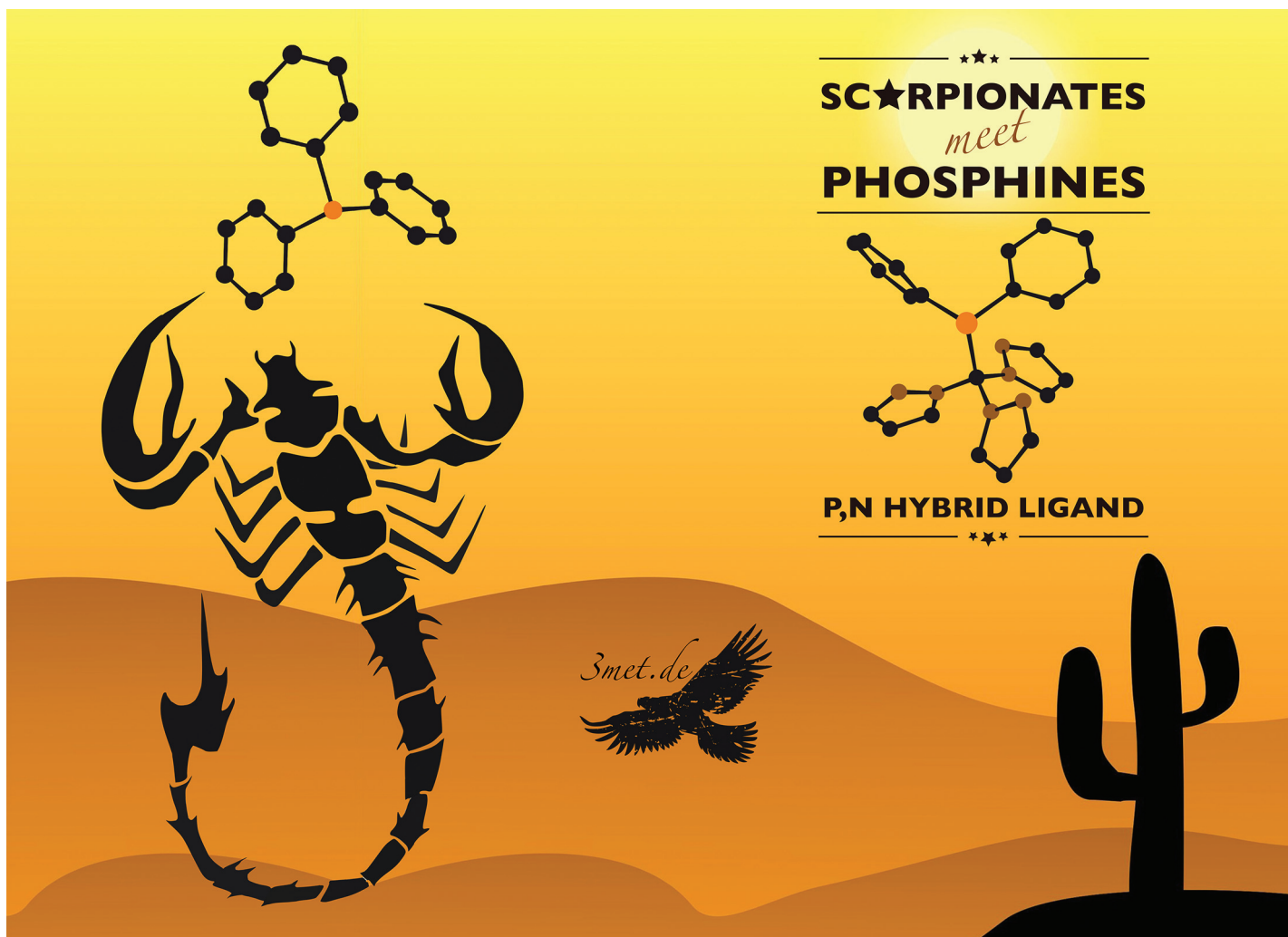


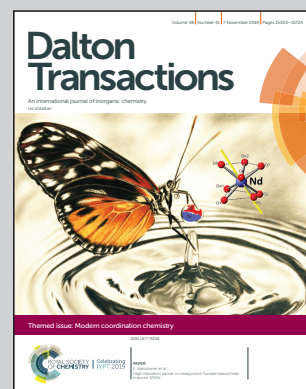


Showcasing research from the group of Frank Breher from the Karlsruhe Institute of Technology (KIT).

Phosphine-functionalised tris(pyrazolyl)methane ligands and their mono- and heterobimetallic complexes

The synthesis, characterisation and reactivity of phosphine-functionalised tris(pyrazolyl)methane ligands are presented. The new P,N hybrid ligands feature characteristics of both "classical" scorpionates and phosphine ligands. Their coordination flexibility and their ability to generate heterobimetallic complexes has been demonstrated.

As featured in:



See Frank Breher *et al.*,
Dalton Trans., 2019, 48, 15397.



Cite this: *Dalton Trans.*, 2019, **48**, 15397

Phosphine-functionalised tris(pyrazolyl)methane ligands and their mono- and heterobimetallic complexes†‡

Hanna E. Wagner, Silvia Hohnstein, Max G. Schußmann, Lukas A. Steppe and Frank Breher *

The synthesis, characterisation and reactivity of new phosphine-functionalised tris(pyrazolyl)methane ligands (TpmPR_2 , **2a–c**, with $\text{R} = \text{Ph}$, $n\text{Bu}$, $i\text{Pr}$) are presented. The reaction of **2a–c** with $[\text{Rh}(\text{CO})_2\text{Cl}]_2$ furnished N,P -heterochelate carbonyl complexes (**3a–c**), which were used to quantify the donor abilities of the ligands via IR spectroscopy. The coordination flexibility was demonstrated by treating representative members of this new ligand class with $[\text{CpRu}(\text{acn})_3][\text{PF}_6]$ ($\text{acn} = \text{acetonitrile}$), $[(\text{tht})\text{AuCl}]$ ($\text{tht} = \text{tetrahydrothiophene}$) and $[\text{Pd}(\text{allyl})\text{Cl}]_2$ providing either a N,N,P -heteroscorpionate complex (**4**) or the P -coordinated complexes (**5,6**) without any involvement of the pyrazolyl entities. With **2a** and $[\text{Pd}(\text{cod})\text{Cl}_2]$, another P,N -heterochelate complex (**7**) was obtained, which served as a precursor for a heterobimetallic complex containing palladium and copper (**8**). Detailed NMR spectroscopic and X-ray crystallographic investigations have been performed on all new complexes.

Received 16th May 2019,
Accepted 5th June 2019
DOI: 10.1039/c9dt02057h
rsc.li/dalton

Introduction

Multidentate ligand systems have been of interest for several decades for the generation of transition metal complexes with specific coordination numbers, nuclearities, or structures. In particular, tripodal *fac*-coordinating N-donor ligands such as tris(pyrazolyl)hydroborates (Tp^{R})¹ or tris(pyrazolyl)methanes ($\text{R}^{\text{R}}\text{Tpm}$)² provide intrinsically well-defined coordination geometries. Scorpionate ligands of this type have long been of interest in the area of bioinorganic, organometallic or coordination chemistry.^{1–3} This ligand family has been extended some years ago to the anionic tris(pyrazolyl)methanide⁴ and -silanide^{4i,5} systems,⁶ which are isoelectronic with the hydroborate ligands, but feature ambidentate characteristics and two possible binding sites, *i.e.* the anionic carbon and silicon or the pyrazolyl ring nitrogen donors. Another possibility for increasing the number of donor atoms in scorpionate-type ligand scaffolds is to attach further donor-containing entities either in the 3-position of the pyrazolyl entities⁷ or at the apical boron (Tp^{R}) or carbon ($\text{R}^{\text{R}}\text{Tpm}$) unit. Ligand systems that are functionalised at the non-coordinating back position of the scorpionate are

commonly referred to as “third generation scorpionates”. As recently stated by Reger *et al.* in a comprehensive review article,⁸ the tris(pyrazolyl)methane ligands are particularly well suited for these studies as the chemistry is based on reaction chemistry at a C–H bond. Functional groups attached to this apical position can either affect the properties of the ligand, like its solubility, or increase the number of possible coordination modes.⁹ Selected examples of backbone-functionalised Tpm hybrid ligands are compiled in Chart 1.

