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# Preparation of Ru-doped $TiO_2$ nanotube arrays through anodizing TiRu alloys for bifunctional HER/OER electrocatalysts†

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In this research, Ru-doped $TiO_2$ nanotube arrays (Ru-TNTA) were prepared by anodizing TiRu alloys, and the effects of annealing temperature, Ru content and test temperature on their performances for the hydrogen evolution reaction (HER) and the oxygen evolution reaction (OER) were investigated. The results show that the unannealed Ru-TNTA (a-Ru-TNTA) exhibits superior activity for the HER, and the Ru-TNTA annealed at 450 °C (c-Ru-TNTA) shows excellent activity for the OER. The Ru content of TiRu impacts the electrochemically active surface area (ECSA) and the charge transfer resistance ($R_{ct}$) significantly. When the Ru content of Ru-TNTA is 6%, its performance is optimal. Moreover, the electrocatalytic activity of Ru-TNTA improves with increasing test temperature, and the overpotentials of a-Ru-TNTA and c-Ru-TNTA at 80 °C are 19 mV and 227 mV (10 mA $cm^{-2}$), respectively. Ru-TNTA exhibits excellent electrocatalytic performance for water splitting and good stability, which provides a new idea for the preparation of advanced bifunctional electrocatalysts for water splitting.

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## 1. Introduction

Hydrogen energy is considered as a clean renewable energy source with high energy density.[1–5] Among the various methods of hydrogen production, water electrolysis has the advantages[4,6] of mature technology, simple equipment, high purity of hydrogen, green pollution-free nature and so on. The electrochemical reactions involved include the hydrogen evolution reaction (HER) at the cathode and the oxygen evolution reaction (OER) at the anode.[7,8] The HER is a two-electron reaction and the OER is a four-electron reaction.[9,10] These two reactions possess a high overpotential and high energy consumption, which is not conducive to efficient and rapid water electrolysis.[11] To improve the utilization rate of electrical energy in water electrolysis, researchers have focused on reducing the overpotential through the use of catalysts. At present, the most advanced electrocatalysts are platinum (Pt), iridium oxide ($IrO_2$) and ruthenium oxide ($RuO_2$).[12–14] $RuO_2$ is a bifunctional catalyst that is much cheaper than Pt and $IrO_2$, but still has a higher cost and poor stability, which limits its large-scale application.

Studies have shown that combining $RuO_2$ with other substances or loading it on a certain carrier can not only reduce costs, but also improve stability. For example, Edison *et al.* prepared carbon encapsulated $RuO_2$ nanorods ($RuO_2$ NRs/C) with an onset overpotential of 99 mV and a Tafel slope of 99.4 mV $dec^{-1}$ for the HER, which is a highly stable HER electrocatalyst.[15] Liu *et al.* loaded $RuO_2$ nanoparticles on $MoS_2$ nanosheets to form an efficient and stable electrocatalyst, and it had a low overpotential of 114 mV at a current density of 10 mA $cm^{-2}$ in 1 M KOH.[16] Lim *et al.* prepared $RuO_2/SnO_2$ electrocatalysts by coating $RuO_2$ on $SnO_2$ support to reduce the cost and enhance the stability of the $RuO_2$ catalyst. $RuO_2/SnO_2$ had high OER electrochemical activity and could reduce the noble metal content by about 50%.[17] Lu *et al.* proved that $RuO_2$-modified $Co_3O_4$ can effectively accelerate the OER kinetics of $Co_3O_4$ in a neutral electrolyte, showing excellent OER activity and stability.[18]

Titanium oxide ($TiO_2$) has the advantages of rich resources, environmental friendliness and strong stability, and the performance of $RuO_2$ can also be improved by combining it with $TiO_2$. For example, Ouyang *et al.* synthesized a $RuO_2$ nanoparticle-decorated $TiO_2$ nanobelt array supported on a titanium plate by hydrothermal growth, cation exchange and subsequent annealing, which exhibited excellent catalytic performance and high stability for the HER at all pH values.[19] Li *et al.* prepared $TiO_2$ through *in situ* fabrication on Ti mesh (TM) and then combined it with defective $RuO_2$ nanoparticles (D-$RuO_2$) to obtain a D-$RuO_2/TiO_2$/TM catalyst. This catalyst showed excellent catalytic activity and superior durability for the HER and OER in 1 M KOH.[20] Naslund *et al.* prepared $TiO_2$-doped $RuO_2$ electrodes for the water oxidation reaction, which showed excellent electrocatalytic activity and durability. The interaction between Ti and Ru played an important role in enhancing the OER activity on the mixed oxide electrodes.[21]

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The morphological structure of $TiO_2$ has a significant effect on electrode performance. $TiO_2$ nanotube arrays (TNTA) grown *in situ* on the surface of titanium sheets by anodic oxidation have a large specific surface area and a large number of active sites. These nanotubes are highly ordered and perpendicular to the substrate, providing a smooth channel for ion transport. Meanwhile, TNTA is tightly bound to the Ti substrate and has low resistance. Therefore, TNTA-based electrodes display excellent performance and have been widely used in photocatalysis,[22,23] sensors[24,25] and solar cells.[26–28]

At present, most of the prepared ruthenium–titanium composite oxides are powder materials. When they are bonded to collectors (*e.g.* nickel foam) using a binder to construct electrodes, the prepared electrodes have a large internal resistance and poor stability. Thus, in this work, to obtain $RuO_2$ electrodes with better performance, Ru-TNTA electrodes were prepared by anodizing TiRu alloys and the performance of Ru-TNTA electrodes for the HER and OER was investigated. The overpotential of the optimized Ru-TNTA for the HER and OER is 41 and 349 mV, respectively (10 mA $cm^{-2}$), and the Tafel slope is 67 and 94 mV $dec^{-1}$, respectively. The overall water splitting voltage of the two-electrode electrolytic cell constructed with Ru-TNTA is only 1.496 V at a current density of 10 mA $cm^{-2}$, which also shows remarkable long-term electrochemical durability.

## 2. Experimental

### 2.1. Preparation of electrodes

A 40 mm × 20 mm × 2 mm TiRu sheet (the atomic ratio of Ru is 6.0%) was ultrasonically cleaned in a mixture of deionized water and ethanol. Using a cleaned TiRu sheet as the anode, a Pt sheet as the cathode, and glycol solution containing 0.25% $NH_4F$ and 10% $H_2O$ as the electrolyte, the anodic oxidation reaction was carried out at 40 °C and 40 V for 3 h. After anodization, the sample was taken out, washed with deionized water and air-dried to obtain Ru-doped $TiO_2$ nanotube arrays (denoted as Ru-TNTA).

For comparison, $TiO_2$ nanotube arrays (TNTA) were prepared by anodizing the Ti sheet at 25 V while keeping the other conditions the same as above.

