



Negative X-ray expansion in cadmium cyanide†

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Cadmium cyanide, Cd(CN)₂, is a flexible coordination polymer best studied for its strong and isotropic negative thermal expansion (NTE) effect. Here we show that this NTE is actually X-ray-exposure dependent: Cd(CN)₂ contracts not only on heating but also on irradiation by X-rays. This behaviour contrasts that observed in other beam-sensitive materials, for which X-ray exposure drives lattice expansion. We call this effect 'negative X-ray expansion' (NXE) and suggest its origin involves an interaction between X-rays and cyanide 'flips'; in particular, we rule out local heating as a possible mechanism. Irradiation also affects the nature of a low-temperature phase transition. Our analysis resolves discrepancies in NTE coefficients reported previously on the basis of X-ray diffraction measurements, and we establish the 'true' NTE behaviour of Cd(CN)₂ across the temperature range 150–750 K. The interplay between irradiation and mechanical response in Cd(CN)₂ highlights the potential for exploiting X-ray exposure in the design of functional materials.

New concepts

In this communication we have identified and quantified the phenomenon of negative X-ray expansion. To the best of our knowledge this represents the first systematic study of X-ray expansion dose effects in an inorganic functional material; we define the X-ray expansion coefficient σ to quantify the effect. Whilst the study X-ray radiation effects is well-established in the field of macromolecular crystallography (MX), it is only starting to be explored in functional inorganic materials and small molecules. In MX the unit cell parameter expands almost linearly with applied dose. We use precise measurement of the exposure times to extract the dose dependence of negative X-ray expansion in cadmium cyanide, we define the X-ray expansion coefficient σ and calculate σ for previously-reported data for proteins. The insights and methodologies presented here act as a starting point for future systematic studies of X-ray induced effects in functional materials and the use X-rays as a design tool in their own right, for example to access functional materials phases that are otherwise inaccessible *via* traditional (*e.g.* heat) treatment alone, or indeed to engineer functionality on the nanoscale *via* lithographic techniques.

1 Introduction

That X-rays can affect the structure, and therefore functionality, of materials is well established. In macromolecular crystallography, the phenomenology of 'radiation damage' is a mature and important field.^{1–3} Conversely, discussions about radiation damage in small molecule crystallography are rarer and only starting to be identified.^{4,5} X-ray-induced effects are somewhat

less well studied in conventional inorganic systems, despite being implicated in a number of interesting phenomena. Examples include decomposition,^{6–8} conductivity enhancement,^{9–13} colour changes,¹⁴ spin-crossover,¹⁵ charge transfer,^{16,17} cell-parameter changes,¹⁸ crystallisation,¹⁹ and amorphisation.^{20,21}

X-ray irradiation has been found to induce changes in the resistivity of magnetoresistive manganites (A_{1–x}B_x²⁺MnO₃).^{9–11} In Pr_{0.7}Ca_{0.3}MnO₃ this metal–insulator (and ferromagnetic to antiferromagnetic) transition occurs below 200 K and involves charge ordering of the Mn³⁺/Mn⁴⁺ cations. Upon application of a magnetic field this converts to a conducting state, which is preserved even upon removal of the field. Analogous effects are observed with X-rays; a decrease in intensity of the charge-ordering reflection occurs as the sample is irradiated. Beyond electronic transitions, amorphous thin films of BaTiO₃—which resist crystallisation on heating above 600 K—are seen to crystallise upon irradiation with high-energy (24 keV) synchrotron radiation.²² This is rationalised in terms of crystallisation arising

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from strain fields due to defects that are generated by X-ray-induced bond breaking. Other studies also document the coupling between X-ray exposure and elastic properties: X-ray irradiation of Barluenga's reagent ($\text{Br}(\text{Coll})_2\text{BF}_4$, Coll = 2,4,6-trimethylpyridine) suppresses the diffuse scattering, which is thought to emerge from short-range correlations between rotational configurations of the $\text{Br}(\text{Coll})_2^+$ and BF_4^- moieties.^{5,23} And X-ray exposure with Cu K α over a period of six hours led to crystallisation of sucrose from amorphous mixtures in 'Jello'.¹⁹

A relatively universal consequence of X-ray-matter interactions is volume expansion of the crystal lattice. In protein crystallography, where the effect has been studied particularly carefully, there is an almost linear dependence of the unit cell volume on the dose of X-ray irradiation.^{24,25} The X-ray dose is routinely monitored in protein crystallography experiments, and is measured in the S.I. unit of Grays (1 Gy = 1 J kg⁻¹). A typical synchrotron experiment deposits energy on the order of MGy, and the 'Henderson limit'—which dictates the approximate maximum dose a typical protein crystal can withstand and still diffract—is 20 MGy. In ref. 25 the X-ray induced expansion is quantified by $S = \frac{\Delta V}{\Delta \text{dose}}$ with values of around 0.02 Å³ Gy⁻¹, which corresponds to a volume change of ~0.3% per MGy of radiation.

The focus of this study is cadmium cyanide, $\text{Cd}(\text{CN})_2$, a cubic coordination polymer best studied for its large and isotropic negative thermal expansion (NTE)—the counterintuitive property in which there is a volume reduction with increasing temperature.^{26,27} It has an ostensibly extremely simple cubic structure with $Pn\bar{3}m$ symmetry, consisting of two interpenetrating cristobalite frameworks.²⁸ The structure is shown in Fig. 1(E). Crystallographically the cyanides exhibit head-to-tail disorder.