As part of our investigations on ambidentate ligands,¹⁰ we pursued the introduction of additional donor functions as additional binding sites in a Tpm-based ligand framework. We

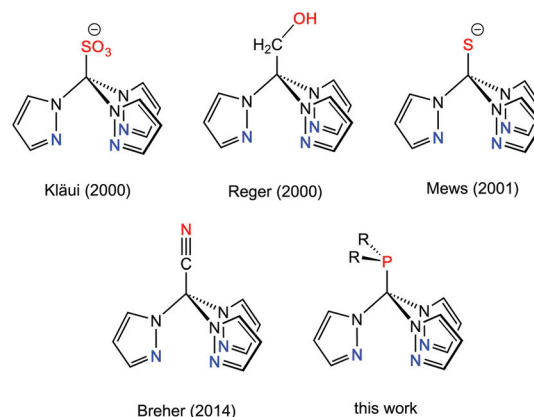


Chart 1 Backbone-functionalised Tpm ligands from the literature.

Institute of Inorganic Chemistry, Karlsruhe Institute of Technology (KIT), Engesserstr. 15, 76131 Karlsruhe, Germany. E-mail: breher@kit.edu

† Dedicated to Prof. Annie Powell on the occasion of her 60th birthday.

‡ Electronic supplementary information (ESI) available: NMR data, X-ray crystal structure data. CCDC 1916045–1916055. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c9dt02057h

report here on neutral, backbone-functionalised Tpm hybrid ligands,¹¹ namely tris(pyrazolyl)methyl phosphines. The latter were found to be versatile hybrid ligands for mono- and heterobimetallic transition metal complexes.

Results and discussion

Synthesis and characterisation of TpmPR₂ (R = Ph, *n*Bu, *i*Pr)

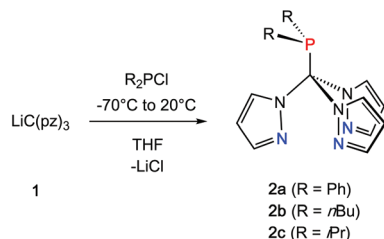
Inspired by the pioneering work of Otero¹² and Burzlaiff¹³ on heteroscorpionate ligands based on bis(pyrazolyl)methanes^{2b} or Reger^{9d} and Kläui^{9a} on tris(pyrazolyl)methanes, freshly prepared lithium tris(pyrazolyl)methanide (**1**) was reacted with R₂PCl. Moisture and air sensitive colourless crystalline pro-

ducts **2a–c** were obtained in analytically pure form in 65–71% isolated yield (Scheme 1).

The ¹H NMR spectra of the three ligands in benzene-*d*₆ show the expected equivalence of all three pyrazolyl rings within one molecule. While the resonances of the phenyl groups in **2a** appear in typical multiplets with a chemical shift of δ = 7.03 ppm (*m*-, *p*-position) and 7.98 ppm (*o*-position), **2b** and **2c** feature diastereotopic protons for the *n*Bu and *i*Pr substituents, as exemplified for **2c** in Fig. 1 (see also ESI†). Due to ¹J_{P,C} couplings, the carbon atoms in apical position of the tris-(pyrazolyl)methyl entities in **2a–c** appear as doublets with chemical shifts between δ = 95.2 and 96.0 ppm and coupling constants of 44.1–55.8 Hz.

As expected for alkyl and aryl phosphines, the ³¹P{¹H} NMR appear at δ = 37.0 ppm for **2b**, δ = 47.2 ppm for **2a** and δ = 59.9 ppm for **2c**, respectively.¹⁴

In order to unambiguously confirm the formation of the products, X-ray structure analyses were conducted on single crystals of **2a–c** (Fig. 2). The compounds crystallise in the triclinic, monoclinic and orthorhombic space groups *P* $\bar{1}$ (**2b**), *P*₂₁/*c* (**2c**) and *Pna*2₁ (**2a**), respectively (ESI†). Relevant structural data are listed in the caption of Fig. 2. As already anticipated from the NMR data, the apical carbon atom of the C(pz)₃ entities are in bound to the PR₂ donor sets. In each case, the three pyrazolyl substituents adopt a “propeller”-like arrangement with the additional nitrogen donor atoms (N_{pz}) pointing in direction of the PR₂ group. Similar effects have been observed before for backbone-functionalised tris(pyrazolyl)silanides such as R₃Si–Si(3,5-Me₂pz)₃ or R₃Sn–Si(3,5-Me₂pz)₃.^{5a} Although these structural aspects presumably can be attributed to steric repulsion between the individual pyrazolyl heterocycles, we cannot exclude additional P...N_{pz} interactions in these compounds. In all three cases, however, The P1–C1 bond lengths of 194.6 pm (**2a**), 192.2 pm (**2b**) and 192.7 pm (**2c**) are slightly longer than regular C–P single bonds (*ca.* 184 pm). It is worth mentioning that these



Scheme 1 Synthesis of **2a–c** (R = Ph, *n*Bu, *i*Pr).

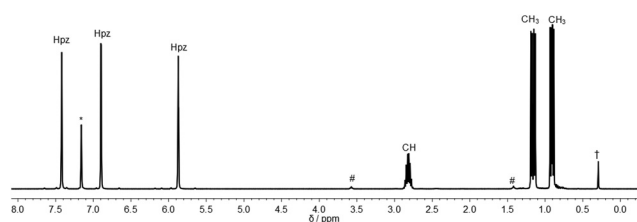


Fig. 1 ¹H NMR spectrum of **2c** (C₆D₆ (*), 300 MHz, # = THF, † = grease).

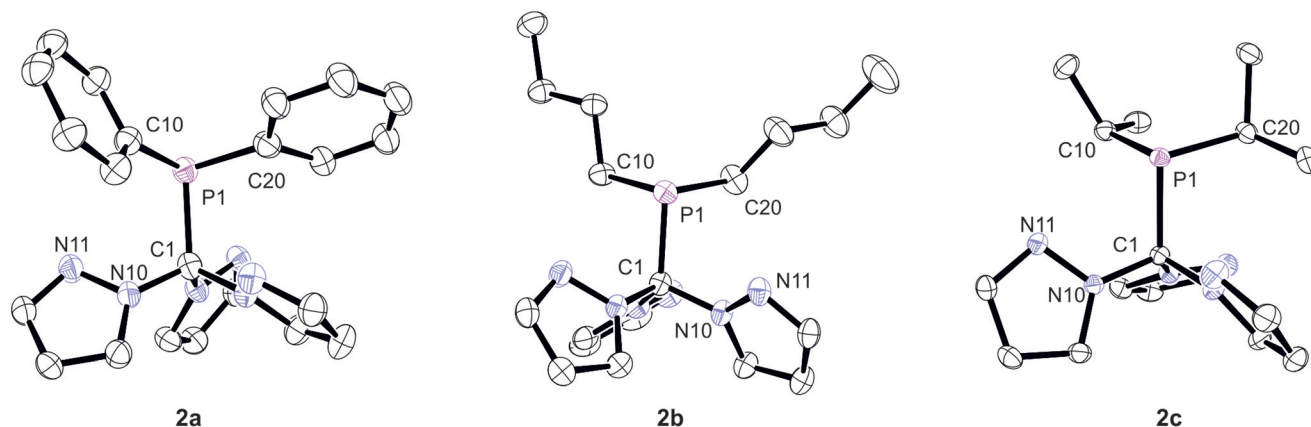


Fig. 2 Molecular structures of the ligands **2a–c**. The thermal ellipsoids are at 30% probability level. Hydrogen atoms have been omitted for clarity. Selected bond lengths [pm] and angles [°]: (**2a**) C1–P1 194.6(5), C1–N10 145.4(6), C1–N20 146.4(6), C1–N30 144.2(6) P1–C10 182.8(5), P1–C20 183.1(6), P1–C1–N10 115.0(3), P1–C1–N20 107.0(3), P1–C1–N30 111.3(3). (**2b**) C1–P1 192.2(13), C1–N10 145.3(16), C1–N20 146.1(14), C1–N30 145.1(16) P1–C10 184.8(16), P1–C20 184.2(16), P1–C1–N20 108.5(8), P1–C1–N10 108.5(8), P1–C1–N30 111.3(9), P1–C1–N20 114.6(9), P1–C1–N30 111.3(9). (**2c**) C1–P1 192.7(14), C1–N10 145.5(18), C1–N20 145.7(17), C1–N30 146.5(16) P1–C10 186.4(15), P1–C20 187.2(16), P1–C1–N10 109.4(9), P1–C1–N20 109.1(9), P1–C1–N30 115.2(9).



