



Cite this: *Chem. Sci.*, 2018, 9, 7362

All publication charges for this article have been paid for by the Royal Society of Chemistry

Received 23rd April 2018
Accepted 17th July 2018

DOI: 10.1039/c8sc01841c

rsc.li/chemical-science

Reversible coordination of N₂ and H₂ to a homoleptic *S* = 1/2 Fe(i) diphosphine complex in solution and the solid state†

Laurence R. Doyle,^a Daniel J. Scott,^a Peter J. Hill,^a Duncan A. X. Fraser,^a William K. Myers,^{ib} Andrew J. P. White,^a Jennifer C. Green^b and Andrew E. Ashley^{ib}*^a

The synthesis and characterisation of the *S* = 1/2 Fe(i) complex [Fe(depe)₂]⁺[BAR₄][−] ([1]⁺[BAR₄][−]), and the facile reversible binding of N₂ and H₂ in both solution and the solid state to form the adducts [1·N₂]⁺ and [1·H₂]⁺, are reported. Coordination of N₂ in THF is thermodynamically favourable under ambient conditions (1 atm; Δ*G*₂₉₈ = −4.9(1) kcal mol^{−1}), while heterogenous binding is more favourable for H₂ than N₂ by a factor of ~300. [1·H₂]⁺[BAR₄][−] represents a rare example of a well-defined, open-shell, non-classical dihydrogen complex, as corroborated by ESR spectroscopy. The rapid exchange between N₂ and H₂ coordination under ambient conditions is unique for a paramagnetic Fe complex.

Introduction

Recent years have seen great interest in the pursuit of well-defined transition metal (TM) complexes capable of catalysing the reduction of N₂ to NH₃ or N₂H₄.¹ One of the most compelling rationales for studying such systems is that their mechanistic details can be more readily discerned, in comparison with the complex proton-coupled electron transfer (PCET) steps operative within nitrogenase enzymes.² The utility of Fe in biological N₂ fixation,³ and the anthropogenic Haber–Bosch process,⁴ has prompted researchers to target Fe complexes as potential synthetic catalysts. These ‘artificial’ nitrogenases employ chemical H⁺ and e[−] sources to reduce N₂ through PCET pathways, although this process competes with proton reduction to H₂, which can preferentially sequester active metal sites.⁵ Facile displacement of H₂ by N₂ is thus an important aspect to maintaining a productive N₂-fixing catalytic cycle, and an understanding of the binding of these small molecules to low-valent Fe centres could lead to more selective and efficient catalysts for the production of azanes. While the coordination chemistry of N₂ and H₂ to Fe(0) and Fe(II) complexes is well-documented,⁶ analogous detailed studies containing Fe(i) are scant,⁷ despite the potential relevance of this low-valent oxidation state in Fe-based synthetic nitrogenases. We recently reported that the Fe(0) bisphosphine complex, Fe(depe)₂(N₂)

[1·N₂; depe = Et₂PCH₂CH₂PEt₂] is a highly selective catalyst for the PCET-mediated conversion of N₂ to N₂H₄ with H⁺ and e[−] equivalents.⁸ Herein we report the related low-spin Fe(i) species [Fe(depe)₂]⁺ ([1]⁺) which is shown to reversibly coordinate N₂ or H₂, with the latter being a rare case of a thoroughly-characterised paramagnetic σ-H₂ complex. Furthermore, this behavior is found to occur both in solution and in the solid-state.

Methods

The ESR spectroscopic methods used for characterizing molecular structure are based on the following spin Hamiltonian, representing the electron Zeeman, nuclear Zeeman, hyperfine and nuclear quadrupole components, respectively.

$$\mathcal{H} = \hat{S}' g \vec{B} + \sum_i \hat{I}'_i g_{N,i} \vec{B} + \sum_i \hat{S}' A_i \hat{I}_i + \sum_k \hat{I}'_k Q_k \hat{I}_k \quad (1)$$

The bold symbols are 3 × 3 tensors (or matrices) and the vector of Cartesian spin operators are defined in appropriate Hilbert space eigenbasis. The third term of the spin Hamiltonian, hyperfine or \mathcal{H}_{HF} is composed of an isotropic component, *a*_{iso}, transformed to a 3 × 3 matrix by, **1** defined as the identity matrix, and an anisotropic component, **T**.

$$\mathcal{H}_{\text{HF}} = \sum_i \hat{S}' [a_{\text{iso},i} \mathbf{1} + \mathbf{T}_i] \hat{I}_i \quad (2)$$

It is the anisotropic tensor component that returns the structural relations of the spin system, based on the summation

^aDepartment of Chemistry, Imperial College London, London SW7 2AZ, UK. E-mail: a.ashley@imperial.ac.uk

^bInorganic Chemistry Laboratory, University of Oxford, Oxford OX1 3QR, UK

† Electronic supplementary information (ESI) available: Full experimental details and characterisation data. CCDC 1451414. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c8sc01841c



of dipolar interactions between the position vectors of the central iron atom and surrounding positions, approximated as ligand nuclei position vectors.⁹

$$T_i = (\mu_0 g_e \mu_B g_N \mu_N / 4\pi\hbar) \times \sum_j \left[f_j (1 + \Delta g_j / g_e) (1 r_{ij}^2 - 3 \vec{p}_i \vec{p}_j') / r_{ij}^5 \right] \quad (3)$$

Eqn (3) includes contribution of all centres, p_j , of spin density f_j at distance r_{ij} to the hyperfine interaction matrix for a single nucleus position, p_i . The principal axis elements in the limit of an axial tensor are $[-T, -T, 2T]$ where the perpendicular component (eqn (4)) is used to estimate a distance in a single point-dipole approximation:

$$T = f_{Fe} g_{eff} \mu_B g_N \mu_N / r^3 \quad (4)$$

Analysis of the nuclear quadrupole term of the spin Hamiltonian of eqn (1) for an $I = 1$ nucleus (relevant to ^2H and ^{14}N in this work) can provide useful information on bonding and electronic structure,¹⁰ and is given by:

$$Q = K \times \text{diag}[-(1 - \eta), -(1 + \eta), 2] \quad (5)$$

where K is the axial quadrupole interaction ($= e^2 q_{zz} Q / 4h$) and η is the orthorhombic asymmetry parameter. In the limit of pure quadrupole frequencies ($\nu_1 = A/2$) the $m_s = +1/2$ manifold frequencies are $\nu_+ = K(3 + \eta)$, $\nu_- = K(3 - \eta)$, and $\nu_0 = 2K\eta$, for a positive hyperfine interaction.¹¹ In the opposite manifold, $m_s = -1/2$, there are two single quantum (sq) peaks and a double quantum (dq) transition, which corresponds to ν_+ in the $m_s = +1/2$ manifold. A formula for the double quantum frequency in the limit of small hyperfine anisotropy is:¹²

$$\nu_{dq} = 2[\nu_{ef\pm}^2 + K^2(3 + \eta^2)]^{1/2} \quad (6)$$

where ν_{ef} is the effective frequency of the Larmor and hyperfine, $\nu_{ef\pm} = \nu_1 \mp A/2$, based on the DC field and unpaired electron.

