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Combined experimental and computational study of the photoabsorption of the monodoped and nondoped nanoclusters $\text{Au}_{24}\text{Pt}(\text{SR})_{18}$, $\text{Ag}_{24}\text{Pt}(\text{SR})_{18}$, and $\text{Ag}_{25}(\text{SR})_{18}^{\dagger}$

Pierpaolo D'Antoni,^a Luca Sementa,^b Sara Bonacchi,^c Mattia Reato,^c Flavio Maran,^b *^{cd} Alessandro Fortunelli^b *^b and Mauro Stener^b *^a

Assessing the accuracy of first-principles computational approaches is instrumental to predict electronic excitations in metal nanoclusters with quantitative confidence. Here we describe a validation study on the optical response of a set of monolayer-protected clusters (MPC). The photoabsorption spectra of $\text{Ag}_{25}(\text{DMBT})_{18}^{-}$, $\text{Ag}_{24}\text{Pt}(\text{DMBT})_{18}^{2-}$ and $\text{Au}_{24}\text{Pt}(\text{SC}_4\text{H}_9)_{18}$, where DMBT is 2,4-dimethylbenzenethiolate and SC_4H_9 is *n*-butylthiolate, have been obtained at low temperature and compared with accurate TDDFT calculations. An excellent match between theory and experiment, with typical deviations of less than 0.1 eV, was obtained, thereby validating the accuracy and reliability of the proposed computational framework. Moreover, an analysis of the TDDFT simulations allowed us to ascribe all relevant spectral features to specific transitions between occupied/virtual orbital pairs. The doping effect of Pt on the optical response of these ultrasmall MPC systems was identified and discussed.

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1. Introduction

Metal nanoclusters protected by a layer of coating ligands with well-defined stoichiometry and chemical structure, often referred to as monolayer-protected clusters (MPC), are compounds that exhibit peculiar physicochemical properties and have emerged as a very important class of nanomaterials for both fundamental studies and potential applications.^{1–31}

Thanks to the versatility of properly devised bottom-up wet-chemistry synthesis and purification procedures, many MPC have been successfully characterized with atomic precision *via* mass spectrometry and single-crystal X-ray crystallography, thereby achieving a level of information that surpasses by far the more elusive bare 'naked' metal clusters originally investigated and characterized in, *e.g.*, molecular beams and inert

dielectric matrices. Knowledge of the crystal geometry and composition has been also instrumental to devise experimental and computational characterization procedures, and this provided the chance of cross-validating methods and results in a rigorous and robust framework. Cross-validation is especially important to develop and refine computational methods aimed to be quantitatively predictive in the study of the MPCs' structure/property relationships, such as first-principles quantum mechanics (QM) methods. For example, although several different theoretical methods are available and have been employed to study MPCs with various degrees of accuracy and computational cost, it is very important to assess their performances and shortcomings in view of choosing the most suitable method for a specific investigation, while keeping the expected error and computational effort under control. Once validated, computational approaches can then be employed with confidence to obtain accurate and predictive information also when experimental information is not available, and ultimately, design new materials with desired properties.

The optical response, particularly the photoabsorption spectrum of MPCs, is one of the those properties for which accurate validation studies are desirable.^{32,33} Indeed, whereas a qualitative agreement between theory and experiment has been generally obtained,^{34–36} a quantitative match is much more challenging and has rarely been achieved.^{37–39} In this context, one of the first quantitatively accurate optical-response calculations for an MPC system was carried out for $\text{Au}_{30}(\text{S}-t\text{Bu})_{18}$.³⁷

^a Dipartimento di Scienze Chimiche e Farmaceutiche, Università di Trieste, Via Giorgieri 1, I-34127 Trieste, Italy. E-mail: stener@units.it

^b CNR-ICCOM & IPCF, Consiglio Nazionale delle Ricerche, Via Giuseppe Moruzzi 1, I-56124 Pisa, Italy. E-mail: alessandro.fortunelli@cnr.it

^c Department of Chemistry, University of Padova, Via Marzolo 1, I-35131 Padova, Italy. E-mail: flavio.maran@unipd.it

^d Department of Chemistry, University of Connecticut, 55 North Eagleville Road, Storrs, Connecticut 06269, USA

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The conclusions of that study were that the necessary conditions for a quantitative match between theory and experiment are: (i) use an accurate MPC geometry, *e.g.*, employing the experimental geometry (from X-ray diffraction) in the calculation to avoid inaccuracies due to QM structure prediction; (ii) employ a time dependent density functional theory (TDDFT) method that uses hybrid exchange–correlation (xc-) functionals (such as the B3LYP^{40,41} xc-functional in both the Kohn–Sham equation and the response kernel).

More recently,⁴² some of us have considered an even more stringent comparison regarding the hybrid B3LYP TDDFT scheme calculation and the low temperature (77 K) absorption spectrum of the Ag₂₄Au(DMBT)₁₈[−] anion. In that work an almost quantitative agreement was attained across the entire energy interval considered in the experimental spectral analysis. In particular, it was confirmed that the use of a hybrid xc-functional is mandatory to properly describe the low energy optical spectral region, where electronic transitions involving metal–ligand bonds play a major role. In contrast, GGA xc-functionals such as PBE⁴³ or even asymptotically correct xc-functionals such as LB94⁴⁴ gave a significantly worse agreement with experiment in the low energy region of the spectrum, although performing better at higher energies, which are dominated by ligand excitations. A previous complete study regarding the choice of the exchange–correlation functional³⁸ (as well the basis set and the geometry) has allowed to assess which are the principal requirements of TDDFT for the calculation of photoabsorption of MPCs. In general it was concluded that asymptotically corrected functionals like LB94 and SAOP perform much better than standard GGA, but the low energy region remains difficult to predict and can be properly described only by hybrid functionals like B3LYP. The accurate prediction of the optical response of MPCs represents a real challenge to theory, due the co-existence and, thus, coupling of three moieties with quite different electronic properties: (1) a metal core in a low-oxidation state; (2) the metal shell of the “staple” motifs that interacts with both metal core and ligands, (3) the ligands (thiolates, *etc.*). Indeed, it has been shown that the B3LYP hybrid xc-functional and kernel enable one to obtain a balanced description of the electronic structure of regions with such a different chemical nature.^{37,38} It is worth stressing that a quantitative prediction of optical excitations in MPCs is particularly important for understanding properties and exploiting these compounds in view of applications, such as sensing and photocatalysis.⁴⁵

The present work develops within this research framework. Our main goal consists in testing and extending the validation of the B3LYP scheme to a larger and more diverse set of clusters than previously investigated. The present study, where we vary the nature of the metal and the ligands, is particularly relevant to assess the predictivity of the previously proposed B3LYP scheme, as it is known³⁷ that the optical properties of MPC are affected by both the metal and the ligand components. We selected three metal clusters that share the same M₂₅ metallic core structure but have different metal composition and ligands: Ag₂₅(DMBT)₁₈[−], Ag₂₄Pt(DMBT)₁₈^{2−} and Au₂₄Pt(SC₄H₉)₁₈ where DMBT is 2,4-dimethylbenzenethiolate and SC₄H₉ is *n*-butylthiolate. These

nanoclusters can be prepared with atomic precision and offer a diverse set of MPC compounds on which to test and validate our theoretical approach.

