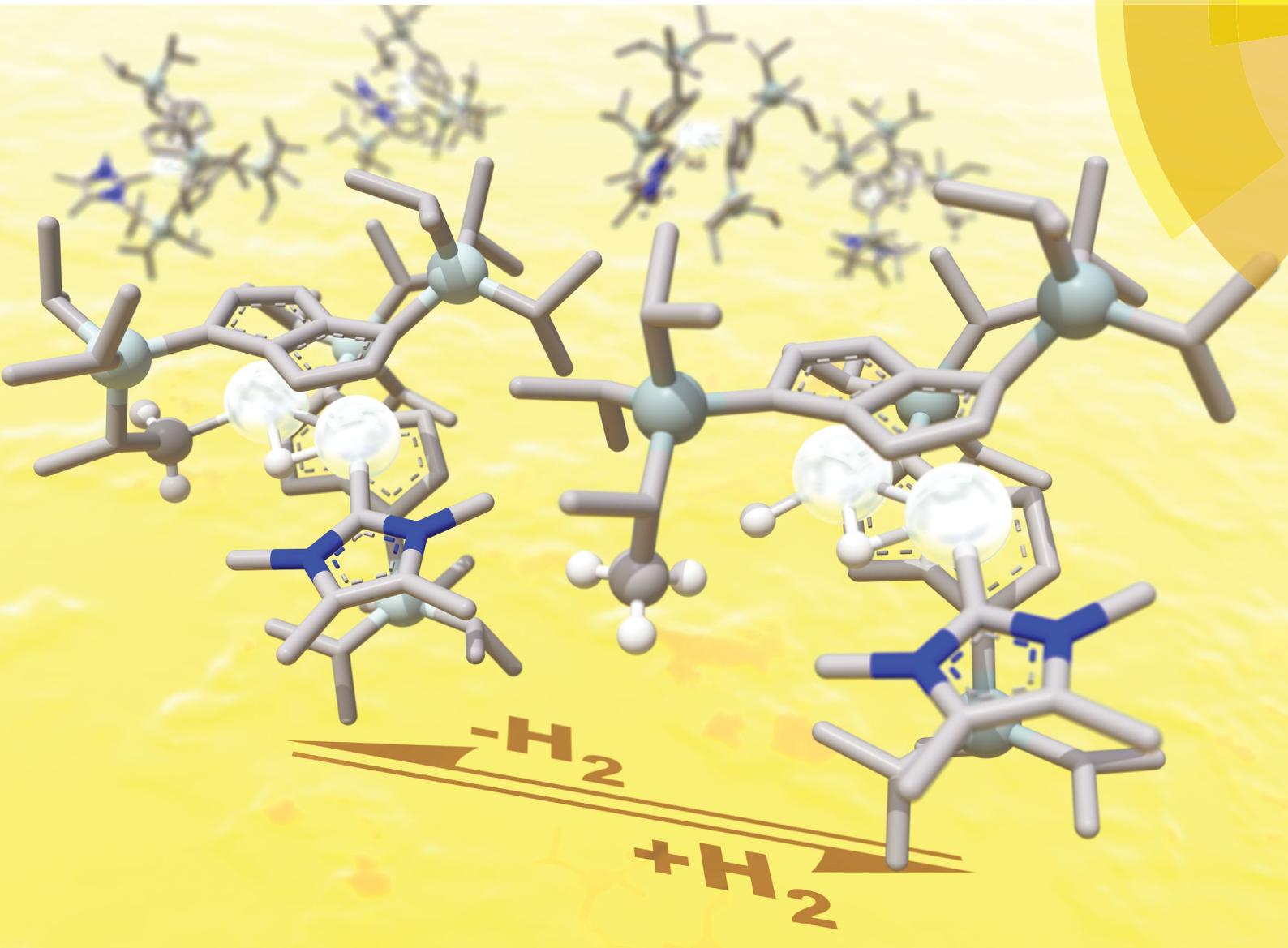


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C–H and H–H activation at a di-titanium centre



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The reaction of the bis(pentalene)dititanium complex $\text{Ti}_2(\mu:\eta^5,\eta^5\text{-Pn}^+)_2$ ($\text{Pn}^+ = \text{C}_8\text{H}_4(1,4\text{-Si}^i\text{Pr}_3)_2$) with the N-heterocyclic carbene 1,3,4,5-tetramethylimidazol-2-ylidene results in intramolecular C–H activation of one of the iPr methyl groups of a Pn^+ ligand and formation of a “tucked-in” bridging hydride complex. The “tuck-in” process is reversed by the addition of hydrogen, which yields a dihydride featuring terminal and bridging hydrides.