Dalton Trans., 2019, 48, 15397–15407 | 15399

interested to studying the coordination flexibility of our new hybrid ligands **2a–c** and, thus, investigated the coordination chemistry of **2a** or **2c** as representatives for the new ligand family in more detail.

Since $[\text{CpRu}(\text{acn})_3][\text{PF}_6]$ (acn = acetonitrile) is a valuable precursor for transferring a $\{\text{RuCp}\}$ entity to tripodal ligands such as Tpm or Tpm d ,^{4n,17} we inspected its reactivity towards **2c** (Scheme 3). Orange crystals of **4** were isolated in 91% yield. ^1H NMR spectroscopic investigations revealed that the hybrid ligand is coordinated in a $\kappa^3\text{N,N',P}$ fashion. The detection of one uncoordinated pyrazolyl ring indicated no interchange of the pz entities in solution in this case. The down field shift of the phosphorous signal in the $^{31}\text{P}\{^1\text{H}\}$ NMR spectrum of **4** is even stronger than for **3a–c** ($\delta = 170.4$ ppm). The molecular structure was established confirming the $\kappa^3\text{N,N',P}$ -hetero-

scorpionate structure (Fig. 4, space group $Pbcn$, ESI†). The coordination sphere of the ruthenium atom is complemented by one $\eta^5\text{-C}_5\text{H}_5$ ligand. The average Ru1–N bond length in **4** was found to amount to 209.1 pm, which is only slightly shorter than in comparable $\kappa^3\text{N-Tpm}$ (211.1 pm)¹⁷ and $\kappa^3\text{N-Tpm}\text{d}$ (210.6 pm)⁴ⁿ complexes published by our group. The Ru1–P1 bond length of 229.3 pm is in good agreement with those observed for other complexes from the literature.¹⁸

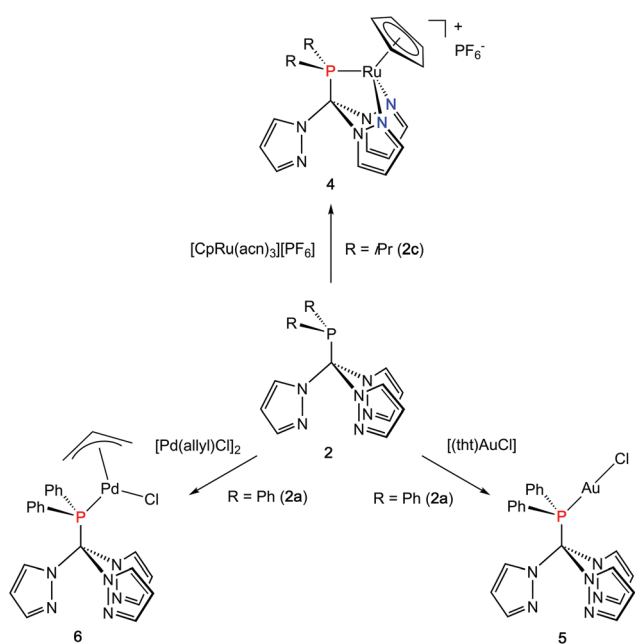
Another metal precursor which was reacted with **2a** was $[\text{Au}(\text{tht})\text{Cl}]$ (tht = tetrahydrothiophene) and complex **5** was obtained in 82% isolated yield as colourless crystalline solid. Considering that Au(I) is a rather soft metal ion and prefers linear coordination, we expected the P donor in **2a** to solely act as 2-electron donor. Indeed, the NMR spectroscopic investigations reveal just one set of signals concerning the pyrazolyl protons. The coordination of the gold atom to the P donor causes a low field shift of the phosphorous atom in the $^{31}\text{P}\{^1\text{H}\}$ NMR spectra from $\delta = 47.2$ ppm in **2a** to $\delta = 78.7$ ppm in **5**. The molecular structure depicted in Fig. 4 (space group $P2_1/n$) illustrates the almost linear coordination of the gold atom ($\text{P1–Au1–Cl1} = 175.5^\circ$). Other structural parameters remain almost unchanged compared to the uncoordinated ligand.

As a representative of group 10 metals, **2a** was reacted with $[\text{Pd}(\text{allyl})\text{Cl}]_2$ to provide a hardly soluble, pale yellow complex **6**, which was isolated in rather low yields of 48% in an analytical pure form. Like in **5**, the phosphorous atom is acting as donor ligand whereas the pyrazolyl rings stay uncoordinated. The downfield shift of the $^{31}\text{P}\{^1\text{H}\}$ chemical shift of $\delta = 89.4$ ppm is even more distinct than for **5**.

Crystals suitable for X-ray crystallography could be obtained from a THF/hexane mixture (Fig. 4, space group $P2_12_12_1$, ESI†). The Pd1–P1 bond length is 232.3 pm and in accordance to similar Pd–P bonds known from literature.¹⁸

Heterobimetallic complex of **2a**

In an effort to prepare heterobimetallic complexes of the new hybrid ligands, various attempts have been made by treating **4–6** with various other metal sources, all of which were



Scheme 3 Synthesis of **4–6**.

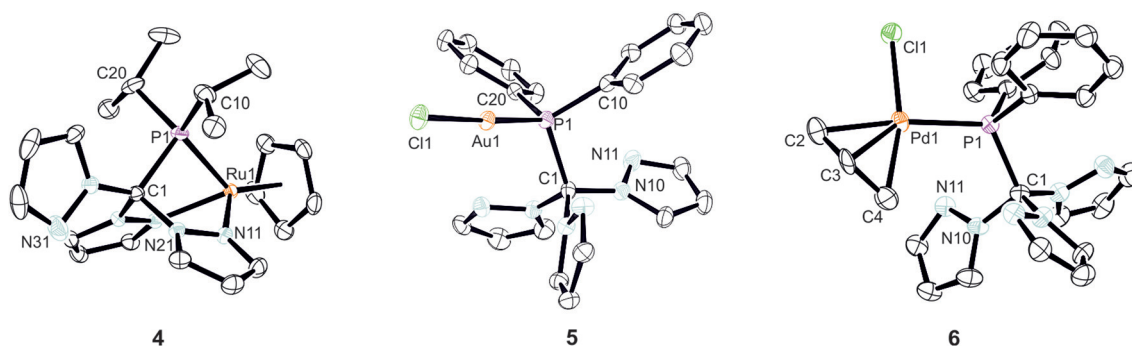
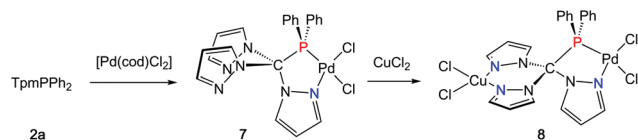


Fig. 4 Molecular structures of **4–6**. The thermal ellipsoids are at 30% probability level. Hydrogen atoms and the anion (**4**) have been omitted for clarity. Selected bond lengths [pm] and angles $^\circ$: **4**: Ru1–P1 229.27(6), Ru1–N11 209.26(18), Ru1–N21 209.00(17), Ru1–Cp $_{\text{cent}}$ 181.9, P1–C1 193.9(2), P1–Ru1–Cp $_{\text{cent}}$ 135.27, N11–Ru1–Cp $_{\text{cent}}$ 145.60, N21–Ru1–Cp $_{\text{cent}}$ 130.00, N11–Ru1–P1 78.93(5), N21–Ru1–P1 82.74(7); **5**: Au1–P1 = 223.38(15), Au1–Cl1 = 227.66(15), C1–P1 = 194.2(6), P1–Au1–Cl1 176.46(5), C1–P1–Au1 = 111.54(18), Au1–P1–C200 = 114.93(19); **6**: Pd1–P1 232.0(2), Pd1–Cl1 235.4(2), P1–C1 196.0(9), P1–Pd1–Cl1 93.97(8), Pd1–P1–C1 116.5(3). (Cp $_{\text{cent}}$ = centroid of the Cp ligand).