Results and discussion

Previously, we showed that $\text{N}_2\text{H}_4/\text{NH}_3$ -producing reactions of $1 \cdot \text{N}_2$ with the acid $[\text{Ph}_2\text{NH}_2]^+[\text{TfO}]^-$ ($\text{TfO} = \text{CF}_3\text{SO}_3$) formed an Fe(i) species, which was shown by X-ray crystallography to be $[\text{Fe}(\text{depe})_2(\eta^1\text{-N}_2)]^+[\text{TfO}]^-$ ($[1 \cdot \text{N}_2]^+[\text{TfO}]^-$).⁸ Variable-temperature ESR spectra of this compound were complicated, however, which we postulated may be due to competitive N_2 vs. $[\text{TfO}]^-$ coordination. In order to suppress the latter, and hence better resolve the behaviour of the $[1 \cdot \text{N}_2]^+$ fragment, we subsequently utilised the more weakly coordinating $[\text{BAR}_4^F]^-$ ($\text{BAR}_4^F = \text{B}[3,5\text{-}(\text{CF}_3)_2\text{C}_6\text{H}_3]_4$). Thus, oxidation of $1 \cdot \text{N}_2$ with 1 eq. $[\text{Cp}_2\text{Fe}]^+[\text{BAR}_4^F]^-$ ($\text{Cp} = \text{C}_5\text{H}_5$) produced a deep-blue solution, from which an intensely blue crystalline solid was obtained following work-up and recrystallisation from Et_2O /pentane under Ar.¹³ A single crystal suitable for X-ray diffraction was subsequently isolated which solved and refined as the 15 valence electron (VE), N_2 -free compound $[\text{Fe}(\text{depe})_2]^+[\text{BAR}_4^F]^-$ ($[1]^+[\text{BAR}_4^F]^-$, Fig. 1(a) and (b)). In



Fig. 1 (a) Synthesis of $[1]^+[\text{BAR}_4^F]^-$. (b) ORTEP diagram of the $[\text{Fe}(\text{depe})_2]^+$ fragment in $[1]^+[\text{BAR}_4^F]^-$; thermal ellipsoids shown at 30% probability, H atoms and $[\text{BAR}_4^F]^-$ counterion omitted for clarity (C and P occupancies are 0.5 as a consequence of a second symmetry-generated orientation of the two depe ligands; this has also been omitted for clarity). (c) CW X-band ESR spectrum of $[1]^+[\text{BAR}_4^F]^-$ (PhMe : DFB 7 : 1; Ar; 40 K); black line: experiment; grey line: simulation. * $[1 \cdot \text{N}_2]^+[\text{BAR}_4^F]^-$ impurity due to trace N_2 in the Ar atmosphere.

the structure the $[1]^+$ and $[\text{BAR}_4^F]^-$ ions are well separated, with no close C or F contacts between the anion and the Fe centre.

The $[1]^+$ cation was found to be disordered and, while the structural model is in good agreement with the X-ray crystallographic data (with the molecular connectivity and absence of an N_2 ligand being conclusive), the interatomic distances are approximate and will not be discussed in detail.¹⁴ The FeP_4 unit is pseudo square planar and exhibits a tetrahedral distortion, as seen by a dihedral angle of $12.41(11)^\circ$ between the two $\text{Fe}(\text{PP})$ [$\text{Fe}(1)\text{P}(1)\text{P}(4)$ and $\text{Fe}(1)\text{P}(1)\text{P}(4i)$] coordination planes, which is very similar to that observed in the approximately square-based pyramidal $[1 \cdot \text{N}_2]^+[\text{TfO}]^-$ ($15.39(9)^\circ$). Evidently, coordination of N_2 results in minimal reorganisation of the $[\text{Fe}(\text{depe})_2]^+$ fragment. While two CH_3 groups from ligand ethyl groups from each depe moiety are directed towards the vacant axial coordination sites of the Fe centre, large $\text{Fe} \cdots \text{H}$ separations suggest an absence of any agostic or anagostic interactions, which is supported by DFT results (*vide infra*).¹⁵

Neither solid samples nor solutions of $[1]^+[\text{BAR}_4^F]^-$ under Ar showed bands attributable to an ν_{NN} stretch in their IR or Raman spectra, further confirming the absence of N_2 in the complex. $[1]^+[\text{BAR}_4^F]^-$ is insoluble in alkanes, PhH and PhMe, yet highly soluble in THF and in the highly polar, non-coordinating 1,2-difluorobenzene (DFB). ^{31}P NMR spectra (DFB, Ar) of $[1]^+[\text{BAR}_4^F]^-$ are silent, whereas very broad paramagnetically-shifted resonances for the $[1]^+$ moiety feature in the ^1H NMR spectrum (see ESI†). The solution-phase magnetic moment (Evans NMR, DFB, 243–298 K, Ar) was found to be $1.75 \mu_B$, and the X-band ESR spectrum (PhMe/DFB glass, Ar, 40 K) revealed a rhombic signal ($g_1 = 2.483$, $g_2 = 2.234$, $g_3 = 1.985$) with an isotropic g -value (g_{iso}) of 2.23, [$g_{\text{iso}} = (g_x + g_y + g_z)/3$; Fig. 1(c)]. These data are consistent with a low-spin ($S = 1/2$, d^7) Fe(i) centre. The strong similarity between the X-band ESR spectrum



of a powdered sample of $[1]^+[\text{BAR}_4^F]^-$ (Ar, 1 bar, 40 K) and solution measurements implies that $[1]^+$ is virtually isostructural in solution and the solid state.

Dissolution of $[1]^+[\text{BAR}_4^F]^-$ in N_2 -saturated solvents (1 atm; DFB or THF) afforded forest-green solutions at room temperature, which became pale yellow upon cooling to -30°C ; either heating these solutions above room temperature, or degassing with Ar, resulted in the rapid reappearance of the characteristic blue colour of $[1]^+[\text{BAR}_4^F]^-$. An IR active ν_{NN} stretch at 2067 cm^{-1} confirmed coordination of N_2 to $[1]^+$ to form $[1\cdot\text{N}_2]^+$, which is intermediate in value between those seen for $1\cdot\text{N}_2$ and the related $\text{Fe(II)} [trans\text{-Fe(H)(N}_2\text{)(depe)}_2]^+$ (1975 cm^{-1} and 2102 cm^{-1} respectively);¹⁶ this trend may be readily accounted for by decreased $\text{Fe} \rightarrow \text{N}_2$ π -backbonding as the oxidation state increases.