Group IV metallocene chemistry has been instrumental in the development of organometallic chemistry,¹ elucidating fundamental aspects of bonding² and reactivity,³ especially *via* the synthesis, isolation and study of low valent metallocene complexes.^{1,4} In this context, the synthesis and isolation of such complexes bearing hydride ligands has been important;⁵ for example, such low or mixed valence hydride complexes have been shown to promote or be involved in the fixation of N_2 to NH_3 .⁶ In this communication, we present the first examples of the synthesis of bridging titanium hydrides under non-reducing conditions, *via* the reaction of bis(pentalene)dititanium complex $\text{Ti}_2(\mu:\eta^5,\eta^5\text{-Pn}^+)_2$ ($\text{Pn}^+ = \text{C}_8\text{H}_4(1,4\text{-Si}^i\text{Pr}_3)_2$) with an N-heterocyclic carbene and subsequent hydrogenolysis.

We have previously reported on the reactivity of $\text{Ti}_2(\mu:\eta^5,\eta^5\text{-Pn}^+)_2$ ($\text{Pn}^+ = \text{C}_8\text{H}_4(1,4\text{-Si}^i\text{Pr}_3)_2$) (**1**) towards a variety of small molecules and π -acceptor ligands.⁷ In order to gain a better insight into the reactivity of (**1**), we decided to study its interaction with strong σ -donor ligands. When (**1**) was treated with an excess of PMe_3 no reaction was observed. However, addition of 1,3,4,5-tetramethylimidazol-2-ylidene (**2**) to (**1**) in toluene at 0 °C resulted in an immediate colour change from crimson red to a dark pine green (Scheme 1).

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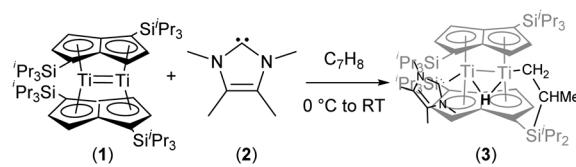
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† Dedicated to Phil on the occasion of his 65th birthday- Lá Breith Sona!

‡ Electronic supplementary information (ESI) available: Full experimental and computational details, and X-ray data. CCDC 1577130–1577132. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c7cc07726b

The formulation of the new complex (**3**) as a “tucked-in” hydride resulting from C–H activation of one of the iPr methyl groups of a Pn^+ induced by addition of the strongly donating NHC (**2**) was initially confirmed by NMR. In particular, the ^1H -NMR spectrum showed two inequivalent pentalene ring environments, a sharp singlet at –7.91 ppm for the bridging hydride (T_1 479 ms), with one of the diastereotopic protons of the “tucked-in” CH_2 group appearing as an overlapping dd at –2.91 ppm whilst the other was largely obscured by the complex aliphatic region of the spectrum at *ca* 1.37 ppm; the coordination of the NHC was confirmed by the observation of a peak at 197.78 ppm in the $^{13}\text{C}\{^1\text{H}\}$ -NMR spectrum.

The structure of (**3**) was confirmed by X-ray diffraction and is shown in Fig. 1. The NHC coordinates to one of the Ti centres (Ti_2 in Fig. 1), while one of the methyl groups on a TIPS substituent has been cyclometallated on the other Ti centre (Ti_1 in Fig. 1) with concurrent formation of a bridging hydride.⁸ The Ti–Ti bond has been retained but lengthened to 2.5610(8) Å (from 2.399(2) Å in (**1**)^{7e}) and is typical of a single bond. The Ti–C(carbene) bond (2.300(2) Å; $\text{Ti}_2\text{-C}1$ in Fig. 1) is within the range of 2.2–2.35 Å reported for other Ti–NHC complexes.⁹ The Ti–H bond lengths (*i.e.* $\text{Ti}_2\text{-H}1$: 1.72(3) Å, $\text{Ti}_1\text{-H}1$: 1.79(3) Å) are identical within esd's and are similar to previously reported monomeric¹⁰ and dimeric^{6d,6b,11} titanium hydrides as well as Ti(IV) alumino-¹² and borohydrides¹³ (1.7–1.9 Å), although it has to be noted that, probably due to the topology of the hydride ligand in (**3**), these Ti–H bond distances fall at the shorter end of the known range. Due to this unique topology, the Ti–H–Ti bond angle (93.4(13)°) approaches a right angle and is the most acute ever observed in dimeric titanium hydrides.^{6b,6d,11}



Scheme 1 Synthesis of a new *syn*-bimetallic hydride titanium cluster.