2. Theoretical method

To predict optical spectra, we employ the complex polarizability TDDFT (polTDDFT) algorithm.⁴⁶ This approach is able to treat very large systems by avoiding the bottleneck of the more traditional Casida diagonalization *via* a direct solution of the response equations, while still providing a very good accuracy when compared with the exact Casida solution.^{38,47} The reader is referred to the original work for a detailed description of the algorithm⁴⁶ and its implementation in the AMS/ADF program.⁴⁸

In practice, the photoabsorption spectrum $\sigma(\omega)$ is calculated point by point, from the imaginary part of the dynamical polarizability $\alpha(\omega)$:

$$\sigma(\omega) = \frac{4\pi\omega}{c} \text{Im}[\alpha(\omega)] \quad (1)$$

This expression is of practical interest when the polarizability is calculated for complex frequency, *i.e.* $\omega = \omega_r + i\omega_i$, where the real part ω_r is the scanned photon frequency (energy) and ω_i is the imaginary part which corresponds to a Lorentzian broadening of the discrete lines and can be interpreted as a pragmatic inclusion of the excited states finite lifetime. The complex dynamical polarizability is calculated by solving the following non-homogeneous linear system, working with a basis set of density-fitting functions:

$$[\mathbf{S} - \mathbf{M}(\omega)]\mathbf{b} = \mathbf{d} \quad (2)$$

In eqn (2) $\mathbf{S}$ is the overlap matrix between fitting functions, $\mathbf{b}$ is the unknown vector with the expansion coefficients $b_\mu(\omega)$ of the time-dependent electron density $\rho_z^{(1)}$ induced by the electromagnetic field, $\mathbf{d}$ is the frequency dependent vector corresponding to the known non-homogeneous term, and finally the elements of the frequency dependent matrix $\mathbf{M}$ are:

$$M_{\mu\nu} = \langle f_\mu | \chi_{\text{KS}}(\omega) K | f_\nu \rangle \quad (3)$$

In eqn (3) χ_{KS} is the Kohn–Sham frequency-dependent dielectric function, K is the kernel, and f_μ and f_ν are the elements of the basis set of density-fitting functions. Note that the matrix element in eqn (3) is between density fitting functions, and therefore the present implementation of the algorithm allows one to employ only density-dependent kernels. For this reason, the density-matrix-dependent Hartree–Fock exchange kernel, which is an ingredient of hybrid xc-functionals, cannot directly be employed in polTDDFT. However, it is still possible to employ hybrid kernels with nonlocal Hartree–Fock exchange components if the hybrid diagonal approximation (HDA) is employed. The HDA method, which we have recently proposed,⁴⁹ to which we refer for further details is a very efficient and robust approximation to TDDFT: by correcting only the diagonal matrix elements for the non-local exchange, while including local-density exchange–correlation terms of the kernel in the off-diagonal elements, the computational complexity is drastically reduced while keeping an excellent accuracy.⁴⁹



3. Experimental and computational details

$[n\text{-Oct}_4\text{N}^+][\text{Ag}_{25}(\text{DMBT})_{18}^-]$, $[n\text{-Oct}_4\text{N}^+]_2[\text{Ag}_{24}\text{Pt}(\text{DMBT})_{18}^{2-}]$, and $\text{Au}_{24}\text{Pt}(\text{SC}_4\text{H}_9)_{18}$ were synthesized according to known procedures,^{50–52} with minor modifications (see ESI†).

Photoabsorption measurements were carried out with a near-infrared VARIAN Cary 60 spectrophotometer with a spectral resolution of 0.5 nm. The spectra were collected over a range of 250–900 nm. $[n\text{-Oct}_4\text{N}^+][\text{Ag}_{25}(\text{DMBT})_{18}^-]$ and $[n\text{-Oct}_4\text{N}^+]_2[\text{Ag}_{24}\text{Pt}(\text{DMBT})_{18}^{2-}]$ were dissolved in a 5:4:2 vol butyronitrile/propionitrile/dichloromethane solution, whereas $\text{Au}_{24}\text{Pt}(\text{SC}_4\text{H}_9)_{18}$ was dissolved in a 1:1 vol tetrahydrofuran/toluene solution; we verified that both solvent combinations form clear glasses at low temperatures. The solutions were adjusted to obtain an absorbance of 0.2 at 400 nm, using a 1 mm optical-path quartz cuvette. The temperature was controlled with an OptistatDN cryostat equipped with a MercuryTC temperature controller (Oxford Instrument). A HiCube 80 Eco, DN 63 ISO-K turbo pumping station (PFEIFFER Vacuum) was used to generate a high vacuum inside the outer vacuum chamber of the cryostat. Before experiment, the sample holder was purged with high-purity helium, which was used as the temperature-exchange gas.

The cluster geometries used to perform TDDFT simulations were taken from the corresponding experimental X-ray structures.⁵¹ The geometries were partially optimized in the ligand coordinates, while freezing the metal and sulfur atoms at their experimental geometry.

Geometry relaxations were performed using the CP2K package⁵³ at the DFT/PBE level⁴³ with the addition of the Grimme-D3 dispersion terms.⁵⁴ DFT simulations, based on the hybrid Gaussian/plane-wave scheme (GPW),⁵⁵ employed double-zeta-valence-plus polarization (DZVP) basis sets⁵⁶ to represent the DFT Kohn–Sham orbitals, GTH pseudopotentials for describing the core-electrons of all the atomic species,⁵⁷ and an auxiliary set of plane-waves whose cut-off was set to 300 Ry.