The thermodynamics of N_2 coordination were obtained from variable temperature UV-vis spectroscopy by monitoring the concentrations of $[1]^+$ and $[1\cdot\text{N}_2]^+$ via their absorption features (λ_{max} (nm) = 618 and 1018, respectively; Fig. 2).¹⁷ N_2 association with $[1]^+$ is accordingly found to be exoergic ($\Delta G_{298} = -4.9(1)\text{ kcal mol}^{-1}$) with, as expected, a favourable enthalpy ($\Delta H^\circ = -13.1(1)\text{ kcal mol}^{-1}$) and an unfavourable entropy ($\Delta S^\circ = -27.6(1)\text{ cal K}^{-1}\text{ mol}^{-1}$) contribution; these values compare well with those for N_2 binding by $(\text{P}_3\text{B})\text{Co} [\text{P} = o\text{-(P}i\text{Pr)}_2\text{C}_6\text{H}_4]$; $\Delta H^\circ = -13.9(7)\text{ kcal mol}^{-1}$, and $\Delta S^\circ = -32(5)\text{ cal K}^{-1}\text{ mol}^{-1}$, which also produces an $S = 1/2$ dinitrogen complex.^{18,19}

The X-band ESR spectra of rapidly freeze-quenched N_2 -saturated DFB or THF solutions (40 K) of $[1]^+[\text{BAR}_4^F]^-$ revealed a pseudo-axial signal which differs markedly from the rhombic signal characteristic of the N_2 -free complex ($g_{\text{iso}} = 2.07$; simulated g tensors = 2.0014, 2.0922, 2.125). Additionally, hyperfine coupling to the four ^{31}P nuclei was resolved for the THF glass giving $A(^{31}\text{P}) = 62.7(1)\text{ MHz}$ (Fig. S5†), which is very similar to $A_{\text{iso}}(^{31}\text{P})$ previously obtained for $[1\cdot\text{N}_2]^+[\text{TfO}]^-$ [$66.2(2)\text{ MHz}$].⁸

Similar spectroscopic observations have been described for the Fe(I) complex $[\text{Fe}(\text{DMeOPrPE})_2(\text{N}_2)]^+$ ($\text{DMeOPrPE} = \text{R}_2\text{-PCH}_2\text{CH}_2\text{PR}_2$; $\text{R} = \text{CH}_2\text{CH}_2\text{CH}_2\text{OMe}$), which were attributed to

an equilibrium between yellow $[\text{Fe}(\text{DMeOPrPE})_2(\text{N}_2)]^+$ and a purple N_2 -bridged bimetallic $S = 1$ species $\{[\text{Fe}(\text{DMeOPrPE})_2]_2(\mu\text{-N}_2)\}^{2+}$; the latter is favoured at higher temperatures and/or low $p(\text{N}_2)$.²⁰ We noted that comparable optical absorptions were observed for $[1]^+$ and $[1\cdot\text{N}_2]^+$ (see Table S1, ESI†), and accordingly 2D-ESR (HYSCORE; HYperfine Sub-level CORrelation) experiments were performed on equivalent $^{14}\text{N}_2$ - and $^{15}\text{N}_2$ -saturated PhMe/DFB (7 : 1) solutions of $[1]^+[\text{BAR}_4^F]^-$, to fully ascertain the solution-phase coordination mode of N_2 to $[1]^+$. As shown in Fig. 3, two ^{14}N signals were clearly detected at the perpendicular field position, in panels a, b, e, f.

The proximal nitrogen, $^{14}\text{N}_\text{p}$, gives intense double quantum dq, dq correlation peaks in the $(-, +)$ quadrant corresponding to the ^{14}N directly bound to Fe(I) centre, leading to peaks at $(\pm 16.2, \pm 12.2)\text{ MHz}$ and, from eqn (6), $A_{\text{iso}}(^{14}\text{N}_\text{p})$ may correspondingly be calculated as $14.3 (\pm 0.1)\text{ MHz}$.²² For the distal nitrogen, $^{14}\text{N}_\text{d}$, field-dependent $^{15}\text{N}_2$ simulations of 4-pulse HYSCORE provide $A_{\text{iso}}(^{14}\text{N}_\text{d}) = 4.4\text{ MHz}$ (seen in ESI, Fig. S7†). A clearly-resolved quadrupole interaction of $^{14}\text{N}_\text{p}$ reveals a very small asymmetry $\eta \sim 0$, which compares well with sp-hybridized ^{14}N found in $\text{N}\equiv\text{N}$ and $[\text{C}\equiv\text{N}]^-$.^{10a} For the $^{14}\text{N}_\text{d}$ (red in Fig. 3), it was found that $K = 0.9\text{ MHz}$. In the case of the $^{14}\text{N}_\text{p}$, the data was insufficient for a more precise determination, and the quadrupole values of $^{14}\text{N}_\text{d}$ were used as an approximation. For ligand atoms directly bonded with covalent character, a point-



Fig. 2 Variable-temperature UV-vis spectra of $[1]^+[\text{BAR}_4^F]^-$ in THF under 1 atm. N_2 ($[\text{BAR}_4^F]^-$ counterions omitted). Thermodynamics of N_2 binding: $\Delta H^\circ = -13.1(1)\text{ kcal mol}^{-1}$, $\Delta S^\circ = -27.6(1)\text{ cal K}^{-1}\text{ mol}^{-1}$; $K_{298} = 4.1(3) \times 10^{-3}\text{ M}^{-1}$.

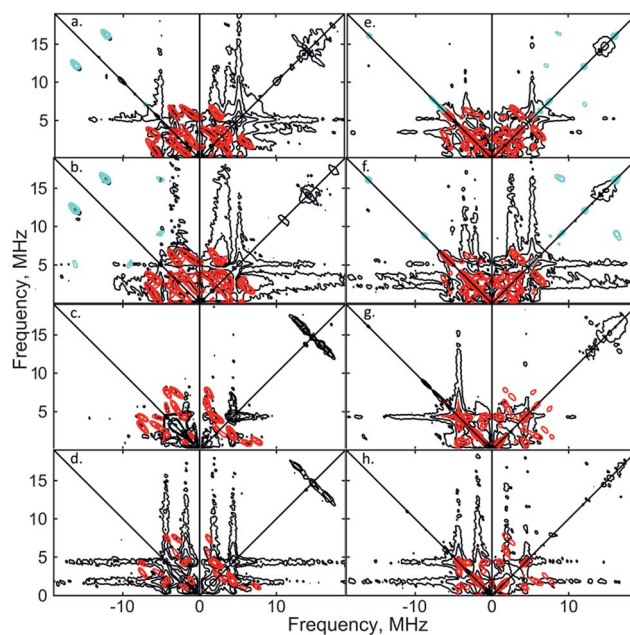


Fig. 3 X-band 2D ESEEM spectra of $2.5\text{ mM } [1\text{-}^{14}\text{N}_2]^+[\text{BAR}_4^F]^-$ in THF, with standard 4-pulse HYSCORE (panels (a–d)), and DONUT HYS-CORE²¹ (panels (e–h)), at two fields (3471 G: (a, b, e and f); 3800 G: (c, d, g and h)) and two τ values (200 ns: (a, c, e and g); 132 ns: (b, d, f and h)) while $\tau_2 = 800\text{ ns}$ for DONUT. The samples were frozen glasses at 20 K and measured with pulse lengths of $\pi/2 = 8\text{ ns}$ and $\pi = 12\text{ ns}$, initial variable delays of 100 ns, and a time step of 20 ns over 200 points in both axes. The microwave frequency was 9.7449 GHz. Experimental data is in black, while separate simulations of two ^{14}N are overlaid in red and cyan.