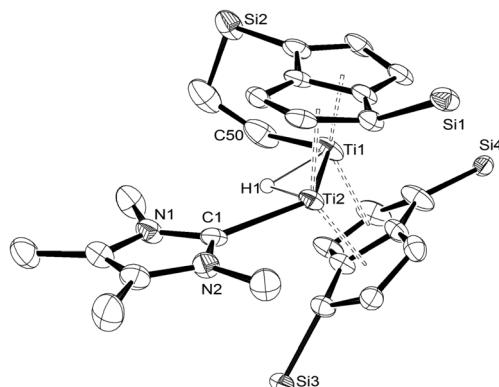


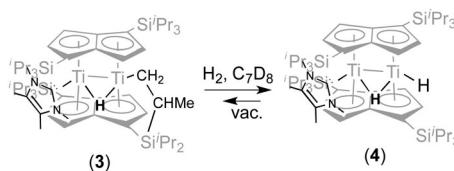
Fig. 1 ORTEP diagram of the molecular structure of (3) displaying 50% probability ellipsoids. ^1Pr groups omitted for clarity.

It is also worth noting that the addition of (2) to (1) results in the formal oxidation of the two Ti centres (*i.e.* from +2 to +3), and employing the CBC model each Ti has a count of 18 e (16 e in (1)).¹⁴

With a view to synthesising a new hydride derivative *via* σ -bond metathesis of the Ti-CH₂ bond in (3) with dihydrogen, 2 bar of H₂ was added to a C₇D₈ solution of (3) in an NMR tube. This indeed resulted in clean conversion of (3) (100% spectroscopic yield) to a new complex (4) (Scheme 2).

Compared to the ^1H NMR spectrum of (3), (4) displays a new, broader hydride peak ($\Delta\nu_{1/2} = 29$ Hz) at -8.82 ppm at room temperature, whilst the signal for the “tucked-in” CH₂ group has disappeared completely; the NHC is still coordinated ($^{13}\text{C}\{^1\text{H}\}$ δ 198.27 ppm). Removal of the H₂ overpressure by freeze-thaw-degassing showed that (4) is persistent in solution, although some regeneration of (3) was observed (Scheme 2). Addition of H₂ to a solution of (3), *via* a Toepler pump, showed that for the conversion of (3) to (4) to occur quickly (minutes) 5 eq of H₂ are required (when 1–2 equivalents of H₂ were added, complete conversion to (4) occurred after *ca.* 1 week). The rate of reaction was also found to be pressure dependent: when (3) was exposed to an atmosphere of 10% H₂ in N₂ at 1.5 bar but in an amount corresponding to only 1 equivalent of H₂ the reaction was again complete in minutes.

Variable temperature ^1H NMR studies showed that the broad hydride peak at -8.82 ppm in (4) becomes fully resolved into a doublet at $0\text{ }^\circ\text{C}$ (with no further change below that temperature and down to $-70\text{ }^\circ\text{C}$) with a T_1 of 310 ms, with the concomitant appearance of a second doublet centred at 2.17 ppm (T_1 336 ms), which is too broad to be observable at room temperature (Fig. 2); these two signals are related by a coupling constant of $J_{\text{HH}} = 11$ Hz. EXSY spectroscopy (in both



Scheme 2 Hydrogenolysis of (3) to afford (4)

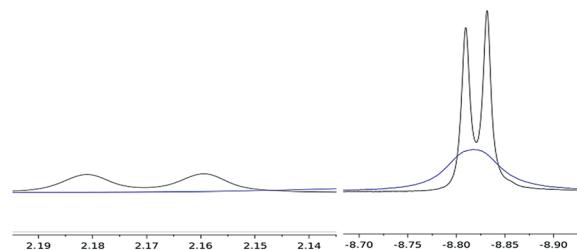


Fig. 2 Hydride peaks in (4) at RT (blue) and at $0\text{ }^\circ\text{C}$ (black).

the presence and absence of an H₂ overpressure) confirmed that these two protons exchange at $30\text{ }^\circ\text{C}$ while at $0\text{ }^\circ\text{C}$ the process is quenched. Thus the peak at -8.82 ppm is assigned to the bridging hydride in (4) and that at 2.17 ppm to the terminal one (Scheme 2).

