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# SO<sub>2</sub> capture and detection with carbon microfibers (CMFs) synthesised from polyacrylonitrile†

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**SO<sub>2</sub> emissions not only affect local air quality but can also contribute to other environmental issues. Developing low-cost and robust adsorbents with high uptake and selectivity is needed to reduce SO<sub>2</sub> emissions. Here, we show the SO<sub>2</sub> adsorption–desorption capacity of carbon microfibers (CMFs) at 298 K. CMFs showed a reversible SO<sub>2</sub> uptake capacity (5 mmol g<sup>−1</sup>), cyclability over ten adsorption cycles with fast kinetics and good selectivity towards SO<sub>2</sub>/CO<sub>2</sub> at low-pressure values. Additionally, CMFs' photoluminescence response to SO<sub>2</sub> and CO<sub>2</sub> was evaluated.**

Sulphur dioxide (SO<sub>2</sub>) is a highly toxic gas that is accountable for severe respiratory illnesses, even at very low concentrations. For example, exposure to small amounts of SO<sub>2</sub> (as low as 1.5 ppm) for only a few minutes can cause momentary incapacity to breathe, and at higher concentrations (above 100 ppm) can cause death.<sup>1</sup>

Different strategies to remove SO<sub>2</sub> (flue gas desulphurisation processes FGD) have been typically used with acceptable results. These include limestone scrubbers (producing calcium sulphite)<sup>2</sup> and even SO<sub>2</sub> fixation (disulfitomercurate).<sup>3</sup> However, these procedures exhibited drawbacks associated with large amounts of wastewater, high toxicity, corrosion of pipelines, and high recuperation fees. Other SO<sub>2</sub> capture alternatives, such as silicas, zeolites, metal oxides, and activated

carbons, have exhibited low SO<sub>2</sub> efficiency.<sup>4,5</sup> Although metal-organic frameworks (MOFs) have demonstrated promising SO<sub>2</sub> capture results, for example, MOF-177 and MIL-101(Cr) showing high SO<sub>2</sub> capture values, the crystal structure of these materials collapsed after being in contact with SO<sub>2</sub>.<sup>6</sup>

Most of the current research on SO<sub>2</sub> has been narrowly focused on capturing this corrosive gas. However, the SO<sub>2</sub> capture is not the only relevant; SO<sub>2</sub> detection is as suitable as the capture and conversion of SO<sub>2</sub>.<sup>7</sup> Efficient materials for SO<sub>2</sub> detection are required to comply with the following characteristics: (i) high chemical stability towards SO<sub>2</sub> under more realistic conditions (60% of relative humidity), (ii) non-dependency on relatively high surface areas, and (iii) high processability.<sup>6</sup> In addition to remarkable chemical and structural stability, such “detector materials” are characterised for showing high SO<sub>2</sub> uptake at low pressure, providing feasible applicability in SO<sub>2</sub> detection devices.<sup>8</sup> Cooper *et al.* demonstrated outstanding SO<sub>2</sub> capture in porous organic cages (POCs) at low pressure.<sup>9,10</sup> Therefore, new porous platforms have appeared as exciting alternatives to capture and detect corrosive and explosive gases. For example, Hiraoka and co-workers reported a functionalised organic nanotube with optimal selective fluorescence properties to detect liquefied petroleum gas.<sup>11</sup>

Carbon materials have been explored for SO<sub>2</sub> capture. Yi *et al.*, tested coconut shell-based activated carbon (SAC) and coal-based activated carbon (CAC), where SAC was the best adsorbent for SO<sub>2</sub>.<sup>12</sup> Muñiz *et al.* performed thermal and chemical treatments to enhance the SO<sub>2</sub> uptake on activated carbon fibres, and they concluded that the superficial functionalities with a basic character seem to be the most important characteristic concerning SO<sub>2</sub> capture.<sup>13</sup> Wang *et al.*, developed a series of N-doped coal-based porous carbons (NCPs) by calcining a mixture of anthracite, MgO, KOH and carbamide at 1073 K; their results showed that the balance between nitrogen doping content and specific surface area (microporosity) improved the number of active adsorption sites of SO<sub>2</sub>.<sup>14</sup> In this context, the carbon microfibers (CMFs) obtained