dipole model can be inaccurate; for $^{14}\text{N}_d$ $T = 1.2$ MHz, and for $^{14}\text{N}_p$ the simulation value was $T = 0.6$ MHz. To confirm that both ^{14}N signals derive from the same molecule (considering that no large ^{15}N hyperfine coupling was observed) close examination of the dq, dq peaks at 3471 G reveals weak multi-nuclear combination frequency peaks such as $\{\text{N}_p^\alpha(\text{dq}), \text{N}_p^\beta(\text{dq}), \text{N}_d^\beta(\text{dq})\}$ (seen in Fig. S7†), a signal class previously reported by Stich *et al.* for $\text{Mn}_2(\text{III/IV})$ Catalase.²³ It is notable that no ‘half-field’ resonance ($g \approx 4$) was observed in the CW-ESR experiments, which would be expected for the hypothetical triplet ($S = 1$) $\{[\text{Fe}(\text{depe})_2(\mu\text{-N}_2)]^{2+}$ due to a formally forbidden $\Delta M_s = 2$ transition, as has been documented for other transition metal diradicals.²⁴ Collectively alongside other spectroscopic data, these observations strongly support the solution-phase assignment of $[\text{1}\cdot\text{N}_2]^+[\text{BAR}_4^F]^-$ as $[\text{Fe}(\text{depe})_2(\eta^1\text{-N}\equiv\text{N})]^+[\text{BAR}_4^F]^-$.

Given that $\{\text{TM}(\sigma\text{-N}_2)\}$ and $\{\text{TM}(\sigma\text{-H}_2)\}$ fragments are related by the same VE count, numerous diamagnetic metal-ligand platforms have been shown to interconvert these species under N_2 and H_2 mixtures;²⁵ however, analogous open-shell examples are extremely rare.²⁶ Furthermore, only two thoroughly characterised paramagnetic dihydrogen ligand complexes have been reported to date: $(\text{P}_3\text{B})\text{Co}(\text{H}_2)^{18}$ and $(\text{P}_3\text{Si})\text{Fe}(\text{H}_2)^7$; both of these are idealised C_3 symmetric, trigonal bipyramidal $S = 1/2$ complexes. In spite of the different coordination geometry and cationic charge, admission of H_2 (1 atm.) to DFB or THF solutions of $[\text{1}]^+[\text{BAR}_4^F]^-$ demonstrated clear reaction (Fig. 4(a)), as evidenced by an immediate colour change to pale green ($\lambda_{\text{max}} = 850$ nm, $\epsilon_{\text{max}} \approx 19$ m² mol⁻¹) and the appearance of a new near-axial ESR signal (X-band, PhMe/DFB, 40 K; $g = [2.000, 2.085, 2.160]$, $g_{\text{iso}} = 2.08$, Fig. S6†), which displays more pronounced hyperfine splitting for g_2 and g_3 than seen for $[\text{1}\cdot\text{N}_2]^+[\text{BAR}_4^F]^-$.²⁷ The most plausible identities of this species are the Fe(I) adduct $[\text{Fe}(\text{depe})_2(\sigma\text{-H}_2)]^+ ([\text{1}\cdot\text{H}_2]^+)$ or the Fe(III) oxidative addition product $[\text{Fe}(\text{depe})_2(\text{H}_2)_2]^+ ([\text{1}(\text{H}_2)_2]^+)$.²⁸ Using the DFT-optimised structures of $[\text{1}\cdot\text{H}_2]^+$ and $[\text{1}(\text{H}_2)_2]^+$ (*vide infra*), the angle (β) between the g_1 principal axis and the Fe–H vectors was calculated to be 16° and 31° , respectively.

Orientation-selective ENDOR has previously been used to differentiate between $\text{Fe}(\text{H}_2)$ and $\text{Fe}(\text{H})_2$ formulations using electron- ^1H dipolar hyperfine coupling interactions, where only



Fig. 4 (a) H_2 coordination by $[\text{1}]^+[\text{BAR}_4^F]^-$. (b) Schematic showing angle β between g_{-1} principal axis and Fe–H bond vectors, relevant to electron- ^1H dipolar coupling for alternative $[\text{1}\cdot(\text{H}_2)]^+$ and $[\text{1}(\text{H}_2)_2]^+$ structures, from reaction of $[\text{1}]^+$ and H_2 . P = PET_2 ; $[\text{BAR}_4^F]^-$ counteranions omitted.

small β angles ($<15^\circ$; see Fig. 4(b)) were able to reproduce the observed experimental lineshape.^{7,29} The 1 : 2 : 1 hyperfine seen at the five negative ^{31}P peaks of g_{min} in CW-ESR of Fig. S6,† was interrogated more closely with H_2/D_2 -saturated solutions of $[\text{1}]^+[\text{BAR}_4^F]^-$ using ^1H Davies ENDOR (Electron-Nuclear DOuble Resonance) and both ^1H & ^2H HYSCORE experiments (see Fig. 5). Fits to data from both techniques used two H hyperfine interaction values, for ^1H $A(^1\text{H}) = [-17.99, -19.93, 26.58]/\text{MHz}$, while the values were scaled by the nuclear g -factor ratio $g_n(^2\text{H})/g_n(^1\text{H}) = 0.1535$ for the ^2H HYSCORE simulation, revealing a dipolar component of $T = 15.2$ MHz. The Fe–H distance, r_{FeH} (Å), can thus be obtained from $r_{\text{FeH}}(\text{Å}) = \sqrt[3]{(\rho_{\text{Fe}})79.06/T}$, using a spin density of $\rho_{\text{Fe}} \sim 0.83$ as remainder of the large isotropic ^{31}P hyperfine interactions;^{30,31} this coupling is consistent with an Fe–H bond distance of 1.64 Å. Considering the spin density at Fe and the coordinating ligand atoms and DFT coordinates (*vide infra*, eqn (3)) was used to fit an angle of $2\beta = \text{H-Fe-H}$, $\beta = 5.5^\circ$, using the empirical principal dipolar values $[-14.2, -16.2, 30.4]$ MHz (written as $[-T(1 - \delta), -T(1 + \delta), 2T]$, with δ as the rhombicity parameter); this value is in excellent agreement with that found for $(\text{P}_3\text{Si})\text{Fe}(\text{H}_2)$ ($\beta = 6^\circ$),⁷ which was interpreted as indicative of partial rotational averaging of the dipolar interaction. However, the angle $\beta = 5.5^\circ$ includes spin density on opposing $^{1,2}\text{H}$, and Morris *et al.*, has shown that rotational motion can reduce H–H dipolar interactions by up to a factor of four, implying that $\beta = 5.5^\circ$ should be considered an upper limit.³²