However, locally, their orientations are highly correlated: each Cd binds preferentially to two C and two N in an ice-like fashion, and cyanide reorientation enables sampling of degenerate ice-like states.^{29–31} Below 150 K there is a symmetry lowering phase transition that is of particular interest both in terms of the NTE—since the symmetry of the low temperature phase reflects the key phonons implicated in the NTE response²⁷—and also in terms of the correlated disorder: the phase transition is conspicuously absent for all *ab initio* studies which ignore the disorder of the cyanides.^{32,33} Recent neutron diffraction measurements of ¹¹⁴Cd(CN)₂ reported a low temperature phase with $I4_1/amd$ symmetry, whereas previous X-ray measurements of single crystals were inconclusive and the difficulty in determining a space group attributed to twinning.^{26,30,31}

There is surprising inconsistency in the literature regarding the magnitude of the thermal expansion coefficient, $\alpha_V = \partial \ln V / \partial T$, with values across the entire range $-35 > \alpha_V > -61.2 \text{ MK}^{-1}$ reported from X-ray diffraction measurements.^{26,30} Consequently, we sought to re-evaluate the nature of the low temperature phase transition using powder X-ray diffraction. It soon became clear that X-rays were unsuitable for unbiased determination of temperature-dependent structural behaviour, owing to X-ray-induced structural effects; these form the basis of this paper. In particular, we show that X-ray radiation leads to unit cell expansion, as well as phase transition reversal and phase selection in $\text{Cd}(\text{CN})_2$. We term the unusual property of contraction on X-ray irradiation 'negative X-ray expansion' (NXE)—by analogy to NTE. The opportunities for tuning material properties with X-rays and the implications for the study of materials using synchrotron X-ray diffraction are discussed.



Fig. 1 Negative X-ray expansion: the evolution of a series of diffraction patterns ($\lambda = 0.82507 \text{ \AA}$) for a sample of $\text{Cd}(\text{CN})_2$, held at (A and B) 200 K and (C and D) 100 K, as a function of X-ray exposure (blue to red). (A) At 200 K $\text{Cd}(\text{CN})_2$ adopts cubic $Pn\bar{3}m$ symmetry and increasing irradiation results in contraction of the unit cell, as seen in (B) the (110) peak moves to progressively larger values of 2θ . (D) With minimal X-ray exposure at 100 K (blue) $\text{Cd}(\text{CN})_2$ has undergone a symmetry-lowering phase transition. On increasing exposure these additional peaks coalesce and those of cubic $\text{Cd}(\text{CN})_2$ are restored. X-ray exposure reverses the phase transition. (E) $\text{Cd}(\text{CN})_2$ crystallises with $Pn\bar{3}m$ symmetry where tetrahedrally coordinated Cd centres (red/pink to highlight interpenetration frameworks) are connected by disordered cyanide ions (C/N grey spheres).



2 Methods

Polycrystalline samples of $\text{Cd}(\text{CN})_2$ were synthesised by slow evaporation as reported previously.^{26,30} Large single crystals were extracted from the solution manually, dried and ground to a polycrystalline powder using a mortar and pestle. Synchrotron powder diffraction data were collected at the high-resolution powder diffraction beamline, I11, at Diamond Light Source. Individual data sets were collected using a Mythen2 Position Sensitive Detector (PSD) with two 5-second scans separated by an angular shift in detector position of 2.5° . The wavelength and intrinsic peak-shape parameters were refined against a known Si 640c NIST standard. Variable temperature (100–500 K) measurements were performed with sample heating from an Oxford Cryostreams attachment. Higher temperature measurements (500–750 K) made use of a Cyberstar hot-air blower attachment. For variable exposure experiments, the exposure time was estimated using a custom-written I11 script, which extracted the opening and closing times of the experimental shutter. In this way the radiation doses applied were carefully controlled by controlling the exposure time of the sample. The width of the primary beam is 2.5 mm; thus a horizontal translation of the capillary 2.5 mm enables pristine $\text{Cd}(\text{CN})_2$ to be irradiated for measurements within the same capillary. Pawley refinements to determine the unit cell parameters were carried out in Topas Academic version 4.1.³⁴

3 Results

3.1 Negative X-ray expansion

We start by showing the effects of X-ray exposure on the structure of ambient temperature cadmium cyanide. Fig. 1(A) shows a waterfall plot of sequential diffraction patterns of a sample of $\text{Cd}(\text{CN})_2$, collected at 200 K. The data are characterised by few, well separated, reflections. This is consistent with the expected cubic symmetry, and the data can be refined using the space-group $Pn\bar{3}m$ with lattice parameter $a \sim 6.3 \text{ \AA}$.^{26,28,35} A magnified view of the (110) reflection at $2\theta \sim 10.5^\circ$ is shown in the inset. The cubic phase is maintained with increasing X-ray exposure.