From these geometries, optical-response TDDFT calculations were performed with the version 2023.01 of the Amsterdam modeling suite (AMS)⁵⁸ set of programs,^{59,60} employing a basis set of Slater type orbitals (STO) functions of Triple Zeta plus polarization (TZP) quality and optimized density fitting functions.⁶¹ Scalar relativistic effects were included *via* the Zero order regular approximation⁶² (ZORA). The hybrid B3LYP xc-functional was employed, which includes a portion of Hartree–Fock non-local exchange and provides consistently the most accurate predictions among DFT approaches. In this work, we employed the resolution of the identity (RI) technique to calculate the B3LYP HDA exchange integrals, as recently implemented in version 2023.01 of AMS.⁶³ Such implementation, known as ‘fitted HDA’, allows a speedup up to a factor of 30 in the calculation of integrals with respect to the previous version that used numerical quadrature. The polTDDFT results were broadened by a Lorentzian function to obtain 0.150 eV of FWHM. An imaginary frequency ω_i of 0.075 eV was employed in eqn (1).

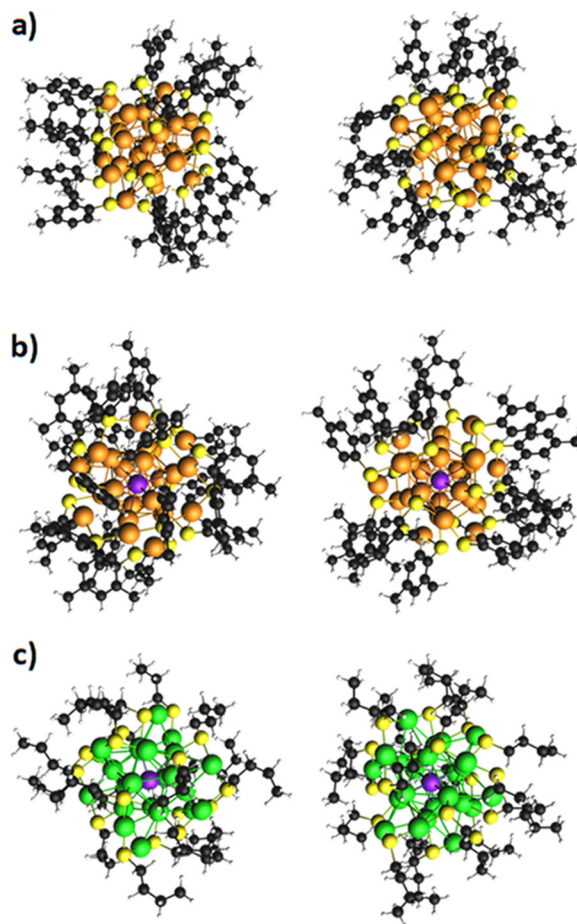


Fig. 1 Geometries of the clusters (a) $\text{Ag}_{25}(\text{DMBT})_{18}^-$, (b) $\text{Ag}_{24}\text{Pt}(\text{DMBT})_{18}^{2-}$, and (c) $\text{Au}_{24}\text{Pt}(\text{SC}_4\text{H}_9)_{18}$.

4. Results and discussion

4.1. Cluster structures

Fig. 1 shows a schematic depiction of the structures of the three clusters, each one taken from two different points of view. The systems are characterized by the typical M_{25} structure which can be briefly described as follows:⁶⁴ (1) one metal atom is in the center of the cluster; (2) 12 metal atoms form an icosahedron around the central atom; (c) a shell of ligands consisting of six ‘staple’ units $-\text{SR}-\text{M}-\text{SR}-\text{M}-\text{SR}-$ protects the inner metal core. In the cluster series here considered: $\text{Ag}_{25}(\text{DMBT})_{18}^-$, $\text{Ag}_{24}\text{Pt}(\text{DMBT})_{18}^{2-}$ and $\text{Au}_{24}\text{Pt}(\text{SC}_4\text{H}_9)_{18}$, the first compound is a monometallic silver cluster, while the next terms correspond to monodoped silver and gold clusters respectively, where the central Au or Ag atom is replaced by a Pt atom. As to the ligand, whereas the first two clusters are protected by the typical aromatic DMBT ligand, SC_4H_9 , is aliphatic and this may have effects on the optical properties. Indeed, it has been shown⁶⁵ that aromatic thiolates can promote a plasmon ‘rebirth’ already at cluster sizes that would not yet be plasmonic. Although the first cluster is anionic, the second is dianionic, and the last one is neutral, they all display a closed-shell electronic structure.



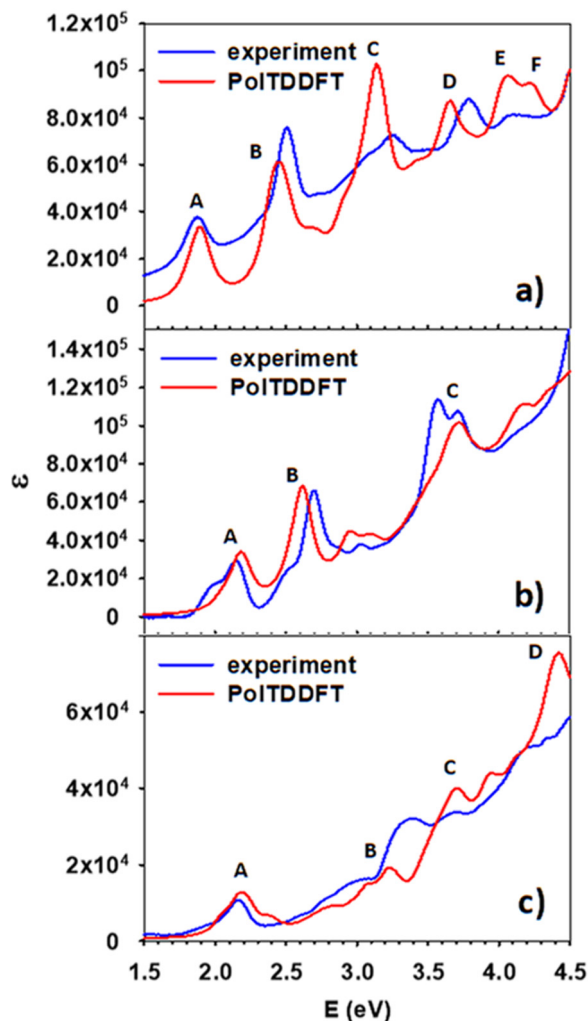


Fig. 2 Comparison between experimental and calculated photoabsorption of the clusters (a) $\text{Ag}_{25}(\text{DMBT})_{18}^-$, (b) $\text{Ag}_{24}\text{Pt}(\text{DMBT})_{18}^{2-}$, and (c) $\text{Au}_{24}\text{Pt}(\text{SC}_4\text{H}_9)_{18}$.