Simulations of ENDOR data for $[\text{1}\cdot\text{H}_2]^+$ in Fig. 6 are consistent with the DFT structure (*vide infra*), in having two classes of

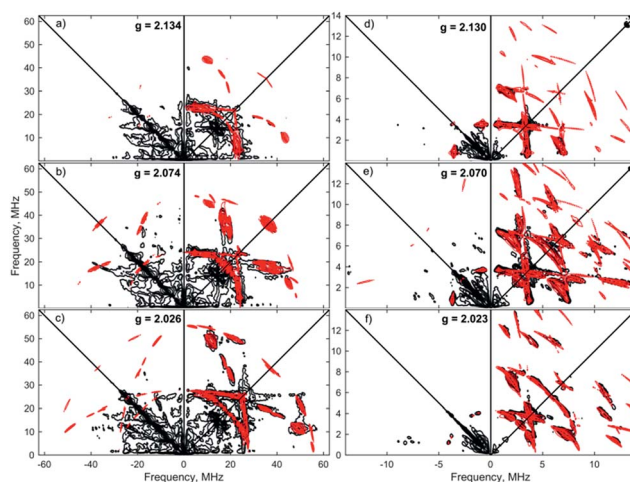


Fig. 5 X-band 4-pulse HYSCORE spectra of $[\text{1}\cdot\text{H}_2]^+[\text{BAR}_4^F]^-$ (panels (a–c)) and $[\text{1}(\text{H}_2)_2]^+[\text{BAR}_4^F]^-$ (panels (d–f)) at 2.5 mM in 7 : 1 PhMe : DFB, each collected at three field positions as indicated by g -value. The samples were frozen glasses at 20 K and measured with pulse lengths for ^2H of $\pi/2 = 12$ ns and $\pi = 8$ ns, $\tau = 204$ ns, initial variable delays of 100 ns and a time step of 20 ns over 300 points in both axes. The pulse lengths for ^1H were $\pi/2 = 8$ ns and $\pi = 12$ ns, $\tau = 144, 140, 132$ ns (top to bottom), initial variable delays of 100 ns and a time step of 8 ns, and 200 points in both axes. The microwave frequency was 9.7411 and 9.7583 GHz, respectively. Experimental data is in black, while simulations are overlaid in red.



^{31}P hyperfine interaction, $a_{\text{iso}}(^{31}\text{P}) = 69.3$ MHz of P in-plane orthogonal to the Fe(I)–H₂ bond and $a_{\text{iso}}(^{31}\text{P}) = 66.3$ MHz for the P bent out-of-plane; furthermore the P–Fe–P depe ligand angle of 85.5° suggests that the symmetry lies closer to 2-fold than 4-fold. Correspondingly, these results provide strong evidence that $[\mathbf{1}\cdot\text{H}_2]^+$ is best described as a non-classical $\sigma\text{-H}_2$ adduct.

Freshly-prepared solutions of $[\mathbf{1}\cdot\text{H}_2]^+$ are sufficiently stable for detailed *in situ* characterisation and, as with solutions under N₂, vacuum/Ar-degassing resulted in regeneration of spectroscopic signals attributed to $[\mathbf{1}]^+$, demonstrating reversible coordination of H₂. Nevertheless, attempts to obtain thermochemical information for H₂ binding were unfortunately frustrated by slow and irreversible formation of the Fe(II) trisphide *trans*-[Fe(depe)₂(H)(H₂)]⁺, $[\mathbf{2}]^+$.³³ Appreciating that ligation of N₂ or H₂ to $[\mathbf{1}]^+$ results in minimal deformation of the [Fe(depe)₂]⁺ core,³⁴ we speculated that binding of these gases might also be reversible in the solid state, as in solution. Gratifyingly, admission of either N₂ or H₂ to powdered samples of $[\mathbf{1}]^+[\text{BAR}_4^{\text{F}}]^-$ led to comparable colour changes (deep blue to yellow or green, respectively), consistent with clean conversion



Fig. 6 X-band Davies ENDOR spectra of $[\mathbf{1}\cdot\text{H}_2]^+[\text{BAR}_4^{\text{F}}]^-$ at 2.5 mM in 7 : 1 PhMe : DFB, collected at nine field positions as indicated by *g*-value along the right side. The sample was a frozen glass at 20 K and measured with pulse lengths of $\pi/2 = 40$ ns and $\pi_{\text{inv}} = \pi = 80$ ns, $\pi = 300$ ns, RF pulse = 16 μs , and 1 μs before and 2 μs after the RF pulse, with stochastic frequency stepping. The microwave frequency was 9.7621 GHz. Experimental data is in black, while simulations of ^{31}P values are red, and simulations of ^1H are in blue (as discussed in main text).

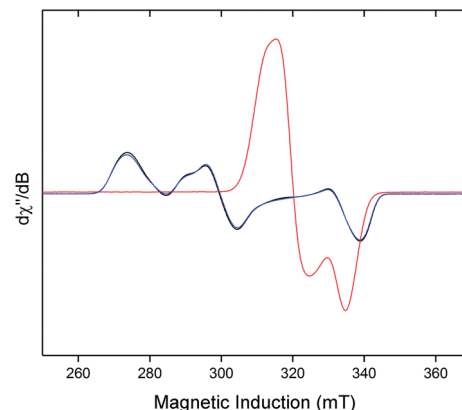


Fig. 7 Reversible binding of H₂ by $[\mathbf{1}]^+[\text{BAR}_4^{\text{F}}]^-$ in the solid-state: CW X-band ESR spectra of powdered $[\mathbf{1}]^+[\text{BAR}_4^{\text{F}}]^-$ under Ar (black), followed by evacuation and subsequent admission of H₂ (1 bar, $[\mathbf{1}\cdot\text{H}_2]^+[\text{BAR}_4^{\text{F}}]^-$, pink), and sequential evacuation to 10^{−3} mbar and backfill with Ar (blue). All spectra recorded at 40 K.

to $[\mathbf{1}\cdot(\text{N}_2/\text{H}_2)]^+[\text{BAR}_4^{\text{F}}]_{(\text{s})}^-$. ESR spectra closely matched those obtained in solution, strongly suggesting that the geometry of the [Fe(depe)₂]⁺ core is preserved in both phases upon N₂ and/or H₂ coordination. Crucially, removal of N₂/H₂ from these samples under vacuum led to complete restoration of the original ESR signal of $[\mathbf{1}]^+$ (Fig. 7 and S2†), confirming that binding is fully reversible (over multiple cycles) in the solid state; importantly, no loss in signal intensity was observed (which would be expected from formation of diamagnetic $[\mathbf{2}]^+$) under H₂. Furthermore, ready exchange of the N₂ ligand for H₂ is achieved *via* simple evacuation of N₂ and replacing with H₂, and *vice versa*.