### 2.2. Characterization

Scanning electron microscopy (SEM, Quanta 450 FEG) and transmission electron microscopy (TEM, JEOL JEM-2100F) were used to analyze the morphological structures of the samples. An X-ray diffractometer (XRD, D8 Discover) and an X-ray photoelectron spectrometer (XPS, ESCALAB-250Xi) were employed to analyze the elemental composition and valence states of the samples.

### 2.3. Electrochemical measurements

All electrochemical measurements were carried out in a typical three-electrode system using an electrochemical workstation (CHI 660e, Chenhua, Shanghai). Linear sweep voltammetry (LSV) tests were performed with Ru-TNTA as the working electrode (test area: 4 $cm^2$), a SCE as the reference electrode and a Pt sheet as the counter electrode at a scan rate of 2 mV $s^{-1}$ in 1 mol $L^{-1}$ KOH. The polarization curves were converted to reversible hydrogen electrode potentials using the equation $E_{RHE} = E_{SCE} + 0.2415 + 0.059 \times pH$ and calibrated using 95% IR compensation. The Tafel plots were obtained by fitting the logarithmic equation of overpotential *vs.* current density in the polarization curve ($\eta = a + b \log j$), where $\eta$ represents the overpotential, $a$ is the Tafel constant, $b$ is the Tafel slope, and $j$ is the current density. Electrochemical impedance spectroscopy (EIS) was performed in the frequency range from 1000 kHz to 0.01 Hz. The electrochemically active surface area (ECSA) was calculated from the electrical double-layer capacitance ($C_{dl}$) using the formula ECSA = $C_{dl}/C_s$, where $C_s$ is the ideal specific capacitance of a smooth planar electrode and the general value in alkaline media is 0.04 mF $cm^{-2}$.[29–32] The long-term stability of the catalyst was tested by chronoamperometry.

For comparison with the commercial electrode materials, Pt/C and $IrO_2$ powders were bought from Beijing Mreda Technology Co., Ltd and the corresponding electrodes were prepared as follows and tested similarly. 5 mg of catalyst powder (20 wt% Pt/C or $IrO_2$) was dispersed in a mixed solution of 350 μL of isopropanol, 150 μL of deionized water and 10 μL of 0.5% Nafion, and a uniform dispersion similar to an ink was obtained after ultrasonic treatment for 1 h. The dispersion was dropped onto a clean TiRu sheet and dried at room temperature to produce a working electrode of the HER (OER).

## 3. Results and discussion

### 3.1. Morphology and structure of Ru-TNTA

SEM images of TNTA and Ru-TNTA are shown in Fig. 1. The nanotubes of TNTA are tightly arranged with an outer diameter

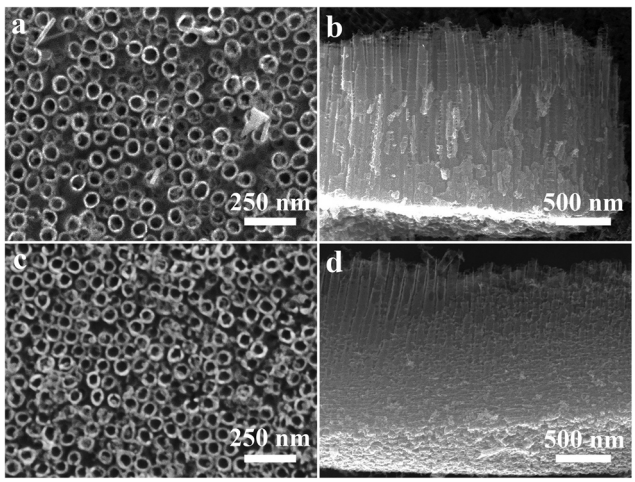


Fig. 1 SEM images of the surface and cross-section of TNTA (a and b) and Ru-TNTA (c and d).

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of about 80 nm (Fig. 1a) and a tube length of about 1586 nm (Fig. 1b). A similar morphology can be observed in Ru-TNTA, where the nanotubes are more tightly bound, but the tube mouths become slightly rougher. The outer diameter and tube length are about 77 nm (Fig. 1c) and 1602 nm (Fig. 1d), respectively. The results in Fig. 1 demonstrate that Ru-TNTA and TNTA have similar tube diameters and lengths.

Fig. 2 shows the XPS spectra of TNTA and Ru-TNTA. As shown in Fig. 2a, in addition to the $C_{1s}$ peak located at around 285.0 eV, Ti, O and Ru elements can be observed in Ru-TNTA, while only Ti and O elements are found in TNTA. Two peaks at 459.0 eV and 464.8 eV in the $Ti_{2p}$ XPS spectra (Fig. 2b) are assigned to $Ti_{2p3/2}$ and $Ti_{2p1/2}$, respectively. The peak spacing between $Ti_{2p3/2}$ and $Ti_{2p1/2}$ is 5.8 eV, revealing that the valence state of Ti element is +4.[20,33–35] In Fig. 2c, the $Ru_{3d}$ peak is located at 281.4 eV, which indicates that Ru in Ru-TNTA is $Ru^{4+}$.[33,36] Fig. 2d shows that the $O_{1s}$ peaks of both TNTA and Ru-TNTA are located at 529.9 eV, suggesting that Ti and Ru in TNTA and Ru-TNTA exist in the form of $TiO_2$ and $RuO_2$. In addition, XPS analysis of the annealed Ru-TNTA was conducted and the results do not change significantly compared with those of the unannealed Ru-TNTA.

Fig. 3 shows the EDS elemental mapping of the Ti, O and Ru elements of Ru-TNTA. It can be seen that the content of Ru is obviously lower than those of Ti and O, and these elements have a substantially uniform distribution in the nanotubes.

The XRD spectra of TNTA and Ru-TNTA are shown in Fig. 4. The diffraction peaks of unannealed TNTA are located at 35.28°, 38.59°, 40.35°, 53.16°, 63.11°, 70.81°, 76.34° and 77.48°, corresponding to the (100), (002), (101), (102), (110), (103), (112) and (201) crystal planes of the Ti substrate (JCPDF#44-1294), respectively. The oxide diffraction peak does not appear, indicating that $TiO_2$ has an amorphous structure. No diffraction peaks of oxides can be found in TNTA annealed

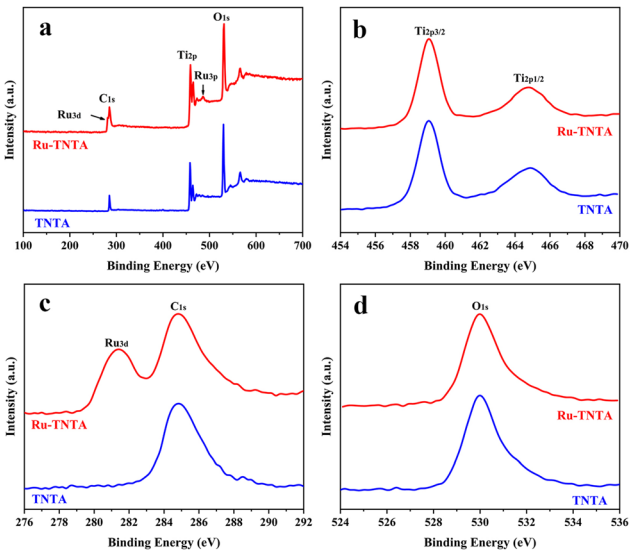


Fig. 2 XPS spectra of TNTA and Ru-TNTA (a: full spectrum, b: $Ti_{2p}$, c: $Ru_{3d}$, and d: $O_{1s}$).

