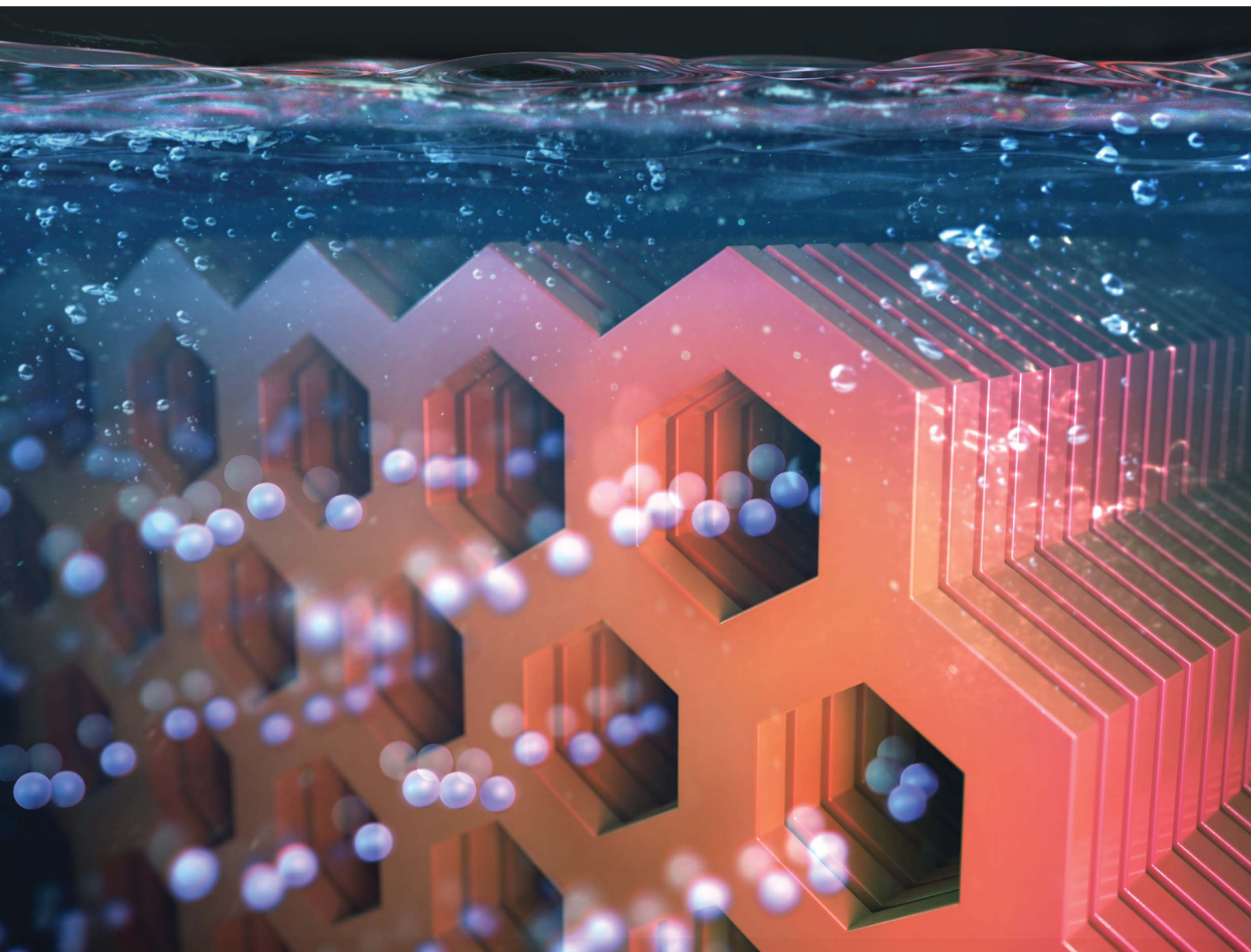


Chemical Science

Volume 11
Number 43
21 November 2020
Pages 11635–11948

rsc.li/chemical-science



ISSN 2041-6539

EDGE ARTICLE

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Cite this: *Chem. Sci.*, 2020, **11**, 11692

All publication charges for this article have been paid for by the Royal Society of Chemistry

A single-ion conducting covalent organic framework for aqueous rechargeable Zn-ion batteries†

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Despite their potential as promising alternatives to current state-of-the-art lithium-ion batteries, aqueous rechargeable Zn-ion batteries are still far away from practical applications. Here, we present a new class of single-ion conducting electrolytes based on a zinc sulfonated covalent organic framework (TpPa-SO₃Zn_{0.5}) to address this challenging issue. TpPa-SO₃Zn_{0.5} is synthesised to exhibit single Zn²⁺ conduction behaviour via its delocalised sulfonates that are covalently tethered to directional pores and achieve structural robustness by its β-ketoenamine linkages. Driven by these structural and physicochemical features, TpPa-SO₃Zn_{0.5} improves the redox reliability of the Zn metal anode and acts as an ionomeric buffer layer for stabilising the MnO₂ cathode. Such improvements in the TpPa-SO₃Zn_{0.5}-electrode interfaces, along with the ion transport phenomena, enable aqueous Zn-MnO₂ batteries to exhibit long-term cyclability, demonstrating the viability of COF-mediated electrolytes for Zn-ion batteries.