ESR spectra of $[\mathbf{1}]^+[\text{BAR}_4^{\text{F}}]^-$ obtained in the presence of a single equivalent of N₂ or H₂ allow for a quantitative comparison of the binding affinities of the two gases; the ratio $[\mathbf{1}\cdot\text{L}]/[\mathbf{1}]^+$ (L = N₂ or H₂; determined by signal intensity at 303 K) is significantly larger for H₂ (44) than N₂ (0.15), revealing that binding of the former is almost 300 times more favourable at ambient temperature.³⁵ For comparison, (P₃B)Co(L) (L = N₂, H₂) are in rapid dissociative equilibrium in solution under similar conditions, whereas (P₃Si)Fe(L) species require several days to interconvert (proposed to proceed *via* an associative mechanism involving partial dechelation of the P₃Si ligand);^{7,18} both demonstrate a preference for H₂ binding, albeit to differing degrees ($K_{\text{H}_2}/K_{\text{N}_2} \approx 2$ and 50 for Co and Fe, respectively). This difference in exchange kinetics was postulated to result from the poorer π -backbonding capability of Co *vs.* Fe which leads to weaker M–L interactions. Hence it is plausible that the ready reversibility of N₂ and H₂ exchange for $[\mathbf{1}]^+$ relative to (P₃Si)Fe could also be due (in part) to poorer π -donation from the former, by virtue of its cationic charge, which is also manifest in the higher ν_{NN} stretch value of the former (2067 *vs.* 2003 cm^{−1}).

Computational calculations

In order to probe the structure and electronic properties of $[\mathbf{1}]^+$, $[\mathbf{1}\cdot\text{N}_2]^+$, and $[\mathbf{1}\cdot\text{H}_2]^+$, density functional calculations were



carried out using the ADF program suite version 2014.1.³⁶ The Slater-type orbital (STO) basis sets were of triple- ζ quality augmented with a one polarization function (ADF basis TZP). Core electrons were frozen (C, N 1s; Fe 2p) in our model of the electronic configuration for each atom. The local density approximation (LDA) by Vosko, Wilk and Nusair (VWN)³⁷ was used together with the exchange correlation corrections of Becke and Perdew (BP86).³⁸

Optimized geometries were ascertained as local minima *via* frequency calculations. Geometry optimisation of base-free $[1]^+$ with $S = 1/2$ resulted in a D_2 structure consistent with X-ray crystallographic data, with an angle of approximately 7° between the two iron-ligand Fe(PP) coordination planes (Fig. 8(a)); fixing the spin state to $S = 3/2$ showed the alternative high-spin structure to be some 0.77 eV (approximately 17.8 kcal mol⁻¹) higher in energy. Geometry optimisation of $[1 \cdot N_2]^+$ results in a structure of C_2 symmetry with a slightly increased angle between the two Fe(PP)₂ planes of 19° (Fig. 8(b)). The coordination geometry around Fe is square-based pyramidal with P–Fe–N_α angles of 94.8° and 95.6° . The calculated stretching frequency for the bound N₂ ligand was 2059 cm⁻¹, which is in good agreement with experiment (2067 cm⁻¹). Geometry optimisation of $[1 \cdot H_2]^+$ (Fig. 8(c)) gave a σ -complex with a H–H bond length of 0.899 Å (*cf.* free H₂: 0.74 Å)⁶ and an average Fe–H distance of 1.61 Å, which correlates well with 2D-ESR data. An alternative $[1 \cdot (H_2)_2]^+$ isomer (Fig. 8(d)), corresponding to the product of H₂ oxidative addition and containing two well-separated hydride ligands (H \cdots H = 1.57 Å and Fe–H distances of 1.51 Å), could also be located, albeit 12 kcal mol⁻¹ higher in energy than $[1 \cdot H_2]^+$ (see Table S2 in ESI for further details†).

Since all compounds had unpaired spins the DFT calculations were unrestricted, with different orbitals for α and β spins. The energies of the α spin electrons, of which there are more, tended to be lower than those of the β spin electrons in corresponding orbitals because of exchange stabilization (Fig. 9).

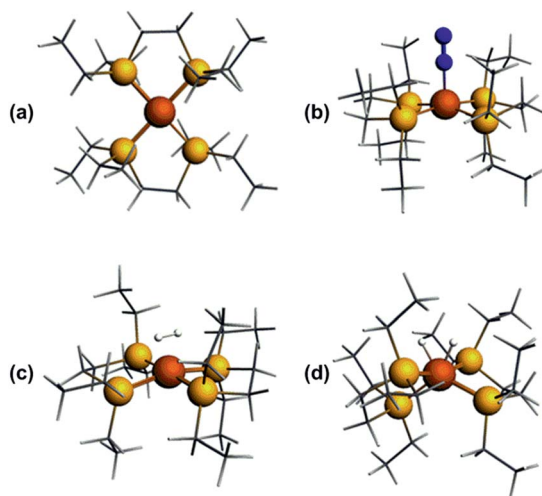


Fig. 8 DFT-optimised structures of (a) $[1]^+$, (b) $[1 \cdot N_2]^+$, (c) $[1 \cdot H_2]^+$, (d) $[1(H)_2]^+$.



Fig. 9 Electronic energy levels and their occupancy for $[1]^+$, $[1 \cdot N_2]^+$ and $[1 \cdot H_2]^+$. The principal Fe character is indicated.

The unpaired electron in the three cases, $[1]^+$, $[1 \cdot N_2]^+$, and $[1 \cdot H_2]^+$, occupies an orbital of primarily d(z^2) character; in the case of $[1 \cdot N_2]^+$ and $[1 \cdot H_2]^+$ this is also hybridized with the Fe 4p(z) orbital (Fig. 10), and is antibonding with respect to the coordinated N₂ or H₂, thus explaining the weak association of these ligands to the $[Fe(depe)_2]^+$ core. Of particular interest is the virtual orbital (z) of $[Fe(depe)_2]^+$, the isosurface of which is shown in Fig. 10; its AO composition is predominantly Fe 4p(z) and P 3p.