4.2. Cluster selection and comparison with experiment

Fig. 2 shows the calculated photoabsorption of the three clusters calculated at the TDDFT level, as described in Sections 2 and 3, and the experimental spectra obtained at 77 K. The intensity scale is absolute for the calculated profiles, whereas the experimental profile intensity has been rescaled to achieve a best match with the calculated profile and compare the intensity of the calculated profile across different clusters. To have a more general perspective, we start by discussing the whole series, and defer the discussion of the single terms of the series to the following sections.

First, we observe that for all three clusters a fairly nice agreement between theory and experiment holds. All the experimental features are properly reproduced by theory, with modest energy shifts of few tenth of eV and a rather accurate distribution of intensities, with only occasional deterioration. These findings provide a strict validation of our theory approach, demonstrating that the fitted HDA B3LYP polTDDFT scheme can attain quantitative accuracy even with

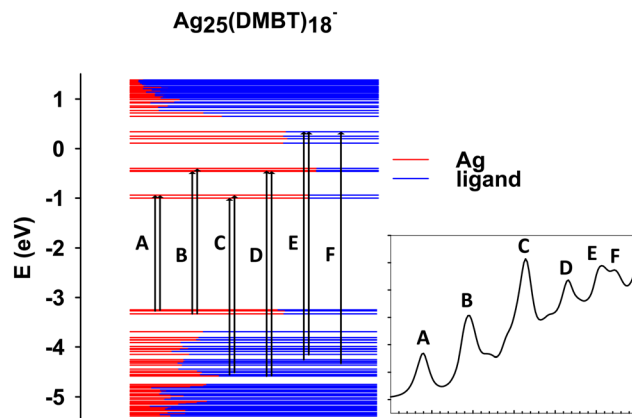


Fig. 3 Scheme with the most important occupied-virtual configurations involved in the transitions in $\text{Ag}_{25}(\text{DMBT})_{18}^-$. The labels refer to specific features as reported in the inset. The color of the levels represents the nature of the molecular orbitals in terms of fragments (see legend).

metal doping and with different ligand nature and electronic effects.

Considering the trend of the spectra along the series, we notice that whereas $\text{Ag}_{25}(\text{DMBT})_{18}^-$ and $\text{Ag}_{24}\text{Pt}(\text{DMBT})_{18}^{2-}$ display a photoabsorption spectrum very rich of spectral features, with many well resolved bands and maxima, the spectrum of $\text{Au}_{24}\text{Pt}(\text{SC}_4\text{H}_9)_{18}$, aside from the region at low energy, is much less structured, consisting mainly of weak shoulders superimposed to a background smoothly increasing with energy.

4.3. $\text{Ag}_{25}(\text{DMBT})_{18}^-$

In this section we focus on the specific analysis of the photoabsorption of $\text{Ag}_{25}(\text{DMBT})_{18}^-$. Fig. 3 shows a schematic picture of the main features (designated by the same capital letters from A to F as in Fig. 2a, as described in the inset), whereas Table 1 provides for each spectral feature the main contributions in terms of electronic configurations. From Fig. 2a it appears that the theory generally underestimates the experimental absorption energy: for the features A, E and F the difference is less than 0.1 eV, whereas a slightly larger disagreement is found for the other features, the largest deviation of 0.13 eV being found for feature D. The shape of the peaks is very well reproduced by theory, but for peak C where a quite sharper shape is predicted: this difference could be attributed to vibrational broadening or to a deficiency in the description of the electronic structure of the cluster. Regarding the assignment of the transitions, peaks A and B have a character very close to HOMO/LUMO excitations, ranging from HOMO–2 to LUMO+4. Interestingly, their energy difference is attributed mainly to the different energy of the virtual orbitals, which are grouped in two close groups separated by about 0.5 eV (see Fig. 3). All these orbitals have large silver contribution (see also the molecular orbital plots in Fig. S1 and S2 of the ESI†) that is especially large in LUMO and LUMO+1. We can describe the character of these orbitals as Ag–S σ -bonds for the HOMO, HOMO–1 and HOMO–2, whereas those from LUMO and LUMO+4 mainly pertain to Ag atoms only. With this picture in mind, we can



Table 1 Analysis of the spectral features of $\text{Ag}_{25}(\text{DMBT})_{18}^{-}$. Label feature, calculated and experimental energy (eV), calculated molar absorption ($\text{M}^{-1}\text{cm}^{-1}$) and assignment

Excitation	Excitation E (eV)	Experimental E (eV)	ϵ	Assignment
A	1.90	1.88	33 555	56.61% HOMO–1 (32% Ag 5s, 21% S 3p, 8% Ag 5p) $\rightarrow$ LUMO+1 (29% Ag 5s, 15% Ag 5p, 4% S 3p, 1% Ag 4d); 30.54% HOMO (28% Ag 5s, 20% S 3p, 7% Ag 5p) $\rightarrow$ LUMO+1
B	2.46	2.51	61 256	21.50% HOMO–2 (31% Ag 5s, 18% S 3p, 7% Ag 5p, 5% Ag 4d) $\rightarrow$ LUMO+2 (25% Ag 5p, 23% Ag 5s, 6% S 3p, 1% Ag 4d); 15.26% HOMO $\rightarrow$ LUMO+4 (26% Ag 5p, 22% Ag 5s, 6% S 3p, 1% Ag 4d); 11.51% HOMO–2 $\rightarrow$ LUMO+4
C	3.16	3.22	102 811	24.56% HOMO–22 (38% S 3p, 6% Ag 4d, 2% Ag 5s) $\rightarrow$ LUMO+1; 16.32% HOMO–20 (27% S 3p, 3% Ag 5s) $\rightarrow$ LUMO+1
D	3.66	3.79	87 273	8.91% HOMO–23 (34% S 3p, 5% Ag 4d, 4% Ag 5s) $\rightarrow$ LUMO+3 (29% Ag 5p, 7% S 3p, 2% Ag 5s, 1% Ag 4d); 8.31% HOMO–22 $\rightarrow$ LUMO+2
E	4.06	4.07	97 827	10.19% HOMO–13 (34% S 3p, 2% C 2p, 1% Ag 4d) $\rightarrow$ LUMO+8 (28% Ag 5p, 9% Ag 5s, 1% C 2p); 8.89% HOMO–12 (32% S 3p, 1% C 2p) $\rightarrow$ LUMO+8
F	4.22	4.15	94 691	32.53% HOMO–16 (29% S 3p, 5% Ag 5s, 4% Ag 4d, 1% C 2p) $\rightarrow$ LUMO+8