Received 15th May 2020

Accepted 23rd September 2020

DOI: 10.1039/d0sc02785e

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Introduction

The forthcoming smart and ubiquitous energy era has inspired the relentless pursuit of advanced power sources with high electrochemical performances and safety that can outperform current state-of-the-art Li-ion batteries.^{1,2} Moreover, the rapid growth of power sources in emerging application fields has brought up new issues on their environmental benignity and cost competitiveness.^{3–5}

Among various power sources explored to date, aqueous Zn-ion batteries have garnered considerable attention as a promising candidate to achieve this challenging goal owing to the use of water-based electrolytes and multielectron redox (Zn^{0/2+})-driven high energy density.^{3–8} Notably, recent advances in electrochemical rechargeability enabled by introduction of mild acidic electrolytes have encouraged potential use of Zn-ion batteries.^{6–13} However, practical application of these batteries has still been staggering mainly due to the lack of suitable electrolytes ensuring interfacial stability with electrodes. In

particular, poor redox reversibility of the Zn metal electrode poses a formidable challenge to the electrochemical performance sustainability.^{14–19}

Single-ion conductors have been investigated as an appealing electrolyte platform beyond conventional liquid electrolytes since their high cation transference number (t_+) can mitigate unwanted interfacial side reactions with electrodes.^{20–22} Despite these advantageous effects, a limited number of single Zn²⁺ conductors have been reported in comparison to monovalent cation (e.g., Li⁺ and Na⁺) conductors. An inorganic single-ion conductor based on ZnPS₃ was reported as a solid-state electrolyte.²³ However, strongly bound Zn²⁺ in the lattices resulted in poor ion transport (e.g., ionic conductivity (σ) = 10^{−8} to 10^{−6} S cm^{−1} at 60 °C). Other approaches include the use of anionic metal-organic frameworks (MOFs) and polymers.^{24–26} Unfortunately, the instability of electrolyte-electrode interfaces has not yet been resolved, making it difficult to enable practical Zn-ion batteries. Thus, an innovative concept of single Zn²⁺ conductors that can secure interfacial compatibility with electrodes as well as reliable ion transport behaviour is urgently needed.

Here, we demonstrate for the first time the use of a single-ion conducting covalent organic framework (COF) as a viable electrolyte strategy for aqueous rechargeable Zn-ion batteries. COFs have been regarded as appealing ion transport media owing to their ordered porous structure, functionalities and structural stability.^{27–33} A zinc sulfonated COF (TpPa-SO₃Zn_{0.5}; Fig. 1a) is synthesised to build well-defined directional channels in which covalently tethered and delocalised sulfonates play key roles in realising single Zn²⁺ transport. In addition, chemically stable β-

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† Electronic supplementary information (ESI) available. See DOI: 10.1039/d0sc02785e

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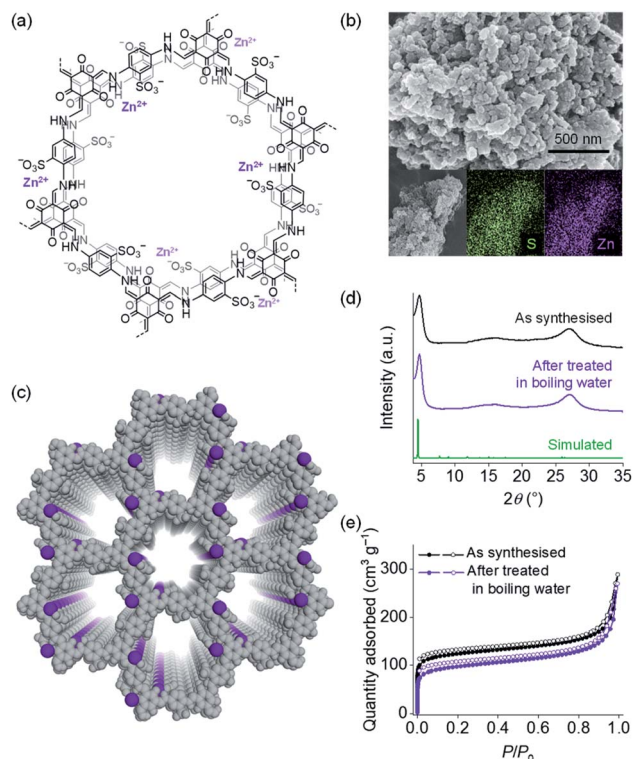


Fig. 1 (a) Chemical structure, (b) SEM and EDS mapping (for S and Zn) images, (c) structural model (Zn^{2+} : purple, TpPa-SO_3^- : grey), (d) XRD patterns and (e) N_2 gas isotherms of $\text{TpPa-SO}_3\text{Zn}_{0.5}$.

ketoenamine linkages are introduced into the framework to achieve structural robustness against water. Benefiting from these structural and physicochemical uniqueness, $\text{TpPa-SO}_3\text{Zn}_{0.5}$ successfully exhibits single Zn^{2+} conduction characteristics ($\sigma = 2.2 \times 10^{-4} \text{ S cm}^{-1}$ at room temperature and $t_{\text{Zn}^{2+}} = 0.91$), thereby allowing sustainable redox of the Zn metal anode. Moreover, $\text{TpPa-SO}_3\text{Zn}_{0.5}$ acts as an ionomeric buffer layer that can suppress structural disruption of the MnO_2 cathode during repeated redox. These advantageous effects of $\text{TpPa-SO}_3\text{Zn}_{0.5}$ on the interfacial compatibility with the electrodes, in combination with its single-ion conduction properties, enable Zn– MnO_2 cells to provide an excellent electrochemical performance with long-term cyclability.

Results and discussion

$\text{TpPa-SO}_3\text{Zn}_{0.5}$ was synthesised *via* the solvothermal reaction of 1,3,5-triformylphloroglucinol (Tp) and 1,4-phenylenediamine-2-sulfonic acid ($\text{Pa-SO}_3\text{H}$) to obtain a sulfonic acid COF ($\text{TpPa-SO}_3\text{H}$; yield: 96%),^{28,34} followed by reaction with zinc acetate to exchange H^+ with Zn^{2+} (yield: 91%; Scheme S1†). The CHN analysis and inductively coupled plasma optical emission spectrometry (ICP-OES) results show that the elemental composition of the synthesised $\text{TpPa-SO}_3\text{Zn}_{0.5}$ matches well with the theoretical composition (e.g., for the Zn content, 10.09 (calcd.) and 9.82 wt% (found); Table S1†). The formation of β -ketoenamine linkages in $\text{TpPa-SO}_3\text{Zn}_{0.5}$ was confirmed by the

characteristic cross-polarisation magic angle spinning ^{13}C nuclear magnetic resonance (CP-MAS ^{13}C NMR) spectrum (C atoms assigned to $-\text{C}=\text{O}$ (184 ppm) and $-\text{HN}-\text{C}=\text{C}-$ groups (147 and 108 ppm); Fig. S1a†). This result was further verified by Fourier transform infrared (FT-IR) spectrometry ($\text{C}=\text{C}$ (1574 cm^{-1}) and $\text{C}-\text{N}$ stretches (1251 cm^{-1}); Fig. S1b†).