Scheme 4 Synthesis of 7 and 8.

not successful. However, sequential treatment of 2a with $[\text{Pd}(\text{cod})\text{Cl}_2]$ and CuCl_2 provided the suitable entry point to the first heterobimetallic complex 8 (Scheme 4).

The initial product (7) of this reaction is a yellow solid and was obtained in very good yields (87%). As evidenced by ^1H and ^{31}P NMR ($\delta = 90.7$ ppm) spectroscopic investigations, the palladium atom is coordinated in a $\kappa^2\text{P}, \text{N}$ -coordination mode leaving two pyrazolyl moieties uncoordinated. The palladium atom in 7 is coordinated in a square planar coordination mode (Fig. 5, space group $P2_1/c$, $\text{ESI}^\dagger$). Due to the different *trans*-influence of phosphorous and nitrogen donors, the

Pd1-Cl1 bond length is slightly longer (237.0 pm) than Pd1-Cl2 (228.4 pm). The Pd1-P1 bond (218.3 pm) is remarkably shorter than the Pd1-P1 bond in complex 6 (232.3 pm).

By reacting 7 with CuCl_2 , the first heterobimetallic complex 8 was successfully isolated and characterised (Fig. 5, space group $P2_1/c$, $\text{ESI}^\dagger$). The coordination sphere of the palladium atom is not affected by the coordination of $\text{Cu}(\text{II})$ since all bond lengths and angles stay almost the same. The average Cu-N and Cu-Cl bond lengths were found to amount to 223.6 pm and 202.1 pm, respectively. The square planar coordination mode at the copper atom is rather distorted as judged from the N21-Cu1-Cl4 angle of only $161.5(2)$ degrees. This distortion is also manifested in the boat-shaped arrangement of the six-membered chelate ring, which, however, has already been described for related copper complexes in the literature.¹⁹

Conclusions

In summary, we have introduced a novel class of phosphine-functionalised Tpm ligands (2a–c), which proved to be a very flexible class of hybrid ligands due to many different coordination modes observed for several transition metal complexes (3–7). Furthermore, we were able to synthesise the first heterobimetallic complex of a tris(pyrazolyl)methyl phosphine containing the metals palladium and copper. Further investigations regarding the cooperativity along the two metals within this complex and the synthesis of even more multimetallic complexes are part of ongoing research in our group.

Experimental section

General methods and materials

All manipulations were carried out with standard Schlenk line and dry-box techniques in a dry argon atmosphere. Methylene chloride and acetonitrile were freshly distilled in an argon atmosphere from calcium hydride. Toluene, diethyl ether and tetrahydrofuran were dried using sodium/benzophenone ketyl. CD_2Cl_2 was vacuum transferred from calcium hydride while C_6D_6 and *thf-d8* were vacuum transferred from sodium/benzophenone ketyl into thoroughly dried glassware equipped with Young Teflon valves.

Tpm^{9b} and $n\text{Bu}_2\text{PCl}^{20}$ were synthesized in accordance to the literature. All other reagents were purchased from Aldrich or Roth and used without further purification.

^1H , ^{13}C , spectra were recorded with Bruker AV 300 and 400 spectrometers in dried deuterated solvents. The chemical shifts are expressed in parts per millions and ^1H and ^{13}C signals are given relative to TMS. Coupling constants J are given in hertz as positive values regardless of their real individual signs. The multiplicity of the signals is indicated as s, d, q, sept or m for singlets, doublets, quartets, septets or multiplets, respectively. The assignments were confirmed as necessary with the use of 2D NMR correlation experiments. Mass spec-

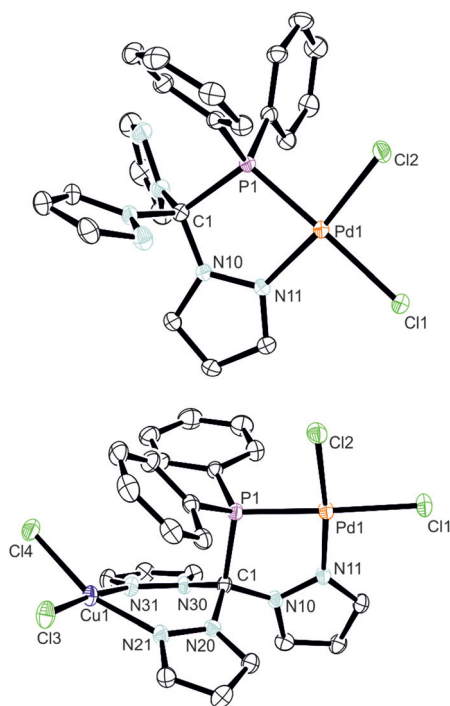


Fig. 5 Molecular structures of 7 and 8. The thermal ellipsoids are at 30% probability level. Hydrogen atoms and solvent molecules (8: acetonitrile) have been omitted for clarity. Selected bond lengths [pm] and angles [°]: 7: $\text{Pd1-P1} = 218.3(6)$, $\text{Pd1-Cl1} = 237.0(6)$, $\text{Pd1-Cl2} = 228.4(6)$, $\text{Pd1-N11} = 201.41(16)$, $\text{P1-C1} = 196.2(2)$, $\text{P1-Pd1-Cl1} = 175.411(19)$, $\text{P1-Pd1-Cl2} = 87.87(2)$, $\text{Cl2-Pd1-Cl1} = 94.91(2)$, $\text{N11-Pd1-P1} = 84.8$, $\text{N11-Pd1-Cl1} = 92.31(5)$, $\text{N11-Pd1-Cl2} = 172.22(5)$. 8: $\text{Pd1-P1} = 219.56(17)$, $\text{Pd1-Cl1} = 236.21(17)$, $\text{Pd1-Cl2} = 227.67(18)$, $\text{Pd1-N11} = 200.5(5)$, $\text{Cl1-Cu1} = 221.46(19)$, $\text{Cu1-N21} = 202.6(5)$, $\text{Cu1-N31} = 201.4(5)$, $\text{P1-C1} = 195.2(6)$, $\text{Pd1...Cu1} = 603.61(14)$, $\text{P1-Pd1-Cl1} = 172.30(6)$, $\text{P1-Pd1-Cl2} = 87.97(7)$, $\text{Cl2-Pd1-Cl1} = 94.47(7)$, $\text{N11-Pd1-P1} = 84.84(15)$, $\text{N11-Pd1-Cl1} = 92.67(15)$, $\text{N11-Pd1-Cl2} = 172.81(15)$, $\text{Cl3-Cu1-Cl4} = 94.63(8)$, $\text{N21-Cu1-Cl3} = 91.32(15)$, $\text{N21-Cu1-Cl4} = 161.49(15)$, $\text{N31-Cu1-N21} = 84.8(2)$, $\text{N31-Cu1-Cl4} = 90.07(15)$, $\text{N31-Cu1-Cl3} = 174.86(15)$.