attribute the A and B features to transitions from Ag–S σ -bonds to the metal core. B is more intense due to the largest metal contribution in the final virtual orbitals. C is the most intense band, whereas in the experiment it is more broadened as already noted. This feature is attributed to transitions starting from rather deep occupied orbitals (HOMO–20, HOMO–21) with major ligand character (Fig. S1, ESI[†]). The nature of the C band can also explain why it is so broadened in the experiment: the mainly ligand character of the occupied orbitals is conceivably more sensitive to their conformational freedom, which in solution is particularly relevant for the ligands but is much less important for the metal core.^{66,67} The next E and F features are less pronounced, with the orbitals involved in such transition localized mainly on the ligands.

4.4. $\text{Ag}_{24}\text{Pt}(\text{DMBT})_{18}^{2-}$

Here we focus on the first doped cluster of the series, namely $\text{Ag}_{24}\text{Pt}(\text{DMBT})_{18}^{2-}$. Fig. 4 shows the scheme of the molecular orbital energy levels together with the main spectral features (A, B and C), whereas the main contributions in terms of electronic configurations are reported in Table 2. Fig. 2b illustrates that the theory reproduces remarkably well the spectral window

here considered, with the position of the three peaks in quantitative agreement with experiment, and the largest deviation for peak C limited to only 0.15 eV. Furthermore, the intensity distribution is quite accurately predicted by theory, and it is also worth noting that the profile shape between the peaks follows very closely the experimental trend. The energy levels from Fig. 4 give a situation qualitatively like the previous pure silver cluster: the A and B bands again start from the same group of occupied orbitals (HOMO, HOMO–1 and HOMO–2) but arrive to two different groups of virtual orbitals, separated by about 0.3 eV (LUMO and LUMO+1 for feature A and LUMO+2 and LUMO+3 and LUMO+4 for feature B). Interestingly, in comparison with $\text{Ag}_{25}(\text{DMBT})_{18}^{-}$, the nature of the transitions is the same but both transitions are blue shifted, which is a consequence of the larger HOMO/LUMO gap as well as the smaller separation between virtual states involved in features A and B. It is worth noting that the calculated shape of the peak C of the present cluster is in very good agreement with the experiment, at variance of previous Ag_{25} cluster where the experimental peak was strongly broadened. We attributed the strong broadening to a consequence of the ligand conformational freedom. Therefore it is interesting to argue if this different behavior could be ascribed to the Pt doping and/or to the charge of the cluster, since the rest of the system (such as the Ag geometry and the ligand nature) is the same. Indeed, the Pt–Ag bonds are much more ‘directional’ and stronger than Ag–Ag bonds, due to the role played by the Pt 5d orbitals, whereas in Ag the 4d orbitals are deeper in energy and play a very minor role in the chemical bond formed by Ag. Moreover, the larger charge of the Pt–Ag cluster should entail a larger ‘pressure’ by the solvent, which should also compress the ligands’ degrees of freedom. This enhanced ‘Pt–Ag bond directionality’ together with a larger environmental pressure should inhibit the conformational freedom of the ligands giving a sharper peak.

In summary, the computational model describes properly the optical behavior, ascribing the observed differences to the presence of the Pt atom. The Pt contribution remains below 10% in all the orbitals considered, except for some that are rather deep, around –3 eV, and do not contribute appreciably to the spectral features. Incidentally, the Pt atom participates mainly with the 5d electrons, which are rather diffuse, thereby

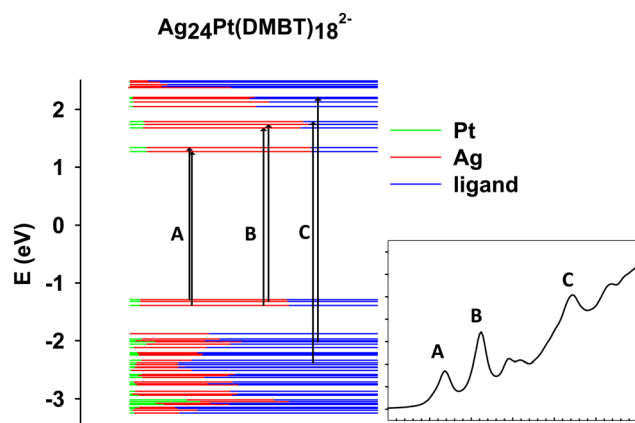


Fig. 4 Scheme with the most important occupied-virtual configurations involved in the transitions in $\text{Ag}_{24}\text{Pt}(\text{DMBT})_{18}^{2-}$. The labels refer to specific features as reported in the inset. The color of the levels represent the nature of the molecular orbitals in terms of fragments (see legend).



Table 2 Analysis of the spectral features of $\text{Ag}_{24}\text{Pt}(\text{DMBT})_{18}^{2-}$. Label feature, calculated and experimental energy (eV), calculated molar absorption ($\text{M}^{-1}\text{cm}^{-1}$) and assignment

Excitation	Excitation E (eV)	Experimental E (eV)	ϵ	Assignment
A	2.18	2.15	34 131	37.37% HOMO (28% Ag 5s, 20% S 3p, 3% Ag 4d, 3% Pt 6p, 2% Ag 5p) $\rightarrow$ LUMO+1 (26% Ag 5s, 19% Ag 5p, 6% Pt 5d); 18.90% HOMO-2 (31% Ag 5s, 18% S 3p, 6% Ag 5p, 2% Pt 6s) $\rightarrow$ LUMO
B	2.62	2.70	68 651	24.62% HOMO-2 $\rightarrow$ LUMO+2 (21% Ag 5p, 19% Ag 5s, 4% Pt 5d, 1% S 3p); 22.00% HOMO-1 $\rightarrow$ LUMO+3 (23% Ag 5s, 25% Ag 5p, 3% Pt 5d, 3% S 3p)
C	3.72	3.58	101 702	10.65% HOMO-13 (29% S 3p, 16% C 2p, 1% Pt 5d) $\rightarrow$ LUMO+4 (22% Ag 5p, 19% Ag 5s, 4% Pt 5d, 2% S 3p); 9.10% HOMO-6 (27% S 3p, 9% Pt 5d, 7% Ag 4d, 3% Ag 5s, 1% C 2p) $\rightarrow$ LUMO+8 (16% Ag 5p, 8% Ag 5s, 1% C 2p)

providing kind of an 'electron reservoir'. We can conclude that for this cluster the spectral difference with respect to the silver cluster can be rationalized in terms of the nature of the central atom.