Scanning electron microscopy (SEM), transmission electron microscopy (TEM) and energy dispersive spectrometry (EDS) results reveal the formation of $\text{TpPa-SO}_3\text{Zn}_{0.5}$ crystallites with a uniform distribution of S and Zn (Fig. 1b and S2†). The structural model of $\text{TpPa-SO}_3\text{Zn}_{0.5}$ (Fig. 1c) was constructed based on a triclinic crystal system (space group = $P1$; Table S2†)^{28,34} in which hexagonal pores are vertically stacked to form a stable eclipsed configuration with a slight offset (Fig. S3a†). The interplanar stacking distance along the c -axis was estimated to be *ca.* 3.4 Å (Fig. S3b†). In addition, the geometry of Zn^{2+} coordinated by three O atoms originating from two $-\text{C}=\text{O}$ and one $-\text{SO}_3^-$ was suggested to be most thermodynamically stable (Fig. S3c†). The charge neutrality of the framework was secured by the localised charge distribution, in which Zn^{2+} and non-coordinating $-\text{SO}_3^-$ produce the electron-deficient and -rich regions, respectively (Fig. S4†). The X-ray diffraction (XRD) pattern of $\text{TpPa-SO}_3\text{Zn}_{0.5}$ shows characteristic peaks at $2\theta = 4.6$ and 27.0° assigned to the (100) and (001) facets, respectively (Fig. 1d, black), which is similar to the simulated pattern obtained from the structural model (Fig. 1d, green). The difference in the peak width and intensity ratio between the experimental and simulated patterns might be due to the small crystallite size and deviation from perfect crystallinity.^{35,36} A series of $\text{TpPa-SO}_3\text{X}$ ($\text{X} = \text{H}, \text{Li}, \text{Zn}_{0.5}$) afforded similar XRD patterns,^{28,34} showing that various positive charge carriers can be paired with the isostructural framework. In addition, the N_2 gas sorption isotherms of $\text{TpPa-SO}_3\text{Zn}_{0.5}$ reveal a porous structure with a Brunauer–Emmett–Teller (BET) surface area of $472 \text{ m}^2 \text{ g}^{-1}$ and a pore size of *ca.* 13 Å (Fig. 1e, black and S5†). This well-ordered porous structure of $\text{TpPa-SO}_3\text{Zn}_{0.5}$ could contribute to facilitating uniform and directional ion transport.^{27–33}

The stability of $\text{TpPa-SO}_3\text{Zn}_{0.5}$ in water was evaluated as a requirement for use in aqueous batteries. After treatment in H_2O at 100°C for 7 days, $\text{TpPa-SO}_3\text{Zn}_{0.5}$ still showed the characteristic XRD pattern (Fig. 1d, purple) and porosity (BET surface area = $365 \text{ m}^2 \text{ g}^{-1}$; Fig. 1e, purple). This good stability in water could be ascribed to the presence of chemically robust β -ketoenamine linkages.^{36–38}

The ionic conductivity of $\text{TpPa-SO}_3\text{Zn}_{0.5}$ was examined by electrochemical impedance spectrometry (EIS) using Zn^{2+} blocking $\text{Ti}||\text{Ti}$ cells. The self-standing pellet of $\text{TpPa-SO}_3\text{Zn}_{0.5}$ used in this analysis (the inset in Fig. 2a, left) was prepared by a cold-pressing method and showed a densely packed morphology (Fig. S6a†). The obtained pellet, in which H_2O (100 wt% of the pellet weight) was incorporated to hydrate Zn^{2+} ,^{24,25} yielded $\sigma = 2.2 \times 10^{-4} \text{ S cm}^{-1}$ at room temperature (Fig. 2a, black and S7†). The temperature-dependent ionic conductivity showed Arrhenius behaviour with an activation energy (E_a) of 0.19 eV, similar to that reported for the hydrated Zn^{2+} conducting MOF.²⁴ To elucidate the role of Zn^{2+} in the COF media, a non-sulfonated COF comprising monophenyl building



