottom, J. F. Stoddart, A. J. P. White and D. J. Williams, *J. Chem. Soc., Perkin Trans. 2*, 1998, 2117–2128.
- 89 K.-J. Chen, Y.-C. Tsai, Y. Suzaki, K. Osakada, A. Miura and M. Horie, *Nat. Commun.*, 2016, **7**, 13321.
- 90 M. Morimoto and M. Irie, *J. Am. Chem. Soc.*, 2010, **132**, 14172–14178.
- 91 F. T. Cheng, R. Y. Yin, Y. Y. Zhang, C. C. Yen and Y. L. Yu, *Soft Matter*, 2010, **6**, 3447–3449.
- 92 L. D. Zhang, H. R. Liang, J. Jacob and P. Naumov, *Nat. Commun.*, 2015, **6**, 7429.
- 93 M. M. Ma, L. Guo, D. G. Anderson and R. Langer, *Science*, 2013, **339**, 186–189.
- 94 C. Yang, W. Wang, C. Yao, R. Xie, X. J. Ju, Z. Liu and L. Y. Chu, *Sci. Rep.*, 2015, **5**, 13622.
- 95 S.-C. Cheng, K.-J. Chen, Y. Suzaki, Y. Tsuchido, T.-S. Kuo, K. Osakada and M. Horie, *J. Am. Chem. Soc.*, 2018, **140**, 90–93.
- 96 S. Guo, K. Matsukawa, T. Miyata, T. Okubo, K. Kuroda and A. Shimojima, *J. Am. Chem. Soc.*, 2015, **137**, 15434–15440.