Synthesis of 3b. TpmPBu₂ (100 mg; 0.279 mmol) and [Rh(CO)₂Cl]₂ (54.2 mg; 0.139 mmol) are dissolved in tetrahydrofuran (5 mL) and a yellow solution is obtained. After stirring over night at room temperature the solution turns orange. After adding 10 mL of hexane, yellow crystals are obtained. Yield: 128 mg (87%). ¹H NMR (400.13 MHz, thf-d₈): δ = 0.85 (t, ³J_{HH} = 7.2 Hz, 6H, CH₃), 1.29–1.36 (m, 4H, CH₂), 1.37–1.44 (m, 2H, CH₂), 1.46–1.58 (m, 2H, CH₂), 1.92–2.13 (m, 4H, CH₂), 6.51 (dd, ³J_{HH} = 2.8 Hz, ³J_{HH} = 1.9 Hz, 3H, 4-*H*_{Pz}), 7.52 (d, ³J_{HH} = 2.8 Hz, 3H, 3-*H*_{Pz}), 8.00 (br, 3H, 5-*H*_{Pz}) ppm. ¹³C{¹H} NMR (100.61 MHz, thf-d₈): δ = 13.96 (s, CH₃), 28.95, (s, CH₂), 28.96 (d, ²J_{CP} = 2.0 Hz, CH₂), 29.19 (d, ¹J_{CP} = 2 Hz), 93.4 (d, ¹J_{CP} = 20.93 Hz, C_{apical}), 108.91 (s, C_{Pz}), 132.97 (s, C_{Pz}), 143.16 (s, C_{Pz}), 191.62 (dd, ¹J_{CrH} = 74.23 Hz, ²J_{CP} = 15.08 Hz, CO) ppm. ³¹P{¹H} NMR (121.49 MHz, thf-d₈): δ = 114.04 (d ¹J_{PrH} = 155.4 Hz) ppm. EA [%] calc. C₁₉H₂₇ClN₆OPRh: C 43.49, H 5.19, N 16.01; found C 43.39, H 4.93, 15.93. APCI-MS (*m/z*): calc. 524.07, found 525.3 [M + H]⁺. FT-IR (solid, ATR, cm⁻¹): 3114 (vw), 3094 (vw), 2955 (vw), 2930 (vw), 2871 (vw), 2235 (vw), 2187 (vw), 2153 (vw), 2072(vw), 2048 (vw), 2019 (vw), 2008 (vw), 1981 (vs), 1517 (vw), 1457 (vw), 1430 (vw), 1419 (vw), 1377 (m), 1325 (w), 1312 (w), 1245 (w), 1226 (vw), 1204 (w), 1191 (vw), 1099 (w), 1086 (m), 1068 (w), 1041 (w), 990 (vw), 963 (vw), 935 (vw), 913(w), 889 (vw), 848 (m), 838 (w), 788 (m), 750 (s), 661 (vw), 637 (vw), 622 (w), 605 (w), 581 (m), 507 (w), 482 (w), 450 (vw), 406 (vw), 381 (vw). Mp (sealed tube under Ar): 148.1–153.9 °C.

Synthesis of 3c. TpmPiPr₂ (165.2 mg; 0.5 mmol) and [Rh(CO)₂Cl]₂ (97.2 mg; 0.25 mmol) are dissolved in tetrahydrofuran (10 mL) and a yellow solution is obtained. The solution is stirred overnight. After adding 10 mL of hexane, yellow crystals are obtained. Yield: 212 mg (85%). ¹H NMR (300 MHz, thf-d₈): δ = 1.03 (dd, ³J_{HP} = 17.2 Hz, ³J_{HH} = 7.0 Hz, 6H, CH-(CH₃)₂), 1.17 (dd, ³J_{HP} = 17.4 Hz, ³J_{HH} = 7.1 Hz, 6H, CH-(CH₃)₂), 3.17 (ds, ²J_{HP} = 10.9 Hz, ³J_{HH} = 7.1 Hz, 2H, CH-(CH₃)₂), 6.52 (dd, ³J_{HH} = 2.8 Hz, ³J_{HH} = 1.9 Hz, 3H, 4-*H*_{Pz}), 7.23 (dd, ³J_{HH} = 2.8 Hz, ³J_{HH} = 0.7 Hz, 3H, 3-*H*_{Pz}), 8.01 (d, ³J_{HH} = 1.9 Hz, 3H, *H*_{Pz}) ppm. ¹³C{¹H} NMR (100.61 MHz, thf-d₈): δ = 19.88 (d, ³J_{CP} = 3.4 Hz, CH-(CH₃)₂), 20.41 (d, ³J_{CP} = 1.9 Hz, CH-(CH₃)₂), 29.10 (dd, ¹J_{CP} = 18.3, ²J_{CrH} = 1.8 Hz, CH-(CH₃)₂), 94.24 (d, ¹J_{CP} = 11.8 Hz, *C*_{apical}), 109.08 (s, *C*_{Pz}), 132.91 (s, *C*_{Pz}), 143.07 (s, *C*_{Pz}), 192.21 (dd, ¹J_{CrH} = 73.5, ²J_{CP} = 14.5 Hz, CO) ppm. ³¹P{¹H} NMR (121.49 MHz, thf-d₈): δ = 143.58 (d ¹J_{PrH} = 156.1 Hz) ppm. EA [%] calc. C₁₇H₂₃ClN₆OPRh: C 41.15, H 4.67, N 16.90; found: C 41.13, H 4.58, N 16.90. APCI-MS (*m/z*): calc. 496.04, found 497.3 [M + H]⁺. FT-IR (solid, ATR, cm⁻¹): 3158 (vw), 3122 (vw), 2958 (vw), 2937 (vw), 2230 (vw), 2174 (vw), 2125 (vw), 2020 (vw), 1978 (s), 1933 (vw), 1517 (vw), 1467 (vw), 1453 (vw), 1434 (vw), 1420 (vw), 1411 (vw), 1395 (w), 1378 (m), 1333 (w), 1319 (m), 1303 (w), 1234 (m), 1201 (w), 1153 (vw), 1102 (m), 1082 (w), 1068 (m), 1043 (m), 1030 (w), 990 (vw), 969 (w), 913 (w), 882 (w), 866 (m), 839 (s), 769 (vs), 745 (vs), 668 (w), 659 (m), 648 (w), 631 (s), 605 (s), 582 (m), 520 (w), 494 (s), 462 (m), 437 (vw), 405 (w), 384 (w). Mp (sealed tube under Ar): 216 °C.

Synthesis of 4. TpmiPr₂ (37.7 mg; 0.115 mmol) and [CpRu(acyn)₃][PF₆] (50 mg; 0.115 mmol) are dissolved in dichloromethane (3 mL) and an orange solution is obtained After

adding 10 mL of hexane as antisolvent orange crystals are obtained. Yield: 67 mg (91%). ^1H NMR (300 MHz, CD_2Cl_2): δ = 0.57 (dd, $^3J_{\text{HP}} = 16.6$ Hz, $^3J_{\text{HH}} = 7.1$ Hz, 6H, $\text{CH}-(\text{CH}_3)_2$), 1.24 (dd, $^3J_{\text{HP}} = 17.5$ Hz, $^3J_{\text{HH}} = 7.0$ Hz, 6H, $\text{CH}-(\text{CH}_3)_2$), 2.63 (dhept, $^2J_{\text{HP}} = 11.7$ Hz, $^3J_{\text{HH}} = 7.1$ Hz, 2H, $\text{CH}-(\text{CH}_3)_2$), 4.88 (s, 5H, H_{CP}), 6.46 (dd, $^3J_{\text{HH}} = 3.0$ Hz, $^3J_{\text{HH}} = 2.2$ Hz, 2H, H_{Pz}), 6.90 (dd, $^3J_{\text{HH}} = 2.7$ Hz, $^3J_{\text{HH}} = 1.8$ Hz, 1H, H_{Pz}), 7.57 (dd, $^3J_{\text{HH}} = 3.0$ Hz, $^3J_{\text{HH}} = 0.7$ Hz, 2H, H_{Pz}), 7.96 (dd, $^3J_{\text{HH}} = 1.9$ Hz, $^3J_{\text{HH}} = 0.5$ Hz, 1H, H_{Pz}), 8.04 (dt, $^3J_{\text{HH}} = 2.2$ Hz, $^3J_{\text{HH}} = 0.5$ Hz, 2H, H_{Pz}), 8.28 (dd, $^3J_{\text{HH}} = 2.8$ Hz, $^3J_{\text{HH}} = 0.5$ Hz, 1H, H_{Pz}). $^{13}\text{C}\{^1\text{H}\}$ NMR (75.48 MHz, CD_2Cl_2): δ = 19.0 (d, $^2J_{\text{CP}} = 4.0$ Hz, $\text{CH}-(\text{CH}_3)_2$), 20.2 (d, $^2J_{\text{CP}} = 5.7$ Hz, $\text{CH}-(\text{CH}_3)_2$), 28.3 (d, $^1J_{\text{CP}} = 12.3$ Hz, $\text{CH}-(\text{CH}_3)_2$), 74.9 (d, $^2J_{\text{CP}} = 2.4$ Hz, C_{CP}), 95.2 (d, $^1J_{\text{CP}} = 5.3$ Hz, C_{apical}), 110.3 (s, C_{Pz}), 111.8 (s, C_{Pz}), 129.6 (s, C_{Pz}), 132.5 (s, C_{Pz}), 143.4 (s, C_{Pz}), 148.6 (s, C_{Pz}) ppm. $^{31}\text{P}\{^1\text{H}\}$ NMR (121 MHz, CD_2Cl_2): δ = -144.4 (sept, $^1J_{\text{PF}} = 710.5$ Hz, PF_6^-), 170.4 (s, TpmP) ppm. ^{19}F NMR (282.38 MHz, CD_2Cl_2) δ = -73.20 (d, $^1J_{\text{FP}} = 710.4$ Hz, PF_6^-) ppm. EA [%] calc. $\text{C}_{21}\text{H}_{28}\text{F}_6\text{N}_2\text{P}_2\text{Ru}$: C 39.32, H 4.40, N 13.10; found: C 39.39, H 4.36, N 12.94. ESI-MS (m/z): calc. 497.12 [$\text{M} - \text{PF}_6^-$] found: 497.2. FT-IR (solid, ATR, cm^{-1}): ν = 388(vw), 411 (vw), 440 (vw), 483(m), 555(vs), 590 (vw), 604 (vw), 655 (w), 668 (w), 738 (w), 763(w), 832 (vs), 875 (vw), 923 (vw), 955 (vw), 985 (vw), 1029 (vw), 1052 (w), 1091 (w), 1108 (vw), 1158 (vw), 1194 (vw), 1211 (vw), 1220 (vw), 1256 (vw), 1320 (w), 1339 (w), 1383 (w), 1406 (vw), 1428 (vw), 1465 (vw), 1529 (vw), 1959 (vw), 1982 (vw), 2022 (vw), 2168 (vw), 2875 (vw), 2968 (vw), 3149 (vw). Mp (sealed tube under Ar): 235 °C (dec.).