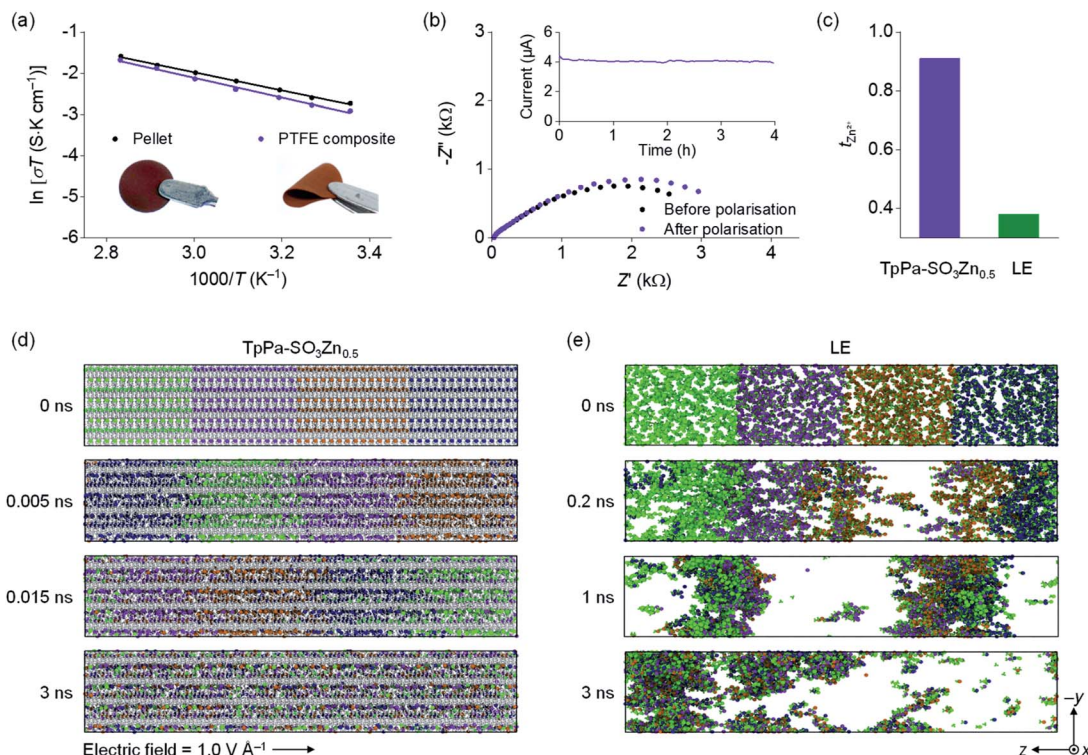


Fig. 2 (a) Arrhenius plots for the ionic conductivity of hydrated TpPa-SO₃Zn_{0.5}. The inset shows optical images of a TpPa-SO₃Zn_{0.5} pellet (left) and a TpPa-SO₃Zn_{0.5}-PTFE composite membrane (right). (b) EIS profiles and time-dependent current profile for a Zn|TpPa-SO₃Zn_{0.5}|Zn cell recorded at 20 mV polarisation. (c) $t_{Zn^{2+}}$ values obtained for TpPa-SO₃Zn_{0.5} (purple) and LE¹⁶ (2 M ZnSO₄ in H₂O; green). Representative snapshots obtained from the MD simulations showing time-dependent ion distributions in (d) TpPa-SO₃Zn_{0.5} and (e) LE. Zn²⁺: coloured diversely for a clear representation of the movement, TpPa-SO₃⁻: grey, SO₄²⁻: green, H₂O: omitted for clarity.

blocks (TpPa)³⁸ was synthesised as a control sample without Zn²⁺ (Fig. S8†). TpPa showed a considerably low conductivity after hydration ($\sigma = 4.3 \times 10^{-6}$ S cm⁻¹; Fig. S9†), verifying that the charge transport in the COF media is predominantly enabled by Zn²⁺.

The mechanically flexible membrane of TpPa-SO₃Zn_{0.5} was prepared by mixing a small amount of polytetrafluoroethylene (PTFE; 5 wt% of the TpPa-SO₃Zn_{0.5} weight). The SEM image of the resultant composite membrane shows that the PTFE webs as a binder tightly entangle the TpPa-SO₃Zn_{0.5} powders (Fig. S6b†), thereby resulting in good flexibility (the inset in Fig. 2a, right). Notably, the ion conduction characteristics of this membrane ($\sigma = 1.8 \times 10^{-4}$ S cm⁻¹ at room temperature, $E_a = 0.20$ eV after hydration; Fig. 2a, purple) are comparable to those of the pellet, indicating that the PTFE binder minimally impedes the ion transport of TpPa-SO₃Zn_{0.5}. It should be noted that the electrochemical tests subsequently shown were performed using this practical composite membrane (identically denoted as TpPa-SO₃Zn_{0.5}).

To demonstrate the single-ion conduction behaviour of TpPa-SO₃Zn_{0.5}, its $t_{Zn^{2+}}$ was examined using a potentiostatic polarisation method according to the following equation:^{39,40}

$$t_{Zn^{2+}} = \frac{I_{ss}(\Delta V - I_0 R_0)}{I_0(\Delta V - I_{ss} R_{ss})}$$

where I_0 and I_{ss} are the currents at the initial and steady states, respectively, and R_0 and R_{ss} are the interfacial resistances before

and after polarisation, respectively. The time-dependent current profile and impedance spectra of a Zn|TpPa-SO₃Zn_{0.5}|Zn cell were recorded before and after polarisation ($\Delta V = 20$ mV) at room temperature. The interfacial resistances were measured as $R_0 = 3360$ and $R_{ss} = 3810$ Ω, during which the current response slightly decreased from $I_0 = 4.40$ to $I_{ss} = 3.91$ μA (Fig. 2b). Accordingly, $t_{Zn^{2+}} = 0.91$ was obtained for TpPa-SO₃Zn_{0.5} (Fig. 2c, purple), which is one of the highest values among those reported for the Zn²⁺ conducting polyanions (Table S3†). It should be noted that a typical liquid electrolyte (LE; 2 M ZnSO₄ in H₂O) gave a low $t_{Zn^{2+}}$ value of 0.38 (Fig. 2c, green).¹⁶

The ion conduction behaviour of TpPa-SO₃Zn_{0.5} was further investigated *via* molecular dynamics (MD) simulations. For this analysis, a TpPa-SO₃Zn_{0.5} model saturated with H₂O and a control model of the LE (2 M ZnSO₄ in H₂O) were constructed (Table S4†). The application of an external electric field (1.0 V Å⁻¹ in the $-z$ -axis direction) to both model systems induced the ion movements (Fig. 2d and e). A notable finding is that a significantly uniform ion distribution was observed in TpPa-SO₃Zn_{0.5} compared with the LE, which could result from the difference in the dynamic behaviour of the anion groups in both systems. In TpPa-SO₃Zn_{0.5}, the -SO₃⁻ group is covalently tethered along the directional pores, forming anionic channels that allowed the uniform Zn²⁺ flux (Fig. 2d). However, freely mobile SO₄²⁻ in the LE formed randomly spread ion clusters with Zn²⁺ due to their strong electrostatic interactions (Fig. 2e). These random ion movements in the LE resulted in the velocity