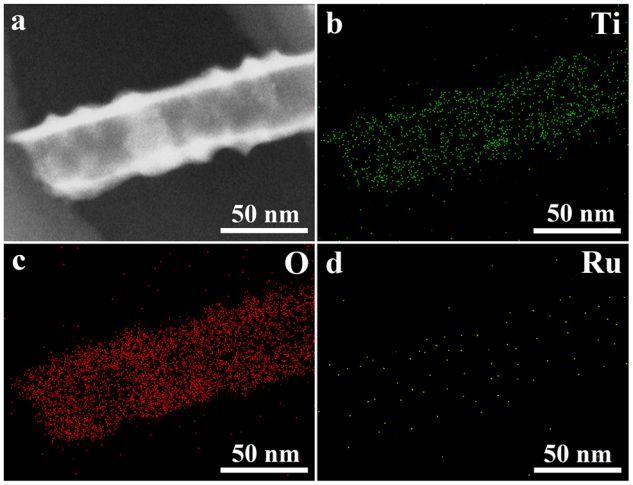


Fig. 3 EDS elemental mapping images of Ru-TNTA (a: TEM image of the scanned area, b: Ti, c: O, and d: Ru).

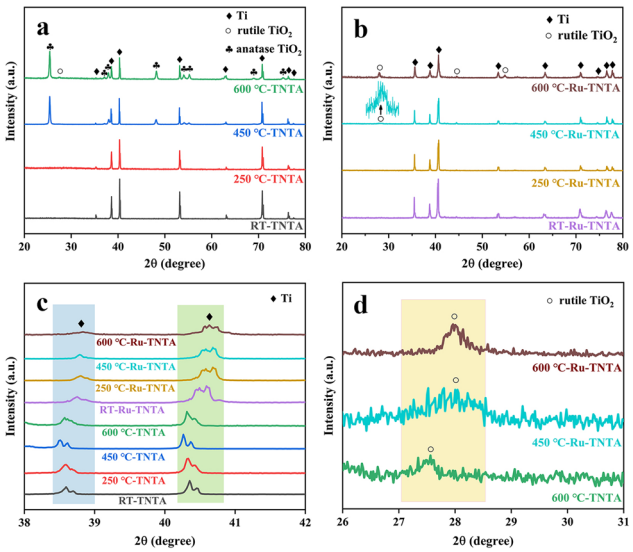


Fig. 4 XRD patterns of TNTA (a) and Ru-TNTA (b) and comparison of the XRD patterns of TNTA and Ru-TNTA (c and d) annealed at different temperatures.

at 250 °C, which demonstrates that $TiO_2$ still has the amorphous structure. When the annealing temperature is increased to 450 °C, the characterized diffraction peaks of oxides at 25.42°, 37.06°, 37.95°, 48.21°, 54.02°, 55.20°, 68.97° and 75.23° are attributed to the (101), (103), (004), (200), (105), (211), (116) and (215) crystal planes of anatase $TiO_2$ (JCPDF#21-1272), respectively. A new diffraction peak at 27.56° is observed when the annealing temperature is increased to 600 °C, corresponding to rutile $TiO_2$ (JCPDF#21-1276), indicating that a part of $TiO_2$ is transformed from the anatase phase to the rutile phase as shown in Fig. 4a.

Fig. 4b shows the XRD spectra of Ru-TNTA at different annealing temperatures. As can be seen from Fig. 4b, after

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drying at room temperature (RT: 20 °C) or annealing at 250 °C, Ru-TNTA only has diffraction peaks of the Ti substrate (35.48°, 38.75°, 40.59°, 53.38°, 63.13°, 70.85°, 74.45°, 76.40° and 77.58°), indicating that the oxide has an amorphous structure. When the annealing temperature is increased to 450 °C, the diffraction peaks of rutile $TiO_2$ appear at 27.96° and 42.15°. As the annealing temperature increases, the diffraction peak intensity of rutile $TiO_2$ gradually becomes stronger, which shows that the crystallization degree of rutile $TiO_2$ is improved.

From the comparison of Fig. 4c and d, it is clear that the diffraction peak of Ti in Ru-TNTA is shifted to a large angle compared to TNTA, indicating the formation of TiRu alloys. The atomic radius of Ru is 1.89 Å and the atomic radius of Ti is 2.00 Å. Due to the fact that the atomic radius of Ru is smaller than that of Ti, a solid solution is formed when Ru is doped into the Ti lattice. At the same time, the Ti lattice shrinks and the crystal plane spacing decreases; thus the diffraction peak of TiRu alloys is shifted to a large angle. Similarly, compared to TNTA, the rutile $TiO_2$ diffraction peaks (27.96° and 42.15°) of Ru-TNTA annealed at 450 °C and 600 °C are also shifted to the right. That is because $TiO_2$ and $RuO_2$ have the same rutile crystal structure; the radius of $Ti^{4+}$ is 0.64 Å and the radius of $Ru^{4+}$ is 0.62 Å.[37] Obviously, the radius of $Ru^{4+}$ is smaller than that of $Ti^{4+}$. Therefore, $Ru^{4+}$ can enter the $TiO_2$ lattice, resulting in the shrinkage of the $TiO_2$ lattice and a decrease in crystal plane spacing, so the diffraction peak is shifted to a large angle. In addition, $TiO_2$ in TNTA annealed at 450 °C has an anatase structure. $TiO_2$ in TNTA annealed at 600 °C appears to be a rutile structure, while $TiO_2$ in Ru-TNTA annealed at 450 °C already has a rutile structure. These results indicate that Ru can promote the formation of rutile $TiO_2$, which is consistent with previous findings. Houskova *et al.* have shown that Ru-doped $TiO_2$ can transform into the rutile phase at lower temperatures compared to undoped $TiO_2$.[38] The diffraction peaks of Ru and its oxides do not appear in all the samples, which may be attributed to the lower doping concentration and the entry of $Ru^{4+}$ into the lattice of $TiO_2$.[39]