Initial attempts to crystallise (4) by standard methods (*i.e.* removal of volatiles and recrystallisation) were frustrated by the preferential isolation of crystalline (3) (as it is less soluble than (4)) with the mother liquor consisting of a mixture of (3) and (4) (*ca.* 20 : 80 by NMR), due to the partial reversibility of the reaction. However, the solid state molecular structure^{8,15} of (4) (Fig. 3) was eventually determined from single crystals grown by cooling slowly a freshly prepared solution of (4) at $-78\text{ }^\circ\text{C}$ under an overpressure (1.5 atm) of H₂, and confirms the spectroscopic assignment.

The Ti-C(carbene) bond length in (4) is $2.291(4)$ Å and is identical to that found in (3). On the other hand, the Ti-Ti bond is slightly shortened in (4) from $2.5610(8)$ Å in (3) to $2.5413(8)$ Å possibly due to the negligible steric requirements of the terminal hydride ligand. The Ti-H(bridging) bond distances (Ti1-H1 = $1.84(5)$ Å; Ti2-H1 = $1.79(5)$ Å) in (4) are similar within esd's and compare with the ones found in (3); the same applies to the Ti-H(terminal) (*i.e.* Ti2-H2 = $1.74(4)$ Å in Fig. 2) bond length. The Ti1-H1-Ti2 bond angle in (4) again approaches 90° ($89(2)^\circ$) and is very similar to that found in (3).

When (3) was treated with an excess of D₂ (5 eq.), the formation of (4-D) was observed, but deuterium was found to be only incorporated in the hydridic positions, and not in the new Me group derived from the previously “tucked-in” CH₂ group (confirmed by ^2H -NMR, DEPT-135 and gHSQC). Hence the reaction of (3) with H₂ to form (4)

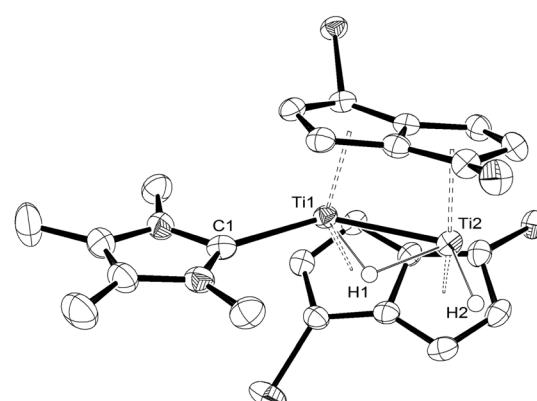


Fig. 3 ORTEP diagram of the molecular structure of (4) displaying 50% probability ellipsoids. ^1Pr groups omitted for clarity.



Table 1 Electronic binding energies (ΔE) and Gibbs energy changes (ΔG) for the reaction $\text{Ti}_2\text{Pn}_2 + \text{NHC} = \text{Ti}_2\text{Pn}_2\text{NHC}$

Ti pentalene dimer	NHC	ΔE (eV)	ΔG^a (kJ mol $^{-1}$)	Ti-Ti-C (°)
$\text{Ti}_2(\text{C}_8\text{H}_6)_2$	$\text{C}_3\text{H}_4\text{N}_2$	−1.31	−66	100
$\text{Ti}_2(\text{C}_8\text{H}_6)_2$	$\text{C}_3\text{Me}_4\text{N}_2$	−1.26	−65	114
$\text{Ti}_2\text{Pn}_2^{\dagger}$	$\text{C}_3\text{Me}_4\text{N}_2$	−0.48	+21	127

^a Calculations are on gas phase species which leads to an overestimation of the entropy change when there is a change in the number of species.

does not go *via* σ -bond metathesis (which would lead to D incorporation in the Me group in the reaction with D_2). Hence the formation of (4) (and also (3)) was probed computationally (ADF:BP/TZP: details are given in the ESI \ddagger). Preliminary studies suggested that sterics were important in determining the reaction energies. For example energy of binding an NHC to a Ti_2Pn_2 dimer depended critically on the substituents. Introduction of the methyl substituents on the NHC made very little difference to its binding energy but the bulky Si^iPr_3 substituents on the pentalene ligands increased the Ti-Ti-C angle forcing the NHC to a less favourable binding position thus decreasing the binding energy significantly (Table 1).