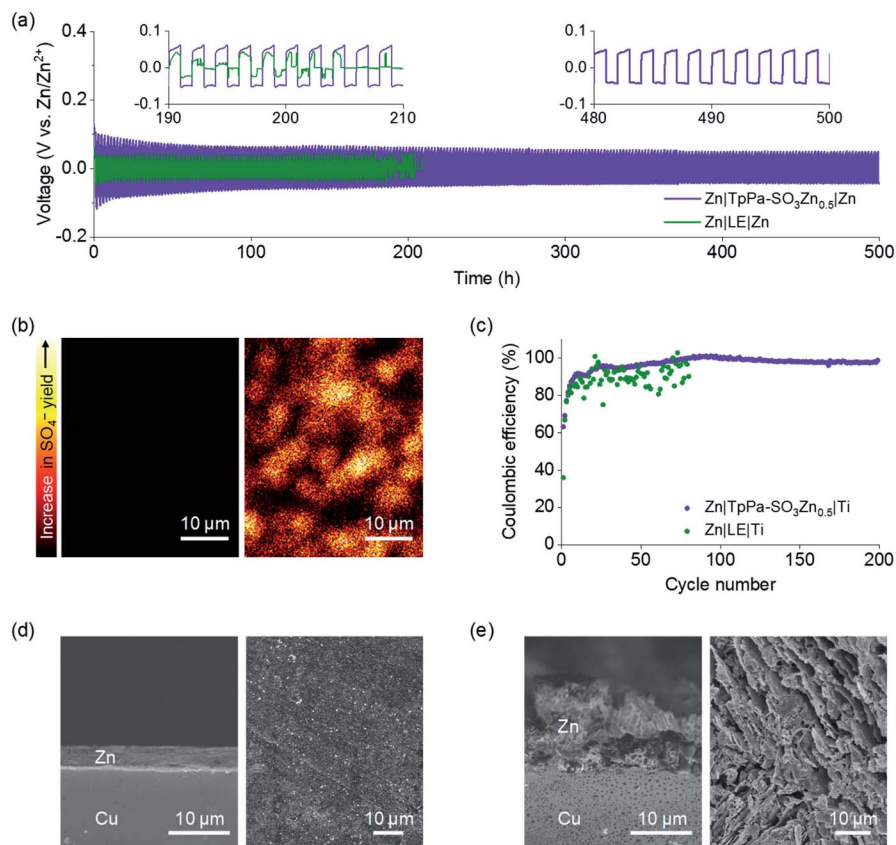


Fig. 3 (a) Galvanostatic voltage profiles of Zn|TpPa-SO₃Zn_{0.5}|Zn (purple) and Zn|LE|Zn cells (LE = 2 M ZnSO₄ in H₂O; green). (b) TOF-SIMS mapping images for SO₄⁻ yielded on Zn metal electrodes in contact with TpPa-SO₃Zn_{0.5} (left) or LE (right) after cycling tests. (c) Coulombic efficiency plots for Zn plating/stripping processes of Zn|TpPa-SO₃Zn_{0.5}|Ti (purple) and Zn|LE|Ti cells (green). Cross-sectional and surface SEM images for electrochemically plated Zn metals in contact with (d) TpPa-SO₃Zn_{0.5} or (e) LE.

distribution to be nearly neutral against the direction of the applied electric field (Fig. S10a†, green). Meanwhile, a shift of the velocity distribution following the direction of the electric field was observed in TpPa-SO₃Zn_{0.5} (Fig. S10a†, purple), correspondingly verifying the characteristic movement of Zn²⁺ in the anionic channels. Moreover, Zn²⁺ in TpPa-SO₃Zn_{0.5} was found to be readily hydrated (*i.e.*, a large fraction of O atoms coordinated to Zn²⁺ was derived from H₂O; Fig. S10b†), indicating the promoted ionic dissociation between the anionic framework and Zn²⁺ that is highly advantageous for facile ion transport. These results verify that the immobilised and delocalised -SO₃⁻ group in TpPa-SO₃Zn_{0.5} plays key roles in enabling fast single-ion conduction.

In addition to the ion conduction characteristics of TpPa-SO₃Zn_{0.5}, its thermal and electrochemical stabilities were investigated. The thermogravimetric analysis (TGA) curve shows no obvious weight loss up to 200 °C under a N₂ atmosphere (Fig. S11†). The linear sweep voltammogram (LSV) shows that TpPa-SO₃Zn_{0.5} is electrochemically stable up to *ca.* 2 V vs. Zn/Zn²⁺ (Fig. S12†), thus fulfilling the requirement for building practical aqueous batteries.

To examine the applicability of TpPa-SO₃Zn_{0.5} to the Zn metal electrode, the galvanostatic cyclability of a Zn|TpPa-SO₃Zn_{0.5}|Zn cell was examined and compared with that of

a Zn|LE|Zn cell (LE = 2 M ZnSO₄ in H₂O) at room temperature (current density = 0.1 mA cm⁻², capacity = 0.1 mA h cm⁻²). The cell containing the LE, although it initially operated with a low overpotential (*ca.* 40 mV), showed fluctuating behaviour after 180 h and eventually failed (Fig. 3a, green), which might result from an electrical short-circuit. In sharp contrast, TpPa-SO₃Zn_{0.5} allowed stable operation of the cell for 500 h, during which the overpotential gradually decreased from *ca.* 100 to 60 mV (Fig. 3a, purple). In addition, the structural ordering of TpPa-SO₃Zn_{0.5} was not disrupted after the cycling test (Fig. S13†).