Fig. 5 shows the TEM and HRTEM images of Ru-TNTA. From Fig. 5a and b, it can be seen that there is no obvious ordered atomic arrangement in the unannealed Ru-TNTA nanotubes, and there are no diffraction spots or diffraction rings in the selected area electron diffraction (SAED) patterns, indicating the amorphous structure of the unannealed sample. The sample annealed at 450 °C is still a nanotube structure (Fig. 5c), and the HRTEM image reveals an ordered crystal structure (Fig. 5d) with a lattice spacing of 0.32 nm (Fig. 5e), which can be ascribed to the (110) planes of rutile $TiO_2$. The lattice spacing is slightly reduced compared to the standard card, which might be due to the doping of Ru. Analysis of the lattice images by inverse fast Fourier transform (IFFT) indicates the presence of distortions and dislocations in the lattice stripes, confirming the existence of defects and lattice distortions in the crystal structure of the sample. No crystalline plane of $RuO_2$ is observed in the sample, meaning that no $RuO_2$ grains are formed. This result is consistent with the XRD analysis. Both $RuO_2$ and $TiO_2$ have a rutile structure,[40] and the radii of $Ti^{4+}$ and $Ru^{4+}$ only differ by 0.02 Å. Therefore, $Ru^{4+}$ is relatively easy to be doped into the lattice of $TiO_2$. On the basis of the above analyses, the $Ru^{4+}$ in Ru-TNTA is uniformly dispersed in the $TiO_2$ lattice.

### 3.2. Performances of the hydrogen evolution reaction

The electrocatalytic performance of TNTA and Ru-TNTA for the HER was investigated by linear sweeping voltammetry (LSV). Fig. 6a shows the polarization curves of TNTA annealed at different temperatures. From Fig. 6a, it is known that at the

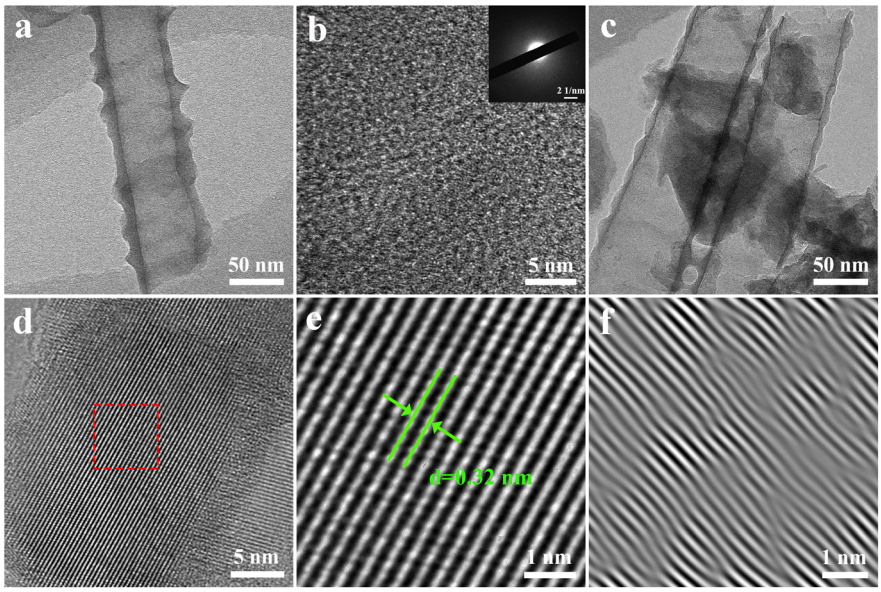


**Fig. 5** TEM (a) and HRTEM (b) images and SAED pattern (inset) of unannealed Ru-TNTA, TEM (c) and HRTEM (d) images of Ru-TNTA annealed at 450 °C and enlarged (e) and IFFT (f) images of selected area.

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same electrode potential, the current density of the electrode decreases as the annealing temperature increases. The current density of unannealed TNTA is the largest, with an overpotential of 591 mV at a current density of 10 mA $cm^{-2}$. The Tafel slope was estimated using the Tafel equation to evaluate HER kinetics.[41] Theoretically, a lower Tafel slope indicates a better electrocatalytic hydrogen evolution ability of the electrode. Fig. 6b shows Tafel plots calculated from the polarization curves in Fig. 6a. After linear fitting, the Tafel slopes of electrodes with four different annealing temperatures (from low to high) are 317 mV $dec^{-1}$, 331 mV $dec^{-1}$, 364 mV $dec^{-1}$ and 374 mV $dec^{-1}$, respectively. The Tafel slope of unannealed TNTA is the smallest, indicating its best HER performance among all the TNTA electrodes. The performance of TNTA electrodes for the HER becomes worse with the increase of annealing temperatures.

Fig. 6c shows the polarization curves of Ru-TNTA annealed at different temperatures. The unannealed Ru-TNTA only requires an overpotential of 41 mV to obtain a current density of 10 mA $cm^{-2}$, which is lower than those of Ru-TNTA annealed at 250 °C (54 mV), 450 °C (81 mV) and 600 °C (111 mV). Therefore, the unannealed Ru-TNTA has the highest catalytic activity for the HER. Along with the increase of annealing temperature, the catalytic activity of Ru-TNTA for the HER becomes worse. Fig. 6d shows the Tafel plots of Ru-TNTA at different annealing temperatures (low to high) with the corresponding Tafel slopes of 67 mV $dec^{-1}$, 81 mV $dec^{-1}$, 86 mV $dec^{-1}$ and 114 mV $dec^{-1}$, respectively. The unannealed Ru-TNTA has the smallest Tafel slope (67 mV $dec^{-1}$), indicating its fast HER kinetics. The performance of Ru-TNTA for the HER becomes worse as the annealing temperature increases. The results in Fig. 6 suggest that the annealing temperature greatly influences the performance of electrodes for the HER. When the annealing temperature changes, the HER performance of TNTA and Ru-TNTA has the same variation tendency, and the unannealed electrodes have the best HER catalytic activity.