4.5. $\text{Au}_{24}\text{Pt}(\text{SC}_4\text{H}_9)_{18}$

The last compound considered in this work consists of a neutral gold cluster protected by alkyl ligands and doped with Pt, $\text{Au}_{24}\text{Pt}(\text{SC}_4\text{H}_9)_{18}$. Fig. 5 shows the molecular-orbital energy levels, with the main features (A, B, C and D) analyzed in Table 3 in terms of leading electronic configurations. From the previous Fig. 2c, it appears that the theory reproduces quantitatively the peak A at low energy (2.2 eV). At higher energy the photoabsorption profile is not as highly structured as in the two silver-based clusters but is much smoother. Only weak shoulders are predicted to show up in the energy range 3.2–3.7 eV, which is in satisfactory agreement with experiment. For this cluster, the theoretical spectrum is blue shifted with respect to the experiment by few tenths of eV. The predicted feature D is unfortunately outside the experimental energy interval, and thus, we do not have an experimental counterpart for comparison. Also for this cluster it is worth noting that, despite a minor disagreement in the energy positions of features B and C, the calculated intensity profile resembles closely the experimental shape, which features a typical 'step' behavior. In

the present work we did not consider the corresponding pure gold clusters, as they have already been the object of previous extensive experimental and theoretical studies,^{48,64,68} and thus we can profitably compare the present results on the Au/Pt nanoalloy with the photoabsorption spectrum of $\text{Au}_{25}(\text{SR})_{18}^-$ measured at room temperature⁶⁸ and the absorption spectrum we previously calculated at the TDDFT B3LYP HDA level.⁴⁹ Note that in ref. 49 the experimental $\text{R} = \text{CH}_2\text{CH}_2\text{C}_6\text{H}_5$ ligand was simplified as $\text{R} = \text{CH}_3$ in the calculations, but keeping the experimental geometry, in such a way that the alkyl ligand nature was preserved (important to describe properly the optical spectrum). This comparison indicates that Pt doping shifts the A feature to higher energy from 1.8 eV in $\text{Au}_{25}(\text{SR})_{18}^-$ to 2.2 eV in $\text{Au}_{24}\text{Pt}(\text{SC}_4\text{H}_9)_{18}$ in both experiment and calculations. In the experimental $\text{Au}_{25}(\text{SR})_{18}^-$ spectrum other less pronounced features appear at 2.75, 3.1 and 3.9 eV, all of them shifted at higher energy in $\text{Au}_{24}\text{Pt}(\text{SC}_4\text{H}_9)_{18}$ by about 0.5–0.6 eV on average. We conclude that Pt doping definitely promotes a blue shift of the spectral features by about 0.4–0.6 eV depending on the specific structure, in both Ag as well as Au clusters.

5. Conclusions

The photoabsorption spectra of three MPC compounds with different alloy nature and ligands: $\text{Ag}_{25}(\text{DMBT})_{18}^-$, $\text{Ag}_{24}\text{Pt}(\text{DMBT})_{18}^{2-}$ and $\text{Au}_{24}\text{Pt}(\text{SC}_4\text{H}_9)_{18}$, where DMBT is 2,4-dimethylbenzenethiolate and SC_4H_9 is *n*-butylthiolate, have been measured at low temperature (77 K) in order to better resolve single optical peaks and compare them with state-of-the-art theoretical simulations at the TDDFT level of theory using hybrid xc-functional and experimentally-derived geometries. For all three cases, we find an excellent match between theory and experiment, with typical deviations being less than 0.1 eV, occasionally reaching 0.2 eV. This result provides a stringent validation of the proposed computational protocol. In addition to validated accuracy, the calculations are analyzed to assign the relevant spectra features, ascribing them to specific transitions between occupied-virtual orbital pairs. It is found that doping by Pt has an important effect on the spectra, shifting the spectral features to higher energy by about 0.4 eV, for both silver and gold clusters. Future developments of the present work would be towards metal clusters containing atoms of different nature with respect to the present ones, as well as ligands different from thiols. This would help to trace more general conclusions regarding the transferability of the method. Moreover

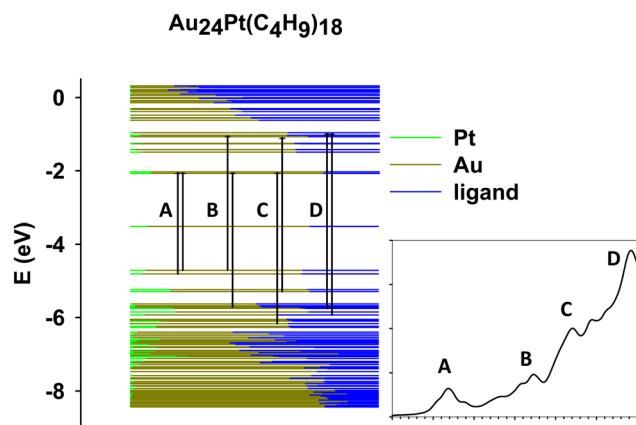


Fig. 5 Scheme with the most important occupied-virtual configurations involved in the transitions in $\text{Au}_{24}\text{Pt}(\text{SC}_4\text{H}_9)_{18}$. The labels refer to specific features as reported in the inset. The color of the levels represent the nature of the molecular orbitals in terms of fragments (see legend).