Such an obvious difference in the stability of the cells was investigated by analysing the Zn metal electrodes after the cycling tests. Fig. 3b shows the mapping images of SO₄⁻ as a representative ion species yielded from time-of-flight secondary ion mass spectrometry (TOF-SIMS) conducted on the Zn metal electrodes of both cells. SO₄⁻ was minimally observed on the Zn metal electrode in contact with TpPa-SO₃Zn_{0.5} (left), exhibiting good interfacial stability. In contrast, randomly proliferated SO₄⁻ was observed on the Zn metal surface in contact with the LE (right), indicating the formation of a deposition phase that could result from the nonuniform ion transport and adverse reactions between mobile SO₄²⁻ and Zn metal. The X-ray photoelectron spectrometry (XPS) and XRD



results correspondingly reveal that the unwanted corrosion products (e.g., $\text{Zn}_4\text{SO}_4(\text{OH})_6 \cdot 5\text{H}_2\text{O}$)⁴¹ were exclusively formed on the Zn metal electrode cycled in the LE (Fig. S14†).

The reversibility of Zn plating/stripping in the two electrolyte systems was compared using the Zn||Ti configurations. During a cycle, Zn metal was electrochemically plated on the Ti working electrode and subsequently stripped out (current density = 0.1 mA cm^{-2} , capacity = 0.1 mA h cm^{-2}). TpPa-SO₃Zn_{0.5} allowed this process to persist over 200 cycles with high reversibility (Fig. 3c, purple), in contrast to the LE that resulted in cycling with severe fluctuation (Fig. 3c, green). The initially observed increase in the Coulombic efficiency of the Zn|TpPa-SO₃Zn_{0.5}|Ti cell might include gradual stabilisation of the TpPa-SO₃Zn_{0.5}-Ti metal interface.¹⁶

The advantages of TpPa-SO₃Zn_{0.5} were further highlighted by monitoring the Zn electroplating behaviour. As shown in the SEM images, densely plated Zn metal with a smooth surface was observed on the Cu substrate in contact with TpPa-SO₃Zn_{0.5} (Fig. 3d), whereas porous and rough Zn metal plating was formed in the LE (Fig. 3e). These results demonstrate the beneficial effects of single-ion transport in TpPa-SO₃Zn_{0.5} on the stable and reversible Zn plating/stripping behaviour.

The applicability of TpPa-SO₃Zn_{0.5} as a new electrolyte in aqueous rechargeable Zn-ion batteries was investigated using the Zn||MnO₂ configurations. α -MnO₂ used as a cathode material was hydrothermally synthesised according to a previous report,¹¹ yielding brown powders that revealed a regular nanorod shape and a characteristic XRD pattern (Fig. S15a and b†). In addition, the normal redox behaviour of α -MnO₂ (cathodic peaks at 1.2–1.4 and anodic peaks at ca. 1.6 V in the cyclic voltammograms (CVs); Fig. S15c†) consistent with the previous reports^{11,13} was observed.

The electrochemical performance of a Zn|TpPa-SO₃Zn_{0.5}-|MnO₂ cell was tested at a current density of 0.6 A g^{-1} at room temperature. Notably, cycling stability over 800 cycles was observed for this cell (Fig. 4a). The cell initially showed a specific capacity of $196.0 \text{ mA h g}^{-1}$ and still delivered $144.0 \text{ mA h g}^{-1}$ at the 800th cycle, representing 73% capacity retention. The discharge curves include two sloping plateaus at 1.2–1.4 V (Fig. 4b), similar to those observed in the CVs. In contrast, the cell containing a typical LE (2 M ZnSO₄ + 0.2 M MnSO₄ in H₂O) showed rapid capacity decay during the cycling (40% capacity retention after the 400th cycle; Fig. S16a†).

Postmortem analyses were conducted on the decent cyclability of the Zn|TpPa-SO₃Zn_{0.5}-|MnO₂ cell. As shown in the SEM image (Fig. 4c), TpPa-SO₃Zn_{0.5} allowed redox stability of the Zn metal anode in the cell, similar to that observed in the Zn plating/stripping tests (*vide supra*; Fig. 3). Moreover, the characteristic nanorod shape of MnO₂ with structural ordering was well-maintained after the cycling (Fig. 4d), indicating that TpPa-SO₃Zn_{0.5} might function as an ionomeric buffer layer to suppress the structural disruption of MnO₂.⁴² The XPS results further show that the formation of decomposition products⁴³ was retarded in the MnO₂ cathode in contact with TpPa-SO₃Zn_{0.5} (Fig. S17†). In contrast, the LE in the cell generated severely pulverised Zn metal and MnO₂ particles (Fig. S16b and c†). These results verify that the good compatibility of TpPa-

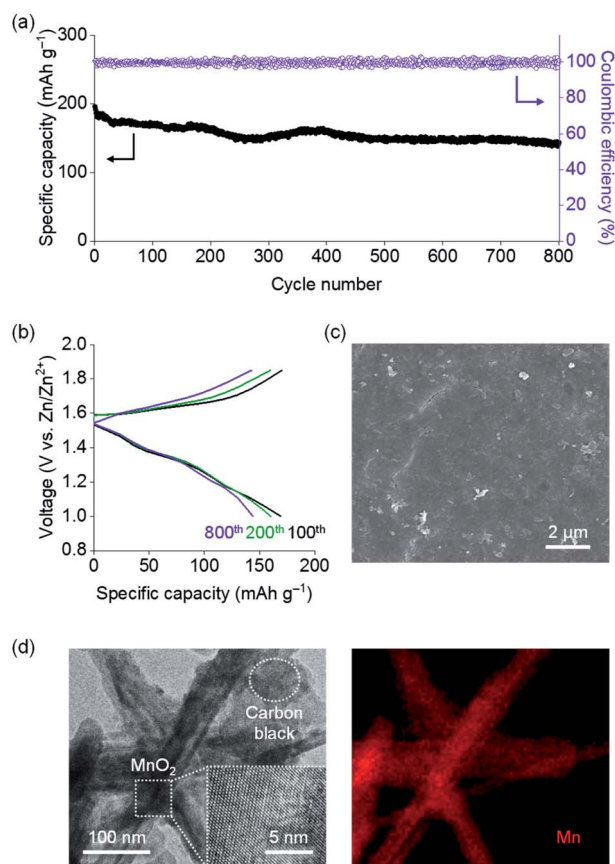


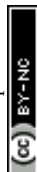
Fig. 4 (a) Cycling performance of a Zn|TpPa-SO₃Zn_{0.5}|MnO₂ cell. (b) Charge/discharge profiles at the 100th, 200th and 800th cycles. (c) SEM image of the Zn metal anode after the 800th cycle. (d) TEM (left) and EDS mapping (for Mn; right) images of the MnO₂ cathode after the 800th cycle.