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by calcination of polyacrylonitrile microfibers (PANMFs)<sup>15</sup> present an opportunity for SO<sub>2</sub> detection due to the following aspects: (i) chemical composition based on nitrogen and oxygen functional groups resulting from the polymer precursor (PAN), (ii) high microporosity controllable depending on calcination temperature, (iii) good thermal stability, and (iv) reversible gas adsorption (e.g., CO<sub>2</sub> or CH<sub>4</sub>). Concerning nitrogen functional groups, there have been identified four groups in the CMFs: N-6 (pyridine-N), N-5 (pyrrolic-N), N-X (pyridine-N-oxide) and N-Q (quaternary-N or graphitic-N).<sup>15</sup> Some of these groups have improved the performance of CMFs in oxygen reduction reactions (ORR) in fuel cells<sup>16</sup> and their gas adsorption properties (CO<sub>2</sub> and CH<sub>4</sub>).<sup>15</sup>

Textural, chemical, and structural characterisation of the CMFs have been reported previously (Fig. S1, ESI†).<sup>17</sup> CMFs were obtained by calcination of PANMFs at 1173 K; this material has a specific surface area of 731 m<sup>2</sup> g<sup>-1</sup>, a total pore volume of 0.348 cm<sup>3</sup> g<sup>-1</sup>, and a microporosity above 70%. An average pore size of 0.78 nm was calculated from the N<sub>2</sub> adsorption isotherm at 77 K, and 0.5 nm was estimated using the CO<sub>2</sub> adsorption isotherm at 273 K.<sup>17</sup> The CMFs average chemical composition is C: 89%, N: 6%, and O: 5%. It is important to mention that the fibrous structure of PANMFs is preserved after calcination with fibre diameters between 200 and 400 nm.

Since CMFs contain several nitrogen sites, which can be potential SO<sub>2</sub>-adsorption sites, we measured the SO<sub>2</sub> adsorption

at 298, 303 and 308 K (Fig. 1a). SO<sub>2</sub> isotherms showed a type-I profile based on IUPAC<sup>18</sup> with a small hysteresis. CMFs showed a maximum uptake of 5.2, 4.9 and 4.6 mmol g<sup>-1</sup> at 1 bar for 298, 303 and 308 K, respectively. This value is higher than several reported in the literature compared to other carbonaceous or inorganic materials in the function of superficial area BET (Fig. S4, ESI†). The three SO<sub>2</sub> adsorption isotherms were used to calculate the isosteric enthalpy of adsorption, obtaining values around -30 kJ mol<sup>-1</sup> (see Fig. S3, ESI†), consistent with a physisorption process and mild regeneration conditions.

Then, cyclability tests were carried out to evaluate the reusability of the material at the conditions where the highest SO<sub>2</sub> capture was obtained. Ten SO<sub>2</sub> adsorption-desorption cycles were performed at 298 K until 1 bar. The amount of SO<sub>2</sub> captured in each cycle is stable, around 5 mmol g<sup>-1</sup> (Fig. 1b). Between each cycle, a vacuum activation process was enough to desorb almost all the SO<sub>2</sub> adsorbed, leading to the slight increase in the baseline and thus, the maximum SO<sub>2</sub> uptake in each cycle.

FTIR-ATR and SEM measurements were performed to characterise the CMFs in the SO<sub>2</sub> capture process (Fig. 2). In the three different stages during the SO<sub>2</sub> uptake (before and after the first adsorption cycle and after ten desorption cycles indicated by pink, yellow and purple colours, respectively, in Fig. 1 and 2), the IR-ATR spectra showed the presence of ester groups between 2250 to 2000 cm<sup>-1</sup>, and coupling C-N stretching and N-H deformation modes of C-N-H groups (amide)