Table 3 Analysis of the spectral features of Au₂₄Pt(SC₄H₉)₁₈. Label feature, calculated and experimental energy (eV), calculated molar absorption (M⁻¹ cm⁻¹) and assignment

Excitation	Excitation <i>E</i> (eV)	Experimental <i>E</i> (eV)	ε	Assignment
A	2.20	2.16	12 804	34.51% HOMO−1 (17% Au 6s, 17% S 3p, 8% Au 5d, 5% Pt 6p) → LUMO+2 (29% Au 6s, 15% Au 6p, 9% Pt 5d, 2% S 3p); 25.04% HOMO (23% S 3p, 15% Au 6s, 11% Au 5d, 7% Pt 6p, 4% Au 6p) → LUMO+2
B	3.24	3.06	19 247	17.44% HOMO → LUMO+8 (25% Au 6p, 9% Au 6s, 6% S 3p, 1% C 2p, 1% H 2s); 15.88% HOMO−6 (30% S 3p, 11% Au 5d, 1% Pt 5d) → LUMO+2
C	3.70	3.40	40 118	16.24% HOMO−14 (20% S 3p, 18% Au 5d, 12% Au 6s) → LUMO+2; 14.65% HOMO−3 (16% Pt 5d, 14% Au 5d, 10% S 3p, 8% Au 6s) → LUMO+7 (30% Au 6p, 13% Au 6s, 3% S 3p)
D	4.42	—	75 502	8.05% HOMO−8 (17% Pt 5d, 16% Au 5d, 14% S 3p, 1% Au 6s) → LUMO+9 (28% Au 6p, 2% Au 6s, 2% S 3p); 6.91% HOMO−10 (26% S 3p, 17% Au 5d, 7% Au 6s, 4% Pt 5d, 1% Pt 7s) → LUMO+9

future inclusion of range-separated xc-functionals, necessary to describe charge transfer excitations, is under consideration.

Conflicts of interest

There are no conflicts of interest to declare.

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References

- R. L. Whetten, J. T. Khoury, M. M. Alvarez, S. Murthy, I. Vezmar, Z. L. Wang, P. W. Stephens, C. L. Cleveland, W. D. Luedtke and U. Landman, *Adv. Mater.*, 1996, **8**, 428.
- P. D. Jadzinsky, G. Calero, C. J. Ackerson, D. A. Bushnell and R. D. Kornberg, *Science*, 2007, **318**, 430.
- J. Akola, M. Walter, R. L. Whetten, H. Häkkinen and H. Grönbeck, *J. Am. Chem. Soc.*, 2008, **130**, 3756.
- M. W. Heaven, A. Dass, P. S. White, K. M. Holt and R. W. Murray, *J. Am. Chem. Soc.*, 2008, **130**, 3754.
- H. Qian, W. T. Eckenhoff, Y. Zhu, T. Pintauer and R. Jin, *J. Am. Chem. Soc.*, 2010, **132**, 8280.
- O. Lopez-Acevedo, H. Tsunoyama, T. Tsukuda, H. Häkkinen and C. M. Aikens, *J. Am. Chem. Soc.*, 2010, **132**, 8210.
- C. Zeng, H. Qian, T. Li, G. Li, N. L. Rosi, B. Yoon, R. N. Barnett, R. L. Whetten, U. Landman and R. Jin, *Angew. Chem.*, 2012, **51**, 13114.
- I. Dolamic, S. Knoppe, A. Dass and T. Bürgi, *Nat. Commun.*, 2012, **3**, 798.
- Y. Negishi, C. Sakamoto, T. Ohyama and T. Tsukuda, *J. Phys. Chem. Lett.*, 2012, **3**, 1624.
- A. Desiredy, B. E. Conn, J. Guo, B. Yoon, R. N. Barnett, B. M. Monahan, K. Kirschbaum, W. P. Griffith, R. L. Whetten, U. Landman and T. P. Bigioni, *Nature*, 2013, **501**, 399.
- S. Mustalahti, P. Myllyperkiö, T. Lahtinen, K. Salorinne, S. Malola, J. Koivisto, H. Häkkinen and M. Pettersson, *J. Phys. Chem. C*, 2014, **118**, 18233.
- A. Dass, S. Theivendran, P. R. Nimmala, C. Kumara, V. R. Jupally, A. Fortunelli, L. Sementa, G. Barcaro, X. Zuo and B. C. Noll, *J. Am. Chem. Soc.*, 2015, **137**, 4610.
- A. Fernando, K. L. D. M. Weerawardene, N. V. Karimova and C. M. Aikens, *Chem. Rev.*, 2015, **115**, 6112.
- W. Kurashige, Y. Niihori, S. Sharma and Y. Negishi, *Coord. Chem. Rev.*, 2016, **320–321**, 238.
- R. Jin, C. Zeng, M. Zhou and Y. Chen, *Chem. Rev.*, 2016, **116**, 10346.
- N. A. Sakthivel, S. Theivendran, V. Ganeshraj, A. G. Oliver and A. Dass, *J. Am. Chem. Soc.*, 2017, **139**, 15450.
- M. Rambukwella, N. A. Sakthivel, J. H. Delcamp, L. Sementa, A. Fortunelli and A. Dass, *Front. Chem.*, 2018, **6**, 330.
- T. Dainese, S. Antonello, S. Bogialli, W. Fei, A. Venzo and F. Maran, *ACS Nano*, 2018, **12**, 7057.
- Z. Lei, J. J. Li, X. K. Wan, W. H. Zhang and Q. M. Wang, *Angew. Chem.*, 2018, **57**, 8639.
- M. A. Abbas, P. V. Kamat and J. H. Bang, *ACS Energy Lett.*, 2018, **3**, 840.
- N. A. Sakthivel, M. Shabaninezhad, L. Sementa, B. Yoon, M. Stener, R. L. Whetten, G. Ramakrishna, A. Fortunelli, U. Landman and A. Dass, *J. Am. Chem. Soc.*, 2020, **142**, 15799.
- C. A. Hosier, I. D. Anderson and C. J. Ackerson, *Nanoscale*, 2020, **12**, 6239.