It is evident that increasing X-ray irradiation is causing a contraction of the unit cell; the diffraction peaks progressively shift towards higher values of 2θ , as seen most clearly for the (110) reflection on the inset. We will come to show that this is not a local heating effect. Whereas in macromolecular crystallography the expectation is that beam damage results in unit cell expansion,²⁴ we see that the opposite is true for $\text{Cd}(\text{CN})_2$. This is consistent with a previous observation for isostructural $\text{Zn}(\text{CN})_2$ where X-ray irradiation at high pressure resulted in unit cell contraction and eventual amorphisation.²⁰ We term this unusual effect of shrinking with increasing irradiation negative X-ray expansion, to highlight the similarities between the unusual temperature-dependent behaviour of these materials and their response to irradiation. The negative thermal expansion in cadmium cyanide is driven by low energy phonons with negative Grüneisen parameters, and this therefore suggests a degree of X-ray–phonon coupling in this material. Similar negative

X-ray expansion was observed in the Prussian blue analogue $\text{Rb}_{0.7}\text{Mn}_{1.15}[\text{Fe}^{\text{III}}\text{CN}_6]\cdot 2\text{H}_2\text{O}$ owing to X-ray-induced electronic, and spin transitions as well as in the zeolitic-imidazolate framework (ZIF) materials ZIF-4, ZIF-62 and ZIF-zni, prior to their amorphisation under irradiation.^{18,21}

Pristine cadmium cyanide displays both negative X-ray expansion as well as phase transition suppression at 100 K. A pristine sample of cadmium cyanide was cooled to 100 K, below the phase transition temperature of $\sim 135 \text{ K}$,^{26,30,31} and held at 100 K for two minutes to allow for thermal equilibration. The diffraction pattern of pristine cadmium cyanide at 100 K is shown in blue in Fig. 1(C and D). There is clear evidence of a symmetry-lowering phase transition with both peak-splitting of the cubic reflections (see $2\theta \sim 10.5^\circ$ and additional peak intensity appearing in between the cubic Bragg peaks for $2\theta \sim 5^\circ$). We were not able to solve the structure of $\text{Cd}(\text{CN})_2$ present in this diffraction pattern, which does not correspond to the low temperature $I4_1/amd$ phase identified using neutron diffraction in ref. 31. The ice-like correlated disorder in this material is likely to lead to a large configurational degeneracy, and this suggests that there may be multiple competing low temperature phases.³⁶ Likewise the long timescale of cyanide reorientations—implicated in the $Pn\bar{3}m$ – $I4_1/amd$ transition reported in ref. 31—raises the possibility we are observing kinetic products in this experiment. Further exploration of this low temperature phase from an experimental and computational point of view will form the basis of a future report. We note simply that the disorder introduces additional complexity and that the phase behaviour is likely to be extremely dependent on synthesis—and experimental—conditions, which were indeed different for the $\text{Cd}(\text{CN})_2$ samples prepared in ref. 31 (ammoniacal synthesis from $\text{Hg}(\text{CN})_2$ and Cd metal) and those synthesised here (aqueous).

3.2 Phase transition

Variation in the diffraction pattern for a given sample was then followed as a function of X-ray exposure, as shown in Fig. 1(C). The sample was cooled to 100 K and irradiated continuously over a period of ~ 200 seconds. After about one hundred seconds of X-ray exposure, there is a clear ascent in symmetry between the initial pattern collected with zero X-ray exposure (blue) and that collected after 200 s (red), which is consistent with the simple cubic diffraction pattern of the $Pn\bar{3}m$ cell [cf. Fig. 1(A)]. What is remarkable here is that X-ray irradiation is able to reverse the phase transition that is observed on cooling. Typically ‘beam damage’ as a result of X-ray irradiation leads to loss of diffraction intensity.^{1,37,38} In this case the intensity appears not to decrease but to accumulate in fewer reflections until they correspond to those of the $Pn\bar{3}m$ phase after around 100 seconds of X-ray irradiation. Once the phase transition back to the cubic phase is complete, continuous irradiation again leads to negative X-ray expansion.

3.3 Is the response to X-rays a localised heating effect?

Having established that increasing X-ray exposure leads to unit cell contraction in $\text{Cd}(\text{CN})_2$, in an effect analogous to increasing



temperature, it is perhaps natural to question whether it is, in fact, a thermal effect. The variation in the lattice parameter with X-ray exposure was determined for temperatures between 160 and 110 K, for a sample containing a mixture of $\text{Cd}(\text{CN})_2$ and $\text{KMn}[\text{Ag}(\text{CN})_2]_3$ as a temperature standard. $\text{KMn}[\text{Ag}(\text{CN})_2]_3$ is a coordination polymer with extreme anisotropic thermal expansion by virtue of its wine-rack-like topology. The thermal expansion coefficients ($\alpha_a = +61(2) \text{ MK}^{-1}$ and $\alpha_c = -60(3) \text{ MK}^{-1}$)³⁹ are three times larger in magnitude than the linear thermal expansion in $\text{Cd}(\text{CN})_2$ ($\alpha_a = -20.4 \text{ MK}^{-1}$)²⁶ so it is expected that any thermal effects would be visible in the variation of lattice parameters. The sample was held at 110 K and the X-ray exposure—corresponding to the number of seconds for which the shutter was open—recorded for each diffraction pattern. At each temperature an initial ~ 50 s of radiation was applied to enable comparison of the $Pn\bar{3}m$ phase in each case. This protocol was then repeated at temperatures in 10 K intervals up to 160 K. At each temperature the capillary was translated such that a pristine sample of $\text{Cd}(\text{CN})_2$ was irradiated. The lattice parameter as a function of exposure was refined *via* Pawley analysis of the diffraction data [see ESI†].