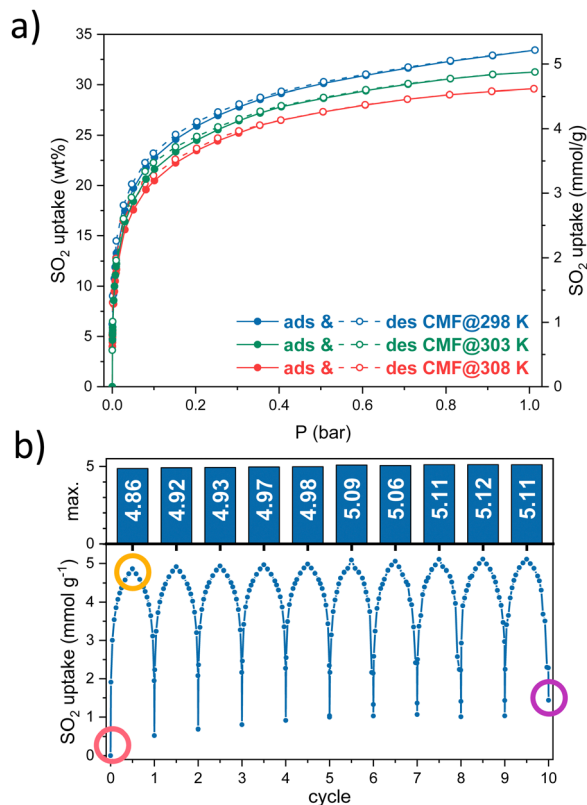


Fig. 1 (a) SO<sub>2</sub> adsorption-desorption isotherms at 298, 303, and 308 K, (b) ten SO<sub>2</sub> adsorption-desorption cycles on CMFs.

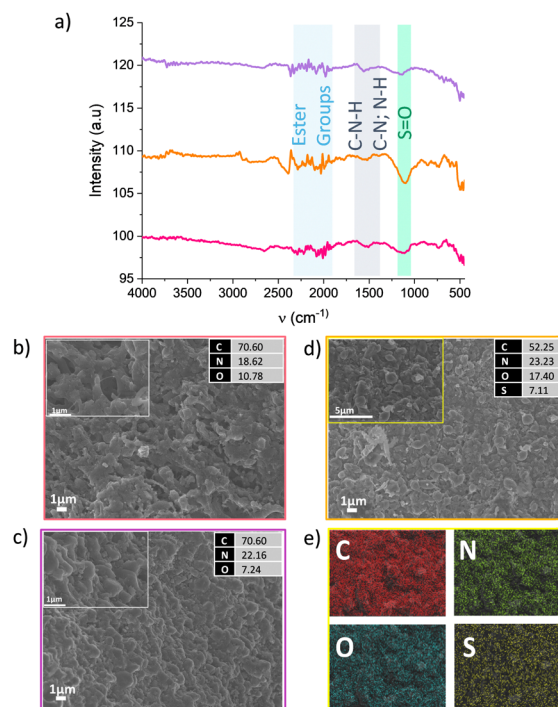


Fig. 2 (a) IR-ATR spectra at different stages according to the points marked in Fig. 1b, pristine sample (pink), after SO<sub>2</sub> uptake (yellow), and after 10 desorption cycles (violet). SEM micrographs with their EDX results on (b) pristine sample, (c) after 10 desorption cycles, (d) after SO<sub>2</sub> uptake, and (e) the EDX mapping results of the rectangular area inside panel (d).