Synthesis of 5. TpmPPH₂ (300 mg, 0.753 mmol) and [(tht)AuCl] (241.4 mg, 0.753 mmol) are dissolved in THF (10 mL) and stirred overnight. The solvent is reduced to about 3 mL THF and hexane is added. The product is obtained as colourless crystals. Yield: 389 mg (82%). ¹H NMR (300.13 MHz, THF-d₈): δ = 6.29 (t, ³J_{HH} = 2.2 Hz, 3H, 4-*H*_{Pz}), 6.80 (d, ³J_{HH} = 2.7 Hz, 3H, 5-*H*_{Pz}), 7.27 (m, 4H, *m-H*_{Ar}), 7.37 (m, 2H, *p-H*_{Ar}), 7.64 (d, ³J_{HH} = 1.8 Hz, 3H, 3-*H*_{Pz}), 8.04 (m, 4H, *o-H*_{Ar}) ppm. ¹³C{¹H} NMR (75.48 MHz, THF-d₈): δ = 93.92 (d, ¹J_{CP} = 50.7 Hz, 1C, C_{apical}) 108.17 (s, 3C, 4-*C*_{Pz}), 121.15 (s, 3C, 5-*C*_{Pz}), 128.83 (d, ³J_{CP} = 11.02 Hz, 4C, *m-C*_{Ar}), 129.55 (d, ¹J_{CP} = 54.3 Hz 1C, C_{ipso}), 132.08 (s, 3C, 3-*C*_{Pz}), 132.25 (d, ⁴J_{CP} = 2.5 Hz, 2C, *p-C*_{Ph}), 136.97 (d, ²J_{CP} = 15.3 Hz, 4C, *o-C*_{Ph}), 143.07 (s, 3C, 3-*C*_{Pz}) ppm. ³¹P{¹H} NMR (121.49 MHz, THF-d₈): δ = 78.60 (s) ppm. MS (EI) *m/z* (%): 398.1 (0.74) [**2a**⁺], 219.9 (24.43) [Tpm⁺]. FT-IR (solid, ATR, cm⁻¹): ν = 379 (vw), 385 (w), 406 (vs), 415 (vs), 437 (s), 446 (vs), 463 (s), 481 (s), 506 (vs), 525 (vs), 534 (m), 547 (w), 562 (w), 572 (m), 584 (m), 600 (m), 622 (w), 653 (w), 686 (w), 709 (w), 744 (m), 761 (w), 853 (w), 880 (w), 977 (vw), 1047 (vw), 1083 (vw), 1096 (vw), 1185 (vw), 1245 (vw), 1322 (vw), 1384 (vw), 1435 (vw), 1479 (vw), 1515 (vw), 1933 (vw), 1965 (vw), 2001 (vw), 2028 (vw), 2071 (vw), 2087 (vw), 2123 (vw), 2177 (vw). EA [%] calc. C₂₂H₁₉AuClN₆P: C 41.89, H 3.04, N 13.32, found: C 41.60, H 2.99, N 13.04. Mp (sealed tube under Ar): 145 °C (decomposition).

Synthesis of 6. TpmPPh₂ (300 mg, 0.753 mmol) and [Pd(allyl)Cl]₂ (137.8 mg, 0.3765 mmol) are dissolved in THF (10 mL) and stirred overnight. The solvent is removed and the residue is washed with 5 ml of Et₂O. The yellow powder is dis-

Crystal data 4. $C_{21}H_{28}F_6N_6P_2Ru$, $M_r = 641.50$, orthorhombic, $Pbcn$ (no. 60), $a = 17.629(4)$ Å, $b = 18.438(4)$ Å, $c = 15.309(3)$ Å, $V = 4976.1(17)$ Å³, $T = 200$ K, $Z = 8$, $\mu(MoK_{\alpha}) = 0.827$, 91 924 reflections measured, 6718 unique ($R_{int} = 0.0415$) which were used in all calculations. The final wR_2 was 0.0821 (all data) and R_1 was 0.0302 ($I > 2(I)$).

Crystal data 5. $C_{22}H_{19}AuClN_6P$, $M_r = 630.82$, monoclinic, $P2_1/n$ (no. 14), $a = 8.7640(18)$ Å, $b = 15.604(3)$ Å, $c = 16.091(3)$ Å, $\beta = 90.76(3)^\circ$, $V = 2200.3(8)$ Å³, $T = 200$ K, $Z = 4$, $\mu(MoK_{\alpha}) = 1.904$, 12 511 reflections measured, 4760 unique ($R_{int} = 0.0548$) which were used in all calculations. The final wR_2 was 0.1086 (all data) and R_1 was 0.0327 ($I > 2(I)$).

Crystal data 6. $C_{25}H_{24}ClN_6PPd$, $M_r = 581.32$, orthorhombic, $P2_12_12_1$ (no. 19), $a = 9.0270(18)$ Å, $b = 15.942(3)$ Å, $c = 17.225(3)$ Å, $V = 2478.8(9)$ Å³, $T = 200$ K, $Z = 4$, $\mu(MoK_{\alpha}) = 1.558$, 9518 reflections measured, 5395 unique ($R_{int} = 0.0552$) which were used in all calculations. The final wR_2 was 0.1489 (all data) and R_1 was 0.0503 ($I > 2(I)$).

Crystal data 7. $C_{22}H_{19}Cl_2N_6PPd$, $M_r = 575.70$, monoclinic, $P2_1/c$ (no. 14), $a = 14.818(3)$ Å, $b = 9.2710(19)$ Å, $c = 16.998(3)$ Å, $\beta = 100.95(3)^\circ$, $V = 2292.7(8)$ Å³, $T = 200$ K, $Z = 4$, $\mu(MoK_{\alpha}) = 1.136$, 15 955 reflections measured, 6154 unique ($R_{int} = 0.0289$) which were used in all calculations. The final wR_2 was 0.0817 (all data) and R_1 was 0.0296 ($I > 2(I)$).