The relative change in lattice parameters on increasing X-ray exposure for both compounds at each temperature is shown in Fig. 2. One would expect the change in unit cell parameters to be clearly visible for $\text{KMn}[\text{Ag}(\text{CN})_2]_3$ if the increased X-ray exposure were generating an increase in temperature. At all temperatures, increasing X-ray exposure has a negligible effect on the lattice parameters of $\text{KMn}[\text{Ag}(\text{CN})_2]_3$ compared to a change in length of up to 0.3% for $\text{Cd}(\text{CN})_2$; from this observation two conclusions can be drawn. Firstly, $\text{KMn}[\text{Ag}(\text{CN})_2]_3$ is structurally unaffected by X-ray exposure. And secondly, the change in unit cell metrics on X-ray exposure in $\text{Cd}(\text{CN})_2$ cannot be attributed to a heating effect caused by X-ray absorption, consistent with existing studies.^{40,41} The

change in lattice parameters is nonetheless significant: irradiation with X-rays for a period of ~ 200 seconds is equivalent to a temperature increase of around 150 K.

It is possible to quantify the dependence of the unit cell parameter on radiation dose, in direct analogy to the thermal expansion coefficient. In ref. 25 the authors use $S = \frac{\Delta V}{\Delta \text{dose}}$. This parameter, however, measures the absolute volume change for a given dose, which does not facilitate comparison between systems. We suggest instead the use of the volume X-ray expansion coefficient σ_V where;

$$\sigma_V = \frac{1}{V} \frac{\Delta V}{\Delta \text{dose}} \quad (1)$$

Typical values for proteins (where $S \sim +0.02 \text{ \AA}^3 \text{ Gy}^{-1}$) correspond to $\sigma_V \sim 3.3 \text{ GGy}^{-1}$, equivalent to a 0.3% volume change per Gy of radiation.²⁵ In order to determine σ_V for $\text{Cd}(\text{CN})_2$ we start by estimating the dose rate in Gy s^{-1} .

The dose rate of radiation experienced by the sample can be estimated using RADDOS-3D, a programme first developed for single-crystal macromolecular samples, and more recently adapted to quantify dose rates for small molecule crystallography (still single crystal) and for small-angle scattering (SAXS) measurements.^{42–44} The dose rate was calculated to be $13\,407 \text{ Gy s}^{-1}$, taking into account the attenuation by the borosilicate capillary ($10 \mu\text{m}$), the beam type (top-hat), and the beam dimensions ($0.8 \times 2.5 \text{ mm}^2$). σ_V for $\text{Cd}(\text{CN})_2$ at 110 K after 225 seconds of irradiation is therefore $\sigma_V \sim -2.98 \text{ GGy}^{-1}$, which is of the same order of magnitude as that of the protein in ref. 25 only negative; further details of the derivation are given as ESI.† Thus $\text{Cd}(\text{CN})_2$ shows a comparable magnitude of response to X-rays as a large, flexible protein only in contraction rather than expansion. Moreover, $\sigma_V \sim -2.98 \text{ GGy}^{-1}$ is likely to severely underestimate the true X-ray expansion coefficient, since this assumes 100% packing density of the sample in the capillary; a more appropriate assumption would be $\sim 50\%$ packing density, which would lower the absorption of X-rays by the sample and therefore the average dose.

3.4 Temperature-dependent X-ray effects

It was also observed that the specific temperature at which the sample is irradiated—and the order in which these stimuli are applied—can also affect the structural properties of cadmium cyanide. A distinct type of phase selection occurs if the sample is irradiated *prior* to cooling. $\text{Cd}(\text{CN})_2$ was measured at 300 K, then cooled to 100 K and left to equilibrate for two minutes before being translated by 2.5 mm in order to measure pristine $\text{Cd}(\text{CN})_2$. A diffraction pattern was measured at this position before translation back to the original position where the ambient data collection had taken place. The resulting diffraction patterns for one sample are shown in Fig. 3, in black, blue and red respectively. Remarkably the two datasets at 100 K are quite different. Irradiation prior to cooling appears to suppress certain low-temperature reflections, whilst having a negligible effect on others (see $2\theta \sim 5^\circ$ vs. $2\theta \sim 8^\circ$, for example). Clearly there is a degree of symmetry lowering in both cases: the



Fig. 2 The relative changes in the lattice parameters (l , where l_0 is the relevant lattice parameter with minimal X-ray exposure at 110 K) of $\text{Cd}(\text{CN})_2$ and $\text{KMn}[\text{Ag}(\text{CN})_2]_3$ as a function of X-ray exposure from blue to red (~ 200 seconds) with $\lambda = 0.82507 \text{ \AA}$. Whilst the lattice parameter of the former is seen to shrink as a function as a function of exposure at all temperatures, those of the latter show no change, as seen by the overlapping symbols, eliminating the possibility of local heating as a mechanism for the negative X-ray expansion in $\text{Cd}(\text{CN})_2$.