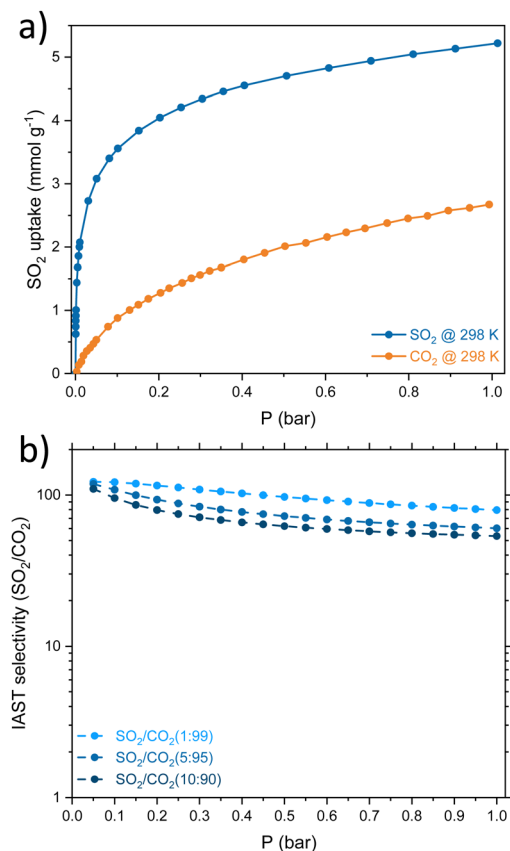


around  $1522\text{ cm}^{-1}$ . For the  $\text{SO}_2$ -saturated CMFs sample, a characteristic band in  $1050\text{ cm}^{-1}$  was identified, indicating the S=O group was present (Fig. 2a).<sup>19,20</sup> These results are in good agreement with the SEM micrographs and EDX analyses. Carbon microfibres morphology only changes when the  $\text{SO}_2$  capture process occurs; the surface showed small globularities protruding from the fibre channels, and the EDX results showed sulphur presence of around 7 wt% and an increment of the oxygen percentage as well (Fig. 2d). The EDX mapping displayed a homogeneous distribution of C, N, O and S on the surface in the area shown inside the  $5\text{ }\mu\text{m}$  scale micrograph in Fig. 2d. The sample showed the same morphology and composition before starting the cycles (Fig. 2b), with the sample pristine and after the last desorption cycle (Fig. 2c). These results indicate that the CMFs are stable against  $\text{SO}_2$  for ten adsorption-desorption cycles.

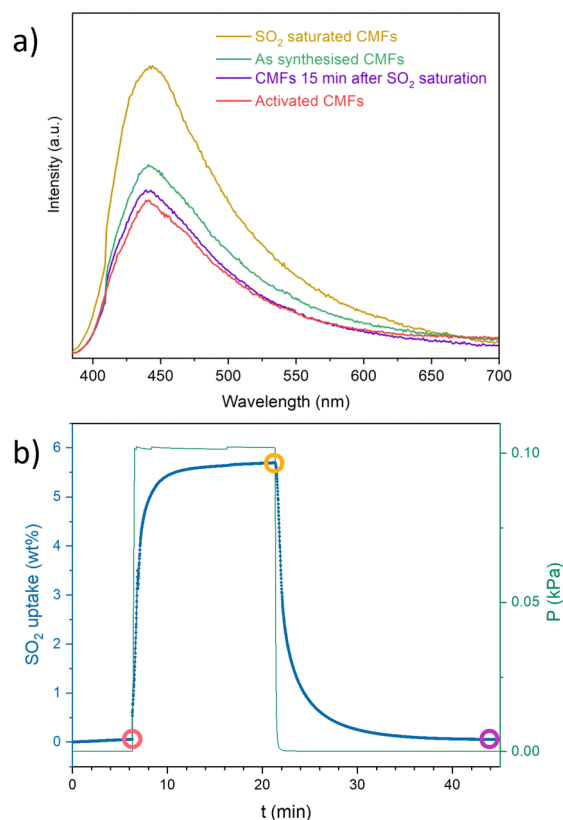
Considering that  $\text{SO}_2$  is often found as a minor component of flue gas mixtures (diluted in  $\text{CO}_2$  and/or  $\text{N}_2$ ), the separation selectivity is a crucial factor to consider. The  $\text{SO}_2/\text{CO}_2$  separation selectivity was determined employing the ideal adsorbed solution theory (IAST) using two monocomponent isotherms of  $\text{SO}_2$  and  $\text{CO}_2$  at 298 K (Fig. 3a, details on ESI†). The CMFs exhibited good selectivity values for the binary mixtures  $\text{SO}_2/\text{CO}_2$  in the low-pressure domain, 122, 118 and 110 for