Crystal data 8. $C_{22}H_{19}Cl_4CuN_6PPd \cdot 2C_2H_5N$, $M_r = 792.25$, monoclinic, $P2_1/n$ (no. 14), $a = 11.268(2)$ Å, $b = 14.819(3)$ Å, $c = 18.659(4)$ Å, $\beta = 98.36(3)^\circ$, $V = 3082.6(11)$ Å³, $T = 200$ K, $Z = 4$, $\mu(MoK_{\alpha}) = 1.704$, 18 940 reflections measured, 6055 unique ($R_{int} = 0.0960$) which were used in all calculations. The final wR_2 was 0.1639 (all data) and R_1 was 0.0626 ($I > 2(I)$).

Crystallographic data (excluding structure factors) for the structures reported in this paper have been deposited with the Cambridge Crystallographic Data Centre as a supplementary publication no. CCDC 1916045–1916055.†

Conflicts of interest

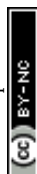
There are no conflicts to declare.

Acknowledgements

We acknowledge financial support from the DFG-funded trans-regional collaborative research centre SFB/TRR 88 “Cooperative effects in homo- and heterometallic complexes (3MET)” (project B4). Support by the Federal Ministry of Education and Research (BMBF Project No. 02NUK020B) is also gratefully acknowledged.

Notes and references

- (a) S. Trofimenko, *Chem. Rev.*, 1993, **93**, 943–980; (b) S. Trofimenko, *Scorpionates: The Coordination Chemistry of Polypyrazolylborate Ligands*, Imperial College Press, London, 1999.
- (a) H. R. Bigmore, S. C. Lawrence, P. Mountford and C. S. Tredget, *Dalton Trans.*, 2005, 635–651; (b) C. Pettinari and R. Pettinari, *Coord. Chem. Rev.*, 2005, **249**, 525–543.
- (a) G. Parkin, in *Adv. Inorg. Chem.*, ed. A. G. Sykes, Academic Press, 1995, vol. 42, pp. 291–393; (b) D. L. Reger, *Coord. Chem. Rev.*, 1996, **147**, 571–595; (c) M. Etienne, *Coord. Chem. Rev.*, 1996, **156**, 201–236; (d) C. Janiak, *Coord. Chem. Rev.*, 1997, **163**, 107–216; (e) D. L. Reger, *Comments Inorg. Chem.*, 1999, **21**, 1–28; (f) N. Kitajima and W. B. Tolman, in *Progress in Inorganic Chemistry*, ed. K. D. Karlin, John Wiley & Sons, Inc., 1995, vol. 43, p. 419; (g) H. V. R. Dias and C. J. Lovely, *Chem. Rev.*, 2008, **108**, 3223–3238; (h) *Adv. Organomet. Chem.*, ed. R. West, A. F. Hill and M. J. Fink, Academic Press, 2008, vol. 56, pp. 1–347; (i) S. Trofimenko, *J. Chem. Educ.*, 2005, **82**, 1715; (j) A. Munetaka and H. Shiro, *Bull. Chem. Soc. Jpn.*, 2002, **75**, 1657–1679; (k) G. Parkin, *Chem. Rev.*, 2004, **104**, 699–768.
- (a) P. K. Byers, N. Carr and F. G. A. Stone, *J. Chem. Soc., Dalton Trans.*, 1990, 3701–3708; (b) P. K. Byers and F. G. A. Stone, *J. Chem. Soc., Dalton Trans.*, 1991, 93–99; (c) S. C. Lawrence, M. E. G. Skinner, J. C. Green and P. Mountford, *Chem. Commun.*, 2001, 705–706; (d) F. Breher, J. Grunenberg, S. C. Lawrence, P. Mountford and H. Rüegger, *Angew. Chem., Int. Ed.*, 2004, **43**, 2521–2524; (e) H. R. Bigmore, S. R. Dubberley, M. Kranenburg, S. C. Lawrence, A. J. Sealey, J. D. Selby, M. A. Zuideveld, A. R. Cowley and P. Mountford, *Chem. Commun.*, 2006, 436–438; (f) I. Krummenacher, H. Rüegger and F. Breher, *Dalton Trans.*, 2006, 1073–1081; (g) H. R. Bigmore, J. Meyer, I. Krummenacher, H. Rüegger, E. Clot, P. Mountford and F. Breher, *Chem. – Eur. J.*, 2008, **14**, 5918–5934; (h) I. Kuzu, I. Krummenacher, I. J. Hewitt, Y. Lan, V. Mereacre, A. K. Powell, P. Höfer, J. Harmer and F. Breher, *Chem. – Eur. J.*, 2009, **15**, 4350–4365; (i) I. Fernández, P. Oña-Burgos, F. Armbruster, I. Krummenacher and F. Breher, *Chem. Commun.*, 2009, 2586–2588; (j) M. G. Cushion, J. Meyer, A. Heath, A. D. Schwarz, I. Fernández, F. Breher and P. Mountford, *Organometallics*, 2010, **29**, 1174–1190; (k) D. Kratzert, D. Leusser, D. Stern, J. Meyer, F. Breher and D. Stalke, *Chem. Commun.*, 2011, **47**, 2931–2933; (l) J. Meyer, I. Kuzu, S. González-Gallardo and F. Breher, *Z. Anorg. Allg. Chem.*, 2013, **639**, 301–307; (m) C. Müller, H. Görls, S. Kriek and M. Westerhausen, *Eur. J. Inorg. Chem.*, 2013, 5679–5682; (n) S. González-Gallardo, I. Kuzu, P. Oña-Burgos, T. Wolfer, C. Wang, K. W. Klinkhammer, W. Kloppe, S. Bräse and F. Breher, *Organometallics*, 2014, **33**, 941–951; (o) C. Müller, A. Koch, H. Görls, S. Kriek and M. Westerhausen, *Inorg. Chem.*, 2015, **54**, 635–645; (p) C. Müller, S. Kriek, H. Görls and M. Westerhausen, *Inorg. Chem.*, 2015, **54**, 2100–2102; (q) R. Lalrempuia, A. Stasch and C. Jones, *Chem. – Asian J.*, 2015, **10**, 447–454; (r) S. Kriek, A. Koch, K. Hinze, C. Müller, J. Lange, H. Görls and M. Westerhausen, *Eur. J. Inorg. Chem.*, 2016, 2332–2348; (s) I. Kuzu, A. Baldes, F. Weigend and F. Breher, *Polyhedron*, 2017, **125**, 74–79.