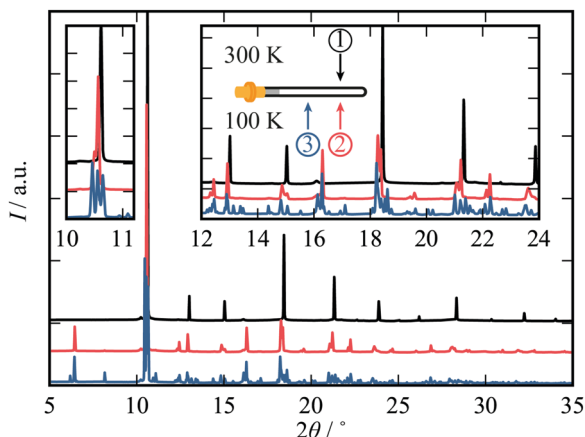


Fig. 3 Diffraction patterns ($\lambda = 0.82484 \text{ \AA}$) collected for a sample of $\text{Cd}(\text{CN})_2$ at 300 K (black) and in the same position at 100 K (red) and at 100 K having been translated (blue = minimal prior exposure). The nature of the phase transition is clearly affected by irradiation prior to cooling.

phase transition is not suppressed completely. This experiment was repeated for three different samples [see ESI†] and in all three cases the same reflections are suppressed by irradiation with X-rays at room temperature. Moreover, these changes are robust with respect to annealing at higher temperature. Interestingly, the diffraction pattern that is observed with prior irradiation at room temperature can be indexed and refined as a single phase with $I4_1/amd$ symmetry, as identified in ref. 31.

So what are the possible mechanisms by which X-rays might interact with $\text{Cd}(\text{CN})_2$? It is known from X-ray studies in macromolecular systems that high energy incoming (primary) electrons lead to the ejection of secondary electrons, which are now thermalised according to the Boltzmann distribution; this can help explain the parallels here with temperature-dependent behaviour.^{1,45} In the case of $\text{Cd}(\text{CN})_2$, it was shown in ref. 31 that NTE arises in part due to a cyanide ordering mechanism driven by cyanide reorientation (as well as more conventional NTE low energy phonons like transverse vibrations of Cd–CN–Cd moieties and framework breathing modes). The cyanide reorientation itself is driven by spin-ice physics: in which the Cd centre preferentially binds to two C and two N of the cyanide ions, mimicking the two-in-two-out ices rules that govern proton or spin configurations in water and spin ice respectively. The key point here is that the spin-ice physics results in a manifold of degenerate correlated disordered states. The configurational landscape is navigated *via* migration of local violations of the ice rules—these defects are the effective magnetic monopoles of spin-ices. The energy barrier associated with such a violation in $\text{Cd}(\text{CN})_2$ —a cyanide flip—was calculated here to be $\Delta \sim 8800 \pm 600 \text{ K} \sim 1 \text{ eV}$. We suggest that X-ray induced cyanide flipping might enable sampling of the spin-ice manifold that would be prohibitively slow for thermally-induced excitations alone.

We further speculate that it is not coincidental that $\text{Cd}(\text{CN})_2$ displays both correlated disorder and X-ray susceptibility, in contrast here to the well-ordered $\text{KMn}[\text{Ag}(\text{CN})_2]_3$ which is structurally unaffected. We draw a parallel between the shallow

energy surface between ice-like states in $\text{Cd}(\text{CN})_2$ and those of proteins. Of other reported functional materials with unusual responses to X-rays many show either correlated disorder or some magnetic or electronic instability.^{9,11,12,15,18,22,46} X-ray irradiation melts the charge order in the colossal magnetoresistance material $\text{Pr}_{0.7}\text{Ca}_{0.3}\text{MnO}_3$ and induces an insulator–metal transition.⁹ A similar reduction in resistivity is observed in the spinel CuIr_2S_4 where X-ray irradiation disorders the lattice dimerisation pattern, preserving the orientation of the dimers whilst destroying the translational long-range order.¹² These are but two examples, but the prospect of post-synthetic control of correlated disorder states and the desired properties deserves further systematic attention, particularly the idea that in this way we may access states that are inaccessible by temperature(/pressure) alone.

3.5 The true thermal expansion coefficient for cadmium cyanide

That X-ray irradiation should lead to contraction of the unit cell may help explain previous experimental discrepancies in the reported values of the thermal expansion coefficient by a factor of two ($\alpha_V = -61.2(12) \text{ MK}^{-1}$ in ref. 26 or $\alpha_V = -36.3(19) \text{ MK}^{-1}$ in ref. 30); the observed effects will have included contributions from both NTE and NXE at once. Numerous factors will affect the measured value of α_V , including the rate and spacing of measurements as well as the direction of measurements with temperature; if measured on heating, for example, one might expect α_V to be larger as the compound effects of NXE and NTE lead to a larger-than-expected contraction of the unit cell. Therefore we resolved to determine a NXE-free thermal expansion coefficient for $\text{Cd}(\text{CN})_2$, also extending our measurements over a broader temperature range than that used in previous studies. We measured X-ray diffraction patterns on heating between 150 K and 750 K, with the sample translated for each temperature point to irradiate a pristine $\text{Cd}(\text{CN})_2$ sample. The sample showed no signs of decomposition; $\text{Cd}(\text{CN})_2$ is stable up to at least 750 K. The lattice parameters were obtained *via* Pawley refinements and α_V obtained using linear strain calculations as implemented in PASCAL [see ESI† for fits to data].⁴⁷ These data are plotted as crosses in Fig. 4(A) and the resulting thermal expansion coefficient for minimal X-ray exposure determined to be $\alpha_V = -55(2) \text{ MK}^{-1}$. Also plotted in Fig. 4(A) are data from ref. 26, 30 and 31 for comparison; note the variation in slopes, which may be attributable to NXE contributions to the apparent thermal expansion behaviour.