The degree to which the tuck-in reaction was favoured in the absence of an NHC was also investigated. The formation of $\text{Ti}_2\text{Pn}_2^{\dagger}(\text{Pn}_2^{\dagger}\text{-H})(\mu\text{-H})$ from $\text{Ti}_2\text{Pn}_2^{\dagger}$ was calculated to have $\Delta E = -0.03$ eV and $\Delta G = 11$ kJ mol $^{-1}$. However, the energies of the observed tuck-in reaction with the methylated NHC present were calculated to be $\Delta E = -0.69$ eV and $\Delta G = 9$ kJ mol $^{-1}$. Thus the presence of the base improves the energetics of the tuck-in reaction. The significant entropy disadvantage in the gas phase would be lessened in solution.

The HOMO of 3 (Fig. 4) shows a Ti-Ti σ -bond. The calculated Ti-Ti distance is 2.56 Å in excellent agreement with experiment. The Ti-H distances are 1.83 Å and the angle at the bridging hydrogen 89°. Such discrepancies from the experimental values are not unusual when comparing distances to bound hydrogen between theory and X-ray diffraction experiments.

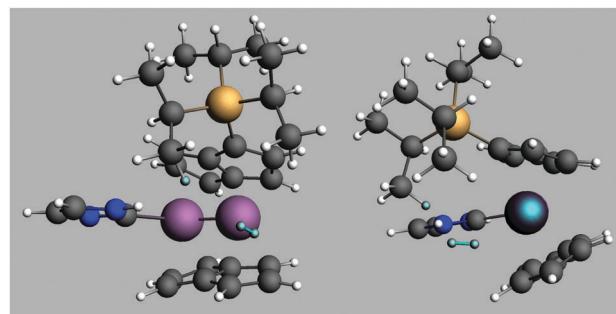


Fig. 5 Structure of the calculated transition state for H_2 addition. The reacting Hs are highlighted.

Addition of H_2 to 3 to form 4 is calculated to have reaction energies $\Delta E = -0.77$ eV and $\Delta G = -31$ kJ mol $^{-1}$. The calculated Ti-Ti distance for 4 is 2.54 Å reproducing the shortening from 3 found experimentally. The Ti-H(terminal) distance is 1.74 Å, the Ti-H(bridging) distances 1.81 and 1.82 Å and the Ti-H-Ti angle unchanged at 89°.

A transition state for this reaction was modelled using just one Si^iPr_3 substituent on one of the pentalene ligands and $\text{C}_3\text{H}_4\text{N}$ as the NHC for computational efficiency. The free energy of activation was estimated as 84 kJ mol $^{-1}$ for such a system. The transition state structure is shown in Fig. 5.

The Ti distance to the previously bridging H is 3.53 Å and the Ti distances to the reacting H_2 are 2.74 and 3.24 Å, the H-H distance being 0.76 Å. Such a geometry indicates that the tuck-in process is reversed before complete H_2 addition, consistent with the lack of deuterium incorporation into ^iPr groups and the conclusion that σ bond metathesis is not in play. It may be that the steric compression induced by the mere approach of the H_2 molecule is sufficient to reverse the tuck-in process, an idea given some credence by the pressure dependence of the reaction of (3) with H_2 (*vide infra*).

In conclusion, we have described the facile preparation of the first example of a *syn*-bimetallic Ti complex (3) featuring a bridging hydride, originating from the C-H activation of a ^iPr substituent induced by addition of the strong Lewis base 1,3,4,5-tetramethylimidazol-2-ylidene. Preliminary studies show that this transformation is also effected by other, effectively “planar” Lewis bases, *e.g.* dimethylaminopyridine (DMAP). The resultant C-H activation product (3) readily reacts with an excess of H_2 to produce very cleanly a unique *syn*-bimetallic di-hydride complex (4) featuring bridging and terminal hydride ligands. Labelling experiments and computational studies strongly suggest that the latter reaction does not proceed *via* a σ -bond metathesis mechanism.

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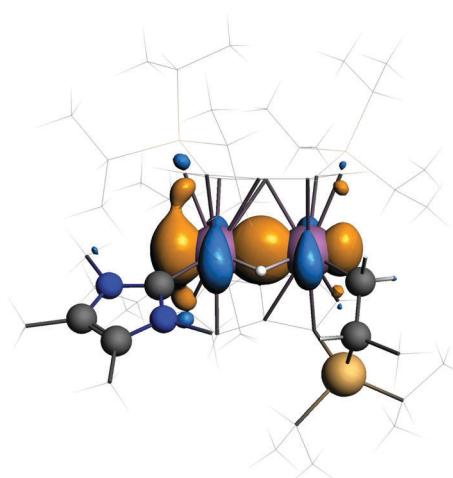


Fig. 4 Isosurface for the HOMO of 3.

Conflicts of interest

There are no conflicts to declare.



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