SO₃Zn_{0.5} with the electrodes, which is ascribed to the unique structural features and single-ion conduction behaviour, significantly contributed to the cycling sustainability of the Zn|TpPa-SO₃Zn_{0.5}-|MnO₂ cell.

The performance of electrochemical cells containing TpPa-SO₃Zn_{0.5} was compared with that of the cells containing other advanced electrolytes (Table S5†). Among the numerous previous approaches, the systems using single Zn²⁺ conductors, including an anionic MOF²⁴ and sulfonated polymers,^{25,26} were exclusively chosen for fair comparison. TpPa-SO₃Zn_{0.5} allowed superior cyclability of the electrochemical cells with a reliable operating voltage compared with the previously reported materials, underscoring that the single-ion conducting COF offers great promise as an attractive electrolyte for aqueous Zn-ion batteries.

Conclusions

We have presented the single-ion conducting COF, TpPa-SO₃Zn_{0.5}, as a viable electrolyte for aqueous rechargeable Zn-ion batteries. TpPa-SO₃Zn_{0.5} featured immobilised and delocalised sulfonates in the directional pores (enabling single Zn²⁺ transport) and the stable β -ketoenamine linkages (allowing



structural robustness against water). The unusual single-ion conduction behaviour of $\text{TpPa-SO}_3\text{Zn}_{0.5}$, which was elucidated by experimental and theoretical studies, allowed Zn plating/stripping reliability. In addition, $\text{TpPa-SO}_3\text{Zn}_{0.5}$ as an ionomeric buffer layer effectively prevented the structural disruption of the MnO_2 cathode during repeated redox. Driven by these improvements in the $\text{TpPa-SO}_3\text{Zn}_{0.5}$ -electrode interfaces and the ion transport phenomena, $\text{TpPa-SO}_3\text{Zn}_{0.5}$ enabled the Zn- MnO_2 cells to provide long-term cycling performance (over 800 cycles at a current density of 0.6 A g^{-1} at an average operating voltage of 1.4 V). Future development of single Zn^{2+} conducting COFs will be directed to material optimisation and engineering with Zn-ion batteries. The COF-mediated electrolytes described herein move us a step closer toward practical use of rechargeable Zn-ion batteries and open a new route for the design of advanced aqueous electrolytes.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by the Basic Science Research Program (2018R1A2A1A05019733 and 2018M3D1A1058624) through the National Research Foundation of Korea (NRF) funded by the Ministry of Science, ICT and Future Planning. This work was also supported by the Samsung Research Funding Centre of Samsung Electronics under project no. SRFC-MA1702-04. KJ acknowledges the financial support from the Basic Science Research Program (2017R1D1A1B03033699 and 2020R1I1A1A01064798) through the NRF funded by the Ministry of Education. SKK acknowledges the financial support from ERC (NRF-2014R1A5A1009799) and computational support from UNIST-HPC.