### 3.3. Performances of the oxygen evolution reaction

The electrocatalytic performance of TNTA and Ru-TNTA for the OER is shown in Fig. 7. Fig. 7a shows the polarization curves of TNTA. Along with the increase of annealing temperature, the current density of the TNTA electrode increases first and then decreases (Fig. 7a). The current density of TNTA annealed at 450 °C is the largest. The overpotential is 1836 mV at a current density of 10 mA $cm^{-2}$, and the Tafel slope is 554 mV $dec^{-1}$ (Fig. 7b). The variation tendency of the OER catalytic performance of Ru-TNTA at different annealing temperatures is the same as that of TNTA (Fig. 7c and a). The highest OER activity is observed for Ru-TNTA annealed at 450 °C, with an overpotential of 349 mV at a current density of 10 mA $cm^{-2}$. According to the Tafel plots in Fig. 7d, the Tafel slopes of Ru-TNTA electrodes dried at room temperature and annealed at 250 °C, 450 °C and 600 °C are 129 mV $dec^{-1}$, 134 mV $dec^{-1}$, 94 mV $dec^{-1}$ and 119 mV $dec^{-1}$, respectively. As shown in Fig. 7c and d, Ru-TNTA annealed at 450 °C possesses the smallest overpotential and Tafel slope, and its OER catalytic performance is the best, which is consistent with the analysis results of TNTA.

It is known from Fig. 6 and 7 that the HER and OER overpotentials of TNTA are larger, while those of Ru-TNTA decrease significantly, indicating that the electrocatalytic activity of Ru-TNTA is much better than that of TNTA. Due to the characteristics of the Ru element, $RuO_2$ has excellent electrocatalytic activity for the HER and OER. The special tube structure of TNTA can not only increase the surface area, but also promote mass transfer, which is conducive to the formation and detachment of hydrogen and oxygen. For Ru-TNTA, $Ru^{4+}$ is highly dispersed in the $TiO_2$ lattice and there are more Ru

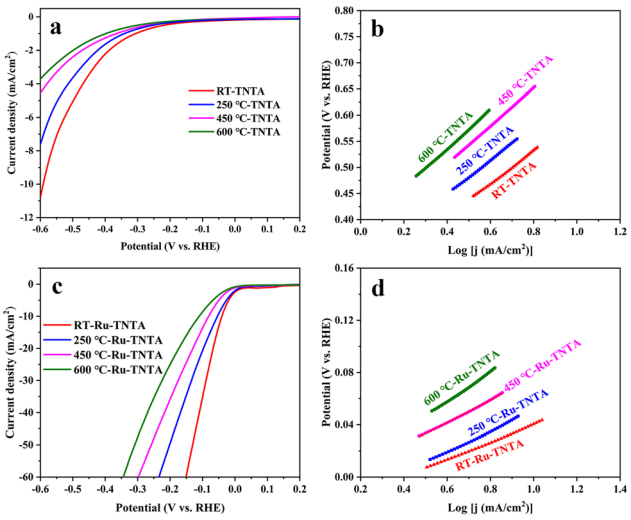


Fig. 6 HER performances of TNTA (a and b) and Ru-TNTA (c and d) annealed at different temperatures.

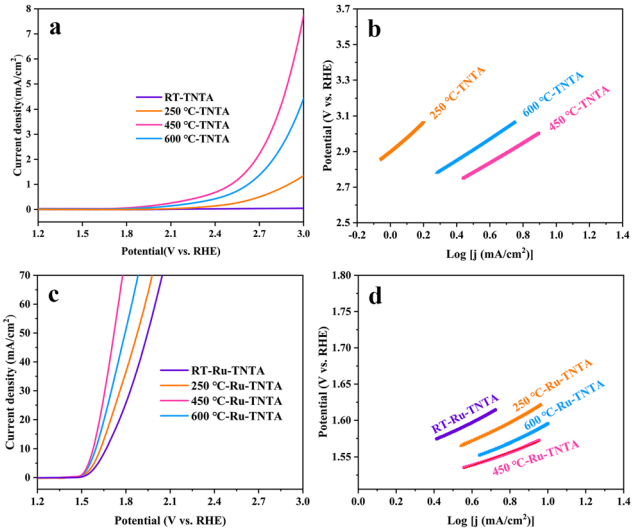


Fig. 7 OER performances of TNTA (a and b) and Ru-TNTA (c and d) annealed at different temperatures.

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active sites on the inner and outer surfaces of the nanotubes, resulting in a significant increase in electron transfer efficiency and a full utilization of Ru; thus the Ru-TNTA electrode exhibits high electrocatalytic activity.[9,42] Moreover, it is known from Fig. 6 and 7 that the annealing temperature has a marked effect on the HER and OER performances of the electrodes. The unannealed Ru-TNTA shows the best HER performance and the Ru-TNTA annealed at 450 °C has the best OER performance.

As shown in Fig. 4, the unannealed TNTA and Ru-TNTA have amorphous structures. There are a large number of hydroxyl groups on the surface of amorphous oxide, which can provide more active sites and make it easy to achieve hydrogen adsorption and detachment. Therefore, the unannealed electrodes exhibit higher HER performance.[43–46] As the annealing temperature increases, the hydroxyl groups will condense and dehydrate, and the surface electronic state of the oxide changes, which is not conducive to the adsorption and detachment of hydrogen, thus resulting in a decrease in the HER catalytic activity.

Along with the increase of annealing temperature, the order degree of atoms increases and the oxide changes from an amorphous structure to a crystalline structure. Meanwhile, the electrical conductivity is improved and the oxide surface is more favorable for oxygen adsorption and detachment. Hence, TNTA and Ru-TNTA after annealing show greater catalytic activity for the OER.[47] The electrodes annealed at 450 °C are almost completely crystallized. When the annealing temperature is higher than 450 °C, the further growth of grains will reduce defects and active sites, and the change of the crystal phase will increase the electron migration resistance, resulting in a decrease in the OER catalytic performance.

### 3.4. Effect of Ru content

*x*Ru-TNTA samples (*x* represents the atomic percentage of Ru in the alloys) were prepared using TiRu alloys with different Ru contents. Fig. 8 shows the HER performance of the amorphous structure sample dried at room temperature (denoted as a-*x*Ru-TNTA) and the OER performance of the crystallized sample annealed at 450 °C (denoted as c-*x*Ru-TNTA). It is known from Fig. 8a and b that the HER performance of a-*x*Ru-TNTA increases first and then decreases along with increasing Ru content. The HER performance of a-6Ru-TNTA is optimal when the Ru content in the alloy is 6%. As shown in Fig. 8c and d, the overpotential of c-*x*Ru-TNTA decreases first and then increases with increasing Ru content. c-6Ru-TNTA displays the best OER catalytic activity when the Ru content in the alloy is 6%. Meanwhile, the Tafel slope of c-6Ru-TNTA is the smallest, meaning that the OER kinetics of c-6Ru-TNTA is the fastest.