We turn now to discuss the newly revised value of α_V and the extended temperature range over which $\text{Cd}(\text{CN})_2$ displays NTE. Cadmium cyanide is most studied in the context of negative thermal expansion materials. Typically NTE materials can be divided into two main classes depending on the mechanism by which NTE is enabled.⁴⁸ The first class, to which $\text{Cd}(\text{CN})_2$ belongs, consists of materials in which NTE is driven by low-energy, volume-reducing phonons and is characterised by modest thermal expansion coefficients ($\alpha_V \sim -20 \text{ MK}^{-1}$) over extended temperature ranges sometimes spanning hundreds of Kelvin. The second class is one in which NTE is driven by an electronic or magnetic instability; such materials have colossal



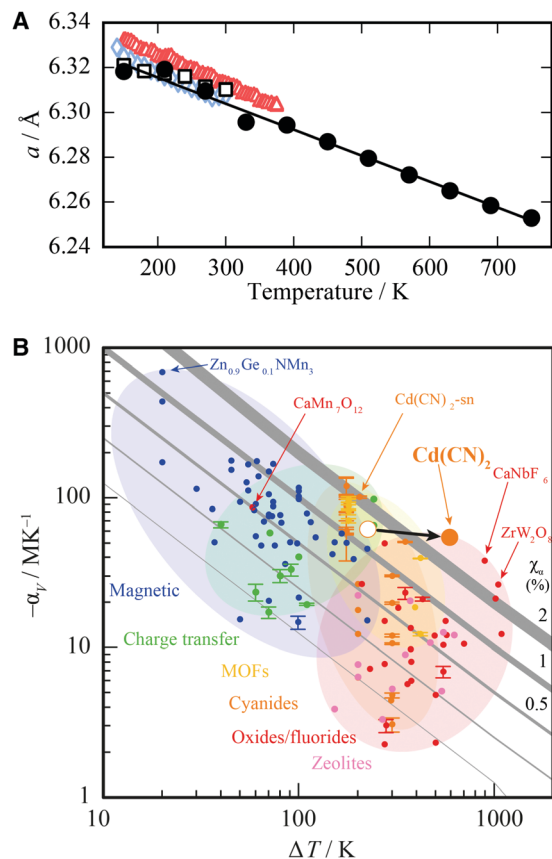


Fig. 4 (A) The cubic lattice parameter, a , of cadmium cyanide collected on heating between 150 K and 750 K for minimal X-ray irradiation, shown as black circles (errors are smaller than the circles). Previously reported data are included for comparison to illustrate the concomitant effects of thermal expansion and X-ray expansion, shown as red triangles (ref. 26), black squares (ref. 30) and blue diamonds (ref. 31). The NXE-minimised thermal expansion coefficient is $\alpha_V = -55(2) \text{ MK}^{-1}$. (B) The Ashby plot of ref. 48 comparing the magnitudes and temperature ranges of NTE for different families of materials, where the NTE capacity $\chi_x = \Delta V/V$ is maximised in the top right corner. The new value of α_V reported here places $\text{Cd}(\text{CN})_2$ alongside CaNbF_6 as the material with the largest NTE capacity.

values of α_V up to values above -400 MK^{-1} , this time over a more limited temperature range $O(\sim 10 \text{ K})$. Cadmium cyanide is remarkable both for the magnitude of its isotropic NTE, but also for the temperature range over which this NTE is maintained. In ref. 48 we collated NTE data for a large number of isotropic NTE materials in order to compare the different systems according to their NTE capacity, χ_x , where $\chi_x = \Delta V/V = -\alpha_V \Delta T$, in an effort to help the materials scientist categorise and select a particular type of NTE material for a given application. These data are reproduced in Fig. 4B with the position of $\text{Cd}(\text{CN})_2$ updated with the results reported here ($\chi_x = 3.3\%$). It now joins the small group of materials, including CaNbF_6 and the canonical ZrW_2O_8 with capacities greater than 2%.^{49–52}

4 Conclusion

To conclude, X-ray radiation has a profound impact on the structure and dynamics of $\text{Cd}(\text{CN})_2$, including negative X-ray

expansion, phase selectivity and hysteresis effects, depending on the thermal and radiation history. These observations are able to explain previous disagreements in the thermal expansion coefficient, α . We determine the true NXE-free $\alpha_V = -55(2) \text{ MK}^{-1}$, and demonstrate that $\text{Cd}(\text{CN})_2$ displays NTE over a large temperature range of 600 K between 150 K and 750 K. The true range may well be higher; our measurements did not extend beyond 750 K, but $\text{Cd}(\text{CN})_2$ had not decomposed by that temperature.

The observations made here for $\text{Cd}(\text{CN})_2$ form part of a broader collection of X-ray-induced effects in functional materials. Whilst ‘radiation damage’ appears to be universal in macromolecular systems, this is not the case for inorganic materials. It could be argued that the observation of X-ray-induced functionality in inorganic systems is often serendipitous, with important ramifications: for some materials the structural behaviour, and therefore interpretations, are influenced by X-rays but the effects attributed to other stimuli. So in the first instance we would like to understand how to minimise X-ray induced effects such that the intrinsic structural behaviour can be studied.