Notes and references

- 1 M. Armand and J. M. Tarascon, *Nature*, 2008, **451**, 652–657.
- 2 M. Li, J. Lu, Z. Chen and K. Amine, *Adv. Mater.*, 2018, **30**, 1800561.
- 3 Z. Liu, Y. Huang, Y. Huang, Q. Yang, X. Li, Z. Huang and C. Zhi, *Chem. Soc. Rev.*, 2020, **49**, 180–232.
- 4 J. Huang, Z. Guo, Y. Ma, D. Bin, Y. Wang and Y. Xia, *Small Methods*, 2019, **3**, 1800272.
- 5 Y. Liang, Y. Jing, S. Gheyhani, K.-Y. Lee, P. Liu, A. Facchetti and Y. Yao, *Nat. Mater.*, 2017, **16**, 841–848.
- 6 B. Tang, L. Shan, S. Liang and J. Zhou, *Energy Environ. Sci.*, 2019, **12**, 3288–3304.
- 7 M. Song, H. Tan, D. Chao and H. J. Fan, *Adv. Funct. Mater.*, 2018, **28**, 1802564.
- 8 G. Fang, J. Zhou, A. Pan and S. Liang, *ACS Energy Lett.*, 2018, **3**, 2480–2501.
- 9 A. Konarov, N. Voronina, J. H. Jo, Z. Bakenov, Y.-K. Sun and S.-T. Myung, *ACS Energy Lett.*, 2018, **3**, 2620–2640.
- 10 D. Kundu, B. D. Adams, V. Duffort, S. H. Vajargah and L. F. Nazar, *Nat. Energy*, 2016, **1**, 16119.
- 11 H. Pan, Y. Shao, P. Yan, Y. Cheng, K. S. Han, Z. Nie, C. Wang, J. Yang, X. Li, P. Bhattacharya, K. T. Mueller and J. Liu, *Nat. Energy*, 2016, **1**, 16039.
- 12 N. Zhang, F. Cheng, J. Liu, L. Wang, X. Long, X. Liu, F. Li and J. Chen, *Nat. Commun.*, 2017, **8**, 405.
- 13 C. Xu, B. Li, H. Du and F. Kang, *Angew. Chem., Int. Ed.*, 2012, **51**, 933–935.
- 14 H. Jia, Z. Wang, B. Tawiah, Y. Wang, C.-Y. Chan, B. Fei and F. Pan, *Nano Energy*, 2020, **70**, 104523.
- 15 J. Shin, J. Lee, Y. Park and J. W. Choi, *Chem. Sci.*, 2020, **11**, 2028–2044.
- 16 Z. Zhao, J. Zhao, Z. Hu, J. Li, J. Li, Y. Zhang, C. Wang and G. Cui, *Energy Environ. Sci.*, 2019, **12**, 1938–1949.
- 17 F. Wang, O. Borodin, T. Gao, X. Fan, W. Sun, F. Han, A. Faraone, J. A. Dura, K. Xu and C. Wang, *Nat. Mater.*, 2018, **17**, 543–549.
- 18 S. Higashi, S. W. Lee, J. S. Lee, K. Takechi and Y. Cui, *Nat. Commun.*, 2016, **7**, 11801.
- 19 H. Qiu, X. Du, J. Zhao, Y. Wang, J. Ju, Z. Chen, Z. Hu, D. Yan, X. Zhou and G. Cui, *Nat. Commun.*, 2019, **10**, 5374.
- 20 K. Jeong, S. Park and S.-Y. Lee, *J. Mater. Chem. A*, 2019, **7**, 1917–1935.
- 21 Z. Zhang, Y. Shao, B. Lotsch, Y.-S. Hu, H. Li, J. Janek, L. F. Nazar, C.-W. Nan, J. Maier, M. Armand and L. Chen, *Energy Environ. Sci.*, 2018, **11**, 1945–1976.
- 22 H. Zhang, C. Li, M. Piszcz, E. Coya, T. Rojo, L. M. Rodriguez-Martinez, M. Armand and Z. Zhou, *Chem. Soc. Rev.*, 2017, **46**, 797–815.
- 23 A. J. Martinolich, C.-W. Lee, I.-T. Lu, S. C. Bevilacqua, M. B. Preefer, M. Bernardi, A. Schleife and K. A. See, *Chem. Mater.*, 2019, **31**, 3652–3661.
- 24 Z. Wang, J. Hu, L. Han, Z. Wang, H. Wang, Q. Zhao, J. Liu and F. Pan, *Nano Energy*, 2019, **56**, 92–99.
- 25 C. Hänsel and D. Kundu, *ACS Omega*, 2019, **4**, 2684–2692.
- 26 B.-S. Lee, S. Cui, X. Xing, H. Liu, X. Yue, V. Petrova, H.-D. Lim, R. Chen and P. Liu, *ACS Appl. Mater. Interfaces*, 2018, **10**, 38928–38935.
- 27 X. Li and K. P. Loh, *ACS Materials Lett.*, 2019, **1**, 327–335.
- 28 K. Jeong, S. Park, G. Y. Jung, S.-H. Kim, Y.-H. Lee, S. K. Kwak and S.-Y. Lee, *J. Am. Chem. Soc.*, 2019, **141**, 5880–5885.
- 29 Z. Li, Z.-W. Liu, Z. Li, T.-X. Wang, F. Zhao, X. Ding, W. Feng and B.-H. Han, *Adv. Funct. Mater.*, 2020, **30**, 1909267.
- 30 Z. Guo, Y. Zhang, Y. Dong, J. Li, S. Li, P. Shao, X. Feng and B. Wang, *J. Am. Chem. Soc.*, 2019, **141**, 1923–1927.
- 31 S. Ashraf, Y. Zuo, S. Li, C. Liu, H. Wang, X. Feng, P. Li and B. Wang, *Chem.-Eur. J.*, 2019, **25**, 13479–13483.
- 32 Q. Xu, S. Tao, Q. Jiang and D. Jiang, *J. Am. Chem. Soc.*, 2018, **140**, 7429–7432.
- 33 Y. Du, H. Yang, J. M. Whiteley, S. Wan, Y. Jin, S.-H. Lee and W. Zhang, *Angew. Chem., Int. Ed.*, 2016, **55**, 1737–1741.
- 34 S. Chandra, T. Kundu, K. Dey, M. Addicoat, T. Heine and R. Banerjee, *Chem. Mater.*, 2016, **28**, 1489–1494.
- 35 E. Vitaku and W. R. Dichtel, *J. Am. Chem. Soc.*, 2017, **139**, 12911–12914.
- 36 S. Mitra, S. Kandambeth, B. P. Biswal, M. A. Khayum, C. K. Choudhury, M. Mehta, G. Kaur, S. Banerjee,



- A. Prabhune, S. Verma, S. Roy, U. K. Kharul and R. Banerjee, *J. Am. Chem. Soc.*, 2016, **138**, 2823–2828.
- 37 S. Kandambeth, K. Dey and R. Banerjee, *J. Am. Chem. Soc.*, 2019, **141**, 1807–1822.
- 38 S. Kandambeth, A. Mallick, B. Lukose, M. V. Mane, T. Heine and R. Banerjee, *J. Am. Chem. Soc.*, 2012, **134**, 19524–19527.
- 39 J. Evans, C. A. Vincent and P. G. Bruce, *Polymer*, 1987, **28**, 2324–2328.
- 40 N. S. Schauser, R. Seshadri and R. A. Segalman, *Mol. Syst. Des. Eng.*, 2019, **4**, 263–279.
- 41 H. Glatz, E. Tervoort and D. Kundu, *ACS Appl. Mater. Interfaces*, 2020, **12**, 3522–3530.
- 42 Y. Zeng, X. Zhang, Y. Meng, M. Yu, J. Yi, Y. Wu, X. Lu and Y. Tong, *Adv. Mater.*, 2017, **29**, 1700274.
- 43 Y. Li, S. Wang, J. R. Salvador, J. Wu, B. Liu, W. Yang, J. Yang, W. Zhang, J. Liu and J. Yang, *Chem. Mater.*, 2019, **31**, 2036–2047.