Time-dependent DFT (TDDFT) was used to calculate the electronic absorption spectra of $[1]^+$, $[1 \cdot N_2]^+$ and $[1 \cdot H_2]^+$ (see

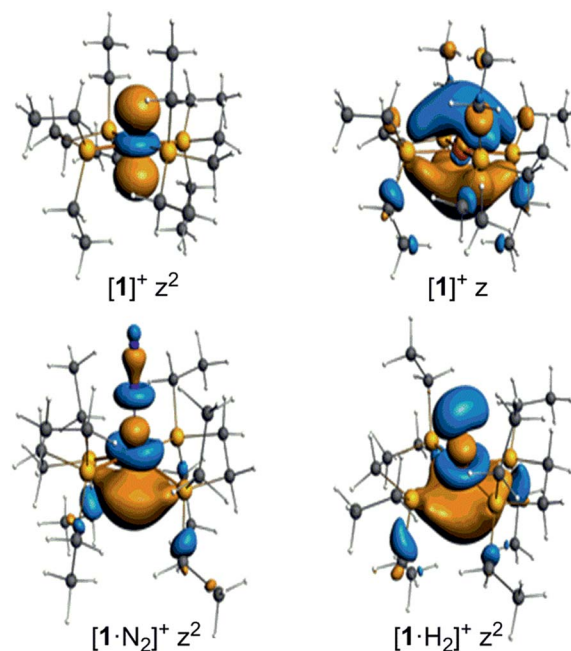


Fig. 10 Selected Kohn–Sham isosurfaces for $[1]^+$, $[1 \cdot N_2]^+$ and $[1 \cdot H_2]^+$ showing the z^2 occupancy of the unpaired electrons and the virtual orbital for $[1]^+$ of primarily $4p_z$ character.



Table S3 and Fig. S15 in the ESI†). The numerical agreement with experiment is only moderate, with the bands being calculated at higher energy than measured experimentally. All three compounds have as their lowest energy bands d–d transitions, with a β spin electron being excited into the hole in the $d(z^2)$ orbital. However those bands of $[1]^+$ have effectively zero oscillator strength on account of its high symmetry. The d–d transitions are of longer wavelength for $[1]^+$ than for $[1 \cdot N_2]^+$ and $[1 \cdot H_2]^+$, which fits with the smaller HOMO–LUMO gap in the former (see Fig. 9). The subsequent set of bands calculated at 587, 548 and 503 nm for $[1]^+$ fit well with the features found at 500 and 618 nm experimentally; they are of significantly higher oscillator strength than the d–d bands of $[1 \cdot N_2]^+$ and $[1 \cdot H_2]^+$ and correspond to excitation from the occupied 3d orbitals into the Fe 4p_z orbital. In $[1 \cdot N_2]^+$ and $[1 \cdot H_2]^+$, the corresponding band is absent as a consequence of ligand binding along the z axis, hence their respective absorption spectra have no analogous feature. Thus, in spite of the lack of numerical agreement for the d–d bands, many features of the calculated spectra give a good account of those observed.

Conclusions

In conclusion, we have synthesised and fully characterised an open-shell, cationic Fe(i) complex, $[1]^+[\text{BAR}_4^F]^-$, and demonstrated the readily reversible coordination of N₂ and H₂ to the $S = 1/2$ Fe centre. Remarkably, this facile exchange between N₂ and H₂ coordination occurs in either solution or solid states under ambient conditions, likely due to the very small structural change in transitioning between square planar $[1]^+$ and square-based pyramidal $[1 \cdot (N_2/H_2)]^+$ species. Furthermore, $[1 \cdot H_2]^+[\text{BAR}_4^F]^-$ is a rare example of a well-defined paramagnetic σ -H₂ complex, as corroborated by ESR spectroscopy and DFT calculations. These results reveal that reversible coordination of these small molecules to open-shell complexes is neither restricted to neutral compounds nor C_3 molecular symmetries, which until now, have been the only known examples. Given the importance of Fe in catalytic N₂ fixation, this work is of significant relevance to the ongoing development and investigation of well-defined transition metal catalysts for N₂ reduction, mediated by H⁺/e[−] sources or (ideally) H₂ as the terminal reductant.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

We thank the EPSRC for providing a PhD studentship (P. J. H) and postdoctoral funding (L. R. D and D. J. S), the Royal Society for a University Research Fellowship (A. E. A.), and the Centre for Advanced ESR spectroscopy (CAESR) at the University of Oxford (EP/L011972/1).

Notes and references

- For recent reviews, see: (a) Y. Tanabe and Y. Nishibayashi, *Chem. Rev.*, 2016, **16**, 1549; (b) K. C. MacLeod and P. L. Holland, *Nat. Chem.*, 2013, **5**, 559; (c) S. Hinrichsen, H. Broda, C. Gradert, L. Sönksen and F. Tuczek, *Annu. Rep. Prog. Chem., Sect. A: Inorg. Chem.*, 2012, **108**, 17; (d) C. Köthe and C. Limberg, *Z. Anorg. Allg. Chem.*, 2015, **641**, 18.
- B. M. Hoffman, D. Lukoyanov, Z.-Y. Yang, D. R. Dean and L. C. Seefeldt, *Chem. Rev.*, 2014, **114**, 4041.
- Y. L. Hu and M. W. Ribbe, *J. Biol. Inorg. Chem.*, 2015, **20**, 435.
- R. Schlögl, in *Catalytic Ammonia Synthesis Fundamentals and Practice*, ed. J. R. Jennings, Plenum Press, New York, 1st edn, 1991, ch. 2, pp. 19–108.
- The higher absolute entropy of N₂ vs. H₂ can offset favourable TM–(N₂) vs. TM–(H₂) bond enthalpies, becoming more important at higher temperatures: A. A. Gonzalez and C. D. Hoff, *Inorg. Chem.*, 1989, **28**, 4295.
- See: G. J. Kubas, *Chem. Rev.*, 2007, **107**, 4152 and references therein.
- Y. Lee, R. A. Kinney, B. M. Hoffman and J. C. Peters, *J. Am. Chem. Soc.*, 2011, **133**, 16366.
- P. J. Hill, L. R. Doyle, A. D. Crawford, W. K. Myers and A. E. Ashley, *J. Am. Chem. Soc.*, 2016, **138**, 13521.
- (a) A. Bachmeier, J. Esselborn, S. V. Hexter, T. Krämer, K. Klein, T. Happe, J. E. McGrady, W. K. Myers and F. A. Armstrong, *J. Am. Chem. Soc.*, 2015, **138**, 5381; (b) A. Schweiger and G. Jeschke, *Principles of Pulsed Electron Paramagnetic Resonance*, Oxford University Press, Oxford, 2003; (c) R. Drago, *Physical Methods for Chemists*, Surfside Scientific Publishers, Gainesville, 1992.
- (a) G. A. Jeffrey and T. Sakurai, *Prog. Solid State Chem.*, 1964, **1**, 380; (b) C. H. Townes and B. P. Dailey, Determination of electronic structure of molecules from nuclear quadrupole effects, *J. Chem. Phys.*, 1949, **17**, 782.
- S. A. Dikanov and Y. D. Tsvetkov, *Electron Spin Echo Envelope Modulation (ESEEM) Spectroscopy*, CRC Press, Inc., Boca Raton, 1992.
- H. L. Flanagan and D. J. Singel, *J. Chem. Phys.*, 1987, **87**, 5606.
- The synthesis of a 'light blue' salt $[1 \cdot N_2]^+[\text{BAR}_4^F]^-$ was briefly described in a recent publication; however only limited characterisation data (under N₂) were provided. In light of the comprehensive study reported herein, we believe that this salt in fact likely comprises a mixture of predominantly $[1]^+[\text{BAR}_4^F]^-$ and $[1 \cdot N_2]^+[\text{BAR}_4^F]^-$. See: J. B. Geri, J. P. Shanahan and N. K. Szymczak, *J. Am. Chem. Soc.*, 2017, **139**, 5952, and related ESI†.
- The disorder arises due to a mismatch between the lower symmetry of the cation (D_2) and the site symmetry of its location in the crystal (S_4), hence only the Fe atom (which sits directly on the latter) is strictly ordered. See ESI† for full crystallographic details.
- M. Brookhart, M. L. H. Green and G. Parkin, *Proc. Natl. Acad. Sci. U. S. A.*, 2007, **104**, 6908.
- L. R. Doyle, P. J. Hill, G. G. Wildgoose and A. E. Ashley, *Dalton Trans.*, 2016, **45**, 7550.