There remain many open questions: the precise mechanism by which X-rays interact with the structure of cadmium cyanide is unclear. Indeed, why are some materials susceptible to X-ray exposure and others not? In this study for example, $\text{KMn}[\text{Ag}(\text{CN})_2]_3$ was unaffected by radiation under the same conditions as $\text{Cd}(\text{CN})_2$. Both are flexible, cyanide-containing materials, but $\text{Cd}(\text{CN})_2$ is extremely sensitive to radiation and $\text{KMn}[\text{Ag}(\text{CN})_2]_3$ is not. Simultaneous spectroscopy and X-ray studies will prove valuable in this regard: UV-Vis measurements to ascertain whether radical formation plays a role; *in situ* Raman or IR spectroscopy to probe changes in bonding with irradiation; EXAFS measurements to understand the effect on local coordination, amongst other techniques. Persistent X-ray-induced effects might also be analysed offline with the above methods, as well as further (^{113}Cd) NMR or neutron scattering measurements. Moreover there is scope for systematic studies—reminiscent of those conducted in macromolecular crystallography, albeit on single crystals as opposed to polycrystalline samples—on the effect of dose, dose rate, X-ray energy, flux, beam-size and temperature. As our understanding of material-X-ray interactions improves, we can start to use X-ray radiation as a design tool in its own right (as opposed to simply a diagnostic one); recent experiments in the X-ray lithography of Metal–Organic framework (MOF) thin films are an exciting new application in that space.⁵³ The systematic studies employed here to identify the negative X-ray expansion in cadmium cyanide are an important starting point.

Conflicts of interest

There are no conflicts to declare.

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References

- 1 E. F. Garman, *Acta Crystallogr., Sect. D: Biol. Crystallogr.*, 2010, **66**, 339–351.
- 2 C. C. F. Blake and D. C. Phillips, *Proceedings of the Symposium on the Biological Effects of Ionizing Radiation at the Molecular Level*, Int. Atomic Energy Agency, Vienna, 1962, pp. 183–191.
- 3 J. R. Helliwell, *J. Cryst. Growth*, 1988, **90**, 259–272.
- 4 J. Christensen, P. N. Horton, C. S. Bury, J. L. Dickerson, H. Taberman, E. F. Garman and S. J. Coles, *IUCrJ*, 2019, **6**, 703–713.
- 5 L. C. F. Morgan, Y. Kim, J. N. Blandy, C. A. Murray, K. E. Christensen and A. L. Thompson, *Chem. Commun.*, 2018, **54**, 9849–9852.
- 6 M. Pravica, L. Bai, D. Sneed and C. Park, *J. Phys. Chem. A*, 2013, **117**, 2302–2306.
- 7 D. Goldberger, E. Evlyukhin, P. Cifligu, Y. Wang and M. Pravica, *J. Phys. Chem. A*, 2017, **121**, 7108–7113.
- 8 E. Evlyukhin, E. Kim, D. Goldberger, P. Cifligu, S. Schyck, P. F. Weck and M. Pravica, *Phys. Chem. Chem. Phys.*, 2018, **20**, 18949–18956.
- 9 V. Kiryukhin, D. Casa, J. P. Hill, B. Keimer, A. Vigliante, Y. Tomioka and Y. Tokura, *Nature*, 1997, **386**, 813–815.
- 10 D. E. Cox, P. G. Radaelli, M. Marezio and S.-W. Cheong, *Phys. Rev. B*, 1998, **57**, 3305–3314.
- 11 V. Kiryukhin, Y. J. Wang, F. C. Chou, M. A. Kastner and R. J. Birgeneau, *Phys. Rev. B*, 1999, **59**, R6581–R6584.
- 12 H. Ishibashi, T. Y. Koo, Y. S. Hor, A. Borissov, P. G. Radaelli, Y. Horibe, S.-W. Cheong and V. Kiryukhin, *Phys. Rev. B*, 2002, **66**, 144424.
- 13 M. Pravica, D. Popov, S. Sinogeikin, D. Sneed, G. Guardala and Q. Smith, *Appl. Phys. Lett.*, 2013, **103**, 224103.
- 14 C. Chen, J.-K. Sun, Y.-J. Zhang, X.-D. Yang and J. Zhang, *Angew. Chem.*, 2017, **129**, 14650–14654.
- 15 D. Papanikolaou, S. Margadonna, W. Kosaka, S.-I. Ohkoshi, M. Brunelli and K. Prassides, *J. Am. Chem. Soc.*, 2006, **128**, 8358–8363.
- 16 K. Ishiji, M. Deguchi, N. Nakajima, T. Matsuda, H. Tokoro, S. Ohkoshi and T. Iwazumi, *J. Phys.: Conf. Ser.*, 2009, **148**, 012046.
- 17 Y. Sekine, M. Nihei, R. Kumai, H. Nakao, Y. Murakami and H. Oshio, *Chem. Commun.*, 2014, **50**, 4050–4052.
- 18 S. Margadonna, K. Prassides and A. N. Fitch, *J. Am. Chem. Soc.*, 2004, **126**, 15390–15391.
- 19 G. Susich, A. O. King and L. M. Dogliotti, *Science*, 1959, **130**, 567–568.
- 20 S. H. Lapidus, G. J. Halder, P. J. Chupas and K. W. Chapman, *J. Am. Chem. Soc.*, 2013, **135**, 7621–7628.
- 21 R. N. Widmer, G. I. Lampronti, N. Casati, S. Farsang, T. D. Bennett and S. A. T. Redfern, *Phys. Chem. Chem. Phys.*, 2019, **21**, 12389–12395.
- 22 Y. Feldman, V. Lyahovitskaya, G. Leitus, I. Lubomirsky, E. Wachtel, V. A. Bushuev, G. Vaughan, Z. Barkay and Y. Rosenberg, *Appl. Phys. Lett.*, 2009, **95**, 051919.
- 23 Y. Kim, E. J. Mckinley, K. E. Christensen, N. H. Rees and A. L. Thompson, *Cryst. Growth Des.*, 2014, **14**, 6294–6301.
- 24 J. Murray and E. Garman, *J. Synchrotron Radiat.*, 2002, **9**, 347–354.
- 25 R. B. G. Ravelli, P. Theveneau, S. McSweeney and M. Caffrey, *J. Synchrotron Radiat.*, 2002, **9**, 355–360.
- 26 A. L. Goodwin and C. J. Kepert, *Phys. Rev. B*, 2005, **71**, 140301.
- 27 M. T. Dove and H. Fang, *Rep. Prog. Phys.*, 2016, **79**, 066503.
- 28 B. F. Hoskins and R. Robson, *J. Am. Chem. Soc.*, 1990, **112**, 1546–1554.
- 29 S. i. Nishikiori, C. I. Ratcliffe and J. A. Ripmeester, *Can. J. Chem.*, 1990, **68**, 2270–2273.
- 30 V. E. Fairbank, A. L. Thompson, R. I. Cooper and A. L. Goodwin, *Phys. Rev. B*, 2012, **86**, 104113.
- 31 C. S. Coates, M. Baise, A. Simonov, J. W. Makepeace, A. G. Seel, R. I. Smith, H. Y. Playford, D. A. Keen, R. Siegel, A. Schmutzler, J. Senker, B. Slater and A. L. Goodwin, 2019, arXiv:1904.05749.
- 32 J. W. Zwanziger, *Phys. Rev. B*, 2007, **76**, 052102.
- 33 P. Ding, E. J. Liang, Y. Jia and Z. Y. Du, *J. Phys.: Condens. Matter*, 2008, **20**, 275224.
- 34 A. A. Coelho, *TOPAS-Academic, version 4.1 (Computer Software)*, 2007.
- 35 E. Shugam and H. Zhdanov, *Acta Physicochim, URSS*, 1945, vol. 20, pp. 247–252.
- 36 J.-L. Kuo, J. V. Coe, S. J. Singer, Y. B. Band and L. Ojamäe, *J. Chem. Phys.*, 2001, **114**, 2527–2540.
- 37 J. M. Holton, *J. Synchrotron Radiat.*, 2009, **16**, 133–142.
- 38 O. Carugo and K. D. Carugo, *Trends Biochem. Sci.*, 2005, **30**, 213–219.
- 39 A. B. Cairns, A. L. Thompson, M. G. Tucker, J. Haines and A. L. Goodwin, *J. Am. Chem. Soc.*, 2012, **134**, 4454–4456.
- 40 M. Witala, J. Han, A. Menzel and K. Nygård, *J. Appl. Crystallogr.*, 2014, **47**, 2078–2080.
- 41 E. H. Snell, H. D. Bellamy, G. Rosenbaum and M. J. van der Woerd, *J. Synchrotron Radiat.*, 2007, **14**, 109–115.
- 42 O. B. Zeldin, M. Gerstel and E. F. Garman, *J. Appl. Crystallogr.*, 2013, **46**, 1225–1230.
- 43 C. S. Bury, J. C. Brooks-Bartlett, S. P. Walsh and E. F. Garman, *Protein Sci.*, 2018, **27**, 217–228.
- 44 J. C. Brooks-Bartlett, R. A. Batters, C. S. Bury, E. D. Lowe, H. M. Ginn, A. Round and E. F. Garman, *J. Synchrotron Radiat.*, 2017, **24**, 63–72.