- 5 (a) F. Armbruster, I. Fernandez and F. Breher, *Dalton Trans.*, 2009, 5612–5626; (b) F. Armbruster, T. Augenstein, P. Oña-Burgos and F. Breher, *Chem. – Eur. J.*, 2013, **19**, 17899–17906; (c) S. Styra, S. González-Gallardo, F. Armbruster, P. Oña-Burgos, E. Moos, M. Vonderach, P. Weis, O. Hampe, A. Grün, Y. Schmitt, M. Gerhards, F. Menges, M. Gaffga, G. Niedner-Schatteburg and F. Breher, *Chem. – Eur. J.*, 2013, **19**, 8436–8446; (d) M. Simon and F. Breher, *Dalton Trans.*, 2017, **46**, 7976–7997; (e) C. Krempner, M. H. Chisholm and J. Gallucci, *Angew. Chem., Int. Ed.*, 2008, **47**, 410–413; (f) H. Li, L. J. Hope-Weeks and C. Krempner, *Chem. Commun.*, 2011, **47**, 4117–4119; (g) B. Carlson, A. J. A. Aquino, L. J. Hope-Weeks, B. Whittlesey, B. Mc Nerney, W. L. Hase and C. Krempner, *Chem. Commun.*, 2011, **47**, 11089–11091; (h) B. Mc Nerney, B. Whittlesey and C. Krempner, *Eur. J. Inorg. Chem.*, 2011, 1699–1702; (i) H. Li, F. Hung-Low and C. Krempner, *Organometallics*, 2012, **31**, 7117–7124; (j) H. Li, A. J. A. Aquino, D. B. Cordes, W. L. Hase and C. Krempner, *Chem. Sci.*, 2017, **8**, 1316–1328.
- 6 (a) A. Steiner and D. Stalke, *J. Chem. Soc., Chem. Commun.*, 1993, 1702–1704; (b) A. Steiner and D. Stalke, *Inorg. Chem.*, 1995, **34**, 4846–4853; (c) I. Krummenacher, C. Oswald, H. Rüegger and F. Breher, *Z. Anorg. Allg. Chem.*, 2007, **633**, 2354–2361; (d) I. Krummenacher, I. Fernández, H. Rüegger, F. Weigend and F. Breher, *Dalton Trans.*, 2009, 5335–5347; (e) M.-Á. Muñoz-Hernández and V. Montiel-Palma, *Inorg. Chim. Acta*, 2009, **362**, 4328–4339; (f) H. R. Simmonds and D. S. Wright, *Chem. Commun.*, 2012, **48**, 8617–8624; (g) K. Zeckert, *Organometallics*, 2013, **32**, 1387–1393; (h) K. Zeckert, S. Zahn and B. Kirchner, *Chem. Commun.*, 2010, **46**, 2638–2640; (i) M. A. Halcrow, *Dalton Trans.*, 2009, 2059–2073; (j) J. Klingele, S. Dechert and F. Meyer, *Coord. Chem. Rev.*, 2009, **253**, 2698–2741.
- 7 M. D. Ward, J. A. McCleverty and J. C. Jeffery, *Coord. Chem. Rev.*, 2001, **222**, 251–272.
- 8 R. F. Semeniuc and D. L. Reger, *Eur. J. Inorg. Chem.*, 2016, 2253–2271.
- 9 (a) W. Kläui, M. Berghahn, G. Rheinwald and H. Lang, *Angew. Chem., Int. Ed.*, 2000, **39**, 2464–2466; (b) D. L. Reger, T. C. Grattan, K. J. Brown, C. A. Little, J. J. S. Lamba, A. L. Rheingold and R. D. Sommer, *J. Organomet. Chem.*, 2000, **607**, 120–128; (c) R. Wanke, P. Smoleński, M. F. C. Guedes da Silva, L. M. D. R. S. Martins and A. J. L. Pombeiro, *Inorg. Chem.*, 2008, **47**, 10158–10168; (d) D. L. Reger and T. C. Grattan, *Synthesis*, 2003, 0350–0356; (e) M. Müller, E. Lork and R. Mews, *Angew. Chem., Int. Ed.*, 2001, **40**, 1247–1249; (f) I. Trapp, S. González-Gallardo, S. Hohnstein, D. Garnier, P. Oña-Burgos and F. Breher, *Dalton Trans.*, 2014, **43**, 4313–4319; (g) T. Augenstein, F. Dorner, K. Reiter, H. E. Wagner, D. Garnier, W. Kloppe and F. Breher, *Chem. – Eur. J.*, 2016, **22**, 7935–7943.
- 10 (a) I. Kuzu, I. Krummenacher, J. Meyer, F. Armbruster and F. Breher, *Dalton Trans.*, 2008, 5836–5865; (b) F.-G. Fontaine, J. Boudreau and M.-H. Thibault, *Eur. J. Inorg. Chem.*, 2008, 5439–5454.
- 11 (a) S. Maggini, *Coord. Chem. Rev.*, 2009, **253**, 1793–1832; (b) W.-H. Zhang, S. W. Chien and T. S. A. Hor, *Coord. Chem. Rev.*, 2011, **255**, 1991–2024.
- 12 (a) A. Otero, J. Fernández-Baeza, J. Tejeda, A. Antiñolo, F. Carrillo-Hermosilla, E. Díez-Barra, A. Lara-Sánchez, M. Fernández-López, M. Lanfranchi and M. A. Pellinghelli, *J. Chem. Soc., Dalton Trans.*, 1999, 3537–3539; (b) A. Otero, J. Fernández-Baeza, A. Antiñolo, F. Carrillo-Hermosilla, J. Tejeda, A. Lara-Sánchez, L. Sánchez-Barba, M. Fernández-López, A. M. Rodríguez and I. López-Solera, *Inorg. Chem.*, 2002, **41**, 5193–5202; (c) A. Otero, J. Fernández-Baeza, A. Antiñolo, J. Tejeda, A. Lara-Sánchez, L. Sánchez-Barba, E. Martínez-Caballero, A. M. Rodríguez and I. López-Solera, *Inorg. Chem.*, 2005, **44**, 5336–5344; (d) A. Otero, J. Fernández-Baeza, A. Antiñolo, J. Tejeda, A. Lara-Sánchez, L. Sánchez-Barba, M. Fernández-López and I. López-Solera, *Inorg. Chem.*, 2004, **43**, 1350–1358; (e) A. Otero, J. Fernández-Baeza, A. Antiñolo, J. Tejeda, A. Lara-Sánchez, L. Sánchez-Barba and A. M. Rodríguez, *Eur. J. Inorg. Chem.*, 2004, 260–266; (f) A. Otero, J. Fernández-Baeza, A. Antiñolo, J. Tejeda and A. Lara-Sánchez, *Dalton Trans.*, 2004, 1499–1510; (g) A. Otero, J. Fernández-Baeza, A. Lara-Sánchez, J. Tejeda and L. F. Sánchez-Barba, *Eur. J. Inorg. Chem.*, 2008, 5309–5326.
- 13 (a) A. Beck, B. Weibert and N. Burzlaff, *Eur. J. Inorg. Chem.*, 2001, 521–527; (b) S. Tampier, R. Müller, A. Thorn, E. Hübner and N. Burzlaff, *Inorg. Chem.*, 2008, **47**, 9624–9641; (c) H. Kopf, B. Holzberger, C. Pietraszuk, E. Hübner and N. Burzlaff, *Organometallics*, 2008, **27**, 5894–5905; (d) H. Kopf, C. Pietraszuk, E. Hübner and N. Burzlaff, *Organometallics*, 2006, **25**, 2533–2546; (e) R. Müller, E. Hübner and N. Burzlaff, *Eur. J. Inorg. Chem.*, 2004, 2151–2159; (f) N. V. Fischer, F. W. Heinemann and N. Burzlaff, *Eur. J. Inorg. Chem.*, 2009, 3960–3965; (g) B. S. Hammes, M. T. Kieber-Emmons, J. A. Letizia, Z. Shirin, C. J. Carrano, L. N. Zakharov and A. L. Rheingold, *Inorg. Chim. Acta*, 2003, **346**, 227–238.
- 14 S. Hohnstein, *Vergleichende Studien zum Koordinationsverhalten tripodaler Liganden gegenüber f-Elementen sowie Synthese und Reaktivitätsstudien von neuartigen P,N-Hybridliganden*, Dissertation, Logos Verlag Berlin GmbH, Berlin, 2016.
- 15 (a) W. Kläui, M. Berghahn, W. Frank, G. J. Reiß, T. Schönherr, G. Rheinwald and H. Lang, *Eur. J. Inorg. Chem.*, 2003, **2003**, 2059–2070; (b) C. Pettinari, F. Marchetti, G. Lupidi, L. Quassinti, M. Bramucci, D. Petrelli, L. A. Vitali, M. F. C. Guedes da Silva, L. M. D. R. S. Martins, P. Smoleński and A. J. L. Pombeiro, *Inorg. Chem.*, 2011, **50**, 11173–11183; (c) F. Marchetti, C. Pettinari, R. Pettinari, A. Cerquetella, L. M. D. R. S. Martins, M. F. C. Guedes da Silva, T. F. S. Silva and A. J. L. Pombeiro, *Organometallics*, 2011, **30**, 6180–6188.
- 16 C. A. Tolman, *Chem. Rev.*, 1977, **77**, 313–348.



- 17 I. Kuzu, D. Nied and F. Breher, *Eur. J. Inorg. Chem.*, 2009, 872–879.
- 18 D. A. Fletcher, R. F. McMeeking and D. Parkin, *J. Chem. Inf. Comput. Sci.*, 1996, **36**, 746–749.
- 19 A. S. Potapov and A. I. Khlebnikov, *Polyhedron*, 2006, **25**, 2683–2690.
- 20 P. Kübler and J. Sundermeyer, *Dalton Trans.*, 2014, **43**, 3750–3766.
- 21 G. M. Sheldrick, *Acta Cryst.*, 2015, **A71**, 3.
- 22 O. V. Dolomanov, L. J. Bourhis, R. J. Gildea, J. A. K. Howard and H. Puschmann, *J. Appl. Crystallogr.*, 2009, **42**, 339–341.