- 23 T. Kawawaki, A. Ebina, Y. Hosokawa, S. Ozaki, D. Suzuki, S. Hossain and Y. Negishi, *Small*, 2021, 2005328.
- 24 J. V. Rival, P. Mymoonna, K. M. Lakshmi Nonappa, T. Pradeep and E. S. Shibu, *Small*, 2021, 2005718.
- 25 T. Dainese, M. Agrachev, S. Antonello, D. Badocco, D. M. Black, A. Fortunelli, J. A. Gascón, M. Stener, A. Venzo, R. L. Whetten and F. Maran, *Chem. Sci.*, 2018, 9, 8796.
- 26 M. Agrachev, M. Ruzzi, A. Venzo and F. Maran, *Acc. Chem. Res.*, 2019, 52, 44.
- 27 M. Agrachev, W. Fei, S. Antonello, S. Bonacchi, T. Dainese, A. Zoleo, M. Ruzzi and F. Maran, *Chem. Sci.*, 2020, 11, 3427.
- 28 E. Boisselier and D. Astruc, *Chem. Soc. Rev.*, 2009, 38, 1759.
- 29 U. H. F. Bunz and V. M. Rotello, *Angew. Chem.*, 2010, 49, 3268.
- 30 Y. Du, H. Sheng, D. Astruc and M. Zhu, *Chem. Rev.*, 2020, 120, 526.
- 31 R. Jin, G. Li, S. Sharma, Y. Li and X. Du, *Chem. Rev.*, 2021, 121, 567.
- 32 A. Tlahuice-Flores, *Phys. Chem. Chem. Phys.*, 2016, 18, 27738.
- 33 C. M. Aikens, *Acc. Chem. Res.*, 2018, 51, 3065.
- 34 K. L. D. M. Weerawardene and C. M. Aikens, *Phys. Chem. C*, 2018, 122, 2440.
- 35 K. L. D. M. Weerawardene and C. M. Aikens, *J. Am. Chem. Soc.*, 2016, 138, 11202.
- 36 G. Barcaro, L. Sementa, A. Fortunelli and M. Stener, *Phys. Chem. Chem. Phys.*, 2015, 17, 27952.
- 37 A. Dass, T. Jones, M. Rambukwella, D. Crasto, K. Gagnon, L. Sementa, M. De Vetta, O. Baseggio, E. Aprà, M. Stener and A. Fortunelli, *J. Phys. Chem. C*, 2016, 120, 6256.
- 38 O. Baseggio, M. De Vetta, G. Fronzoni, D. Toffoli, M. Stener, L. Sementa and A. Fortunelli, *Int. J. Quantum Chem.*, 2018, 118, e25769.
- 39 H.-Ch. Weissker, H. Barron Escobar, V. D. Thanthirige, K. Kwak, D. Lee, G. Ramakrishna, R. L. Whetten and X. López-Lozano, *Nat. Commun.*, 2014, 5, 3785.
- 40 A. D. Becke, *J. Chem. Phys.*, 1993, 98, 5684.
- 41 P. J. Stephens, F. J. Devlin, C. F. Chabalowski and M. J. Frisch, *J. Phys. Chem.*, 1994, 98, 11623.
- 42 M. Medves, L. Sementa, D. Toffoli, G. Fronzoni, K. R. Krishnadas, T. Bürgi, S. Bonacchi, T. Dainese, F. Maran, A. Fortunelli and M. Stener, *J. Chem. Phys.*, 2021, 155, 084103.
- 43 J. P. Perdew, K. Burke and M. Ernzerhof, *Phys. Rev. Lett.*, 1996, 77, 3865.
- 44 R. Van Leeuwen and E. J. Baerends, *Phys. Rev. A*, 1994, 49, 2421.
- 45 S. Linic, P. Christopher and D. B. Ingram, *Nat. Mater.*, 2011, 10, 911.
- 46 O. Baseggio, G. Fronzoni and M. Stener, *J. Chem. Phys.*, 2015, 143, 024106.
- 47 O. Baseggio, M. De Vetta, G. Fronzoni, M. Stener, L. Sementa, A. Fortunelli and A. Calzolari, *J. Phys. Chem. C*, 2016, 120, 12773.
- 48 O. Baseggio, M. De Vetta, G. Fronzoni, M. Stener and A. Fortunelli, *Int. J. Quantum Chem.*, 2016, 116, 1603.
- 49 M. Medves, L. Sementa, D. Toffoli, G. Fronzoni, A. Fortunelli and M. Stener, *J. Chem. Phys.*, 2020, 152, 184102.
- 50 M. S. Bootharaju, C. P. Joshi, M. R. Parida, O. F. Mohammed and O. M. Bakr, *Angew. Chem., Int. Ed.*, 2016, 55, 922.
- 51 X. Liu, J. Yuan, C. Yao, J. Chen, L. Li, X. Bao, J. Yang and Z. Wu, *J. Phys. Chem. C*, 2017, 121, 13848.
- 52 W. Fei, S. Antonello, T. Dainese, A. Dolmella, M. Lahtinen, K. Rissanen, A. Venzo and F. Maran, *J. Am. Chem. Soc.*, 2019, 141, 16033.
- 53 J. Hutter, M. Iannuzzi, F. Schiffmann and J. VandeVondele, *Wiley Interdiscip. Rev.: Comput. Mol. Sci.*, 2014, 4, 15.
- 54 S. Grimme, J. Antony, S. Ehrlich and H. Krieg, *J. Chem. Phys.*, 2010, 132, 154104.
- 55 G. Lippert, J. Hutter and M. Parrinello, *Theor. Chem. Acc.*, 1999, 103, 124.
- 56 J. VandeVondele and J. Hutter, *J. Chem. Phys.*, 2007, 127, 114105.
- 57 S. Goedecker, M. Teter and J. Hutter, *Phys. Rev. B*, 1996, 54, 1703.
- 58 <https://www.scm.com/product/ams>.
- 59 E. J. Baerends, D. E. Ellis and P. Ros, *Chem. Phys.*, 1973, 2, 41.
- 60 C. Fonseca Guerra, J. G. Snijders, G. Te Velde and E. J. Baerends, *Theor. Chem. Accounts*, 1998, 99, 391.
- 61 M. Medves, G. Fronzoni and M. Stener, *J. Comput. Chem.*, 2022, 43, 1923.
- 62 E. Van Lenthe, E. J. Baerends and E. J. G. Snijders, *J. Chem. Phys.*, 1993, 99, 4597.
- 63 P. D'Antoni, M. Medves, D. Toffoli, A. Fortunelli, M. Stener and L. Visscher, *J. Phys. Chem. A*, 2023, 127, 9244.
- 64 M. W. Heaven, A. Dass, P. S. White, K. M. Holt and R. W. Murray, *J. Am. Chem. Soc.*, 2008, 130, 3754.
- 65 L. Sementa, G. Barcaro, O. Baseggio, M. De Vetta, A. Dass, E. Aprà, M. Stener and A. Fortunelli, *J. Phys. Chem. C*, 2017, 121, 10832.
- 66 M. Monti, G. Brancolini, E. Coccia, D. Toffoli, A. Fortunelli, S. Corni, M. Aschi and M. Stener, *J. Phys. Chem. Lett.*, 2023, 14, 1941.
- 67 M. Monti, M. F. Matus, S. Malola, A. Fortunelli, M. Aschi, M. Stener and H. Häkkinen, *ACS Nano*, 2023, 17, 11481.
- 68 M. Zhu, C. M. Aikens, F. J. Hollander, G. Schatz and R. Jin, *J. Am. Chem. Soc.*, 2008, 130, 5883.