- 45 M. Weik, R. B. Ravelli, G. Kryger, S. McSweeney, M. L. Raves, M. Harel, P. Gros, I. Silman, J. Kroon and J. L. Sussman, *Proc. Natl. Acad. Sci. U. S. A.*, 2000, **97**, 623–628.
- 46 H. L. B. Boström, A. B. Cairns, L. Liu, P. Lazor and I. E. Collings, *Dalton Trans.*, 2020, **49**, 12940–12944.
- 47 M. J. Cliffe and A. L. Goodwin, *J. Appl. Crystallogr.*, 2012, **45**, 1321–1329.
- 48 C. S. Coates and A. L. Goodwin, *Mater. Horiz.*, 2019, **6**, 211–218.
- 49 J. C. Hancock, K. W. Chapman, G. J. Halder, C. R. Morelock, B. S. Kaplan, L. C. Gallington, A. Bongiorno, C. Han, S. Zhou and A. P. Wilkinson, *Chem. Mater.*, 2015, **27**, 3912–3918.
- 50 B. R. Hester, J. C. Hancock, S. H. Lapidus and A. P. Wilkinson, *Chem. Mater.*, 2017, **29**, 823–831.
- 51 J. S. O. Evans, T. A. Mary, T. Vogt, M. A. Subramanian and A. W. Sleight, *Chem. Mater.*, 1996, **8**, 2809–2823.
- 52 T. A. Mary, J. S. O. Evans, T. Vogt and A. W. Sleight, *Science*, 1996, **272**, 90–92.
- 53 M. Tu, B. Xia, D. E. Kravchenko, M. L. Tietze, A. J. Cruz, I. Stassen, T. Hauffman, J. Teyssandier, S. De Feyter, Z. Wang, R. A. Fischer, B. Marmiroli, H. Amenitsch, A. Torvisco, M. D. J. Velásquez-Hernández, P. Falcaro and R. Ameloot, *Nat. Mater.*, 2021, **20**, 93–99.