- 17 THF was selected for these experiments as the solubility of N_2 is known in this medium, in contrast to DFB. See: R. Battino, T. R. Rettich and T. Tominaga, *J. Phys. Chem. Ref. Data*, 1984, **13**, 563.
- 18 D. L. M. Suess, C. Tsay and J. C. Peters, *J. Am. Chem. Soc.*, 2012, **134**, 14158.
- 19 The solution-phase free energy of N_2 binding to a $[P_4(\text{amine})_2] Fe(I)$ -tetraphosphine complex has been recently disclosed (see: D. E. Prokopchuk, E. S. Wiedner, E. D. Walter, C. V. Popescu, N. A. Piro, W. S. Kassel, R. M. Bullock and M. T. Mock, *J. Am. Chem. Soc.*, 2017, **139**, 9291). In contrast to $[1]^+$, N_2 binding was found to be endoergic ($\Delta G_{298} = +0.4(5)$ kcal mol $^{-1}$), which may arise from their use of poorer aryl-phosphine donors vs. alkyl-phosphines in $[1]^+$, translating to a weaker Fe- N_2 bond.
- 20 C. G. Balesdent, J. L. Crossland, D. T. Regan, C. T. López and D. R. Tyler, *Inorg. Chem.*, 2013, **52**, 14178.
- 21 D. Goldfarb, V. Kofman, J. Libman, A. Shanzer, R. Rahmatouline, S. Van Doorslaer and A. Schweiger, *J. Am. Chem. Soc.*, 1998, **120**, 7020.
- 22 R. D. Britt, J. L. Zimmermann, K. Sauer and M. P. Klein, *J. Am. Chem. Soc.*, 1989, **111**, 3522.
- 23 (a) T. A. Stich, J. W. Whittaker and R. D. Britt, *J. Phys. Chem. B*, 2010, **114**, 14178; (b) Simultaneous simulations of both ^{14}N in Fig. S7† are important for a full description of the signal in Fig. 3, yet these are beyond the scope of the present work.
- 24 S. S. Eaton, K. M. More, B. M. Sawant and G. R. Eaton, *J. Am. Chem. Soc.*, 1983, **105**, 6560.
- 25 J. Ballmann, R. F. Munhá and M. D. Fryzuk, *Chem. Commun.*, 2010, **46**, 1013.
- 26 R. H. Crabtree, *Chem. Rev.*, 2016, **116**, 8750.
- 27 Use of D_2 in place of H_2 did not lead to any change in the 1H NMR spectrum, nor did any new resonances appear in the 2H NMR spectrum (other than D_2 itself). These data are consistent with H/D being directly bound to the paramagnetic Fe centre.
- 28 Yellow/green crystals form upon cooling (-35 °C) saturated solutions of $[1]^+[BAr_4^F]^-$ in Et_2O under N_2/H_2 , respectively. However, attempts to obtain structural data from these (presumably of $[1 \cdot N_2/H_2]^+[BAr_4^F]^-$) were thwarted by a rapid loss in crystallinity during handling.
- 29 W. A. Gunderson, D. L. M. Suess, H. Fong, X. Wang, C. M. Hoffmann, G. E. Cutsail III, J. C. Peters and B. M. Hoffman, *J. Am. Chem. Soc.*, 2014, **136**, 14998.
- 30 O. M. Usov, P. S.-T. Choi, J. P. Shapleigh and C. P. Scholes, *J. Am. Chem. Soc.*, 2006, **128**, 5021.
- 31 Exact determination of the fraction of spin on the central Fe(I) ion and the fractions of spin on ligating ^{31}P nuclei is non-trivial due to the nature of the particular s-orbital and p-orbital character of the molecular orbital, and these differ by a factor 36.5 for implications of ligand spin density. See: J. R. Morton and K. F. Preston, *J. Magn. Reson.*, 1969, **30**, 577.
- 32 (a) R. H. Morris and R. Wittebort, *Magn. Reson. Chem.*, 1997, **35**, 243; (b) G. A. Facey, T. P. Fong, D. Gusev, P. M. Macdonald, R. H. Morris, M. Schlaf and W. Xu, *Can. J. Chem.*, 1999, **77**, 1899.
- 33 Diamagnetic *trans*- $[Fe(\text{depe})_2(H)(H_2)]^+[BAr_4^F]^-$ was identified by 1H and $^{31}P\{^1H\}$ NMR spectroscopy. See: M. Bautista, K. A. Earl, R. H. Morris and A. Sella, *J. Am. Chem. Soc.*, 1987, **109**, 3780; M. V. Baker, L. D. Field and D. J. Young, *J. Chem. Soc., Chem. Commun.*, 1988, 546. The rate of conversion was observed to be somewhat variable (depending on solvent, trace impurities *etc.*), and promoted by elevated temperatures and/or extended periods. The mechanism is currently under investigation.
- 34 Computational data for $[1 \cdot H_2]^+$ predict a similar dihedral angle between Fe(PP) planes (16.9°) as that for $[1 \cdot N_2]^+$ (18.9°). DFT calculations on $[1]^+$ and $[1 \cdot N_2]^+$ were in good agreement with experimental X-ray crystallographic data (see ESI†).
- 35 Unfortunately the difficulty of ensuring a consistent temperature throughout the headspace of the sample tube, and the inability to accurately deconvolute the overlapping ESR signals of $[1]^+$, $[1 \cdot N_2]^+$ and $[1 \cdot H_2]^+$ (which become increasingly isotropic at elevated temperatures, likely due to phase change behaviour), precluded a solid-state thermodynamic analysis of binding.
- 36 Theoretical Chemistry, *ADF2016, SCM*, Vrije Universiteit, Amsterdam, The Netherlands, <http://www.scm.com>.
- 37 S. H. Vosko, L. Wilk and M. Nusair, *Can. J. Phys.*, 1980, **58**, 1200.
- 38 A. D. Becke, *Phys. Rev. A: At., Mol., Opt. Phys.*, 1988, **38**, 3098.