To investigate the reason for the effect of Ru content on the electrocatalytic performance of Ru-TNTA, the ECSAs and EIS spectra of the electrodes with different Ru contents were tested. Fig. S1† shows the CV curves of TNTA and *x*Ru-TNTA in 1 M KOH at different sweep speeds. The current density *vs.* sweep speed curves (Fig. 9a) of all electrodes can be obtained

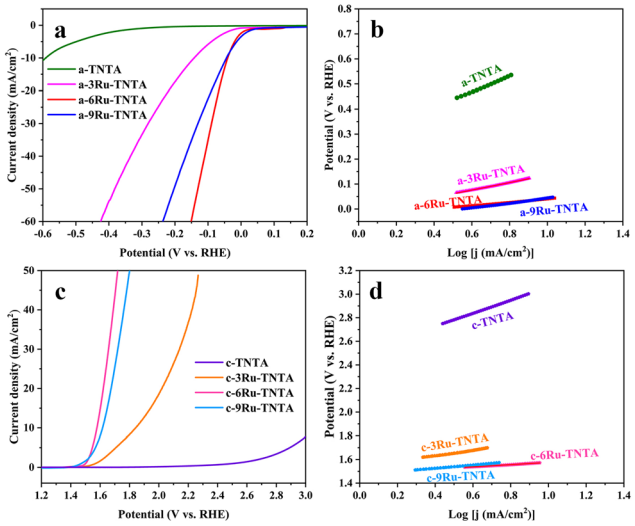


Fig. 8 HER performance of a-*x*Ru-TNTA (a and b) and OER performance of c-*x*Ru-TNTA (c and d).

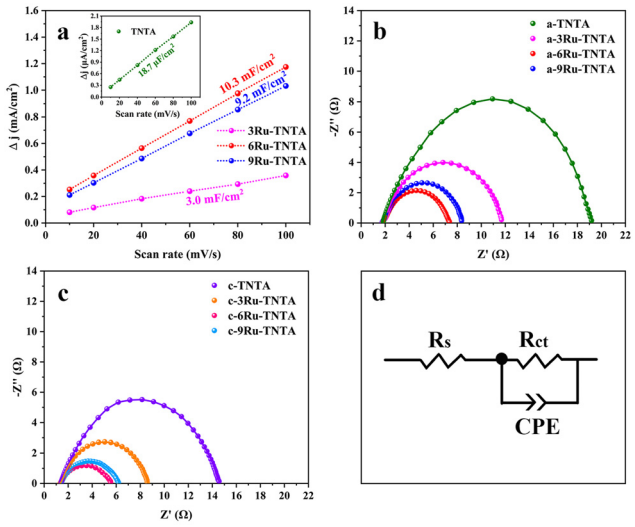


Fig. 9 The current density *vs.* sweep speed curves of *x*Ru-TNTA (a), Nyquist plots of a-*x*Ru-TNTA for the HER (b) and c-*x*Ru-TNTA for the OER (c) and the respective equivalent circuits used to fit the data (d).

from Fig. S1,† and all the curves are approximately straight lines, which means that all electrodes exhibit an electric double-layer capacitance under the test conditions. The slope of each fitting curve in Fig. 9a represents $C_{dl}$. The $C_{dl}$ values of TNTA, 3Ru-TNTA, 6Ru-TNTA and 9Ru-TNTA are 18.7 μF cm$^{-2}$, 3.0 mF cm$^{-2}$, 10.3 mF cm$^{-2}$ and 9.2 mF cm$^{-2}$, and the corresponding ECSAs are 0.47 cm$^2$, 75.0 cm$^2$, 257.5 cm$^2$ and 230.0 cm$^2$, respectively. The ECSAs of all electrodes suggest that doping Ru into TNTA can significantly increase the electrochemically active surface area. Moreover, the Ru content in Ru-TNTA has a significant effect on ECSA. With increasing Ru content, the ECSA of electrodes first increases and then decreases, and a maximum value appears at a Ru content of

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6% (6Ru-TNTA). Fig. 9b–d show the EIS spectra of a-*x*Ru-TNTA for the HER (Fig. 9b) and c-*x*Ru-TNTA for the OER (Fig. 9c) with different Ru contents and the corresponding equivalent circuits used to fit the data (Fig. 9d). As can be seen in Fig. 9b, the charge transfer resistances for the HER of a-TNTA, a-3Ru-TNTA, a-6Ru-TNTA and a-9Ru-TNTA are 17.36 Ω, 9.33 Ω, 5.43 Ω and 6.20 Ω, respectively. This indicates that the $R_{ct}$ of each electrode is obviously different and Ru doping can considerably reduce the $R_{ct}$ of the electrodes. With increasing Ru content, the $R_{ct}$ of electrodes first decreases and then increases slightly, and the $R_{ct}$ of a-6Ru-TNTA is the smallest. The $R_{ct}$ in the EIS spectra of c-*x*Ru-TNTA for the OER (Fig. 9c) has a similar variation trend to that in Fig. 9b, where c-6Ru-TNTA exhibits the smallest $R_{ct}$ of about 4.14 Ω. The results of Fig. 9b and c are consistent with those of Fig. 9a.

The above results could possibly be due to two main reasons. One is that the conductivity of $RuO_2$ is noticeably better than that of $TiO_2$, so doping $RuO_2$ in $TiO_2$ can improve its conductivity; the second is that the radii of $Ti^{4+}$ and $Ru^{4+}$ are different. Doping TNTA with Ru causes distortion of the $TiO_2$ lattice and generation of defects, resulting in an increase in the surface area of the electrode. The combined effect of the two factors will increase the ECSA of *x*Ru-TNTA. With the increase in Ru content, the amount of $RuO_2$ in TNTA increases, so the conductivity of the electrodes improves and the ECSA expands. However, when the Ru content in Ru-TNTA is too high, there are too many defects. Moreover, the integrity and continuity of the $TiO_2$ phase are severely damaged, which will cause deterioration of conductivity along with a concomitant decrease in the ECSA. According to the results of the ECSA and EIS, doping TNTA with Ru can significantly increase the ECSA and decrease the $R_{ct}$ of the electrodes. The larger ECSA represents more catalytically active sites and the smaller $R_{ct}$ represents the faster charge transfer rate and electrochemical process, leading to the improvement of the catalytic activity of the electrodes. At the same time, the Ru content in Ru-TNTA affects the ECSA and $R_{ct}$. Thus, 6Ru-TNTA with the largest ECSA and the smallest $R_{ct}$ displays optimal electrocatalytic performance.