1%, 5% and 10% of  $\text{SO}_2$  at 0.05 bar, respectively (Fig. 3b). The IAST selectivity result is comparable with similar superficial area BET adsorbents such as zeolite Y ( $180,930\text{ m}^2\text{ g}^{-1}$ ),<sup>21</sup> Mg-gallate ( $321,576\text{ m}^2\text{ g}^{-1}$ ),<sup>22</sup> Co-gallate ( $143,494\text{ m}^2\text{ g}^{-1}$ ),<sup>22</sup> DMOF-TM ( $169,900\text{ m}^2\text{ g}^{-1}$ ),<sup>23</sup> MIL-160 ( $128,1170\text{ m}^2\text{ g}^{-1}$ ),<sup>24</sup> Cu-ATC ( $114,600\text{ m}^2\text{ g}^{-1}$ ),<sup>25</sup> NbOFFIVE-Cu-TPA ( $78,1179\text{ m}^2\text{ g}^{-1}$ ).<sup>26</sup> Granted, the  $\text{SO}_2$  uptake of CMFs falls short in front of benchmark materials. However, these results invite us to explore another application of the CMFs,  $\text{SO}_2$  detection, where the reversible adsorption and selectivity are relevant.

Photoluminescence experiments were carried out on the CMFs using a  $\lambda_{\text{ex}} = 370\text{ nm}$  after exposure to an  $\text{SO}_2$ -saturated atmosphere (details on ESI†). The PL intensity increased by about 50% after the  $\text{SO}_2$  exposure, compared to the activated sample (Fig. 4a). This switch-on emission decreased over time: after 15 min of exposure, the emission returned to the value of the reference sample. However, when the sample is not activated and has been left in contact with the environment, the signal increases, indicating that it detects other molecules, such as  $\text{H}_2\text{O}$  or  $\text{CO}_2$ . To evaluate this hypothesis, PL measurements were performed by saturating the CMFs with  $\text{CO}_2$  and  $\text{H}_2\text{O}$  separately (Fig. S6, ESI†). The results showed a positive response for carbon dioxide but not for water. The presence of specific functional groups on carbon materials as the nitrogen-



**Fig. 3** (a) Comparison of the  $\text{SO}_2$  and  $\text{CO}_2$  adsorption isotherms at 298 K on CMFs. (b) IAST selectivity of  $\text{SO}_2/\text{CO}_2$  on MCFs for different concentrations of the binary mixture.



**Fig. 4** (a) Photoluminescence CMFs spectra ( $\lambda_{\text{ex}} = 370\text{ nm}$ ) of as synthesised, activated,  $\text{SO}_2$  saturated and after the exposure to  $\text{SO}_2$  samples. (b) Profile of adsorption-desorption kinetics of  $\text{SO}_2$  on CMFs (gravimetric experiment with controlled  $\text{SO}_2$  atmosphere).



bearing active sites in the CMFs (in the form of  $\text{NH}_2$ , for example, where nitrogen acts as an electron donor) may enhance the interaction with  $\text{SO}_2$  and the resulting photoluminescent response. The interaction between these gases with free electron pairs favours light absorption and subsequent emission for detection.<sup>27</sup> For the  $\text{SO}_2$  interaction, the observed reversibility agrees with the observed adsorption-desorption kinetics of  $\text{SO}_2$  obtained by gravimetric experiments (Fig. 4b). However, even though a similar PL intensity was observed for  $\text{CO}_2$  exposed sample compared to  $\text{SO}_2$ , the PL emission of the  $\text{CO}_2$  exposed sample remained after several hours, indicating a slow desorption of this gas molecule (Fig. S7, ESI†).

In summary,  $\text{SO}_2$  adsorption-desorption capacity at room temperature and 1 bar of CMFs was around  $5 \text{ g mol}^{-1}$ . It maintained good chemical and morphological stability during 10 adsorption-desorption cycles of  $\text{SO}_2$  and a good  $\text{SO}_2/\text{CO}_2$  selectivity, achieving a reasonable degree of reuse. When evaluating the photoluminescence of the material, it was determined that it can detect  $\text{SO}_2$  and  $\text{CO}_2$  but not  $\text{H}_2\text{O}$  and that  $\text{SO}_2$  desorption is faster than  $\text{CO}_2$ . CMFs may be functionalised to improve their textural properties,  $\text{SO}_2$  uptake and selectivity overall.

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## Conflicts of interest

There are no conflicts to declare.

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