For comparison with the electrocatalytic performance of *x*Ru-TNTA, the HER and OER performances of TiRu alloys with different Ru contents (denoted as TiRu*x*, where *x* represents the atomic percentage of Ru in the alloy) were also tested. The relevant results are shown in Fig. 10. It can be seen from Fig. 10a and b that the HER performance of the Ti sheet is poor, and its current density increases slowly. The overpotentials of TiRu3, TiRu6 and TiRu9 are 254 mV, 188 mV and 169 mV (10 mA $cm^{-2}$), and the Tafel slopes are 248 mV $dec^{-1}$, 220 mV $dec^{-1}$ and 187 mV $dec^{-1}$, respectively. As shown in Fig. 10c and d, with increasing Ru content in the TiRu alloys, the overpotentials and Tafel slopes of TiRu*x* electrodes decrease, indicating enhanced and stabilized HER performance. Furthermore, the HER performance of TiRu*x* is considerably lower than that of a-*x*Ru-TNTA. From Fig. 10e–h, the variation tendency of the OER performance of TiRu*x* alloys is the same as that in Fig. 10a–d. Meanwhile, the performance of

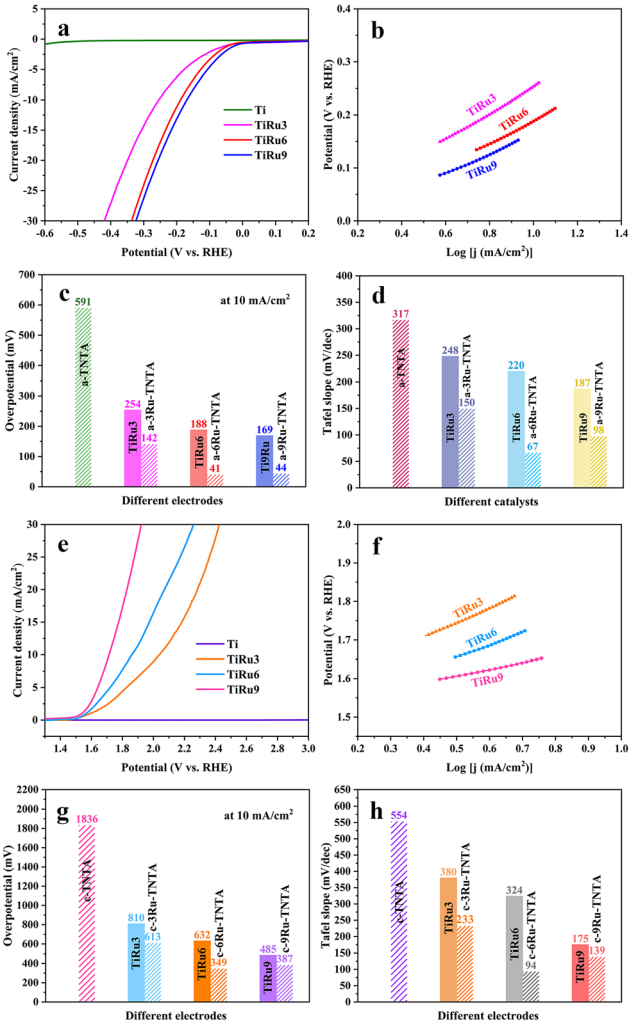


**Fig. 10** HER performance of TiRu*x* (a and b) and the comparison histogram with a-*x*Ru-TNTA (c and d) and OER performance of TiRu*x* (e and f) and the comparison histogram with c-*x*Ru-TNTA (g and h).

TiRu*x* for the OER is significantly inferior to that of the corresponding c-*x*Ru-TNTA.

The significant difference between TiRu*x* and *x*Ru-TNTA shown in Fig. 10 can be ascribed to their different surface areas. The surface area of TiRu*x* approximates the physical area of the alloy (about 4 $cm^2$), while the electrochemically active surface area of *x*Ru-TNTA is several times larger than that of TiRu*x* and more active sites can be provided. Besides, differences in the properties of metallic Ru and $RuO_2$ can have an impact on the electrocatalytic activity of TiRu*x* and *x*Ru-TNTA.

### 3.5. Effect of test temperature

Fig. 11 shows the performances of a-Ru-TNTA (*i.e.*, a-6Ru-TNTA) for the HER and c-Ru-TNTA (*i.e.*, c-6Ru-TNTA) for the OER at different test temperatures. For the HER, the overpotentials of a-Ru-TNTA at room temperature, 40 °C, 60 °C and 80 °C are 41 mV, 34 mV, 28 mV and 19 mV (10 mA $cm^{-2}$), and

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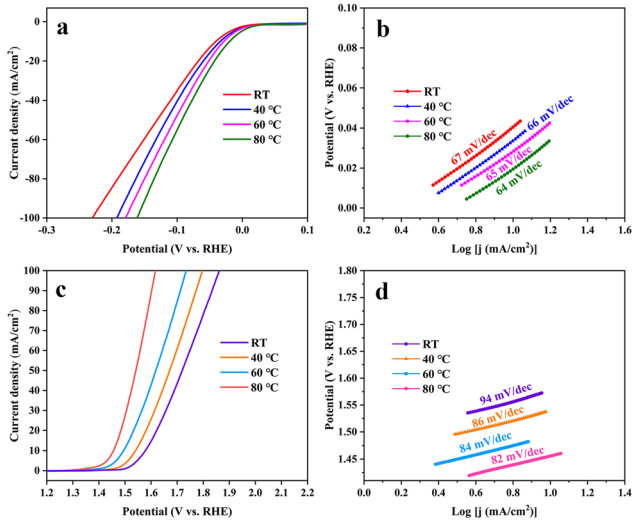


Fig. 11 HER performance of a-*x*Ru-TNTA (a and b) and OER performance of c-*x*Ru-TNTA (c and d) at different test temperatures.

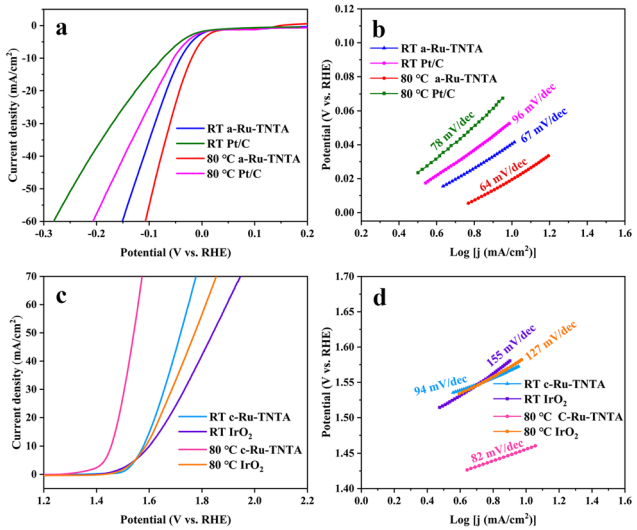


Fig. 12 HER performances of a-Ru-TNTA and commercial Pt/C (a and b) and OER performances of c-Ru-TNTA and $IrO_2$ (c and d) at room temperature and 80 °C test temperature.

the corresponding Tafel slopes are 67 mV dec$^{-1}$, 66 mV dec$^{-1}$, 65 mV dec$^{-1}$ and 64 mV dec$^{-1}$, respectively (Fig. 11a and b). For the OER, the overpotentials of c-Ru-TNTA at room temperature, 40 °C, 60 °C and 80 °C are 349 mV, 311 mV, 265 mV and 227 mV (10 mA cm$^{-2}$), respectively, and the corresponding Tafel slopes are 94 mV dec$^{-1}$, 86 mV dec$^{-1}$, 84 mV dec$^{-1}$ and 82 mV dec$^{-1}$, respectively (Fig. 11c and d). As the test temperature increases, the overpotential and Tafel slope of Ru-TNTA gradually become lower, demonstrating that the test temperature can affect the HER and OER. At a higher reaction temperature, the reactant activity is improved and the reaction activation energy, interfacial tension and mass transfer resistance are reduced, which is more beneficial for the HER and OER, and therefore the overpotentials of electrodes are lowered. Moreover, the electrocatalytic activity of Ru-TNTA also changes with different test temperatures.

### 3.6. Comprehensive performances of the optimized electrodes

The performance comparison of the optimized Ru-TNTA with commercial Pt/C and $IrO_2$ is shown in Fig. 12. As can be seen from Fig. 12a and b, the HER performance of commercial Pt/C at the test temperature of 80 °C ($\eta_{10}$ = 54 mV, Tafel slope = 78 mV dec$^{-1}$) is better than that at room temperature ($\eta_{10}$ = 74 mV, Tafel slope = 96 mV dec$^{-1}$), but both are lower than that of a-Ru-TNTA (80 °C: $\eta_{10}$ = 19 mV, Tafel slope = 64 mV dec$^{-1}$; RT: $\eta_{10}$ = 41 mV, Tafel slope = 67 mV dec$^{-1}$). The electrocatalytic performance of a-Ru-TNTA for the HER is considerably superior to those of recently reported electrodes (Table S1†).

From Fig. 12c and d, the OER performance of $IrO_2$ at 80 °C ($\eta_{10}$ = 357 mV, Tafel slope = 127 mV dec$^{-1}$) is improved compared to that at room temperature ($\eta_{10}$ = 371 mV, Tafel slope = 155 mV dec$^{-1}$), but the change is not substantial. The performance of c-Ru-TNTA for the OER at room temperature ($\eta_{10}$ = 349 mV, Tafel slope = 94 mV dec$^{-1}$) is superior to that of $IrO_2$. Moreover, its performance for the OER is substantially improved at 80 °C ($\eta_{10}$ = 227 mV, Tafel slope = 82 mV dec$^{-1}$). The performance of c-Ru-TNTA for the OER is superior to those of previously reported electrodes (Table S2†).

The stability measurement results are given in Fig. S2.† After 5000 CV cycles, the HER polarization curve of a-Ru-TNTA almost overlaps with the initial curve (Fig. S2a†), and the OER polarization curve of c-Ru-TNTA deviates marginally from the initial curve, with its increasing overpotential of only 14 mV (10 mA cm$^{-2}$) and a smaller activity loss (Fig. S2b†). During the electrolysis testing at a constant voltage for 72 hours continuously at room temperature, the current density of a-Ru-TNTA is nearly invariable (Fig. S2c†) while that of c-Ru-TNTA is slightly reduced (Fig. S2d†). These results demonstrate that both a-Ru-TNTA and c-Ru-TNTA possess outstanding stability.

Based on the above results, Ru-TNTA exhibits excellent HER (a-Ru-TNTA) and OER activities (c-Ru-TNTA), and can be used as a bifunctional electrocatalytic electrode. Fig. 13a shows the current density *vs.* voltage curves for water splitting with different cathodes and anodes. At room temperature, the a-Ru-TNTA||c-Ru-TNTA electrode group exhibits a voltage of 1.496 V at a current density of 10 mA cm$^{-2}$, which is significantly lower than those of a-Ru-TNTA||a-Ru-TNTA (1.545 V), c-Ru-TNTA||c-Ru-TNTA (1.567 V) and Pt/C||$IrO_2$ (1.604 V). At the test temperature of 80 °C, the voltage of the a-Ru-TNTA||c-Ru-TNTA electrode group (1.480 V) is still markedly lower than that of Pt/C||$IrO_2$ (1.563 V). Fig. 13b shows the stability measurement of the a-Ru-TNTA||c-Ru-TNTA electrode group. There is no apparent decay in current density after 72 hours of continuous overall water splitting at room temperature. The results in Fig. 13 demonstrate that the a-Ru-TNTA||c-Ru-TNTA electrode group has excellent water splitting performance and good stability, much superior to those reported in previous literature (Table S3†).

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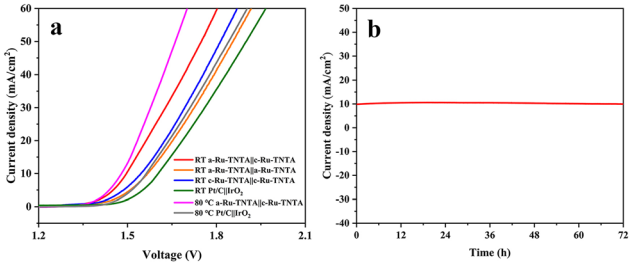


**Fig. 13** Current density *vs.* voltage curves for water splitting with different cathodes and anodes (a) and stability of the a-Ru-TNTA||c-Ru-TNTA electrode group (b).

## 4. Conclusions

In conclusion, Ru-TNTA electrodes were prepared from TiRu alloys by anodic oxidation, and their performances for the HER and OER are closely related to the annealing temperature, Ru content and testing temperature. The unannealed Ru-TNTA (a-Ru-TNTA) with an amorphous structure has more hydroxyl groups on the oxide surface, and the adsorption and detachment of hydrogen are easier, exhibiting excellent HER activity. Ru-TNTA annealed at 450 °C (c-Ru-TNTA) has a rutile structure and its oxide surface is more favorable for oxygen adsorption and detachment, exhibiting superior OER activity. The Ru content in Ru-TNTA has an obvious effect on the ECSA and $R_{ct}$. The lowest $R_{ct}$ and the largest ECSA contribute to the excellent electrochemical performance of Ru-TNTA when the Ru content is 6%. At room temperature, the overpotential of a-Ru-TNTA is 41 mV for the HER and 349 mV for c-Ru-TNTA for the OER, and at 80 °C they are reduced to 19 mV and 227 mV (at 10 mA $cm^{-2}$), which are lower than those of commercial Pt/C and $IrO_2$, respectively. The overall water splitting voltage of the a-Ru-TNTA||c-Ru-TNTA electrode group is also remarkably smaller than that of Pt/C||$IrO_2$. In addition, Ru-TNTA exhibits excellent electrochemical stability and durability, and is a bifunctional electrocatalytic electrode with promising applications.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

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